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# Difference in reaction mechanism between ZnZrO<sub>x</sub> and InZrO<sub>x</sub> for CO<sub>2</sub> hydrogenation†

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Oxide solid-solution catalysts, such as Zn-doped ZrO<sub>2</sub> (ZnZrO<sub>x</sub>) and In-doped ZrO<sub>2</sub> (InZrO<sub>x</sub>), exhibit distinctive catalytic capabilities for CH<sub>3</sub>OH synthesis via CO<sub>2</sub> hydrogenation. We investigated the active site structures of these catalysts and their associated reaction mechanisms using both experimental and computational approaches. Electron microscopy and X-ray absorption spectroscopy reveal that the primary active sites are isolated cations, such as Zn<sup>2+</sup> and In<sup>3+</sup>, dissolved in tetragonal ZrO<sub>2</sub>. Notably, for Zn<sup>2+</sup>, decomposition of the methoxy group, which is an essential intermediate in CH<sub>4</sub> synthesis, is partially suppressed because of the relatively high stability of the methoxy group. Conversely, the methyl group strongly adsorbs on In<sup>3+</sup>, facilitating the conversion of the methoxy species into methyl groups. The decomposition of CH<sub>3</sub>OH is also suggested to contribute to CH<sub>4</sub> synthesis. These results highlight the generation of CH<sub>4</sub> as a byproduct of the InZrO<sub>x</sub> catalyst. Understanding the active site structure and elucidating the reaction mechanism at the atomic level are anticipated to contribute significantly to the future development of oxide solid-solution catalysts.

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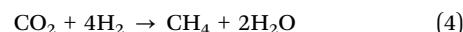
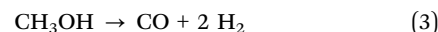
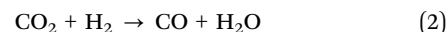
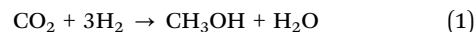
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## Introduction

Carbon dioxide hydrogenation is recognized as a key technology for the utilization of CO<sub>2</sub>. Hydrogen is generated by water electrolysis using renewable energy sources. The produced H<sub>2</sub> reacts with CO<sub>2</sub> to create raw materials for chemicals and fuels. This process has the potential to contribute substantially to the establishment of a sustainable society, provided the feasibility of obtaining cost-effective H<sub>2</sub> is considered. Methanol is often regarded as the building block in C1 chemistry. Essentially, CO<sub>2</sub>-to-CH<sub>3</sub>OH hydrogenation (eqn (1)) occurs at approximately 250 °C using Cu-based catalysts, such as Cu/ZnO/Al<sub>2</sub>O<sub>3</sub>,<sup>1–3</sup> and

the resulting CH<sub>3</sub>OH molecules are subsequently transformed into hydrocarbons and oxygenates at acid sites in solid acid catalysts at 400 °C and higher temperatures.<sup>4</sup> When reactors designed for CH<sub>3</sub>OH synthesis and conversion reactions are connected in series, it is essential to recognize that the former reaction typically determines the overall rate. In the context of CO<sub>2</sub> hydrogenation, additional challenges arise owing to the simultaneous formation of CO byproducts through the reverse water gas shift reaction (eqn (2)) and/or CH<sub>3</sub>OH decomposition (eqn (3)), along with the CH<sub>3</sub>OH synthesis reaction. Thermodynamically, the dominance of CO byproducts occurs under conditions where the reaction temperature is high.<sup>5</sup> Furthermore, depending on the catalyst, CH<sub>4</sub> by-production via CO<sub>2</sub> methanation (eqn (4)) may also occur. To enhance the efficiency of the entire system, it is necessary to improve the yield of CH<sub>3</sub>OH.



To address this limitation, numerous researchers have explored bifunctional catalysts that incorporate both the CH<sub>3</sub>OH synthesis and conversion reactions. These catalysts facilitate both reactions simultaneously in a one-pass process:

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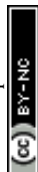
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the CH<sub>3</sub>OH produced is swiftly intercepted by a solid acid catalyst, generating hydrocarbons and oxygenates and effectively suppressing CO byproducts. Notably, integrating CH<sub>3</sub>OH synthesis catalysts with solid acid catalysts substantially enhances the overall yield of CH<sub>3</sub>OH-derived products, including CH<sub>3</sub>OH, hydrocarbons, and oxygenates.<sup>6</sup> However, the considerable difference between the reaction temperatures of these two processes poses a challenge. Thus, optimization of the catalyst design involves either increasing the temperature for CH<sub>3</sub>OH synthesis or reducing the temperature for CH<sub>3</sub>OH conversion.

Metal-oxide catalysts have emerged as promising agents for CO<sub>2</sub>-to-CH<sub>3</sub>OH hydrogenation. Notably, CH<sub>3</sub>OH synthesis over these catalysts occurs at higher temperatures (above 300 °C) than over traditional Cu-based catalysts. The combination of metal oxides and solid acid catalysts enables the one-pass hydrogenation of CO<sub>2</sub> to hydrocarbons and oxygenates. Some reported metal oxide catalysts for high-temperature CH<sub>3</sub>OH synthesis include Zn-doped ZrO<sub>2</sub>,<sup>7–13</sup> In-doped ZrO<sub>2</sub>,<sup>14,15</sup> In<sub>2</sub>O<sub>3</sub>,<sup>14–16</sup> and Ga-doped ZrO<sub>2</sub>,<sup>17</sup> among others. For Zn-doped ZrO<sub>2</sub>, isolated [Zn<sub>o</sub>] clusters activate H<sub>2</sub> molecules, whereas Zr<sup>4+</sup> sites activate CO<sub>2</sub>, emphasizing the crucial role of the interfacial sites between Zn and Zr in high-temperature CH<sub>3</sub>OH synthesis.<sup>8</sup> This reaction over Zn-doped ZrO<sub>2</sub> proceeds *via* a formate route, where the surface formate species on the Zr<sup>4+</sup> sites serve as essential intermediates, eventually hydrogenating to methoxy species on Zr<sup>4+</sup> and CH<sub>3</sub>OH.<sup>7,8,12</sup> In the case of In<sub>2</sub>O<sub>3</sub>, as initially reported by Ye *et al.*,<sup>18</sup> the reaction over In<sub>2</sub>O<sub>3</sub>-based catalysts is expected to follow the formate route.<sup>18–21</sup> However, the mechanism of In<sub>2</sub>O<sub>3</sub>-based catalysts appears to be slightly different from that of Zn-doped ZrO<sub>2</sub> because In<sub>2</sub>O<sub>3</sub>-based catalysts produce a small amount of CH<sub>4</sub>, unlike Zn-doped ZrO<sub>2</sub>.<sup>22</sup> Our study aimed to investigate the differences between the Zn-doped ZrO<sub>2</sub>- and In<sub>2</sub>O<sub>3</sub>-based catalysts. Similar to traditional methanation catalysts, the catalyst may undergo CO<sub>2</sub> methanation, in which a surface formate species is hydrogenated to CO and the resulting CO is then methanated.<sup>23,24</sup> Alternatively, the surface methoxy species or adsorbed CH<sub>3</sub>OH molecules may be directly hydrogenated to CH<sub>4</sub>.<sup>25</sup> In this study, we prepared model catalysts of Zn- and In-doped ZrO<sub>2</sub> and examined the differences in their reaction mechanisms.

