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# Polydentate chalcogen bonding: anion trapping with a water-stable host compound carrying Se–CF<sub>3</sub> functions†

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Bidentate and tetradentate chalcogen bonding host systems with SeCF<sub>3</sub> functions as  $\sigma$ -hole donors in close proximity at the alkyne functions of 1,8-diethynylanthracene and its *syn*-dimer were prepared in quantitative yield by tin-selenium exchange reactions of the corresponding trimethylstannyl precursors with ClSeCF<sub>3</sub>. The bidentate system shows chalcogen bonding interactions with THF, but does not bind halide ions. The tetradentate system cooperatively chelates chloride, bromide and iodide ions with its four C $\equiv$ C–SeCF<sub>3</sub> units by rotating the four  $\sigma$ -holes towards the halide ion. The structures of these halide ion adducts were determined by X-ray diffraction. The hydrobromide and -iodide salts of the ethyl derivative of Schwesinger's phosphazene superbases served as halide salts with very weakly coordinating cations.

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## Introduction

Poly  $\sigma$ -hole host compounds with the ability to cooperatively bind nucleophiles can be more active than their monodentate derivatives because they benefit from a comparable effect to the chelate effect which is well known for polydentate donor ligands.<sup>1–6</sup> Preorganization of the host compound can further enhance the binding abilities.<sup>7,8</sup> Consequently, such compounds are of great interest for processes such as catalysis,<sup>8–10</sup> anion recognition in host–guest chemistry<sup>11–13</sup> or anion transport.<sup>14–16</sup> Halogen bonding<sup>17,18</sup> is the best known member of the family of  $\sigma$ -hole interactions, but in recent years compounds utilizing pnictogen or chalcogen bonding have come to the fore.<sup>15,16,19–24</sup> Since heavier and thus more easily polarizable atoms are typically better  $\sigma$ -hole donors, most chalcogen bonding systems contain selenium or tellurium.<sup>6,22</sup> However, there are as well some efficient host systems featuring sulfur.<sup>25–27</sup> The  $\sigma$ -holes can become stronger lone-pair acceptors by using cationic  $\sigma$ -hole functions<sup>11,13,24</sup> or by attaching perfluorinated substituents (typically polyfluorinated arenes)<sup>2,28</sup> or cationic heterocycles.<sup>4,5,27</sup> For example,

Te(C<sub>6</sub>F<sub>5</sub>)<sub>2</sub> has been shown to be a more potent catalyst compared to Se(C<sub>6</sub>F<sub>5</sub>)<sub>2</sub> for Reissert-type reactions with an halide-abstraction as a rate-limiting step, and Sb(C<sub>6</sub>F<sub>5</sub>)<sub>2</sub>Ph outperforms Sb(C<sub>6</sub>F<sub>5</sub>)Ph<sub>2</sub>.<sup>28</sup> By choosing a suitable organic backbone, the directionality of multiple  $\sigma$ -hole donor sites can be predetermined to achieve both cooperativity and selectivity. Although the application of cooperative chalcogen bonding is only partially explored, there are some impressive bidentate examples that demonstrate intriguing new possibilities for anion binding and catalysis.

For example, Taylor *et al.* were able to demonstrate the advantages of the bidentate system **A** over a similar monodentate host.<sup>2</sup> Matile's benzodiselenazoles **B** have conformationally immobilized  $\sigma$ -holes on the selenium atoms, making systems like **B** potent catalysts.<sup>26</sup> Hosts such as **C** of Huber and co-workers use sulfur or selenium and were used as cationic chalcogen bonding systems in the activation of neutral compounds.<sup>27</sup> Gabbaï introduced the also cationic host **D**, in which the catalytic activity is enhanced by the presence of a telluronium function (Scheme 1).<sup>24</sup>

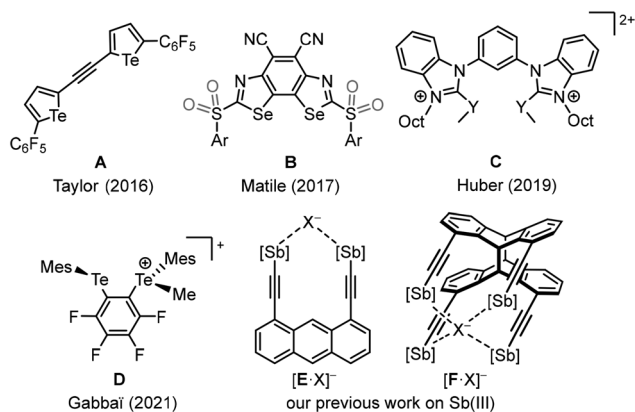
It is evident, that the design of the host system is of major importance for the activity of the respective compound. Recently, we have established diethynylanthracene<sup>29</sup> and its *syn*-photodimer<sup>1</sup> as potent organic backbones for bi- and tetradentate pnictogen bonding by attaching C $\equiv$ C–Sb(C<sub>2</sub>F<sub>5</sub>)<sub>2</sub> functions. They can cooperatively bind various nucleophiles such as chalcogenides or halide ions, and the tetradentate host **F** has been shown to clearly outperform the bidentate system **E**.<sup>1,29</sup> Two major drawbacks of these host systems are their high sensitivity towards moisture and the steric overcrowding in the tetradentate host **F**. Therefore, we have here applied the

