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Advanced 1D heterostructures based on nanotube templates and molecules†

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Recent advancements in materials science have shed light on the potential of exploring hierarchical assemblies of molecules on surfaces, driven by both fundamental and applicative challenges. This field encompasses diverse areas including molecular storage, drug delivery, catalysis, and nanoscale chemical reactions. In this context, the utilization of nanotube templates (NTs) has emerged as promising platforms for achieving advanced one-dimensional (1D) molecular assemblies. NTs offer cylindrical, crystalline structures with high aspect ratios, capable of hosting molecules both externally and internally (Mol@NT). Furthermore, NTs possess a wide array of available diameters, providing tunability for tailored assembly. This review underscores recent breakthroughs in the field of Mol@NT. The first part focuses on the diverse panorama of structural properties in Mol@NT synthesized in the last decade. The advances in understanding encapsulation, adsorption, and ordering mechanisms are detailed. In a second part, the review highlights the physical interactions and photophysics properties of Mol@NT obtained by the confinement of molecules and nanotubes in the van der Waals distance regime. The last part of the review describes potential applicative fields of these 1D heterostructures, providing specific examples in photovoltaics, luminescent materials, and bio-imaging. A conclusion gathers current challenges and perspectives of the field to foster discussion in related communities.

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1 Introduction

"Synthesized from different growth processes, SWCNTs are now produced commercially in large quantities with narrow distributions of diameters, each defining a quasi-1D surface perimeter space with a central pore sized between 0.2 nm to more than few nm."

Driven by practical challenges and potential technologies, the topic of hierarchical assemblies of molecules on complex surfaces has attracted significant interest in material science. This field has branched out over the years into different topics, such as molecular storage, drug delivery, molecular organization and aggregation, catalysis and chemical reactions at the nanoscale, to name just a few. In the context of advanced nanostructures, complex micro/nanoporous systems, such as zeolites and metal organic frameworks (MOFs), have been thoroughly studied as media to template molecules into well-defined shapes and architectures.[1,2] Depending on their composition, these hosts have the advantage of giving rather homogeneous distributions of pore sizes, which is key for guiding or engineering molecular adsorption into more complex 2D and 3D architectures. Because of the directionality of the molecular dipole, the formation of molecular chains of molecules forming head-to-tail assemblies, similar to the wagons on a train, are of particular interest. Their intrinsic anisotropy and the specific intermolecular couplings of such molecules are at the origin of various and intriguing photophysical phenomena. Due to a lack of the high anisotropy needed to restrict directions, it has been difficult to driving or templating the assembly into such one-dimensional (1D) shapes using known 3D crystalline templates. Although interesting strategies leading to the self-assembly of molecules into 1D shapes can be found,[3,4] the natural tendency of molecules to rather favour compact 2D and 3D structures makes these schemes complex to generalize by simple assembly and very hard to implement using top-down approaches such as lithography and mechanical means. A universal scheme to drive assemble in truly 1D shapes remains nowadays elusive. This review focus on recent advances on the use of nanotubes to drive hierarchical assemblies of molecules into 1D heterostructures of different shapes and compositions. Using nanotube-based templates, and in particular single-wall carbon nanotubes (SWCNTs), researchers have derived a plethora of 1D assemblies of molecules spanning lengths from micrometres down to the nano scale and exhibiting different architectures. These works have highlighted some of the most exciting and surprising hybrid structures and exposed their optical properties. This review intends to cover some of the most significant studies of the recent years as reviewed by the authors during a meeting in Paris.

Since the widespread report of their discovery in the 90's,[5] single wall carbon nanotubes (SWCNTs) have been used as a 1D template for assembling molecules and inorganic compounds into well-defined shapes, thanks to their hollow, crystalline and cylindrical architectures. The SWCNTs present high aspect ratios and can induce weak or strong molecular attachment on the outside (exohedral) and the inside (endohedral) of their constitutive wall. In both cases, the molecular adsorption produces unique 1D-like assemblies that are distinct from that obtained with other templates. These assemblies are interesting in large part due to the unique structures of various sources of nanotubes available commercially. Synthetized from different growth processes, SWCNTs are nowadays produced in large quantity with a narrow distribution of diameters, which define a perimeter space of the quasi-1D surface and the central pore having sizes between 0.2 nm and few nm. Early works have shown that the nanotube template can encapsulate other nanotubes, giving the so-called double-walled nanotube for example, and other structures such as fullerene, graphene nanoribbons and linear carbon chains. As discussed in a recent review, these 1D heterostructures have interesting properties derived from the specific arrangement of the periodic inner structures.[6] There is in addition a long list of organic and inorganic materials that were successfully encapsulated into SWCNTs (see Fig. 1c). These includes gases (H_2, CH_4), water, inorganic compounds, such as uranyl nitrate, and various kinds of organic molecules, including dyes and drugs, *etc.*

The advantage of the carbon nanotubes as a 1D host comes from the combined richness and specificity of the diameters available, which are linked to the chiral angle of its crystalline lattice (see Fig. 1a).

Indeed, more than 160 possible (n,m) structures of SWCNTs of $\leq$2 nm diameter can be derived from the lattice, giving a plethora of structures to choose from for tailored assembly. When addressed individually, two different chiralities can provide extremely small differences of available space for molecular confinement and this level of control on the assembly is probably unmatched compared to other templates. That is, the cavity is highly tunable depending on the nanotube species; diameter increments of about 0.01 nm between adjacent chiralities are accessible for studies. This richness of nanotube chiralities gives an impressive set of available templates for 1D assemblies, but the sources of nanotubes generally contain statistical mixtures of chiralities. The structural heterogeneity within a batch of nanotubes is therefore a hurdle. This complexity is further enhanced because the chiral mixtures contain nanotube species that can be either semiconducting or metallic.[7] The outstanding challenge of sorting nanotubes with selected chirality has, however, been significantly advanced, with impressive progress made *via* nanotube processing.[8,9] Scientists now have access to samples of SWCNTs with narrow chiral distributions or uniform properties. Such recent breakthroughs in the synthesis, characterization, and purification of SWCNTs by diameter and chirality, as recently reviewed in ref. 10 (see Fig. 1b) have further refined the material available for making nanotube templates for molecular adsorption. The situation is such that the main obstacles of using nanotube templates for 1D assemblies have been lowered significantly. This review highlights in this context the recent progress on fabricating the molecular hybrids with SWCNTs (Mol@SWCNT) and discusses their properties as well as the

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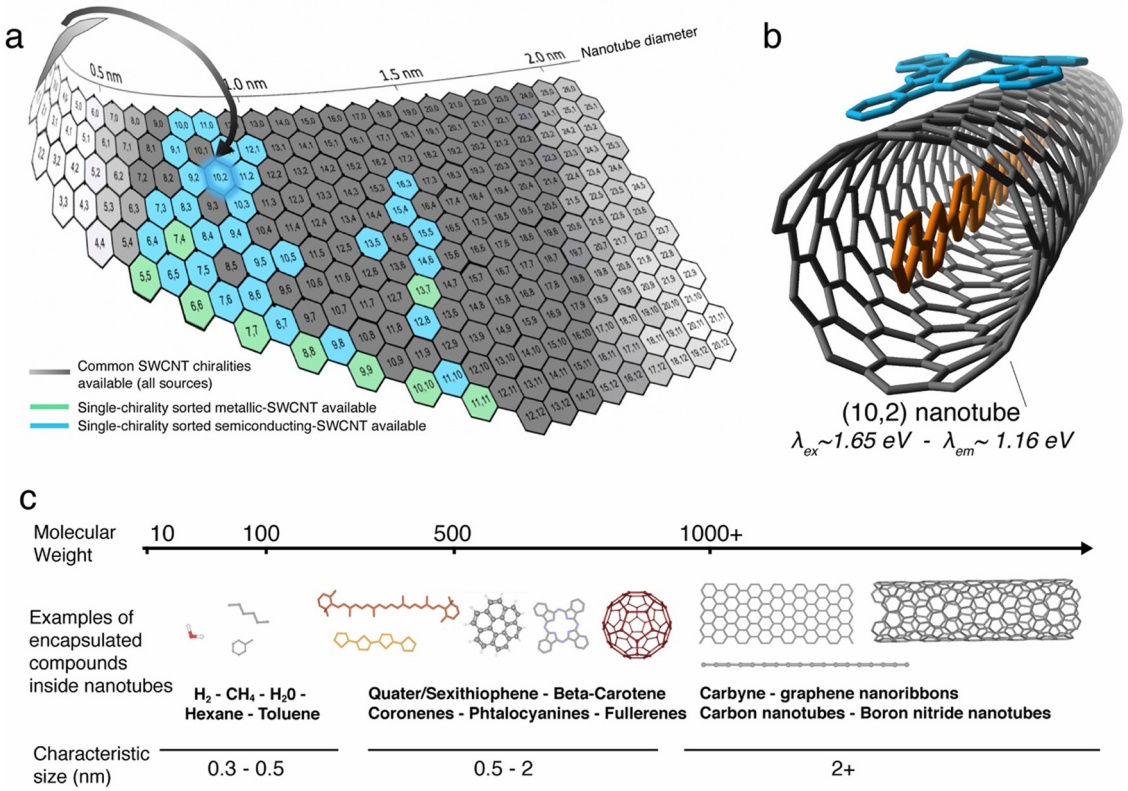

Fig. 1 Nanotubes as an *ad hoc* template for molecular adsorption. (a) Nomenclature description of the SWCNT (*n*,*m*) chiralities from a rolled graphene sheet with indicative values of the nanotube diameter. The grey scale represents the common SWCNT chiralities available from the different synthesis techniques such as CoMoCat, HiPCo, plasma torch *etc*. The chiralities highlighted in blue and in green represent the one available after common post-synthesis sorting processing with a purity higher than 99%. (b) Schematic representation of a Mol@NT system, composed by a chain of quaterthiophene encapsulated inside a semiconducting (10,2) nanotube with a Cu-phtalocyanine molecule adsorbed on the outside. (c) Examples of encapsulated compounds inside carbon nanotubes as a function of their molecular weight and characteristic size.

new horizons that pursuable using this 1D templating approach.

2. Confinement and ordering of molecules assisted by a nanotube template

2.1. Molecular ordering inside SWCNT and encapsulation mechanisms

2.1.1. Stacking configurations: observations and theoretical simulations. The formation of guest/host nanohybrids based on nanotube filled with organic compounds having a molecular mass ranging from 200 to 1000 Da is peculiar because the sizes of such molecules are comparable to a typical nanotube diameter, *i.e.* 0.5–2 nm. Consequently, the encapsulation of the molecules and the formation (or not) of molecular stacking in a 1D assembly obtained by simple diffusion are strongly impacted by the van der Waals (vdW) interactions between molecules and by the high confinement ascribed to the nanotube host. (Note that comparative description of encapsulation methods and characterization techniques for Mol@NT systems is available in the ESI.†) As an example, Fig. 2 presents HRTEM images of various organic molecules confined inside carbon nanotubes, from C_{60} (Fig. 2a and b) to graphene nanoribbons (Fig. 2f) displaying different stacking configurations dependent on the specific molecule and nanotube diameter.[11–16] An interesting trend is that the 1D confinement of elongated dye molecules, such as quaterrylene or quarter-/sexi-thiophene (4T–6T), induce their overall alignment of the molecules all along the nanotube axis, leading to a head-to-tail aggregation state of the molecules. Contrastingly, when the 1D confinement is applied on disc-like molecules, such as coronenes, a face-to-face aggregation is preferentially induces, in which the molecules are tilted with a ~90° angle with respect to the nanotube axis (Fig. 2g).

Elongated molecules stack into single or multiple columns depending on the nanotube geometry and the encapsulation conditions (Fig. 2c–e). Probing such configurations beyond microscopy methods, multiple research groups use Raman spectroscopy of the molecules to acquire evidence of encapsulation and to better understand the stacking trends in Mol@SWCNTs. For example, the amplitude of the shift of the radial breathing modes (RBM) in the Raman scattering spectra of SWCNTs was found to trend with the number of molecules stacked inside the nanotube as a function of

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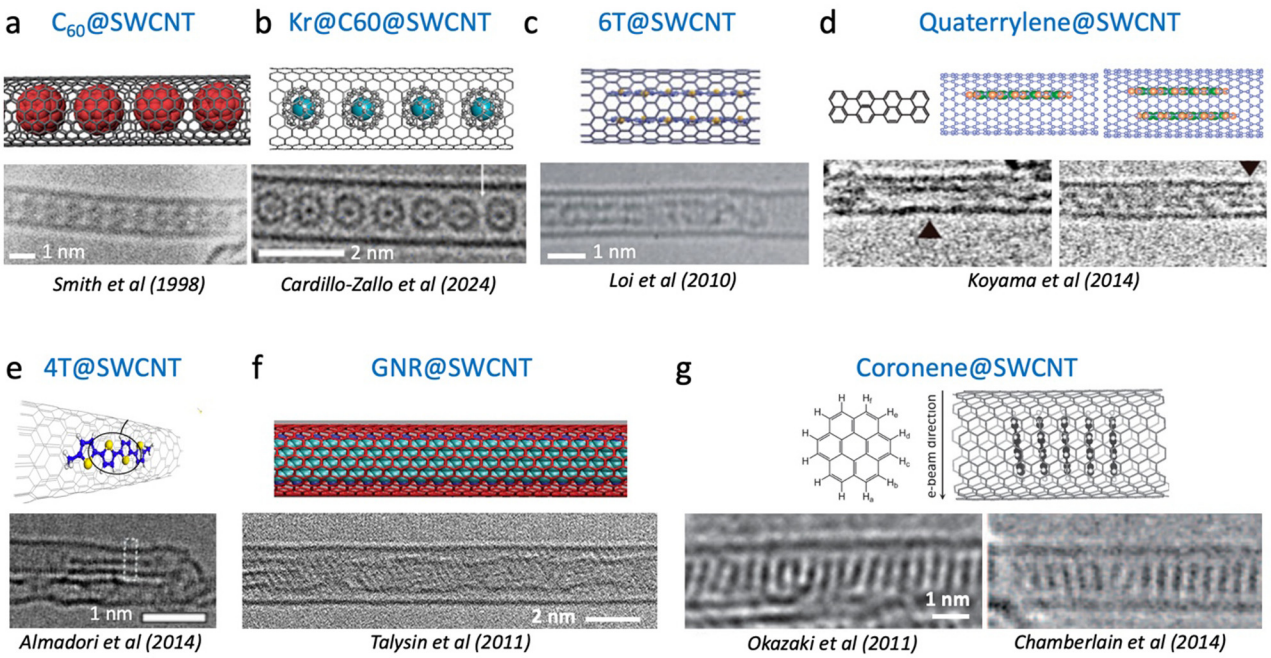

Fig. 2 High resolution transmission electron microscopy (HRTEM) images and illustrations of different stacking geometries of organic molecules encapsulated inside carbon nanotubes. (a) C60 fullerenes inside SWCNT. Scale bar shows 1 nm. Adapted with permission from ref. 17 and 18 first published in 1998 in ref. 19. (b) Kr@C60@SWCNT, adapted with permission from Cardillo-Zallo, *ACS Nano* 2024, **18**, 4, 2958–2971, Copyright 2024 American Chemical Society, ref. 18 and 20. (c) Double chain of alpha-sexithiophene inside SWCNT. Adapted from *Adv. Mater.* 2010, **22**, 1635–1639. Copyright 2010 Wiley-VCH.[14] (d) Single and double rows of quaterrylene molecules aligned along the SWCNT axis. Adapted with permission from *J. Phys. Chem. C* 2014, **118**, 37, 21671–21681.[12] Copyright 2014 American Chemical Society. (e) Identification of single quaterthiophene inside a SWCNT. Adapted with permission from *J. Phys. Chem. C* 2014, **118**, 33, 19462–19468. Copyright 2014 American Chemical Society.[13] (f) Graphene nanoribbons inside a SWCNT (GNR@SWCNT) from Talysin *et al.*, *Nano Lett.* 2011, **11**, 10, 4352–4356, Copyright 2011 American Chemical Society.[21] (g) Coaxial stacking of coronene molecules perpendicular to the SWCNT axis, adapted with permission from *Angew. Chem., Int. Ed.* 2011, **50**(21), 4853–4857, and *Small* 2015, **11**, No. 5, 622–629. Copyright 2011 Wiley-VCH.[15,16]

the diameter (see Fig. 3a). In the case of 4T@SWCNTs, two different plateaus are observed between 0.8–1.1 nm and between 1.45–1.65 nm, which are attributed to the formation of a single aggregate or pair aggregates into the hollow core of the nanotubes. This assumption is consistent with TEM experiments.[13]

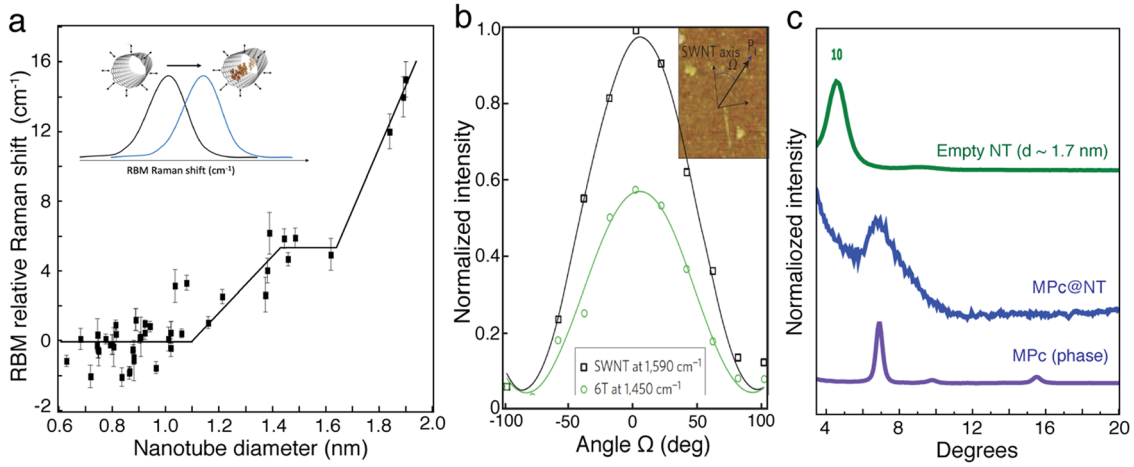

Fig. 3 (a) Radial breathing mode (RBM) relative up-shifts after encapsulation of 4T molecules inside nanotubes (NT) of different diameters. (b) Variation in intensity of two Raman peaks (at 1590 cm^{-1} and 1440 cm^{-1} for the semiconducting SWNT and 6T modes, respectively) of an individual 6T@s-SWNT *versus* laser polarization angle (Ω) at λ = 532 nm. The inset is an atomic force microscopy image of the SWNT and a description of the polarization angle. Adapted from Gaufrès *et al.*, *Nature Photon* 2014 **8**, 72–78, copyright 2013 Springer Nature.[22] (c) Diffraction pattern of pristine and functionalized nanotubes with a diameter distribution centered around 1.7 nm, together with the α MPc phase. The 10 peak of the nanotube bundles is located at 4.57°. These positions correspond to an inter reticular distance of 19.3 Å.

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On aligned or single nanotube samples, polarized Raman spectroscopy, taking advantage of the highly polarized resonant Raman signal of SWCNTs, has also been used as an internal reference for probing the transition dipole alignment of elongated molecules. For example, polarized micro Raman spectroscopy performed on individual SWCNTs filled with sexithiophene molecules (6T) (see Fig. 3b) showed that the intensities of the 6T molecular mode at 1450 cm^{-1} and of the SWCNT longitudinal mode at 1590 cm^{-1} have the same polarization-dependency. This result provides strong evidence that the transition dipole moments of the molecules are strongly aligned along the nanotube axis.[23]

The 1D character of the molecular organization was also probed by X-ray diffraction measurements on phthalocyanine molecules (MPc) encapsulated inside larger diameter SWCNTs (1.4 nm $< d <$ 2.1 nm). The evidences are summarized in the Fig. 3c.[24] Two main features are observed after the encapsulation: the MPc molecules are: the disappearance of the contributions associated to empty SWCNTs and the appearance of a new broad peak at around 6.9°. This value coincides with the diffraction peak of the MPc molecules' α phase, which corresponds to the distance between molecules in the [200] direction. Thus, this additional peak suggests the 1D stacking of the MPc molecules encapsulated inside the carbon nanotubes.

The results presented in Fig. 2 and 3 highlights the complexity of the interactions between molecules and between molecules and the nanotube sidewall. To shine a light on these interactions, *ab initio* methods have been used to identify and understand many specific observed behaviors. In this perspective, several groups have calculated the binding energy for π-conjugated systems encapsulated inside SWCNT. For example, it was reported that the optimal distance between a 6T chain and the nanotube wall is roughly 4 Å and that, depending on the diameter, the more favourable configuration is a single or double pair of parallel 6T, with a preferential adsorption of the 6T on the nanotube sidewall.[14] It was also showed that depending on the diameter, the more favourable configuration is a single or double pair of parallel 6T, with a preferential adsorption of the 6T on the nanotube sidewall. This explains the observation of a narrow distribution of the molecule-wall distance that is well centered at 3–4 Å and a wide distribution of the adjacent distances for molecule-molecule spacing, in the range between 4 and 8 Å. Similarly, for terthiophene that the host–guest interactions dominates in large diameter nanotubes compared to the molecules size, and leading to the adsorption of the molecules on the wall.[25] Within a smaller-diameter nanotube, host–guest interactions were hypothesized to compete with the guest–guest interactions and promote a final stacking that is strongly dependent on the diameter. It was also shown that the host–guest interactions are sufficient to deform the encapsulated π-conjugated oligomers molecular structure.[26]

In a separate study, using dispersion-corrected density functional theory (DFT), Yumura *et al.* calculated the potential energy surfaces of the formation of a second row of asymmetric and polarized *p*-(dimethylamino)-*p*′-nitrostilbene (DANS) molecules inside a SWCNT.[27] The study shows that the insertion of DANS into a nanotube with a diameter of approximately 1.0 nm spontaneously proceeds through attractive host–guest interactions. Various metastable states and a global optimization structure appear to evolve from one end to the middle of a tube. Additional potential energy surface calculations on the insertion of another DANS molecule into the nanotube already containing one row of molecules show the significant roles of intermolecular interactions, which can lead to high energy barriers with values that strongly depends on the relative distribution of the polarized charge inside the nanotubes, *i.e.* the charge repartition along the nanotube axis.

The behavior of short molecules inside SWCNTs has also been studied by DFT. For example, Almadori *et al.* have reported DFT calculations of phtalocyanine molecules inside SWCNTs, which indicate that the minimum of the bonding energy was obtained for an angle of 32° between the molecular ring and the nanotubes axis. This represents an intermediate case between polythiophene@SWCNTs (0°) and coronene@SWCNT (90°).

2.1.2. Kinetic aspects and encapsulation mechanisms. The previous section focused on the final stacking state of the molecules inside the nanotubes at thermodynamic equilibrium. To tailor desirable aggregation, however, controlling the kinetic and dynamic adsorption mechanisms underlying the encapsulation process is crucial. For molecules that are small compared to the size of the nanotube cavity, *e.g.* Kr, Ar or CH_4, studies of encapsulation in 0.7–2 nm diameter SWCNTs have shown evidence of capillary condensation and multilayer formation, giving a nearly ideal two-dimensional (2D) adsorption behaviour inside the nanotube cavity.[28,29] For large molecules, defined in this context as approaching the cavity size, however, greater confinement and vdW interactions further restrain molecular motions and force the encapsulation to become one-dimensional (1D).

Capitalizing on the very strong, ''giant'', Raman scattering of aligned sexithiophene molecules (6T) inside SWCNT (6T@SWCNTs),[23] the filling fraction of 6T in a SWCNT (R_{6T}) was quantitatively estimated by measuring the Raman ratio $R_{6T} = I_{6T}/I_G$, (where I_{6T} is the Raman 6T-band intensity at 1450 cm^{-1} and I_G is the G-band intensity at 1590 cm^{-1} of the carbon nanotubes) from a statistical assembly of individualized SWCNTs.[30] (see Fig. 4a). By varying the conditions of the encapsulation process (temperature and 6T concentrations) the encapsulation isotherms of this system in liquid phase were obtained (see Fig. 4b). Below a specific 6T concentration (C1 = 2.10^{-7} M), R_{6T} follows the same evolution for the two temperatures studied, which indicates a direct (spontaneous) encapsulation mechanism in these conditions. This is consistent with previous simulations of large molecules encapsulation inside carbon nanotubes.[26,31] Above the concentration threshold C1, a second step is revealing of another process, which appears to be thermally activated and associated to the formation of a pair-aggregate (Fig. 4c).

