

## PAPER

[View Article Online](#)  
[View Journal](#) | [View Issue](#)Cite this: *Dalton Trans.*, 2024, **53**, 8948

## Orthopalladated imidazolones and thiazolones: synthesis, photophysical properties and photochemical reactivity†

Darius Dumitraş,<sup>a</sup> David Dalmau,<sup>b</sup> Pilar García-Orduña,<sup>b</sup> Alexandra Pop,<sup>a</sup> Anca Silvestru<sup>a</sup> and Esteban P. Urriolabeitia<sup>\*b</sup>

The reaction of  $\text{Pd}(\text{OAc})_2$  with (Z)-5-arylidene-4-(4H)-imidazolones (**2a–e**) and (Z)-4-arylidene-5(4H)-thiazolones (**3a–e**) in trifluoroacetic acid results in the corresponding orthopalladated dinuclear complexes (**4a–e**, imidazolones; **11a–d**, thiazolones) with trifluoroacetate bridges through regioselective C–H activation at the *ortho* position of the 4-arylidene group. Compound **4e**, which contains an imidazolone substituted at 2- and 4-positions of the arylidene ring with methoxide groups and exhibits strong push–pull charge transfer, is an excellent precursor for the synthesis of fluorescent complexes with green yellowish emission and remarkable quantum yields. Breaking the bridging system with pyridine yields the mononuclear complex **5e** ( $\Phi_F = 5\%$ ), while metathesis of trifluoroacetate ligands with chloride leads to the dinuclear complex **6e**, also a precursor of fluorescent complexes by breaking the chloride bridging system with pyridine (**7e**,  $\Phi_F = 7\%$ ), or by substitution of chloride ligands with pyridine (**8e**,  $\Phi_F = 15\%$ ) or acetylacetonate (**9e**,  $\Phi_F = 2\%$ ). In addition to notable photophysical properties, dinuclear complexes **4** and **11** also exhibit significant photochemical reactivity. Thus, irradiation of orthopalladates **4a–c** and **11a–c** in  $\text{CH}_2\text{Cl}_2$  with blue light (465 nm) proceeds via [2 + 2] photocycloaddition of the C=C double bonds of imidazolone and thiazolone ligands, yielding the corresponding cyclobutane-bridging diaminotruillic derivatives **10a–c** and **12a–c**, respectively.

Received 11th March 2024,

Accepted 1st May 2024

DOI: 10.1039/d4dt00730a

[rsc.li/dalton](http://rsc.li/dalton)

## Introduction

The use of light as a reagent to promote new properties, such as luminescence, and new reactivities, such as photochemistry, has been known for over 100 years.<sup>1</sup> Although there has been sustained interest and development since then, it has only been in the last two decades that an unprecedented development has been experienced, and continues so,<sup>2–8</sup> due to its numerous and important applications in lighting, imaging, or photoinduced synthetic processes.<sup>9–11</sup> In both cases, whether

inducing luminescence or promoting chemical processes through light is considered, molecules capable of absorbing light, transitioning to an excited state, and subsequently emitting light or initiating a chemical reaction are required.

Among the different photoreactive molecular platforms, azlactone-type substrates, such as oxazolones, have demonstrated excellent properties from both the photophysical and photochemical perspectives. They exhibit intense fluorescence in the solid state,<sup>12,13</sup> absorption of two photons,<sup>14</sup> and non-linear optical (NLO) properties.<sup>15</sup> Additionally, from a photochemical standpoint, they undergo 2 + 2 photocycloadditions and other light-induced processes, resulting in diaminotruillic and diaminotruillic-type cyclobutanes that are highly interesting due to their pharmacological properties.<sup>16–19</sup>

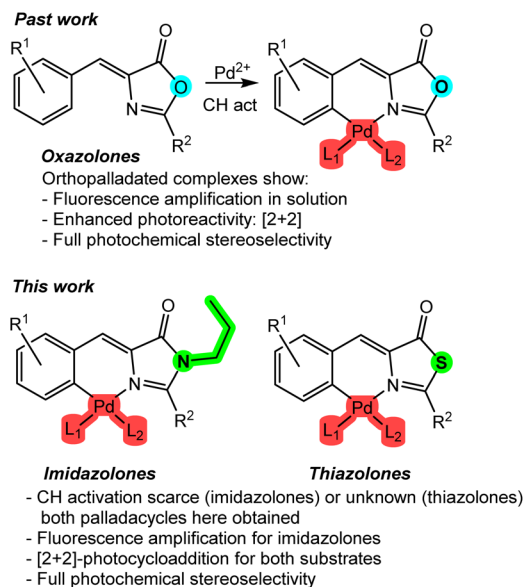
It is evident that the incorporation of a transition metal into the skeleton of a specific organic ligand will have a profound effect on both its photophysical and photochemical properties. This fact is particularly relevant and highly impactful in the case of oxazolones for several reasons. Regarding their photophysical properties, the incorporation of a metal such as palladium as an intramolecular lock between the 4-arylidene ring and the heterocycle<sup>20</sup> allows for the amplification of the fluorescence in solution of certain substrates by several orders

<sup>a</sup>Supramolecular Organic and Organometallic Chemistry Centre, Department of Chemistry, Faculty of Chemistry and Chemical Engineering, Babeş-Bolyai University, Arany Janos 11, 400028 Cluj-Napoca, Romania

<sup>b</sup>Instituto de Síntesis Química y Catálisis Homogénea, ISQCH, CSIC-Universidad de Zaragoza, Pedro Cerbuna 12, 50009 Zaragoza, Spain. E-mail: [esteban.u.a@csic.es](mailto:esteban.u.a@csic.es)

† Electronic supplementary information (ESI) available: Complete experimental section with references, copies of the NMR spectra of all new prepared species, absorption and excitation–emission spectra of imidazolones **2** and complexes **5e–9e**, computational details and Cartesian coordinates of all calculated species. CCDC 2332942–2332944, 2336747 and 2336748. For ESI and crystallographic data in CIF or other electronic format see DOI: <https://doi.org/10.1039/d4dt00730a>





**Fig. 1** Past work: consequences of the incorporation of Pd into the molecular skeleton of oxazolones on their photophysical properties and photochemical reactivity. This work: results obtained from the palladation of imidazolones and thiazolones concerning their luminescence and photochemical reactivity.

of magnitude (Fig. 1, past work),<sup>21–25</sup> because well-characterized non-radiative deactivation pathways, such as the hula-twist or *Z-E* isomerizations, are restricted.<sup>26–30</sup> Regarding photochemical reactivity, the formation of dinuclear palladacycles is the basis for the building of molecular templates that enable a stereoselective [2 + 2] photocycloaddition in solution, resulting in diaminotruilic-type cyclobutanes as single diastereoisomers,<sup>31–34</sup> in sharp contrast to what occurs with free-state oxazolones.<sup>16</sup> Therefore, the significant influence of the incorporation of a metal into the skeleton of these species is evident.

In this work, the study of the impact of the incorporation of Pd on the photophysical and photochemical properties of two families of compounds structurally related to oxazolones has been addressed. These families include imidazolones and thiazolones (Fig. 1, this work). Imidazolone-type substrates are of particular interest from the fluorescence perspective, as they are chromophores of GFP (green fluorescent protein) and have well-established luminescence properties.<sup>35–37</sup> Similar to what has been reported for oxazolones in previous paragraphs, they completely lose fluorescence in solution.<sup>26–30</sup> There are very few precedents for fluorescence recovery, and they typically use only boron as an intramolecular lock.<sup>38–40</sup> The use of Pd for these substrates is virtually unexplored.<sup>21</sup> On the other hand, contrary to what might seem, the structural similarity of imidazolones and thiazolones to oxazolones does not unequivocally imply the same reactivity. For example, there is only one isolated example of [2 + 2] photocycloaddition involving free imidazolones,<sup>41</sup> and no photocycloadditions are known under other conditions. Referring to thiazolone-type species, they are

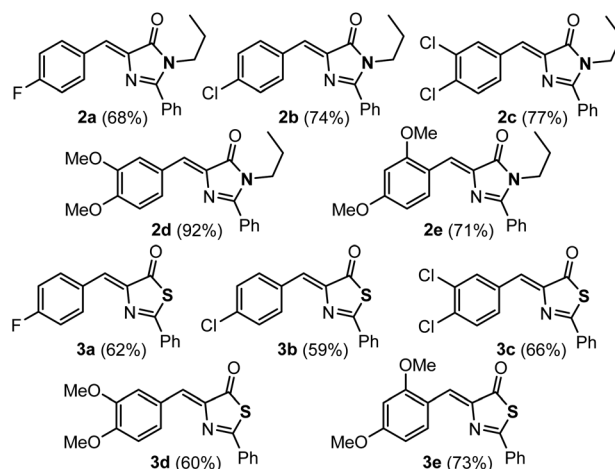
even less studied than imidazolones and oxazolones. In this case, there are no fluorescence studies, no incorporation of any metal into their molecular skeleton has been described, and their photochemical reactivity is also very little explored.<sup>42</sup>

## Results and discussion

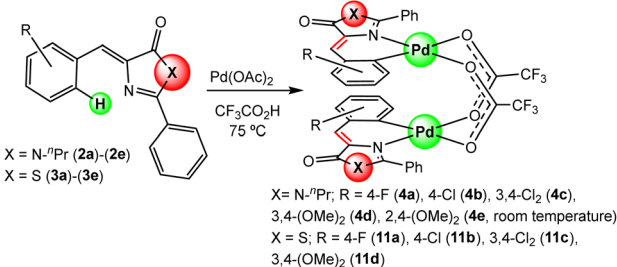
### Synthesis and characterization of orthopalladated complexes from imidazolones 2 and thiazolones 3

The imidazolones 2a–2e and thiazolones 3a–3e shown in Fig. 2 have been synthesized from the corresponding oxazolones 1a–1e,<sup>43–46</sup> following reported procedures.<sup>21,42,47–50</sup> Representatives with both electron-donating and electron-withdrawing substituents at different positions of the arylidene ring have been chosen to encompass a broader structural variety.

