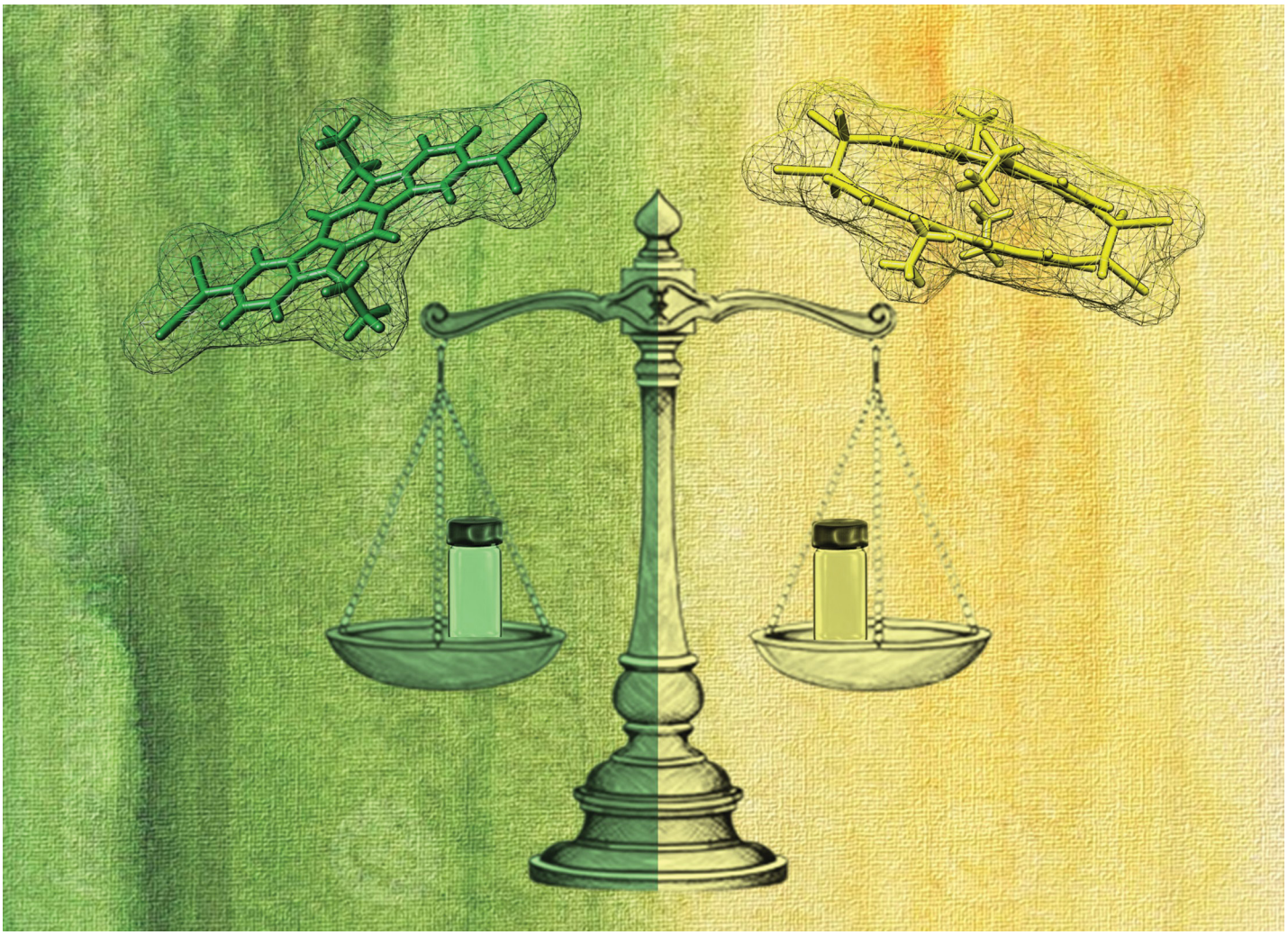

Showcasing research from the groups of Prof. M. Carmen Ruiz Delgado at the University of Málaga, Prof. František Hartl at the University of Reading and Prof. Hongxiang Li at the East China University of Technology.

Tuneable stimuli-responsive behaviour, spectroscopic signatures and redox properties of indolo[3,2-*b*]carbazole-based diradicals

Structural changes stabilising (or destabilising) diradical systems have been rationalised in this work aimed at identifying new design strategies for stimuli-responsive materials. To this end, a new indolo[3,2-*b*]carbazole compound containing terminal dicyanomethylene groups in the meta-position, and its para-isomer reference, demonstrate how their chemical reactivity, physical properties and redox behaviour are affected by the substitution pattern and the core elongation and conjugation.

