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Bisquinoline-based fluorescent cadmium sensors

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Rational molecular design afforded fluorescent Cd^{2+} sensors based on bisquinoline derivatives. Introduction of three methoxy groups at the 5,6,7-positions of the quinoline rings of BQDMEN (*N,N'*-bis(2-quinolylmethyl)-*N,N'*-dimethylethylenediamine) resulted in the reversal of metal ion selectivity in fluorescence enhancement from zinc to cadmium. Introduction of bulky alkyl groups and an *N,N*-bis(2-quinolylmethyl)amine structure, as well as replacement of one of the two tertiary amine binding sites with an oxygen atom and the use of a 1,2-phenylene backbone significantly improved the Cd^{2+} specificity. The fluorescent cadmium ion selectivity could be explained by the differential binding with Cd^{2+} and Zn^{2+} , and the formation of a bis(μ -chloro) dinuclear cadmium complex in contrast to the mononuclear zinc complex.

1. Introduction

Detection and quantification of toxic heavy metal ions are currently important objectives in regulation of environmental pollutants. Fluorescence sensing provides high selectivity and sensitivity toward targeted metal ions *via* a rapid analytical protocol using relatively inexpensive equipment. One of the most important issues in the detection of toxic metal ions is discrimination of Cd^{2+} from naturally abundant Zn^{2+} , both of which are group 12 elements in the periodic table and exhibit only a 21 pm difference in their ionic radii. Since recent human activities in industry increase the exposure of cadmium to air, water and soil, continuous development of fluorescent probes for strict detection of cadmium, especially in environmental water, is in high demand. Many research groups have been investigating quinoline-based molecules for this purpose considering that the long coordination distances and softness of the quinoline nitrogen atom in comparison with conventional pyridine ligands would be suitable for selective binding with Cd^{2+} .^{1–10} In this Frontier article, several modifications of the molecular structure of BQDMEN (*N,N'*-bis(2-quinolylmethyl)-*N,N'*-dimethylethylenediamine (**1**), Fig. 1)¹¹ for better Cd^{2+} -selectivity in the fluorescence response are discussed (Table 1). A preceding Frontier article by the present author, dealing with a related hexadentate tetrakisquinoline

ligand TQEN (*N,N,N',N'*-tetrakis(2-quinolylmethyl)ethylenediamine, Fig. 1), contains other examples for fluorescence sensing of Zn^{2+} , Cd^{2+} , Hg^{2+} and phosphate species.¹²

2. Polymethoxy substitution on the quinoline ring

The BQDMEN **1** has been reported as a tetranitrogen ligand supporting mononuclear and dinuclear metal complexes including Fe^{2+} , Co^{2+} , Ni^{2+} , Pd^{2+} , $\text{Cu}^{2+/+}$, Ru^{2+} , Cr^0 and $\text{Mn}^{3+/2+}$ centres.^{13–20} We utilized this skeleton for fluorescent Zn^{2+} sensors and improved the fluorescence quantum yield by introducing a methoxy group to the quinoline rings.¹¹ The fluorescence turn-on mechanism includes the inhibition of photo-induced electron transfer (PeT) and chelation enhanced fluorescence (CHEF).^{21–23} Thus-obtained 6-MeOBQDMEN (**2**) (Fig. 1) exhibits sufficient fluorescence intensity and Zn^{2+} selectivity ($\phi_{\text{Zn}} = 0.28$ and $I_{\text{Zn}}/I_{\text{Cd}} = 2.2$ in DMF– H_2O (1 : 1)†) for intracellular study. On the other hand, the unsubstituted BQDMEN **1** exhibits dim fluorescence ($\phi_{\text{Zn}} = 0.017$) but slightly higher Zn^{2+} selectivity ($I_{\text{Zn}}/I_{\text{Cd}} = 3.9$) than **2**. Although we have known that further methoxy substitution does not improve the quantum yield from our study on TQEN derivatives,²⁴ we expected that the fluorescence metal ion selectivity could shift (or even reverse) to the Cd^{2+} side. This estimation is indeed true, in which TriMeOBQDMEN **3** (Fig. 1) exhibits Cd^{2+} selectivity ($I_{\text{Zn}}/I_{\text{Cd}} = 0.22$) with similar fluorescence quantum yield ($\phi_{\text{Cd}} = 0.29$) to the Zn^{2+} complex with **2** (Table 1).²⁵ Interestingly, the extension of the carbon chain in the mole-

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†The DMF– H_2O (1 : 1) was used as a solvent for spectroscopic measurements in the first half (Sections 2 and 3) of this article.



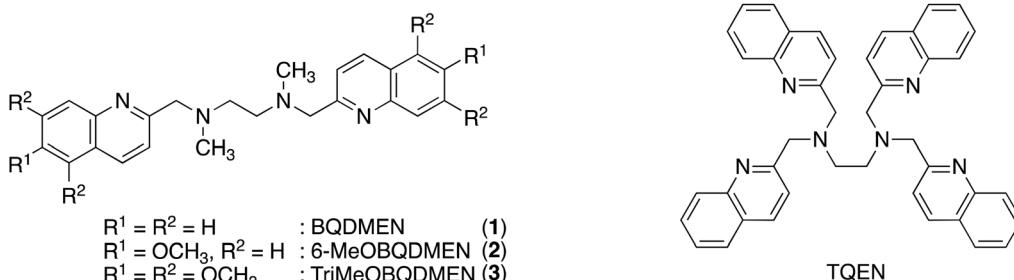


Fig. 1 Structure of BQDMEN derivatives 1–3 and TQEN.