## Experimental

### Chemicals

Amorphous ZrO<sub>2</sub> was supplied by Daiichi Kigenso Kagaku Kogyo (product name: NND). Zn(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O, ZnO, In<sub>2</sub>O<sub>3</sub>, and quartz sand were purchased from Fujifilm Wako. In(NO<sub>3</sub>)<sub>3</sub>·*n*H<sub>2</sub>O was purchased from Kanto Chemical.

### Catalyst preparation

ZnZrO<sub>x</sub> and InZrO<sub>x</sub> were prepared using an incipient wetness impregnation method. The amorphous ZrO<sub>2</sub> was impregnated with a metal nitrate aqueous solution, dried at 110 °C for 12 h,

**Table 1** Abbreviation, composition, specific surface area, and total pore volume of ZnZrO<sub>x</sub> and InZrO<sub>x</sub>

Catalyst	Abbreviation	Zn/Zr molar ratio <sup>a</sup> [–]	In/Zr molar ratio <sup>a</sup> [–]	SSA <sup>b</sup> [m <sup>2</sup> g <sup>–1</sup> ]	Vm <sup>c</sup> [10 <sup>–2</sup> cm <sup>3</sup> g <sup>–1</sup> ]
ZnZrO <sub>x</sub>	ZnZr46	0.46		40	0.14
InZrO <sub>x</sub>	InZr09		0.087	82	0.24
	InZr22		0.22	66	0.15

<sup>a</sup> Determined by XRF. <sup>b</sup> Specific surface area determined by N<sub>2</sub> adsorption. <sup>c</sup> Total pore volume determined by N<sub>2</sub> adsorption. *p*/*p*<sub>0</sub> = 0.98.

and calcined at 500 °C for 3 h. The abbreviations and compositions of the prepared catalysts are summarized in Table 1.

### Characterization

X-ray fluorescence (XRF), X-ray diffraction (XRD), N<sub>2</sub> adsorption and desorption, X-ray absorption spectroscopy (XAS), and scanning transmission electron microscopy (STEM) analyses were performed. The detailed procedure is summarized in the ESI.†

### Performance test

**CO<sub>2</sub> hydrogenation.** A stainless-steel fixed-bed tubular reactor was used for CO<sub>2</sub> hydrogenation performance tests. Each catalyst was compressed at 20 MPa to form tablets. The tablet was crushed and sieved to obtain catalyst grains of 1.18–1.7 mm. The grains (300 mg) were placed in the catalyst bed. The reaction temperature was controlled using a K-type thermocouple inserted into the catalyst bed, and an electric furnace installed outside the catalyst bed. A reaction gas mixture (CO<sub>2</sub>/H<sub>2</sub>/N<sub>2</sub> = 1/3/1) was introduced into the catalyst bed at a flow rate of 30 mL<sub>STP</sub> min<sup>–1</sup>. The reaction pressure was set to 1.0 MPa because it is restricted to 1.0 MPa by the High-Pressure Gas Safety Act of Japan. The outlet gas composition was analyzed using a gas chromatograph (SHIMADZU GC-2014) with a thermal conductivity detector (TCD) and flame ionization detector (FID). According to eqn (5)–(7), the CO<sub>2</sub> conversion, space-time yield of species *i* [mL min<sup>–1</sup> g<sup>–1</sup>], and yield of species *i* were obtained.

$$\text{CO}_2 \text{ conversion} = \frac{F_{\text{in,CO}_2} - F_{\text{out,CO}_2}}{F_{\text{in,CO}_2}} \quad (5)$$

$$\text{Space – time yield of } i \text{ species} = \frac{F_{\text{out},i}}{w} \quad (6)$$

$$\text{Yield of } i \text{ species} = \frac{F_{\text{out},i}}{F_{\text{in,CO}_2}} \quad (7)$$

In these equations, *F*<sub>in,*i*</sub> and *F*<sub>out,*i*</sub> are the inlet and outlet gas flow rates of species *i* [mL<sub>STP</sub> min<sup>–1</sup>], respectively, and *w* is the catalyst weight [g]. The carbon yield was 100% ± 0.5%.

**Methanol decomposition.** A silicate glass fixed-bed tubular reactor (i.d. 7.5 mm) was used for CH<sub>3</sub>OH decomposition performance tests. Each catalyst was compressed at 25 MPa to form tablets. The tablet was crushed and sieved to obtain catalyst grains of 0.85–1.4 mm. Grains (500 mg) were placed in a catalyst bed. The reaction temperature was controlled using a K-type thermocouple, which was inserted into the catalyst bed,



and an electric furnace was installed outside the catalyst bed. Liquid CH<sub>3</sub>OH, supplied by a high-performance liquid chromatography pump (SHIMADZU LC-20AT) with a constant flow rate (0.038 mL min<sup>-1</sup> as a liquid, corresponding to 21.06 mL<sub>STP</sub> min<sup>-1</sup> as a gas), was introduced into a vaporizer heated at 130 °C with N<sub>2</sub> (flow rate: 110 mL<sub>STP</sub> min<sup>-1</sup>) and H<sub>2</sub> (flow rate: 38 mL<sub>STP</sub> min<sup>-1</sup>) streams. The gas mixture was fed into the catalyst bed. The reaction was conducted under atmospheric pressure. The outlet gas composition was analyzed using gas chromatographs equipped with a TCD (SHIMADZU GC-8A) for CO quantification and an FID (SHIMADZU GC-2014) for CH<sub>4</sub> and dimethyl ether (CH<sub>3</sub>OCH<sub>3</sub>) quantification. According to eqn (8), the space-time yield of the product of species *i* (CH<sub>4</sub>, CO, and CH<sub>3</sub>OCH<sub>3</sub>) was obtained.

$$\text{Space-time yield of } i \text{ species} = \frac{F_{\text{out},i} \times N}{W} \quad (8)$$

where  $F_{\text{out},i}$  is the outlet gaseous flow rate of *i* species [mL<sub>STP</sub> min<sup>-1</sup>], and *N* is the number of carbon atoms in a molecule of each product.