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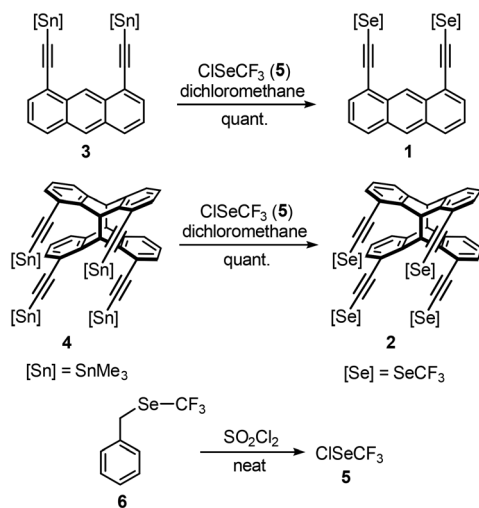
**Scheme 1** Mono- and bidentate chalcogen bonding host systems and complexes of our pnictogen bonding host systems investigated previously (Ar = C<sub>6</sub>H<sub>4</sub>tBu; Y = S, Se; X<sup>-</sup> = Cl<sup>-</sup>, Br<sup>-</sup>, I<sup>-</sup>; [Sb] = Sb(C<sub>2</sub>F<sub>5</sub>)<sub>2</sub>).

concept of chalcogen bonding to these backbones by attaching C≡C–SeCF<sub>3</sub> functions, thus avoiding steric repulsion while precluding decomposition by hydrolysis.

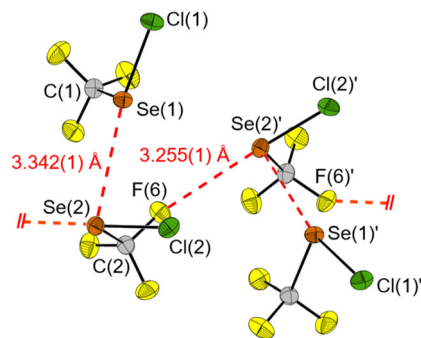
## Results and discussion

The selanyl host compounds **1** and **2** were synthesized by reacting the stannylated precursors **3**<sup>30</sup> and **4**<sup>31</sup> with ClSeCF<sub>3</sub> (**5**, Scheme 2). This was synthesized according to literature<sup>32</sup> by chlorinating benzyltrifluoromethyl selenide (**6**)<sup>33</sup> with SO<sub>2</sub>Cl<sub>2</sub>. The removal of SO<sub>2</sub> was achieved by isothermal distillation, but this purification is not necessary since the reaction of **5** with the stannylated precursors was found to be unperturbed by SO<sub>2</sub> and an excess of **5**.

To get a first indication of the σ-hole donor activity, we purified **5** (m.p. ≈ –100 °C) in order to determine its solid state structure by *in situ* crystallization and X-ray diffraction (Fig. 1).



**Scheme 2** Synthesis of selanyl chloride **5** and exchange reaction to afford the monomeric host system **1** and dimeric host system **2**.



**Fig. 1** Molecular structure of **5** in the solid state. To emphasize the infinite chain structure, additional molecules marked with ' are depicted, which were generated by symmetry (+x, 3/2 – y, 1/2 + z). The red dotted lines mark Se...F distances below the sum of van der Waals radii.<sup>35</sup> Ellipsoids are set at 50% probability. Selected distances [Å] and angles [°]: Se(1)–C(1) 1.966(2), Se(2)–C(2) 1.967(2), Se(1)–Cl(1) 2.167(1), Se(2)–Cl(2) 2.151(1); C(1)–Se(1)–Cl(1) 94.1(1), C(2)–Se(2)–Cl(2) 94.8(1), Cl(1)–Se(1)–Se(2) 167.8(1), Cl(2)–Se(2)–F(6)' 177.5(1).

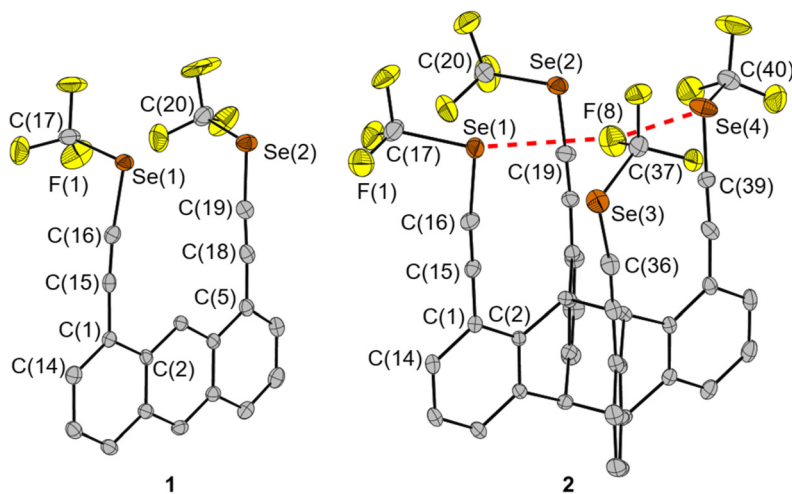
There are two independent molecules in the asymmetric unit. One of them forms infinite chains of intermolecular σ-hole interactions (Se(2)⋯F(6)). The other molecules are laterally connected to this chain by Se...Se-type chalcogen bonding interactions, where Se(1) acts as a σ-hole donor and Se(2) as a lone pair donor.