Table 1 Fluorescence properties of Cd^{2+} probes

	λ_{ex} (nm)	λ_{em} (nm)	I_{Cd}/I_0 (eq.)	$I_{\text{Zn}}/I_{\text{Cd}}$ (eq.)	K_d (metal ion) (M)	ϕ_{Cd}	Ref.
TriMeOBQDMEN (3) ^a	339	472	15 (1)	0.22 (1)	$\sim 10^{-8}$ (Cd) $\sim 10^{-7}$ (Zn)	0.29	25
3 ^b	335	458	25 (3)	0.24 (3)	3×10^{-7} (Cd) 3×10^{-6} (Zn)	—	29
TriMeOBQDIEN (5) ^a	339	459	23 (20)	0.04 (20)	4×10^{-6} (Cd) 5×10^{-5} (Zn)	0.37	26
TriMeOBQDBEN (6) ^a	337	459	27 (3)	0.20 (3)	$\sim 10^{-7}$ (Cd) 8×10^{-6} (Zn)	0.33	26
TriMeOBQDMPHEN (9) ^b	341	467	51 (1)	0.26 (1)	8×10^{-7} (Cd) 6×10^{-6} (Zn)	0.43	27
TriMeOBQDIPHEN (10) ^b	343	480	427 (40)	0.05 (40)	1×10^{-3} (Cd) 4×10^{-2} (Zn)	—	27
TriMeO- <i>N,N</i> -OBQMAE (11) ^b	338	470	10 (3)	0.34 (3)	1×10^{-5} (Cd) 8×10^{-5} (Cd)	0.28	29
TriMeO- <i>N,N</i> -BQMAE (12) ^b	340	465	12 (3)	0.26 (3)	$\sim 10^{-7}$ (Cd) 2×10^{-5} (Zn)	0.23	29
TriMeOBQMOA (14) ^b	341	464	90 (5)	0.02 (5)	1×10^{-5} (Cd)	0.18	30

^a In DMF-H₂O (1 : 1). ^b In DMF-HEPES buffer (1 : 1).

cular skeleton from ethylenediamine to propanediamine also shifted the fluorescence response toward the Cd^{2+} side.

The fluorescence titration revealed that TriMeOBQDMEN (3) binds an equimolar amount of metal ions. The Cd^{2+} complex with 3 exhibits significantly long fluorescence lifetime ($\tau = \sim 30$ ns) in comparison with the Zn^{2+} complex (~ 20 ns), indicating the different complex structure and/or fluorescence pathway. X-ray crystallography revealed the structures of the bis(μ -chloro) dinuclear cadmium complex and mononuclear zinc complex with BQDMEN (1) (Fig. 2). ESI-MS also confirmed the formation of $[(\mu\text{-Cl})_2\text{Cd}_2(3)_2](\text{ClO}_4)_2$ species under the spectral measurement conditions. The metal binding affinity of 3 is high enough to saturate the ligand in the presence of 1 equiv. of Cd^{2+} and Zn^{2+} under our experimental conditions (Table 1). These observations clearly indicate that the BQDMEN ligand discriminates Cd^{2+} and Zn^{2+} by differential complexation, and the introduction of three methoxy groups on the quinoline rings successfully highlights the difference in the complex structure *via* fluorescence outputs. The next step would be an improvement of the insufficient Cd^{2+} selectivity of 3. Several modifications of the molecular structure were explored to selectively reduce the Zn^{2+} -induced fluorescence signal.

3. Introduction of bulky alkyl groups

Two methyl groups of TriMeOBQDMEN (3) attached to the aliphatic nitrogen atoms were replaced with isopropyl groups to afford TriMeOBQDIEN (*N,N*-bis(5,6,7-trimethoxy-2-quinolylmethyl)-*N,N*-diisopropylethylenediamine (5), Fig. 3).²⁶ This sterically hindered bisquinoline derivative exhibits weak metal binding affinity (Table 1) and shows characteristic fluorescence titration profiles responding to Cd^{2+} and Zn^{2+} (Fig. 4). The highly fluorescent Cd^{2+} complex with 5 ($\phi_{\text{Cd}} = 0.37$) was formed by the addition of ~ 5 equiv. of Cd^{2+} and was stable in the excess amount of metal ions; however, the gradual addition of Zn^{2+} to 5 exhibited small fluorescence enhancement around 5 equiv. of metal ions, and then, the fluorescence significantly decreased in the presence of an excess amount of zinc salt, $\text{Zn}(\text{ClO}_4)_2$. The monomethoxy derivative (4) (Fig. 3) also exhibits similar profiles. After careful investigations, this observation was explained by the selective decomplexation of the zinc complex by protons generated by excess $\text{Zn}(\text{ClO}_4)_2$ used in the titration experiment. $\text{Cd}(\text{ClO}_4)_2$ did not drop the pH significantly, and the use of $\text{Zn}(\text{NO}_3)_2$ or DMF-HEPES buffer (1 : 1) as a solvent did not disrupt the Zn^{2+} -induced fluorescence even in the presence of a large amount of metal ions. Nevertheless, the