**Calculation.** We performed spin-polarized density functional theory (DFT) calculations using the vienna *ab initio* simulation program,<sup>26,27</sup> which adopts the projector-augmented wave method.<sup>28</sup> We used the generalized-gradient approximation for DFT calculations with the Perdew–Burke–Ernzerhof<sup>29</sup> exchange–correlation functional. The plane-wave cut-off was 400.0 eV unless otherwise noted. The total energy for each self-consistent field step converged to an accuracy of 10<sup>-5</sup> eV. Gaussian smearing was used with a width of 0.1 eV.

We calculated the electronic energy under 3D periodic boundary conditions. The optimization of a unit cell of bulk tetragonal (*t*)-ZrO<sub>2</sub> results in cell lengths of  $a = b = 3.623$  Å and  $c = 5.280$  Å using a larger plane-wave cutoff of 600.0 eV, which agrees with previous studies.<sup>30,31</sup> The initial configuration of the *t*-ZrO<sub>2</sub> (101) surface was made by cleaving the optimized bulk ZrO<sub>2</sub>. Along the surface parallel direction, the cell lengths are 12.806 Å × 10.868 Å, and the employed Monkhorst–Pack<sup>32</sup> *k*-point mesh is 2 × 2. Along the surface normal direction, the cleaved surface has two ZrO<sub>2</sub> layers (72 atoms) to follow the computational model in ref. 31 because ref. 31 (two-layer model) and ref. 7 (three-layer model) yielded comparable energy profiles (the validity of our model is discussed in the ESI†); the width of the vacuum region is 15 Å, and one *k*-point is used. During the geometry optimization, the coordinates of the atoms, except for the bottom layer, were optimized. An electric dipole correction was considered.<sup>33</sup>

The InZrO<sub>x</sub> surface was prepared by removing one O atom and replacing one Zr atom with an In atom from the top layer of a stoichiometric ZrO<sub>2</sub> surface. The Zr atoms on the surface are classified into two types, and the replaced Zr atom acts as a strong Lewis site with longer Zr–O bonds.<sup>34,35</sup> Fig. S1 (ESI†) shows the positions of In and one oxygen vacancy (V<sub>O</sub>). For comparison, ZnZrO<sub>x</sub> was prepared using the same method. The V<sub>O</sub> was situated next to the In atom, similar to a previous study of the ZnZrO<sub>x</sub> surface.<sup>31</sup> The geometries of the surfaces were optimized, and the adsorbates were placed on the

optimized surfaces. The adsorption energies are calculated using the following eqn (9):

$$E_{\text{ads}} = E(\text{substrate} + \text{adsorbate}) - E(\text{substrate}) - E(\text{adsorbate}) \quad (9)$$

where  $E(\text{substrate} + \text{adsorbate})$ ,  $E(\text{substrate})$ , and  $E(\text{adsorbate})$  are the energies of the adsorbed complex, the substrate, and the adsorbate, respectively. According to this definition, a lower adsorption energy indicates stable adsorption.

The climbing-image nudged elastic band method<sup>36</sup> was used to calculate the potential surface and activation energies. Free energy (*G*) was computed using the normal mode analysis. In the calculations of the translational, rotational, and vibrational enthalpic and entropic contributions, the gas-phase molecules were approximated as ideal gases. The further details are summarized in the ESI.†

## Results and discussion

### CO<sub>2</sub> hydrogenation

We conducted CO<sub>2</sub> hydrogenation performance tests using ZnZr46, InZr9, and InZr22. We have previously reported that the Zn/Zr ratio of ZnZrO<sub>x</sub> does not affect its catalytic performance on CO<sub>2</sub>-to-CH<sub>3</sub>OH hydrogenation. Therefore, this time, we adopted ZnZr46 as a representative ZnZrO<sub>x</sub>.<sup>8</sup> Fig. S2 (ESI†) shows plots of CO<sub>2</sub> conversion at different reaction temperatures. At 250 °C, CO<sub>2</sub> conversion is approximately 1%, irrespective of the catalyst type. For ZnZr46, the CO<sub>2</sub> conversion gradually increases from 1% to 4% with increasing reaction temperature. For InZr09 and InZr22, the increase in CO<sub>2</sub> conversion is more substantial. Particularly, at 325 °C, the CO<sub>2</sub> conversion for InZr09 and InZr22 is more than twice as high as that of ZnZr46, reaching 8%.

Fig. 1 illustrates the space-time yields (STY) of CH<sub>3</sub>OH, CO, and CH<sub>4</sub> for ZnZr46, InZr09, and InZr22, with the right axes indicating the yields. In this experiment, the identified carbon-containing products are limited to CH<sub>3</sub>OH, CO, and CH<sub>4</sub>; therefore, the total STY corresponds to the CO<sub>2</sub> conversion rate, as shown in Fig. S2 (ESI†). For ZnZr46, an increase in the reaction temperature enhances the STY of CH<sub>3</sub>OH and CO. At

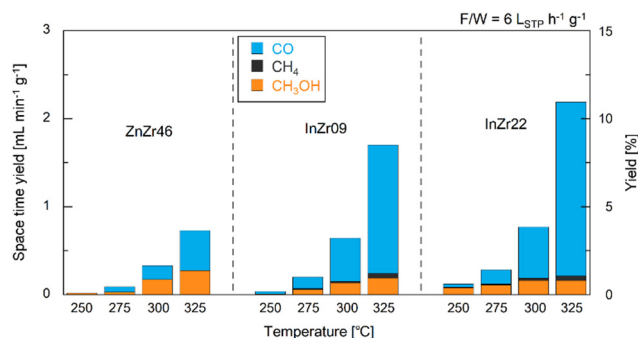


Fig. 1 Space–time yields and yields of CH<sub>3</sub>OH, CO, and CH<sub>4</sub> over ZnZr46, InZr09, and InZr22 for CO<sub>2</sub> hydrogenation. F/W = 6 L<sub>STP</sub> h<sup>-1</sup> g<sup>-1</sup>. Pressure = 1.0 MPa. Gas composition: CO<sub>2</sub>/H<sub>2</sub>/N<sub>2</sub> = 1/3/1.





325 °C, the STY of CH<sub>3</sub>OH and CO reach 0.27 and 0.46 mL min<sup>-1</sup> g<sup>-1</sup>, respectively. CH<sub>4</sub> production is not observed. According to our previous study,<sup>37</sup> we have confirmed that below 400 °C, Zn-doped ZrO<sub>2</sub> produces CH<sub>3</sub>OH and CO, but not CH<sub>4</sub>. For comparison, we also assessed the performance of commercially available ZnO particles at 325 °C. However, no CH<sub>3</sub>OH is observed (Fig. S3a, ESI†). This result indicates that ZnZr46, unlike ZnO, possesses active sites that are effective for CH<sub>3</sub>OH production.