The orthopalladation of the whole set of imidazolones 2 and thiazolones 3 has been attempted through CH bond activation processes, by the reaction of 2 or 3 with Pd(OAc)<sub>2</sub> (1 : 1 molar ratio; OAc = acetate) in trifluoroacetic acid CF<sub>3</sub>CO<sub>2</sub>H.<sup>20</sup> The incorporation of Pd into the different molecular scaffolds was successful in all attempted cases, except in the case of thiazolone 3e, giving the corresponding dimers with carboxylate bridges 4a–4e or 11a–11d as shown in Fig. 3.



**Fig. 2** Imidazolones 2a–2e and thiazolones 3a–3e used in this work.



**Fig. 3** Synthesis of the orthopalladated derivatives 4a–4e (imidazolones) and 11a–11d (thiazolones).



The reaction was carried out by heating for 2 h at 75 °C for most of the cases. Only in the case of imidazolone **2e** the reaction was performed at room temperature for 17 h; partial conversion was observed at room temperature at shorter reaction times, while decomposition occurred if heat was applied. For thiazolone **3e** several different reaction conditions were attempted (temperatures, reaction times, solvents), but the reaction failed in all cases. The yields of isolated compounds **4a–4e** and **11a–11d** are in most of the cases higher than 78%, showing that the orthopalladation of azlactones through C–H bond activation is a general process.

In dimers **4a–4e** and **11a–11d**, the incorporation of Pd occurred regioselectively at the *ortho* position of the phenyl ring of the arylidene group, following the trend observed in oxazolones.<sup>20</sup> In the cases where two inequivalent accessible positions are present (positions 2 and 6; **2c**, **2d**, **3c**, **3d**), palladation took place at the less sterically hindered position (position 6), as deduced from their <sup>1</sup>H and <sup>13</sup>C NMR spectra. The dinuclear nature of these derivatives was inferred from the observation of peaks in their high-resolution mass spectra (HRMS) with the stoichiometry and isotopic distribution corresponding to the  $[M + Na]^+$  cations, where M is the dinuclear complex shown in Fig. 3 (see the ESI†). The relative arrangement of the two orthopalladated fragments is *transoid*, as deduced from the observation of a single signal in the <sup>19</sup>F NMR spectra, because this configuration minimizes intramolecular repulsions.

The determination of the crystal structures of complexes containing orthopalladated imidazolones **4c** (Fig. 4), and thiazolones **11a** (Fig. 5) and **11b** (Fig. 6) provides additional structural information. In all the studied cases, the asymmetric unit is composed of half of the molecule, and then, the dinuclear molecules are formed by two symmetry-related monomers, linked through the carboxylate bridges. The three structures shared common structural features: in all of them the expected open-book structure is observed, placing the two coordination planes containing palladium in close proximity, with an inter-metallic distance of 3.0398(2) Å, 3.1569(4) Å and 3.0953(3) Å, in **4c**, **11a** and **11c**, respectively. Although the immediate

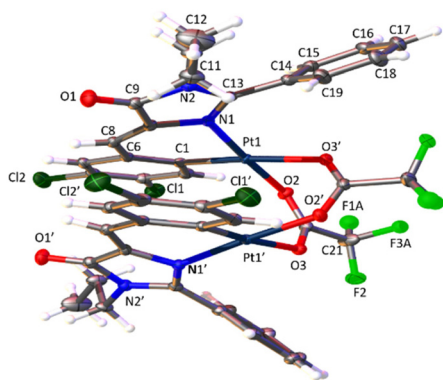


Fig. 4 Molecular structure of cyclopalladated **4c**. Thermal ellipsoids are drawn at 50% probability. Symmetry code: 1 – x, y, 3/2 – z.

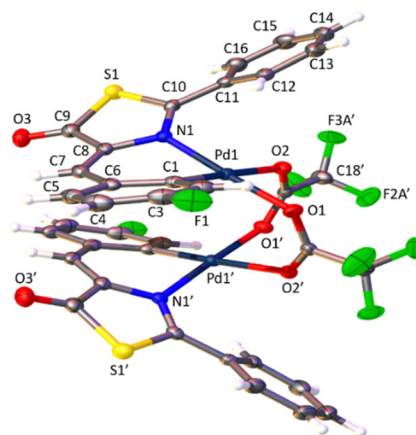


Fig. 5 Molecular structure of cyclopalladated **11a**. Thermal ellipsoids are drawn at 50% probability. For clarity, only the major component of the disordered fragment has been depicted. Symmetry code: 1 – x, y, 1/2 – z.

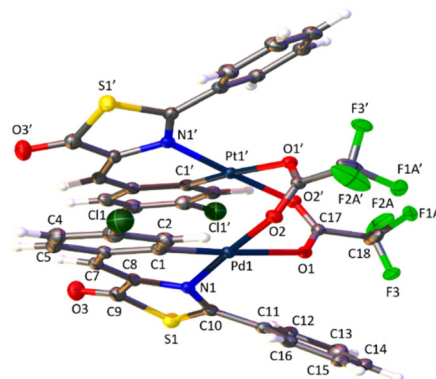


Fig. 6 Molecular structure of cyclopalladated **11b**. Thermal ellipsoids are drawn at 50% probability. For clarity, only the major component of the disordered fragment have been depicted. Symmetry code: 1 – x, y, 3/2 – z.

environments around Pd are square-planar in all three structures, the molecules are highly distorted. Thus, the imidazolone moiety in **4c** and thiazolone groups in **11a** and **11b** deviate significantly from planarity, adopting a U-shaped form, and similarly, they largely deviate from coplanarity with respect to the square-planar metal coordination environments. Both types of distortions are analogous to those recently described in mononuclear complexes derived from oxazolones,<sup>22,24</sup> suggesting that they are inherent to this type of structural skeleton and not dependent on the nuclearity of the resulting complexes or the ancillary ligands of palladium. It is noteworthy that the free imidazolone ligands exhibit remarkable coplanarity of the arylidene and heterocycle rings,<sup>51–55</sup> which is lost upon metal incorporation.

Structures involving unsaturated 4-arylidene-5(4*H*)-thiazolones have not been reported to date by X-ray analysis; therefore no proper comparisons can be made. In each complex, the corresponding imidazolone or thiazolone acts as a C<sup>^</sup>N-



chelating ligand, and the *transoid* relative arrangement of the two orthopalladated ligands is confirmed, as was deduced from the observation of single peaks in the  $^{19}\text{F}$  NMR spectra. Despite the distortion, the internal structural parameters of the orthopalladated arylidene-imidazolone fragment show no significant deviations from the uncomplexed situation, and in general, no relevant variations in bond distances and angles are observed compared to related structural situations found in the literature.<sup>56</sup>

### Synthesis, characterization and photophysical properties of complexes 5e–9e

A preliminary screening with the orthometallated complexes of imidazolones **4a–4e** and thiazolones **11a–11d** shows that the most promising reactivity from the perspective of preparing fluorescent complexes is realized in the derivative of the dinuclear complex **4e**, as shown in Fig. 7. The complexes **5e–9e** exhibit fluorescence, albeit with different intensities, as will be discussed later. Attempts to develop reactivity similar to that shown in Fig. 7 with other imidazolone derivatives (*e.g.*, **4d**) did not lead to fluorescent complexes (results not shown). Moreover, the reactivity of thiazolone (**11d**) in reactions analogous to those shown in Fig. 7 resulted in all cases in complex mixtures of products that also did not exhibit fluorescence. Although the results of both reactivities are not shown here, it has been possible to characterize the mononuclear derivatives resulting from the trifluoroacetate bridge-breaking with pyridine in **4d** and **11d** through X-ray diffraction (CCDC codes 2336747 and 2336748†).

Given that in analogous structural situations, both pyridine and trifluoroacetate, chloride, and acetylacetonate ligands have proven to be good ancillary ligands for the development of fluorescence in orthopalladated azlactones, covering an appropriate scope of variation in electronic (*trans* influence) and steric properties, these same ligands have been selected to complete the coordination sphere of Pd for comparative purposes. As depicted in Fig. 7, the dinuclear derivative **4e** reacts with pyridine (1:2 molar ratio) to yield the mononuclear derivative **5e**. Treatment of **4e** with LiCl (1:4 molar ratio) in MeOH produces the dinuclear chloride-bridging complex **6e**, which subsequently reacts with pyridine (1:2 molar ratio) in

$\text{CH}_2\text{Cl}_2$  to form mononuclear **7e**, with  $\text{AgClO}_4$  and pyridine (1:2:4 molar ratio) in  $\text{CH}_2\text{Cl}_2$ /acetone, to give the cationic bis-pyridine derivative **8e**, or with  $\text{Ti}(\text{acac})_3$  (1:2 molar ratio; acac = acetylacetonate) affording the neutral mononuclear complex **9e**. The characterization of all prepared complexes **5e–9e** is straightforward from their NMR and HRMS data, as shown in the ESI.† Complexes **5e**, **7e**, and **8e** exhibit dynamic behaviour in solution at room temperature, which is stopped upon lowering the temperature to 233 K. At this temperature, the molecules behave as static during the response time of the NMR. In general, the ligand arrangement follows the antisymbiotic effect shown by the soft Pd(II) center,<sup>57</sup> with the pyridine ligand *cis* to the Pd–C bond.<sup>58</sup>

As for the photophysical properties, palladacycles typically exhibit low-efficient emission, only detected at low temperatures, due to the thermal deactivation of the excited state to the low-lying metal-centered orbitals (MC; d–d\*) which experience non-radiative decay.<sup>25,59–61</sup> On the grounds of previous studies performed in oxazolones where fluorescence amplification was achieved by using Pd as an intramolecular lock,<sup>24</sup> we have investigated here the emissive properties of orthopalladated analogs with imidazolones **5e–9e**, which result to be fluorescent. Photophysical data of **2a–2e** and **5e–9e** including their absorption, emission maxima, and emission quantum yields ( $\Phi_{\text{PL}}$ ) in solution at room temperature are summarized in Table 1.