By modelling the results in Fig. 4c using a modified Langmuir isotherm, the formation enthalpy for the insertion of a

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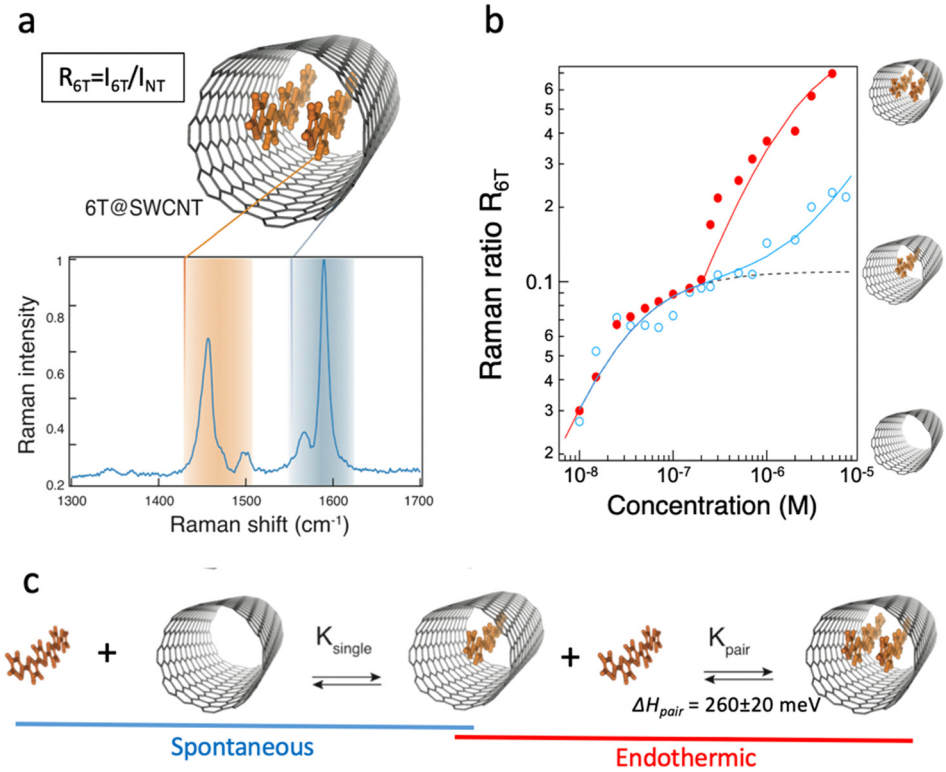

Fig. 4 (a) Schematics of a single pair-aggregate and a typical Raman spectrum of 6T@SWCNTs at an excitation wavelength of 532 nm. The intensity is normalized with the signal of the G-band of the SWCNTs. (b) Raman ratio R_{6T} at 532 nm of 6T molecules encapsulated inside SWCNTs as a function of the concentration of 6T in toluene at 30 °C (blue) and 115 °C (red). (c) Encapsulation scenario for single and paired aggregates of 6T into SWCNTs. The two sequential processes are controlled by the equilibrium constant K_{single} and K_{pair}. Adapted with permission from *ACS Nano* 2016, **10**, 11, 10220–10226. Copyright 2016 American Chemical Society.[30]

second row of molecules inside the SWCNTs (having a diameter distribution centered at 1.35 nm) was determined to be ΔH_{pair} = 260 ± 20 meV. The formation of pair aggregates also occurs homogeneously in SWCNTs to up to 140 microns in length, evidencing easy diffusion of the molecules along the inside of the nanotubes. These preliminary works are appealing for further experiments on the dynamic aspect of the encapsulation and the understanding of the entangled regimes of molecular diffusion and adsorption equilibrium in complex 1D heterostructures.

2.2. Adsorption of (porphyrin) molecules on the outside of carbon nanotube sidewalls

The SWCNT exterior can also template the coupling of the nanotube with molecules. This configuration lifts the constraint of the size of the molecule of interest. For instance, molecules such as porphyrins are too large to be encapsulated in commonly utilized diameter SWCNT. However, they can stack on the outside of the nanotube. In this part, we will focus on one specific way to reach the grafting of molecules on the nanotube sidewall: the functionalization of nanotubes in micellar suspensions.

The use of ultrasonication and surfactants to individualize and stabilize a SWCNT in a micelle is now widely used since their first report in 2002.[32] In 2008, Ziegler and co-workers also showed that the incorporation of water-immiscible organic solvents inside the micelles containing SWNTs was possible.[33] The presence of a small amount of organic solvent in the vicinity of the nanotubes can be evidenced by fluorescence spectroscopy, causing a shift of the nanotubes fluorescence lines due to a solvatochromic effect (see Fig. 5a).

Swelling the micelle using an organic solvent opens the way to use it as a vector to bring molecules of interest in the vicinity of the nanotube sidewall. For example, the functionalization of CoMoCat carbon nanotubes with tetraphenyl porphyrin (TPP) using a micelle swelling approach based on sodium cholate and dichloromethane was reported.[34] The presence of the TPP stacked on the wall of the SWCNTs resulted in a clear red-shift of the TPP absorption features (420 to 440 nm) and a concurrent shift as well to the nanotube transitions (Fig. 6).

In this micelle swelling process, the microscopic mechanism is related to the exchange of the surfactant molecules and the chromophore with specific kinetic parameters; the bigger is the diameter the easier it is to stack the molecules on the sidewall of the nanotube[35] (Fig. 6b). The kinetics of the reaction can be controlled by playing with the relative proportions of surfactants in co-surfactant suspensions.[36] For example, the surfactant exchange timescale in these hybrid micellar system could be tuned from a multiple days timescale to as brief as a

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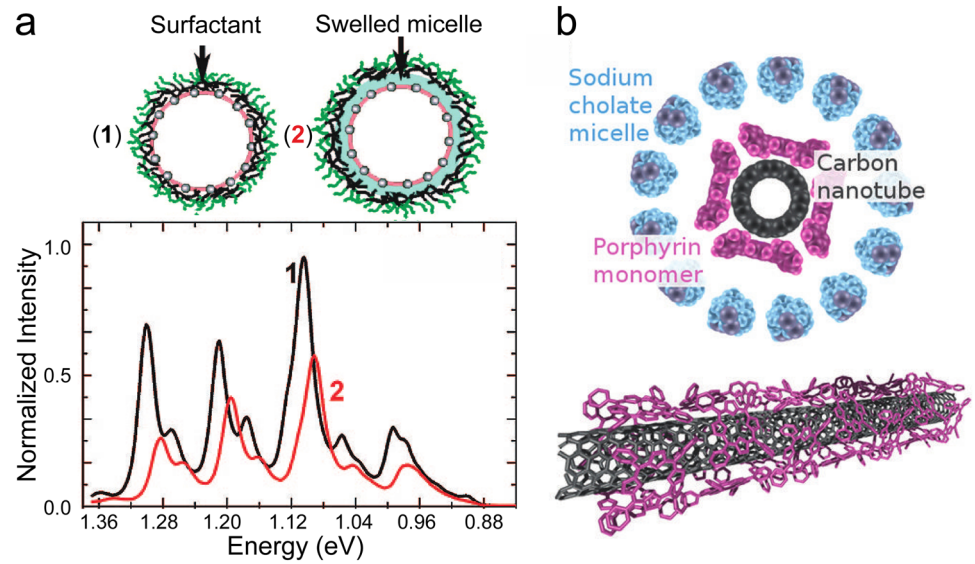

Fig. 5 (a) Swelling of the hydrophobic core of the micelle surrounding a SWCNT (up) NIR fluorescence spectra of an sodium dodecylbenzene sulfonate (SDBS)-coated SWCNT suspension after addition of *o*-dichlorobenzene (ODCB) (bottom), adapted with permission from *J. Am. Chem. Soc.* 2008, **130**(48), 16330–16337, Copyright 2008 American Chemical Society.[33] (b) Schematic representation of the micelle wrapped and functionalized compounds, side and 3D views.

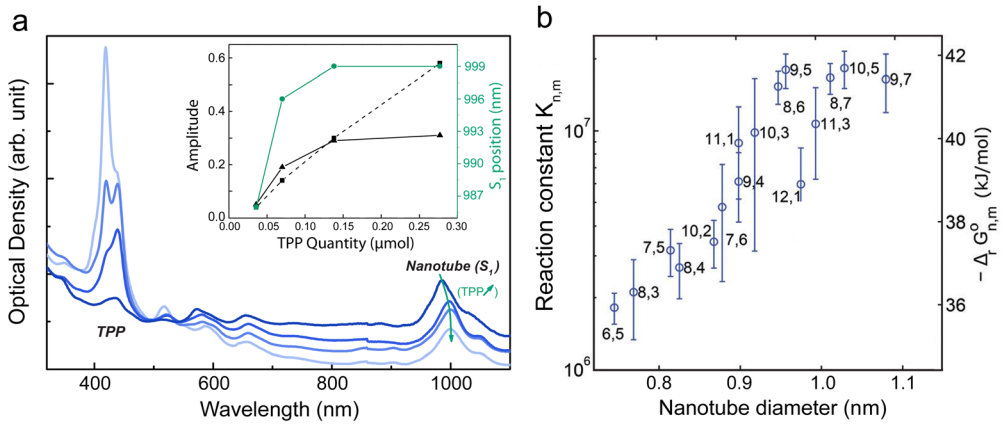

Fig. 6 (a) Optical absorption spectra of SWNT/TPP complexes for different concentrations of porphyrin with a DCM/water volume ratio of 34% 0.04 μmol, 0.07 μmol, 0.14 μmol and 0.27 μmol (from dark blue to light blue), within a DCM/water solution of volume ratio of 2 : 1. The curves are vertically translated to match at 490 nm (background correction). Inset: Amplitude of the band at 420 nm (dashed line) and at 438 nm (black line) as a function of the quantity of porphyrin; shift of the S11 band (grey line) as a function of the quantity of porphyrin. Adapted with permission from *ChemPhysChem* 2010, **11**, 1667–1672, Copyright 2010 Wiley-VCH.[34] (b) Reaction constant (log scale) for several (*n*,*m*) nanotube species as a function of the diameter of the species. Reproduced from ref. 35 with the permission of Royal Society of Chemistry.

minute. Interestingly, porphyrin molecules self-organize with their O_x direction along the nanotube axis,[37] underlining that, even if the 1D stacking outside the nanotube is weaker compared to inside the nanotube, a specific templating of molecules can be achieved simply by an exohedral stacking on the nanotube sidewall.

The micelle swelling method has also been extended to enable more complex chemical functionalizations, using the micelle core as a nano platform to perform organic chemistry reactions. Through the realization of polymerization processes inside the micelle core itself.[38] The dimerization of thiol groups from thiol functionalized porphyrins, put into contact with nanotubes through the micelle swelling method, form a cross-linked network of porphyrin on the nanotube surface (Fig. 7). The functionalized nanotubes can then be dispersed in solvents without loss of the organic shell allowing an easy manipulation of the objects. This strategy has been extended to other types of molecules. Examples include putting thermo-responsive acrylamide polymers or polystyrene around nanotubes, allowing the fabrication of core–shell structures where

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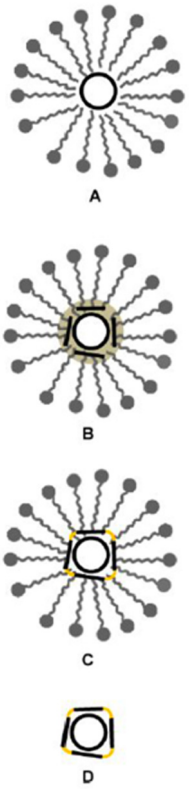

Fig. 7 Schematic representation of the different steps of functionalization: (A) the nanotubes are dispersed and individualized in micelles; (B) the porphyrins in an organic solvent are introduced on the nanotube surfaces; (C) after solvent evaporation, the reaction to cross-link the porphyrin is performed leading to a stable organic shell around the nanotubes; and (D) the nanotube derivatives are purified *via* filtration and extensive washing to remove the surfactants, reagents, and unbound porphyrins. Adapted with permission from *Chem. Mater.* 2013, **25**, 13, 2700–2707. Copyright 2013 American Chemical Society.[38]

the active core is the nanotube and the polymer shell protects the nanotube from its local environment.[39,40]

These porphyrin/nanotubes hybrids form a class of systems with complex intermolecular coupling leading to specific optical processes such as energy transfer. These excitonic effects are presented in Section 4.2.

2.3. Water & solvent phases in CNTs

"The access to chirality-controlled sources of SWCNTs with closed or open ended termination is a key point to study the 1D confinement of small molecules inside SWCNTs such as water."

As discussed previously in this review, the endohedral environment of SWCNT samples can be empty or filled with various molecules having a size-scale similar to the nanotube cavity pore size, such as polythiophene, phtalocyanine or beta-carotene. In this subsection we focus on the confinement of small molecules compared to the nanotube hollow core, *e.g.* solvent-like organic molecules and water molecules. We address the particular importance of closed-ended, empty SWCNTs for comparison to filled SWCNTs, and the reported effects of filling both on the confined molecules and the SWCNTs properties.

As early as 2000,[41] molecular dynamics simulations predicted that water would enter into the SWCNTs endohedral cavity despite the assumed hydrophobic nature of their interior, and that confinement inside a SWCNT would invoke unusual behavior of the encapsulated water molecules such as ultra-fast diffusion[42–44] or atypical phase diagrams.[45–48] A comprehensive review on the molecular modelling of confined water molecules can be found in ref. 49. Building excitement in this research direction, ultrafast transport through a nanotube was experimentally demonstrated in 2006 in DWCNTs;[50] shortly thereafter the potential of water-filled CNTs for use as nanovalves to control gas flow was also validated.[51] Since then, much research effort has been devoted to water-filled CNTs and their applications, with a focus on understanding and engineering efficient (or selective) ion transport through the filled SWCNTs pores.[52,53] By probing the encapsulation-induced changes in water directly by nuclear magnetic resonance (NMR),[54–56] XRD,[57–59] IR spectroscopy[60,61] and neutron scattering,[62,63] the effect of the confinement on the water molecules could be investigated. Such studies revealed phenomena such as multi-ring stacked water structures inside broader diameter SWCNTs,[60] a step-wise filling of CNTs depending on the surrounding water vapor pressure,[59] and the formation of *n*-gonal ice structures at lower temperature.[47] However, each of the techniques used in these early contributions were not able to differentiate the effects of any SWCNT specific chiral structure surrounding the encapsulated molecules in the absence of chirality separated SWCNT samples. Additionally, most of these techniques also could not provide quantitative information on other parameters of interest for filled SWCNTs, *e.g.*, the percentage of filling inside the SWCNTs. Broadly, this question about quantifying the fractions of closed/opened SWCNTs in a given sample, *i.e.* the fraction of filled/unfilled SWCNTs, is crucial to quantitatively understand the impact of the filling on the SWCNT properties. Indeed, it is consistently observed that water filling of SWCNTs occurs spontaneously and quickly given any opening in the SWCNT carbon lattice. Literature can be found on how to realize solutions with open or closed SWCNTs by combining surfactant wrapping and density-gradient ultra-centrifugation of SWCNTs.[64–69] To estimate the ratio of opened/closed SWCNT populations, calibrated optical spectroscopy such as the Raman scattering can be used to monitor the radial breathing mode (RBM) of the SWCNTs.[64,70]

Having purely empty and purely water-filled samples available,[68] it became possible to investigate the influence of encapsulated water on the electronic properties of the SWCNTs in more details. In comparison to the optical properties of empty SWCNTs, water-filling clearly results in severe broadening and a red-shifting of the electronic transitions, as well as a strong quenching of the photoluminescence emission of the semiconducting SWCNTs (Fig. 8). These differences are observed both at the bulk dispersion level, and at the level of a single nanotubes.[70] The same effects were observed on the

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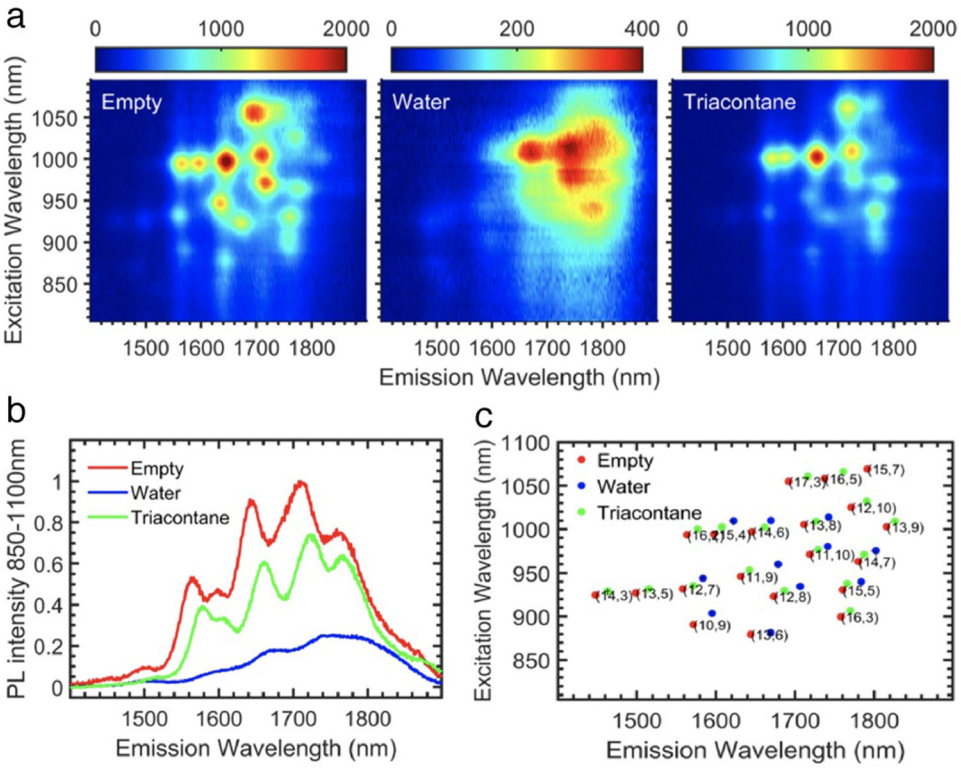

Fig. 8 (a) Experimental PLE maps of empty (left), water-filled (middle) and triacontane-filled (right) SWCNT samples, normalized over absorption so that the actual PL intensity can be directly compared (note the 5× difference in color scale of the water-filled sample). (b) Integrated PL intensities of empty, water-filled and triacontane-filled SWCNTs over the excitation range of 850–1100 nm, both demonstrating the drastic quenching and broadening of the PL form the water-filled SWCNTs in comparison to the empty SWCNTs, and the approaching similar intensity of alkane-filled SWCNTs to empty SWCNTs. (c) Peak positions of empty, water-filled and triacontane-filled SWCNTs obtained by fitting the above PLE maps, showing the electronic shift of the peak positions upon water-filling. Figure obtained with permission from a combination of *Carbon* 2015, **95**, 442–451 and *ACS Nano* 2021, **15**(2), 2301–2317, Copyright 2021 American Chemical Society.[73,74]

vibrational and electronic transitions of long (≈7 to 10 μm) suspended individual SWCNTs after opening the individually suspended SWCNTs by laser heating and controlling the water vapor pressure in the chamber.[71] The shift of the RBM of the SWCNTs due to water filling, for a wide range of individual SWCNT chiralities, was found to depend on the SWCNT diameter in the large-diameter regime.[72] In addition, water-filling can also strengthen the SWCNTs and make them more robust with respect to radial compression.[73]

One interesting aspects that were predicted theoretically is for water molecules behave quite differently under confinement, templating new forms of ice structures to occur with precise structure and phase transition temperature depending strongly on the diameter of the surrounding SWCNTs[44–48] Such new phase transitions, were first experimentally observed in large diameter SWCNTs (1.6–2.4 nm) by XRD and NMR.[75] However, the diameter dependence (or chirality-dependence) could not be revealed due to the broad diameter distributions considered. However, reversible freezing-thawing transitions of encapsulated water in different diameter nanotubes were observed for a few individualized large-diameters SWCNTs, resulting in shifts of the RBM vibration by 2 to 5 cm^{-1}.[76] These larger shifts were attributed to variations in the compressibility due to the phase change. Independently, and through the alternate strategy of utilizing separated populations, A quasi-phase transition within a one-dimensional chain of water molecules in (6,5) SWCNTs was also observed; at low temperature, the water dipole moments were ferro-electrically oriented with all dipoles pointing along the axis of the SWCNTs.[77] Such phase transitions could also be observed at distinct temperatures for larger diameter SWCNTs.[78] Fig. 9 summarizes these findings in a diameter-dependent temperature plot showing the complexity of the diameter-dependence of these phase transitions.

Although water-filling of CNTs is scientifically interesting due to its analogy with biological membranes and new physics of confinement, it is often a strong hurdle for other studies or when preparing SWCNT dispersions for use in applications. A main factor in this is from its detrimental impact on the optical properties of the SWCNTs, which as noted above are red-shifted and reduced in fluorescence emission efficiency. While the use of empty SWCNTs is an obvious route to avoid these effects, they are simply not present in sufficient fractional quantity in most SWCNT soots for many applications.

One idea to overcome these hurdles is to instead controllably fill open-ended SWCNTs with an inert molecule that will

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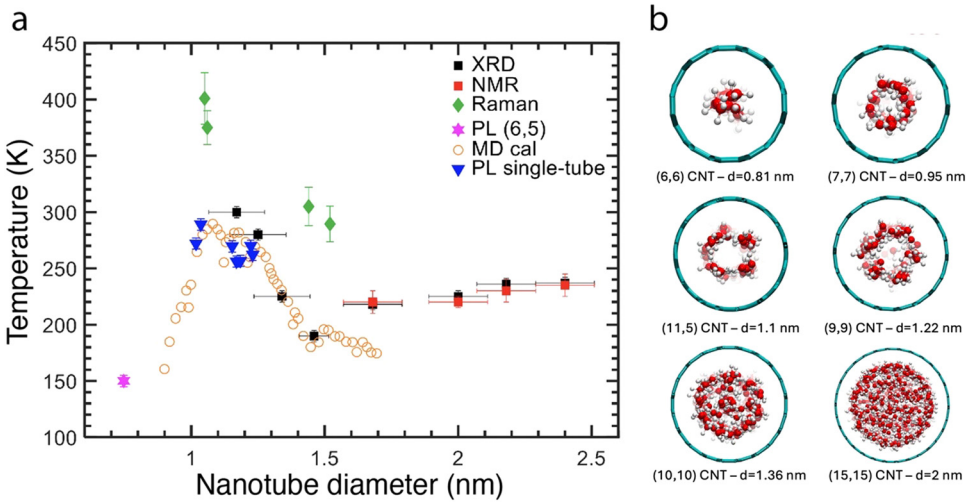

Fig. 9 (a) Experimental and theoretical solid–liquid phase transition temperatures of water confined inside SWCNTs as a function of the SWCNT diameter. Behavior extractable from macroscopic parameters, asymptoting to the bulk value of T = 273.15 K at 1 atm pressure in an unconfined system, is expected for pore diameters >4 nm.[48] Disruption to the molecular ordering of water due to confinement below this length scale is predicted to result in complex, pore diameter-dependent phase transition temperatures. Reported values from calculations and experiments include extreme deviations, both positive and negative, from the bulk melting point. Data is collected from different references: MD calculations (orange open circles) are from ref. 47, XRD (black squares) and NMR (red squares) experimental data points are obtained on SWCNT samples with a broad diameter distribution in ref. 75. Green diamonds are obtained from Raman spectra of the RBMs of individual SWCNTs.[76] Blue triangles are obtained from PL experiments on individual SWCNTs[78] and the magenta star is a data point obtained by microscopic PL experiments on a sorted (6,5) SWCNT sample.[79] Note that the latter is not assigned as a real liquid–solid phase transition, but as a quasi-phase transition in the orientation of the molecular dipole moments of the water molecules as this diameter only allows for a single chain of water molecules to be encapsulated. Similar complexity in melting point temperatures can also be expected for other confined molecules. (b) Illustration of different water states in SWCNT as a function of diameter, adapted from ref. 80.

be stably contained inside the SWCNT cavity prior to dispersion. Such an approach was first purposely attempted by Campo *et al.*,[81] for the formation of alkane-filled SWCNTs, although solvent molecules and other organic molecules desired for their active material properties had been inserted into SWCNTs as early as the late 1990s.

Excitingly, these alkane@SWCNT (and similar compounds) display optical absorbance spectra that are generally blue-shifted, with narrowed linewidths, compared to water-filled SWCNTs, although still red-shifted compared to the peak transition wavelengths of empty SWCNTs (Fig. 10). For the fluorescence, a substantial enhancement relative to water-filled SWCNTs was observed, approaching but not quite equalling that of empty nanotubes for the longest length alkane samples (Fig. 8c).[74]

Nearly unique to the Campo *et al.* effort, however, was the further dispersion of the linear alkane and similar saturated hydrocarbon molecule filled SWCNTs using bile salt surfactants in aqueous solution, and the application of dispersion-based SWCNT purification processes such as rate-zonal ultracentrifugation and aqueous two polymer phase extraction (ATPE).[82] Later, this effort was extended by separately filling of two different diameter range SWCNT populations with molecules of different bulk dielectric constant.[74] The dispersion and separation of semiconducting species and the effects of the internal dielectric constant on the observed solvatochromism and fluorescence intensity of the SWCNTs were also reported.[74,83,84]

As an additional complexity, the insertion of alkane molecules into particularly small diameter nanotubes induces radial expansion strain-based shifts in the SWCNT optical spectra, adding further complications to analysis.[84,85] Similar to water encapsulation, computational modelling of a linear alkane insertion into a (6,5) found the process to occur spontaneously at room temperature for a fully open SWCNT end, with the alkane molecule adopting an extended configuration within its energetically sampled distribution to fit into the cavity space.[84] While not yet experimentally reported for individual (n,m)s, calculations for the packing of a linear alkane (paraffin) have been reported for larger diameter (n,m) SWCNT structures,[86] and report reduced melting points for the alkanes due to the confinement. Efforts to study confinement effects both on the arrangement and density of filler molecules inside the nanotube, and on the nanotubes themselves, are obvious next progressions characterization in these non-water molecule@SWCNT systems.