The tin compounds **3** and **4** were reacted with **5** analogously to a protocol of a tin-selenium exchange reaction,<sup>34</sup> affording the bi- and tetradentate chalcogen bonding host systems **1** and **2** in quantitative yield. Since **5** was applied neat, the workup to obtain a product of reasonable purity consisted only of removing the excess of **5** as well as the generated chlorotrimethylstannane under reduced pressure. This synthetic protocol allows for a simple and straightforward functionalization with C≡C–SeCF<sub>3</sub> units, which should be readily transferable to other polyalkynes.

The solid state structures of **1** and **2** were determined by single crystal X-ray diffraction (Fig. 2). In both structures, there is approximately a right angle between the alkyne and the trifluoromethyl group (C–Se–C<sup>F</sup> angles between 94.4(1)° and 96.8(1)°), which also implies a right angle between the alkyne and the direction of the σ-hole on each selenium atom. Since the direction of the σ-hole can change by rotating along the alkyne axis, all four σ-holes in **2** can point right at the central cavity, making **2** a tetradentate host system. Also of interest are the Se...F distances, which are shorter than the sum of their van der Waals radii ( $\sum d_{vdw}(\text{Se}, \text{F}) = 3.37 \text{ \AA}$ ),<sup>35</sup> as they indicate attractive interactions.

There are two remarkably short intramolecular Se...F contacts that indicate chalcogen bonding interactions. These two are contacts between the fluorine atoms of a CF<sub>3</sub> group and the selenium atoms of neighboring SeCF<sub>3</sub> functions. They show some directionality by adopting F...Se–C<sup>F</sup> angles of 164.8(2)° and 158.3(1)°. The relatively large deviations from linearity allow these to be classified as weaker chalcogen





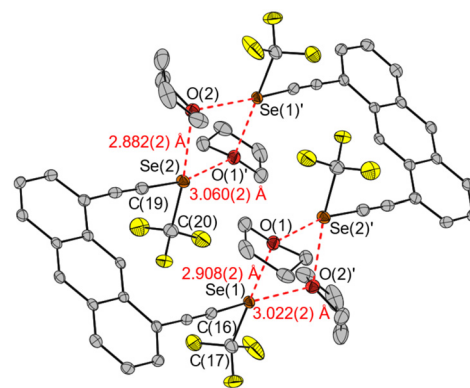
**Fig. 2** Molecular structures of **1** and **2** in the solid state. Hydrogen atoms and, in the case of **2**, solvent molecules (benzene) and minor occupied disordered parts are omitted for clarity. The red dotted lines mark Se...F distances below the sum of van der Waals radii.<sup>35</sup> Ellipsoids are set at 50% probability. Selected distances [Å] and angles [°] of **1**: Se(1)...Se(2) 4.396(1), Se(1)–C(17) 1.960(4), Se(1)–C(16) 1.824(3), Se(2)–C(20) 1.954(4), Se(2)–C(19) 1.829(3); C(14)–C(1)–C(15) 121.2(3), C(2)–C(1)–C(15) 118.4(2), C(1)–C(15)–C(16) 174.2(3), C(15)–C(16)–Se(1) 173.4(3), C(16)–Se(1)–C(17) 96.3(1); of **2**: Se(1)...F(8) 3.196(2), Se(4)...F(8) 3.296(2), Se(1)...Se(2) 4.217(2), Se(2)...Se(4) 3.680(2), Se(3)...Se(4) 5.769(1), Se(1)...Se(3) 3.892(2); C(16)–Se(1)–C(17) 96.8(1), C(19)–Se(2)–C(20) 94.4(1), C(36)–Se(3)–C(37) 95.6(6), C(39)–Se(4)–C(40) 95.2(3), F(8)–Se(1)–C(17) 164.8(2), F(8)–Se(4)–C(40) 158.3(1).

bonding interactions.<sup>36</sup> On a larger scale, **1** forms staggered layers caused by aryl stacking interactions, while molecules of **2** arrange back-to-back to form alternating layers of perfluorinated SeCF<sub>3</sub> groups and organic backbones.

One of our goals was to improve the stability of  $\sigma$ -hole host systems compared to the previously reported bis- and tetrastibanyl host systems, which are highly sensitive to air and moisture. In fact, the selenyl compounds **1** and **2** proved to be stable when stored for at least six months exposed to air and moisture. Solutions of the compounds can even be heated in the presence of concentrated aqueous hydrochloric acid or sodium hydroxide solution without showing decomposition or hydrolysis (see ESI† more information).

By crystallizing compound **1** from THF, we obtained structural insight into the dimer [1·2 THF]<sub>2</sub> by X-ray diffraction (Fig. 3). Here, all four Se...O distances are below the sum of the van der Waals radii ( $\sum d_{vdw}(Se,O) = 3.42$  Å),<sup>35</sup> confirming the donation of both lone pairs of each oxygen atom into one of the two  $\sigma$ -holes of each selenium atom. The Se...O distances associated with the  $\sigma$ -holes opposite to the CF<sub>3</sub> groups are shorter than those opposite to the alkyne groups, demonstrating that the extremely electron-withdrawing CF<sub>3</sub> groups to induce a “deeper”  $\sigma$ -hole and thus a stronger chalcogen bonding interaction. In addition, the O...Se–C<sup>F</sup> and O...Se–C<sup>alkyne</sup> angles are approximately linear (ranging from 172.3(1)° to 177.0(1)°), further indicating typical features of chalcogen bonding.<sup>6,20</sup>