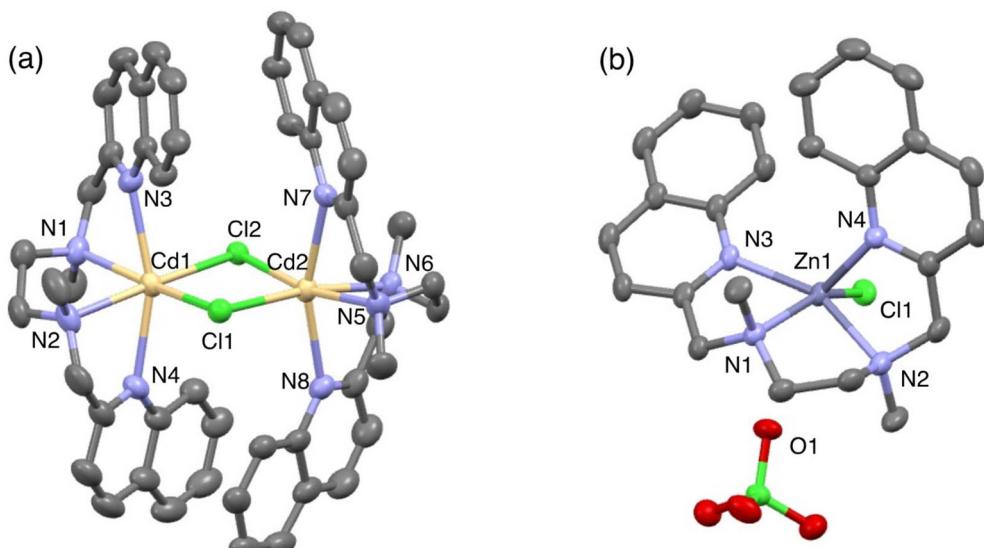


Fig. 2 Perspective view of (a) $[(\mu\text{-Cl})_2\text{Cd}_2(1)_2]^{2+}$ (CSD refcode CIBKOR) and (b) $[\text{Zn}(1)\text{Cl}]\text{ClO}_4$ (CSD refcode CIBKUX). Reproduced from ref. 25 with permission from the Royal Society of Chemistry.

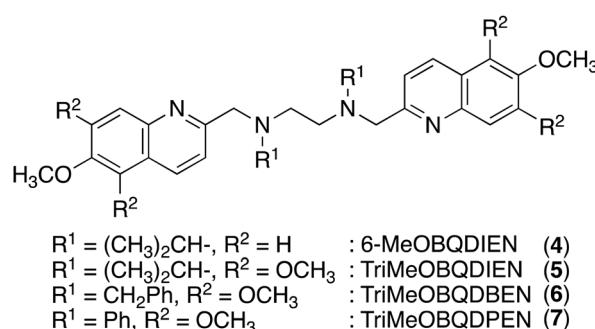


Fig. 3 Structure of BQDMEN derivatives 4–7.

excellent Cd²⁺ specificity in fluorescence enhancement achieved in 5 ($I_{\text{Zn}}/I_{\text{Cd}} = 4\%$ in the presence of 20 equiv. of metal ions) in DMF–H₂O (1 : 1) is of significant interest.

Although X-ray crystallography employing 4 afforded only mononuclear Cd²⁺ complexes, the long fluorescence lifetime ($\tau = \sim 30$ ns) for the Cd²⁺ complex with 5 suggested the possible formation of bis(μ -chloro) dicadmium species assembled in the excited state even in the isopropyl derivative. The corresponding benzyl (TriMeOBQDBEN (6)) and phenyl (TriMeOBQDPEN (7)) derivatives (Fig. 3) resulted in similar metal ion selectivity and fluorescence stability with 3 and an extremely poor fluorescence response with any metal ions, respectively.

4. Introduction of a benzene skeleton

Instead of the introduction of bulky substituents, replacement of an ethylenediamine backbone of BQDMEN (1) with a 1,2-

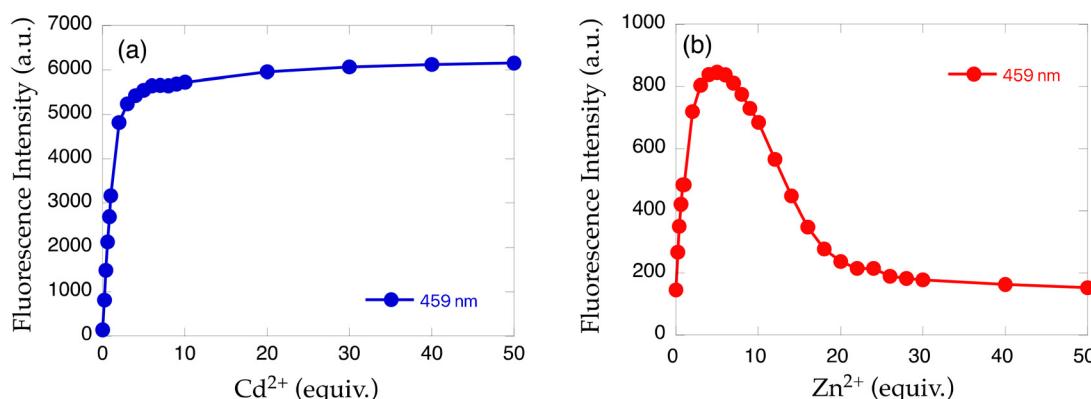


Fig. 4 Fluorescence intensity plot of 5 with increasing amounts of (a) Cd²⁺ and (b) Zn²⁺.²⁶