In InZr09, the CH<sub>3</sub>OH STY is slightly lower than that of ZnZr46, whereas the CO STY is higher (Fig. 1). The increased CO<sub>2</sub> conversion rate (Fig. S2, ESI†) of InZr09 is attributed to the enhanced generation of CO. Additionally, the production of CH<sub>4</sub> is observed. When comparing InZr09 and InZr22, no significant differences are observed in the STY trends. Additionally, the performance of commercially available In<sub>2</sub>O<sub>3</sub> is evaluated at 325 °C (Fig. S3a, ESI†). Despite its lower space velocity, the yield of In<sub>2</sub>O<sub>3</sub> is lower than that of InZr09 and InZr22, suggesting the presence of additional active sites on InZr09 and InZr22 that activate CO<sub>2</sub>. However, these sites are not present in In<sub>2</sub>O<sub>3</sub>. Notably, In<sub>2</sub>O<sub>3</sub> generates CO and CH<sub>3</sub>OH while producing a negligible amount of CH<sub>4</sub>. Therefore, In<sub>2</sub>O<sub>3</sub> does not possess active sites that generate CH<sub>4</sub>, whereas InZr09 and InZr22 possess active sites that facilitate CH<sub>4</sub> production.

## Structure

In this study, we examined the catalyst structure to reveal the role of active sites. First, we measured the specific surface area (SSA) and total pore volume by N<sub>2</sub> adsorption (Table 1). The SSA of ZnZr46, InZr09, and InZr22 is lower than that of amorphous ZrO<sub>2</sub> (approximately 200 m<sup>2</sup> g<sup>-1</sup>),<sup>35</sup> which is the raw material. This result was attributed to sintering at 500 °C during catalyst preparation. Fig. 2 shows the high-angle annular dark field scanning transmission electron microscopy (HAADF-STEM) images of ZnZr46, InZr09, and InZr22. Regardless of the catalyst type, all elements are distributed uniformly. The size of the observed particles is approximately 15 nm. Consequently, no differences are observed between the morphologies of the two InZrO<sub>x</sub> catalysts.

The crystal structure was examined using X-ray diffraction (XRD) (Fig. 3 and Fig. S3b, ESI†). For amorphous ZrO<sub>2</sub>, a broad peak is observed at approximately 31°, indicating a lack of a distinct crystal structure. For ZnZr46, strong peaks are observed at 31°, 35°, and 51°, which correspond to *t*-ZrO<sub>2</sub>. In contrast to the conventional anticipation of the formation of monoclinic (*m*-)ZrO<sub>2</sub> from amorphous ZrO<sub>2</sub> at 500 °C, calcination of amorphous ZrO<sub>2</sub> results in the formation of *t*-ZrO<sub>2</sub>. This unexpected outcome signifies the incorporation of Zn<sup>2+</sup> species into the ZrO<sub>2</sub> lattice, leading to the establishment of a solid solution of ZnZrO<sub>x</sub> with a tetragonal crystal structure, which agrees with our previously reported observations.<sup>8</sup> The doped Zn<sup>2+</sup> species coordinated with the O atoms to form isolated [ZnO<sub>a</sub>] clusters (refer to ESI† and Fig. S4). The additional peaks at 32°, 34°, 36°, 48°, and 56° indicate the presence of ZnO. Therefore, ZnO precipitates when the Zn/Zr molar ratio exceeds 0.19 because the ZnZrO<sub>x</sub> solid solution reaches its limit when additional Zn species are incorporated.<sup>8</sup> Consequently, ZnZr46 contains both

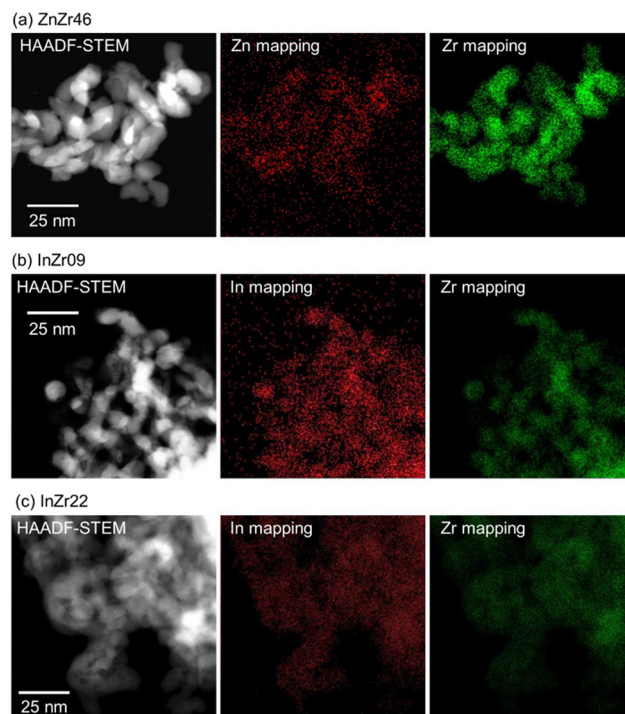


Fig. 2 HAADF-STEM images and elemental maps of (a) ZnZr46, (b) InZr09, and (c) InZr22.

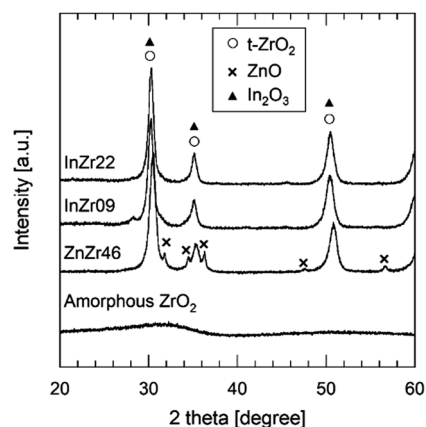
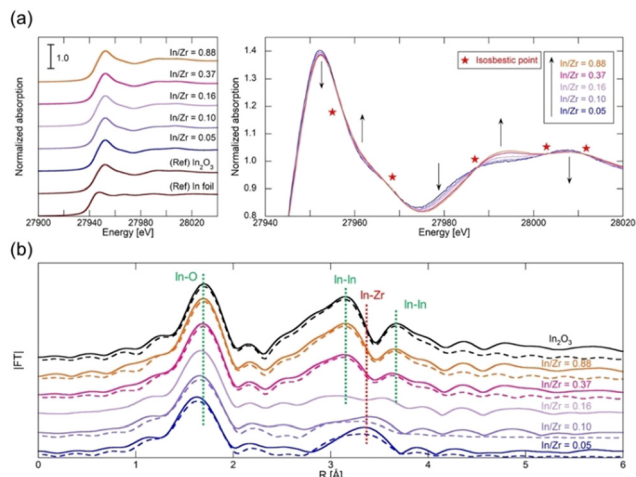


Fig. 3 XRD patterns of amorphous ZrO<sub>2</sub>, ZnZr46, InZr09, and InZr22.