The next step was to demonstrate the ability of the compounds to bind equimolar amounts of halide ions (Scheme 3). Neither the selenyl compounds **1** or **2** nor the cation of the corresponding halide salt showed any significant shift in the <sup>1</sup>H, <sup>19</sup>F, <sup>31</sup>P or <sup>77</sup>Se NMR spectra, thus there is no clear evidence for the existence of an interaction from NMR spec-

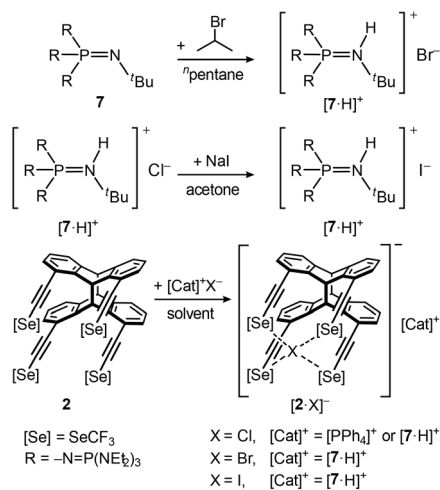


**Fig. 3** Molecular structure of [1:2 THF]<sub>2</sub> in the solid state. Hydrogen atoms are omitted for clarity. The red dotted lines mark Se...O distances below the sum of van der Waals radii.<sup>35</sup> Ellipsoids are set at 50% probability. Atoms marked with ' are generated by symmetry (1 – x, 1 – y, 1 – z). Selected distances [Å] and angles [°]: Se(1)...Se(2) 4.782(1), Se(1)–C(16) 1.844(3), Se(1)–C(17) 1.969(3), Se(2)–C(19) 1.838(3), Se(2)–C(20) 1.972(3); C(16)–Se(1)–C(17) 92.7(1), C(19)–Se(2)–C(20) 93.0(1), C(16)–Se(1)–O(2') 177.0(1), C(17)–Se(1)–O(1) 172.3(1), C(19)–Se(2)–O(1') 165.0(1), C(20)–Se(2)–O(2) 172.9(1), Se(1)–O(1)–Se(2') 79.5(1), Se(2)–O(2)–Se(1') 80.6(1).

troscopy in solution. Changing the concentration did also not increase this chemical shift difference and the solvent cannot be varied at will due to the limited solubility of the compound. Because the overall shifts are so small (<sup>1</sup>H, <sup>19</sup>F and even <sup>77</sup>Se), it was not possible to determine an association constant *K*<sub>a</sub> from the binding isotherms, as we did for the antimony systems in our previous work.

Consequently, we attempted to obtain structural evidence. However, the production of single crystals suitable for X-ray





**Scheme 3** Synthesis of the phosphazene base hydrohalides  $[7\text{-H}]\text{Br}$  and  $[7\text{-H}]\text{I}$  and adduct formation of host system **2** with various halide salts.

diffraction was quite challenging, requiring numerous attempts to find suitable crystallization conditions, including testing different cations. Thus, for the bidentate system **1**, we found no evidence for adduct formation with halides, despite the use of weakly coordinating cations.

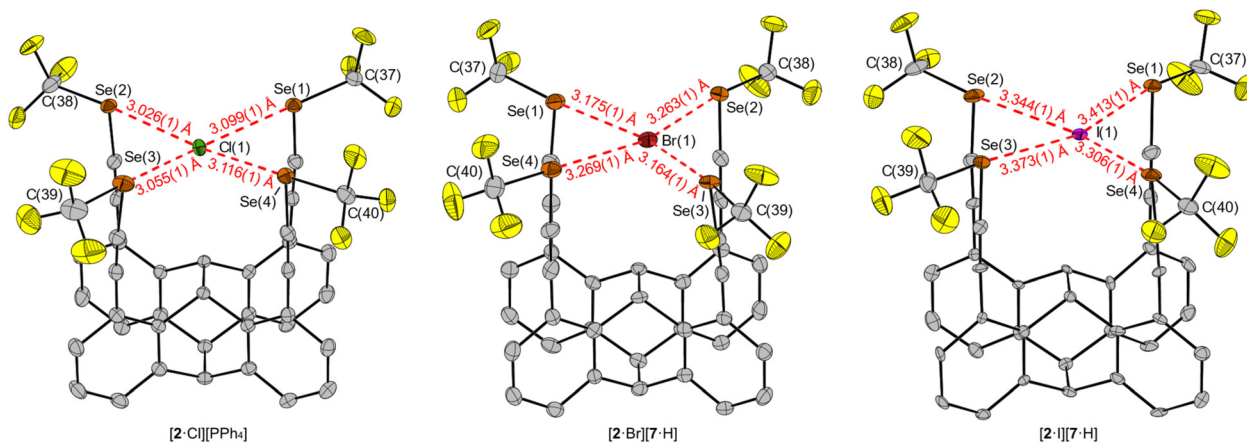
In contrast, for the tetradentate host **2**, the chloride ion adduct  $[2\text{-Cl}]^-$  could be crystallized when **2** was combined with tetraphenylphosphonium chloride. However, the same recipe did not work for tetraphenylphosphonium bromide and iodide. Therefore, we moved on to even weaker coordinating cations. The hydrohalides of phosphazene superbases as reported by Schwesinger *et al.*<sup>37</sup> are ideally suited for this

purpose due to their very low electrophilicity. The ethyl derivative of the phosphazene base **7**, recently introduced by Hoge *et al.*, has an even larger volume,<sup>38</sup> further optimizing the ion separation. Its hydrochloride  $[7\text{-H}]\text{Cl}$  proved to be a suitable alternative for the crystallization of the chloride adduct  $[2\text{-Cl}]^-\text{[Cat]}^+$  ( $[\text{Cat}]^+ = [7\text{-H}]^+$  instead of  $[\text{PPh}_4]^+$ ).