Most imidazolones **2** have the strongest absorption maxima in the UV region of the spectrum, due to a  $\pi$ – $\pi^*$  charge transfer from the 4-arylidene ring to the imidazolone heterocyclic moiety.<sup>12,62</sup> In previously reported orthopalladated imidazolones the maxima were red-shifted to the blue-green region with respect to free imidazolones.<sup>21</sup> We have observed here a similar trend for **5e–9e** when compared with **2e**. Therefore, the absorption spectra of **5e–9e** in  $\text{CH}_2\text{Cl}_2$  solution resemble each other and display an intense absorption in the blue region (445–454 nm). The bands with the highest absorbance correspond to intraligand charge transfer transition (ILCT) of the imidazolone fragment, as explained previously. The quantum yield  $\Phi_{\text{PL}}$  values range from 2 to 15% for complexes **5e**, **7e**, **8e** and **9e**, which have methoxide substituents at the two *meta* positions of the arylidene ring with respect to the Pd–C bond

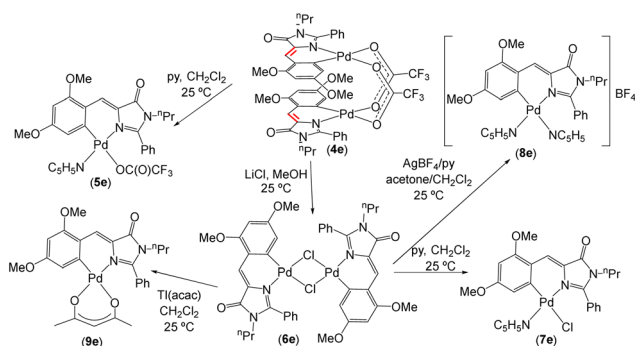


Fig. 7 Synthesis of the fluorescent imidazolone derivatives **5e–9e**.

Table 1 Absorption and emission maxima, and fluorescence quantum yields for free imidazolones **2a–2e** and complexes **5e–9e**

| Compound <sup>a</sup> | $\lambda_{\text{abs}}$ (nm) | $\lambda_{\text{em}}$ (nm) | $\Phi_{\text{PL}}$ (%) |
|-----------------------|-----------------------------|----------------------------|------------------------|
| <b>2a</b>             | 378                         | 458                        | <1                     |
| <b>2b</b>             | 380                         | 463                        | <1                     |
| <b>2c</b>             | 380                         | 468                        | 5                      |
| <b>2d</b>             | 400                         | 462                        | <1                     |
| <b>2e</b>             | 406                         | 468                        | <1                     |
| <b>5e</b>             | 445                         | 514                        | 5                      |
| <b>7e</b>             | 454                         | 529                        | 7                      |
| <b>8e</b>             | 451                         | 522                        | 15                     |
| <b>9e</b>             | 452                         | 518                        | 2                      |

<sup>a</sup> Measured in  $10^{-5}$  M  $\text{CH}_2\text{Cl}_2$  solution at room temperature.





(Fig. 7), while it is close to zero for the starting imidazolone **2e** (<1%). Therefore, a true recovery of the fluorescence has been achieved. The change of ligands attached to the Pd atom has a significant influence on  $\Phi_{\text{PL}}$ , similar to that determined in oxazolones.<sup>24</sup> In the present case, the complex **9e** with the acetylacetonate group as the ancillary ligand exhibits the lowest  $\Phi_{\text{PL}}$  (2%), followed by complex **5e** with TFA and pyridine (5%), quite similar to that of complex **7e** with TFA and Cl (7%). As in related complexes with palladated oxazolones,<sup>24</sup> the presence of two pyridines in *cis* arrangement as ancillary ligands for the Pd (complex **8e**) provides the highest performance, reflected in a quantum yield value as high as 15%.

Aiming to provide a detailed explanation about the role of Pd and the ancillary ligands in the recovery and amplification of the fluorescence observed on going from **2e** to **5e–9e**, the photophysical properties of the most important examples were modeled using computational methods (DFT and TDDFT). The absorption properties of compounds **5e**, **7e**, and **8e** were calculated using the M06-2X functional, as described in the Computational Details section (see Experimental). The obtained results for **5e**, **7e** and **8e** are shown in Fig. 8.

The three studied compounds exhibit similar characteristic behavior, for which the absorption properties correspond to the excitation from the ground state ( $S_0$ ) to the first excited singlet state ( $S_1$ ), predominantly involving a HOMO to LUMO transition. In our previous studies in oxazolones,<sup>24</sup> we observed a correlation between the participation of Pd and the orthopalladated carbon in the HOMO orbital with the obtained quantum yield  $\Phi_{\text{PL}}$ . We aimed to verify if this pattern extended to complexes with imidazolones. Fig. 8 depicts the HOMO orbitals of the products, showing the percentage of the participation of the Pd orbitals in black parentheses and that for the metallated carbon in red. The comparison of complexes **5e** and **7e** reveals that the differences in Pd and C participation are negligible, as are the differences in  $\Phi_{\text{PL}}$ . Moreover, the lowest participation of both percentages in Pd and palladated C are observed for **8e**, which correlates straightforward with the highest value of  $\Phi_{\text{PL}}$  measured for this compound. Therefore, it seems that once again, data suggest a correlation between the participation of the Pd–C fragment in the HOMO

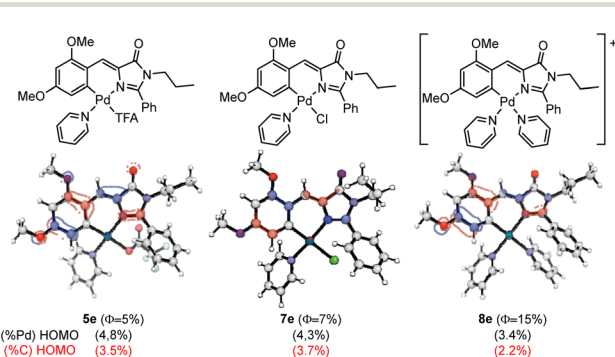


Fig. 8 Schematic structures of **5e**, **7e** and **8e** (top), HOMO orbitals (middle) with the corresponding Pd contributions (black) and orthopalladated C atom (red) to HOMO in parentheses (bottom).

orbital and the value of the quantum yield measured experimentally, this behavior being parallel to that found for complexes with oxazolones.

## [2 + 2]-Photocycloaddition of dimers **4** and **11**: synthesis of 1,3-diaminotruxillic derivatives **10** and **12**

Given the interesting photochemical reactivity previously observed in 4-arylidene-5(4*H*)-oxazolones to yield bis-amino acids of diaminotruxillic and diaminotruxinic types,<sup>16–19</sup> and considering the continued interest in this type of cyclobutanes due to their exceptional pharmacological properties in the treatment of type 2 diabetes,<sup>63–66</sup> we have attempted to extend this reactivity to free imidazolone substrates (**2**). However, irradiation of (**2a**)–(**2e**) in a variety of solvents with blue light (465 nm) does not lead to any detectable transformation (Fig. 9a), while irradiation using near-ultraviolet light (365 nm) gradually decomposes the sample, and it is not possible to distinguish cyclobutane-type products in either case. Since the incorporation of Pd into the imidazolone skeleton produces a clear red-shift in absorption, with absorption maxima in the blue region for compounds **4**, and dinuclear orthopalladated oxazolones have also shown clear photochemical reactivity in the visible region,<sup>31–34</sup> the irradiation of dinuclear complexes (**4a**)–(**4e**) with blue light has been attempted. The result of the reaction is shown in Fig. 9b. In the case of thiazolones (**3**), there is a single report of [2 + 2] photocycloaddition in non-complexed substrates,<sup>42</sup> but there are no precedents in the case of orthopalladated thiazolones. For those reasons, irradiation of dimers (**11**) has been carried out as well.

The irradiation with blue light (465 nm, Kessil lamps, 40 W) of  $\text{CH}_2\text{Cl}_2$  solutions of dinuclear complexes (**4a**)–(**4e**), containing orthometallated imidazolones and bridging trifluoroacetate ligands, proceeds through [2 + 2] photocycloaddition of the two C=C double bonds of the imidazolones and the selective formation of the corresponding cyclobutanes (**10a**)–(**10c**) as the sole diastereoisomers (epsilon  $\epsilon$ -isomer). The selectivity of the reaction is dictated by molecular constraints, with the dimer with an open-book structure acting as a molecular template. The optimized reaction time is 24 h, notably shorter than that observed for analogous oxazolones, usually in the

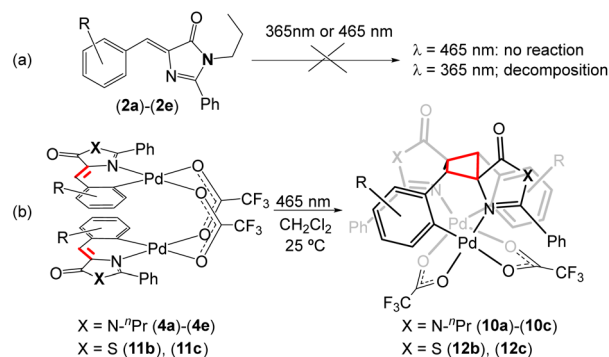


Fig. 9 Blue light-promoted [2 + 2]-photocycloaddition of orthopalladated imidazolones (**4**) and thiazolones (**11**) to give cyclobutanes (**10**) and (**12**), respectively.



range of 48 h–72 h.<sup>33,34</sup> The reaction proceeds with complete conversion to a single product in the presence of electron-withdrawing substituents (F **4a**; Cl, **4b**, **4c**) but does not occur (zero conversion) in the case of methoxide substituents (**4d**, **4e**). Despite this limitation, given the complete absence of reactivity in the irradiation of free imidazolones (**2**) with visible light, the incorporation of Pd into imidazolone provides an advantageous method for accessing new diaminotruaxillic-type bis-amino acids. Similarly, the irradiation of CH<sub>2</sub>Cl<sub>2</sub> solutions of dinuclear complexes with orthometallated thiazolones (**11b**) and (**11c**) with blue light (465 nm) results in the [2 + 2] photocycloaddition of the exocyclic C=CH fragments and the formation of cyclobutanes (**12b**) and (**12c**), as shown in Fig. 9b. In the case of complex (**11a**) a mixture of compounds was obtained – the expected cyclobutane (**12a**) and another unidentified species which could not be separated from (**12a**). Due to this fact (**12a**) was not fully characterized. In all other described cases, the reaction products (**10**) and (**12**) are isolated in pure form and with very good yields (even quantitative) by simple evaporation of the solutions and collection of the formed solid. Likewise, all isolated products (**10**) and (**12**) are photostable towards further irradiation in the visible region, and also are stable against thermal [2 + 2] retro-cycloaddition.

