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A density functional theory study on interactions in water-bridged dimeric complexes of lignin†

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Lignin is the main plant cell wall component responsible for recalcitrance in the process of lignocellulosic biomass conversion into biofuels. The recalcitrance and insolubility of lignin in different reaction media are due in part to the hydrogen bonds and π interactions that hold syringyl (S) and guaiacyl (G) units together and promote the formation of stable water-bridged dimeric complexes (WBDCs): S...G and S...S, in native lignin. The current understanding of how each type of interaction influences the stability of these complexes within lignin native cell walls is still limited. Here, we found by DFT calculations that hydrogen bonding is more dominant than π -stacking interaction between aromatic rings of WBDCs. Although there is a stronger interaction of hydrogen bonds between subunits and water and higher π -stacking interaction in the S...S complex compared to the S...G complex, the former complex is less thermodynamically stable than the latter due to the entropic contribution coming from the methoxy substituents in the S-unit. Our results demonstrate that the methoxylation degree of lignin units does not significantly influence the structural geometries of WBDCs; if anything, an enhanced dispersion interaction between ring aromatics results in quasi-sandwich geometries as found in "coiled" lignin structures in the xylem tissue of wood. In the same way as that with ionic liquids, polar solvents can dissolve S-lignin by favorable interactions with the aliphatic hydroxyl group in the α -position as the key site or the aromatic hydroxyl group as the secondary site.

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1. Introduction

Lignin, one of the main components of lignocellulosic biomass, is a crucial structural component preserving the integrity of the plant cell wall, imparting stiffness and strength of vascular plants, enabling the transport of water and solutes through the tracheary elements in the vasculature system, and affording physical barriers against invasions of phytopathogens, and

other environmental stresses.^{1–3} In nature, this component is primarily produced by enzymatic dehydrogenation and oxidative coupling of three hydroxycinnamyl alcohols or monolignols, namely, *p*-coumaryl alcohol, coniferyl alcohol, and sinapyl alcohol (Fig. 1), which are the origin of the *p*-hydroxyphenyl (H), guaiacyl (G), and syringyl (S) units following incorporation into the lignin polymer chains.^{4–8} Thus, lignin is a plant-originated polymer with a three-dimensional supramolecular structure of dimethoxylated (S), monomethoxylated (G), and non-methoxylated (H) phenylpropanoid and acetylated units linked by various types of C–O and C–C bonds, including α -O-4, β -O-4, β -5, 5-5, 4-O-5, β -1 and β - β linkages.⁹ Lignins generally have β -O-4 (alkyl aryl ether) as the majority bond, accounting for 45–60% of monomer connectivity.^{1,10,11} Based on its polyphenolic structure, lignin is a large-scale renewable feedstock composed of aromatics from biomass and can be used as a potential source of phenolic compounds to substitute for petroleum-derived chemicals.^{12–15}

The formulation of lignin and the ratio of the three units change with the type of cell and type of plant. For example, grasses are built up from H, G, and S units; softwood lignin (SL) or G-lignin essentially consists of G units; hardwood lignin (HL) of G and S units.^{16–18} G units constitute approximately 90% of SL from the gymnosperm plant, while roughly equal proportions of G units and S units appear in HL from the angiosperm plant.^{18–20}

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† Electronic supplementary information (ESI) available: Comparison of WBDCs inferred by radial distribution functions from MD simulations and DFT calculations. Optimized structures of S...G without water using DFT calculations. Structures of WBDCs: S...G and S...S showing all bonding critical points estimated using AIM's theory. Comparison of the geometric characteristics of WBDCs using DFT calculations and MD simulations. Electrostatic properties of hydrogen bonds in WBDCs. The main electron donor–acceptor interactions in π -stacking for WBDCs and their second-order perturbation stabilization energies at the M06-2X/6-31++g(d,p) level. Energetic contribution of π - π * and π - σ * interactions in WBDCs. Values of thermodynamic properties of units (S, G, and water) and WBDCs (S...S and S...G) calculated at the M06-2X/6-31++g(d,p) level. See DOI: <https://doi.org/10.1039/d4cp00312h>

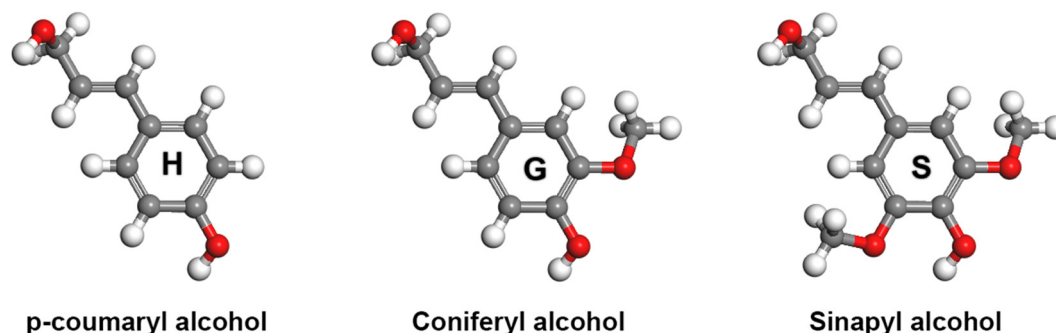
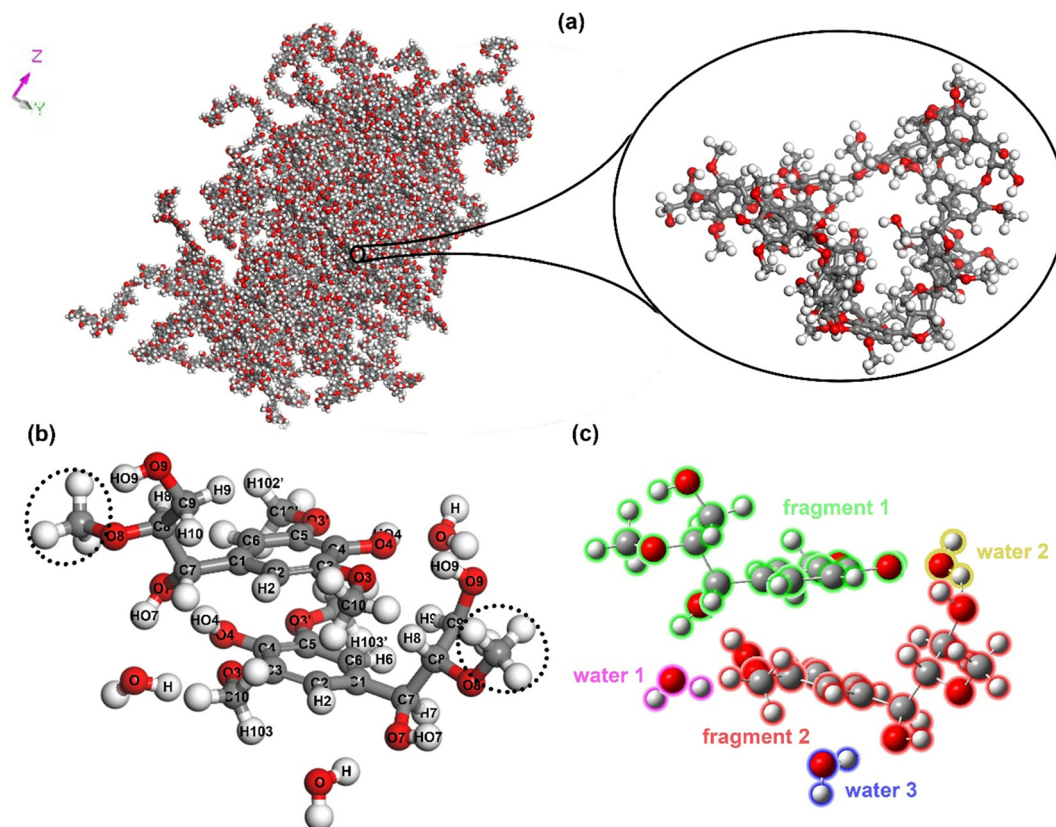


Fig. 1 Three cinnamyl alcohols: *p*-coumaryl, coniferyl, and sinapyl alcohols (referred to as monolignols) that respectively give rise to the H, G, and S units in lignin macromolecules. The $-OCH_3$ group is located at the *meta*-positions (positions 3 and 5) of the phenylpropanoid structure in coniferyl and sinapyl alcohols.

However, out of this classification, some species diverging from angiosperm plants can synthesize S-lignin having a high content of syringyl units.²¹ The hydrogen bonds (H-bonds) between neighboring hydroxyl groups and ether groups, as well as π - π interactions between aromatic groups, are responsible for the associated and complicated structures of lignin.²² These interactions hold lignin units together^{23,24} and promote the formation of

stable water-bridged dimeric complexes (WBDCs), as shown in Fig. 2.^{25–27} In fact, the existence of WBDCs: $G \cdots S$ and $S \cdots S$ in native lignin is a consequence of the association of monolignols and its radicals in aqueous solution during their (bio)polymerization^{2,28–31} rendering lignols in folded and extended-chain structures mediated by hydrogen bonding interactions between close HO4–O4 and HO7–O7 hydroxyl groups



(Fig. 2a),⁸ which participate in “coiled” structures found in the xylem tissue of tracheary elements of wood.^{32,33}

A fundamental understanding of how non-covalent interactions influence lignin stability at the microstructural level (e.g., in WBDCs) is key in the search for robust solvents to deconstruct lignocellulosic biomass or to prepare biopolymers by lignin aggregation in nanoparticles/colloids for environmentally friendly various applications.^{9,36–44} Like solubilization and aggregation phenomena, the structural stability of lignin is connected to the strength of H-bonds and π - π interactions, which in turn is linked to the number of methoxy groups ($-\text{OCH}_3$) in their units.^{25,38,45} From a π interaction perspective, lignins rich in G units (less methoxylated) could be better π -stacked than those rich in S units (more methoxylated), thus allowing better self-aggregation with poorer solubility.³⁸ In contrast, some studies^{45,46} showed that lignins rich in S units collapse better due to the hydrophobic effect, so they would have less solubility than lignins rich in G units. Apart from the hydrophobic effect and π interactions, hydrogen bonding is another type of interaction considered in the structural stability and solubility of lignin. Kubo *et al.*,²⁵ using simple model compounds of lignins with S units in aqueous media, revealed “strong” H-bonds between aromatic groups, which could be a factor hindering solubility and favoring self-aggregation of lignins rich in S units compared to those rich in G units. The above is an example, based only on physical interactions, that the degree of lignin aggregation and its stability in a solvation media depends on the competition between the repulsion force, intermolecular hydrogen bonding as well as hydrophobic interactions (π - π interactions) between lignin's subunits.³⁸