In order to provide access to the difficult-to-reach bromide and iodide adducts of **2**, we developed a synthesis of the corresponding hydrohalides  $[7\text{-H}]\text{Br}$  and  $[7\text{-H}]\text{I}$  (Scheme 3). The preparation of the hydrobromide  $[7\text{-H}]\text{Br}$  was achieved by deprotonation of 2-bromopropane with superbases **7**, while  $[7\text{-H}]\text{I}$  could be synthesized by an anion exchange of the hydrochloride  $[7\text{-H}]\text{Cl}$  with sodium iodide under formation of sodium chloride. Suitable single-crystals of  $[7\text{-H}]\text{Br}$  and  $[7\text{-H}]\text{I}$  were grown by cooling a solution in dichloromethane and acetone, respectively. The solid state structures of  $[7\text{-H}]\text{Br}$  and  $[7\text{-H}]\text{I}$  are shown in Fig. S35 of the ESI.† Similar to the structure of  $[7\text{-H}]\text{Cl}$ ,<sup>38</sup> the P–N bond lengths including the  $\text{N(H)}^+\text{Bu}$  group are significantly longer than the other three P–N bonds in both structures.

With the salts  $[7\text{-H}]\text{X}$ , we succeeded in preparing the adducts of the tetradentate host **2**:  $[2\text{-X}]^-\text{[Cat]}^+$  (X = Cl, Br, I). Single crystals were grown by cooling saturated solutions of equimolar mixtures of  $[7\text{-H}]\text{X}$  and **2** in benzene/dichloromethane (12 : 1). Elemental analyses of the formed precipitates confirmed a 1 : 1 stoichiometry of the bulk materials of  $[2\text{-X}]^-\text{[Cat]}^+$  (X = Cl, Br, I;  $[\text{Cat}]^+ = [7\text{-H}]^+$ ,  $[\text{PPh}_4]^+$ ). The <sup>1</sup>H, <sup>77</sup>Se and <sup>19</sup>F NMR spectra of these adducts in CDCl<sub>3</sub> are mostly the sum of the signals of host **2** and the respective salt, and show almost no chemical shifts.

The solid state structures of all halide adducts show Se<sub>4</sub>X motifs, in which the halogen atom is surrounded by four sel-



**Fig. 4** Molecular structures of  $[2\text{-Cl}][\text{PPh}_4]$ ,  $[2\text{-Br}][7\text{-H}]$  and  $[2\text{-I}][7\text{-H}]$  in the solid state. Hydrogen atoms, solvent molecules (benzene) and the cations are omitted for clarity. In case of  $[2\text{-Br}][7\text{-H}]$  and  $[2\text{-I}][7\text{-H}]$ , the crystal partially contained the adduct  $[2\text{-OH}]^-$  (more information on this is provided in the ESI†). Ellipsoids are set at 50% probability. The red dotted lines mark Se...X distances below the sum of van der Waals radii.<sup>35</sup> Selected distances [Å] and angles [°] of  $[2\text{-Cl}][\text{PPh}_4]$ : C(37)–Se(1) 1.970(2), C(38)–Se(2) 1.970(2), C(39)–Se(3) 1.967(2), C(40)–Se(4) 1.968(2); C(37)–Se(1)⋯C(1) 177.3(1), C(38)–Se(2)⋯C(1) 174.5(1), C(39)–Se(3)⋯C(1) 178.3(1), C(40)–Se(4)⋯C(1) 173.6(1); of  $[2\text{-Br}][7\text{-H}]$ : C(37)–Se(1) 1.980(6), C(38)–Se(2) 1.969(6), C(39)–Se(3) 1.981(6), C(40)–Se(4) 1.975(6); C(37)–Se(1)⋯Br(1) 174.3(2), C(38)–Se(2)⋯Br(1) 173.4(2), C(39)–Se(3)⋯Br(1) 173.2(2), C(40)–Se(4)⋯Br(1) 176.8(2); of  $[2\text{-I}][7\text{-H}]$ : C(37)–Se(1) 1.973(9), C(38)–Se(2) 1.965(8), C(39)–Se(3) 1.960(9), C(40)–Se(4) 1.977(9); C(37)–Se(1)⋯I(1) 171.2(3), C(38)–Se(2)⋯I(1) 174.7(3), C(39)–Se(3)⋯I(1) 174.2(2), C(40)–Se(4)⋯I(1) 172.9(2).



