



## Correction: Quantifying the ion coordination strength in polymer electrolytes

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Correction for 'Quantifying the ion coordination strength in polymer electrolytes' by Rasmus Andersson et al., *Phys. Chem. Chem. Phys.*, 2022, 24, 16343–16352, <https://doi.org/10.1039/D2CP01904C>.

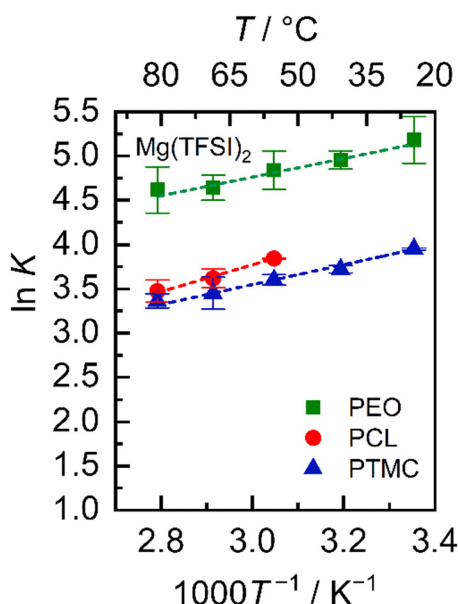
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The authors would like to make changes to Table 1 and Fig. 4c in the published manuscript. The corrections pertain to the data for the Mg(TFSI)<sub>2</sub> systems, where the wrong concentration of TFSI ions was used for the original calculations. The revised versions of Fig. 4c and the relevant part of Table 1 are shown here.

**Table 1** Extracted enthalpy, entropy and derived Gibbs free energy at 25 °C for the dissociation of Mg(TFSI)<sub>2</sub> in PEO, PCL and PTMC

Salt	Polymer	$\Delta H^0/\text{kJ mol}^{-1}$	$\Delta S^0/\text{J mol}^{-1} \text{K}^{-1}$	$\Delta G^0/\text{kJ mol}^{-1}$
Mg(TFSI) <sub>2</sub>	PEO	−8.8	13.3	−12.7
	PCL	−12.6	−6.5	−10.7
	PTMC	−9.4	1.4	−9.8



**Fig. 4** (c) Van't Hoff plots of the equilibrium constant  $K$  as a function of the inverse temperature for the dissociation equilibrium of Mg(TFSI)<sub>2</sub> in PCL, PTMC and PEO. Dashed lines represent the linear fit to the data.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