2.4. Inorganic compounds inside CNTs

Despite the discussion above of inserting organic molecules into SWCNTs, inorganic compounds were the first class of materials to be inserted intentionally within nanotubes, and especially carbon nanotubes. This research field of inserting inorganic compounds is indeed as old as the research on carbon nanotubes themselves, with pioneering works from Ajayan[87] and others as recently reviewed, by the Green's group.[88] There are multiple possible routes for filling

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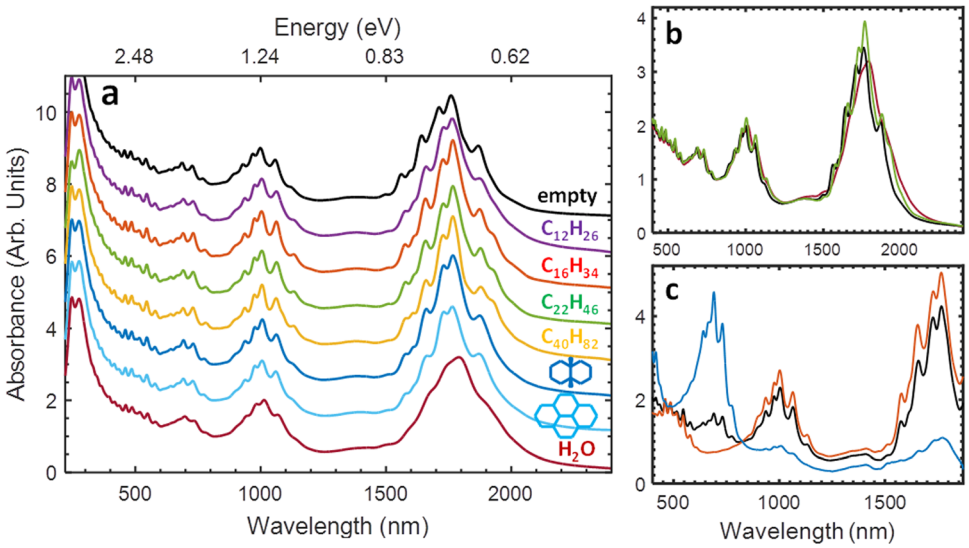

Fig. 10 (a) Absorbance spectra of arc-discharge SWCNTs dispersed in aqueous sodium deoxycholate (DOC) solutions with different endohedral fillings (vertical offset 1 unit/spectra). From top: empty core, dodecane-filled, hexadecane-filled, docosane-filled, tetracontane-filled, *cis*-decalin-filled, perhydropyrene-filled, and water-filled. The optical transition peaks exhibited by the alkane-filled nanotubes are similar to each other, while intermediate in position and breadth to the empty and water-filled nanotubes. (b) Absorbance spectra of empty (black), docosane-filled (green) and water-filled (red) nanotubes from panel (a), without vertical offset. Narrower optical transition linewidths concentrate the optical intensity of the SWCNT peak features. (c) Absorbance spectra of aqueous two-phase extraction isolated metallic (blue) and semiconducting (red) SWCNT populations of octadecane-filled SWCNTs (black). All spectra are normalized at their 810 nm valley. Reproduced with permission from ref. 81.

nanotubes,[89,90] but in the specific case of inorganic materials, the main ones are the capillary filling by immersion into either a concentrated solution or the molten compound itself. Because many inorganic compounds can withstand the required thermal treatment, as opposed to most organic compounds (thermal degradation), direct filling using a molten salt is particularly common. Moreover, opening and filling generally occur simultaneously in such cases, albeit depending on the chemical reactivity between the nanotube and the filling material at high temperature.[87,91] Finally, filling may also be obtained incidentally during the synthesis of the nanotubes; however, such approaches usually lead to the formation of rather defective multiwalled CNTs with a large diameter distribution and low filling rates.[92–94] After briefly reviewing the main results described in the literature, we focus here on the interplay between the inner channel of the nanotube and the filling material and its impact on the structural modifications of the latter (compared to the bulk) and conclude with some examples of possible applications of nanotubes filled with inorganic compounds.

2.4.1. Filling nanotubes with inorganic compounds from solutions. The solution filling method is usually obtained by incubation of a concentrated solution of the desired material (or more generally a precursor) with the nanotubes. This is a very flexible and useful alternative method for filling nanotubes, particularly when the physical properties of the targeted material to be encapsulated are not compatible with other filling techniques. Initially developed and extensively used by the Greens' group,[95] this method was often used to fill nanotubes with oxides (more rarely with metals), such as multiwalled CNT with Fe, Co, Ni, Uranium oxides,[96] $CoFe_2O_4$,[97] SnO_2,[98] or metals such as α-Fe[99] or Ag.[100] More recently, CdS was incorporated from Cadmium acetate,[101] as well as SbSI and SbSeI, directly from the elements using an original sonochemical approach.[102] It is also possible to fill nanotubes with colloidal suspensions of nanocrystals, such as tungsten polyoxometalate.[103]

2.4.2. Filling nanotubes with inorganic compounds from molten compounds. Although first proposed by Ajayan,[87] this technique was mainly developed by Sloan and co-workers since 1998, initially working with eutectic salts mixtures.[104] Because this technique is rather easy to use, it was quickly extensively used to fill CNT with a very large set of different inorganic compounds including many halides[88] which will not be described in details here. A recent notable advance in the field is the demonstration of the possibility of closing the MWCNTs after filling *via* simple heat treatment above 900 °C,[105] previously also observed for SWCNTs.[64] The main advantages and drawbacks of this approach are shared with all other filling routes: because the filling is performed after synthesis, it is possible to select the nanotubes to be filled (CNT, Boron nitride nanotubes, *etc.*). The main issue concerns the cleaning of the samples after filling, which is required to remove all the material outside the nanotubes, especially the material trapped in the interstitial areas inside CNT bundles. For these reasons, the characterization of the samples after filling needs to be performed very carefully in order to assess with a high level of confidence the actual nature of the final filling material. Filling of nanotubes with inorganic compounds may also be performed in the gas phase. Among many successful examples

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are filling with $ZrCl_4$,[106] Selenium[107] and Rhenium oxides.[108] Notably, filling has also been demonstrated for iodine atoms chains starting from different metal iodides, and especially with Nickel iodide.[109]

2.4.3. Structural modifications. An expected challenge for characterizations is that the crystal structure of the encapsulated material rarely coincides with that of the same material in bulk due to the confinement of the CNT cavity. This strongly depends on both the crystal structure of the material (either very simple such as a cubic one, or more complex in the case of layered compounds) and the diameter of the inner channel: the narrower the latter is, the higher the impact. Distortions usually occur mainly perpendicular to the axis of the nanotube because there is no space restriction along the axis. In the case of very narrow nanotubes (*ca.* 1–2 nm inner diameter), the limited available space for crystallization makes it impossible for all atoms to be surrounded by the same number of neighbours than that they normally have in the bulk structure. This leads to reduced coordination of the atoms and is likely to induce some modifications of their electronic properties. A case study example is the filling of SWCNT by cubic Potassium iodide (KI) reported by Sloan *et al.*[110] in which the coordination of K^+ and I-ions is reduced from 6 : 6 to 4 : 4.

Although the confinement effect of the nanocrystals is easily understood and accepted, another situation may occur where not only the encapsulated material is distorted, but the container as well. This may occur *a priori* only in the case of SWCNTs.[111] One of the best illustrations is that, for similar inner diameters, filling is observed for the SWCNTs and not for the DWCNTs[112] because the distortion is not possible for two walls at the same time.

2.4.4. Examples of potential applications of nanotubes filled with inorganic compounds. Many modifications of the electronic properties of the filling material and/or of the nanocrystal@nanotube nanocomposites have been reported in the literature. It is for example possible to increase the electrical conductivity of CNTs by doping with I_2 or Copper chloride in order to prepare transparent conducting electrodes.[113] Following Raman characterisations revealing charge transfer between the filling material (Se,[114] HgTe[115]) and the nanotubes, it was shown that the interactions between the core and the shell could be used to tune the electrical transport properties of double-walled CNTs.[116] The impact on the magnetic properties of nanotubes was also demonstrated in the case of filling with magnetic compounds. Recent demonstrations were proposed with the possibility of enhancing magnetic interactions in carbon systems through filling with Gadolinium highlighted potential applications to the development of spin-based quantum computing elements based on 1D channels[117] or applications to nano-mechatronics and spintronics.[118] Applications of such filled nanotubes are also proposed in catalysis,[119] biology and medicine,[120] environmental remediation[121] and for gas sensors.[122] Complementary review article focusing on synthesis and properties of CNT filled with metallic compounds can be found in ref. 123 and 124.

3. 1D nanoreactor: *in situ* transformation inside CNTs

The extreme confinement formed by ''tight'' crystalline walls and the high 1D aspect ratio make carbon nanotubes an attractive ''nanoreactor'' system for inducing and controlling chemical reactions on the nanoscale. Examples include the quasi-1D polymerization of phosphorus, sulfur, and polythiophene compounds,[125–128] and the synthesis of carbon nanoribbons from precursors such as fullerenes.[129–131] In this section, we focus our review on the very special 1D carbon chains known as carbynes and the in-tube synthesis of nanotubes and graphene nanoribbons.[132,133]

3.1. Confined carbyne chain synthesized inside CNTs

Studies on fullerenes (0D), CNTs (1D), graphene and other newly-discovered low-dimensional materials (2D) have evolved the investigation of nanomaterials toward lower dimensions. Even though CNTs behave as 1D materials, the ultimate goal in 1D would be true atomic chains. As an example of true 1D atomic chains, synthesis and Raman spectra of confined carbyne (CCs) are briefly discussed in the following.

3.1.1. Synthesis of confined carbyne. Short carbon chains in the forms of polyyne or cumulene have been studied for more than a century. However, their length is still limited to a few nanometers, *e.g.*, the longest polyyne consists of of from 44 and 48 to 52 and 68.[1–4] In addition to using carbon nanotubes as hosts for different molecules, *e.g.*, C_{60} and C_{70},[87] their hollow inner space was also used for the synthesis of novel 1D materials including sulfur chains, CsI chains, and linear carbon chains (LCCs, including CC).[126,134,135] The longest LCCs grown inside double-walled carbon nanotubes (DWCNTs) contain more than 6000 carbon atoms, confirmed to be CC,[135] which is more than two orders of magnitude longer than the longest polyyne previously reported.

LCCs can be synthesized mainly by two methods: arc-discharge and heat treatment. The first LCCs were obtained during the synthesis of multi-walled carbon nanotubes (MWCNTs) by arc-discharge. Normally, the arc-discharge between a graphite anode and a graphite cathode, in a chamber filled with inert gas, results in the obtention of highly-crystallized MWCNTs with ultra-thin inner tubes.[136] By introducing a specific ratio of hydrogen gas into the reaction chamber, LCCs were synthesized inside the innermost tubes, forming a hybrid structure of LCCs@MWCNTs[137] (Fig. 11). The carbon chains can be terminated with the hemisphere of the nanotube or the wall of the surrounding nanotube when the innermost tube is too large. Later on, the optimization of the synthesis of the LCCs@MWCNTs relied on various gases and current between the electrodes to control the temperature and the temperature gradient around the electrodes. For example, the preparation of LCCs@MWCNTs by the arc-discharge method has been tried in different inert gases or mixed with hydrogen, or even in liquid nitrogen.[138–140] Cooling and magnetic fields can be used to change the temperature gradient

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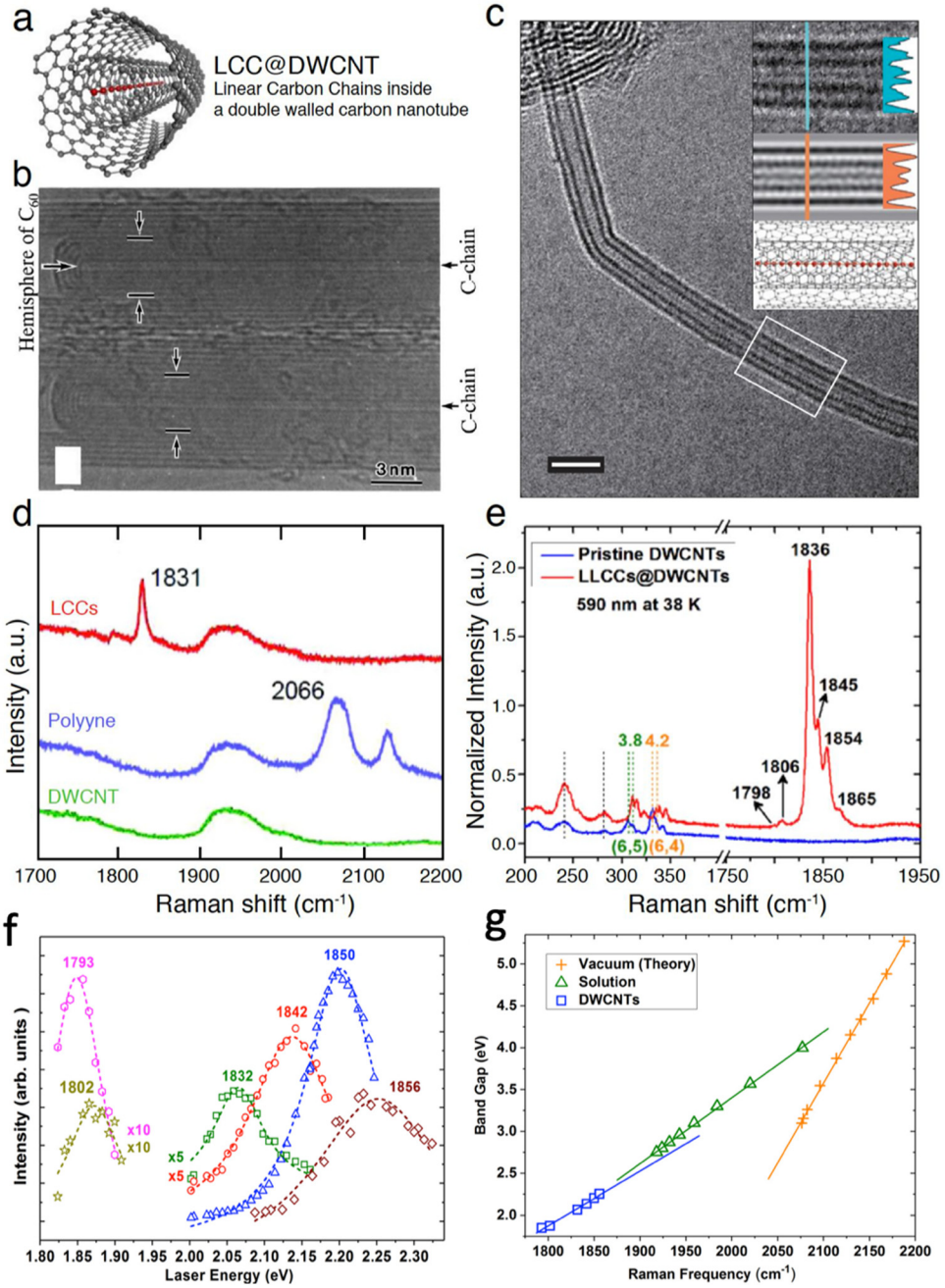

Fig. 11 (a) Illustration of a linear carbon chain inside a double walled carbon nanotube. Adpated from Zhang *et al.*[142] HRTEM images of (b) a LCC inside a MWCNTs and (c) a CC inside a DWCNTs (scale bar 2 nm). The figures are adapted from the ref. 135 and 137. Raman spectra of (d) DWCNTs (green), polyyne (blue), and LCCs (red); (e) DWCNTs (blue) and LCCs@DWCNTs (red). (f) Raman intensity as a function of laser energy for the CCs at different Raman frequencies. (g) Optical energy gap as a function of Raman frequency for the polyynes and CCs in different environments. The figures are adapted from the ref. 143 and 144.

around the electrodes, thus improving the synthesis of LCCs@MWCNTs.[141]

Heat treatment is commonly used for the purification of the CNTs, and it was found that short LCCs consisting of several carbon atoms could be formed between the inner and outer walls of DWCNTs after annealing the sample in inert gas.[145] At first, it was believed that CCs were synthesized inside DWCNTs during their annealing at 1550 °C, but this hypothesis was disputable due to a lack of, direct observations.[146,147] Recently, CCs have been prepared by annealing in high vacuum using DWCNTs made from chemical vapor deposition.[135] The annealing in high vacuum of DWCNTs made from chemical vapor deposition resulted in a much-improved CCs yield, rendering it possible to observe CCs directly using HRTEM, undoubtedly confirming the hybrid structure of CC@DWCNT (Fig. 11c). Compared to the Raman spectrum of the bulk sample, the

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coincided Raman spectrum from an individual CC@DWCNT also rules out the possibility of the carbon chains between the walls of the DWCNTs. Therefore, the status of the CCs has been finally solved. In addition, other than annealing above 1400 °C, heat treatment of filled polyynes inside DWCNTs at lower temperature below 1000 °C can combine the polyynes into long LCCs.[143]

The heat treatment can also be realized by methods other than furnace-based heating, for example, laser heating or field emission. Recently, CCs and also inner carbon nanotubes can be grown inside SWCNTs under laser annealing in vacuum by using lasers with power of tens of mW.[148,149] In addition, CCs were formed during the testing of field emission when using SWCNTs as a cathode material.[150]

Although the synthesis of CCs has been successfully performed, precisely controlling the length of the CCs is still a big challenge. Furthermore, the yield should be further improved, especially for long chains.[151,152] An interesting perspective is the synthesis of isotopically-labelled confined carbyne through the encapsulation of a isotopically-labelled liquid precursor. As shown in Fig. 11d–f, isotope CCs with around 11.9% of ^{13}C labelling were achieved using 13C-methanol as precursor, enabling the possibility to engineer the properties of the encapsulated CCs. Later, ^{13}C labelling ratio was increased to be 28.8% by using ^{13}C–C_{60} as precursor.[153]

3.1.2. Raman spectra of confined carbyne. In general, the longer the carbon chain, the lower the Raman frequence. The bond length alternation (BLA) between the single carbon–carbon bond and the triple bond decreases with the increased carbon chain length, until the point where the length does no longer define the BLA, such as in carbyne.[154] As shown in Fig. 11d, the Raman spectral mode of polyynes usually locates above 2000 cm^{-1}.[143] For confined CCs inside CNTs, which have longer length, compared to the polyynes, the characteristic Raman mode of CCs with multi-components appears at a lower frequency, normally between 1700 and 1900 cm^{-1} (Fig. 11e). When the polyynes or LCCs in general are confined inside the CNTs, their Raman modes shift to lower frequency due to the interaction between the polyynes/CCs and the CNTs.[155] A TERS study on individual CCs (see Section 5.3 for details) showed that for Raman frequencies between 1770 cm^{-1} and 1835 cm^{-1}, the bond-length alteration of CCs is determined by the interaction with the host nanotubes, while no dependence on the chain length was observed, indicative of true carbyne.[156]

As demonstrated in Fig. 11f, the resonance Raman profiles of the CCs at different frequencies reveal the energy gap of the CCs. The energy gaps of polyynes in gas phase and in solvents, as well as the energy gaps of confined CCs inside DWCNTs, obtained by absorption or resonance Raman spectroscopies, are summarized in Fig. 11g. In general, the longer the chain, the lower the band gap, until true carbyne is reached. The energy gap is also closely related to the environment and shift linearly with Raman frequency, because both the energy gap and the Raman frequency of CCs depend on the bond length alternation between C–C single and triple bonds. Thus, the energy gap of any types of LCCs can be estimated quite accurately by the frequency of the Raman mode.[144] When investigating the resonance Raman profile beyond the energy gap of the CCs, as shown in Fig. 12a–c, the resonance Raman profile each time shows resonances corresponding to two electronic transitions and two vibronic series of the second optical transition.[157] A resonant Raman study of an individual confined CC revealed an excited sate lifetime of around 5 fs,[158] whereas time-resolved resonance Raman scattering revealed that the exciton relaxation dynamics of the CCs occurred on a hundred of picoseconds timescale.[159]

3.2. Nanotubes and endohedral nanoribbons

Different from the gapless graphene, graphene nanoribbons (GNRs), as another 1D carbon nanomaterials, get a tunable band gap due to the quantum confinement in the lateral direction. Such band gap is closely related to the edge structure and the width of the GNRs.[160] Especially, GNRs with an armchair structure of their edges (AGNRs) present a band gap inversely proportional to the width of the GNRs, as shown in Fig. 13b, which can be classified into $n = 3p + 1$, $3p$, and $3p - 1$ species in order of reduced band gap, where n is the number of dimer lines across the ribbon width as defined in Fig. 13a and p is an integer.

Several methods including cutting the graphite/graphene, unzipping the CNTs, and chemical vapor deposition have been developed to obtain GNRs with sufficiently narrow widths.[161–163] However, up to now only on-surface synthesis (polymerizing the precursor molecules) and confined synthesis (reacting the molecules inside CNTs) result in both width and edge controlled GNRs.[164,165] Although the on-surface synthesis can prepare the GNRs with various edge structure and width, it is still limited by complexities regarding the design of the precursor molecule and most importantly the mass production. In contrast, confined synthesis not only enables the same polymerization inside CNTs as the on-surface synthesis but also allows to decompose and react the filled molecules into GNRs with their width limited by the diameter of the CNTs.[166,167]

Confined synthesis of GNRs can be realized by polymerizing the polyaromatic hydrocarbons similar to the on-surface synthesis. As illustrating in Fig. 13c, coronene or perylene were used as precursor molecules, which were polymerized into GNRs *via* annealing. However, from the Raman spectra taken on the annealed samples (Fig. 13d), it is clear to see that other types of GNRs exist in the sample except the expected ones. In principle, more selectively-precise synthesis could be achieved by applying those specifically designed brominated molecules, which are also used in the on-surface synthesis. Sulfur-terminated GNRs were also prepared by using functionalized fullerenes as precursors,[168,169] but control over the edge structure could not be obtained (Fig. 13e). By contrast, Ferrocene molecules used as precursors were shown to lead to GNRs with determined width and edge structure,[170] (Fig. 13e). Also prefilling CNTs with the solvent trichlorobenzene and subsequent thermal conversion resulted in the formation of specifically 6-AGNRs, which can only be explained by the CNT diameter

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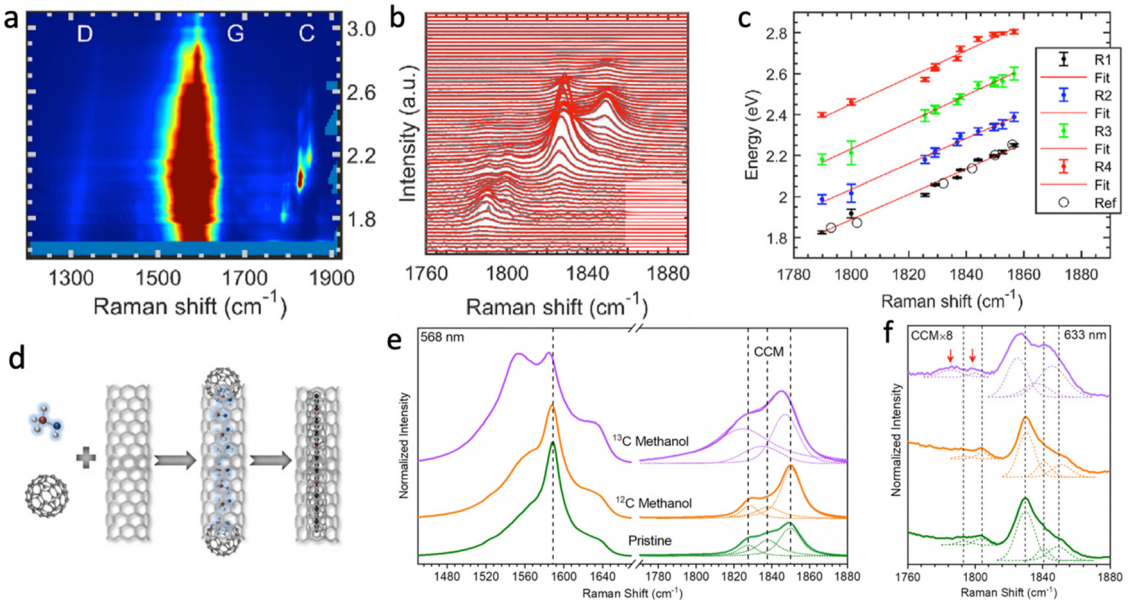

Fig. 12 (a) Resonance Raman intensity map (red being the most intense color) of CC@DWCNTs in the range of the G-band of the CNTs (1600 cm^{-1}) and the Raman modes of the encapsulated CCs (1700–1900 cm^{-1}), adapted from ref. 157. (b) Raman spectra (black) and corresponding fits (red) zoomed in on the CC modes as a function of laser excitation energy, adapted from ref. 157. (c) Optical energy gap (black filled circles) and additional resonances (blue, green and red) of the CCs as a function of Raman frequency and corresponding best linear fits through the data points (red lines). Figures are adapted from ref. 157. Additional data points (black open circles) are obtained from ref. 144. (d) The Scheme of the isotopic labelling of CC. The Raman spectra and peak analysis of annealed pristine, ^{12}C methanol, and ^{13}C-labelled methanol filled SWCNTs excited by (e) 568 and (f) 633 nm lasers. The figures are adapted from the ref. 152 and 157.

controlling the formation of such GNRs.[171] More recently, similar thermal conversions of trichlorobenzene inside BNNTs resulting in the formation AGNRs in BNNTs.[172] As shown in Fig. 13f, Raman features when excited with a laser wavelength of 633 nm belong to the n = 6 arm-chair GNRs. while when excited by a laser with wavelength of 568 nm, n = 7 arm-chair GNRs were observed in the sample.[173] In addition, the high intensity of the Raman modes of the GNRs compared to that of the CNTs points out the high the high yield of the grown AGNRs. Indeed, wavelength-dependent Raman studies of similarly synthesized AGNRs could correlate the electronic resonances with the Raman vibrational frequencies of 5-AGNR, 6-AGNR and 7-AGNRs synthesized within the CNTs.[165,170] Although progresses have been made, the precise synthesis *via* confinement in CNTs still needs further investigations, especially for the other nonexplored GNRs, *e.g.*, n = 4 arm-chair GNRs, zigzag GNRs, as well as heterodoped and isotopic GNRs. There is plenty of room in the future for both synthesis and property studies with application in electronic transport or photonics, for example.[174]

3.3. Inorganic compounds transformation inside CNTs

Inorganic compounds may not be possible to insert directly within nanotubes due to their incompatibility with the different filling routes available in the literature.[89] This incompatibility may be due a poor solubility of the compound in solution or a too high melting point. Other limitations may occur in case of a limited thermal stability (decomposition before melting), or a possible direct chemical reaction between the nanotube itself and the filling material: this may be the case for example while filling CNT with oxides (carboreduction reaction).[91]

The best strategy is to identify a suitable precursor that can first be placed inside the nanotubes before being transformed into the final desired compound. This transformation is usually performed by heat treatment, because this can easily be applied to a bulk sample and does not require any chemical to be able to get inside the nanotubes – which is challenging for nanotubes with a narrow inner channel (less than 1–2 nm). The chosen heating atmosphere depends on both the stability of the nanotubes (CNT are generally air-sensitive above 250 °C, while boron nitride nanotubes are more stable in air). For example, air (or oxygen diluted in an inert gas such as N_2 or Ar) is used to generate oxides while hydrogen is used to reduce compounds to the metallic state. Many other kinds of transformations (thermal degradation or decomposition) may occur while heating in inert atmosphere, depending on the chemistry of the compound of interest. We review here examples of such *in situ* transformations and discuss the limitations of this strategy.