enium atoms in a roughly rectangular planar fashion. Fig. 4 shows the structure of  $[2\cdot\text{Cl}]^-[\text{PPh}_4]^+$ , which is similar to that of  $[2\cdot\text{Cl}]^-[\text{7}\cdot\text{H}]^+$ , which is why the latter is described in the ESI.† The unusual coordination geometry found in the  $[2\cdot\text{X}]^-$  anions is comparable to the halide ion adducts of the tetrastibanyl compound reported previously by us.<sup>1</sup> However, the tetraselanyl compound **2** seems to be better suited for the chelation of halide ions in its central cavity, since the comparably small  $\text{C}\equiv\text{C}-\text{Se}(\text{CF}_3)$  units are much more sterically unhindered than the larger  $\text{C}\equiv\text{C}-\text{Sb}(\text{C}_2\text{F}_5)_2$  units. Due to the ability of the  $\text{C}\equiv\text{C}-\text{Se}(\text{CF}_3)$  units to rotate about the  $\text{C}\equiv\text{C}$ -axis, the  $\sigma$ -holes of each selenium atom are directed right towards the centrally chelated halide ion in all three adducts, resulting in collinear  $\text{X}\cdots\text{Se}-\text{C}^{\text{F}}$  motifs (angles vary between  $171.2(3)^\circ$  and  $178.3(1)^\circ$ ). The four trifluoromethyl groups are oriented outward, forming a cross with the halide ion in their center. In each structure, the halide ion is near the geometric center of the four selenium atoms ( $d(\text{X}\cdots\text{Se}_4\text{-centroid})$  Cl: 0.395(1), Br: 0.393(1), I: 0.438(1) Å). Each  $\text{X}\cdots\text{Se}$  distance is below the respective sum of the van der Waals radii ( $\sum d_{\text{vdW}}(\text{Se},\text{X})$ : Cl = 3.65, Br = 3.73, I = 3.88 Å),<sup>35</sup> which also confirms four attractive  $\sigma$ -hole interactions. In contrast to the  $\text{Sb}-\text{C}^{\text{F}}$  bonds in the tetrastibanyl compound,<sup>1</sup> we observed no significant stretching of the  $\text{Se}-\text{C}^{\text{F}}$  bonds due to an electron donation into the  $\sigma^*(\text{Se}-\text{C})$  orbitals.

For the adducts  $[2\cdot\text{Br}][\text{7}\cdot\text{H}]$  and  $[2\cdot\text{I}][\text{7}\cdot\text{H}]$ , small occupancies of  $\text{OH}^-$  ions (23% and 25%) at the halide positions were observed in the single crystals by X-ray diffraction. This is most likely due to traces of  $[\text{7}\cdot\text{H}]\text{OH}$  in the bromide and iodide salts used, but it appears that these  $\text{OH}^-$  ions are preferentially incorporated into the crystals rather than the mostly microcrystalline or amorphous bulk material of the adducts. No such problems were encountered for  $[2\cdot\text{Cl}][\text{7}\cdot\text{H}]$  (which is described in the ESI†).

## Conclusions

In essence we have demonstrated a quantitative-yielding approach to a bidentate and the first tetradentate chalcogen bonding host system through a tin-selenium exchange reaction with  $\text{ClSeCF}_3$ . The introduction of  $\text{SeCF}_3$  groups into molecules is of particular interest for material or life sciences and can be achieved by various methods,<sup>39</sup> but such a tin-selenium exchange reaction has only been reported for  $\text{PhSeCl}$ .<sup>34</sup>

The ability of the  $\text{SeCF}_3$  group to function as a chalcogen bonding donor is already evident in the molecular structures of both the free host **2** and  $\text{ClSeCF}_3$  in the form of intermolecular  $\text{Se}\cdots\text{F}$  and  $\text{Se}\cdots\text{Se}$  contacts. The bidentate system **1** forms an adduct with THF,  $[\text{1}\cdot\text{2 THF}]_2$ , in which **1** uses of both types of its  $\sigma$ -holes. Remarkable NMR shifts of the host-guest complexes could not be observed in the  $^1\text{H}$ ,  $^{19}\text{F}$ ,  $^{31}\text{P}$  or  $^{77}\text{Se}$  NMR spectra, but the molecular structures for a series of halide adducts of the type  $[2\cdot\text{X}]^-$  clearly demonstrated the affinity of the tetradentate host for anionic guests. Advantageous in this context was the use of halide salts of extremely weak coordinating cations, which we found in the

hydrohalides of  $[\text{7}\cdot\text{H}]\text{Br}$  and  $[\text{7}\cdot\text{H}]\text{I}$  of the Schwesinger base  $[(\text{Et}_2\text{N})_3\text{PN}]_3\text{PN}^t\text{Bu}$  (**7**). In the halide adducts, **2** aligns all four  $\sigma$ -holes of its  $\text{C}\equiv\text{C}-\text{SeCF}_3$  units toward the trapped central halide ion, forming unique planar  $\text{Se}_4\text{X}$  motifs. All  $\text{X}\cdots\text{Se}-\text{CF}_3$  chalcogen bonding units are linear and the  $\text{X}\cdots\text{Se}$  distances are less than the sum of van der Waals radii. Although there is evidence that the  $\sigma$ -holes at the opposite of alkyne substituents are more activated than those opposite of  $\text{CF}_3$  groups,<sup>40</sup> in our system only those towards  $\text{CF}_3$  groups are addressed, probably due to the cooperative effects. Because the compounds are stable in the presence of water, air, and even strong aqueous acids, they have potential for applications in the fields of catalysis and sensing. Initial experiments have shown a catalytic activity of these compounds for the transfer hydrogenation of quinoline derivatives with Hantzsch ester, a common benchmark reaction for such  $\sigma$ -hole compounds.<sup>25,26</sup>

## Author contributions

J. L. Beckmann: investigation, methodology, validation, visualization, writing (original draft), N. Tiessen: investigation (supporting synthesis), B. Neumann: investigation (SCXRD), H.-G. Stammer: investigation (SCXRD), N. W. Mitzel and B. Hoge: funding acquisition, project administration, supervision, reviewing and editing.