The current understanding of how each type of these interactions impacts the stability of complexes within lignin native cell walls is still limited. To our knowledge, few studies have considered the interaction between lignin units in aqueous media. Earlier studies investigated the association between monolignol radicals and monomers to reveal the lignin biosynthesis mechanism rather than the interaction between phenolic substructures constituting the natural lignin polymer. Li and Eriksson²⁸ performed molecular dynamics (MD) simulations to examine the association of some monolignol dimers – coumaryl and coniferyl alcohols – in water. However, the study of such interactions between coniferyl and sinapyl alcohols (Fig. 1), which constitute primary alcohols in HL subunits, was not addressed. Chen and Sarkanen⁴⁷ analyzed interactions between aromatic moieties in lignin chains, but some complexes involving lignin residues alone were probed. In a recent study,²⁶ we determined by molecular dynamics the most stable conformation of the hydrated G \cdots G complex into crumpled globules of softwood lignin. However, a quantitative understanding of non-covalent interaction strength governing its structural stability was not provided. Such an understanding could be obtained more efficiently by electronic structure calculations due to the complex nature of non-covalent interactions involving different types of energetic contributions such as electrostatics, charge transfer, orbital repulsion, electron dispersion, and so on.^{48–50}

Electronic structure calculations using density functional theory (DFT) have been used to understand the solubility of different lignin types in organic solvents and green reaction media,^{9,42,51–54} as well as the inter/intramolecular interactions between components of the lignocellulosic biomass.^{28,55,56} Nevertheless, lignin is a complex aromatic biopolymer and a rather expensive one using which quantum mechanical DFT calculations are performed. From a total of approximately 200 density functionals used in DFT,⁵⁷ three types of hybrid functionals are most commonly used in lignin or monolignols: B3LYP,^{9,22,28,55,56,58–63} Minnesota's suite (e.g. M05-2X, M06-2X)^{52,56,64–66} and B97-family (e.g. B97-D, ω B97M-V, ω B97D-2).^{22,51,54,67,68} The M05-2X and M06-2X hybrid meta-density functional series have broader accuracy than previous popular functionals (e.g. B3LYP) and provide an adequate representation of nonbonded interactions such as hydrogen bonding and dispersion forces,^{69–71} which are important in molecular systems containing phenolic moieties as lignin. These functionals account for “medium-range” electron correlation, sufficient to describe the dispersion interactions within many complexes.⁷² DFT calculations at the M05-2X/6-31+G(d,p) level of theory were performed to estimate interaction energies between lignol radicals and aromatic substructures in lignin and thus demonstrate that a template polymerization process is reasonable during lignin biosynthesis.⁴⁷ Hohenstein *et al.*⁷³ demonstrated that the M06-2X functional is better than the M05-2X functional at capturing “medium-range” (≤ 5 Å) electron correlation, allowing accurate description of the stacking interactions when the electron clouds of each monomer for stacked base pairs are in relative proximity to one another. The M06-2X/6-31+G(d) level of theory was able to reproduce stacking interaction energies in benzene sandwich dimers substituted with the $-\text{OCH}_3$ group comparable to CCSD(T)/aug-cc-pVTZ, but at a drastically reduced computational cost.^{74,75}

Some analysis using DFT calculations can be performed to obtain more information about the nature of non-covalent interactions, such as natural bond orbital (NBO) analysis,^{76,77} atoms in molecules (AIM) theory,^{77–81} the reduced density gradient (RDG) method,^{82,83} and the Shubin Liu's energy decomposition scheme (EDA-SBL).^{25,84–86} These analyzes can be performed together with DFT calculations because the key quantity estimated is the quantum-mechanical electron density (ρ), from which all chemical properties can, in principle, be obtained.⁸⁷ Some of these analyses have recently been performed to understand the key interactions between a variety of solvents with lignocellulose biomass components,^{9,42,54} however they have not been used to investigate interactions between lignin units in cell walls so far.

The primary goal of this work is to corroborate, from a quantum mechanical modeling point of view, the formation of stable WBDCs: S \cdots G and S \cdots S in native lignin, which influences conformation, stability and dissolution of lignin. Another purpose is to examine all energetic components engaged with non-covalent interactions affecting the structural stability of these complexes; for hydrogen bonds: electrostatic and charge-transfer, and for π - π interaction: electrostatic, London

dispersion and repulsion, which are much difficult to study by MD simulations. Our interest is unraveling how the strength of these interactions depends on the methoxylation degree of lignin and which of them governs the WBDC interaction networks. Here, we address the study of complexes with S units because of the importance of the syringyl constituent in facilitating overall lignin degradation for more efficient materials and energy production from angiosperm than from gymnosperm wood.^{88–90} Beyond intermolecular interactions, this study finds the relationship between these interactions and thermodynamic parameters (enthalpy and entropy) of stability in WBDCs, which are important in the rational design of new solvents for the dissolution of lignin rich in S-units. Herein, if some assumption is made about the solvent design based on DFT calculations, it applies more to the depolymerization of lignin that has many S-units (hardwood or syringyl lignins) in its composition and not to softwood lignin, which has almost no S-units and it is more branched.

2. Computational methodology

2.1. Optimization of geometries

All calculations were performed using Gaussian 09 (revision D.01) package⁹¹ with the M06-2X functional and 6-31++g(d,p) basis set at a temperature of 298 K and pressure of 1 atm. Although a medium-size basis set, this basis set shows a good estimation of interaction energy for π - π stacked complexes. For example, DFT calculations at the M05-2X/6-31++g(d,p) level of theory yielded a mean unsigned error of just 1.2 kcal mol⁻¹ relative to the best estimates for nucleobase π - π stacking energies.^{47,92} The nature of π -stacking interactions in various models of DNA base pairs,^{93,94} structure and energies of dilignols,⁶⁰ hydrogen bonding network in woody biomass⁵⁵ and the mechanism of lignin solubilization in ionic liquids (ILs)^{9,53} were also satisfactorily described using this basis set. Using 6-31G enables an excellent compromise between the computational cost and the accuracy of the computational results.⁹ In principle, + and ++ diffuse functions might be indistinctly used in the basis sets for modeling lignocellulose materials since they do not significantly affect the calculation of interaction energy, H-bonding, and O-H stretching frequencies.^{54,95} However, diffuse polarization functions on the second-row elements (d,p) must always be used in the basis set to successfully describe the interactions of components in woody biomass.⁵⁵

Before optimization, the WBDC geometries, S \cdots G and S \cdots S, were created starting from π - π stacked conformation for G \cdots G dimers found *via* MD simulation (structure in Fig. 2 by removing -OCH₃ groups).²⁶ Thus, to obtain the methoxylated complexes S \cdots G and S \cdots S, one and two -OCH₃ groups were introduced into G units, respectively. This study was performed for simple lignin complexes having a high content of -OCH₃ in their units because the S-unit shows less crosslinking in the native lignin's network.³⁰ For comparison, we leave the same number of water molecules in all WBC geometries.

The electronic energy minimization of these complexes was carried out in the gas phase employing a tight self-consistent field convergence criterion and ultrafine integration grid. This grid, having 99 radial shells and 590 angular points per shell, is recommended for optimizations of molecules with many soft modes, such as methyl rotations, making such optimizations more reliable.⁹⁶ The stability of the optimized WBDCs was certified by harmonic vibrational frequency analyses. Harmonic vibrational frequencies were computed for all WBDC optimized structures in order to characterize the stationary points. No imaginary frequencies were obtained, confirming that each geometry has a minimum on the potential energy surface.

It is important to stress that the aim of this work is not to derive new structures for WBDCs starting from randomly different configurations but rather to corroborate that there may be structures whose geometries are like those estimated with MD simulations²⁶ trapped in a nearby local minimum of energy landscapes estimated with DFT calculations.

2.2. Calculation of thermodynamic properties

The interaction energy (ΔE , eqn (1) and (2)) of the different configurations has been expressed as the electronic energy difference of WBDCs: $E^{S\cdots G}$ and $E^{S\cdots S}$ (Fig. 2a) and the corresponding optimized isolated geometries (E^G , E^S and E^{water}).⁹⁷ Zero-point energy corrections and basis superposition set error (BSSE) were included in the ΔE calculation. The BSSE-correction scheme was based on the counterpoise (CP) method developed by Boys and Bernardi⁹⁸ to account for basis set superposition errors resulting from using finite basis sets.

$$\Delta E^{S\cdots G} = 627.591 [E^{S\cdots G} - (E^G + E^S + 3E^{\text{water}})] \text{ in kcal mol}^{-1} \quad (1)$$

$$\Delta E^{S\cdots S} = 627.591 [E^{S\cdots S} - (2E^S + 3E^{\text{water}})] \text{ in kcal mol}^{-1} \quad (2)$$

The more negative ΔE , stronger interactions between units make up the complex.

The interaction enthalpy (ΔH), entropy (ΔS) and Gibbs energy (ΔG) of the formation of these complexes were calculated employing the same formulation used to estimate the interaction energy (eqn (1) and (2)). The Gibbs energy (G) and enthalpy (H) for all optimized configurations were obtained by adding the thermal corrections to the electronic energy calculated at the theory level. The volume variation (ΔV) during the formation of the complex was determined using the relationship $(\Delta H - \Delta E)/P$, where P ($= 1$ atm) is the pressure. A more negative ΔG correspond to a higher thermodynamic stability of the formed complex.

2.3. Topological analysis of bonding interactions

2.3.1 AIM theory. AIM theory has been used to analyze bonding characteristics in chemical systems regarding bond critical points (BCPs).^{99,100} In short, the theory is founded on the location of BCPs between atoms in which the electron density gradient ($\Delta\rho$) is zero, being possible to classify their nature according to the value of its Laplacian ($\nabla^2\rho$).⁸¹

This classification is possible due to the existent relationship between $\nabla^2\rho$ and the bond interaction energy by a local expression of the virial theorem $\left(\frac{\hbar^2}{4m}\right)\nabla^2\rho = V(r) + 2G(r)$,^{81,101} where $G(r)$ is the electronic kinetic energy density, which is always positive, and $V(r)$ is the electronic potential energy density that must be negative.¹⁰² A negative $\nabla^2\rho$ shows the excess potential energy at the BCP, which indicates that electronic charges are concentrated in the inter-nuclear region, and, therefore, shared by two nuclei by covalent bonding.¹⁰³ On the other hand, a positive $\nabla^2\rho$ shows the excess kinetic energy at the BCP, which reveals that electronic charges are depleted between nuclei as in conventional hydrogen bonding and ionic bonds.^{9,80} However, the extremely strong H-bonds can also have negative Laplacian values. For conventional H-bonds in neutral molecules (closed-shell molecules), the electron density and Laplacian value should be positive within the ranges of 0.002–0.035 a.u. and 0.024–0.139 a.u., respectively.⁸⁰ The Multiwfn software¹⁰⁴ was employed to determine all BCPs in the hydrogen bonding network of each WBDCs studied. A Python script was used to extract all BCPs that fulfil the requirements for conventional H-bonds.