Among the first examples of *in situ* transformations after filling are from the Green's group, when they initially filled CNT with different nitrates (Fe, Co, Ni, U and Cu/Au, $CoFe_2O_4$, Re) before annealing in He atmosphere to generate the corresponding oxide and then in H_2 atmosphere to reduce oxides to the corresponding metals.[96,99,108,175–177] The temperature reduction ranges usually between 200° to 600° depending on the filled compound.

Reduction may also be performed using chemicals such as alcohols (methanol) or glycols (ethylene glycol, propylene

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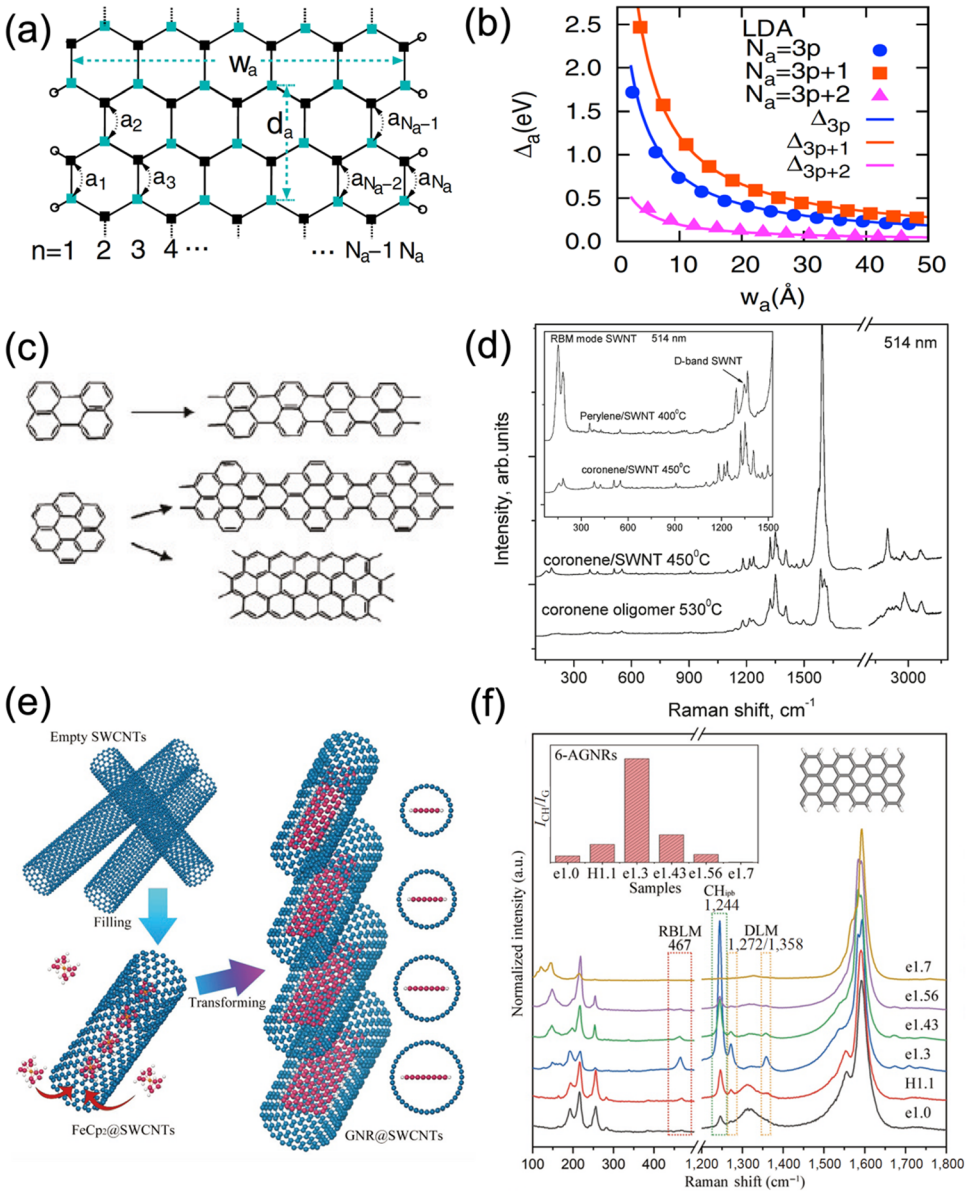

Fig. 13 (a) Schematic of a n = 11 armchair GNR with width of w_a and 1D unit cell distance of d_a passivated by hydrogen atoms (empty circles). (b) The band gaps of n = Na armchair GNRs as a function of width (w_a) obtained from first-principles calculations. (a) and (b) adapted with permission from Son *et al.*, *Phys. Rev. Lett.* 2006, **97**(21), 216803, Copyright 2006 American Physical Society.[160] (c) Suggested structures of GNRs formed using coronene or perylene precursors. (d) Raman spectra of oligomer obtained by annealing the coronene power and the GNRs@CNTs transformed from the coronene and perylene encapsulated CNTs. (c) and (d) adapted with permission from Chamberlain *et al.*, *ACS Nano* 2012, **6**(5), 3943–3953, Copyright 2012 American Chemical Society.[168] (e) Schematics of ferrocene molecules filled inside CNTs and transformed into GNRs with different widths inside CNTs with various diameters. (f) Raman spectra of the 6-AGNRs inside CNTs with marked diameters (the numbers after the alphabets e or H). The insets show the model of a n = 6 armchair GNR and the relative intensity of the C–H in-plane-bending mode to the G-mode in different samples. The figures are adapted from the references. Adapted with permission from ref. 170.

glycol) as proposed by the group of Rao.[178] The question of the completion of such reactions was not discussed for a long time, people usually assuming a total transformation. This is the case especially for the reduction with H_2 gas, which is very small and assumed to easily diffuse into most matrices To prepare nano-magnets confined inside double-wall CNT, different iodides of magnetic metals (Fe, Co, Ni) were reduced. However, the reduction of these iodides by hydrogen was found to not be as fast as expected. This of course depends on the temperature of the heat treatment and a compromise must be found between heating at a lower temperature for a longer time or at a higher temperature for a shorter time. If the temperature is too high, the reduced metal may be released from the nanotube due to its poor wettability of the CNT inner channel, leading to metal nanoparticles decorating the outer walls instead of the desired metal nanoparticles/nanowires inside the nanotubes. The reduced filling material may also be pushed out the nanotubes in case of gas release during the reaction (for

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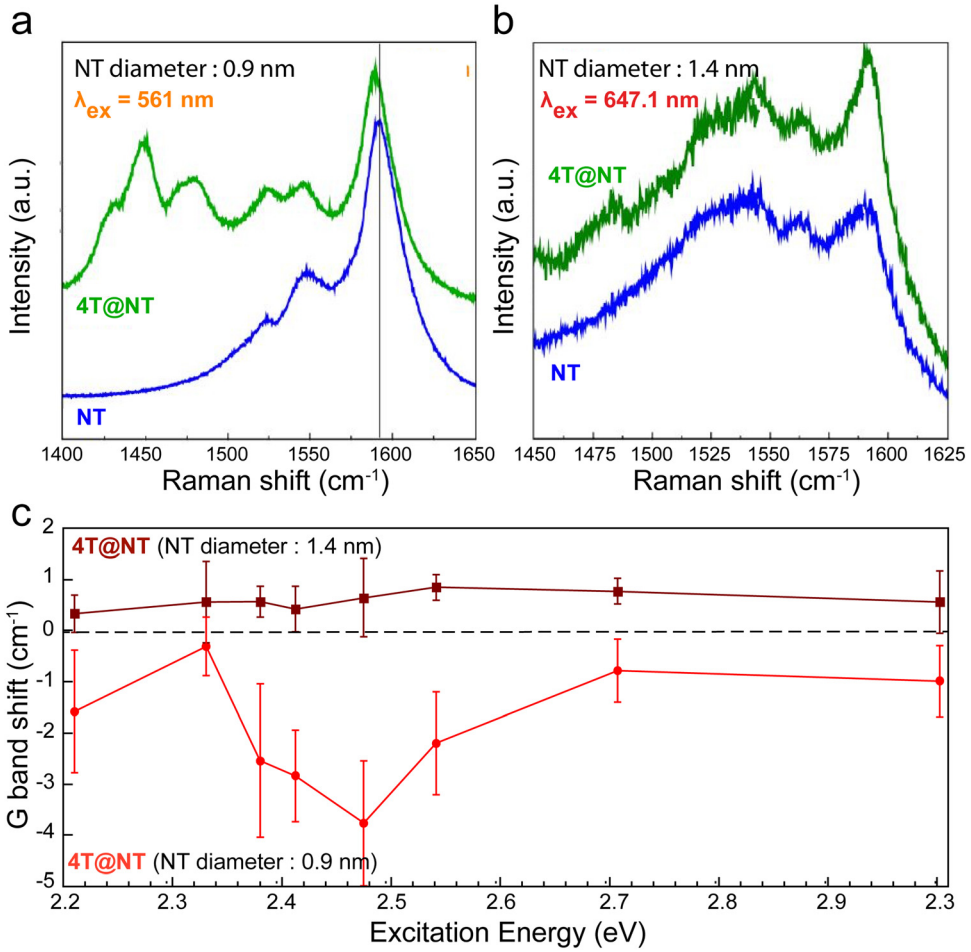

Fig. 14 Raman spectra of pristine (black curve) and hybrid (blue curve) nanotubes having a diameter of 0.9 nm (a) or 1.4 nm (b) and excited close to their resonance excitation wavelengths, 561 nm and 647.1 nm, respectively. Hybrid spectra are offset for clarity. (c) Raman G^+-band shifts of semiconducting NT after 4T confinement as a function of the excitation wavelength for 4T@NT with small diameters (ϕ = 0.9 nm, red circle and curve) and 4T@NT (ϕ = 1.4 nm, wine square and curve). Adapted with permission from ref. 187.

example water vapour during the reduction of an oxide). Heating in hydrogen often leads to only partial transformation of the precursor when the inner diameter is very narrow, with a gradient of composition between the fully reduced metal to the initial metal iodide, progressing from the opening by which H_2 gas gets into the nanotube towards the inside the nanotube.[109] The thermal degradation of Nickel iodide, in particular, could be used to fill nanotubes with Iodine adopting different structures depending on the inner diameter of the host nanotube.

Although organometallics are a popular source of precursors for the direct synthesis of metal-filled nanotubes, there are few examples in the literature of filling with such molecules and further transformation into other compounds (metal oxides, metals). In particular, acetylacetonates have been used to fill CNTs with Nickel after thermal decomposition above 500 °C.[179]

Solvothermal treatments have also been proposed to fill CNTs with CdS nanoparticles. In this case, the CNT were first filled with Cadmium acetate before heating in an autoclave at 180 °C in DMSO (to provide sulphur). This strategy is interesting because the high pressure available in the autoclave must favour the capillary filling of the nanotubes and probably increases the filling rate.[101]

Finally, it is important to mention the possible incidental *in situ* rearrangement or decomposition of precursors within the TEM during the characterization of filled CNT. Among the first examples were the *in situ* rearrangement of SnO nanocrystals encapsulated within CNTs[180] and the decomposition of $ZrCl_4$-filled CNT to form $ZrCl_x$ clusters. The decomposition of metal iodides and subsequent release of iodine was also observed in similar conditions.[109]

4. Tuning CNT optical properties using molecular assemblies

The proximity of the electronic surface of the nanotubes and the molecules, often in the range of the van der Waals distance, creates a favorable soil to physical interactions and coupling. A parallel can be done with 2D heterostructures based on nanomaterials such as graphene and other semiconducting layered

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materials. Among the large panorama of possible interactions between 0D/1D and 1D/1D nanobjects, this section focuses on charge and energy transfers to tune the nanotubes' host properties. Dye molecules confinement into SWCNTs is a non-covalent functionalization that leads to new hybrid nanosystems with a new electronic structures that must be properly understood.

4.1. Charge transfers and doping

The confinement of organic molecules, which are then in close proximity with the inner wall of the host CNT, enables different physical interactions such as $\pi\cdots\pi$ or $CH\cdots\pi$ interactions, charge transfer, or hybridization. For instance, the encapsulation of molecules that have either a high electron affinity or a small ionization energy can lead to a charge transfer (CT) in between the two sub-systems.

Here, we review how to probe such phenomena in acceptor/donor molecules inside SWCNT using Raman and photoluminescence spectroscopies. For example, the Raman spectra of quaterthiophene molecules (4T) encapsulated in semiconducting (0.9 nm diameter) and metallic (1.4 nm diameter) SWCNTs both exhibit the G^- and G^+ bands of the SWCNTs (between 1500 and 1595 cm^{-1})[181–186] and the C=C stretching vibrations of the confined 4T molecules at 1440–1530 cm^{-1} (Fig. 14a and b). However, a clear shift of a few cm^{-1} of the G-band is observed for the semiconducting SWCNTs. By comparing the relative G^+ band shifts of the encapsulated SWCNT with respect to non-encapsulated semiconducting SWCNT as a function of the excitation wavelengths for the two distinctive SWCNT diameters, a noteworthy and wavelength dependent downshift of the G-band frequency for the 0.9 nm diameter SWCNT is observed (Fig. 14c).

Because semiconducting nanotubes having smaller diameters do not undergo electron–phonon coupling,[188] the G^+ band downshift comes from a weakening of the force constant, due to electron transfer from the confined molecules to the NT walls. The magnitude of this shift can be derived from the following relationship adapted from ref. 189 or 190:

$$\Delta\omega_{static}\,(cm^{-1}) = -804 \times \rho_c - 5126 \times {\rho_c}^2 - 176\,790 \times {\rho_c}^3 - 1657 \times {\rho_c}^{3/2} \quad (1)$$

where ρ_c is the number of electrons transferred per carbon atom of the nanotube.[190]

According to eqn (1), shifts in the range from -0.28 cm^{-1} to -1.6 cm^{-1} would correspond to $\rho_c \sim (1.1 \pm 0.7) \times 10^{-3}$.

Considering that the 4T molecule is around 2 nm in length and calculating the number of carbon atoms in a NT of the same length with a diameter of 9 Å, we then calculate that ρ_{4T}, the number of electron withdrawn from one 4T molecule, is equal to 0.25 ± 0.15, assuming that only one 4T molecule row is inserted.[13] Finally, adapting the equation below from ref. 191 the Fermi level shift can be estimated by the following equation:

$$\Delta E_F(eV) = 51.77 \times \rho_c \times d\,(nm) \quad (2)$$

where d is the CNTs diameter. Thus, $\Delta E_F = 0.05 \pm 0.03$ eV for 4T@NT09.

Furthermore, an important amplification of the downshift for excitation energies around 2.5 eV is evidenced. As the resonance profile of the confined 4T molecules overlaps this energy range, we can reasonably assume an extra photo-induced charge transfer, leading to an additional contribution in the G-band downshift (from -2.2 cm^{-1} to -3.7 cm^{-1}), corresponding to $\rho_c \sim (3.2 \pm 0.8) \times 10^{-3}$ and $\rho_{4T} = 0.73 \pm 0.17$. Here, $\Delta E_F = 0.15 \pm 0.04$ eV for 4T@NT09 at the resonance.[187]

For larger diameter tubes (d = 1.4 nm), the persistent upshift indicates that dynamic effects (electron–phonon coupling) govern the G-band behaviour. Using calculations on graphene,[189] on individual semiconducting carbon nanotubes[188] or observations realized on an ensemble of carbon nanotubes,[182] the amplitude of the upshifts observed on Fig. 14c, around +0.5 cm^{-1}, can be induced by a charge transfer of $\rho_c \sim (0.45 \pm 0.15) \times 10^{-3}$ and $\rho_{4T} = 0.05 \pm 0.005$ (assuming that three columns of 4T molecules can stack in the SWCNT section of such a diameter[13]). The Fermi level shift is then estimated as $\Delta E_F = 0.035 \pm 0.005$ eV in this case.

The shifts of Fermi level and charge transfer values presented here are consistent with the measurements performed on an ensemble of SWCNTs.[182] Furthermore a decrease by a factor of 2 of the coupling parameter $-1/q$, coupled to Fermi level shift $\Delta E_F \sim 0.2$ eV, a has been already observed on individual SWCNT of comparable diameter.

Charge transfers and Fermi level shift in mol@SWCNTs also may strongly impact the excitonic processes and resonance profile and hence, modulate the light emission properties of the overall nanohybrids.

The photoluminescence excitation (PLE) maps of unfilled SWCNTs and of SWCNTs filled with donor quaterthiophene molecules (4T@SWCNT) are shown in Fig. 15a and b. The reference PLE map (Fig. 15a) indicates the presence of (6,5) and (9,4) SWCNTs with diameters of 0.75 nm and 0.9 nm, respectively. The 4T molecules do not fit in the (6,5) SWCNTs while there is enough space in the (9,4) SWCNTs to form one row of encapsulated 4T molecules. Interestingly, the (6,5) subpopulation shows an PL intensity ratio ($R = I_{filled\text{-}SWCNT}/I_{unfilled\text{-}SWCNT}$) close to unity, while a $R \sim 5.5$ is observed for the (9,4) SWCNTs.

This PL increase indicates that the encapsulation of 4T molecules provide a way to modulate the emission properties of the smaller diameter (9,4) SWCNTs, which has also been seen for ferrocene or alkane encapsulation.[81,192] Such enhancement of the intensity could be attributed to the weak electron transfer from the encapsulated molecules to the SWCNT, which compensate the intrinsic p-type doping of pristine nanotubes, induced, for example by curvature[193] or defects.[194] Thus, the charge transfer moves back the Fermi level closer to its expected position for "defect free" tubes,[192] increasing PL efficiency as the volume available for molecules in the SWCNT increases.

In contrast, encapsulating an electron acceptor chromophore such as TCNQ leads to a very important decrease of the

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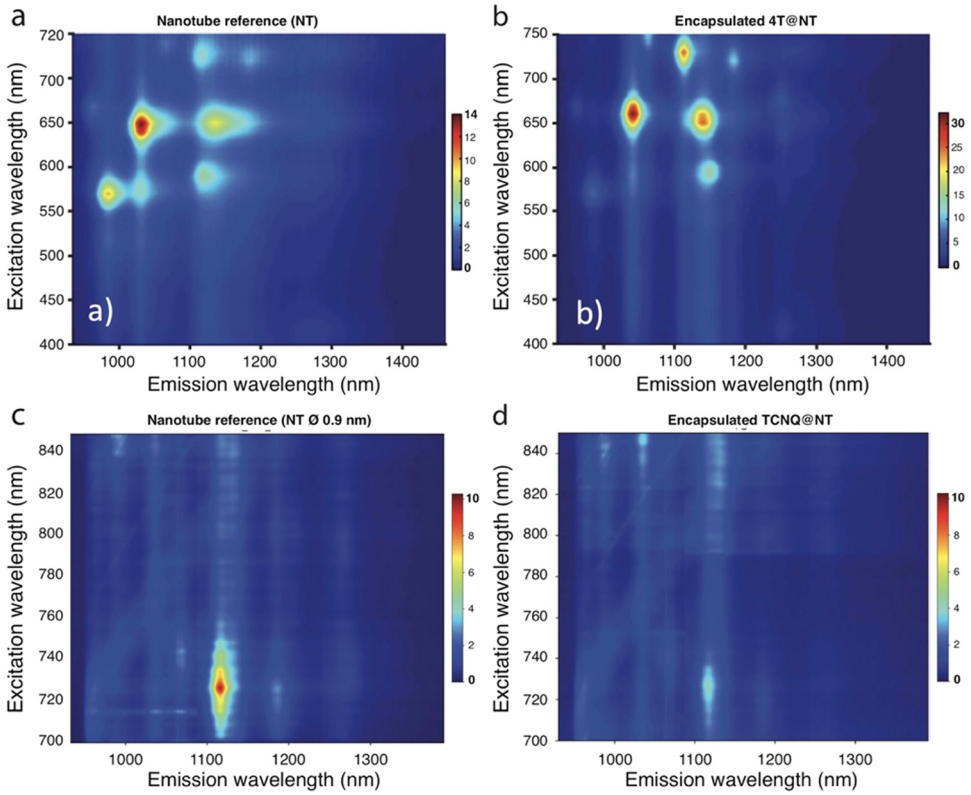

Fig. 15 PLE map: PL signal intensity *versus* emission (*x*-axis) and excitation (*y*-axis) of NT09 (a), 4T@NT09 (b) NT09 (c) and TCNQ@NT09 (d), Note that here the intensities have been normalized by adjusting the absorption intensity on the E_{22} transitions of both samples.

PL intensity (Fig. 15c and d). This behavior is consistent with the above assumption but, in this case, the Fermi Level is even more downshifted, which allows new de-excitation processes and thus reduces the PL efficiency. These assumptions were confirmed by optical studies on individual hybrids with confined electron donor (4T, TMPD) and acceptor (TCNQ, F4TCNQ) molecules.[195]

4.2. Energy transfers, sensitization and quenching

As will be described in more detail in Part 6, the endo- and exohedral functionalisation of SWCNTs with chromophores is essential to enhance or extend the light absorption of SWCNTs in other wavelength ranges and improve the photovoltaic device performance, for example.[196,197] To achieve this photosensitization through energy transfers engineering (Fig. 16), which is distinct from charge transfer, several approaches can be found in the literature and can be classified into 3 main strategies: (i) covalent functionalisation and non-covalent (ii) endo- and (iii) exohedral functionalisation.