## Data availability

The data published in this contribution are available as ESI,† submitted with the manuscript.

## Conflicts of interest

There are no conflicts to declare.

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## References

- J. L. Beckmann, J. Krief, Y. V. Vishnevskiy, B. Neumann, H.-G. Stammer and N. W. Mitzel, *Chem. Sci.*, 2023, **14**, 13551.
- G. E. Garrett, E. I. Carrera, D. S. Seferos and M. S. Taylor, *Chem. Commun.*, 2016, **52**, 9881.



- 3 (a) M. G. Sarwar, B. Dragisic, S. Sagoo and M. S. Taylor, *Angew. Chem., Int. Ed.*, 2010, **49**, 1674; (b) D. Bulfield and S. M. Huber, *Chem. – Eur. J.*, 2016, **22**, 14434; (c) S. H. Jungbauer, D. Bulfield, F. Kniep, C. W. Lehmann, E. Herdtweck and S. M. Huber, *J. Am. Chem. Soc.*, 2014, **136**, 16740.
- 4 P. Wonner, A. Dreger, L. Vogel, E. Engelage and S. M. Huber, *Angew. Chem., Int. Ed.*, 2019, **58**, 16923.
- 5 P. Wonner, L. Vogel, M. Düser, L. Gomes, F. Kniep, B. Mallick, D. B. Werz and S. M. Huber, *Angew. Chem., Int. Ed.*, 2017, **56**, 12009.
- 6 L. Vogel, P. Wonner and S. M. Huber, *Angew. Chem., Int. Ed.*, 2019, **58**, 1880.
- 7 (a) M. H. H. Voelkel, P. Wonner and S. M. Huber, *Chem. Open*, 2020, **9**, 214; (b) R. L. Sutar, E. Engelage, R. Stoll and S. M. Huber, *Angew. Chem., Int. Ed.*, 2020, **59**, 6806.
- 8 S. H. Jungbauer and S. M. Huber, *J. Am. Chem. Soc.*, 2015, **137**, 12110.
- 9 M. Breugst and J. J. Koenig, *Eur. J. Org. Chem.*, 2020, 5473.
- 10 A. C. Keuper, K. Fengler, F. Ostler, T. Danelzik, D. G. Piekarski and O. García Mancheño, *Angew. Chem., Int. Ed.*, 2023, **62**, e202304781.
- 11 H. Zhao and F. P. Gabbaï, *Nat. Chem.*, 2010, **2**, 984.
- 12 (a) S. M. Walter, F. Kniep, L. Rout, F. P. Schmidtchen, E. Herdtweck and S. M. Huber, *J. Am. Chem. Soc.*, 2012, **134**, 8507; (b) E. Navarro-García, B. Galmés, M. D. Velasco, A. Frontera and A. Caballero, *Chem. – Eur. J.*, 2020, **26**, 4706; (c) A. Brown and P. D. Beer, *Chem. Commun.*, 2016, **52**, 8645; (d) J. Pancholi and P. D. Beer, *Coord. Chem. Rev.*, 2020, **416**, 213281; (e) A. Docker, C. H. Guthrie, H. Kuhn and P. D. Beer, *Angew. Chem., Int. Ed.*, 2021, **60**, 21973; (f) J. Y. Lim and P. D. Beer, *Chem*, 2018, **4**, 731.
- 13 R. Hein, A. Docker, J. J. Davis and P. D. Beer, *J. Am. Chem. Soc.*, 2022, **144**, 8827.
- 14 (a) L. E. Bickerton, A. Docker, A. J. Sterling, H. Kuhn, F. Duarte, P. D. Beer and M. J. Langton, *Chem. – Eur. J.*, 2021, **27**, 11738; (b) S. Benz, M. Macchione, Q. Verolet, J. Mareda, N. Sakai and S. Matile, *J. Am. Chem. Soc.*, 2016, **138**, 9093.
- 15 L. M. Lee, M. Tsemperouli, A. I. Poblador-Bahamonde, S. Benz, N. Sakai, K. Sugihara and S. Matile, *J. Am. Chem. Soc.*, 2019, **141**, 810.
- 16 G. Park and F. P. Gabbaï, *Chem. Sci.*, 2020, **11**, 10107.
- 17 (a) G. R. Desiraju, P. S. Ho, L. Kloo, A. C. Legon, R. Marquardt, P. Metrangolo, P. Politzer, G. Resnati and K. Rissanen, *Pure Appl. Chem.*, 2013, **85**, 1711; (b) G. Cavallo, P. Metrangolo, R. Milani, T. Pilati, A. Priimagi, G. Resnati and G. Terraneo, *Chem. Rev.*, 2016, **116**, 2478.
- 18 (a) P. Metrangolo and G. Resnati, Halogen Bonding – Fundamentals and Applications, in *Structure and Bonding*, Springer, Berlin, 2010, vol. 126; (b) P. Metrangolo and G. Resnati, Halogen Bonding I: Impact on Materials Chemistry and Life Sciences, in *Top. Curr. Chem.*, Springer, Berlin, 2015, vol. 358; (c) Halogen Bonding II: Impact on Materials Chemistry and Life Sciences, in *Top. Curr. Chem.*, Springer, Berlin, 2015, vol. 359.