The observation of a single set of signals in the NMR spectra of (**10**) and (**12**) shows that these derivatives are obtained as unique isomers. The NMR signal set indicates that the two orthopalladated ligands, as well as the two carboxylate groups, behave as chemically equivalent, suggesting that (**10**) and (**12**) have C<sub>2</sub> symmetry. This symmetry, along with the *transoid* relative configuration of the orthometallated ligands determined in their respective precursors (**4**) and (**11**) (Fig. 4–6), implies that the photocycloaddition proceeds with 1,3-head-to-tail C–C coupling, and the regio- and stereoselectively obtained isomer is the epsilon  $\epsilon$ -isomer. Key data supporting the formation of the cyclobutane skeleton in (**10**) and (**12**) include the disappearance of signals assigned to the vinyl fragment C=CH in their <sup>1</sup>H and <sup>13</sup>C NMR spectra, as well as the appearance of new signals in the 4.9–5.1 ppm (CH) region of the <sup>1</sup>H NMR spectra and in the 59 ppm (CH) and 71 ppm (C) regions (62 ppm and 84 ppm, respectively in the case of thiazolones) in the <sup>13</sup>C NMR spectra. These changes are in perfect agreement with the Csp<sup>2</sup>–Csp<sup>3</sup> hybridization change occurring in the photocycloaddition.

## Experimental

Full details about materials and measurements, crystallography and computational methods are given in the ESI.†

### Synthetic procedures

**General synthesis of new imidazolones 2a–2e.** All (*Z*)-4-arylidene-1-propyl-2-phenyl-5(4*H*)-imidazolones **2a–e** have been prepared from the corresponding oxazolones **1a–1e** following methods reported in the literature.<sup>21,47</sup> This is exemplified here with the synthesis of **2a**. The synthesis and characterization of **2b–2e** are given in the ESI.†

**(*Z*)-4-(4-Fluorobenzylidene)-1-(propyl)-2-phenyl-imidazol-5(4*H*)-one (2a).** (*Z*)-4-(4-Fluorobenzylidene)-2-phenyl-5(4*H*)-oxazolone **1a** (2.11 g, 7.9 mmol) was suspended in pyridine (6 mL) and treated with *n*-propylamine (0.65 mL, 7.9 mmol) with strong stirring for 20 minutes at 25 °C. A white precipitate formed. Then, bis(trimethylsilyl)acetamide (BSA, 3.85 mL, 15.8 mmol) was added, and the resulting mixture was stirred while heating at 110 °C for 15 h. The resulting dark suspension was allowed to cool to room temperature and treated with 15 mL of ethyl acetate. This mixture was washed with brine (5 × 15 mL), dried over anhydrous MgSO<sub>4</sub>, filtered to remove any remaining solid, and evaporated to dryness. The resulting brown oil was treated with 10 mL of *n*-hexane giving **2a** as an orange solid, which was filtered, washed with additional *n*-hexane (10 mL) and dried by suction. Orange solid. Obtained: 1.67 g, yield: 68%. <sup>1</sup>H NMR (CDCl<sub>3</sub>, 400.16 MHz):  $\delta$  8.24 (dd, 2H, <sup>3</sup>J<sub>HH</sub> = 8.5 Hz, <sup>3</sup>J<sub>HH</sub> = 5.75 Hz, H<sub>2</sub>/H<sub>6</sub>, C<sub>6</sub>H<sub>4</sub>), 7.78 (dd, 2H, <sup>3</sup>J<sub>HH</sub> = 8.1 Hz, <sup>4</sup>J<sub>HH</sub> = 1.6 Hz, H<sub>o</sub>, C<sub>6</sub>H<sub>5</sub>), 7.58–7.52 (m, 3H, H<sub>m</sub>, H<sub>p</sub>, C<sub>6</sub>H<sub>5</sub>), 7.20 (s, 1H, =CH), 7.10 (dd, 2H, <sup>3</sup>J<sub>HH</sub> = 8.7 Hz, H<sub>3</sub>/H<sub>5</sub>, C<sub>6</sub>H<sub>4</sub>), 3.74 (t, 2H, <sup>3</sup>J<sub>HH</sub> = 7.5 Hz, NCH<sub>2</sub>), 1.59 (sext, 2H, <sup>3</sup>J<sub>HH</sub> = 7.4 Hz, CH<sub>2</sub>), 0.84 (t, 3H, <sup>3</sup>J<sub>HH</sub> = 7.4 Hz, CH<sub>3</sub>). <sup>13</sup>C{<sup>1</sup>H} NMR (125.77 MHz, CDCl<sub>3</sub>)  $\delta$  171.72 (C=O), 163.89 (d, <sup>1</sup>J<sub>CF</sub> = 252.86 Hz, CF, C<sub>6</sub>H<sub>4</sub>), 163.00 (CN), 138.67 (C=), 134.69 (d, <sup>3</sup>J<sub>CF</sub> = 8.34 Hz, C<sub>2</sub>/C<sub>6</sub>, C<sub>6</sub>H<sub>4</sub>), 131.46 (C<sub>p</sub>, C<sub>6</sub>H<sub>5</sub>), 130.78 (d, <sup>4</sup>J<sub>CF</sub> = 3.13 Hz, C<sub>i</sub>, C<sub>6</sub>H<sub>4</sub>), 129.88 (C<sub>i</sub>, C<sub>6</sub>H<sub>5</sub>), 128.99 (C<sub>m</sub>, C<sub>6</sub>H<sub>5</sub>), 128.40 (C<sub>o</sub>, C<sub>6</sub>H<sub>5</sub>), 127.41 (d, <sup>5</sup>J<sub>CF</sub> = 1.34 Hz, =CH), 115.99 (d, <sup>2</sup>J<sub>CF</sub> = 21.88 Hz, C<sub>3</sub>/C<sub>5</sub>, C<sub>6</sub>H<sub>4</sub>), 43.46 (NCH<sub>2</sub>), 22.68 (CH<sub>2</sub>), 11.11 (CH<sub>3</sub>). <sup>19</sup>F NMR (376.49 MHz, CDCl<sub>3</sub>)  $\delta$  –108.42 (s, CF). HRMS (ESI<sup>+</sup>) *m/z* calcd for C<sub>19</sub>H<sub>18</sub>FN<sub>2</sub>O [M + H]<sup>+</sup>: 309.1403, found: 309.1405.

**General synthesis of thiazolones 3a–3e.** All (*Z*)-4-arylidene-2-phenyl-5(4*H*)-thiazolones **3a–3e** have been prepared from the corresponding oxazolones **1a–1e** following reported methods.<sup>42,49,50</sup>

**General method for the orthopalladation of imidazolones 2 and thiazolones 3.** All orthopalladation reactions have been performed in the same way, following methods published in the literature.<sup>20</sup> In some instances, these procedures have been adapted to a particular case, which will be detailed. The general method is exemplified here for the synthesis of **4a**. The synthesis and characterization of **4b–4e** and **11a–11d** are given in the ESI.†

**Synthesis of orthopalladated 4a.** To a suspension of Pd(OAc)<sub>2</sub> (0.218 g, 0.97 mmol) in CF<sub>3</sub>CO<sub>2</sub>H (4 mL), imidazolone **2a** (0.300 g, 0.97 mmol) was added. The resulting suspension was heated at 75 °C for 2 h. During this time, the color of the suspension changes from orange–brown to yellow. After the reaction time, the suspension was cooled to room temperature and treated with water (25 mL). The precipitated solid of **4a** was filtered, washed with additional water (3 × 15 mL) until the smell of CF<sub>3</sub>CO<sub>2</sub>H was not detected and dried by suction. Yellow solid. Obtained: 0.412 g, yield 81%. <sup>1</sup>H NMR (CDCl<sub>3</sub>, 500.13 MHz):  $\delta$  7.52 (t, 2H, <sup>3</sup>J<sub>HH</sub> = 7.5 Hz, H<sub>p</sub>, C<sub>6</sub>H<sub>5</sub>), 7.50 (s, 2H, =CH), 7.44 (t, 4H, <sup>3</sup>J<sub>HH</sub> = 8.0 Hz, H<sub>m</sub>, C<sub>6</sub>H<sub>5</sub>), 7.26 (dd, <sup>3</sup>J<sub>HH</sub> = 8.4 Hz, <sup>4</sup>J<sub>HF</sub> = 6.1 Hz, 2H, H<sub>2</sub>, C<sub>6</sub>H<sub>3</sub>), 7.17 (d, broad, 4H, <sup>3</sup>J<sub>HH</sub> = 8.4 Hz, H<sub>o</sub>, C<sub>6</sub>H<sub>5</sub>), 6.91 (ddd, 2H, <sup>3</sup>J<sub>HH</sub> = 8.0 Hz, <sup>3</sup>J<sub>HF</sub> = 8.0 Hz, <sup>3</sup>J<sub>HH</sub> = 2.4 Hz, H<sub>3</sub>, C<sub>6</sub>H<sub>3</sub>), 6.86 (dd, 2H, <sup>3</sup>J<sub>HF</sub> = 10.5 Hz,