As featured in:

See Hongxiang Li, František Hartl, M. Carmen Ruiz Delgado *et al.*, *Phys. Chem. Chem. Phys.*, 2024, **26**, 26238.

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Tuneable stimuli-responsive behaviour, spectroscopic signatures and redox properties of indolo[3,2-*b*]carbazole-based diradicals†

Irene Badía-Domínguez,[a] Deliang Wang,[b] Rosie Nash,[c] Víctor Hernández Jolín,[a] David Collison,[d] Muralidharan Shanmugam,[d] Hongxiang Li,*[be] František Hartl *[c] and M. Carmen Ruiz Delgado *[a]

During the last decade, there has been an increasing interest in the rationalisation of how structural changes stabilise (or destabilise) diradical systems. Demonstrated herein is that indolocarbazole (ICz) diradicals, substituted with dicyanomethylene (DCM) groups, are useful motifs for dynamic covalent chemistry by self-assembling from isolated monomers to cyclophane structures. The comparison of ICz-based systems substituted with DCM groups in *para*- or *meta*-positions (***p*-ICz-CN** and ***m*-ICz-CN**) and their short-chain carbazole analogues (***p*-Cz-CN** and ***m*-Cz-CN**) may identify new potential design strategies for stimuli-responsive materials. The principal objectives of this investigation are the elucidation of (i) the connection between diradical character and the cyclophane stability, (ii) the spatial disposition of the cyclophane structures, (iii) the monomer/cyclophane interconversion both in solution and solid state in response to external stimuli and (iv) the impact that the different π-conjugation and electronic communication between the DCM terminals exerts on the electronic adsorption of the diradicals and their redox behavior. The spontaneous nature of the cyclophane structure is supported by the negative relative Gibbs free energies calculated at 298 K and experimentally by UV-Vis and Raman spectroscopy of the initial yellow solid powder. The conversion to monomeric species having diradical character was demonstrated by variable-temperature (VT) EPR, UV-Vis, Raman and IR measurements, resulting in appreciable chromic changes. In addition, electrochemical oxidation and reduction convert the cyclophane dimer **(*m*-ICz-CN)$_2$** to the monomer monocations and dianions, respectively. This research demonstrates how the chemical reactivity and physical properties of π-conjugated diradicals can be effectively tuned by subtle changes in their chemical structures.

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1. Introduction

Open-shell organic conjugated materials quickly attracted the attention of scientists because of their new chemical and physical properties with great potential in flexible organic electronics,[1–4] spintronics,[5,6] stimuli-responsive materials,[7–17] spin-crossover materials[18–21] or molecular self-assembly.[22,23] Open-shell organic materials, known as radicals (classified as monoradicals, diradicals and other high-spin radicals), are small organic molecules with unpaired electrons. These molecules can build cyclophane structures by forming C–C σ-bonds between radical centres longer than typical C(sp^3)–C(sp^3) bonds.[24–26] Therefore, these systems can show reversible association–dissociation behaviour by switching between closed-shell molecules and unpaired electron systems. These electronic spin-state changes would result in (thermo-, mechano-, photo-)chromic behaviour caused by an external stimulus, which makes it a very attractive tool for the design

[a] *Department of Physical Chemistry, University of Málaga, Campus de Teatinos s/n, Málaga, 229071, Spain. E-mail: carmenrd@uma.es*
[b] *Key Laboratory of Synthetic and Self-assembly Chemistry for Organic Functional Materials, Shanghai Institute of Organic Chemistry, Chinese Academy of Sciences, Shanghai, 200032, China*
[c] *Department of Chemistry, University of Reading, Whiteknights, Reading RG6 6DX, UK. E-mail: f.hartl@reading.ac.uk*
[d] *Photon Science Institute and Department of Chemistry, The University of Manchester, Manchester M13 9PL, UK*
[e] *Key Laboratory for Advanced Materials, Feringa Nobel Prize Scientist Joint Research Center, Frontiers Science Center for Materiobiology and Dynamic Chemistry, Institute of Fine Chemicals, School of Chemistry and Molecular Engineering, East China University of Science and Technology, Shanghai, 200237, China. E-mail: lihongxiang@ecust.edu.cn*

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of stimuli-responsive materials. Thanks to this property, organic mono- and diradicals have emerged as essential building blocks in dynamic covalent chemistry (DCC).[27–31]