2.3.2 RDG method. RDG is a very popular method for studying weak interactions.⁸² RDG, a form of electron density gradient norm function, is defined as follows:

$$\text{RDG}(\mathbf{r}) = \frac{1}{2(3\pi^2)^{1/3}} \frac{|\nabla\rho(\mathbf{r})|}{\rho(\mathbf{r})^{4/3}} \quad (3)$$

where, ρ and $\nabla\rho$ represent the electronic density and electronic density gradient, respectively. In order to recognize the nature of the interactions, three eigenvalues ($\lambda_1 \leq \lambda_2 \leq \lambda_3$) of the Hessian matrix of ρ are measured. The maximum electronic density is located at nuclei; hence, all the eigenvalues are negative. Two eigenvalues are negative for covalent interactions, whereas the third is positive. For bonded and non-bonded interactions, the λ_2 eigenvalue is vital, and therefore $\lambda_2 < 0$ and $\lambda_2 > 0$, respectively. According to Bader's theory of AIM,^{81,105} when λ_2 is negative, it indicates the presence of attractive interactions ((3, -1) type critical point), while when λ_2 is positive, it reveals repulsive forces as those found at the center of aromatic rings ((3, +1) type critical point). Thus, a plotting of RDG iso-surfaces against $\text{sign}(\lambda_2)\rho$ allows knowing where weak interaction occurs, but also intuitively captures the type of the interaction.^{9,42,82,105} RDG iso-surfaces are constructed with color codes *i.e.*, red, blue, and green, representing the steric repulsion, hydrogen bonding, and non-covalent interaction, respectively.

2.4. Energetics of bonding interactions

2.4.1 NBO analysis. This analysis complements results obtained using the AIM theory, interpreting hydrogen bonds and π - π interaction regarding charge transfer.^{9,106} For example, the hydrogen bonding is based on a charge transfer between filled 'donor' (*e.g.*, lone electron pair) and vacant 'acceptor' (*i.e.*, π^* , σ^*) orbitals leading to delocalization of electrons during the

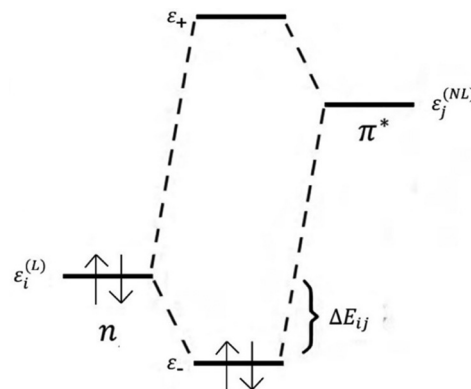


Fig. 3 2e-stabilizing interaction between a filled donor orbital $\varepsilon_i^{(L)}$ and vacant acceptor orbital $\varepsilon_j^{(NL)}$, leading to energy lowering ΔE_{ij} .¹⁰⁷ Here, L and NL stand for Lewis-type and non-Lewis type natural bonding orbitals (NBOs), respectively. The σ , π donor orbitals and lone electron pairs (Lp) of an atom in one molecule can overlap with the vacant σ^* or π^* acceptor orbital of another molecule. The figure illustrates the interaction ($n \rightarrow \pi^*$) between the lone electron pair (n) as the donor orbital and π^* as the vacant acceptor orbital leading to mixing (delocalization) by connecting the NL type NBOs with L NBOs.

mixing of these orbitals. This quantum phenomenon causes an overall energy lowering ('stabilization energy', Fig. 3), which can be quantified according to second-order perturbation theory as follows:

$$E(2) = \Delta E_{ij} = \frac{-q_i |F_{ij}|^2}{(\varepsilon_j^{(NL)} - \varepsilon_i^{(L)})} \quad (4)$$

where q_i is the occupancy of the donor orbital, and $\varepsilon_i^{(L)}$ and $\varepsilon_j^{(NL)}$ are the respective donor and acceptor orbital energies. F_{ij} is the off-diagonal NBO Fock matrix element, related to orbitals superposition integral.¹⁰⁷ The 'stabilization energy', ΔE_{ij} is commonly abbreviated as $E(2)$.

In the NBO analysis, $E(2)$ is used to describe the intensity of orbital interaction between electron donors and electron acceptors. The higher the value of $E(2)$, the more the electrons transfer from the donor, increasing the bond strength.⁹ The Gaussian 09 package was used to make such an analysis employing pop = nbo and nboread keywords. The visualization of NBO was carried out with the Multiwfn software.

2.4.2 EDA-SBL decomposition scheme. To investigate the nature of interactions involved in stabilizing WBDCs, we followed an energy decomposition scheme proposed by Shubin Liu.^{84–86} According to this scheme, the total energy density functional ($E[\rho]$) of an atom and molecule can alternatively be partitioned as

$$E[\rho] = E_s[\rho] + E_e[\rho] + E_q[\rho] \quad (5)$$

where $E_s[\rho]$, $E_e[\rho]$, and $E_q[\rho]$ stand for the independent energy contribution from the steric, electrostatic, and quantum effects, respectively. The steric term is simply the energy derived using the Weizsäcker kinetic functional (eqn (6)), which corresponds to the exact kinetic energy under the assumption that the electrons in the present system are non-interacting bosons.⁸⁴

$E_s[\rho]$ is nowhere negative in space and extensive. That is, the larger the system, the larger the steric energy.

$$E_s[\rho] = \frac{|\nabla\rho(\mathbf{r})|^2}{8\rho(\mathbf{r})} \quad (6)$$

The electrostatic term ($E_e[\rho]$, eqn (7)) is the sum of all classical Coulomb interactions of the particles in the system.

$$\begin{aligned} E_e[\rho] &= E_J + E_{N-E} + E_{N-N} \\ &= \iint \frac{\rho(\mathbf{r}_1)\rho(\mathbf{r}_2)}{r_{12}} d\mathbf{r}_1 d\mathbf{r}_2 - \int \rho(\mathbf{r}) \sum_A \frac{Z_A}{|\mathbf{r} - \mathbf{R}_A|} d\mathbf{r} \\ &\quad + \sum_{A>B} \frac{Z_A Z_B}{R_{AB}} \end{aligned} \quad (7)$$

where E_J , E_{N-E} and E_{N-N} represent classical interelectron Coulomb repulsion, nuclear–electron attraction, and nuclear–nuclear repulsion, respectively. $\rho(\mathbf{r})$ and z stand for electron density at \mathbf{r} distance and nuclear charge, respectively. $E_e[\rho]$ is calculated from point charges or multipoles distributed over the monomers.

The quantum term ($E_q[\rho] = E_{XC}[\rho] + E_{\text{Pauli}}[\rho]$) is the energy purely caused by the quantum effect, which can be broken down into exchange–correlation energy ($E_{XC}[\rho]$) and Pauli energy ($E_{\text{Pauli}}[\rho]$). According to the Kohn and Sham (KS) formalism, $E_{XC}[\rho]$ can be calculated as follows:⁵⁴

$$E_{XC}[\rho] = \frac{1}{2} \iint \frac{\rho(\mathbf{r}_1)\rho(\mathbf{r}_2)}{|\mathbf{r}_1 - \mathbf{r}_2|} [\tilde{g}(\mathbf{r}_1, \mathbf{r}_2) - 1] d\mathbf{r}_1 d\mathbf{r}_2 \quad (8)$$

where $\tilde{g}(\mathbf{r}_1, \mathbf{r}_2) = \int_0^1 g(\mathbf{r}_1, \mathbf{r}_2, \lambda) d\lambda$ and $g(\mathbf{r}_1, \mathbf{r}_2, \lambda)$ stand for the electron pair correlation function in a range of interaction strengths (λ). The shape of the $\tilde{g}(\mathbf{r}_1, \mathbf{r}_2)$ function is precisely determined using the DFT functional used in calculations. For hybrid meta-density functionals such as M06-2X used in this work, $\tilde{g}(\mathbf{r}_1, \mathbf{r}_2)$ depends on ρ , $\nabla\rho$, and $\nabla^2\rho$ (included in the kinetic energy density). $E_{XC}[\rho]$ describes the long-range attractive dispersion energetic component (e.g., London-dispersion forces) acting between aromatic rings, which is negative. This dispersion energy depends on explicit electron–electron correlations,¹⁰⁸ which are associated with overlapping of π - π

and π - σ orbitals.¹⁰⁹ $E_{\text{Pauli}}[\rho]$ is also known as exchange repulsion and it is determined by subtracting the Weizsäcker kinetic energy (eqn (6)) from the exact kinetic energy of the non-interacting electron system defined using KS-DFT theory.¹⁰⁵ From a physical point of view, it represents a short-range repulsion effect between electrons in occupied orbitals, which is always positive.

The contribution of each type of energy to the overall interaction energy was calculated from $\frac{E_i \Delta E^{\text{complex type}}}{\left| \sum_i E_i \right|}$; where i represents the energy type: steric, electrostatic, exchange–correlation, and Pauli repulsion, and the complex type is S··G or S··S. Here $\left| \sum_i E_i \right| = |E[\rho]|$.