$^4J_{\text{HH}} = 2.5$  Hz,  $\text{H}_5$ ,  $\text{C}_6\text{H}_3$ ), 3.46 (ddd, 2H,  $^2J_{\text{HH}} = 14.0$  Hz,  $^3J_{\text{HH}} = 9.2$  Hz,  $^3J_{\text{HH}} = 5.7$  Hz,  $\text{NCH}_2$ ), 3.19 (ddd, 2H,  $^2J_{\text{HH}} = 14.4$  Hz,  $^3J_{\text{HH}} = 9.4$  Hz,  $^3J_{\text{HH}} = 6.1$  Hz,  $\text{NCH}_2$ ), 1.28, 1.14 (m, 4H,  $\text{CH}_2$ ), 0.63 (t, 6H,  $^3J_{\text{HH}} = 7.4$  Hz,  $\text{CH}_3$ ).  $^{13}\text{C}\{^1\text{H}\}$  NMR (125.77 MHz,  $\text{CDCl}_3$ )  $\delta$  165.24 (C=O), 164.87 (CN), 164.52 (q,  $^2J_{\text{CF}} = 38.5$  Hz,  $\text{CF}_3\text{CO}_2$ ), 161.44 (d,  $^1J_{\text{CF}} = 259.7$  Hz, CF), 138.75 (C=), 134.90 (=CH), 134.15 (d,  $^2J_{\text{CF}} = 8.7$  Hz  $\text{C}_2$ ,  $\text{C}_6\text{H}_3$ ), 132.67 ( $\text{C}_p$ ,  $\text{C}_6\text{H}_5$ ), 129.14 ( $\text{C}_m$ ,  $\text{C}_6\text{H}_5$ ), 128.69 ( $\text{C}_o$ ,  $\text{C}_6\text{H}_5$ ), 128.33 (d,  $^3J_{\text{CF}} = 2.6$  Hz,  $\text{C}_6$ ,  $\text{C}_6\text{H}_3$ ), 126.51 (d,  $^4J_{\text{CF}} = 1.6$  Hz  $\text{C}_i$ ,  $\text{C}_6\text{H}_3$ ), 126.36 ( $\text{C}_i$ ,  $\text{C}_6\text{H}_5$ ), 121.00 (d,  $^3J_{\text{CF}} = 22.0$  Hz,  $\text{C}_5$ ,  $\text{C}_6\text{H}_3$ ), 114.6 (q,  $^1J_{\text{CF}} = 287.9$  Hz,  $\text{CF}_3$ ), 112.76 (d,  $^2J_{\text{CF}} = 22.0$  Hz,  $\text{C}_3$ ,  $\text{C}_6\text{H}_3$ ), 43.67 ( $\text{NCH}_2$ ), 21.82 ( $\text{CH}_2$ ), 10.94 ( $\text{CH}_3$ ).  $^{19}\text{F}$  NMR (470.55 MHz,  $\text{CDCl}_3$ )  $\delta$  -74.40 (s,  $\text{CF}_3$ ), -106.10 (dt,  $^3J_{\text{FH}} = 10.6$  Hz,  $^4J_{\text{FH}} = 7.0$  Hz, CF). HRMS ( $\text{ESI}^+$ )  $m/z$  calcd for  $\text{C}_{42}\text{H}_{32}\text{F}_8\text{N}_4\text{NaO}_6\text{Pd}_2$  [ $\text{M} + \text{Na}$ ] $^+$ : 1075.0161, found: 1075.0107.

### Reactivity of orthopalladated 4e, synthesis of luminescent mononuclear and dinuclear derivatives 5e–9e

**Synthesis of orthopalladated 5e.** A suspension of orthopalladated 4a (0.100 g, 0.088 mmol) in  $\text{CH}_2\text{Cl}_2$  (10 mL) was reacted with pyridine (14.2  $\mu\text{L}$ , 0.176 mmol) for 30 minutes at room temperature. During this time the initial suspension gradually dissolved. After the reaction time any remaining solid was filtered, and the resulting solution was evaporated to dryness, giving 5e as a deep red solid. Obtained: 0.086 g, yield: 93%. Compound 5e was characterized in solution at 233 K as a mixture of *cis*- and *trans*-isomers (1 : 1 molar ratio). Signals at room temperature are broad and preclude a correct characterization.  $^1\text{H}$  NMR ( $\text{CD}_2\text{Cl}_2$ , 400.13 MHz, 233K):  $\delta$  9.73 (d,  $^3J_{\text{HH}} = 5.6$  Hz, 1H,  $\text{H}_o$ , py), 8.52 (d,  $^3J_{\text{HH}} = 5.2$  Hz, 1H,  $\text{H}_o$ , py), 8.10 (s, 1H,  $\text{H}_v$ ), 8.00 (m, 2H,  $\text{H}_o$ , Ph), 7.85–7.72 (m, 4H,  $\text{H}_o + \text{H}_p$ , py), 7.67–7.58 (m, 3H,  $\text{H}_v + \text{H}_o$  Ph), 7.55 (m, 2H,  $\text{H}_p$ , Ph), 7.46–7.37 (m, 2H,  $\text{H}_m$ , py), 7.36–7.22 (m, 4H,  $\text{H}_m$  py +  $\text{H}_m$  Ph), 7.06–6.95 (m, 2H,  $\text{H}_m$ , Ph), 6.18 (s, br, 1H,  $\text{C}_6\text{H}_2$ ), 6.16 (s, br, 1H,  $\text{C}_6\text{H}_2$ ), 5.66 (s, br, 1H,  $\text{C}_6\text{H}_2$ ), 5.64 (s, br, 1H,  $\text{C}_6\text{H}_2$ ), 3.89, 3.87 (2s, 6H, 2 OMe), 3.66 (t,  $^3J_{\text{HH}} = 7.6$  Hz, 2H,  $\text{NCH}_2$ , Pr), 3.59 (t,  $^3J_{\text{HH}} = 7.6$  Hz, 2H,  $\text{NCH}_2$ , Pr), 3.51, 3.50 (2s, 6H, 2 OMe), 1.45 (q,  $^3J_{\text{HH}} = 7.7$  Hz, 2H,  $\text{CH}_2$ , Pr), 1.33 (q,  $^3J_{\text{HH}} = 7.6$  Hz, 2H,  $\text{CH}_2$ , Pr), 0.73 (t,  $^3J_{\text{HH}} = 7.6$  Hz, 3H,  $\text{CH}_3$ , Pr), 0.65 (t,  $^3J_{\text{HH}} = 7.7$  Hz, 3H,  $\text{CH}_3$ , Pr).  $^{13}\text{C}\{^1\text{H}\}$  NMR ( $\text{CD}_2\text{Cl}_2$ , 125.77 MHz, 233K)  $\delta$  165.51, 165.12 (C=N), 162.47, 161.94 (C=O), 161.11, 160.82, 160.17, 159.29 (C–O,  $\text{C}_3 + \text{C}_5$ ,  $\text{C}_6\text{H}_2$ ), 153.13, 152.64 ( $\text{C}_o$ , py), 148.76, 148.25 (=C), 139.47, 137.96 ( $\text{C}_p$ , py), 138.62, 132.10, 131.67, 131.41 (=CH +  $\text{C}_p$  Ph), 129.60, 129.24, 128.93, 128.48 ( $\text{C}_o + \text{C}_m$ , Ph), 127.52, 127.46, 127.19, 126.45 ( $\text{C}_1 + \text{C}_2$ ,  $\text{C}_6\text{H}_2$ ), 126.69 (2C,  $\text{C}_m$ , py), 115.77, 115.56 ( $\text{C}_i$ , Ph), 114.75, 114.24 ( $\text{C}_6$ ,  $\text{C}_6\text{H}_2$ ), 94.48, 94.34 ( $\text{C}_4$ ,  $\text{C}_6\text{H}_2$ ), 55.92, 55.76, 55.48, 55.34 (OMe), 43.55 (2 $\text{NCH}_2$ , Pr), 22.68, 22.12 ( $\text{CH}_2$ , Pr), 10.95, 10.85 ( $\text{CH}_3$ , Pr). Peaks assigned to  $\text{CF}_3\text{CO}_2$  in the APT spectrum could not be assigned, despite the use of low temperatures and long accumulation trials.  $^{19}\text{F}$  NMR (470.55 MHz,  $\text{CDCl}_3$ )  $\delta$  -74.51, -74.94. HRMS ( $\text{ESI}^+$ )  $m/z$  calcd for  $\text{C}_{28}\text{H}_{25}\text{F}_3\text{N}_3\text{O}_5\text{Pd}$  [ $\text{M}$ ] $^+$ : 646.0781, found: 646.0801.

**Synthesis of orthopalladated 6e.** To a suspension of 4e (1.11 g, 0.976 mmol) in MeOH (15 mL) at room temperature, LiCl (0.340 g, 1.952 mmol) was added. The resulting mixture

was stirred for 30 min, then filtered. The orange solid collected of 6e was washed with additional MeOH (5 mL) and  $\text{Et}_2\text{O}$  (3  $\times$  5 mL) and dried by suction. Orange solid. Obtained: 0.599 g, yield 63%. This resulting compound too was insoluble in the usual NMR solvents; therefore it was characterized by IR and HRMS. HRMS ( $\text{ESI}^+$ )  $m/z$  calcd for  $\text{C}_{42}\text{H}_{42}\text{Cl}_2\text{N}_4\text{O}_6\text{Pd}_2$  [ $\text{M}$ ] $^+$ : 980.0551, found: 980.0567. IR ( $\nu$ ,  $\text{cm}^{-1}$ ): 1710 ( $\nu\text{C=O}$ , heterocycle), 1570 ( $\nu\text{C=N}$ , heterocycle), 334 ( $\nu\text{Pd-Cl}$ ).