the ZnZrO<sub>x</sub> solid solution and ZnO. Previous observations indicate minimal differences in catalyst performance among ZnZrO<sub>x</sub> samples with different Zn/Zr molar ratios, implying that the [ZnO<sub>a</sub>] isolated clusters on the ZnZrO<sub>x</sub> solid solution serve as crucial active sites, whereas the ZnO particles remain inert for CO<sub>2</sub> hydrogenation to methanol.<sup>8</sup>

The XRD patterns of InZr09 and InZr22 (Fig. 3) reveal peaks exclusively corresponding to *t*-ZrO<sub>2</sub>. The presence of *t*-ZrO<sub>2</sub> instead of *m*-ZrO<sub>2</sub> suggests the incorporation of In species into the ZrO<sub>2</sub> lattice, resulting in the formation of an InZrO<sub>x</sub> solid solution. Notably, the peaks derived from tetragonal ZrO<sub>2</sub> and In<sub>2</sub>O<sub>3</sub> coincide, making it difficult to unequivocally confirm the





**Fig. 4** (a) K-edge XANES spectra of In foil,  $\text{In}_2\text{O}_3$ , and  $\text{InZrO}_x$  with In/Zr molar ratios of 0.05, 0.10, 0.16, 0.37, and 0.88. (b) Fourier transforms of  $k^3$ -weighted extended X-ray absorption fine structure (EXAFS) oscillations measured at room temperature near the In K-edge of  $\text{InZrO}_x$ .  $k$  range: 3–12  $\text{\AA}^{-1}$ .  $R$  range: 1–4  $\text{\AA}$ . Solid line: experimental; dotted line: EXAFS model.

existence of  $\text{In}_2\text{O}_3$  through XRD analysis alone. As elucidated earlier, the presence of the In species is conclusively determined using XRF (Table 1).

We conducted K-edge XAS measurements on various  $\text{InZrO}_x$  samples with different In/Zr molar ratios (0.05, 0.10, 0.16, 0.37, and 0.88) to explore the structural characteristics of the In species. Our primary focus was to determine the relationship between the In/Zr ratio and coordination structure of In. Fig. 4a shows the In K-edge X-ray absorption near edge structure (XANES) spectra, with In foil and  $\text{In}_2\text{O}_3$  included as the reference materials. The  $\text{InZrO}_x$  spectra closely resemble those of  $\text{In}_2\text{O}_3$ , characterized by a distinctive white line at 27 952 eV. A detailed examination of the  $\text{InZrO}_x$  XANES spectrum reveals several isosbestic points (Fig. 4a), indicating the presence of two distinct In species in  $\text{InZrO}_x$ . The relative proportions of the In species also change with variations in the In:Zr ratio. In Fig. 4b, the In K-edge radial structure functions (RSFs) of  $\text{In}_2\text{O}_3$  and  $\text{InZrO}_x$  are shown, and Table 2 summarizes the fitting

parameters. The RSF of  $\text{In}_2\text{O}_3$  exhibits three peaks, representing the nearest-neighbor In–O distance (at 1.7  $\text{\AA}$ ) and the nearest-neighbor and second-nearest In–In distances (at 3.1 and 3.7  $\text{\AA}$ ). For  $\text{InZrO}_x$  at In/Zr ratios of 0.88 and 0.37, the RSFs exhibit these characteristic peaks. The In–O peaks shift to the left as the ratio decreases from 0.37 to 0.05, indicating a gradual change in the In coordination structure. Simultaneously, the In–In peaks at 3.1 and 3.7  $\text{\AA}$  gradually disappear, and a new peak emerges at 3.3  $\text{\AA}$ . The peak at 3.3  $\text{\AA}$  corresponds to the nearest-neighbor In–Zr distance, determined from the optimized  $\text{InZrO}_x$  structure (Fig. S1, ESI†). In the sample with an In/Zr ratio of 0.16, the peaks derived from In–In and In–Zr coexist, making it difficult to analyze the coordination structure of the In species by fitting. Consequently, when the In/Zr ratio is low,  $\text{InZrO}_x$  predominantly contains  $\text{In}^{3+}$  species doped into  $t\text{-ZrO}_2$ . Conversely, when the In/Zr ratio is high,  $\text{InZrO}_x$  primarily contains  $\text{In}_2\text{O}_3$ . These results suggest that the solid solubility limit of the In species is approximately 0.10.

As detailed in the  $\text{CO}_2$  hydrogenation section,  $\text{InZr09}$  and  $\text{InZr22}$  exhibit distinct active sites that are absent in  $\text{In}_2\text{O}_3$ . The catalytic performances of  $\text{InZr09}$  and  $\text{InZr22}$  are notably similar, as observed for the STYs of  $\text{CH}_3\text{OH}$ ,  $\text{CO}$ , and  $\text{CH}_4$  (Fig. 1). XAS measurements (Table 2) further reveal that when the In/Zr ratio is below 0.10, the prevalent In species in  $\text{InZrO}_x$  are primarily In species doped with  $t\text{-ZrO}_2$ . Collectively, these findings strongly suggest that the catalytic performances of  $\text{InZr09}$  and  $\text{InZr22}$  are governed by the presence of doped In species.

### Reaction mechanism analysis by DFT

As described earlier, it is evident that the sites responsible for producing  $\text{CH}_3\text{OH}$  and  $\text{CH}_4$  are Zn and In species incorporated into  $\text{ZrO}_2$ . Through DFT calculations, we aimed to delve into reactions at the atomic level, which are challenging to observe experimentally.

Fig. 5 shows the free energy surfaces for  $\text{CH}_4$  and  $\text{CH}_3\text{OH}$  formation on the  $\text{InZrO}_x$  surface. Fig. 6 and 7 show the structures of the reaction intermediates in the formate,  $\text{CO}$ ,

**Table 2** Parameters calculated by fitting the EXAFS signals of  $\text{In}_2\text{O}_3$  and  $\text{InZrO}_x$  shown in Fig. 4b