The covalent functionalisation of SWCNTs can be achieved through a variety of different chemical reaction pathways and choice of molecules. Covalent functionalisation leads to long-term bonding, which is advantageous for many applications. However, the attachment of functional groups using standard covalent strategies results in the conversion of the underlying sp^2 carbon atom into its sp^3 state, lowering the conjugation and, hence, disturbing the long-range transport properties. Moreover, it is essential to design suitable linkers to ensure the optimization of the distance and orientation of the chromophore with respect to the SWCNTs for efficient energy transfers.[198,199] To limit the number of added defects to the SWCNT walls, DNA wrapping was used as a protective coating to then selectively attach CdSe/ZnS core shell quantum dots (QDs) to the carboxylic end-defects of SWCNTs, by using an amidation reaction.[200] The steady-state PL intensity of the quantum dots was quenched by a factor of 2 upon their attachment to the SWCNTs, indicating electronic coupling between the two. This electronic coupling was further evidenced by single-QD-SWCNT emission studies, where QDs-SWCNTs displayed reduced blinking with respect to non-functionalised QDs. A different covalent strategy that preserves the π-conjugation of the SWCNT walls was also developed, allowing for high-density functionalisation (up to 1 out of 25 carbon atoms),[201,202] and showing great promise for the design of stable and efficient energy-transfer nanohybrids. In this approach, a highly reactive azidodichloro-triazine species is created *in situ* to bridge onto two nearby-lying carbon atoms of the nanotubes. This strategy results in the regeneration of the π-conjugation of the whole system and the full electronic integration of the triazine group into the conjugated network, without creating trap/scattering centers for the nanotubes' excitons. Because the chlorine atoms in the azidodichloro-triazine molecule can be easily substituted, this molecule is an ideal building block for further non-perturbative

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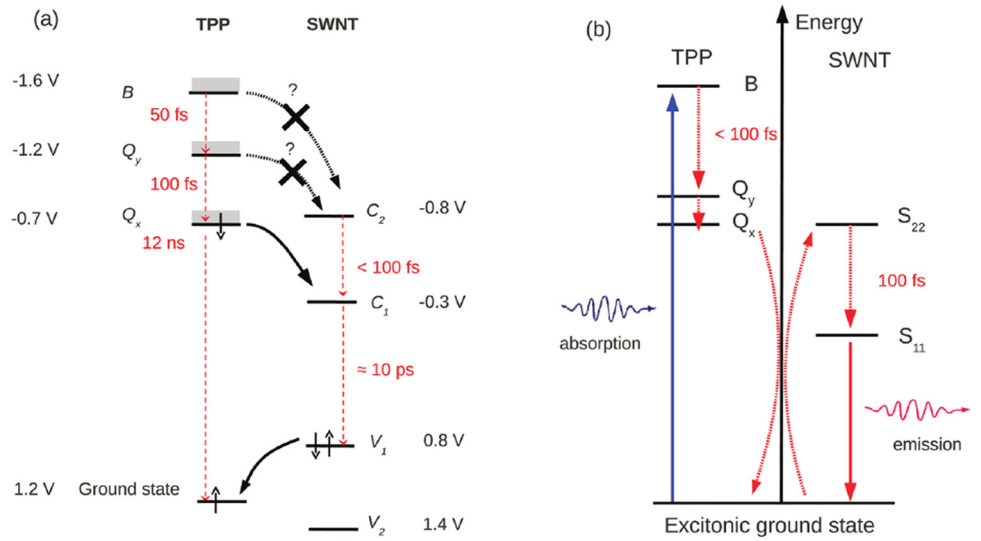

Fig. 16 Scheme use to show possible energy transfer mechanisms (a) in an electronic description through double electron transfer, (b) in an excitonic description.

attachment. For example, the replacement of the chlorine atoms with thiol derivatives ensured the covalent attachment of gold nanoparticles onto the nanotube sidewall, enhancing their photoluminescence emission.[201,202] Such plasmon-assisted emission enhancement of CNTs could previously only be observed in non-covalent systems.[203,204]

The non-covalent functionalisation of SWCNTs by either π-stacking on the outer walls or by encapsulating chromophores inside the SWCNTs, has the advantage of not perturbing the π-conjugated structure of the SWCNTs while still providing a very close interaction between the chromophores and SWCNTs (of the order of the van der Waals radius). Although the encapsulation of chromophores for energy-transfer nanohybrids is beneficial since the SWCNT wall protects the photochemical and thermodynamic stability of the dyes in comparison to external interfacial adsorption,[205] the typical size of the chromophores prevents the use of small SWCNT diameters.[82] Nevertheless, endohedral functionalisation, providing a very close interaction between the encapsulated molecules and the SWCNTs, has been shown to provide almost always close to 100% energy transfer efficiencies.[12,14,23,205–211] Similar near unity EET yield have been reported for exohedral functionalization when the dye is directly π-stacked onto the nanotube side-wall, for example in the case of tetraphenyl porphyrin.[212] However, this is not always the case for exohedral approaches. Depending on the manner of the exohedral functionalisation, either through direct π-stacking with the SWCNT walls,[36,213–217] by using intermittent anchor molecules such as *e.g.* pyrene[218,219] or perylene[220] that provide the strong π-stacking with the SWCNT walls, or by attaching the chromophores to a polymer that wraps around the SWCNTs,[221] very different energy transfer efficiencies can be observed. In addition, a morphological mismatch between the relative orientation of the chromophore and the CNT leads to altered energy-transfer efficiency, even if their relative separation is within the Förster radius.[222,223] One novel strategy is to first coat the SWCNTs with porphyrin molecules through π-stacking, and subsequently polymerising of these porphyrin molecules to stabilize the exohedral functionalisation.[38]

The attachment configuration strongly affects not only the strength and origin of the interaction between molecules and the nanotubes but also the chemo-physical nature of the building units themselves. For example, the molecular photochromic switch spiropyran profoundly changes its chemical behaviour and optical response depending on how it is complexated with carbon nanotubes. In its isolated, pristine form, when excited with ultraviolet light, the colourless spiropyran undergoes a conformational change, turning into merocyanine.[224] The merocyanine emission is not quenched by the presence of the nanotubes if non-covalently attached onto the tubes with a relative separation of *ca.* 1 nm.[225,226] However, it is completely quenched when π-stacked onto the nanotube sidewalls.[227] When spiropyran gets covalently attached onto the nanotubes *via* the triazine-based covalent approach,[201,202] the uninterrupted conjugation between the molecule and the nanotubes drastically changes the response of the system upon UV light illumination: the π-electron released upon spiropyran-to-merocyanine isomerization is not confined into the merocyanine anymore and is released onto the nanotubes. Instead of emitting, then, the molecule yields light-triggered doping of the carbon nanotubes, affecting the position of their Raman G-band and the efficiency of the photoluminescence emission.[201,202] The emission of the carbon nanotube is completely quenched at the single nanotube level. This light-triggered, reversible modulation of the nanotube's emission can be exploited to achieve super-resolution microscopy in the near-infrared, where biological tissues are transparent.[228]

Different experimental techniques can be employed to investigate the mechanism and efficiency of the energy transfer between chromophores and SWCNTs. Typically, a combination

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of different spectroscopic techniques is employed to distinguish energy- from charge-transfer, unveil the physical mechanism of the energy transfer and, most importantly, measure the efficiency of the transfer.

Raman spectroscopy is known as an efficient way to probe charge transfer combined with energy transfer in hybrid systems;[195,201,202,229] yet, the main technique to evidence energy transfer however is the wavelength-dependent fluorescence-excitation (PLE) spectroscopy, in which the emission of SWCNTs is probed as a function of their higher-order excitation bands. When energy transfer is occurring, the encapsulated or adsorbed molecules absorb the excitation light (at wavelengths that the SWCNTs do not preferentially absorb) and subsequently transfer their energy to the SWCNTs that then afterwards emit light, resulting in an additional band that can be observed in such PLE maps. This additional band occurs at the same emission wavelength of the SWCNTs but at the excitation wavelength of the dyes.[209,211,230,231] Representative PLE maps of such nanohybrids, in comparison with the non-functionalised starting materials, are presented in Fig. 17. Excitation energy transfer (EET) peaks in the PLE maps can be clearly distinguished, either for externally adsorbed porphyrins on (6,5) SWCNTs[212] (Fig. 17a and b) or for encapsulated squarylium molecules inside a broad-diameter SWCNT sample[209] (Fig. 17c and d).

To determine the energy transfer efficiency, a rough estimation can be obtained by comparing the PL counts of the excitation energy band with that of the SWCNTs. However, several unknown parameters are needed to make this estimation inaccurate, such as the absorption cross-section of the SWCNTs, the absorption cross-section of the dyes (that can be altered due to the stacking on or encapsulation inside the SWCNTs), and their relative concentrations.

Besides providing a direct evidence for EET between the dye and the nanotubes, a remarkable outcome of these PLE maps in the case of exohedral functionalization with porphyrins has been to provide an original means to benchmark the absorption cross-section of the different nanotube chiral species to the absorption cross-section of the porphyrin on their Soret band. Indeed, the ratio in intensity of the EET peaks and the intrinsic SWCNT emission peak can yield direct information on the absolute absorption cross-section of the SWCNTs, when the number of adsorbed porphyrin molecules and the energy transfer yield are known:[230]

In addition, when measuring the EET peaks for encapsulated molecules, a minimal encapsulation diameter can be estimated from the onset of the observation of this energy transfer in the PLE maps; for smaller diameters SWCNTs, no EET peaks are observed while they are observed for larger diameters SWCNTs.[209,210]

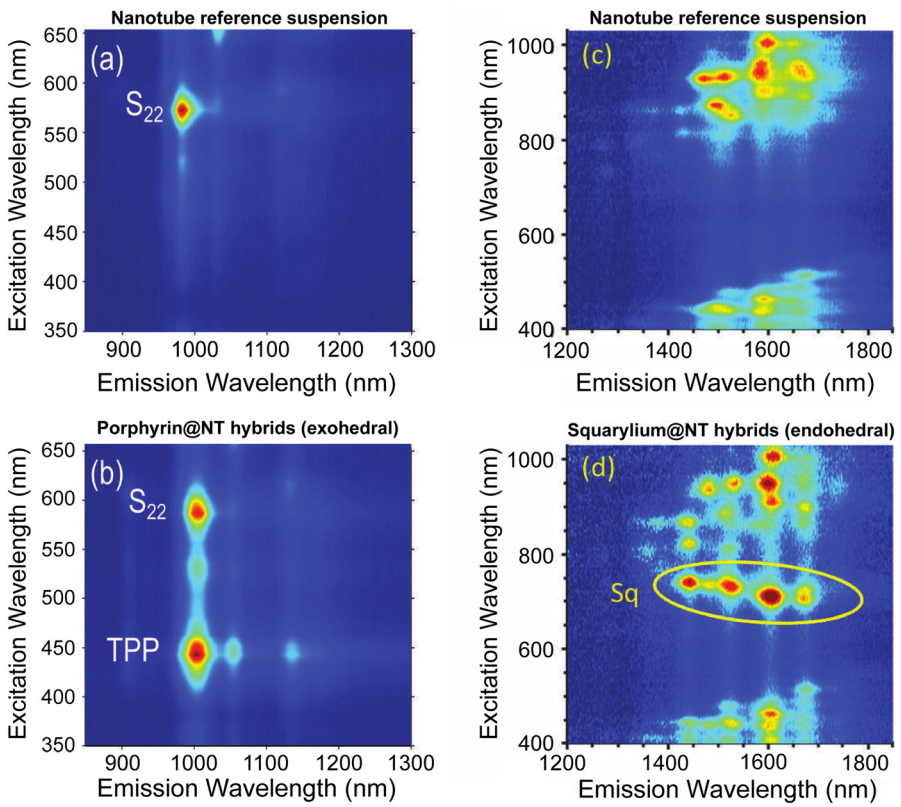

Fig. 17 Excitation Energy transfer observed through PLE spectroscopy of functionalised SWCNTs ((b) and (d)) by comparing the PLE maps with non-functionalised SWCNTs ((a) and (c)). Additional bands can be clearly observed, corresponding to specific absorption of the adsorbed (b) and encapsulated (d) dye molecules and emission from the SWCNTs. Adapted with permission from ref. 212 and 209.

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Nevertheless, it is readily seen in Fig. 17d that for the squarylium dye (SQ) encapsulated inside SWCNTs, the EET peaks corresponding to different SWCNT diameters (with different emission wavelengths) have their maximum resonance at different excitation wavelengths.[209] This peculiar behaviour was attributed to a different molecular stacking of the dye molecules inside SWCNTs of different diameters. This became even more clear after subsequent sorting of the SWCNT samples into single- or few-chirality distributions, allowing to directly access the change of the dye absorption after encapsulation in different chiralities (Fig. 18).[232] The diameter-dependence of the excitation energy of the encapsulated dye molecules can be understood by taking into account the geometric freedom available for the dye molecules within SWCNTs with different diameters [see also Part 2.1 on dyes aggregation inside nanotubes]. In small diameters, the molecules can only stack in a single-file arrangement, hence resulting in the formation of J-like aggregates and yielding a large red-shift with respect to the dye molecules freely suspended in solution. For larger diameters, the dyes can take a staggered or multi-file arrangement, yielding smaller or even larger red-shifts depending on the exact geometric arrangement of the dyes.[209]

If energy transfer takes place, the molecules' fluorescence will be reduced. Thus, if no other non-radiative decay pathways are created due to the interaction with the SWCNTs, the fluorescence quenching of the adsorbed/encapsulated molecules can give information on the energy transfer efficiency. Recent works have demonstrated a near-to-complete quenching of the PL efficiency of encapsulated/adsorbed molecules, with values ranging from 99.9%[210] up to 99.999%[23,209,212] It is now generally accepted that lower reported PL quenching values for encapsulated species,[12,14] originate from trace amounts of non-encapsulated molecules still present in the suspension, highlighting the importance of sample purification after encapsulation, as described in detail in the ESI† File. Indeed, even for SWCNTs encapsulated inside SWCNTs, (such as in DWCNTs), the PL is drastically quenched by several orders of magnitudes (PL quantum yield to be 10^{-6}),[233] and energy transfer was observed from inner to outer SWCNTs as a cause of this PL quenching.[234] By contrast, as discussed in Part 5.1, molecules encapsulated within boron-nitride nanotubes display preserved and tunable luminescence.

For exohedrally non-covalently adsorbed molecules, it is inherently more difficult to determine the PL quenching factors because non-adsorbed molecules in the suspension are hard to avoid. Thereby, only a lower limit for the PL quenching factor can be estimated. For SWCNTs functionalized with porphyrins and subsequently solubilized in surfactant, the lower limit of the PL quenching of the porphyrins was determined to be from 99.5% to 99.9% depending on the trace amounts of non-interacting (thus luminescent) porphyrins in surfactant micelles[212,217,235,236] (Fig. 19a).

More detailed information on the efficiency, kinetics and mechanism of the energy transfer comes from time-resolved absorption experiments, as PL quenching alone cannot distinguish between energy transfer or other non-radiative decay pathways.

For non-covalently-attached tetraphenyl porphyrins, an extremely fast energy transfer to the SWCNTs is observed, experimentally determined to be below 100fs (the detection limit), from transient absorption measurements (Fig. 19b). An enhanced bleaching of the nanotube excitonic transition when the hybrid is excited on the porphyrin transition is also observed simultaneously.[237,238] This bleaching behaviour shows that the nanotubes are photo-sensitized by the porphyrin and that the accelerated decay of the porphyrin is due to EET rather than to non-radiative quenching. This ultrafast energy transfer is understood in the framework of a Förster mechanism between a point donor and a one-dimensional acceptor separated by a distance of a few ångstroms only. Förster energy transfer is usually a resonant mechanism. Here, however, the significant coupling means that the TPP molecules can transfer their energy through their Q bands to the weak band-to-band absorption of the SWCNTs in between the excitonic

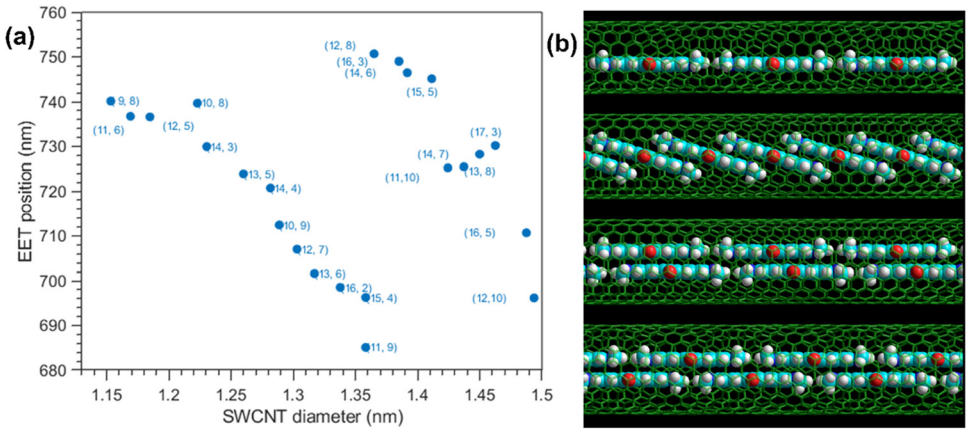

Fig. 18 (a) Position of the EET peak extracted from 2D spectral analysis of the PLE maps of different squarylium-filled and chirality-sorted SWCNT samples. The pink shaded area represents the minimal encapsulation diameter range while the purple area represents the diameter range after which a more complex behaviour is found due to multiple possible arrangement in larger diameter SWCNTs. (b) Qualitative models for possible stacking geometries of squarylium molecules resulting in different EET peak positions. Composed from ref. 209 and 232.

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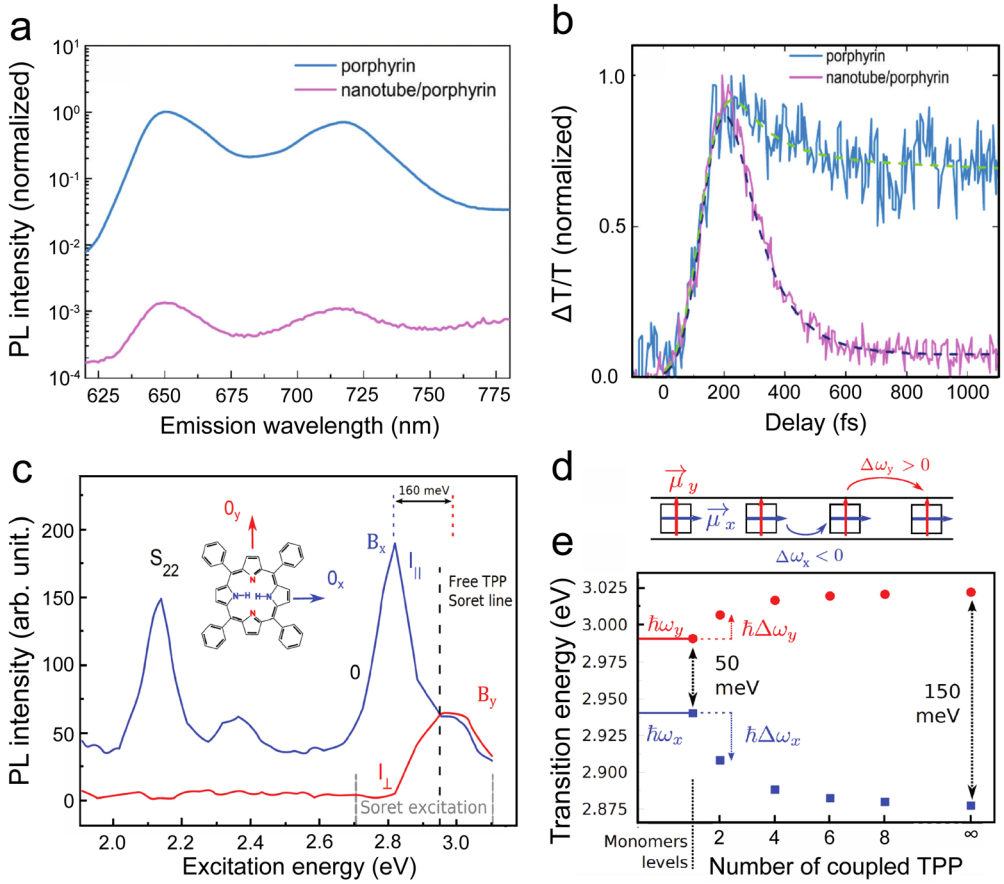

Fig. 19 (a) Quenching of the PL of porphyrins/SWCNT hybrids (red line) with respect to the non-interacting porphyrins (blue line) yielding a quenching factor of the order of 1000.[212] (b) Transient absorption measurements on the Soret transition of the porphyrin showing an extremely fast energy transfer from the porphyrins to the SWCNTs (red line together with exponential fit (doted line)); for comparison the decay of the porphyrin reference sample is shown in blue. (c) Polarized photoluminescence excitation spectrum. The blue line represents the luminescence spectrum for an excitation polarization parallel to that of the emission, whereas the red line corresponds to a cross polarization. The data show a splitted Soret band at 2.8 eV and 2.9 eV, which is interpreted in the framework of Davydov splitting for a layer of self-organized porphyrin molecules on the nanotube outer wall. (d) Schematic of an array of porphyrins at the surface of a carbon nanotube, with their h_x dipoles aligned to the nanotube axis; (e) evolution of the low energy $\hbar\omega_x$ and the high energy $\hbar\omega_y$ bands as a function of the number of porphyrin monomers in the aggregate. Adapted with permission from Delport *et al.*, *Nano Lett.* 2017, **17**(11), 6778–6782 Copyright 2017 American Chemical Society (ref. 37), and from ref. 212.

transitions.[176] Another interesting feature of this large coupling between the dyes and the SWCNT is the strong absorption anisotropy that the hybrids inherit from the nanotube due to the self-organization of the TPP array on the nanotube wall.[37,239] This absorption anisotropy results in peculiar polarized PLE spectra (Fig. 19c–e) and opens interesting perspective for the self-organization of non-covalently-attached molecules using carbon nanotubes as a template. In addition, the molecule/molecule interaction can be studied in a configuration partially controlled by the nanotube geometry.

Extremely fast and efficient energy transfer has also been observed for encapsulated molecules. For quaterrylene-encapsulated SWCNT hybrids, the EET time constants were found to depend on the specific nanotube diameter range of the samples in which the quaterrylene molecules were encapsulated. This difference was then interpreted to arise from different molecule-wall distances for different molecular arrangements.[12] For squarylium dyes encapsulated inside SWCNTs, transient absorption experiments (where the dye was pumped and the SWCNT absorption was probed) showed an instantaneous decrease of the SWCNT absorption of 15%, evidencing that EET took place within the 190fs time resolution of the experimental setup.[231,237] This estimation is however difficult to do in a multi-chirality sample, due to the strongly overlapping absorption peaks of SWCNTs. More recently, transient absorption experiments were performed on SWCNT samples with a narrow diameter distribution, so that both the dynamics of the encapsulated dyes and the SWCNTs could be characterized[209] (Fig. 20a). A clear ground-state bleach of the dye absorption and an enhanced bleaching of the SWCNT absorption in the S_{11} and S_{22} optical transitions of the SWCNTs was observed for the dye-filled sample in comparison to the reference sample. Probing the exciton creation and decay dynamics of the (10,8) SWCNT, this increase in intensity appears to be instantaneously on a time scale similar or shorter than the instrumental response (~200 fs, see ref. 209). By

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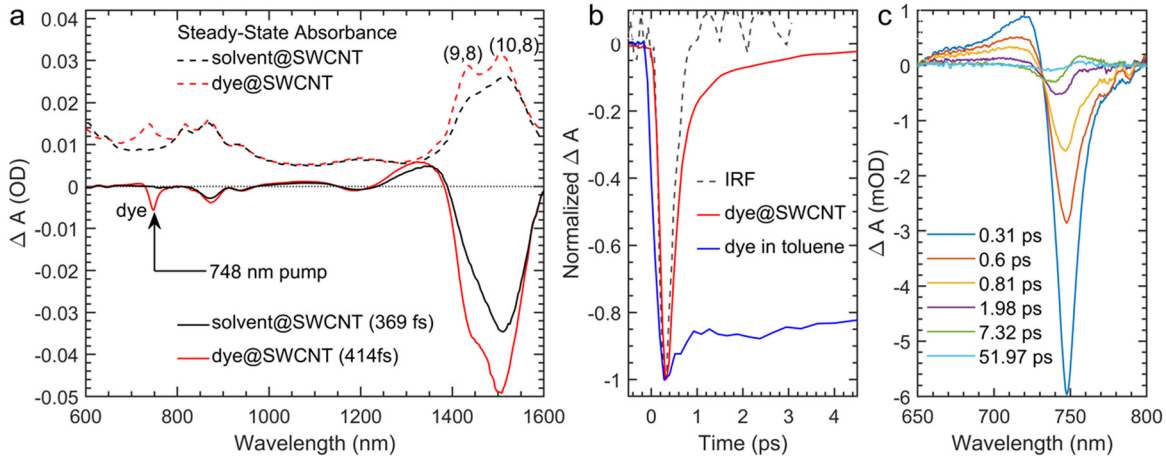

Fig. 20 Transient absorption spectroscopy of diameter-sorted squarylium-filled SWCNTs pumped at 748 nm (in the Squarylium dye absorption band). (a) Peak transient absorption spectra at early times (solid lines) and steady-state absorption spectra (dashed lines) of dye-filled SWCNTs (red traces) and the reference sample filled with DCM (black traces). (b) Kinetic slices of the Squarylium dye (SQ) dissolved in toluene (blue) or encapsulated inside the SWCNTs (red) and the instrument response function (IRF). (e) Selection of transient absorption spectra for short (sub-ps), medium (ps) and long (tens of ps) time scales. Figures reproduced from ref. 209.

probing the kinetics of the encapsulated molecule (Fig. 20b and c) reveals that the encapsulated dye molecules show a multi-component decay with a short component (205 fs) and different longer components. At early times, the ground-state bleach peaks at 748 nm and the excited-state absorption peaks at 720 nm, which can be associated to the EET itself (as it is occurring on the same timescale). Later, however, the transient spectrum turns into a derivative-like spectrum with a crossing point at ~750 nm, that lasts for several tens of ps, denoting a shifted absorption spectrum of the dye.[209]

From the transient absorption experiments, which unlike PL quenching studies track the entire excited-state population, the changes in the rate coefficients can also be used measure the energy transfer efficiency.[36,209,212] In addition, the enhanced transient absorption intensities observed for dye-filled SWCNTs indicate the number of additionally added excitons on the SWCNTs from the dye absorption, therefore providing a complementary measure for the energy-transfer efficiency.[209,231] Each of these methods have their advantages and disadvantages and therefore a combination of these techniques is required to properly estimate the energy transfer efficiency in these nanohybrid samples.

fluorescence background. For example, the encapsulation mechanism of dyes inside SWCNTs has been studied by capitalizing on the quantitative and polarized properties of the Raman scattering of the confined dyes (see Part 2.1).[30,240] Nevertheless, the nearly quantitative quenching of the dyes fluorescence drastically limits fundamental studies and applications based on confined 1D aggregates of fluorescent dyes for photonics. In this context, boron nitride nanotubes (BNNT) have been identified as a promising host template for dyes because of their dielectric nature. Indeed, the slight ionicity of the B–N bond leads to an asymmetric charge distribution and π-bond localization, as opposed to the delocalized structure attributed to SWCNTs. This opens up the band gap, leaving BNNTs as insulators or wide-gap semiconductors of ~5.5 eV, that is mainly independent of diameter and chirality.[241] (Fig. 21) Because of this different electronic behaviour, dye fluorescence is not expected to be quenched in BNNTs, opening the way for the design of fluorescent nano-hybrids.