**Synthesis of orthopalladated 7e.** To a suspension of 6e (0.100 g, 0.102 mmol) in  $\text{CH}_2\text{Cl}_2$  (10 mL) pyridine (16.4  $\mu\text{L}$ , 0.204 mmol) was added, and the resulting mixture was stirred for 30 minutes at room temperature. Then, any remaining solid was filtered, and the resulting solution was evaporated to dryness, giving 7e as a deep red solid. Obtained: 0.091 g, yield: 98%. Compound 7e was characterized in solution at 233 K as a mixture of the *C-trans-Cl* and *C-trans-py* isomers (4 : 1 molar ratio). Only the major isomer could be fully characterized due to extensive overlapping of peaks. Signals at room temperature are broad and preclude a correct characterization.  $^1\text{H}$  NMR ( $\text{CD}_2\text{Cl}_2$ , 400.13 MHz, 233K):  $\delta$  8.78 (dd,  $^3J_{\text{HH}} = 4.8$  Hz,  $^4J_{\text{HH}} = 1.6$  Hz, 2H,  $\text{H}_o$ , py), 8.18 (s, 1H,  $\text{H}_{\text{vinyl}}$ ), 8.00 (m, 2H,  $\text{H}_o$ , Ph), 7.85 (t,  $^3J_{\text{HH}} = 7.5$  Hz, 1H,  $\text{H}_p$ , py), 7.51 (tt,  $^3J_{\text{HH}} = 8$  Hz,  $^4J_{\text{HH}} = 1.2$  Hz, 1H,  $\text{H}_p$ , Ph), 7.41 (m, 2H,  $\text{H}_m$ , py), 7.20 (d,  $^4J_{\text{HH}} = 2.4$  Hz, 1H,  $\text{H}_6$ ,  $\text{C}_6\text{H}_2$ ), 6.98 (t,  $^3J_{\text{HH}} = 8$  Hz, 2H,  $\text{H}_m$ , Ph), 6.18 (d,  $^4J_{\text{HH}} = 2.4$  Hz, 1H,  $\text{H}_4$ ,  $\text{C}_6\text{H}_2$ ), 3.89, 3.87 (2s, 2 OMe), 3.60 (t,  $^3J_{\text{HH}} = 7.2$  Hz, 2H,  $\text{NCH}_2$ , Pr), 1.33 (q,  $^3J_{\text{HH}} = 7.5$  Hz,  $\text{CH}_2$ , Pr), 0.66 (t,  $^3J_{\text{HH}} = 7.2$  Hz,  $\text{CH}_3$ , Pr).  $^{13}\text{C}\{^1\text{H}\}$  NMR ( $\text{CD}_2\text{Cl}_2$ , 100.4 MHz, 233 K)  $\delta$  165.39 (C=N), 161.00 (C=O), 160.19, 159.90 (C–O,  $\text{C}_3 + \text{C}_5$ ,  $\text{C}_6\text{H}_2$ ), 153.14 ( $\text{C}_o$ , py), 150.25 (=C), 139.10 ( $\text{C}_p$ , py), 132.03, 131.38 (=CH +  $\text{C}_p$  Ph), 128.98 ( $\text{C}_o$ , Ph), 127.93, 126.51 ( $\text{C}_1 + \text{C}_2$ ,  $\text{C}_6\text{H}_2$ ), 125.34 ( $\text{C}_m$ , Ph), 124.58 ( $\text{C}_m$ , py), 117.30 ( $\text{C}_6$ ,  $\text{C}_6\text{H}_2$ ), 115.95 ( $\text{C}_i$ , Ph), 94.78 ( $\text{C}_4$ ,  $\text{C}_6\text{H}_2$ ), 55.76 (2C overlapped, OMe), 43.45 ( $\text{NCH}_2$ , Pr), 22.25 ( $\text{CH}_2$ , Pr), 10.89 ( $\text{CH}_3$ , Pr). HRMS ( $\text{ESI}^+$ )  $m/z$  calcd for  $\text{C}_{26}\text{H}_{26}\text{N}_3\text{O}_3\text{Pd}$  [ $\text{M} - \text{Cl} + \text{H}$ ] $^+$ : 534.1004, found: 534.1004.

**Synthesis of orthopalladated 8e.** To a suspension of 6e (0.100 g, 0.102 mmol) in  $\text{CH}_2\text{Cl}_2$  (5 mL) and acetone (1 mL),  $\text{AgBF}_4$  (0.040 g, 0.204 mmol) was added. The resulting mixture was stirred at room temperature for 30 min with exclusion of light, and then filtered over Celite to remove the precipitated AgCl. The clear solution thus obtained was treated with pyridine (32.8  $\mu\text{L}$ , 0.408 mmol) and further stirred at room temperature for additional 30 minutes. Then, the resulting solution was evaporated to dryness, affording an orange solid characterized as 8e. Obtained: 0.138 g, yield 97%.  $^1\text{H}$  NMR ( $\text{CDCl}_3$ , 400.13 MHz, 233K):  $\delta$  8.79 (d,  $^3J_{\text{HH}} = 5.4$  Hz, 2H,  $\text{H}_o$ , py), 8.29 (s, 1H, =CH), 8.12 (d,  $^3J_{\text{HH}} = 5.0$  Hz, 2H,  $\text{H}_o$ , py), 7.73 (t,  $^3J_{\text{HH}} = 7.7$  Hz, 2H,  $\text{H}_p$ , py), 7.52–7.35 (m, 7H,  $\text{H}_m$  py,  $\text{H}_o + \text{H}_m + \text{H}_p$ , Ph), 7.04 (dd,  $^3J_{\text{HH}} = 7.8$  Hz,  $^3J_{\text{HH}} = 6.0$  Hz, 2H,  $\text{H}_m$  py), 6.15 (d,  $^3J_{\text{HH}} = 2.1$  Hz, 1H,  $\text{H}_6$ ,  $\text{C}_6\text{H}_2$ ), 5.78 (d,  $^3J_{\text{HH}} = 2.2$  Hz, 1H,  $\text{H}_4$ ,  $\text{C}_6\text{H}_2$ ), 3.88 (s, 3H, OMe), 3.63 (t,  $^3J_{\text{HH}} = 7.6$  Hz, 2H,  $\text{NCH}_2$ ), 3.51 (s, 3H, OMe), 1.39 (q,  $^3J_{\text{HH}} = 7.7$  Hz, 2H,  $\text{CH}_2$ ), 0.69 (t,  $^3J_{\text{HH}} = 7.4$  Hz, 3H,  $\text{CH}_3$ ). This product was too insoluble at low temperature to obtain reliable  $^{13}\text{C}$  NMR data.  $^{19}\text{F}$  NMR (470.55 MHz,  $\text{CDCl}_3$ )  $\delta$  -152.26 (s,  $\text{BF}_4$ ). HRMS ( $\text{ESI}^+$ )  $m/z$  calcd for  $\text{C}_{21}\text{H}_{21}\text{N}_2\text{O}_3\text{Pd}$  [ $\text{M} - 2\text{py}$ ] $^+$ : 455.0587, found: 455.0606.





**Synthesis of orthopalladated 9e.**  $\text{Ti}(\text{acac})$  (0.062 g, 0.204 mmol) was added to a stirring suspension of the orthopalladate of (Z)-4-(2,4-dimethoxybenzylidene)-2-phenyl-1*H*-thiazol-5(4*H*)-one (Cl) (0.1 g, 0.102 mmol) in  $\text{CH}_2\text{Cl}_2$  (5 mL) at room temperature. The mixture was stirred for 30 minutes, then the precipitated  $\text{TiCl}$  was removed by filtration. The clear solution was evaporated to dryness, and the solid residue was washed with cold *n*-hexane ( $3 \times 3$  mL) and dried by suction. Orange solid. Obtained: 0.088 g, yield 82%.  $^1\text{H}$  NMR ( $\text{CDCl}_3$ , 500.13 MHz):  $\delta$  8.22 (s, 1H, =CH), 7.87 (dd, 2H,  $^3J_{\text{HH}} = 7.6$  Hz,  $^4J_{\text{HH}} = 1.9$  Hz,  $\text{H}_o$ ,  $\text{C}_6\text{H}_5$ ), 7.56–7.47 (m, 3H,  $\text{H}_p$ ,  $\text{C}_6\text{H}_5$ ,  $\text{H}_m$ ,  $\text{C}_6\text{H}_5$ ), 7.03 (d, 1H,  $^4J_{\text{HH}} = 2.3$  Hz,  $\text{H}_3$ ,  $\text{C}_6\text{H}_2$ ), 6.20 (d, 1H,  $^4J_{\text{HH}} = 2.3$  Hz,  $\text{H}_5$ ,  $\text{C}_6\text{H}_2$ ), 5.05 (s, 1H, CH, acac), 3.90 (s, 3H,  $\text{OCH}_3(\text{C}_4)$ ), 3.84 (s, 3H,  $\text{OCH}_3(\text{C}_2)$ ), 3.69 (t, 2H,  $^3J_{\text{HH}} = 7.7$  Hz,  $\text{NCH}_2$ ), 1.94 (s, 3H,  $\text{CH}_3$ , acac), 1.51 (sext,  $^3J_{\text{HH}} = 7.6$  Hz,  $^3J_{\text{HH}} = 7.7$  Hz, 2H,  $\text{CH}_2$ ), 1.18 (s, 3H,  $\text{CH}_3$ , acac), 0.76 (t, 3H,  $^3J_{\text{HH}} = 7.4$  Hz,  $\text{CH}_3$ ).  $^{13}\text{C}\{^1\text{H}\}$  NMR (125.77 MHz,  $\text{CDCl}_3$ )  $\delta$  187.30 (C=O, acac), 185.40 (C=O, acac), 166.16 (CN), 161.58, ( $\text{C}_4$ ,  $\text{C}_6\text{H}_2$ ), 160.42 (C=O), 159.99 ( $\text{C}_2$ ,  $\text{C}_6\text{H}_2$ ), 150.68 ( $\text{C}_i$ ,  $\text{C}_6\text{H}_2$ ), 131.51 (=CH), 131.17 ( $\text{C}_p$ ,  $\text{C}_6\text{H}_5$ ), 129.13 ( $\text{C}_i$ ,  $\text{C}_6\text{H}_5$ ), 128.92 ( $\text{C}_o$ ,  $\text{C}_6\text{H}_5$ ), 128.68 (C=), 126.39 ( $\text{C}_m$ ,  $\text{C}_6\text{H}_5$ ), 116.23 ( $\text{C}_6$ ,  $\text{C}_6\text{H}_2$ ), 110.77 ( $\text{C}_3$ ,  $\text{C}_6\text{H}_2$ ), 99.05 (CH, acac), 95.88 ( $\text{C}_5$ ,  $\text{C}_6\text{H}_2$ ), 55.74 ( $\text{OCH}_3(\text{C}_2)$ ), 55.50 ( $\text{OCH}_3(\text{C}_4)$ ), 43.70 ( $\text{NCH}_2$ ), 27.48 ( $\text{CH}_3$ , acac), 26.22 ( $\text{CH}_3$ , acac), 22.34 ( $\text{CH}_2$ ), 11.12 ( $\text{CH}_3$ ). HRMS ( $\text{ESI}^+$ )  $m/z$  calcd for  $\text{C}_{26}\text{H}_{29}\text{N}_2\text{O}_5\text{Pd}$   $[\text{M} + \text{H}]^+$ : 555.1111, found: 555.1112.