3. Results and discussion

3.1. Geometric characteristics and interaction energy of the complexes

The optimized structures of WBDCs (Fig. 4) by DFT were like the starting structures guessed by molecular dynamics calculations²⁶ in terms of dimer ring orientation (Fig. S1, ESI†). This indicates that new parameterizations of the CHARMM force field for lignin³⁴ adequately reproduce structural features of lignin dimers with hydrogen bonding. DFT dimerization energy landscapes of studied complexes may be well captured by the CHARMM force field, as observed for pairs of ring-containing amino acids in proteins when comparing CHARMM27 or OPLS-AA force fields with the MP2 method.¹⁰⁸ If there is any structural discrepancy between the DFT and MD results, it is attributed to the inability of force fields based on fixed partial charges centered on the atoms to reproduce the effects related to the partially covalent character of hydrogen bonding, such as higher-order multipoles and electronic polarization. On the other hand, it is also noteworthy that the initial structures used in the optimization by DFT were derived from MD simulations using radial distribution functions, which provide average distances

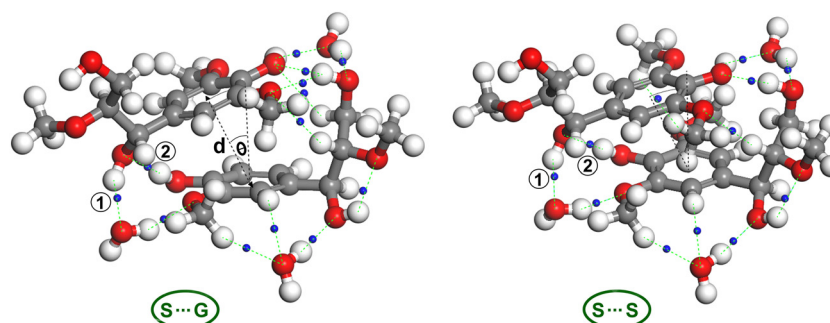


Fig. 4 Optimized water-bridged dimeric complex (WBDC) structures: S··G and S··S using DFT at the M06-2X/6-31++g(d,p) theory level. Carbon, oxygen, and hydrogen atoms are represented by gray, red, and white balls, respectively. In these WBDCs, the bonding critical points (BCPs) of hydrogen bonds (H-bonds) are illustrated by blue balls connecting with dashed green lines between oxygen and hydrogen atoms. In dashed black lines, the distance (d) from centroids and interplanar angle (θ)¹⁰⁸ between aromatic rings. The BCPs with circled number 1 are H-bonds formed between HO7 of fragment 1 and oxygen of water 1 ([1] O7H··O [W1]), while the BCPs with circled number 2 are H-bonds formed between O7 of fragment 1 and HO4 of fragment 2 ([1] O7··HO4 [2]). Fig. 2 illustrates the nomenclature of atoms in each fragment and the numbering of water molecules in these WBDCs.

Table 1 Geometric characteristics of water-bridged dimeric complexes (WBDCs) and thermodynamic parameters of their formation estimated using DFT

WBDCs	Number of methoxy groups	d^a [Å]	θ^a [°]	ΔE^b [kcal mol ⁻¹]	ΔH^b [kcal mol ⁻¹]	$\Delta V \times 10^{-4}$ [cm ⁻³]	$T\Delta S^c$ [kcal mol ⁻¹]	ΔG^b [kcal mol ⁻¹]
S··G	3	3.60	23.74	-40.33	-49.29	-2.17	-43.66	-5.63
S··S	4	3.52	23.11	-40.63	-50.53	-2.39	-45.70	-4.83

^a d and θ mean the distance and interplanar angle between aromatic rings (Fig. 4a). ^b The values of E , H , S and G for each optimized configuration involved in the formation of WBDCs are assembled in Table S5 (ESI). ^c T means the temperature and is equal to 298 K.

between water molecules and hydroxyl groups so that the positions of the H-bonds can be somewhat different between DFT and MD (Fig. S1, ESI†).

A glance at Fig. 4 reveals that the methoxylation degree of subunits in WBDCs does not significantly influence their molecular geometries. For example, the estimated distance of the center of mass (d) and interplanar angle (θ) between aromatic rings slightly decreases with the increasing number of $-\text{OCH}_3$ groups in WBDCs (Table 1). For the studied WBDCs, the d values agree with those found for benzene sandwich dimers (~ 3.8 Å, Fig. 5a)¹⁰⁹ and double veratrylglycerol- β -guaiacyl ether dimers (3.643 Å).⁴³ The value of d is somewhat higher for the S··G complex than for the S··S complex, which indicates a more favorable interaction in the latter complex, as shown by interaction energy (ΔE) values in Table 1. The above agrees with theoretical results, wherein it was demonstrated that $-\text{OCH}_3$ substituents in lignin led to the strengthening of lignin/lignin effective interactions in aqueous media.⁴⁵ Also, the ΔE results obtained by DFT in this study are in line with those estimated by MD for lignin constituents in water,²⁸ in which the estimated potential energy of interactions between coumaryl (H) dimers was lower (-41.4 kcal mol⁻¹) than that between coniferyl (G) dimers (-80.8 kcal mol⁻¹). On the other hand, the hydration favors interunit interactions in these complexes since the values of ΔE calculated for these complexes without water were lower (e.g., $\Delta E = -17.34$ kcal mol⁻¹ for S··G) compared to those reported here (Table 1). However, even without water molecules, these complexes can be stabilized by intermolecular hydrogen bonding as $\text{O7}\cdots\text{HO4}$ and $\text{O4}\cdots\text{HO9}$ (Fig. S2 in the ESI†).

The values of θ around 23–29° (Table 1) suggest the existence of parallel-displaced geometry in the π -stacking,^{110,111} showing a quasi-sandwich (Fig. 5) character. This geometry justifies transition structures acting in the growing mechanism of dilignols to oligolignols in the plant cell.⁸ It has been suggested that the coniferyl alcohol radical may be oriented in the transition state during cross-coupling so that its aromatic ring is parallel to the guaiacyl ring at the growing end by the formation of a π -complex like a “sandwich transition state”.¹¹²

3.2. Calculation of thermodynamic properties in the complex formation

Although the interaction energy (ΔE) appears to be a dominant force in the stabilization of the complexes through the highly exothermic enthalpic component (ΔH) (Table 1), calculations of Gibbs energy (ΔG) of the formation of these complexes must be performed to harshly judge their stability. The calculated ΔG values are collected in Table 1, where it can be seen that ΔG for the S··G complex is slightly more negative (-5.63 kcal mol⁻¹) than for the S··S complex (-4.83 kcal mol⁻¹), revealing a superior thermodynamic stability for the S··G complex. The above can be rationalized on the basis of entropic contributions coming from the methoxy group to the Gibbs energy ($\Delta G = \Delta H - T\Delta S$) of complex formation. Despite more favorable interactions (with more negative ΔH) during the formation of the S··S complex, it shows a more negative ΔS (Table 1) during its formation, which causes a less negative ΔG for the S··S complex compared to the S··G complex. That is, the increase of S-unit entropy (Table S5, ESI†) coming from the methoxy substituent provokes a more accentuated decrease in ΔS in the

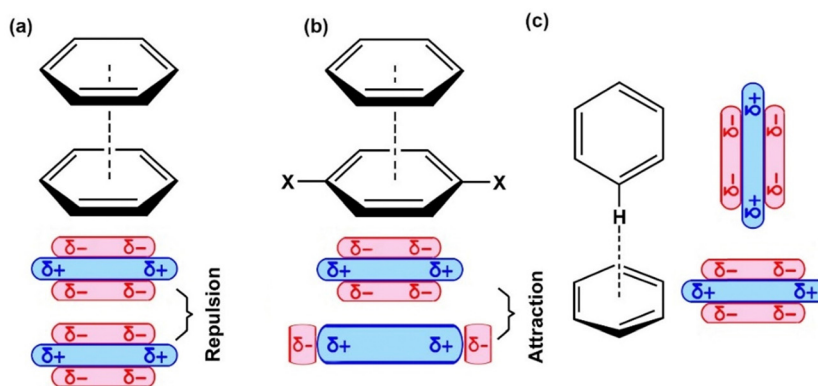


Fig. 5 (a) and (b) Cofacial “sandwich” π -stacked configurations, (c) edge-to-face T-shaped geometry.^{114,117} The delocalized π electrons impart a quadrupole moment with partial negative charge (δ^-) above both aromatic faces and a partial positive charge (δ^+) on hydrogen atoms around the periphery.^{49,118} (b) A favored cofacial “sandwich” π -stacked configuration between electron-rich and electron-deficient aromatic rings, which is accompanied by close contact between π orbitals of adjacent molecules.¹¹⁸ X symbolizes electron-withdrawing substituting groups.

formation of the S··S complex. Such a diminution of ΔS is in turn accompanied by favored interactions between the S··S complex's units, leading to a slightly larger decrease in ΔV (or contraction) for the S··S complex, as shown in Table 1. This outcome is in consonance with the lower value of distance (d) between aromatic rings for the S··S complex (3.52 Å) compared to the S··G (3.60 Å) complex. The higher thermodynamic stability of S··G in relation to S··S justifies the greater recalcitrance in biomass dissolution found for lignins containing higher contents of guaiacil (G) units.⁸⁸ On the other hand, taking into account the thermodynamic stability of these complexes, it can be inferred that solvents with some polarity (between ethanol and DMSO)³⁵ can favor the dissolution of lignin rich in S units, as found for hardwood lignin.⁴² Possibly, it is due to the interaction of hydrogen bonding between lignin units with the solvent, which decreases the entropy ($S^{\text{S} \cdots \text{S}}$) and increases the enthalpy ($H^{\text{S} \cdots \text{S}}$) of S-units into complexes compared to the dissociated state (e.g., S^{S}), thus disfavoring complexation (or increase in ΔG) with improved lignin solvation.

Despite the fact that the entropic component drives the stability of WBDCs under ambient conditions, the interaction energy is an opposite force and not much less important than the entropy, which is worth investigating. In fact, as before-mentioned, the interaction energy influences the geometry and stability of these complexes. An explanation of the interaction forces driving the formation of parallel-displaced or sandwich geometries in π -stacking is still unclear, and various models have been proposed for specific systems. According to the “ π /polar or electrostatic model” developed by Hunter and Sanders,¹¹³ sandwich geometries might be stabilized by decreasing quadrupolar electrostatic interactions between aromatic rings competing with London dispersion that favor cofacial π -stacking.^{49,113} However, this intuitive electrostatic model has been criticized for considering only the electrostatic forces as dominant and not considering other phenomena, such as charge penetration, to explain parallel-displaced structures in π -stacking.¹¹⁴ Also, the quadrupolar repulsion cannot account for the structures adopted by larger polycyclic aromatics and geometries of larger graphene analogues, which may be better understood *via* Pauli repulsion models whose origin is quantum. Another proposed model, the “direct-interaction model”,^{74,75,115,116} which conflicts with the electrostatic model, states that substituent effects in the benzene sandwich dimer are relatively unimportant in the π -system of the substituted benzene and rather dominate the electron-withdrawing character of the substituent in the σ -bonds.