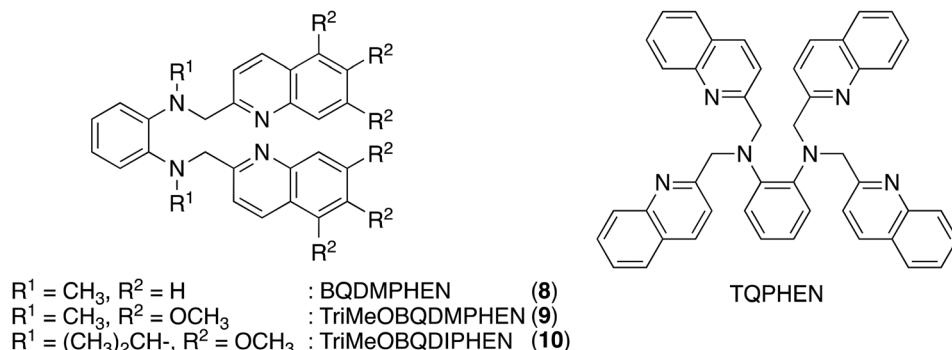


Fig. 5 Structure of BQDMPHEN derivatives 8–10 and TQPHEN.

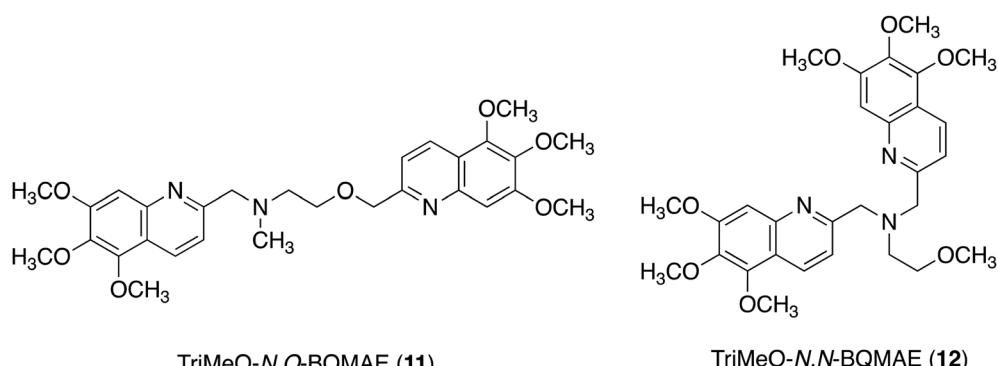


Fig. 6 Structure of BQMAE derivatives 11 and 12.

phenylenediamine also reduces metal binding affinity because the anilinic nitrogen atoms of BQDMPHEN (*N,N*-bis(2-quinolylmethyl)-*N,N*'-dimethyl-1,2-phenylenediamine (8), Fig. 5) have weaker basicity than the aliphatic nitrogen binding site of 1.²⁷ This approach has been successful for the TQEN-based tetrakisquinoline compound, TQPHEN (*N,N,N',N'*-tetrakis(2-quinolylmethyl)-1,2-phenylenediamine, Fig. 5), which scarcely binds to Zn²⁺ under the spectral measurement conditions and therefore exhibits Cd²⁺-specific fluorescence enhancement.²⁸ Upon addition of metal ions in DMF-HEPES buffer (1 : 1),‡ the ligand 8 exhibited similar fluorescence enhancement with Cd²⁺ and Zn²⁺ ($I_{\text{Zn}}/I_{\text{Cd}} = 0.90$ in the presence of 1 equiv. of metal ion), but apparent Cd²⁺ selectivity ($I_{\text{Zn}}/I_{\text{Cd}} = 0.26$ in the presence of 1 equiv. of metal ions) was achieved in the trimethoxy derivative (TriMeOBQDMPHEN (9), Fig. 5). The insufficient metal ion selectivity was further improved by introducing isopropyl groups as discussed above, affording high fluorescent Cd²⁺ specificity in TriMeOBQDIPHEN (10) ($I_{\text{Zn}}/I_{\text{Cd}} = 0.05$ in the presence of 40 equiv. of metal ions) (Fig. 5).

‡ The DMF-HEPES buffer (1 : 1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) was used as a solvent for spectroscopic measurements in the second half (Sections 4–6) of this article, considering the pH change in the presence of large amount of metal salts.