In 2016, Niskanen *et al.* demonstrated that large diameter BNNTs (~50 nm) can be used as delivery vectors for fluorescent dyes and drugs in cells,[242] experimentally confirming the conservation of the dye fluorescence upon encapsulation. In 2019,

5. Tuning molecules properties using nanotubes as host-templates

5.1. Carbon *versus* boron nitride as nanotube host-template

It was demonstrated in the previous sections that the typical larger bandgap of the encapsulated dyes in the visible range (2–3 eV) compared to that of the semiconducting nanotubes in the NIR leads to effective energy transfers that both readily quench the dyes fluorescence and sensitize the nanotube host. This quenching enabled the study of the Raman fingerprint of the encapsulated dyes that is usually embedded in the huge

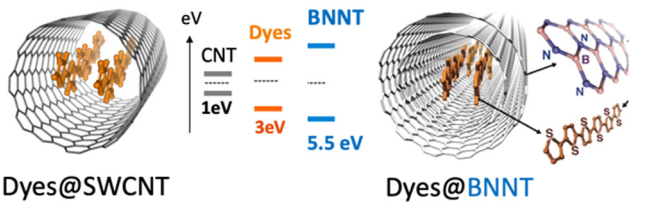

Fig. 21 Illustration of double row of sexithiophene encapsulated inside a SWCNT (left) and in a double walled boron nitride nanotube (right). The energy band gap (average) of semiconducting SWCNT and BNNT are compared to the HOMO–LUMO band of common dyes in the visible range.

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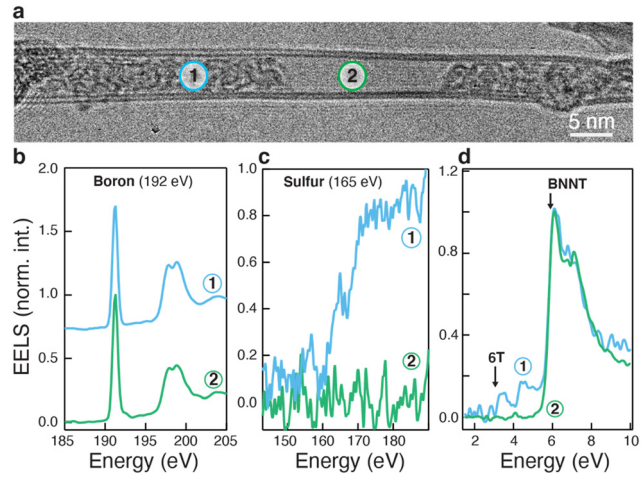

Fig. 22 (a) HRTEM image of a partially filled 6T@BNNT suspended on a Molybdenum-SiO_2 TEM grid. (b) and (c) Chemical analysis of selected areas (1) and (2) in (b) using core-loss STEM-EELS at the boron K-edge (192 eV) and sulfur L2,3-edge (165 eV). (d) Measurements of the energy levels in the same two areas by STEM-EELS in the 1–10 eV energy range. Data reproduced with permission from Allard *et al., Adv. Mater.* 2020, **32**(29), 2001429, Copyright 2020 Wiley.[243]

Allard *et al.* demonstrated that small diameter BNNTs ($\sim$3 nm) can act as smart dye templates for both preserving and stabilizing the fluorescence of dyes and driving specific aggregation states through 1D confinement.[243] The encapsulation is done in a supersaturated solution in an appropriate solvent, in which the BNNTs and encapsulating solution are refluxed at a specific temperature for a given time. This procedure is similar to the liquid-phase method used to encapsulate SWCNTs.[22] The success of the encapsulation was assessed using TEM (Fig. 22). Nano electron energy loss spectroscopy (EELS) data confirmed the chemical nature of the encapsulated material, 6T dye molecules in this case, the core loss signal at 165 eV evidencing the sulfur atoms that are present in 6T, and the threshold at 2.7 eV in the low loss range confirming the dielectric fingerprint of 6T molecules.

Dye molecules in BNNTs exhibit different conformation and aggregation states as a function of the inner diameter of the template (Fig. 23a). The fluorescence spectra, (Fig. 23b) of the dyes@BNNTs hybrids are, with respect to the monomer spectra, broadened, red-shifted and depend on the excitation energy. This characteristic PL signature is most likely indicative of the presence of a broad population of emitters, in which each dye/BNNT combo has a distinctive emission behaviour depending on the diameter of the BNNT and the aggregation state of the dye. It was already shown that SWCNTs templates structure of the assembly of dyes for which specific transition dipole moments alignments, such as J-aggregation are observed.[23] However, the relative narrowness of the diameter distribution of SWCNTs inhibits the formation of more complex, mixed H-like, J-like and HJ-states.[244] Coupled with the aforementioned energy transfer phenomena between SWCNTs and dyes, this impedes on exploiting and studying the richness of the molecular photophysical processes involved in confined 1D systems. Dyes@BNNTs hybrids are, in this context, an ideal model system still under investigation.

Similar to SWCNTs,[23,205] BNNTs sidewalls act as barriers protecting the dyes from external chemical reactions and oxidizing environment, leading to increased photo-stability (Fig. 23d). This conserved fluorescence allows the design of non-fading fluorescent hybrids, complementary to the SWNTs-based Raman hybrids described above, in which the signal can be tailored by the choice of the dye and the diameter of the BNNT.[243,245]

5.2. Non-linear optics from encapsulated dyes

"The 1D confinement of molecules in CNTs can lead to supramolecular arrangements not achievable in traditional bulk materials. Such original ordering of polarizability tensors can be particularly useful in second order nonlinear optics (NLO)".

Second order nonlinear optics (NLO) are widely used in effects such as second harmonic generation and other parametric wavelength conversion, as well as in electro-optics used in optical data communications. Hence, vast research efforts have been focused at developing more efficient NLO materials based on organic compounds.[246–251] Typically, push–pull chromophores, consisting of a conjugated chain with a donor and acceptor group at either end (which have both a large hyperpolarizability β and a large dipole moment) are used. One of the key problems in the design of organic NLO materials is that this large dipole moment also results in dipole–dipole interactions favoring a pairwise antiparallel stacking of the molecules when incorporated in a macroscopic material (Fig. 24a), thus cancelling the NLO response at the macroscopic scale (the vast majority of push–pull dyes crystallizes centrosymmetrically). Traditionally (among other approaches[252]), the main strategy to overcome this is to orient the dipolar molecules by electric field poling in a polymer matrix while heated above its glass transition temperature.[250,253] However, this approach intrinsically results in a thermodynamically unstable state, and long term degradation of the obtained alignment remains difficult to avoid.

An elegant way to solve this central problem is to encapsulate the molecules in a nanoscopic channel just the right size to force the elongated molecules into a 1D array. In that case, the Coulomb interaction naturally favors the ideal head-to-tail alignment of the dipolar NLO molecules (Fig. 24b), resulting in a constructive, coherent addition of the molecular NLO responses. This principle has previously been pursued by encapsulating organic NLO dyes in zeolites,[254–256] but drawbacks include the low filling fractions achieved, the limited processability, and the fact that dye arrays in different channels are not necessarily aligned in the same way (this relies on additional forces such as surface charges).

These remaining issues could be solved by using SWCNTs rather than zeolites.[210] The advantage of this approach is that the aligned molecular arrays are obtained individually as dye@SWCNT hybrids, which form from chemically inert yet solution processible building blocks. Note that because of their

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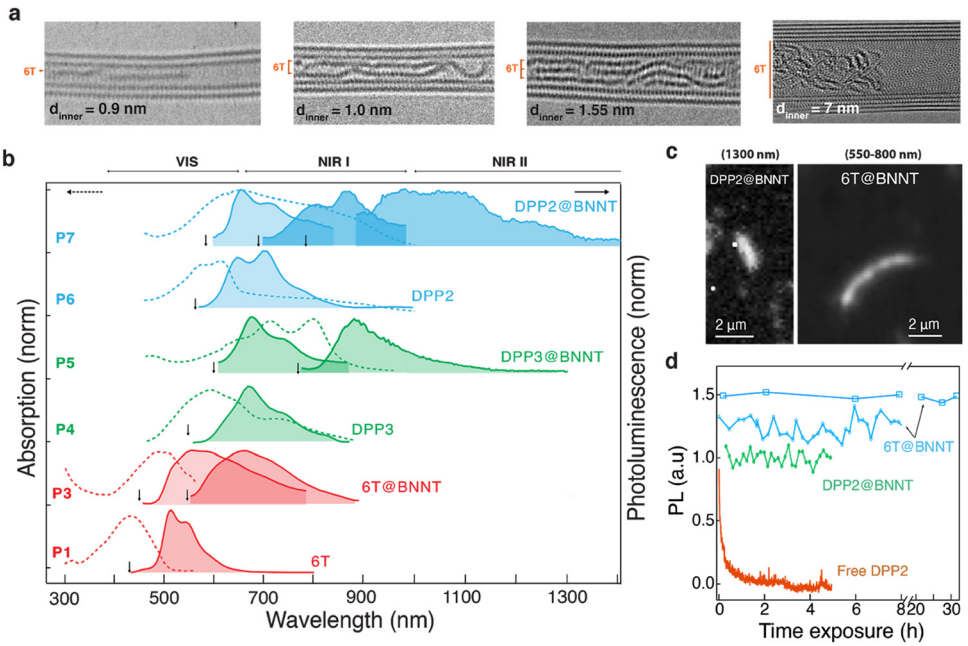

Fig. 23 (a) HRTEM images of 6T molecules inside BNNTs of different diameters. The HRTEM and STEM-EELS data are recorded at 80 kV (b) Absorption (dashed) and fluorescence (continuous) spectra at room temperature of various dyes and corresponding Dyes@BNNTs in solution. The excitation wavelength used for each PL spectrum is indicated by a vertical black arrow. (c) Left panel: Fluorescence image recorded at 1300 nm (1 nm of spectral interval) from a DPP2@BNNT deposited on a Si/SiO_2 substrate (λ_{ex} = 800 nm). The acquisition time is 10 s at a fluence of 1.0 μW μm^{-2}. Right panel: Integrated luminescence imaging of a typical microbundle of a 6T@BNNT deposited on a Si/SiO_2 substrate (λ_{ex} = 532 nm). The time acquisition is 0.2 sec and the laser fluence is 1.2 μW μm^{-2}. (b) Time evolution of the integrated PL of isolated bundles of 6T@BNNTs (blue) and DPP2@BNNTs (green) on a Si/SiO_2 surface, compared to free DPP2 (orange) photoexcited under a fluence of 1.2 μW μm^{-2} at λ_{ex} = 532 nm (long pass filter at 533 nm). Data reproduced with permission from Allard *et al.*, *Adv. Mater.* 2020, **32**(29), 2001429, Copyright 2020 Wiley.[243]

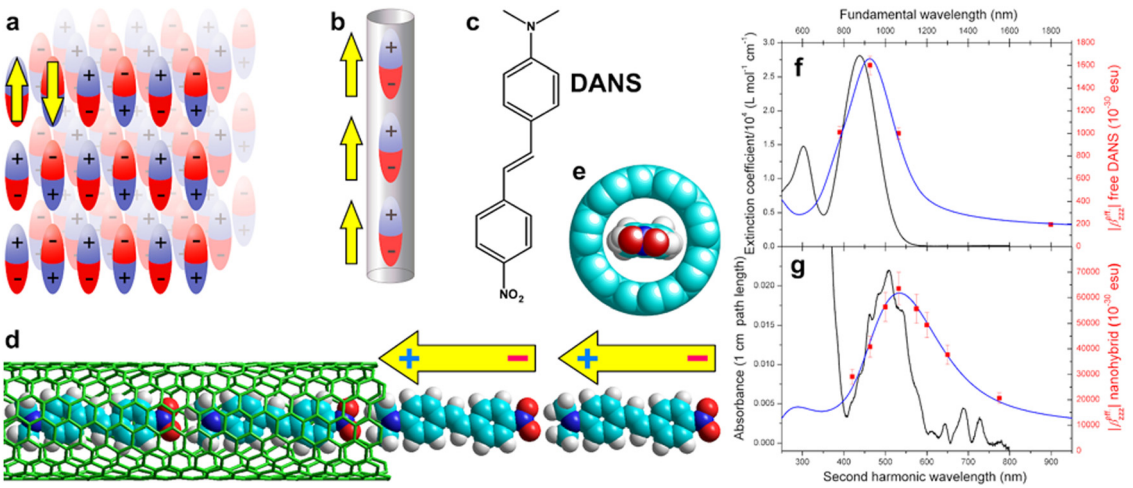

Fig. 24 Stacking of dipolar molecules in different dimensions. (a) In two or three-dimensional bulk materials, Coulomb interactions tend to favor a pairwise antiparallel ordering, thus cancelling asymmetric properties such as the NLO response. (b) In one-dimensional confinement, a polar head-to-tail order is naturally favored, as oppositely charged 'heads' and 'tails' attract each other, leading to a constructive addition of the molecular NLO responses (and dipole moments). (c) Chemical structure of *p*,*p*′-dimethylaminonitrostilbene (DANS), a prototypical NLO molecule. (d) Theoretical model of these NLO molecules in a (9,7) single-walled carbon nanotube (diameter, 1.09 nm), showing that the internal channel just fits a single file of molecules. (e) Axial view of the same structure, with all atoms represented with their van der Waals radii, showing more clearly the relative size of DANS molecules and the SWCNT internal channel. (f) Wavelength dependent NLO response of the free dye and (g) of the dye@SWCNT nanohybrids. Adapted with permission from Cambré *et al.*, *Nat. Nanotechnol.* 2015, **10**(3), 248–252, Copyright © 2015, Springer Nature Limited.[210]

high symmetry, *non-filled* CNTs exhibit only relatively modest second order NLO responses (albeit nonzero, and sizeable, due to their chiral hence non-centrosymmetric nature)[257] but have very large third order nonlinearity.[258–262]

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In an initial study,[210] a well-known NLO dye (*p*,*p*′-dimethylaminonitrostilbene), DANS[263] (Fig. 24c–e) was inserted in SWCNTs from solution using a reflux method (see also ESI† File). Optical absorption measurements, when compared with reference samples made using closed SWCNTs, gave the first clue of the presence of the dye inside the SWCNTs, although far more detailed, chirality dependent information was obtained from PLE and multi-wavelength Raman spectroscopy. A critical SWCNT diameter for NLO dye filling of ~1.02–1.08 nm was demonstrated (Fig. 25), in line with molecular models. This critical diameter was inferred through shifts of the E_{11} and E_{22} transitions and diameter-dependent energy transfer in the PLE map, as well as through shifts of the RBMs in the Raman spectra.

Nonlinear optical characterization by hyper-Rayleigh scattering (*i.e.* incoherent second harmonic light scattering, HRS) allowed the demonstration and the quantification of the head-to-tail alignment. Resonant, wavelength-dependent HRS characterization helped in attributing the NLO response to the encapsulated dye, as opposed to the SWCNTs themselves (Fig. 24f and g). For randomly-oriented molecules in liquid solution, HRS is a fully incoherent process and its intensity is proportional to $N\langle\beta^2\rangle$ where N is the number of molecules and $\langle\beta^2\rangle$ is an orientational average of the appropriate β components. If n dye molecules are aligned in a configuration with head-to-tail stacking (with unchanged β of each molecule), the second-harmonic field amplitudes of dyes in the aggregate add up coherently and the intensity rises quadratically with n. However, the intensities from different aggregates (randomly oriented in the solution) still add up incoherently. Hence the overall intensity is proportional to $\frac{N}{n}\left\langle(n\beta)^2\right\rangle = nN\langle\beta^2\rangle$, which is n times larger intensity than if the molecules were randomly oriented. Thus, the average domain size (n) of perfectly head-to-tail aligned molecules can be determined by comparing the HRS intensity of a dispersion of the dye@SWCNT hybrids to that of a solution of the free dye. Arrays with total hyperpolarizabilities and dipole moments that are ~55–70 times those of the free molecules were observed, translating to domain sizes of ~55–70 molecules. In addition, as the intensity is proportional to β^2, the second harmonic light intensity per array was ~5000 times superior than that of the free molecules. These domain sizes were in line with expectations from a simple model for dipole–dipole interactions, but were also believed to be limited by the length of the (relatively short) SWCNTs used in that initial study. Further improvements can be envisaged by using more gentle tube opening procedures to obtain longer open SWCNTs. The highest enhancements were obtained when using SWCNTs with a smaller average diameter, in line with the idea that only in the thinner tubes just fitting a single file of molecules the ideal head-to-tail alignment is obtained. For larger diameters very diverse stacking arrangements are possible,[264] leading to either constructive or destructive interference of NLO responses. The diversity of possible stackings and resulting inter-molecular coupling has also been demonstrated for different dyes such as symmetric squaraine derivative (see also Part 4.2).[27,30,209] It should be noted that in ref. 210 the main enhancement in intensity was obtained from the

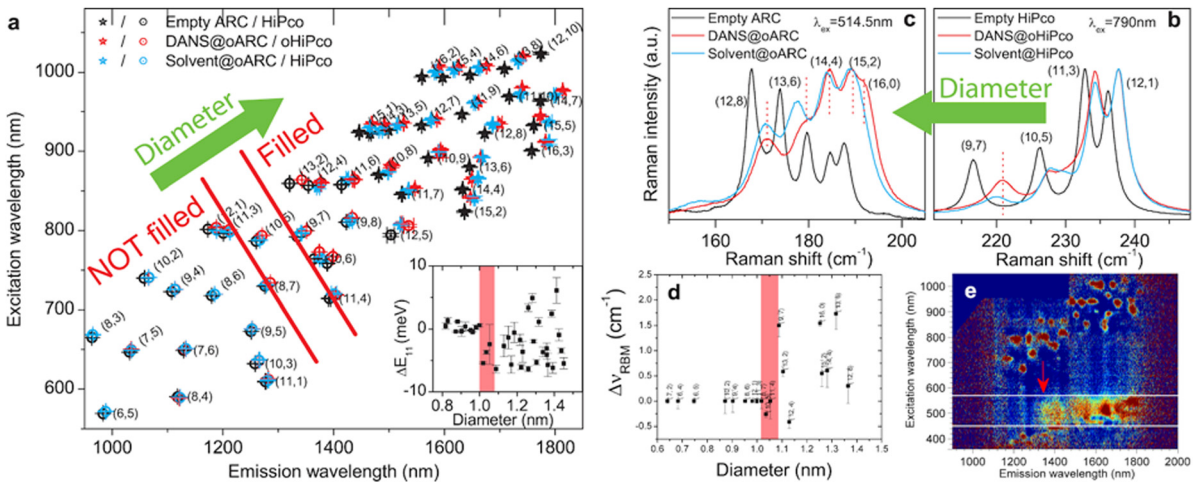

Fig. 25 Raman and PLE shifts and energy transfer, all showing the critical diameter for encapsulation of the NLO molecule DANS. (a) PLE peak positions from 2D fits, showing that for small diameters, no difference is observed between the solvent-filled and DANS-filled SWCNTs. The inset presents the electronic shifts, with a clear onset starting at 1.02 nm. The intermediate region, *i.e.* partial filling, is indicated by the red band. (b) and (c): Raman spectra of empty (black), DANS-filled (red) and solvent-filled (blue) SWCNTs excited at two different wavelengths. (b) clearly shows that for the slightly thinner (12,1) and (11,3) tubes, no difference is observed between the DANS@oSWNT and solvent@oSWNT samples, while for the (9,7) tube a clear shift can be observed, as well as for larger diameter tubes excited at 514.5 nm (c). Vertical lines indicate peak positions of DANS-filled SWNTs. For the intermediate diameter range, only partial filling is observed ((10,5) in (b)). (d) Results of Raman fits (to data using multiple wavelengths) presenting the vibrational shifts also showing the onset of encapsulation at ~1.02–1.08 nm. (e) Difference PLE map (intensity of DANS@SWCNT – D_2O@SWCNT) showing an energy transfer band at excitation wavelengths corresponding to the dye absorption (around 500 nm), which shows a clear onset at SWCNT emission wavelengths of ~1350 nm, corresponding to a critical diameter for filling of ~1.08–1.10 nm. Adapted with permission from Cambré *et al.*, *Nat. Nanotechnol.* 2015, **10**(3), 248–252, Copyright © 2015, Springer Nature Limited.[210]

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coherent addition of molecular NLO responses as discussed above, but that a small additional enhancement through the improvement of the molecular hyperpolarizability itself, due to intermolecular interactions in the J-aggregate-like assembly, was also noted. Theoretical work[265,266] has modelled the cooperative NLO enhancement in these molecular arrays, and while the calculations in ref. 266 overestimate the enhancement effect of intermolecular coupling compared to the experimental results in ref. 210, they do indicate the potential for further improvement in appropriate conditions. Future work should also focus on implementing these nanohybrids in macroscopically aligned NLO materials. Their huge total dipole moment promises easy and good alignment in electric fields, and their size will ensure that the obtained alignment will be extremely stable.

5.3. Near-field nano-optics on encapsulated molecules

Because the dimensions of molecules and the diameters of SWCTs (on the order of 1 nm) are smaller than the characteristic length scale of visible light (500 nm), far-field studies of molecular-CNT hybrids are restricted to the detection of entire hybrids and measurements on bulk samples containing many tubes. Electromagnetic near fields overcome this restriction in resolution,[267,268] enabling optical studies with a resolution reaching one nm and less.[269,270] This section reviews the study of molecular-CNT hybrids in optical near fields.

Metal nanostructures support strong electromagnetic resonances that originate from the collective oscillations of free electrons in the metal. These excitations are called localized surface plasmon resonances or simply plasmons for short.[267,268] About half of the plasmon energy is stored in the energy of the excited electrons, and the other half in the electromagnetic near field close to the metal surface. This field is tailored and further focused in nanoparticle agglomerates or so-called plasmonic oligomers, which are controlled arrangements of nanoscale metallic building blocks. In the void between two or more particles, the electromagnetic near fields are orders of magnitudes stronger than the far field that was used to excite the plasmon. The enhanced field increases the cross section of optical processes like light scattering and emission as well as non-linear processes like second-harmonic generation. Particularly strong, however, is the plasmonic enhancement for Raman scattering,[271,272] for which enhancements of up to 10^8–10^{10} were observed.[273,274] For historical reasons, this effect is known as surface-enhanced Raman scattering (SERS).[275,276]

Nanotubes that are filled with molecules are excellent probes to study the fundamentals of SERS and plasmonic enhancement, Fig. 26.[23,208,277–279] The nanotube's walls prevent direct interaction between the molecules and the metal, which may otherwise change the molecules' properties.[280] Aligning of the molecules inside the tubes, in addition, allows the control of the orientation and placement of the molecular probe in the near field of the plasmonic nanostructure.[23,208] Alpha-sexithiophene (6T) molecules in single-walled carbon nanotubes were used to study plasmonic enhancement by a gold nanodimer (Fig. 26a).[208] The tubes were deposited in the dimer void by dielectrophoresis thereby carrying the molecules into

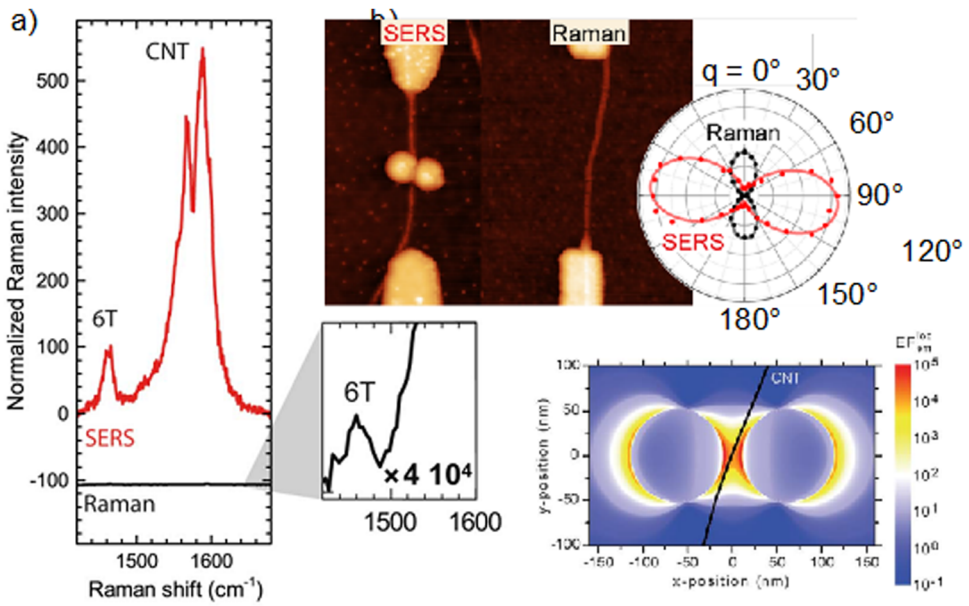

Fig. 26 SERS enhancement of 6T@CNT in a plasmonic gold nanodimer. (a) Comparison of the SERS (red) and Raman (black) intensity measured at 638 nm yielding an experimental SERS enhanced of 8×10^4. (b) AFM images of 6T@CNT in a nanodimer (SERS) and the reference 6T@CNT without plasmonic enhancement (Raman). Right: Polarization dependence of the SERS (red) and Raman (black) intensity. θ measures the polarization of the incoming and outgoing light (they were kept parallel) with respect to the vertical axis in the AFM images. (c) Simulated SERS enhancement factors for 638 nm excitation and the nanodimer geometry of the device used in panel (a) as obtained through AFM and SEM images. The total calculated enhancement factor is 800. Adapted from ref. 277.