Since the overall nature of substituent effects on π stacking involves the interplay of several competing factors, including electrostatics, dispersion, and direct substituent-ring interactions, a careful investigation should be performed to unravel which of them is predominant in WBDC interaction networks. However, the picture of the interaction in WBDCs goes beyond such interactions that only participate in π -stacking. The WBDC geometries (Fig. 4) may result from intermolecular H-bonding interactions between the water molecules and the units, as well as between the units themselves, overcoming repulsive interactions

Table 2 Electron density properties and geometric characteristics of hydrogen bonds (H-bonds) at bond critical points (BCPs) in water-bridged dimeric complexes (WBDCs)

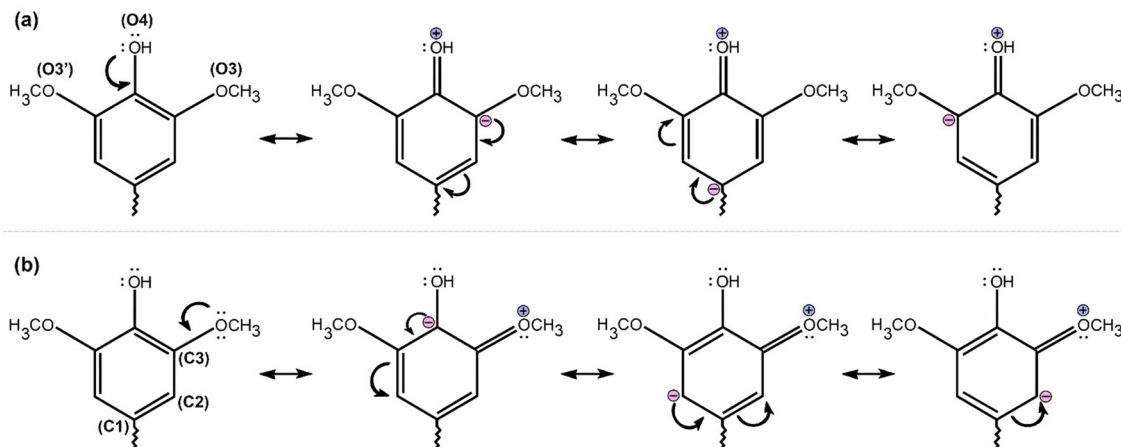
WBDCs	H-Bonds ^a	Length [Å]	Angle [°]	ρ_{BCP}	$\nabla^2 \rho_{\text{BCP}}$	$H_{\text{BCP}} \times 10^{-3}$ [a.u.]
S··G	[1] O7H··O [W1]	1.859	148.94	0.0302	0.0995	−0.0160
	[1] O7··HO4 [2]	1.838	171.14	0.0318	0.0978	−0.0230
	[1] O4H··O [W2]	2.072	119.38	0.0201	0.0729	−0.2050
	[1] O4··HO9 [2]	2.012	131.83	0.0227	0.0784	−0.2080
	[1] O3··HO9 [2]	2.398	146.13	0.0093	0.0357	0.7580
	[1] O4··H9 [2]	2.559	110.15	0.0102	0.0388	1.0950
	[1] O3··H8 [2]	2.596	141.11	0.0084	0.0276	0.6118
	[2] O3··H [W1]	1.924	141.28	0.0258	0.0868	−0.2591
	[2] O9··H [W2]	1.962	150.23	0.0251	0.0755	−0.8722
	[2] O7H··O8 [2]	1.997	117.29	0.0256	0.0986	0.2999
	[2] O7··H [W3]	1.897	166.27	0.0261	0.0866	−0.3020
	[2] H103··O [W3]	2.349	154.76	0.0122	0.0394	0.3487
	[2] H2··O [W3]	2.453	123.88	0.0096	0.0410	1.6779
S··S	[1] O7H··O [W1]	1.852	150.78	0.0308	0.1006	−0.03094
	[1] O7··HO4 [2]	1.866	171.29	0.0299	0.0908	−0.19783
	[1] O4H··O [W2]	1.966	131.89	0.0243	0.0816	−0.28417
	[1] O4··HO9 [2]	1.941	137.45	0.0252	0.0878	−0.02348
	[1] H103'··O3' [2]	2.452	120.01	0.0112	0.0389	0.79105
	[1] O3··H8 [2]	2.369	96.74	0.0125	0.0378	0.09377
	[2] O3··H [W1]	1.907	129.72	0.0266	0.0903	−0.1382
	[2] O9··H [W2]	1.971	145.84	0.0250	0.0754	−0.9600
	[2] O7H··O8 [W2]	1.952	118.74	0.0277	0.1058	0.1942
	[2] O7··H [W3]	1.900	164.16	0.0256	0.0861	−0.3397
	[2] H103··O [W3]	2.382	151.28	0.0156	0.0376	0.4226
	[2] H2··O [W3]	2.411	132.25	0.0102	0.0433	0.0016

^a The numbers 1 and 2 in brackets represent the fragments 1 and 2 from which the oxygen or hydrogen atoms in the H-bonds come from. Here the notations [W1], [W2], and [W3] also indicate fragments, but they are referred to as water molecules: water 1, water 2, and water 3, respectively (Fig. 2).

(e.g., Pauli repulsion) between aromatic rings. Just considering that π - π interaction energies for the benzene dimer in stable geometries (parallel displacement or T-shape) have a contribution of around 2.7 kcal mol^{−1},^{74,119,120} and that the calculated overall interaction energy is approximately 40 kcal mol^{−1} (Table 1), other energetic contributions are expected to predominate in the association of units in WBDCs. In the next sections, we examined each type of these interactions engaged in the stability of WBDCs.

3.3. Topological analysis of hydrogen bonding interactions

The AIM theory was used to analyze hydrogen bonding characteristics to obtain information about the interactions in these studied complexes. Using this theory, different bond critical points (BCPs) were found for all studied WBDCs (Fig. S3, ESI†). From these BCPs, those related to conventional H-bond type (3, −1) could be obtained, as shown in Fig. 4. The values of electron density (ρ_{BCP}), its Laplacian ($\nabla^2 \rho_{\text{BCP}}$), and energy density (H_{BCP}) for these H-bonds are listed in Table 2. We can observe that the values of ρ_{BCP} and $\nabla^2 \rho_{\text{BCP}}$ lie in the relative proposed ranges for conventional H-bonds, and strongest H-bonds like O7H··O and O7··HO4 have the highest values for S··G and S··S complexes. From a geometric point of view, these H-bonds show shorter bond lengths and angles close to 180° (Table 2), which favor strong hydrogen donor–acceptor interactions.^{9,121,122} These H-bonds show negative values of



Scheme 1 Resonance structures caused by the positive mesomeric (M+) effect of hydroxyl groups (a) and methoxy $-\text{OCH}_3$ (b) in S units of lignin. The effect draws into lone pairs of electrons from the O3, O3' and O4 atoms towards the π electron cloud conferring negative partial charges on carbon atoms of the benzene ring. According to our calculations, almost all benzene carbon atoms showed negative or low positive partial charges in S units of WBDCs (results not shown), showing a +M combined effect of these electron-donating groups.

H_{BCP} , which reveals electronic shared interactions or covalent within the WBDCs.¹⁰² They have partially covalent and partially electrostatic properties and can be classified as strong H-bonds.¹²³ The strongest H-bonds at O7 and HO4 interactions in these complexes are consistent with theoretical findings reported for lignin dissolution with ILs.⁹ Zhang *et al.*, using the guaiacyl glycerol- β -guaiacyl ether (GG) model with imidazolium-based ionic liquids, found that the α -OH group (in our case O7–HO7 in WBDCs) is the key site when ion pairs interact with GG, while the p-OH (O4–HO4 in WBDCs) is the secondary site favorable for H-bonds.

On the other hand, the weakest H-bond interactions correspond to those formed between hydrogen atoms attached to carbon atoms in a unit and some type of oxygen from water (*e.g.*, W3) or neighbor unit: O3 \cdots H8, O4 \cdots H9, H103' \cdots O3', H103 \cdots O, and H2 \cdots O. These H-bonds have positive values of H_{BCP} , which discloses the electrostatic nature of them. This is a consequence of low partial positive charges, ranging from 0.101 to 0.405 (Table S2, ESI[†]) of aliphatic hydrogen bond donors: H8, H9, H103, and H103', by successive Sigma bond polarizations on carbon skeleton. This polarization is mediated by the inductive ($-I$) effect of electron-withdrawing oxygen atoms (O3, O3', O8, and O9) attached to carbon atoms adjacent (C10, C10', C8, and C9) to these hydrogen atoms.

For the case of H2 atoms, belonging to the aromatic ring and involved in H2 \cdots O hydrogen bonds, its low partial positive charge is caused by the positive mesomeric effect (+M) (Scheme 1). The $-\text{OCH}_3$ substituent added at the C3 position of the aromatic ring increases the negative charge on the C2 carbon (from -0.373 in S \cdots G to -0.436 in S \cdots S), decreasing the partial positive charge on the H2 atom (from 0.109 in S \cdots G to 0.103 in S \cdots S, Table S2, ESI[†]). Actually, H2 \cdots O is the weakest bond of all H-bonds having an electrostatic attraction force (Table S2, ESI[†]) of -318 and 330 pN for S \cdots G and S \cdots S, respectively, which is in the range of -75.8 to -521 pN found for a diversity of planar complexes with H-bonds dominated by

electrostatic interactions.¹²⁴ The geometry of this H-bond is similar to that found between the anions of ILs and GG, wherein there is an anchoring of the anion (in our case water) between the $-\text{OCH}_3$ group and the α -OH position of GG.⁹ Curiously, the O3 \cdots HO9 hydrogen bond, which should be of moderate strength in the S \cdots G complex, it is a weak H-bond considering low ρ_{BCP} (Table 2) and electrostatic force values (Table S2, ESI[†]). The weakness of this H-bond can be explained in light of that lone electron pairs of O3 atoms are delocalized by the +M effect (Scheme 1),¹²⁵ resulting in low partial negative charges (-0.264 , Table S2, ESI[†]) on the O3 atom.