Sample	Path	CN	$\sigma^2$ [ $\text{\AA}^2$ ]	$\Delta E_0$ [eV]	$R$ [ $\text{\AA}$ ]	$R$ factor
$\text{In}_2\text{O}_3$	In–O	$5.5 \pm 0.3$	$0.007 \pm 0.001$	$8.0 \pm 0.7$	$2.16 \pm 0.01$	0.002
	In–In	$5.1 \pm 0.8$	$0.005 \pm 0.001$	$7.0 \pm 0.9$	$3.35 \pm 0.01$	
	In–In	$3.8 \pm 1.2$	$0.005 \pm 0.002$	$7.2 \pm 1.9$	$3.84 \pm 0.01$	
$\text{InZrO}_x$ In/Zr = 0.88	In–O	$6.2 \pm 0.5$	$0.007 \pm 0.001$	$7.2 \pm 0.9$	$2.16 \pm 0.01$	0.003
	In–In	$5.2 \pm 1.2$	$0.006 \pm 0.001$	$5.8 \pm 1.4$	$3.36 \pm 0.01$	
	In–In	$3.6 \pm 1.8$	$0.006 \pm 0.003$	$6.2 \pm 3.1$	$3.83 \pm 0.02$	
$\text{InZrO}_x$ In/Zr = 0.37	In–O	$6.1 \pm 0.4$	$0.007 \pm 0.001$	$6.8 \pm 0.7$	$2.16 \pm 0.01$	0.003
	In–In	$4.3 \pm 0.9$	$0.006 \pm 0.001$	$5.0 \pm 1.3$	$3.36 \pm 0.01$	
	In–In	$2.7 \pm 1.3$	$0.005 \pm 0.003$	$4.9 \pm 3.0$	$3.83 \pm 0.02$	
$\text{InZrO}_x$ In/Zr = 0.10	In–O	$5.0 \pm 0.7$	$0.007 \pm 0.001$	$7.5 \pm 1.6$	$2.14 \pm 0.01$	0.034
	In–Zr	$8.1 \pm 3.8$	$0.017 \pm 0.004$	$5.2 \pm 2.8$	$3.53 \pm 0.02$	
	In–In	$2.7 \pm 1.3$	$0.005 \pm 0.003$	$4.9 \pm 3.0$	$3.83 \pm 0.02$	
$\text{InZrO}_x$ In/Zr = 0.05	In–O	$4.5 \pm 0.5$	$0.007 \pm 0.001$	$6.2 \pm 1.4$	$2.12 \pm 0.01$	0.022
	In–Zr	$6.7 \pm 2.1$	$0.013 \pm 0.002$	$6.4 \pm 2.1$	$3.57 \pm 0.02$	

Notation: CN, coordination number;  $\sigma$ , Debye–Waller factor;  $\Delta E_0$ , increase in the threshold energy;  $R$ , distance. Confidence intervals = 68%.

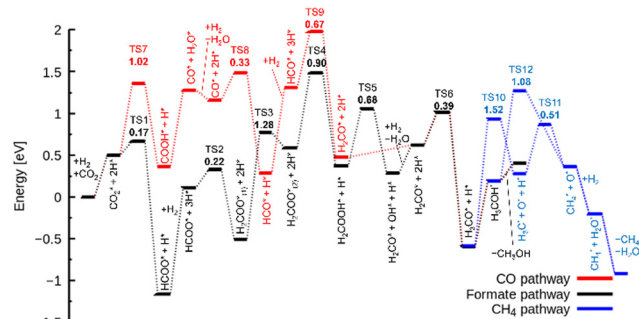


Fig. 5 Free energy surfaces of the InZrO<sub>x</sub> surface for formate (black), CO (red), and CH<sub>4</sub> (blue) pathways. The values in the bold face and \* indicate the activation free energies of the TSs and adsorbed species, respectively.

and CH<sub>4</sub> pathways, respectively. Fig. S5 (ESI<sup>†</sup>) shows the structures of transition states (TSs) in Fig. 5.

In the formate pathway, CO<sub>2</sub> adsorbs onto the surface Zr atoms as a bidentate carbonate species, whereas H<sub>2</sub> dissociates and adsorbs onto the surface Zr and O atoms. The adsorbed CO<sub>2</sub> is protonated by the neighboring H atom to form HCOO (formate). The formate group receives a H atom to form H<sub>2</sub>COO. One O atom in H<sub>2</sub>COO changes its bond with the surface Zr atom to an In atom with the largest barrier ( $\Delta G^\ddagger$ ) of 1.28 eV in the formate pathway. The O atom on the In atom in H<sub>2</sub>COO accepts the H atom to form H<sub>2</sub>COOH, which then decomposes into H<sub>2</sub>CO (formaldehyde), OH, and H species. Protonation of formaldehyde yields a methoxy group (H<sub>3</sub>CO), which is converted into CH<sub>3</sub>OH by further protonation.

In the CO pathway, the H atom on the In atom hydrogenates adsorbed CO<sub>2</sub> to yield COOH (carboxyl group). This carboxyl group decomposes into CO and water. The CO molecule successively changes to HCO, H<sub>2</sub>CO (formaldehyde) and finally, to H<sub>3</sub>CO to join the formate pathway. In Fig. 1, CO production increases at higher temperatures partly because CO<sub>2</sub> hydrogenation to CH<sub>3</sub>OH is an exothermic reaction while the reverse

water gas shift reaction is an endothermic reaction, making the chemical equilibrium shift to the right side in CO<sub>2</sub> + H<sub>2</sub>  $\rightleftharpoons$  CO + H<sub>2</sub>O and enhancing CO production at higher reaction temperatures. However, it is noted that we showed in Fig. S6 (ESI<sup>†</sup>) by changing the space velocity that CH<sub>3</sub>OH can be converted to CO, contributing to CO production. Fig. 5 shows that the decomposition of COOH to CO and H<sub>2</sub>O, which undergoes also at zirconia surfaces,<sup>38</sup> is endothermic by 0.92 eV, and it is found to be barrierless. This suggests that after CO is formed from the decomposition of COOH, the backward reaction can proceed without a barrier, allowing CO to readily transition back to a stable COOH. Additionally, the adsorption energy in Table S1 (ESI<sup>†</sup>) is the highest for a CO molecule among the reaction intermediates. This indicates that a CO molecule can desorb from the catalyst surface easier than the other intermediates,<sup>7</sup> and the CO pathway does not proceed further.

In the CH<sub>4</sub> pathway, Fig. 5 suggests two possible routes: decompositions of methoxy group and CH<sub>3</sub>OH. In the case of methoxy group decomposition, the methyl group in the methoxy group is transferred to the surface of the In atom for stable adsorption, with a C–In distance of 2.69 Å. However, this adsorption is unstable on the Zn atom of the ZnZrO<sub>x</sub> surface, and the methyl group changes to a methoxy group (Fig. S7, ESI<sup>†</sup>) because the methyl group forms a bond with the O atom next to Zn. The barrier of the methyl group transfer is 1.52 eV on the InZrO<sub>x</sub> surface, which is higher than those in the other steps of this study, such as 1.28 eV in the formate pathway. It is also higher than those in CH<sub>3</sub>OH formation from CO<sub>2</sub> on ZnZrO<sub>x</sub> surfaces.<sup>7,31</sup> This suggests that CH<sub>4</sub> production is unfavored compared to CH<sub>3</sub>OH. This result is consistent with the lower experimental yield of CH<sub>4</sub> than that of CH<sub>3</sub>OH. Therefore, methyl group transfer is the rate-determining step in CH<sub>4</sub> formation. This result supports the experimental evidence of CH<sub>4</sub> formation on the InZrO<sub>x</sub> surface, indicating that stable adsorption of a methyl group on the InZrO<sub>x</sub> surface is important for CH<sub>4</sub> production.