**General method for the [2 + 2]-photocycloaddition of orthopalladated complexes 4 and 11. Synthesis of cyclobutanes 10 and 12.** All photocycloadditions have been performed in the same way, following methods published in the literature.<sup>31–34</sup> In some instances these procedures have been adapted to a particular case, which will be detailed. The general method is exemplified here for the synthesis of **10a**. The synthesis and characterization of **10b**, **10c**, **12b** and **12c** are given in the ESI.†

**Synthesis of cyclobutane derivative 10a by [2 + 2] cycloaddition of orthopalladated 4a.** A stirred suspension of orthopalladated **4a** (0.300 g, 0.28 mmol) in 10 mL  $\text{CH}_2\text{Cl}_2$  was irradiated with blue light (465 nm; Kessil lamp, 40 W) for 24 h. During the reaction, the initial suspension gradually dissolved. After the reaction time, any remaining solid was filtered through a Celite bed. The resulting clear solution was evaporated to dryness, giving **10a** as a yellow solid. Obtained: 0.298 g, yield 99%.  $^1\text{H}$  NMR ( $\text{CDCl}_3$ , 500.13 MHz):  $\delta$  7.65 (t, 2H,  $^3J_{\text{HH}} = 7.6$  Hz,  $\text{H}_p$ ,  $\text{C}_6\text{H}_5$ ), 7.54 (t, 4H,  $^3J_{\text{HH}} = 7.9$  Hz,  $\text{H}_m$ ,  $\text{C}_6\text{H}_5$ ), 7.39 (d, 4H,  $^3J_{\text{HH}} = 7.0$  Hz,  $\text{H}_o$ ,  $\text{C}_6\text{H}_5$ ), 6.75 (dd, 2H,  $^3J_{\text{HH}} = 9.5$  Hz,  $^4J_{\text{HF}} = 2.5$  Hz,  $\text{H}_2$ ,  $\text{C}_6\text{H}_3$ ), 6.72–6.64 (m, 4H,  $\text{H}_3$ ,  $\text{C}_6\text{H}_3$ ,  $\text{H}_5$ ,  $\text{C}_6\text{H}_3$ ), 4.94 (s, 2H, CH cyclobut), 3.61 (ddd, 2H,  $^2J_{\text{HH}} = 14.7$  Hz,  $^3J_{\text{HH}} = 8.1$  Hz,  $^3J_{\text{HH}} = 6.7$  Hz,  $\text{NCH}_2$ , Pr), 3.26 (ddd, 2H,  $^2J_{\text{HH}} = 14.1$  Hz,  $^3J_{\text{HH}} = 9.0$  Hz,  $^3J_{\text{HH}} = 6.1$  Hz,  $\text{NCH}_2$ , Pr), 1.45–1.28 (m, 4H,  $\text{CH}_2$ , Pr), 0.69 (t, 6H,  $^3J_{\text{HH}} = 7.4$  Hz,  $\text{CH}_3$ , Pr).  $^{13}\text{C}\{^1\text{H}\}$  NMR (125.77 MHz,  $\text{CDCl}_3$ )  $\delta$  179.36 (C=O), 172.91 (C=N), 159.45 (d,  $^1J_{\text{CF}} = 252.9$  Hz, CF), 138.17 (d,  $^2J_{\text{CF}} = 5.9$  Hz,  $\text{C}_6$ ,  $\text{C}_6\text{H}_3$ ), 132.65 ( $\text{C}_p$ ,  $\text{C}_6\text{H}_5$ ), 129.11 ( $\text{C}_m$ ,  $\text{C}_6\text{H}_5$ ), 129.04 ( $\text{C}_5$ ,  $\text{C}_6\text{H}_3$ ), 128.41 ( $\text{C}_o$ ,  $\text{C}_6\text{H}_5$ ), 126.75 ( $\text{C}_i$ ,  $\text{C}_6\text{H}_5$ ), 125.42 (d,  $^4J_{\text{CF}} = 2.6$  Hz,  $\text{C}_i$ ,  $\text{C}_6\text{H}_3$ ), 120.40 (d,  $^3J_{\text{CF}} = 19.4$  Hz,  $\text{C}_2$ ,  $\text{C}_6\text{H}_3$ ), 112.12 (d,  $^2J_{\text{CF}} = 22.3$  Hz,  $\text{C}_3$ ,  $\text{C}_6\text{H}_3$ ), 71.04 (C cyclobut), 58.87 (CH cyclobut), 44.16 ( $\text{NCH}_2$ ), 21.76 ( $\text{CH}_2$ ), 10.89 ( $\text{CH}_3$ ).  $^{19}\text{F}$  NMR (470.55 MHz,

$\text{CDCl}_3$ )  $\delta$  -75.00 (s, 3F,  $\text{CF}_3$ ), -112.82 (tt,  $^3J_{\text{FH}} = 14.9$  Hz,  $^4J_{\text{FH}} = 9.1$  Hz, 1F, CF). HRMS ( $\text{ESI}^+$ )  $m/z$  calcd for  $\text{C}_{42}\text{H}_{32}\text{F}_8\text{N}_4\text{NaO}_6\text{Pd}_2$   $[\text{M} + \text{Na}]^+$ : 1075.0156, found: 1075.0154.

## Conclusions

The incorporation of Pd into the molecular framework of 4-arylidene-imidazolones (**2**) and thiazolones (**3**) through CH bond activation processes generates platforms with remarkable photophysical and photochemical properties. In the case of imidazolones, new systems have been developed where the fluorescence of the free ligands is amplified, achieving quantum yields of orthopalladated complexes of up to 15%. This fluorescence amplification has been explained based on TD-DFT calculations and is related to the involvement of Pd in the HOMO orbitals. However, no apparent amplification occurs in the studied cases of thiazolones. Regarding photochemistry, it has been shown that irradiation in the visible region (blue light), both for orthopalladated imidazolones and thiazolones, proceeds through the [2 + 2] photocycloaddition of exocyclic C=C double bonds, leading to the stereoselective generation of the corresponding diaminotruillics, bis-amino acids with a cyclobutane-type core. In summary, the incorporation of Pd into oxazolone, imidazolone, and thiazolone-type heterocycles causes similar changes in photophysical and photochemical properties, resulting in an amplification of the luminescence of specific substrates and a clear increase in photochemical reactivity due to the templating effect of Pd.

## Author contributions

Darius Dumitras: synthesis, chemical characterization (investigation, validation, data curation, formal analysis, methodology), writing and correction (visualization, writing – original draft, review and editing); David Dalmau: synthesis, chemical characterization, measurement of the photophysical properties (investigation, data curation, validation), computational calculations (formal analysis, methodology), writing and correction (visualization, writing – original draft, review and editing); Pilar García Orduña: measurement of the crystal structures (investigation, validation), writing and correction (visualization, writing – original draft, review and editing); Alexandra Pop: supervision, writing and correction (visualization, review and editing); Anca Silvestru: supervision, funding acquisition, resources, writing and correction (visualization, review & editing); Esteban P. Urriolabeitia: original idea (conceptualization), project administration, supervision, funding acquisition, resources, writing and correction (visualization, writing, original draft, review and editing). All authors analyzed and discussed the results and reviewed the manuscript.

## Conflicts of interest

There are no conflicts to declare.





## Acknowledgements

E. P. U. and D. Dalmau thank the Spanish Government (Grant PID2019-106394GB-I00, funded by MCIN/AEI/10.13039/501100011033) and Gobierno de Aragón-FSE (Spain, research groups Química Inorgánica y de los Compuestos Organometálicos E17\_23R) for funding. D. Dalmau thanks Gobierno de Aragón-FSE for a PhD fellowship. D. Dumitraş is grateful for the financial support from Babes–Bolyai University in the form of a Special Scholarship for Scientific Activity. We acknowledge the support of the publication fee by the CSIC Open Access Publication Support Initiative through its Unit of Information Resources for Research (URICI).