In general, intermolecular H-bonds (*e.g.*, H103 \cdots O[W3], O7 \cdots H[W3], O4H \cdots O[W2], H2 \cdots O[W3]) related to hydration of units are stronger in the S \cdots S complex than in the S \cdots G complex, due to the +M effect of extra $-\text{OCH}_3$ group in the aromatic ring (Scheme 1). This effect augments the partial positive charges on O4H and H103 and decreases that on H2, whereas the partial negative charge increases on the O7 atom. The oxygen atoms (O[W2] and O[W3]) of water molecules involved in these H-bonds appear to increase the negative charge due to the polarization effect by surrounding units, leading to a net augmentation of the attractive electrostatic forces and diminution of distance for these H-bonds (Table S2, ESI[†]). The better hydration observed in the S \cdots S complex caused by H-bond interactions is in line with theoretical evidence about S-units that shows higher interactions of hydrogen bonding in water than in G-units.³⁵ However, on analyzing the values of ρ_{BCP} (Table 2) corresponding to H-bonds involved with interunit interactions (O7 \cdots HO4, O4 \cdots HO9, O3 \cdots HO9, O4 \cdots H9, O3 \cdots H8, and H103' \cdots O3'), there is no clear trend to claim that the S \cdots S complex has stronger interaction than the S \cdots G complex (Table 1). That is, when the values of ρ_{BCP} for each of these interactions are compared individually, sometimes they appear to be higher for S \cdots S and other times for S \cdots G as shown in Table 2. In this case, it is necessary to apply another auxiliary method, as will be seen in the next sections,

Table 3 The main electron donor–acceptor interactions in hydrogen bonds (H-bonds) for water-bridged dimeric complexes (WBDCs) and their second-order perturbation stabilization energies ($E(2)$) at the M06-2X/6-31++g(d,p) level

WBDCs	H-Bond ^a	Donor (i)	Acceptor (j)	$E(2)$ [kcal mol ^{−1}]
S··G	[1] O7H··O [W1]	Lp O	σ^* O7–H	15.74
	[1] O7··HO4 [2]	Lp O7	σ^* O4–H	17.47
	[1] O4H··O [W2]	Lp O	σ^* O4–H	6.17
	[1] O4··HO9 [2]	Lp O4	σ^* O9–H	7.56
	[1] O3··HO9 [2]	Lp O3	σ^* O9–H	1.72
	[1] O4··H9 [2]	Lp O4	σ^* C9–H9	0.51
	[1] O3··H8 [2]	Lp O3	σ^* C8–H8	0.60
	[2] O3··H [W1]	Lp O3	σ^* O–H	1.66
	[2] O9··H [W2]	Lp O9	σ^* O–H	2.39
	[2] O7H··O8 [2]	Lp O8	σ^* O7–H	3.28
	[2] O7··H [W3]	Lp O7	σ^* O–H	6.38
	[2] H103··O [W3]	Lp O	σ^* C10–H103	6.49
	[2] H2··O [W3]	Lp O	σ^* C2–H2	0.36
S··S	[1] O7H··O [W1]	Lp O	σ^* O7–H	19.65
	[1] O7··HO4 [2]	Lp O7	σ^* O4–H	15.20
	[1] O4H··O [W2]	Lp O	σ^* O4–H	7.63
	[1] O4··HO9 [2]	Lp O4	σ^* O9–H	8.43
	[1] H103'··O3' [2]	Lp O3'	σ^* C10'–H103'	0.98
	[1] O3··H8 [2]	Lp O3	σ^* C8–H8	2.48
	[2] O3··H [W1]	Lp O3	σ^* O–H	3.13
	[2] O9··H [W2]	Lp O9	σ^* O–H	2.48
	[2] O7H··O8 [2]	Lp O8	σ^* O7–H	3.85
	[2] O7··H [W3]	Lp O7	σ^* O–H	6.78
	[2] H103··O [W3]	Lp O	σ^* C10–H103	2.94
	[2] H2··O [W3]	Lp O	σ^* C2–H2	0.68

^a The numbers 1 and 2 in brackets, respectively, represent the fragments 1 and 2 from which the oxygen or hydrogen atoms in the H-bonds come from. Here the notations [W1], [W2], and [W3] also indicate fragments, but they are referred to as water molecules: water 1, water 2, and water 3, respectively (Fig. 2).

to obtain detailed information about the energy values of each hydrogen bonding. Then, the sum of these energy values encompassing H-bonds related to an interaction type (e.g., interunit or hydration) can inform their contribution to the interaction overall energy.

3.4. NBO analysis

3.4.1 Energetics of hydrogen bonding. To estimate the energetic contribution of each interaction to the overall interaction

network in WBDCs, hyper-conjugative donor–acceptor interactions in hydrogen bonding and π -stacking interactions involving aromatic rings were calculated. In addition to the energy of each interaction, NBO analysis also gives the electron-related properties. Table 3 shows the main donor–acceptor interactions between atoms involved in H-bonds of WBDCs and their second-order perturbation stabilization energies, $E(2)$, calculated using eqn (4). The values of $E(2)$ denote the strength of lone pair donor–antibonding Sigma acceptor interaction ($n \rightarrow \sigma^*$). The larger the value, the stronger the interaction will be. From $E(2)$ values listed in Table 3, it can be concluded that some of the H-bonds are strong, and some are weak, as previously demonstrated by AIM theory. Among these, the strongest H-bond in the S··G complex is formed between O7 and HO4 from two units ([1] O7··HO4 [2]), while the strongest H-bond in the S··S complex is formed between water oxygen and HO7 ([1] O7H··O [W1]). The variation of interaction strength in these bonds comes from the competitiveness of O7 as a hydrogen acceptor or a hydrogen donor. In the S··G complex, there is a better overlap of orbitals between the lone-pair of O7 oxygen atoms and the anti-bonding Sigma orbital of O4–H bonds, indicating that O7 acts as an electron donor in the charge transfer when forming this H-bond (Fig. 6a). However, the presence of an additional –OCH₃ group in the S··S complex contributes to a shorter distance (1.852 Å) between the oxygen of water 1 ([W1] O) and HO7 hydrogen, thus allowing an increased overlap of orbitals between the lone-pair of [W1] O and anti-bonding Sigma orbital of the HO7–O7 bond, as illustrated in Fig. 6b. According to Grabowski's classification,¹²⁶ the stabilization energy of these bonds (17.47 and 19.65 kcal mol^{−1}, Table 1) is in the range of 15 to 40 kcal mol^{−1}, which allows classifying O7··HO4 and O7H··O as strong hydrogen bonds, as above corroborated by AIM theory.

Summing up all $E(2)$ values of H-bonds between water molecules and units in the WBDCs (e.g., for S··G: O7H··O, O4H··O, O3··H, O9··H, O7··H, H103··O, and H2··O), we have, respectively, 39.19 and 43.29 kcal mol^{−1} for S··G and S··S complexes, which indicates a better interaction of water molecules with the S··S complex. Thus, a greater number of S-units would allow a more effective degree of hydration in lignin

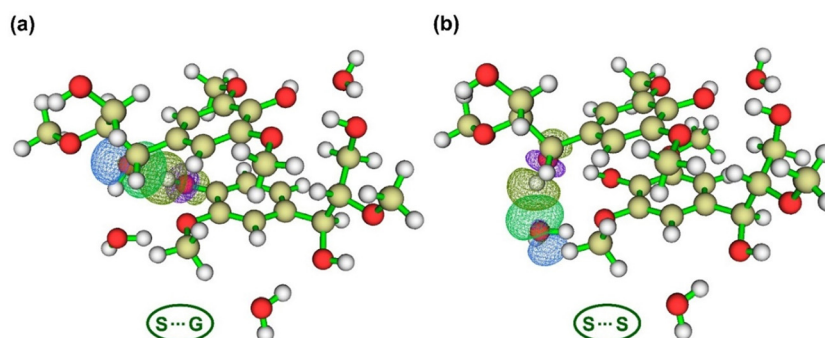


Fig. 6 Natural bond orbital interaction in water-bridged dimeric complexes: S··G and S··S calculated at the M06-2X/6-31++g(d,p) theory level (green: isovalue = 0.05 and blue: isovalue = −0.05). The H-bond is considered a charge transfer interaction ($n \rightarrow \sigma^*$) between lone electron pair (Lp) and antibonding Sigma (σ^*). (a) Shows the orbital interaction between Lp of O7 in fragment 1 and σ^* of O4–H bond in fragment 2 ([1] O7··HO4 [2]). (b) Depicts the orbital interaction between Lp of water 1's oxygen and σ^* of O7–H bond ([1] O7H··O [W1]).

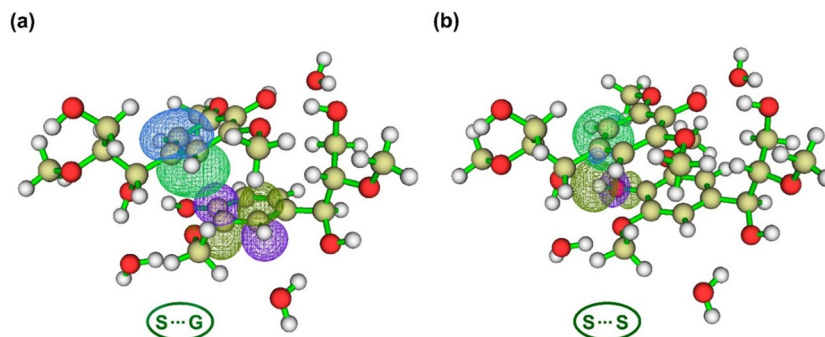


Fig. 7 Natural bond orbital interaction in water-bridged dimeric complexes: S...G and S...S calculated at the M06-2X/6-31++g(d,p) theory level (green: iso-value = 0.05, blue: iso-value = -0.05). (a) The orbital interaction between π of the C1-C2 bond in fragment 1 (π C1-C2 [1]) and π^* of C2-C3 bond in fragment 2 (π^* C2-C3 [2]). (b) The orbital interaction between π of the C1-C6 bond in fragment 1 (π C1-C6 [1]) and σ^* of the O4-HO4 bond in fragment 2 (σ^* O4-HO4 [2]).

complexes, which is in consonance with a higher bound water/free water ratio found in HL (more methoxylated lignin) than in SL¹²⁷ and lower absolute aqueous solvation free energy for S-lignin.^{35,45} Thus, the increase of the S-unit content in lignin would impede polymer extension and solvation (dissolution) with solvents, due possibly to strong interactions between units and water making lignin structures more compact.³⁵ On the other hand, by summing up all $E(2)$ values for H-bonds related to interactions between units (*e.g.*, O7...HO4, O4...HO9, O3...HO9, O4...H9, O3...H8, and H103'...O3'), we have 27.86 and 27.09 kcal mol⁻¹ for S...G and S...S, respectively. This suggests that interunit interactions in the S...G complex are slightly stronger than in the S...S complex, which shows that hydration is one of the phenomena contributing to favor interactions in WBDCs of lignin. However, another type of interaction, such as π interaction between aromatic rings, might also contribute to the overall interaction energy of WBDCs, which should be analyzed in detail.