Considering the extremely weak metal binding affinity of the isopropyl ligand 10 ($K_{\text{d}}(\text{Cd}) = 1 \times 10^{-3}$ M), the methyl counterpart 9 ($K_{\text{d}}(\text{Cd}) = 8 \times 10^{-7}$ M) is more suitable for practical use (Table 1). This high metal binding affinity causes diminished Cd²⁺/Zn²⁺ selectivity estimated from the ratio of binding constants ($K_{\text{d}}(\text{Zn})/K_{\text{d}}(\text{Cd}) = 8$ and 32 for 9 and 10, respectively); however, the high fluorescence quantum yield ($\phi_{\text{Cd}} = 0.43$) and wide pH window (pH = 4–10) of 9 are worth exploring. The structures of Cd²⁺ and Zn²⁺ complexes with 8 are both mononuclear but, here again, the long fluorescence lifetime ($\tau = \sim 30$ ns) for Cd²⁺ complexes with 9 and 10 suggests the excimer-like fluorescence mechanism including quinoline–quinoline interactions in the excited state.

4. Replacement of a nitrogen atom with an oxygen atom

Another strategy to reduce the metal binding affinity of the tetranitrogen ligand TriMeOBQDMEN (3), aiming at the improvement of fluorescent Cd²⁺ selectivity, is the replacement of one of the two aliphatic ethylenediamine nitrogen atoms of 3 with an oxygen atom. Thus-designed 2-aminoethanol-based N3O1 ligand, TriMeO-*N,O*-BQMAE (*N,O*-bis(5,6,7-trimethoxy-2-quinolylmethyl)-2-methylaminoethanol (11), Fig. 6) exhibited two

order lower metal binding affinity and slightly decreased Cd^{2+} preference in fluorescence enhancement in comparison with 3 ($I_{\text{Zn}}/I_{\text{Cd}} = 0.34$ and 0.24 for 11 and 3, respectively, in the presence of 3 equiv. of metal ions in DMF-HEPES buffer (1:1)) (Table 1).²⁹ Interestingly, TriMeO-*N,N*-BQMAE (12) (Fig. 6), the regioisomer of 11, in which both quinoline moieties are attached to the same aliphatic nitrogen atom of the aminoethanol skeleton, exhibits one order higher metal binding affinity than 11 and similar fluorescent Cd^{2+} selectivity to 3 ($I_{\text{Zn}}/I_{\text{Cd}} = 0.26$ for 12). The fluorescence quantum yield of 12 ($\phi_{\text{Cd}} = 0.23$) is slightly smaller in comparison with those of 3 (0.29) and 11 (0.28) under the same experimental conditions. This difference is also reflected by the fluorescence lifetimes, where long fluorescence lifetimes ($\tau = \sim 30$ ns) with a possible intramolecular excimer emission for the Cd^{2+} complex were observed only for 3 and 11. For the *N,N*-isomer 12, no difference in fluorescence lifetimes of Zn^{2+} and Cd^{2+} complexes was detected ($\tau = \sim 18$ ns), indicating the monomer emission for both metal complexes with this ligand. The ligand structure of 12 strictly controls the metal binding affinity, possibly due to the difference in the complex structures, affording the fluo-

rescent Cd^{2+} selectivity. It is important to note that the two order difference in dissociation constants between Cd^{2+} and Zn^{2+} complexes was achieved in the *N,N*-bisquinoline structure of 12 (Table 1). The corresponding *N,N*-isomer for 3 has not been examined due to problems in its synthesis.

5. Use of a benzene skeleton with an oxygen binding site

Finally, the N3O1 ligand with a 1,2-phenylene skeleton was explored. Considering the synthetic accessibility, only the *N,N*-isomer was examined here. Thus-designed TriMeOBQMOA (*N,N*-bis(5,6,7-trimethoxy-2-quinolylmethyl)-2-methoxyaniline (14), Fig. 7) exhibits excellent Cd^{2+} specificity in fluorescence enhancement ($I_{\text{Zn}}/I_{\text{Cd}} = 0.02$ in the presence of 5 equiv. of metal ions) with moderate metal binding affinity ($K_{\text{d}}(\text{Cd}) = 1 \times 10^{-5}$ M) and fluorescence quantum yield ($\phi_{\text{Cd}} = 0.18$) (Table 1).³⁰ The binding affinity of 14 with Zn^{2+} is extremely weak. The LOD (limit of detection) for Cd^{2+} was estimated to

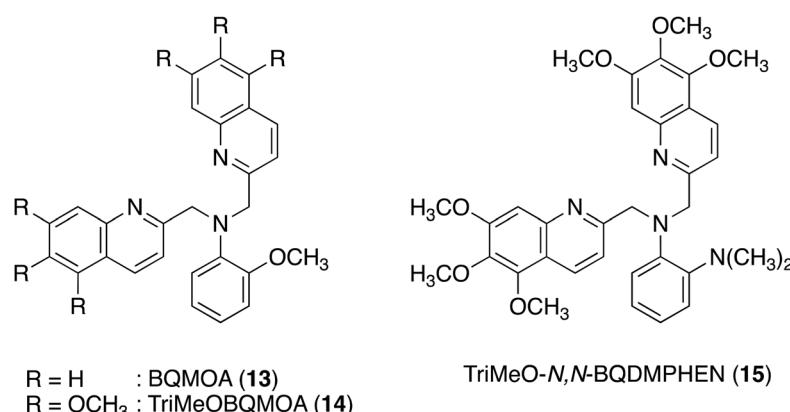


Fig. 7 Structure of BQMOA and BQDMPHEN derivatives 13–15.