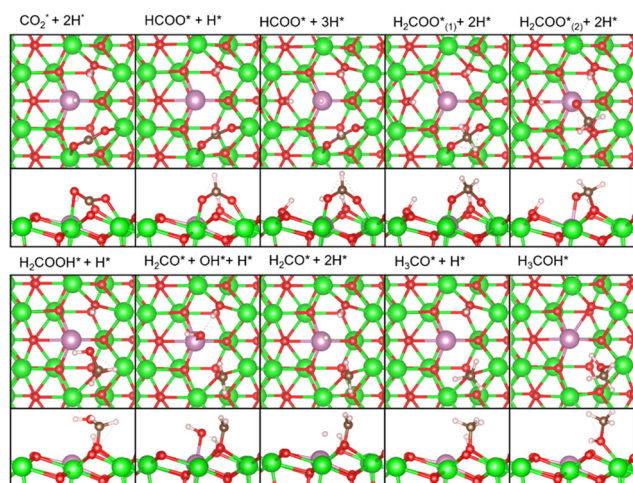


Fig. 6 Top and side views of the reaction intermediates in the formate pathway of Fig. 5. Color code: green: Zr, red: O, purple: In, white: H, and brown: C. The figures are made by VESTA.<sup>39</sup>

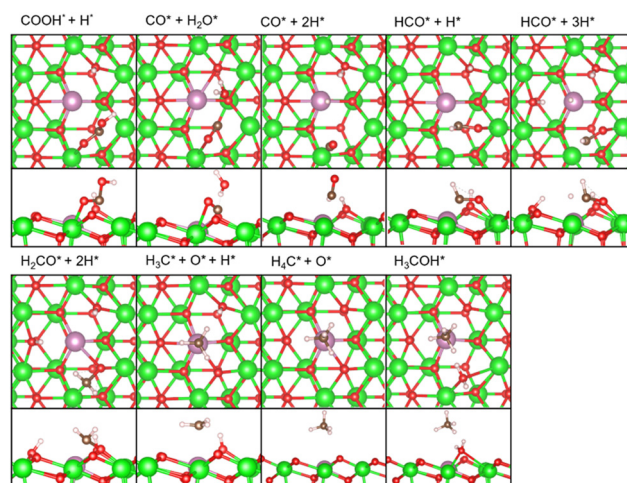


Fig. 7 Top and side views of the reaction intermediates in the CO and CH<sub>4</sub> pathways in Fig. 5. Color code: green: Zr, red: O, purple: In, white: H, and brown: C. The figures are made by VESTA.<sup>39</sup>





The methyl group adsorbed on the In atom accepts a proton from the neighboring OH group to form CH<sub>4</sub>. The remaining O atom adsorbed at the V<sub>O</sub> site combines with the H<sub>2</sub> molecule to form H<sub>2</sub>O. The H<sub>2</sub>O and CH<sub>4</sub> molecules finally desorb from the surface into the gas phase, completing CH<sub>4</sub> formation. In the last step, the InZrO<sub>x</sub> surface regains the surface V<sub>O</sub> sites.

In the case of CH<sub>3</sub>OH decomposition (CH<sub>3</sub>OH → CH<sub>4</sub> + O\*) without the transfer of the methyl group on the InZrO<sub>x</sub> surface, the activation free energy is 1.08 eV that is lower than that in the case of methoxy group decomposition, yet CH<sub>3</sub>OH desorbs with the free energy difference of only 0.21 eV. This suggests that CH<sub>3</sub>OH decomposition can be competitive but is unfavored over CH<sub>3</sub>OH desorption. It is noted that the decomposition of a CH<sub>3</sub>OH molecule on the ZnZrO<sub>x</sub> surface requires a higher barrier (2.60 eV) as illustrated in Fig. S8 (ESI<sup>†</sup>), indicating that CH<sub>4</sub> production on the ZnZrO<sub>x</sub> surface is unfavored.

Here, we discuss the density of states (DOS) shown in Fig. 8 to understand the adsorption of methyl groups on the InZrO<sub>x</sub> surface. The mid-gap (also called the impurity) band for the InZrO<sub>x</sub> surface in Fig. 8a lies between the valence band maximum of ZrO<sub>2</sub> and the conduction band minimum. The mid-gap band primarily arises from the In atom, and the *s* and 5*p<sub>z</sub>* orbitals mainly contribute to the DOS of the In atom, as shown in Fig. 8d (the *z* coordinate is along the surface-normal direction). In contrast, the DOS of the ZnZrO<sub>x</sub> surface in Fig. 8b has no midgap band. The *d* band of a Zn atom in Fig. 8e dominates the DOS of the Zn atom between −10 and 0 eV.

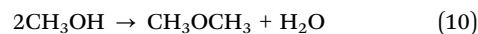
The projected DOSs of In and C atoms on the InZrO<sub>x</sub> surface with an adsorbed methyl group in Fig. 8c also shows a mid-gap

state. Fig. 8f shows the midgap state consisting of the 5*p<sub>z</sub>* orbital of the In atom and 2*p<sub>z</sub>* orbital of the C atom. Therefore, the 5*p<sub>z</sub>* orbital of the In atom and the 2*p<sub>z</sub>* orbital of the C atom overlapped with the adsorbed methyl group on the InZrO<sub>x</sub> surface, and the distinct band structures of the doped metals determined the adsorption strength of the methyl group on the InZrO<sub>x</sub> and ZnZrO<sub>x</sub> surfaces. A methyl group was readily adsorbed on the In atom on the InZrO<sub>x</sub> surface owing to the superior overlap of orbitals; however, the adsorption of a methyl group on the Zn atom on the ZnZrO<sub>x</sub> surface is unstable, as shown in Fig. S7 (ESI<sup>†</sup>). Therefore, In-doping of ZrO<sub>2</sub> enhances CH<sub>4</sub> formation.

### CH<sub>3</sub>OH conversion

We conducted a test of the CH<sub>3</sub>OH conversion reaction to directly observe how each catalyst activates the CH<sub>3</sub>OH molecules (Fig. 9). Taking into consideration the reaction condition of CO<sub>2</sub> hydrogenation, unreacted H<sub>2</sub> is assumed to coexist with CH<sub>3</sub>OH during CO<sub>2</sub> hydrogenation. Hence, for the CH<sub>3</sub>OH conversion test, CH<sub>3</sub>OH and H<sub>2</sub> were supplied simultaneously.

ZnZr46 generates CO and CH<sub>3</sub>OCH<sub>3</sub> from CH<sub>3</sub>OH. The former is attributed to the product of the CH<sub>3</sub>OH decomposition reaction, which represents the reverse of the CH<sub>3</sub>OH synthesis *via* CO hydrogenation (eqn (3)). In contrast, the latter is related to CH<sub>3</sub>OH dehydration (eqn (10)) at the Brønsted acid site of the catalyst.