3.4.2 Energetics of π stacking. Considering all types of π interactions in both complexes (Tables S3 and S4, ESI[†]), the total $E(2)$ of π -stacking was 2.9 kcal mol⁻¹ for the S...S complex, almost double that found for the S...G (1.54 kcal mol⁻¹) complex. By comparing these $E(2)$ values with those found for total H-bonds (70.33 for S...G and 74.23 kcal mol⁻¹ for S...S, Table 3), it can be concluded that the stabilization energies by π interactions do not play a major role compared to hydrogen bonding ones. Furthermore, according to the total $E(2)$ value, it can be claimed that the π -stacking in the S...S complex is much stronger than in the S...G complex, which is due to better overlapping between π and σ^* bond orbitals (Fig. 7b).

The shorter distance (d) and lower interplanar angle (θ) between centroids in the S...S complex, with a more parallel (stacked) geometry, enables a better overlap by the contribution of dispersion forces.¹²⁸ Because of the higher polarizability of electrons in the S...S complex (348.4 a.u.) than in the S...G complex (332.2 a.u.), the dispersion forces are stronger¹²⁹ in the first complex. As is well known, higher polarizability enables electrons to more easily redistribute and allows the formation of stronger poles, which creates a higher attractive London

dispersion energy (eqn (9)).¹³⁰

$$E_{AB}^{\text{disp}} \approx -\frac{3}{2} \frac{I_A I_B}{I_A + I_B} \frac{\alpha^A \alpha^B}{R^6} \quad (9)$$

where E_{AB}^{disp} is the stabilization energy produced by the attractive London dispersion forces between two molecules, A and B. α^A and α^B mean their polarizability volumes, while I_A and I_B are their first ionization potentials. R is the separation distance between A and B.

As found in organometallic complexes,¹³⁰ the attractive London dispersion forces might benefit electron transfer from a Lewis base (*e.g.*, π donor orbitals) to a σ^* antibonding orbital in the S...S complex, in which there is a contribution of 1 kcal mol⁻¹ promoted solely by the interaction between π C1-C6 [1] and σ^* O4-HO4 [2] (Fig. 7b). Here, instead of an appropriate π stacking interaction, a local direct interaction is observed between a substituent (*e.g.*, hydroxyl group) of one aromatic ring and the nearby vertex of the other ring¹¹⁵ (*e.g.*, C1-C6) by π - σ^* orbital interaction. Possibly this local interaction is predominant in the S...S complex because -OCH₃ is an electron-withdrawing group, which manifests on the σ -bonding of the aromatic ring, thereby the σ -withdrawing character dominates the effects of substituents in WBDCs.^{74,75} In addition to this effect, the electrostatic attraction in H-bonds holding the rings closer in the S...S complex would allow stronger local direct interactions between substituents of a ring and the nearby vertex of the other ring,^{33,129} as discussed above. There may be certain cooperativity of hydrogen bonding and π -stacking in WBDCs. Unlike the S...S complex, π - σ^* interactions are not abundant in the S...G complex (Fig. S4, ESI[†]), but rather predominate π - π^* interactions. The most predominant interaction in the S...G complex was π C1-C2 [1] and π^* C2-C3 [2] (Fig. 7a) with the value of $E(2)$ of 0.55 kcal mol⁻¹.

All results show that π interactions are relatively much weaker than H-bonds, but they contribute positively to the hydration of units, thus favoring interactions between units in WBDCs. At a supramolecular level, the combined interaction of hydrogen bonding and π -stacking stabilize lignin xylem tissue of tracheary elements of wood whose main function is water conduction.^{32,33} Based on the similar structures found in

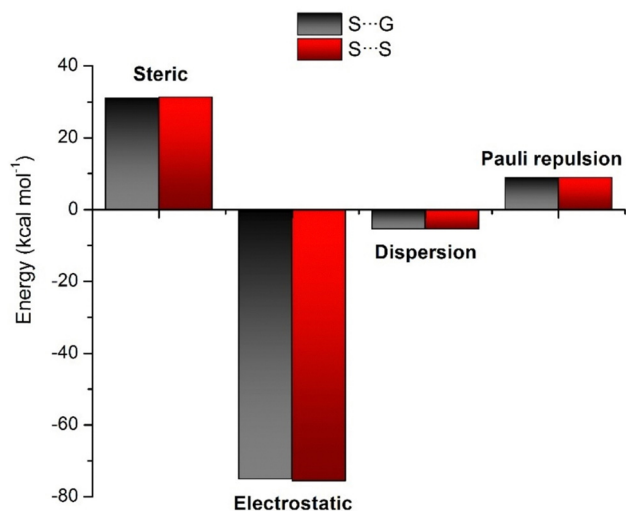


Fig. 8 Energetic contributions of steric, electrostatic, and quantum (attractive dispersion and Pauli repulsion) forces to the overall interaction energy of water-bridged dimeric complexes: S···G and S···S.

the complexes forming quasi-sandwich geometries between the centroids of the aromatic ring, it is quite likely that S-G lignins and S-lignins have similar “coiled” structures in the xylem tissue of wood. However, S-lignins are expected to show more coiled spiral conformation than S-G lignins, which have a higher content of β - β and β -5 bonds.¹³¹

3.5. Energy decomposition by the EDA-SBL scheme

Up to now, we have only considered attractive forces (H-bonds and London dispersion) in the interaction networks of WBDCs, but there are others forces involved in them, such as steric and Pauli repulsion corresponding to the electronic interaction between orbitals. Such repulsive effects of these forces can be measured by decomposing the total interaction energy into different types of energy using an energy partition scheme, such as EDA-SBL. Fig. 8 shows the breakdown of the molecular total energy of S···G and S···S complexes into steric, electrostatic, exchange–correlation, and Pauli dispersion energetic

components using the EDA-SBL scheme. Calculating the energy difference of each contribution in Fig. 8, we found that the addition of one $-\text{OCH}_3$ group to the S···G complex to produce the S···S complex provides an increase in steric and Pauli repulsion in 0.22 and 0.08 kcal mol⁻¹, while there is also an increase in attractive electrostatic and dispersion forces in 0.58 and 0.03 kcal mol⁻¹, respectively. On the one hand, the results show that the electrostatic term is dominant and controls the geometric preference of WBDCs in a quasi-sandwich conformation instead of other interactions coming from dispersion and repulsion due to electronic delocalization of the aromatic rings.⁴⁹ The net contribution of the stacking interactions to the overall interaction energy is surprisingly small in these complexes, as observed in H-bonded “zipper” complexes.¹³²

On the other hand, the better electrostatic stabilization of the S···S complex relative to the S···G complex comes precisely from the additional $-\text{OCH}_3$ substituent group, which explains the higher interaction energy for S···S than for S···G (Table 1). In fact, by summing the found differences values by the EDA-SBL scheme, an energy difference between S···G and S···S equal to -0.31 kcal mol⁻¹ is obtained, which corresponds to the difference calculated between interaction energies (-0.3 kcal mol⁻¹, Table 1) of these complexes. Assuming that there exists substituents additivity on interaction energy of WBDCs, as found in benzene dimers,⁷⁵ one can validate this energy difference between WBDCs. By *ab initio* calculations, the interaction energy of benzene sandwich dimers when substituted with a $-\text{OCH}_3$ group relative to the unsubstituted benzene dimer was in the range of -0.3 to -0.6 kcal mol⁻¹,⁷⁴ while the computed difference (-0.31 kcal mol⁻¹) in our work falls in this range.

All the above results indicate that attractive electrostatic forces and London dispersion forces are positive contributors, while steric and Pauli repulsion forces are negative contributors in the interaction networks of WBDCs. The attractive electrostatic forces (Table S2, ESI†) are the major source of interactions in these dimers by H-bonds, overcoming the steric and Pauli repulsion effects (Fig. 8). The greater number of $-\text{OCH}_3$ substituents in the S···S complex enhances the London dispersion energy¹³³ due to the increase in polarizability of the aromatic ring (eqn (9)), but also

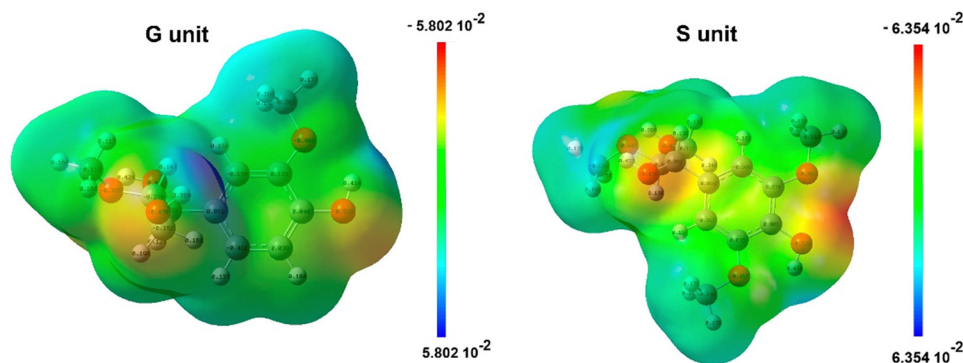


Fig. 9 Molecular electrostatic potential map for guaiacyl (G) and syringyl (S) units in the gaseous phase calculated at the M06-2X/6-31+g(d,p) theory level. The S unit has a richer π electron cloud than the G unit in the center of the substituted rings because $-\text{OCH}_3$ is a π donor group,¹²⁵ which increases the π -electron density by conjugating the lone electron pair of oxygen with the π -cloud of the aromatic ring by resonance effects (the resonance parameter for $-\text{OCH}_3$ is -0.56).⁷⁴