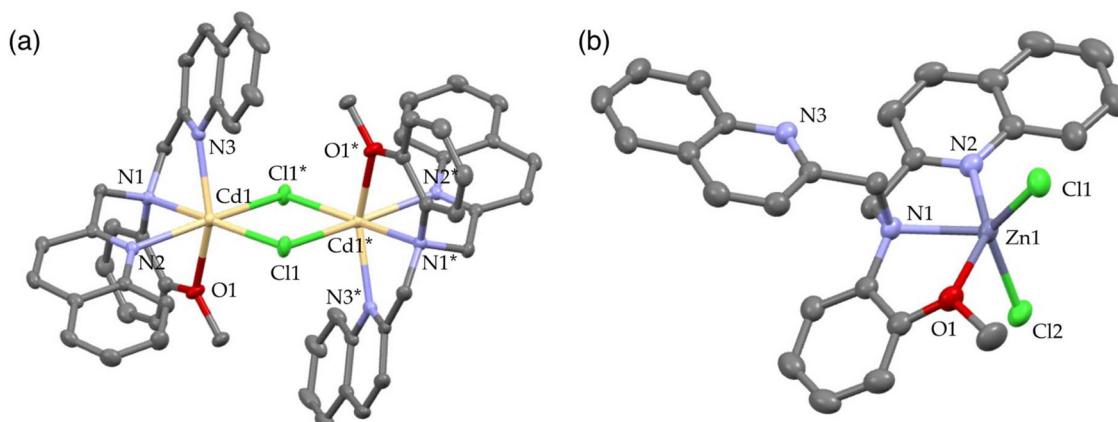


Fig. 8 Perspective view of (a) $[(\mu\text{-Cl})_2\text{Cd}_2(13)_2]^{2+}$ (CSD refcode YOMQOK) and (b) $[\text{Zn}(13)\text{Cl}_2]$ (CSD refcode YOMQUQ). Adapted from ref. 30 with permission from American Chemical Society, copyright 2024.



be 25 nM, which is lower than the environmental limit of water in Japan (3 ppb, 27 nM).

Both Cd^{2+} and Zn^{2+} complexes with BQMOA (**13**) (Fig. 7) were structurally characterized by X-ray crystallography (Fig. 8) where the bis(μ -chloro) dinuclear cadmium complex ($[(\mu\text{-Cl})_2\text{Cd}_2(\mathbf{13})_2]^{2+}$) and mononuclear zinc complex ($[\text{Zn}(\mathbf{13})\text{Cl}_2]$) with an uncoordinated quinoline moiety were revealed. These complex structures clearly explain the reason for the short fluorescence lifetime ($\tau = \sim 13$ ns) for the Cd^{2+} and Zn^{2+} complexes

with **14** because of the lack of interquinoline stacking interaction even in the bis(μ -chloro) dicadmium complex. The fluorescent Cd^{2+} specificity is a result of the extremely weak binding affinity of this ligand with Zn^{2+} .

The importance of chloride ions in HEPES buffer was also disclosed in this study. As shown in Fig. 9, chloride ions exclusively enhance the fluorescence intensity of **14** in the presence of Cd^{2+} *via* enhanced complexation, but no such effect was observed for Zn^{2+} with **14** and dimethylamino derivative

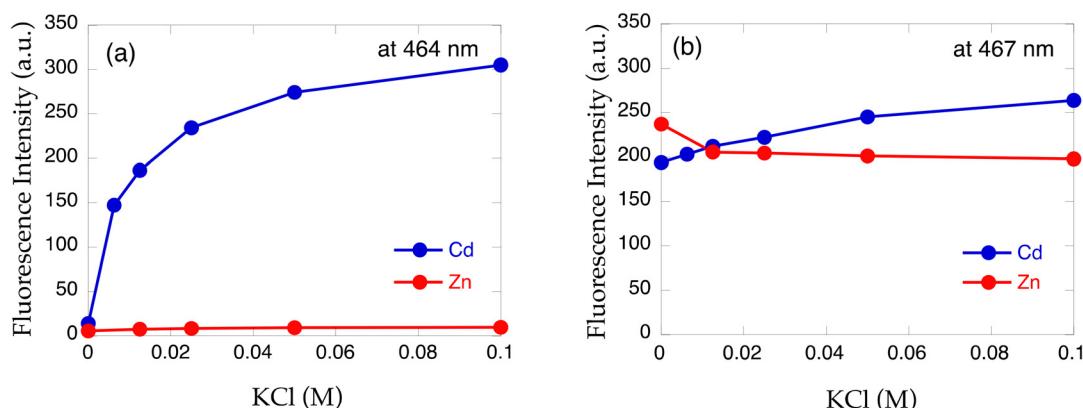


Fig. 9 Fluorescence intensity plot of (a) **14** and (b) **15** in the presence of Cd^{2+} (blue) or Zn^{2+} (red) with increasing amount of KCl. Adapted from ref. 30 with permission from American Chemical Society, copyright 2024.