## References

- G. Ciamician, *Science*, 1912, **36**, 385–394.
- N. Hoffmann, *Photochem. Photobiol. Sci.*, 2012, **11**, 1613–1641.
- J. C. Scaiano, *Chem. Soc. Rev.*, 2023, **52**, 6330–6343.
- M. Hertzog, M. Wang, J. Mony and K. Börjesson, *Chem. Soc. Rev.*, 2019, **48**, 937–961.
- L. Marzo, S. K. Pagire, O. Reiser and B. König, *Angew. Chem., Int. Ed.*, 2018, **57**, 10034–10072.
- H. E. Bonfield, T. Knauber, F. Lévesque, E. G. Moschetta, F. Susanne and L. J. Edwards, *Nat. Commun.*, 2020, **11**, 804–807.
- Y. Markushyna and A. Savateev, *Eur. J. Org. Chem.*, 2022, e202200026.
- E. Schroeder and P. Christopher, *ACS Energy Lett.*, 2022, **7**, 880–884.
- F. Gomollón-Bel and J. García-Martínez, *Angew. Chem., Int. Ed.*, 2023, **62**, e202218975.
- F. Gomollón-Bel and J. García-Martínez, *Nat. Chem.*, 2022, **14**, 113–114.
- F. Gomollón-Bel, *Chem. Int.*, 2020, **42**, 3–9.
- S. Icli, H. Icli, S. Alp, H. Koc and A. Mckillop, *Spectrosc. Lett.*, 1994, **27**, 1115–1128.
- K. Ertekin, S. Alp, C. Karapire, B. Yenigül, E. Hendem and S. Icli, *J. Photochem. Photobiol. A*, 2000, **137**, 155–161.
- C. Rodrigues, I. Mariz, E. Maçôas, C. Afonso and J. Martinho, *Dyes Pigm.*, 2012, **95**, 713–722.
- C. Rodrigues, I. Mariz, E. Maçôas, C. Afonso and J. Martinho, *Dyes Pigm.*, 2013, **99**, 642–652.
- A. García-Montero, A. M. Rodríguez, A. Juan, A. H. Velders, A. Denisi, G. Jiménez-Osés, E. Gómez-Bengoa, C. Cativiela, M. V. Gómez and E. P. Urriolabeitia, *ACS Sustainable Chem. Eng.*, 2017, **5**, 8370–8381.
- S. Sierra, M. V. Gómez, A. I. Jiménez, A. Pop, C. Silvestru, M. L. Marín, F. Boscá, G. Sastre, E. Gómez-Bengoa and E. P. Urriolabeitia, *J. Org. Chem.*, 2022, **87**, 3529–3545.
- S. Sierra, R. López, E. Gómez-Bengoa, L. R. Falvello and E. P. Urriolabeitia, *Org. Biomol. Chem.*, 2023, **21**, 3203–3213.
- S. Sierra, D. Dalmau, J. V. Alegre, A. Pop, C. Silvestru, M. L. Marín, F. Boscá and E. P. Urriolabeitia, *Int. J. Mol. Sci.*, 2023, **24**, 7583–7602.
- D. Roiban, E. Serrano, T. Soler, G. Aullón, I. Grosu, C. Cativiela, M. Martínez and E. P. Urriolabeitia, *Inorg. Chem.*, 2011, **50**, 8132–8143.
- S. Collado, A. Pueyo, C. Baudequin, L. Bischoff, A. I. Jiménez, C. Cativiela, C. Hoarau and E. P. Urriolabeitia, *Eur. J. Org. Chem.*, 2018, 6158–6166.
- E. Laga, D. Dalmau, S. Arregui, O. Crespo, A. I. Jiménez, A. Pop, C. Silvestru and E. P. Urriolabeitia, *Molecules*, 2021, **26**, 1238–1257.
- C. García-Sanz, A. Andreu, B. de las Rivas, A. I. Jiménez, A. Pop, C. Silvestru, E. P. Urriolabeitia and J. M. Palomo, *Org. Biomol. Chem.*, 2021, **19**, 2773–2783.
- D. Dalmau, O. Crespo, J. M. Matxain and E. P. Urriolabeitia, *Inorg. Chem.*, 2023, **62**, 9792–9806.
- D. Dalmau and E. P. Urriolabeitia, *Molecules*, 2023, **28**, 2663.
- J.-S. Yang, G.-J. Huang, Y.-H. Liu and S.-M. Peng, *Chem. Commun.*, 2008, 1344–1346.
- S. R. Meech, *Chem. Soc. Rev.*, 2009, **38**, 2922–2934.
- Q. Zhang, X. Chen, G. Cui, W. H. Fang and W. Thiel, *Angew. Chem., Int. Ed.*, 2014, **53**, 8649–8653.
- A. Usman, O. F. Mohammed, E. T. J. Nibbering, J. Dong, K. M. Solntsev and L. M. Tolbert, *J. Am. Chem. Soc.*, 2005, **127**, 11214–11215.
- X. He, A. F. Bell and P. J. Tonge, *FEBS Lett.*, 2003, **549**, 35–38.
- D. Roiban, E. Serrano, T. Soler, C. Cativiela, I. Grosu and E. P. Urriolabeitia, *Chem. Commun.*, 2009, 4681–4683.
- E. Serrano, A. Juan, A. García-Montero, T. Soler, F. Jiménez-Márquez, C. Cativiela, M. V. Gomez and E. P. Urriolabeitia, *Chem. – Eur. J.*, 2016, **22**, 144–152.
- C. Carrera, A. Denisi, C. Cativiela and E. P. Urriolabeitia, *Eur. J. Inorg. Chem.*, 2019, 3481–3489.
- P. Sánchez, A. Pop, C. Silvestru, E. Laga, A. I. Jiménez, C. Cativiela and E. P. Urriolabeitia, *Beilstein J. Org. Chem.*, 2020, **16**, 1111–1123.
- O. Shimomura, *Angew. Chem., Int. Ed.*, 2009, **48**, 5590–5602.
- M. Chalfie, *Angew. Chem., Int. Ed.*, 2009, **48**, 5603–5611.
- R. Y. Tsien, *Angew. Chem., Int. Ed.*, 2009, **48**, 5612–5626.
- L. Wu and K. Burgess, *J. Am. Chem. Soc.*, 2008, **130**, 4089–4096.
- M. S. Baranov, K. A. Lukyanov, A. O. Borissova, J. Shamir, D. Kosenkov, L. V. Splichenko, L. M. Tolbert, I. V. Yampolsky and K. M. Solntsev, *J. Am. Chem. Soc.*, 2012, **134**, 6025–6032.
- M. S. Baranov, K. M. Solntsev, N. S. Baleeva, A. S. Mishin, S. A. Lukyanov, K. A. Lukyanov and I. V. Yampolsky, *Chem. – Eur. J.*, 2014, **20**, 13234–13241.
- P. Naumov, J. Kowalik, K. M. Solntsev, A. Baldrige, J.-S. Moon, C. Kranz and L. M. Tolbert, *J. Am. Chem. Soc.*, 2010, **132**, 5845–5857.
- S. Sierra, D. Dalmau, S. Higuera, D. Cortés, O. Crespo, A. I. Jimenez, A. Pop, C. Silvestru and E. P. Urriolabeitia, *J. Org. Chem.*, 2021, **86**, 12119–12140.



- 43 J. Plöchl, *Chem. Ber.*, 1883, **16**, 2815.
- 44 E. Erlenmeyer, *Justus Liebigs Ann. Chem.*, 1893, **275**, 1.
- 45 H. E. Carter, Ch. 5 of the book series *Organic Reactions, Azlactones*, 1946, **3**, 198.
- 46 Y. S. Rao and R. Filler, Oxazoles, in *The Chemistry of Heterocyclic Compounds*, ed. I. J. Turchi, John Wiley & Sons, Inc., New York, 1986, ch. 3, vol. 45, pp. 363–691.
- 47 M. Muselli, L. Colombeau, J. Hédouin, C. Hoarau and L. Bischoff, *Synlett*, 2016, 2819–2825.
- 48 R. Filler and Y. S. Rao, *J. Org. Chem.*, 1962, **27**, 3730–3731.
- 49 H. Behringer and H. W. Stein, *Chem. Ber.*, 1949, **82**, 209–212.
- 50 H. Behringer and J. B. Jepson, *Chem. Ber.*, 1952, **85**, 138.
- 51 G. Bhattacharjya, G. Savitha and G. Ramanathan, *CrystEngComm*, 2004, **6**, 233.
- 52 B. K. Rajbongshi and G. Ramanathan, *J. Chem. Sci.*, 2009, **121**, 973.
- 53 G. Bhattacharjya, G. Savitha and G. Ramanathan, *J. Mol. Struct.*, 2005, **752**, 98–103.
- 54 C. I. C. Esteves, I. da S. Fonseca, J. Rocha, A. S. Silva and S. Guieu, *Dyes Pigm.*, 2020, **177**, 108267.
- 55 W.-T. Chuang, C.-C. Hsieh, C.-H. Lai, C.-H. Lai, C.-W. Shih, K.-Y. Chen, W.-Y. Hung, Y.-H. Hsu and P.-T. Chou, *J. Org. Chem.*, 2011, **76**, 8189–8202.
- 56 A. Guy-Orpen, L. Brammer, F. H. Allen, O. Kennard, D. G. Watson and R. Taylor, *J. Chem. Soc., Dalton Trans.*, 1989, S1–S83.
- 57 R. G. Pearson, *Inorg. Chem.*, 1973, **12**, 712.
- 58 A. J. Deeming, I. P. Rothwell, M. B. Hursthouse and L. New, *J. Chem. Soc., Dalton Trans.*, 1978, 1490–1496.
- 59 M. Ghedini, I. Aiello, M. La Deda and A. Grisolia, *Chem. Commun.*, 2003, 2198.
- 60 P.-K. Chow, W.-P. To, K.-H. Low and C.-M. Che, *Chem. – Asian J.*, 2014, **9**, 534.
- 61 F. Neve, A. Crispini, C. Di Pietro and S. Campagna, *Organometallics*, 2002, **21**, 3511.
- 62 S. Icli, A. O. Doroshenko, S. Alp, N. A. Abmanova, S. I. Egorova and S. T. Astley, *Spectrosc. Lett.*, 1999, **32**, 553.
- 63 Z. Cong, Q. Zhou, Y. Li, L.-N. Chen, Z.-C. Zhang, A. Liang, Q. Liu, X. Wu, A. Dai, T. Xia, W. Wu, Y. Zhang, D. Yang and M.-W. Wang, *Proc. Natl. Acad. Sci. U. S. A.*, 2022, **119**, e2200155119.
- 64 C. de Graaf, D. Donnelly, D. Wootten, J. Lau, P. M. Sexton, L. J. Miller, J.-M. Ahn, J. Liao, M. M. Fletcher, D. Yang, A. J. H. Brown, C. Zhou, J. Deng and M.-W. Wang, *Pharmacol. Rev.*, 2016, **68**, 954–1013.
- 65 M. Méndez, H. Matter, E. Defossa, M. Kurz, S. Lebreton, Z. Li, M. Lohmann, M. Löhn, H. Mors, M. Podeschwa, N. Rackelmann, J. Riedel, P. Safar, D. S. Thorpe, M. Schäfer, D. Weitz and K. Breitschopf, *J. Med. Chem.*, 2020, **63**, 2292–2307.
- 66 B. Manandhar and J.-M. Ahn, *J. Med. Chem.*, 2015, **58**, 1020–1037.