Apart from ZrO<sub>2</sub>, which possesses acid sites, Brønsted acid sites are formed by doping metal into ZrO<sub>2</sub>.<sup>40</sup> Moreover, the CH<sub>4</sub> production is minimal, with a selectivity below 1%. Even when using commercially available ZnO, the CO selectivity (93%) is much higher than that for CH<sub>4</sub> (3%) and CH<sub>3</sub>OCH<sub>3</sub> (4%). For InZr22, along with CO and CH<sub>3</sub>OCH<sub>3</sub>, a significant amount of CH<sub>4</sub> is generated (selectivity = 15%). Notably, the results for In<sub>2</sub>O<sub>3</sub> revealed a CH<sub>4</sub> selectivity of 65%, although the

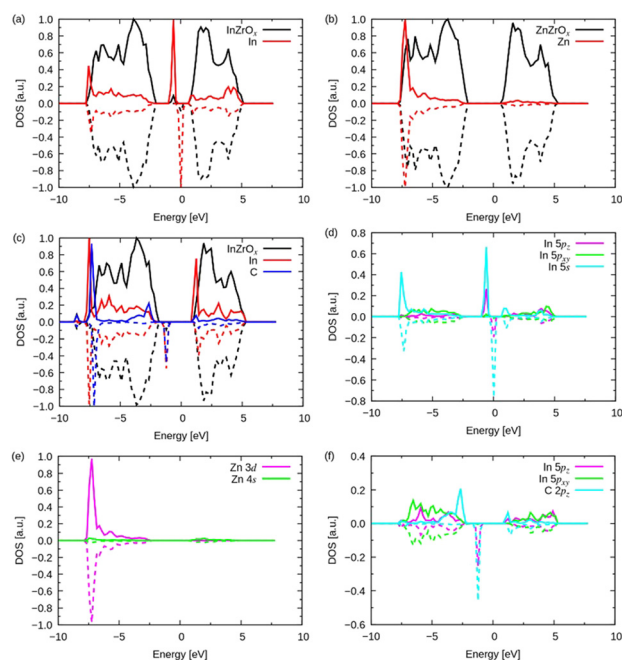


Fig. 8 Density of states (DOS) of neat (a) InZrO<sub>x</sub>, (b) ZnZrO<sub>x</sub> surfaces, and (c) methyl groups on the InZrO<sub>x</sub> surface. The solid and dashed lines denote the up and down spins, respectively. Each DOS is normalized for clarity. (d)–(f) Projected DOSs corresponding to panels (a)–(c), where the *z* coordinate is along the surface normal direction.

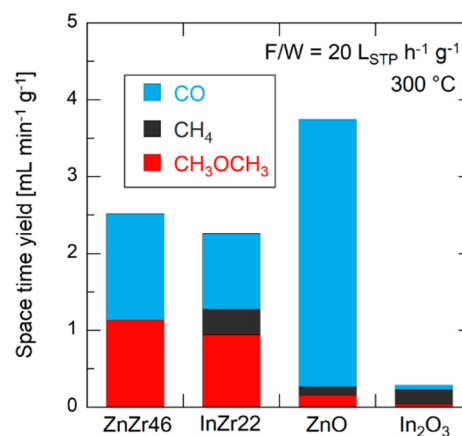


Fig. 9 Space time yields of CO, CH<sub>4</sub>, and CH<sub>3</sub>OCH<sub>3</sub> over ZnZr46, InZr22, ZnO, and In<sub>2</sub>O<sub>3</sub> for the methanol conversion reaction. F/W = 20 L<sub>STP</sub> h<sup>−1</sup> g<sup>−1</sup>. Temperature = 300 °C. Pressure = 0.1 MPa. Gas composition: CH<sub>3</sub>OH/H<sub>2</sub>/N<sub>2</sub> = 13/22/65.



yield is extremely small compared to that of the other catalysts. These results align with the computational chemistry predictions, indicating that  $\text{CH}_3\text{OH}$  is directly converted to  $\text{CH}_4$  on  $\text{In}^{3+}$ . As shown in Fig. S3 (ESI<sup>†</sup>), during the  $\text{CO}_2$  hydrogenation reaction with  $\text{In}_2\text{O}_3$ ,  $\text{CH}_3\text{OH}$  is observed, but  $\text{CH}_4$  remains undetected. Therefore, the results in Fig. S3 (ESI<sup>†</sup>) and Fig. 9 appear to be contradictory. We expect that the absence of  $\text{CH}_4$  in the  $\text{CO}_2$  hydrogenation reaction with  $\text{In}_2\text{O}_3$  is linked to the inhibition of the In-methoxy bond formation owing to the presence of  $\text{CO}_2$  in the reaction gas. However, we were unable to clarify these differences beyond the reaction tests. We plan to explore and elucidate this distinction in future studies.

## Conclusions

This study focused on oxide solid-solutions of  $\text{ZnZrO}_x$  and  $\text{InZrO}_x$  for  $\text{CO}_2$ -to- $\text{CH}_3\text{OH}$  hydrogenation and clarified the differences in their reaction mechanisms. For both catalysts, the metal cations ( $\text{In}^{3+}$  and  $\text{Zn}^{2+}$ ) incorporated into the  $t\text{-ZrO}_2$  lattice serve as crucial active sites for  $\text{CH}_3\text{OH}$  synthesis. In the case of the  $\text{ZnZrO}_x$  catalyst,  $\text{CO}$  is produced as a byproduct along with  $\text{CH}_3\text{OH}$ . Conversely, the  $\text{InZrO}_x$  catalyst produces a small amount of  $\text{CH}_4$ , in addition to  $\text{CH}_3\text{OH}$  and  $\text{CO}$ . Remarkably,  $\text{CH}_4$  is produced by passing  $\text{CH}_3\text{OH}$  and  $\text{H}_2$  through the  $\text{InZrO}_x$  catalyst. We performed a DFT study on  $\text{CH}_4$  formation on  $\text{InZrO}_x$  by calculating the free energy surface and density of states (DOS). Our calculations suggest a possible reaction pathway, highlighting the methyl group transfer with a competitive activation barrier, which is consistent with the experimental findings. The  $\text{CH}_3\text{OH}$  decomposition was also suggested to contribute to the  $\text{CH}_4$  production. A methyl group can only be adsorbed on the  $\text{InZrO}_x$  catalyst due to the better overlap of DOS between the adsorbate and substrate than on the  $\text{ZnZrO}_x$  catalyst.

## Author contributions

This manuscript was written through the contributions of all authors. All authors have given approval to the final version of the manuscript.

## Conflicts of interest

There are no conflicts of interest to declare.

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