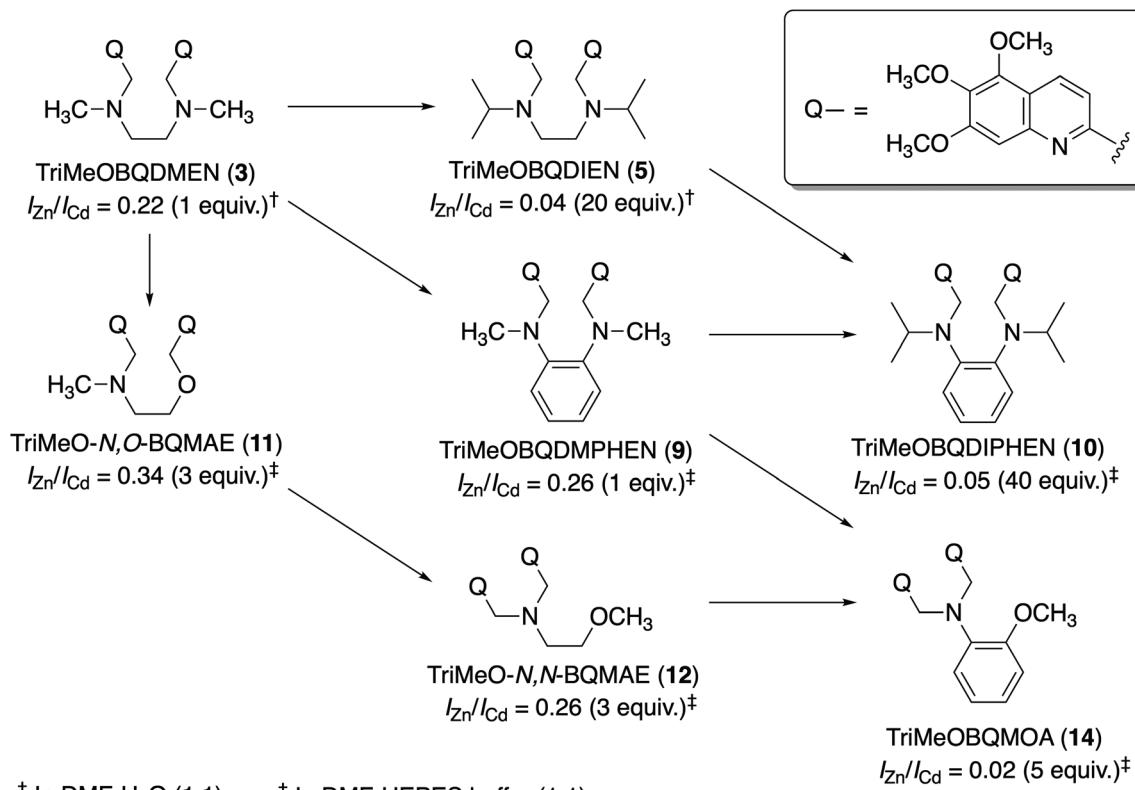


Fig. 10 Summary of the structural modifications of Cd^{2+} probes.

TriMeO-*N,N*-BQDMPHEN (*N,N*-bis(5,6,7-trimethoxy-2-quinolylmethyl)-*N',N'*-dimethyl-1,2-phenylenediamine (**15**), Fig. 7) with Cd²⁺ and Zn²⁺. Not only the moderate metal binding affinity but also the appropriate size and coordination environment of the metal binding cavity of the methoxy derivative **14** in the presence of chloride ions are necessary for fluorescence discrimination of Cd²⁺ from Zn²⁺.

6. Conclusions

Based on the structure of BQDMEN (**1**), a set of fluorescent Cd²⁺ sensors were rationally designed (Fig. 10). The extensive use of polymethoxy-substituted quinolines as a metal binding motif and chromophore provides several unique characteristics suitable for strict discrimination of Cd²⁺ from Zn²⁺ via fluorescence signals. Unprecedented fluorescence enhancement mechanisms including intramolecular excimer emission from the bis(μ -chloro) dinuclear cadmium complex were clarified by X-ray crystallography. Tiny changes in the molecular structure of the fluorescent probe largely affect the structure and stability of the resulting metal complexes, affording significant improvement in fluorescent metal ion selectivity and optical properties. In particular, the size and flexibility of the metal binding pocket with a carefully organized coordination geometry are crucial for selective accommodation of the target metal ion. This Frontier article may provide a useful dataset for future molecular design of new fluorescent sensors targeting a wide variety of metal ions.