- 19 (a) O. Hassel, *Science*, 1970, **170**, 497; (b) L. Turunen, J. H. Hansen and M. Erdélyi, *Chem. Rec.*, 2021, **21**, 1252.
- 20 C. B. Aakeroy, D. L. Bryce, G. R. Desiraju, A. Frontera, A. C. Legon, F. Nicotra, K. Rissanen, S. Scheiner, G. Terraneo, P. Metrangolo and G. Resnati, *Pure Appl. Chem.*, 2019, **91**, 1889.
- 21 P. Scilabra, G. Terraneo and G. Resnati, *Acc. Chem. Res.*, 2019, **52**(5), 1313.
- 22 K. T. Mahmudov, M. N. Kopylovich, M. F. C. Da Guedes Silva and A. J. L. Pombeiro, *Dalton Trans.*, 2017, **46**, 10121.
- 23 (a) J. Zhang, J. Wei, W.-Y. Ding, S. Li, S.-H. Xiang and B. Tan, *J. Am. Chem. Soc.*, 2021, **143**, 6382; (b) A. Frontera and A. Bauza, *Int. J. Mol. Sci.*, 2021, **22**, 12550.
- 24 B. Zhou and F. P. Gabbaï, *Organometallics*, 2021, **40**, 2371.
- 25 S. Benz, J. López-Andarias, J. Mareda, N. Sakai and S. Matile, *Angew. Chem., Int. Ed.*, 2017, **56**, 812.
- 26 S. Benz, J. Mareda, C. Besnard, N. Sakai and S. Matile, *Chem. Sci.*, 2017, **8**, 8164.
- 27 P. Wonner, T. Steinke and S. M. Huber, *Synlett*, 2019, 1673.
- 28 S. Benz, A. I. Poblador-Bahamonde, N. Low-Ders and S. Matile, *Angew. Chem., Int. Ed.*, 2018, **57**, 5408.
- 29 J. L. Beckmann, J. Krief, Y. V. Vishnevskiy, B. Neumann, H.-G. Stammer and N. W. Mitzel, *Angew. Chem., Int. Ed.*, 2023, **62**, e202310439.
- 30 J.-H. Lamm, J. Glatthor, J.-H. Weddeling, A. Mix, J. Chmiel, B. Neumann, H.-G. Stammer and N. W. Mitzel, *Org. Biomol. Chem.*, 2014, **12**, 7355.
- 31 P. Niermeier, K. A. M. Maibom, J.-H. Lamm, B. Neumann, H.-G. Stammer and N. W. Mitzel, *Chem. Sci.*, 2021, **12**, 7943.
- 32 E. Magnier and C. Wakselman, *Collect. Czech. Chem. Commun.*, 2002, **67**, 1262.
- 33 T. Billard, S. Large and B. R. Langlois, *Tetrahedron Lett.*, 1997, **38**, 65.
- 34 H. Lang, H. Keller, W. Imhof and S. Martin, *Chem. Ber.*, 1990, **123**, 417.
- 35 A. Bondi, *J. Phys. Chem.*, 1964, **68**, 441.
- 36 K. T. Mahmudov, A. V. Gurbanov, V. A. Aliyeva, M. F. C. G. Da Silva, G. Resnati and A. J. Pombeiro, *Coord. Chem. Rev.*, 2022, **464**, 214556.
- 37 (a) R. Schwesinger, H. Schlemper, C. Hasenfratz, J. Willaredt, T. Dambacher, T. Breuer, C. Ottaway, M. Fletschinger, J. Boele, H. Fritz, D. Putzas, H. W. Rotter, F. G. Bordwell, A. V. Satish, G.-Z. Ji, E.-M. Peters, K. Peters, H. G. von Schnering and L. Walz, *Liebigs Ann.*, 1996, **1996**, 1055; (b) R. Schwesinger and H. Schlemper, *Angew. Chem., Int. Ed. Engl.*, 1987, **26**, 1167; (c) R. Schwesinger, C. Hasenfratz, H. Schlemper, L. Walz, E.-M. Peters, K. Peters and H. G. von Schnering, *Angew. Chem., Int. Ed. Engl.*, 1993, **32**, 1361.
- 38 R. F. Weitkamp, B. Neumann, H.-G. Stammer and B. Hoge, *Angew. Chem., Int. Ed.*, 2019, **58**, 14633.
- 39 (a) C. Ghiazza, T. Billard and A. Tlili, *Chem. – Eur. J.*, 2017, **23**, 10013; (b) A. Tlili, E. Ismalaj, Q. Glenadel, C. Ghiazza and T. Billard, *Chem. – Eur. J.*, 2018, **24**, 3659; (c) X.-L. Chen, S.-H. Zhou, J.-H. Lin, Q.-H. Deng and J.-C. Xiao, *Chem. Commun.*, 2019, **55**, 1410.
- 40 V. Mamane, P. Peluso, E. Aubert, R. Weiss, E. Wenger, S. Cossu and P. Pale, *Organometallics*, 2020, **39**, 3936.