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the plasmonic hotspot.[281,282] The plasmonic enhancement is manifested impressively in the polarization dependence of the Raman effect[277] (Fig. 26a and b). Whereas in standard Raman scattering the light is polarized along the nanotube axis, SERS is strongest for polarization along the nanodimer.[208,277,279,281,282] Dyes@CNTs are superior probes for such studies, because of the well-known symmetries of the molecular vibrations. A similar polarization dependence in unfilled CNTs led to speculations on the activation of non-totally symmetric and symmetry-forbidden Raman modes,[283] although this idea was invalidated in a thorough study of polarized SERS enhancement.[279]

6T@CNTs were also used to quantify SERS enhancement factors compared to conventional Raman scattering, showing that plasmonic enhancement is intrinsically stronger than predicted by near-field theory.[208] Remarkably, the increase in the 6T cross section (10^5–10^6) was two orders of magnitude larger than predicted from the electromagnetic enhancement by the plasmonic nanostructure (10^3–10^4).[208] Similar discrepancies were also observed in SERS of molecules without nanotube encapsulation, but typically reasoned to arise from chemical interaction between the molecule and the gold.[273,284,285] Possible explanations are collective effects like coherent scattering in SERS and strong light–matter interaction between the dye and the plasmonic cavity.[285] Strong light–matter interaction may also explain a remarkable narrowing of the SERS resonance (70 meV) when compared to resonant elastic scattering by the plasmonic excitation (290 meV) and even to the 6T Raman resonance in the reference sample (130 meV).

By exploiting plasmonic enhancement was used to study encapsulated species in CNTs were studied with high spatial resolution using tip-enhanced Raman scattering (TERS)[135,156,158,286] where a sharp metal tip acts as an optical antenna that converts propagating lights fields into localized near fields and *vice versa*.[268,287,288] The spatial resolution of TERS as a vibrational optical scanning probe microscopy technique is determined by the extension of the near field, *i.e.*, 10–20 nm for a metallic tip on a semiconducting substrate and down to 1 nm for a metallic tip some Angstrom above a metallic substrate (STM configuration).[269,288] Because of its superior resolution, TERS can be used to characterize confined molecules, the nanotube host, and their mutual interaction by Raman spectroscopy. This feature is particularly useful for dyes@CNTs with filling factors below one, where filled and empty CNT segments alternate over length scales below the diffraction limit, and for molecular compounds with length-dependent properties such as confined carbyne (see also Part 3.1).

Confined carbynes have one Raman active phonon, a fully symmetric vibration along the chain axis termed C-mode,[289] which is well separated from the Raman modes of the host CNTs. The frequency of the C-mode informs on the atomic configuration of the carbyne chain (termed bond-length alteration), which is determined by the length of the chain and the interaction with the nanotube host (for example through van-der-Waals interaction and charge transfer). Initial observations of a spatially continuous TERS signal of carbyne inside isolated DWCNTs (Fig. 27c) suggested that confined carbyne chains may reach lengths up to several hundreds of nanometers.[135,286] However, the relative contributions of chain length and nanotube host to the chain's atomic configuration remained unclear. The correlation between the C-mode frequency and the RBM of the host nanotube was also demonstrated (Fig. 27e).[156] The C-mode frequency was independent of the chain length measured by TERS for chains longer that 30 nm (more than 230 atoms). The C-mode frequency between 1770 and 1835 cm^{-1} was explained entirely by the interaction of the chain with the host nanotube. Using TERS to measure the length of individual carbynes chains in combination with resonant Raman spectroscopy revealed that carbyne confined inside carbon nanotubes have a Raman scattering cross section of 10^{-22} cm^2 sr^{-1} per atom, making it the strongest Raman scatterer ever reported.[290] This high Raman scattering cross section makes carbynes promising candidates for nanoscale thermometry *via* Stokes/anti-Stokes Raman spectroscopy.[291]

Filled nanotubes are excellent probes to study the fundamentals of near-field optical spectroscopy. At the same time,

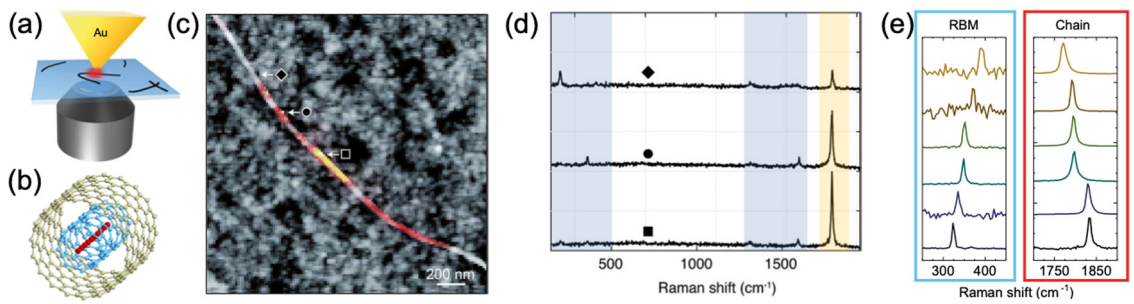

Fig. 27 (a) TERS on carbyne@CNTs (b) Sketch of carbyne chain confined inside a double-walled carbon nanotube. (c) AFM image of a small bundle of DWCNTs overlaid with a TERS image of encapsulated carbyne chains. TERS reveals several carbyne chains with different lengths. (d) Raman spectra of carbyne chains and the nanotube host at three different locations indicated by symbols in (c). (e) TERS spectra of DWCNTs and encapsulated carbyne chains reveal a clear correlation between the RBM of the inner carbon nanotube and the C-mode frequency of the encapsulated carbyne chain. (a), (b) and (e) adapted with permission from Heeg *et al. Nano Lett.* 2018, **18**(9), 5426–5431 Copyright 2018 American Chemical Society, ref. 156 and (d) from ref. 286.

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SERS and TERS can both be used to measure weak signals from encapsulated molecules and study the spatial distribution of the guest species. An exciting development to further improve TERS for dye@CNTs is to exploit of resonant Raman scattering in TERS by performing TERS with tunable laser excitation.[292] Such a system would allow probing of the optical response of the encapsulated and the host species and thus the study of guest–host interaction in the optical response with nanoscale resolution.

6. Towards applications

6.1. Molecules@NT and their use in bio-imaging and drug delivery

Nanotubes are widely investigated as attractive material for biomedical applications due to their unique properties such as their high aspect ratio, nanometric size, modifiable surface and possible biocompatibility. Notably, their 1D architecture is favourable to tissue and cell diffusion[293] when compared to other promising, nano-based biomaterials like quantum dots. Carbon nanotubes possess intrinsic optical properties well-suited for bio-imaging. Their PL emission in the NIR has been well-documented, but the emission intensity is highly dependent on the crystalline quality of the nanotube, which severely limits the extent of available functionalization. Carbon nanotubes also exhibit a strong Raman response exploitable for biological applications. For example, by tailoring C_{13}/C_{12} carbon isotope compositions and monitoring the associated G-band position, multiplexing of up to five Raman nanoprobes was demonstrated in cancer cells and tissues.[294] However, the low specificity of this Raman signature impedes the diversification of the nanoprobes library and decreases the multimodal capabilities.

In a complementary manner, the hollowness of the nanotubes inner channel offers interesting perspectives due to the diversity of materials available for encapsulation, while still leaving the surface free for further, unconstrained functionalization. Nanotubes have been filled with various materials such as contrast agents for computed tomography and magnetic resonance imaging,[295–301] magnetic materials,[302–305] organic dyes and therapeutic compounds, to name a few.

Organic dyes are particularly promising for optical imaging, whether as multiplexed Raman probes in SWNTs, or as stable fluorescence probes in Boron Nitride Nanotubes (BNNTs). Visible fluorescence from dye/SWNTs hybrids is seldom reported, either due to dye/SWNT energy transfer[30,209,231] or to the presence of metallic SWNTs providing non-radiative de-excitation pathways. Free from fluorescence background, encapsulated dyes can then be detected by their Raman scattering signature. Raman spectroscopy is a particularly interesting modality due to the narrowness of its emission peaks and the uniqueness of the Raman signature associated with each compound. However, Raman scattering intensity is usually weak with respect to fluorescence which severely limits its sensitivity for biomedical detection. Encapsulating organic molecules, such as sexithiophene (6T), in SWNTs can increase the Raman cross-section of the aggregates over the threshold for single-object detection.[22] Furthermore, the large quantity of available dyes enables the creation of a library of different nanohybrids, in which the encapsulated dyes are protected from degradation notably *via* reaction with reactive oxygen species.[205] This, combined with appropriate sidewall modification, allows the design of robust and functionalizable nanoprobes tailored for multiplexed detection.

Dyes@SWCNTs were first demonstrated as Raman tags in *Candida Albicans* cells. The SWCNTs were subsequently functionalized with PEG-NH2 to allow solubility in physiological saline solutions and *Candida Albicans* cells were non-specifically tagged following an incubation and washing protocol. Monochromatic Raman images, acquired at a characteristic Raman scattering frequency for βcar, showed the homogeneous localization of the tags on the yeast cells. Further expanding on this topic, multiplexing was then demonstrated using two dyes@SWCNTs on *Candida Albicans*. Both tags distributed mostly evenly on all cells, with some cells displaying a slight enrichment of one type of tag.[306] These studies display the capability of Raman imaging as a multiplexed detection modality for dye@SWCNTs hybrids. The versatility and robustness of the Raman signal was also assessed by targeting specific biological functions. In that case, βcar@SWCNTs probes were functionalized with a biotinylated PEG for the detection of streptavidin, exploiting the strong biotin/streptavidin interaction. Using Raman imaging, it was shown that the probe was selective for the regions where streptavidin was present. Furthermore, good co-localization of the Raman peaks specific to beta-carotene and to SWCNTs demonstrated the reproducibility of the encapsulation process.

Because of the aforementioned advantages, other types of similar nanotubes, such as BNNTs, were investigated as capsules for the encapsulation of organic dyes. However, in the case of BNNTs, the fluorescence of the dye is not quenched, which allows the possibility of photostable fluorescence imaging (see Section 5.1). Large diameter BNNTs ($\sim$50 nm), coated with glycine, were first used as transporters for contrast agents enabling the crossing of cell-permeant and non-cell-permeant dyes, respectively Rhodamine 6G and Lucifer Yellow, beyond the cellular membrane while encapsulated in BNNTs.[242] Dyes encapsulated in small diameter BNNTs ($\sim$3 nm) were also used as a digestive tube contrast agent for *daphnia pulex*.[243] Upon the encapsulation of sexithiophene (6T) and diketopyrrolopyrole derivatives DPP2 and DPP3, the three different color probes were used to illuminate the digestive tract of the *daphnia*, and the signal could be tracked, both *in vivo* and *post-mortem*, for over 12 hours without significant decrease in signal intensity (Fig. 28a and b). Presence of the dyes in *daphniids* was confirmed through their Raman signature, leading to an increased detection specificity (Fig. 28c). Furthermore, *daphniids* incubated with free dyes died within a few minutes of incubation, whereas *daphniids* incubated with dyes@BNNTs survived after incubation times ranging from a few minutes to 36 hours. This illustrates the protection effect of the BNNT which is two-fold:

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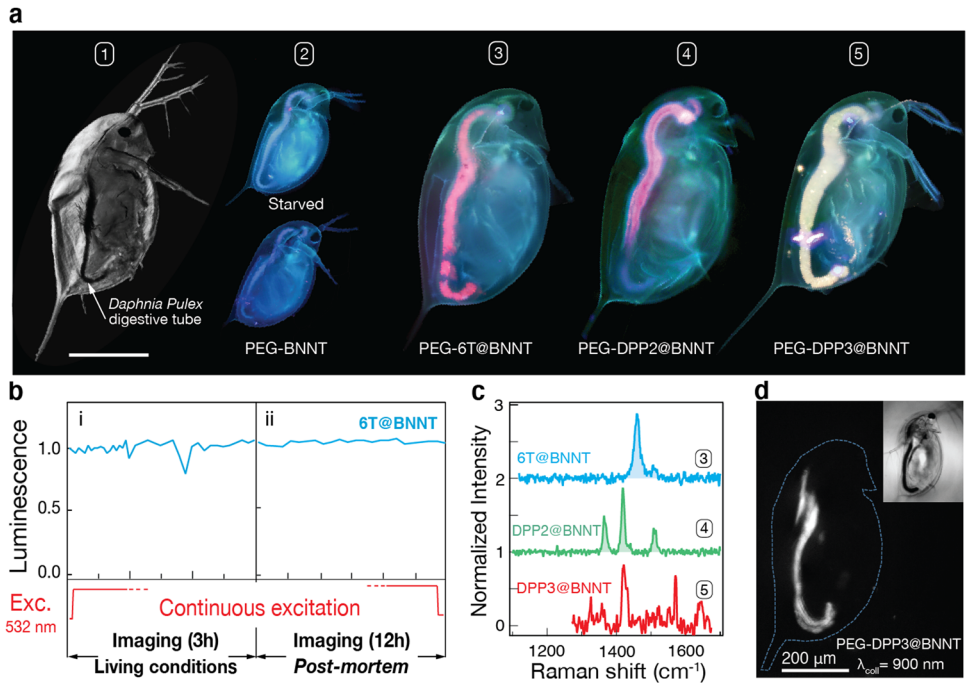

Fig. 28 Dyes@BNNTs as hypermodal nanoprobes illustrated for Vis-NIR bio-imaging of *Daphnia Pulex*. (a) Z-reconstructed confocal optical image of a Daphnia immobilized in Leica medium (#1). PL images recorded through a fluorescence microscope mounted with a UBG filter and 10X objective lens of living starved Daphnia (top #2) and after incubation in a solution of unfilled BNNTs (bottom #2), 6T@BNNTs (#3), DPP2@BNNTs (#4) and DPP3@BNNTs (#5). (b) Luminescence intensity time-lapses of dyes@BNNTs, (i) in living conditions, (ii) in post-mortem conditions. (c) Raman spectra at λ_{ex} = 532 nm taken in the region of the digestive tube of Daphnia for #3 and at λ_{ex} = 633 nm for #4 and #5. (d) PL image of a living Daphnia collected at a wavelength of 900 nm using an excitation at λ_{ex} = 532 nm. Inset: Optical image of the Daphnia. The scale bar is 200 µm for images #1, #3–#5 and 400 µm for #2. Data reproduced with permission from Allard *et al.*, *Adv. Mater.* 2020, **32**(29), 2001429, Copyright 2020 Wiley.[243]

on one-hand, the dyes are shielded from undesirable chemical reactions leading to degradation and photobleaching (Part 5.1). On the other hand, bio-organisms are also protected from relative toxic effects from the dyes. Additionally, encapsulation can shift the emission of the dyes to the NIR II and NIR I spectral windows, which is ideal for *in vivo* biological detection due to decreased tissue scattering, absorption and autofluorescence.[307] An example of imaging in the NIR I window with DPP3@BNNT is shown in Fig. 28d, in which the signal from the digestive track of the daphnia is obtained using a long-pass filter with a cutoff at 900 nm. Large red-shift in emission, attributable to J-aggregate formation, have been previously reported for self-assembled nitroazo or cyanine dyes and used for *in vivo* bio-imaging in the NIR II.[308,309] BNNTs of appropriate diameters could act as templates towards the J-aggregation of dyes for which this aggregation state is naturally unfavourable, unlocking their use as NIR bio-imaging tools.

The internalization of PEG-6T@BNNT in human hepatoblastoma cells (HUH6) was also studied by confocal hyperspectral fluorescence microscopy (Fig. 29a) and the stability of the probe against a conventional cellular membrane label (DiA) was assessed by two-photon microscopy. After 30 min of irradiation, the DiA fluorophore used to reveal the cellular membrane faded whereas strong signal could still be observed from the 6T@BNNT hybrid, confirming the enhanced photostability of the nanoprobe (Fig. 29a). This high photostability in cellular environments enables applications such as tracking of individual dyes@BNNT was monitored for up to 30 min (Fig. 29c).

The hollow interior of nanotubes also provides opportunities for drug loading and targeted delivery. As with dyes, encapsulation allows protection of the molecule from degradation, which, in a drug delivery context, inhibits early inactivation of the therapeutic compound before the attainment of the site of interest. Because of the two-fold protection mentioned previously, encapsulation can also prevent unwanted side effects to the host organism during circulation to the target sites. The surface, then free for modification, can be functionalized with different molecules for increased biocompatibility and targeting specificity.

The potential of carbon nanotubes as loading and delivery systems was investigated theoretically for drugs such as anti-cancer agents,[310–314] antibiotics,[315] antiretrovirals[316,317] and others,[318–320] as well as for DNA and other biomolecules.[321–324] A number of experimental studies focused on the drug-delivery potential of large, multi-walled CNTs (MWCNTs), in which molecules are typically less stable than in smaller diameter CNTs. MWCNTs filled with the anti-cancer drug carboplatin were effective in inhibiting the growth of bladder cancer cells whereas control, unfilled MWCNTs had barely any effect.[325] MWCNTs loaded with another anti-cancer drug, oxaliplatin, and functionalized with PEG-600 showed delayed cytotoxicity in human colon adenocarcinoma cells compared

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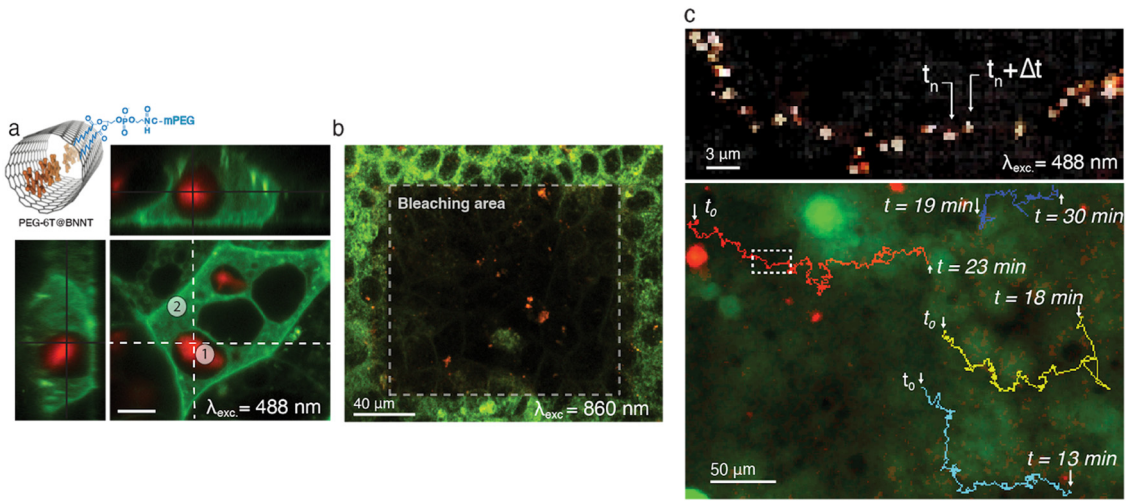

Fig. 29 Imaging of Dyes@BNNTs in a HuH6 cells mat. (a) Hyperspectral and confocal fluorescence imaging of liver cells incubated with PEG-6T@BNNT for 24 hours and with DiA for 10 min before imaging. The red and green channels correspond to the integrated intensity from 600 nm to 780 nm and 490 nm to 520 nm, respectively. (b) Hyperspectral two-photon fluorescence image of fixed hepatoblastoma cells previously incubated for 24 hours with PEG-6T@BNNTs. The dashed line highlights a bleaching test area where the DiA and the PEG-6T@BNNTs were illuminated in two-photon imaging conditions at λ_{ex} = 860 nm for 30 min. The red and green channels correspond to the integrated intensity from 645 nm to 715 nm and from 500 nm to 530 nm, respectively. (c) Top: Tracking of a single PEG-6T@BNNT nanoprobe in confluent HuH6 cells using a superposition of luminescence images extracted every 30 s from a time-lapse datacube (Δt = 650 ms, λ_{ex} = 488 nm). Bottom: Full trajectory reconstruction using a tracking algorithm of the four PEG-6T@BNNTs detected within the field of view. Data reproduced with permission from Allard *et al.*, *Adv. Mater.* 2020, **32**(29), 2001429, Copyright 2020 Wiley.[243]

to oxaliplatin alone.[326] Diameter-dependent release was also highlighted; for cisplatin (CDDP) in smaller diameter MWCNTs enabled slower, more controlled release.[327] As opposed to MWCNTs, SWCNTs have a smaller size and large aspect ratio favoring cell and tissue penetration. As such, efforts have also been deployed for the use of these carriers for drug delivery. Cisplatin (CDDP)-filled ultra-short (US) SWNTs wrapped with pluronic surfactant were found to cause enhanced cytotoxicity to different breast cancer cell lines as compared to both free cisplatin and empty US-SWNTs. In that case, release of CDDP is obtained through sidewall defects and openings as well as through nanotube ends, since the harsh shortening procedure reduces sidewall integrity in US-SWNTs. The addition of the pluronic surfactant hinders the early release of the cisplatin by wrapping over openings in the nanotube.[328] A similar approach was employed while exploiting the enhanced permeability and retention (EPR) effect of tumorous tissues in *in vivo* breast cancer xenografts.[329] Whilst encapsulated in US-SWNTs, cisplatin demonstrated longer circulation time compared to free cisplatin, which resulted in an increased accumulation at the tumor site.

Due to predicted improved biocompatibility,[330] a number of theoretical studies also investigated the feasibility of encapsulation of various therapeutic compounds in boron nitride nanotubes. For example, it was shown that the encapsulation of caffeine into (14,0) BNNTs is unfavourable due to its rigid molecular structure and to repulsive Coulombic nitrogen–nitrogen interactions between the BNNT and caffeine. However, dopamine is readily encapsulated.[331] The most probable physisorption site for carboplatin was calculated to be inside the BNNT cavity, with the surrounding water molecules not affecting the encapsulated carboplatin.[332] Similarly, the anti-cancer drug, gemcitabine (GMC) tend to locate inside (18,0) BNNTs. In that particular case, the release of the drug was activated using heterofullerene ($C_{48}B_{12}$) as a replacing agent. GMC was released from the nanotube upon encapsulation of $C_{48}B_{12}$ due to stronger van der Waals interactions stabilizing the $C_{48}B_{12}$/BNNT system.[333] The feasibility of drug delivery with large diameter BNNTs were also used to deliver curcumin, which, upon release, decreased the level of inflammatory markers in activated microglia, indicating that the loaded curcumin was protected from degradation and retained its biological activity.[242]

Encapsulation of organic dyes in small diameter BNNTs can thus template aggregation of organic dyes, unlocking their use as PL contrast agents down to the NIR. Similarly, encapsulation of dyes in SWCNTs produces a complementary optical signature based on Raman scattering, suited for distinct bio-imaging applications. The inner cavity is also attractive for drug loading, allowing delayed, more precise delivery. In all cases, the nanotube provides two-fold protection whilst facilitating chemical processing and functionalization. The obtained nano-hybrids are stable and robust, with easily modulated and diversified prospective applications due to a multitude of molecules and dyes that can be readily encapsulated.

6.2. Thin films with fluorescence anisotropy based on Mol@BNNTs

Due to their intrinsic polarized optical properties, carbon nanotubes have been used since decades for inducing optical

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anisotropy in nematic liquid or thin films for VIS to THz applications.[334–337] Nanotubes host template is also a promising materials for aligning molecules and achieving strong polarization effect at the nanoscale. For example, Gaufrès *et al.* reported highly polarized Raman scattering pattern of confined molecules inside individual SWCNT has been reported.[22] Jakubek *et al.* has also reported on the fluorescence of polymers, such as rra-P3HT, wrap along the outer wall of BNNTs and showed that the emission is polarized and useful for quality assessment of the BNNT itself.[338,339]

The fluorescence polarization properties of an elongated dye molecule, α-sexithiophene (6T), encapsulated inside a BNNT (6T@BNNT) has also been recently explored.[340] Aberration-corrected High Resolution Transmission Electronic Microscopy (ac-HRTEM) and high-resolution fluorescence imaging of many individual 6T@BNNTs show evidence of remarkably well-aligned molecules along the BNNT axis for inner diameter below 3 nm. The optical results also establish that these individual 6T@BNNTs emit highly polarized fluorescence. (Fig. 30).

At the same time, the BNNT can be easily oriented within a flexible and transparent matrix by stretching. For example, stretched matrix made of PMMA loaded with dispersed and aligned 6T@BNNTs display strong fluorescence anisotropy with an enhanced anisotropy factor ($r = \frac{I_{0-0} - GI_{0-90}}{I_{0-0} - 2GI_{0-90}}$, where with G, the optical factor correction of the set-up (G = 1.12), (I_{0-0}, I_{0-90}) and the normalized intensity of the fluorescence, parallelly polarized with respect to the incident polarization, respectively) with an anisotropy enhancement factor of 2,7 between the stretched and unstretched film. (Fig. 31). These results clearly evidence that the BNNTs structure acts, because of their ratio aspect, as a powerful 1D template for aligning organic molecules from the nano to the macroscopic scales.