Data availability

No primary research results, software or code have been included and no new data were generated or analysed as part of this review.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 X.-L. Tang, X.-H. Peng, W. Dou, J. Mao, J.-R. Zheng, W.-W. Qin, W.-S. Liu, J. Chang and X.-J. Yao, *Org. Lett.*, 2008, **10**, 3653–3656.
- 2 L. Xue, G. Li, Q. Liu, H. Wang, C. Liu, X. Ding, S. He and H. Jiang, *Inorg. Chem.*, 2011, **50**, 3680–3690.
- 3 M. Mameli, M. C. Aragoni, M. Area, C. Caltagirone, F. Demartin, G. Farruggia, G. De Filippo, F. A. Devillanova, A. Garau, F. Isaia, V. Lippolis, S. Murgia, L. Prodi, A. Pintus and N. Zaccheroni, *Chem. – Eur. J.*, 2010, **16**, 919–930.
- 4 X. Zhou, P. Li, Z. Shi, X. Tang, C. Chen and W. Liu, *Inorg. Chem.*, 2012, **51**, 9226–9231.
- 5 H. Tian, B. Li, J. Zhu, H. Wang, Y. Li, J. Xu, J. Wang, W. Wang, Z. Sun, W. Liu, X. Huang, X. Yan, Q. Wang, X. Yao and Y. Tang, *Dalton Trans.*, 2012, **41**, 2060–2065.
- 6 Y. Li, H. Chong, X. Meng, S. Wang, M. Zhu and Q. Guo, *Dalton Trans.*, 2012, **41**, 6189–6194.
- 7 Y. Ma, F. Wang, S. Kambam and X. Chen, *Sens. Actuators, B*, 2013, **188**, 1116–1122.
- 8 X.-J. Jiang, M. Li, H.-L. Lu, L.-H. Xu, H. Xu, S.-Q. Zang, M.-S. Tang, H.-W. Hou and T. C. W. Mak, *Inorg. Chem.*, 2014, **53**, 12665–12667.
- 9 E. Hrishikesan, R. Manjunath and P. Kannan, *J. Solution Chem.*, 2016, **45**, 907–919.
- 10 D. Udhayakumari, *J. Mol. Struct.*, 2023, **1287**, 135715.
- 11 Y. Mikata, A. Yamashita, A. Kawamura, H. Konno, Y. Miyamoto and S. Tamotsu, *Dalton Trans.*, 2009, 3800–3806.
- 12 Y. Mikata, *Dalton Trans.*, 2020, **49**, 17494–17504.
- 13 B. Rieger, A. S. Abu-Surrah, R. Fawzi and M. Steiman, *J. Organomet. Chem.*, 1995, **497**, 73–79.
- 14 A. S. Abu-Surrah, U. Thewalt and B. Rieger, *J. Organomet. Chem.*, 1999, **587**, 58–66.
- 15 A. S. Abu-Surrah, *Asian J. Chem.*, 2002, **14**, 1251–1256.
- 16 Y. Mikata, H. So, A. Yamashita, A. Kawamura, M. Mikuriya, K. Fukui, A. Ichimura and S. Yano, *Dalton Trans.*, 2007, 3330–3334.
- 17 L. S. Morris, M. P. Girouard, M. H. Everhart, W. E. McClain, J. A. van Paridon, R. D. Pike and C. Goh, *Inorg. Chim. Acta*, 2014, **413**, 149–159.
- 18 M. Saga, T. Anamushi, W. Miyahara, S. Yamazaki and K. Saito, *Anal. Sci.*, 2015, **31**, 185–189.
- 19 N. Singh, J. Niklas, O. Poluektov, K. M. Van Heuvelen and A. Mukherjee, *Inorg. Chim. Acta*, 2017, **455**, 221–230.
- 20 R. V. Ottenbacher, A. G. Medvedev, A. A. Nefedov and K. P. Bryliakov, *Inorg. Chem. Commun.*, 2023, **156**, 111282.
- 21 A. P. de Silva, H. Q. N. Gunaratne, T. Gunnlaugsson, A. J. M. Huxley, C. P. McCoy, J. T. Rademacher and T. E. Rice, *Chem. Rev.*, 1997, **97**, 1515–1566.
- 22 B. Daly, J. Ling and A. P. de Silva, *Chem. Soc. Rev.*, 2015, **44**, 4203–4211.
- 23 D. Escudero, *Acc. Chem. Res.*, 2016, **49**, 1816–1824.
- 24 Y. Mikata, M. Wakamatsu, A. Kawamura, N. Yamanaka, S. Yano, A. Odani, K. Morihiro and S. Tamotsu, *Inorg. Chem.*, 2006, **45**, 9262–9268.
- 25 Y. Mikata, M. Tanaka, S. Yasuda, A. Tsuruta, T. Hagiwara, H. Konno and T. Matsuo, *Dalton Trans.*, 2023, **52**, 7411–7420.
- 26 Y. Mikata, S. Yasuda, T. Hagiwara, H. Konno and S. Shoji, *Inorg. Chim. Acta*, 2024, **571**, 122218.



27 Y. Mikata, K. Kawakami, M. Nagaoka, S. Shoji, H. Konno and T. Matsuo, *Inorg. Chim. Acta*, 2024, **565**, 121968.

28 Y. Mikata, A. Kizu and H. Konno, *Dalton Trans.*, 2015, **44**, 104–109.

29 Y. Mikata, A. Tsuruta, H. Koike, S. Shoji and H. Konno, *Molecules*, 2024, **29**, 369.

30 Y. Mikata, N. Tosaka, S. Yasuda, Y. Sakurai, S. Shoji, H. Konno and T. Matsuo, *Inorg. Chem.*, 2024, **63**, 8026–8037.