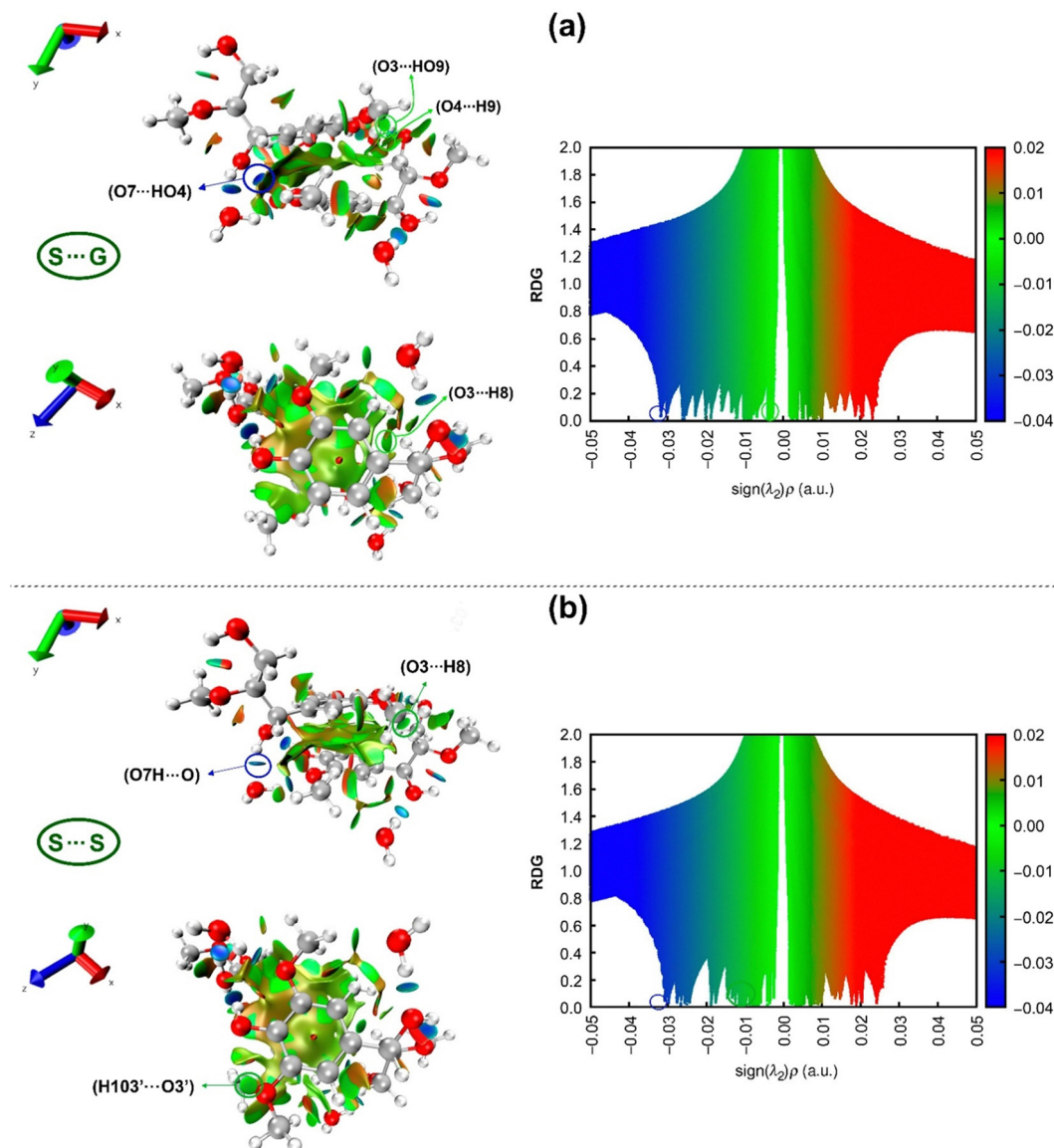


Fig. 10 Non-covalent interaction (NCI) plots (left) of (a) S...G, (b) S...S complexes and corresponding scatter plots (right). The iso-surfaces are colored on a blue-red scale according to values of $\text{sign}(\lambda_2)\rho$, ranging from -0.04 to 0.02 a.u. Blue, green, and red indicate strong attractive, pi-stacking, and repulsion interactions, respectively.

there is an increase in the Pauli repulsion energy because $-\text{OCH}_3$ is a π donor group. In other words, the additional $-\text{OCH}_3$ substituent in the S...S complex relative to the S...G complex provides a slight increase of electron density in the internuclear region of aromatic rings (Fig. 9), hindering the stacking interaction as has been evidenced in the π /polar or electrostatic model.^{49,74} Herein, although the role of dispersion interactions is overweighted by performing the calculations in the gas phase,¹²⁹ such interactions show to be even smaller than Pauli repulsion, which indicates that quantum effects contribute to the destabilization in WBDCs.

3.6. RDG analysis of the complexes

The reduced density gradient (RDG) analysis is employed as another useful method to further study non-covalent interactions, which can supply more reliable data compared to the

AIM method and visualize them. The studied non-covalent interactions are revealed with RDG vs. the sign of the second Hessian eigenvalue ($\text{sign}(\lambda_2)\rho$),⁸² as shown in Fig. 10.

RDG scatter points in Fig. 10 indicate H-bonding attractive interactions at a negative scale (blue color) less than -0.01, van der Waals interactions (green color) at a scale from -0.01 to 0.01,¹³⁴ and steric repulsions at the positive scale of $\text{sign}(\lambda_2)\rho$ greater than 0.01 "red color". Although the difference between these two RDG scatters appears to be subtle, some differentiating details between S...G and S...S complexes can be drawn from them. The hydrogen bond AIM analysis states that the strongest intermolecular H-bond for the S...G complex is O7...HO4, while for the S...S complex, it is O7H...O, which are visualized with spikes at values of $\text{sign}(\lambda_2)\rho$, respectively of -0.0318 and -0.0308 a.u. These spikes are marked by blue

circles on the right-hand side of Fig. 10, corresponding to blue circled disk-shaped blocks in the structures on the left-hand side of this figure. The more negative values of $\text{sign}(\lambda_2)\rho$ for $\text{O7H}\cdots\text{O}$, $\text{O4H}\cdots\text{O}$, $\text{O3}\cdots\text{H}$, $\text{H103}\cdots\text{O}$ and $\text{H2}\cdots\text{O}$ define a higher spike density zone in the RDG plot for the $\text{S}\cdots\text{S}$ complex, which is an indication of its higher hydration capacity than the $\text{S}\cdots\text{G}$ complex, in consonance with NBO analysis. Around -0.01 a.u., it is possible to observe a higher spike density for the $\text{S}\cdots\text{S}$ complex than for the $\text{S}\cdots\text{G}$ complex due to the attractive stronger $\text{O3}\cdots\text{H8}$ interunit interaction located at -0.0125 a.u. and the appearance of a new interaction ($\text{H103}'\cdots\text{O3}'$) at -0.0112 a.u. binding two S units in the $\text{S}\cdots\text{S}$ complex.

Interestingly, introducing a $-\text{OCH}_3$ group in the $\text{S}\cdots\text{G}$ complex turns off the weak $\text{O3}\cdots\text{HO9}$ and $\text{O4}\cdots\text{H9}$ interactions at -0.0093 and -0.0102 a.u., respectively (Table 2). This is caused by the +M effect of the $-\text{OCH}_3$ group, which depletes the negative charge on O3 and O4 atoms and decreases the attractive electrostatic forces in such interactions (Table S2, ESI†). It is important to highlight that although $\text{O3}\cdots\text{H8}$, $\text{H103}'\cdots\text{O3}'$, $\text{O3}\cdots\text{HO9}$ and $\text{O4}\cdots\text{H9}$ interactions are classified as van der Waals interactions according to the value of $\text{sign}(\lambda_2)\rho$ (around -0.01), they are likely “Keesom”¹³⁵ interactions given by dipole/dipole forces governed by electrostatic effects. Finally, analyzing the steric repulsion at values of $\text{sign}(\lambda_2)\rho > 0.01$ in RDG scatter points (the red region in Fig. 10), it can be noted that in the $\text{S}\cdots\text{S}$ complex, there is a slightly higher density of spikes than in the $\text{S}\cdots\text{G}$ complex. The above reveals that the additional $-\text{OCH}_3$ group in the $\text{S}\cdots\text{S}$ complex induces steric repulsions between carbon (e.g., C10) of this group and carbons belonging to the aromatic ring (e.g., C2) (Fig. 2), besides the intrinsic electronic repulsion between centers of the aromatic rings (Pauli repulsion), as indicated by the red ellipsoid in the structures of the left panel in Fig. 10.

4. Conclusions

We have carried out DFT calculations at a reliable theory level (M06-2X/6-31+G(d,p)) on $\text{S}\cdots\text{G}$ and $\text{S}\cdots\text{S}$ complexes forming “coiled” structures into globular conformations of lignin. Regarding dimer ring orientation, the optimized structures of WBDCs by DFT were like the starting structures achieved by molecular dynamics studies. This indicates that new parameterizations of the CHARMM force field for lignin adequately reproduce structural features of lignin dimers with hydrogen bonding. The analysis methods (AIM, NBO, EDA-SBL, and RDG) suggest that in the $\text{S}\cdots\text{S}$ complex there is a stronger intermolecular interaction than in the $\text{S}\cdots\text{G}$ complex due to the interaction of the H-bonds with the water molecules acting as the glue between the units and $\pi-\sigma^*$ local direct-interaction between the substituent of one aromatic ring and the nearby vertex of the other ring. The hydrogen bonding interactions are predominant relative to π -stacking interactions, which include London dispersion and Pauli repulsion. Electrostatic interactions control the quasi-sandwich geometries between centroids

of aromatic rings in these complexes and not by $\pi-\pi$ interactions, in consonance with other studies of aromatic systems in the literature. However, despite the more favorable interactions in the $\text{S}\cdots\text{S}$ complex, it shows a less thermodynamic stability than the $\text{S}\cdots\text{G}$ complex due to the entropic contributions coming from the presence of the methoxy group in the S-unit. From a physical interaction point of view, our computational results suggest that S-G lignins are more stable or recalcitrant than S-lignins in biomass dissolution, even when branched interunit bonds (e.g., $\beta-\beta$ and $\beta-5$) were not considered in our simple models. The similarity of WBDC structures indicates that S-G lignins and S-lignins are likely to have similar coiled structures in the xylem tissue of wood, but the stronger intermolecular interaction between S units mediated by H-bonds in S-lignins leads to somewhat more coiled and compact structures. Thus, as a biorefinery strategy to fraction lignin rich in S units might be used in polar solvents to replace intermolecular hydrogen bonds between S units and water molecules due to the increased hydration capacity of the $\text{S}\cdots\text{S}$ complex compared to the $\text{S}\cdots\text{G}$ complex. In the same way that with ionic liquids, polar solvents can dissolve S-lignin through favorable interactions with the aliphatic hydroxyl group in the α -position as the key site or the aromatic hydroxyl group as the secondary site. Future DFT calculations should be performed with these complexes using solvents different from water to evaluate which solvent system is better for dissolving them.

Conflicts of interest

The authors declare no conflicts of interest.

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