6.3. Dyes@NT for photovoltaic

6.3.1. SWCNT for photovoltaic. It is the unique combination of desirable electronic, optical, and physiochemical properties that make SWCNTs an attractive candidate for next generation optoelectronics and in particular photovoltaics.[197,341–346] These include excellent stability towards degradation in ambient, humid, hot or ultraviolet radiation conditions,[347–349] a wide electrochemically stable window,[345,350] a 1D structure with high surface area, ultrafast charge transport mobility ($\sim 10^4$ cm^2 V^{-1} s^{-1})[349,351] and the presence of multiple excitonic transitions in the infrared, visible and ultraviolet spectral ranges. To date SWCNTs have been applied to almost all optoelectronics components.[352–374] It is therefore unsurprising that a plethora of photovoltaic architectures involving SWCNTs can be found in the literature. These include photo-electrochemical cells,[352,375–377] Si:CNT heterojunctions,[341,342,378–381] dye sensitized solar cells,[382–384] single layer,[362,363,385] bilayer-[386–388] and bulk-heterojunction[374,389] organic photovoltaics.

6.3.2. Dye-sensitised carbon nanotube photovoltaics. "The goal of CNT sensitization with chromophores has always been to improve light absorption and thereby the photocurrent of a device"

For photovoltaics the goal of sensitization with chromophores, be that organic dyes[16,211,231,390–392] porphyrins and phtalocyanine derivatives,[352,375–377] polyoxometalates[393] or azobenzene complexes,[391] has always been to improve light absorption and thereby the photocurrent of a device. CNT-based devices are no different. Ideally, photosensitization is performed with a chromophore located in the spectral range where the CNTs have a low extinction coefficient.[231] In fact, sensitization is not expected to be beneficial in all cases. For example, in Si:CNT heterojunctions, where an ultrathin nanotube mat is placed in interaction on a silicon substrate, incident solar radiation mostly passes through the nanotube mat and is absorbed by the silicon, creating electron–hole pairs that diffuse to a depletion region set up at the carbon CNT:silicon interface.[341,342] In this architecture the CNTs are generally assumed not to contribute to the photocurrent[394–396] but rather act as an efficient hole transport layer where a balance between low sheet resistance and high optical transparency is important.[397–400] Here, sensitization of the CNT layer would only serve to reduce the amount of light reaching the silicon substrate by reducing the transmittance of the CNT film and introduce scattering sites at the CNT:Si interface. A similar argument can be made for dye sensitized solar cells, where CNTs are also commonly used as transparent conducting electrodes and where exciton dissociation occurs entirely at a dye/TiO_2 interface. In a CNT:Si heterojunctions small but measurable photocurrent contribution from the nanotubes was inferred using transient absorption spectroscopy.[401] Endohedral sensitization may be a convenient avenue to further study this contribution without degrading the CNT:Si interface. However, because most potential fillers absorb in the visible range, where excitons are already formed efficiently in silicon, any gains to the photocurrent are expected to be modest at best and would need to be balanced against a reduced absorption in the silicon. It is more likely that Si:CNT heterojunctions and dye sensitized solar cells will benefit from endohedral doping of the CNTs with molecules like tetrathiafulvalene (TTF), tetracyanoquinodimethane (TCNQ)[402,403] and fullerenes[404–406] which afford

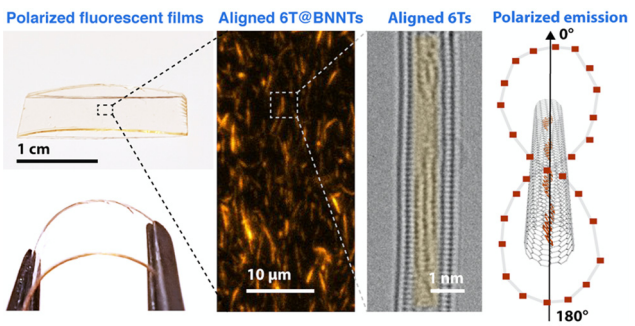

Fig. 30 Polarized fluorescent and transparent films based on aligned 6T emitters inside assemblies of aligned boron nitride nanotubes. Reproduced from ref. 340.

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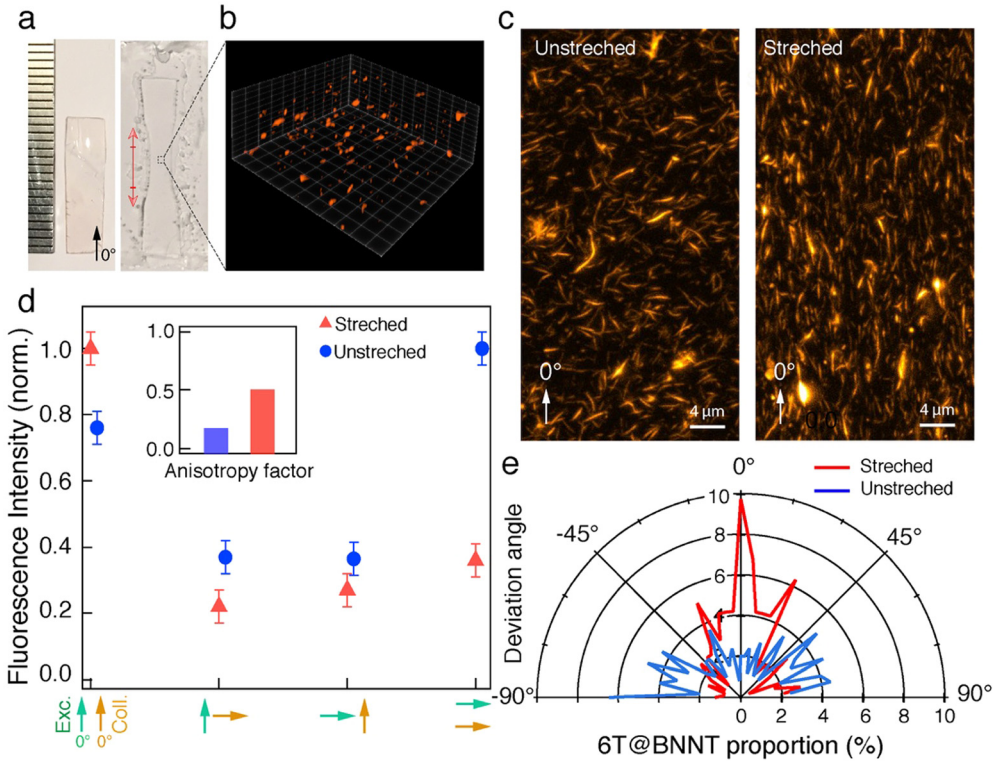

Fig. 31 Polarized fluorescence from transparent thin films based on 6T@BNNT and PMMA. (a) Image of a centimetre wide PMMA films before and after stretching (left and right). (b) Partial rendering of a 'z-stack' of a PMMA-6T@BNNT film acquired with confocal fluorescence microscopy, in a 40 μm × 40 μm × 30 μm volume. (c) Z-projection of the stretched and unstretched film datacubes. (d) Normalized fluorescence intensity recorded from the film, when unstretched and stretched, for different parallel, orthogonal and crossed polarization conditions of the excitation and collection, as highlighted by green and orange arrows, respectively. Inset: Anisotropy factor for the unstretched (left) and stretched sample (right) estimated from the data presented in the same panel. (e) Polar representation of the 6T@BNNT percentage as a function of the deviation angle from the direction of stretch (0°). Reproduced from ref. 340.

the opportunity to alter the band gap energies[407,408] or carrier concentration[402] of the CNTs.[409]

In chromophore-sensitized SWCNT photovoltaics, the CNT is expected to take one of two roles: as a conductive conduit to improve the transport of charge carriers in existing light harvesting stacks, such that energy, hole or electron transfer from the chromophore is primarily driven by its contact to other organic molecules (Fig. 32a) or as a donor or acceptor in the photoactive layer, contributing to the photocurrent and sensitization is used to spectrally extend the absorption of the SWCNTs (Fig. 32b). In the first case, intimate contact between the chromophore and a third material is necessary, reducing the appropriateness of endohedral *vs.* exohedral functionalization. However, in the second case, endohedral filling is clearly advantageous because it isolates the sensitizer from the donor/acceptor pair necessary for nanotube exciton dissociation, thereby reducing charge trapping and scattering at the heterojunction interface and ensuring high efficiency. To date, despite the extensive spectroscopic studies discussed previously, the number of photovoltaic devices containing SWCNTs that are modified with chromophores, including devices where the measured photocurrent is the sum of both the nanotube's and the chromophore's absorption and devices containing endohedrally-modified SWCNTs, is limited.

Early experiments on SWCNT/chromophore hybrids were still performed on as-grown SWCNTs composed of a broad distribution of semiconducting and metallic species.[345,377] SWCNTs were dispersed with conjugated diazo dyes such as congo red and trypan blue,[350] and pyrene and its derivatives,[352,376,377,412] poly(phenyleneethynylene)[385] and *N*-(1-pyrenyl)maleimide.[413] At this early stage of development, photoelectrochemical devices were a convenient proof of principle. SWCNTs dispersed by charged pyrenes were particularly attractive because they facilitated the subsequent electrostatic functionalization with charged porphyrins and phthalocyanines.[382–384] Usually, photoelectrochemical cells are made in a layer by layer approach on ITO electrodes.[414] A single layer of SWCNT/Pyrene^{+}/ZnP^{8-} lead to monochromatic incident photon-to-charge-carrier (IPCE) of 4.2% whereas multilayers reached 8.5%.[377,415] Subsequent works have shown the efficiency of nanotube functionalization to improve dramatically the film photocurrents. Py/SWNT films incorporating metal-free (H_2Pc) and zinc phthalocyanines (ZnPc) chemically modified with a pyrene linker had stable and reproducible photocurrents with IPCE values as large as 15 and 23%, respectively without and with an applied bias of +0.1 V.[352] Likewise, SWCNTs functionalized with 4-(trimethylsilyl)ethynylaniline to which a ZnPc derivative bearing an azide group was

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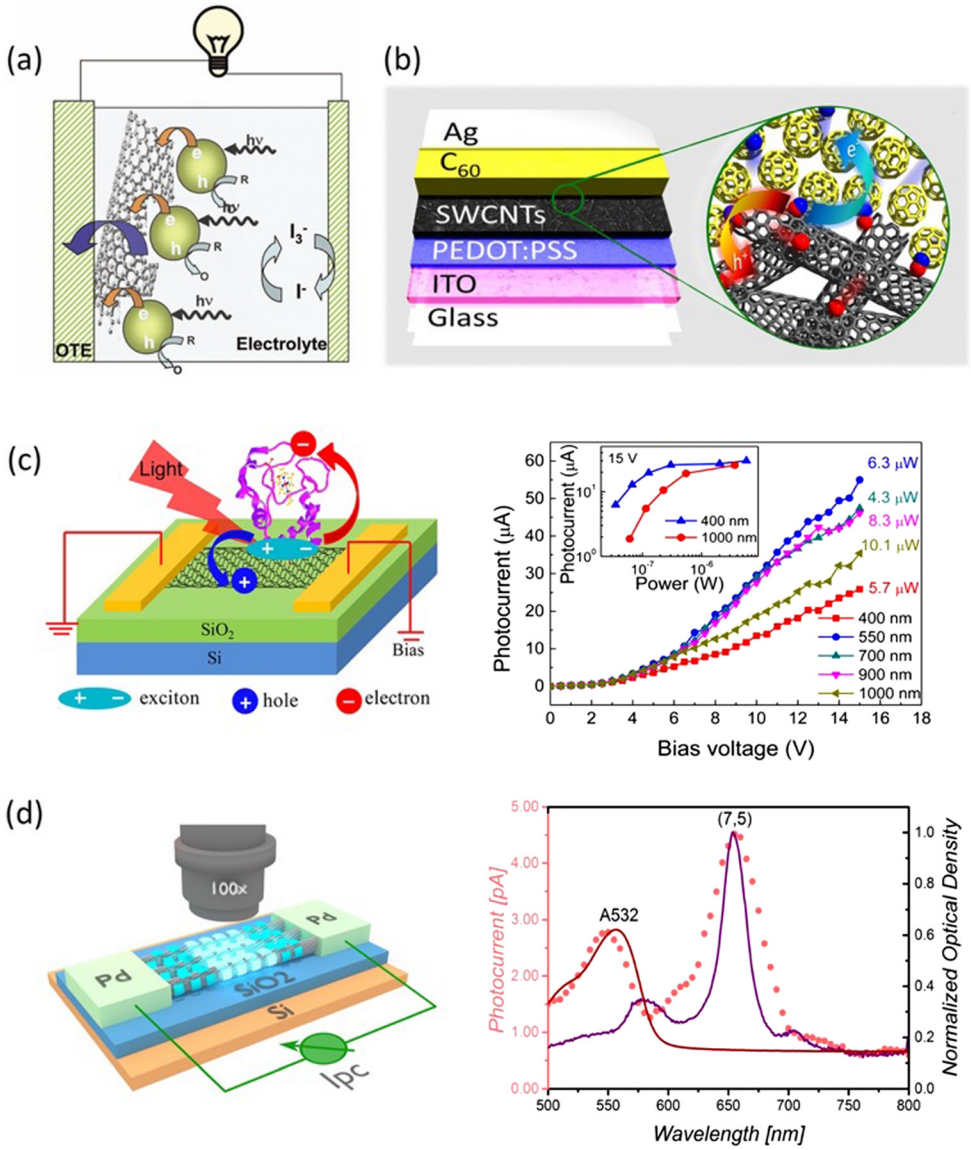

Fig. 32 Strategies to employ chromophore modified carbon nanotubes: (a) in a photoelectrochemical cell by excitation of light harvesting assemblies anchored to the nanotubes, panel reproduced with permission from Kamat *et al.*, *Nano Today* 2006, **1**(4), Copyright © 2006 Elsevier, ref. 410 or (b) as a donor or acceptor in the photoactive layer, where dye modification enhances light absorption. Panel reproduced with permission from Pfohl *et al. Adv. Energy Mater.* 2016, **6**(1), 1501345 © 2015 Wiley-VCH, ref. 372 (c) broad band light detection, as shown with permission from Gong *et al.*, *Acs Appl. Mater. Interfaces* 2017, **9**(12), 11016–11024, Copyright 2017 American Chemical Society.[411] and (d) the measurement of photocurrent from dye-modified (A532) (7,5) SWCNTs in the work of Alam *et al.*[390]

attached reached IPCE values of 17.3%, which was 30% larger than for SWCNTs alone.[218] In related work, electrodes fabricated from SWCNTs covalently sensitised with commercially available N3 dye ((*cis*-bis(4,4)-dicarboxyl-2,2-bipyridine) dithiocyanato ruthenium(II)) showed a modest photoresponse.[415] In all cases, the carbon nanotubes were assumed not to contribute to the photocurrent but played the role of electron acceptors. Because of the energetic position of the ITO conduction band, a thermodynamically driven electron transfer from the reduced SWCNT is favoured. The redox couple in the electrolyte was responsible for regenerating the oxidized porphyrins or phthalocyanines.

The first single layer organic photovoltaic devices containing CNTs were metal/insulator/metal (MIM) devices and quickly combined with chromophores. In pioneering work, Kymakis *et al.*[363] interfaced SWCNTs with the polymer P3OT in a layer stack consisting of ITO/P3OT/P3OT-SWCNTs/Al. The interaction of the polymer with the CNTs allowed for charge separation of the photogenerated carriers in the polymer and efficient electron transport through the CNTs.

In light of the large exciton binding energies of 0.3–1 eV[416–418] for semiconducting SWCNTs, the strong exciton quenching ability of metallic nanotubes,[389] and the possibility of bundled SWCNTs to lead to charge trapping between the different bandgap

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species,[388] it is unsurprising that these early devices, which used unsorted CNTs, failed to measure any discernible photocurrent from the nanotubes themselves. However, since then, the efficacy at which CNTs can be processed and sorted according to their structure,[419–423] has come a long way.[424] Highly selective separation techniques have been developed and it is now possible to fabricate devices from type-selective SWCNTs.[197,425–433]

Capitalizing on these sorting techniques, SWCNTs wrapped with PFO were mixed with different electron and hole acceptor compounds such as C_{60}, [6,6]-phenyl C_{61} butyric acid methyl ester ($PC_{61}BM$) or poly(3-hexylthiophene) (P3HT) to form a type-II heterojunctions with sufficient energetic offset to induce the dissociation of excitons from the nanotubes and generate photocurrent.[386] Power conversion efficiencies of 1% with C_{60} and 3% with $PC_{71}BM$ were achieved based on this approach.[374,389] For the first time external quantum efficiency (EQE) measurements irrefutably correlated device photocurrent to the absorption of a nanotube film. Since these first reports the research teams around Strano,[388] Bao,[434,435] Blackburn,[373,436] Flavel,[372,387] Hersam[389,437] and Brabec[438] to name a few, have all worked on improving various aspects of similar devices but to date no-one has presented a chromophore-modified variant.

This lack of devices is mostly related to materials challenges in finding the correct combination of nanotube chirality/diameter, chromophore and acceptor/donor system. In the early work of Bindl *et al.* the authors also measured the internal quantum efficiency (IQE) of five different nanotubes (7,5), (7,6), (8,6), (8,7) and (9,7) at a SWCNT:C_{60} interface.[196,386] They observed a reduction in IQE of 91% for (7,5), which is the smallest diameter of the five species, to below 30% for (9,7) which has the largest, and suggested that excitons from a CNT with a diameter of 1 nm or larger are no longer efficiently dissociated. This observation is associated with the requirement of a minimum energetic offset between the LUMO of the nanotube and that of its acceptor in order for efficient exciton dissociation to take place. When using C_{60} as an acceptor, a diameter limit of 0.95 nm, corresponding to the (8,6) SWCNT, was suggested.[387] Marcus theory further showed that a LUMO offset of 130 meV is optimal, which is best satisfied for much smaller diameter ($\sim$0.8 nm) species such as (8,3), (6,5) and (9,1) in combination with C_{60}.[436] The diameter-dependent bandgap of SWCNTs therefore imposes a specific selection of (n,m) species which match with the fixed LUMO band of the acceptor molecule. Large diameter SWCNTs ($D_t \approx 1.665$ nm) can be used in combination with fullerene-based acceptors with high electron affinity.[436] However, given that most research groups use commercially available acceptors and polymers to sort the CNTs, which are inherently selective to species below 1 nm, the available endohedral volume is far below that required for usual chromophore encapsulation, (1.4–2 nm).[205,231,392] It is also important to consider that even if chromophores could be placed inside CNTs, most absorb light between 200–500 nm, which is a spectral region where C_{60}, $PC_{71}BM$, *etc.* already have a high extinction coefficient and thus any gains may be marginal. Furthermore, co-location of a chromophore inside a CNT does not guarantee a measurable photocurrent contribution. In fact, the 1D endohedral arrangement of coronenes resulted in an increase in the exciton binding energy no energy transfer to the CNT was observed.[16] The only known example of dye/CNT/fullerene being assembled into an architecture similar to that used for solar cells involved a ferrocenyl dye inside a 1.4 nm diameter CNT, which was subsequently modified with a fullerodendron outside.[392] Although the diameters of the CNTs were too large for a photocurrent to be measured from them, electron transfer from the dye to the C_{60} through an intermediate S_{22} band of the SWCNT was observed. This sparks interesting questions regarding the transparency of the carbon nanotube wall to charge carriers in combination with other organic layers CNT-FETs have been functionalized with chromophores as light sensors, optically activated switches, memories and to measure photocurrent.[229,439–442] In these hybrid systems, photogenerated carriers are separated by either an externally applied gate bias or by internal fields at the Schottky barrier induced at the SWCNT-metal interface.[371,390,391,440,442–450]

The only known example of a device-integrated nanohybrid to extend the optical property of the CNT is a CNT-FET made exclusively from (7,5) nanotubes to which three different fluorescent dyes, ATTO 532, 565 and 610 were drop cast.[390] With the aid of a photocurrent spectroscopy setup and a supercontinuum light source the authors had a much large incident photon flux and were able to show simultaneous photocurrent generation from the nanotube and dyes. Förster energy transfer was found to be responsible for the photocurrent, but the dye was significantly photobleached due to the light intensity used for measurement. Using endohedral functionalization could reduce this effect, leaving significant room for improvement in the future. Complementary review articles focusing on electronic and electro-chemistry application of filled CNTs can be found in ref. 18 and 451.

7. Looking forward

7.1. Perspectives

Strategies that leverage nanotubes as architectural scaffolds have diversified in the last decade to include a large array of 1D molecular assemblies of varied compositions and architectures. As reviewed here, a sizable spectrum of molecular organizations at the nanoscale have been prepared and tested. This field has developed new fascinating ways to engineer physical and chemical interactions between 1D hosts and molecules with potential groundbreaking applications in bio-imaging, light harvesting, photonics, to name a few. This review highlights, however, an impressive landscape of possible interactions between 0D/1D and 1D/1D optically actives nano-objects for nano-optoelectronics. The combination of intermolecular forces and electronic interactions has allowed to unravel, in the van der Waals, new regimes for efficient charge and energy transfer and other thermally activated processes in these 1D heterostructures. This is especially useful to actively tune the

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physical properties of the host nanotubes and optimize fluorescence quantum yields, charge transfer doping, or Fermi level alignment. On the other hand, the nanotubes provide a crystalline and chemically stable environment, which are effective media to stabilize 1D molecular aggregation that can lead to confinement and high ordering of molecules. The approach has offered examples of bright J-aggregation states and highly ordered dipolar organizations, which are of considerable interest for bio-imaging, none-linear optics, and other photonics applications. Further advances in complex encapsulation processes could provide access to an even wider library of assemblies, which can serve as building blocks for more complex and multifunctional 1D heterostructures. In particular, the possibility of combining semiconductors, metals, and optical emitters together at the nanoscale opens fascinating opportunities for shaping multifunctional materials and metamaterials. Mixed dimensionality heterostructures based on Mol@NT and 2D materials will also bring unprecedented degrees of freedom for designing complex materials build novel functionalities. As a result, technological functions could be improved and further miniaturized with, on the one hand, a smaller quantity of materials and lower costs, on the other hand, improved performance.

As a platform for advanced nanochemistry, the nanotube-based 1D heterostructures are built around a nanocontainer that acts as a barrier against reactive chemicals and other environmental factors inducing degradation. Capitalizing on the robustness of the nanotube host, progress in achieving advanced synthesis inside the nanotube reactor shows exciting potential in chemical engineering, such as the formation and stabilization of exotic phases, isolation of unstable compounds and reactive materials that have hitherto been difficult to produce in ambient conditions. In addition, the synergistic properties of 1D heterostructures made solely of naturally abundant compounds, such as carbon, sulfur, and cheap metals (*e.g.* Fe), can develop into viable alternatives for key technologies where rare earths, precious metals, or toxic elements are still being used nowadays, albeit problematic societal acceptance and negative impact of their use on life and ecosystems.

7.2. Challenges

Effort towards developing more sophisticated 1D heterostructures faces multiple challenges from synthesis strategies to characterization. Indeed, combining nanotubes of different composition, diameters, and wall thickness with molecular species requires one to deal with multiple interactions and complex behaviors. Outer and inner wall functionalization has also to be considered to facilitate processing in batches and scaled-up production for subsequent use in technologies. Access to standardized nanotubes as starting materials is therefore paramount for 1D heterostructures adoption in applications. Despite the continuous progress in sorting of carbon nanotubes by chirality, diameter and length, the access of low cost and standardized quality of single chiralities of nanotubes remains an outstanding issue to be addressed. The supply challenge is exacerbated for the specific case of BNNTs, which come from suppliers as a complex mixture of diameters and wall thicknesses. Compared to the commercially available CNT, only few post-synthesis purification/sorting methods of BNNTs have been reported up to now. The large BNNT bandgap ($\sim$5.5 eV) also makes the optical characterization of material's composition challenging to measure. Developing efficient techniques for sorting BNNTs by length, diameter and number of walls is therefore essential before one could reach better control over the optical properties of the BNNT-based heterostructures.

To take full advantage of Mol@NTs capabilities, it will be essential to control the molecular encapsulation process into the nanotubes. (Note that comparative description of encapsulation methods and characterization techniques for Mol@NT systems is available in the ESI.†) Considering the impact on properties, methods that can position and tailor the interaction between similar or different molecules inside and outside the nanotubes are to be searched for. Scheme to drive 1D assemblies along the nanotube axis will also need to be developed in the future. A thorough understanding of the elementary mechanisms behind the encapsulation/adsorption of molecules during encapsulation will help reach a quasi-deterministic control of the entry and diffusion of molecules along the nanotube wall. The specific architecture of the nanotube template represents a strongly anisotropic and complex nanoporous system. The adsorption and diffusion mechanisms of molecules and their aggregates inside/outside the nanotubes are still not fully understood. The impact of the encapsulation intermediates, such as organic solvents and of the presence functional groups naturally grafted at the defective ends of the nanotubes are likely to change significantly the thermodynamic map of the elementary adsorption events. In addition, future work will need to disentangle the roles of phonons, strains, dielectric screening, and aggregation effects on the optoelectronic properties of the 1D heterostructures.

Therefore, a long and exciting way remains before we fully master the synthesis and properties of 1D heterostructures. These challenges call for innovative experiments and theoretical studies to gain further insight into molecular diffusion inside confined spaces and into new opportunities in surface science topics related to 1D templating and high confinement.

Data availability

No primary research results, software or code have been included and no new data were generated or analysed as part of this review.

Conflicts of interest

There are no conflicts to declare.

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