Some of us have recently demonstrated how external stimuli impact the molecular structure and supramolecular organization, and thus upon the resulting optical and electronic properties of a family of carbazole (Cz) and indolocarbazole-based (ICz) diradicals with dicyanomethylene (DCM) groups.[7,10,14,15,32] The beginning of the interest in this research was based on a Cz-based system substituted with dicyanomethylene (DCM) groups, *para*-connected (***p*-Cz-CN**).[7,10] This molecule formed long σ-bonds between the radical centres of the four different monomer units, resulting in a tetrameric cyclophane. The formation/dissociation of these long C–C bonds implies an equilibrium between the tetrameric structure and the isolated monomers in response to external conditions, giving rise to strong chromic changes (*i.e.*, from the white colour corresponding to the cyclophane to purple ascribed to the diradical monomer). The next research line was based on a similar Cz-based compound with DCM groups *meta*-substituted (***m*-Cz-CN**) which created cyclophanes structures (*i.e.*, a trimer and a tetramer).[7,10] The comparison of these two Cz-based systems (***p*-Cz-CN** and ***m*-Cz-CN**) represented a magnificent test case to study the impact of the substitution pattern on the formation of stimuli-responsive cyclophanes.[14] Later on, the influence of the elongation of the conjugated backbone was investigated through the characterization of an indolo[3-2,*b*]carbazole (ICz) compound substituted with DCM groups *para*-linked (***p*-ICz-CN**).[15] This molecule showed a dimeric cyclophane structure with two co-facial ICz units. Comparing the ***p*-ICz-CN** compound with its short-chain analogue ***p*-Cz-CN**, the extension of the conjugated chain increased the diradical character leading to stronger cyclophane structure with shorter C–C bonds and attractive π–π interactions.

Herein, a new ICz compound containing the terminal DCM groups in the *meta*-positions (***m*-ICz-CN**) was synthesized and characterized to demonstrate how its chemical reactivity and physical properties are modulated by DCM substitution when compared to its *para*-substituted homologue (***p*-ICz-CN**) (Fig. 1). This work aims to experimentally reveal the differences previously identified at the theoretical level between the ***m*-ICz-CN** and ***p*-ICz-CN** systems.[32] Interestingly, the *meta* substitution pattern implies an unusual closed-shell (CS) molecule together with the presence of a zwitterionic resonance structure which will be determinant for the molecular and optoelectronic properties.[33–36] The main objectives of this study are the investigation of (i) the connection between diradical character and the cyclophane stability, (ii) the spatial disposition of the cyclophane structures, (iii) the monomer/cyclophane interconversion both in solution and solid state in response to external stimuli and (iv) how the difference in π-conjugation and electronic communication between the DCM termini affects the electronic absorption and the redox behavior. To this end, a combined experimental and theoretical approach was developed to reach the proposed goals by linking electronic, vibrational (IR and Raman) and EPR spectroscopic and spectro-electrochemical techniques with density functional theory (DFT) calculations.

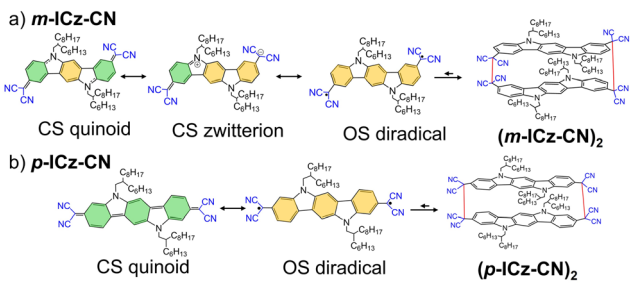

Fig. 1 Structures of (a) ***m*-ICz-CN** and (b) ***p*-ICz-CN** open-shell (OS) diradicals and their closed-shell (CS) quinoid/zwitterion resonance structures. The cyclophane dimeric arrangements of (a) ***m*-ICz-CN** and (b) ***p*-ICz-CN** are also shown.

2. Results and discussion

2.1 Synthesis and experimental characterization

The detailed synthesis of **(*m*-ICz-CN)$_2$** is shown in Scheme S1 (ESI†). Based on the presence of the zwitterionic resonance structure of the ***m*-ICz-CN** monomer, a higher reactivity compound was expected which could promote a more feasible formation of a dimeric cyclophane structure. HPLC (high performance liquid chromatography) and GPC (gel permeation chromatography) results have shown that **(*m*-ICz-CN)$_2$** has a similar retention time as **(*p*-ICz-CN)$_2$**,[15] indicating that they exhibit similar aggregated structures (Fig. S1a and b, ESI†). In addition, the calculated BET surface area (Fig. S1c, ESI†) through the nitrogen adsorption–desorption isotherm has a value of 4.7 $m^2 g^{-1}$. This is in line with the creation of a reduced cavity structure in the aggregate, pointing out the formation of a cyclophane structure where two ICz units are co-facially coupled. Variable-temperature ^{1}H NMR spectra showed that **(*m*-ICz-CN)$_2$** displayed clear hyperfine peak splitting at ambient temperature; as the temperature increased, the hyperfine splitting of the signals disappeared (Fig. S2, ESI†). The original ^{1}H NMR spectrum was recovered when the solution was cooled back to ambient temperature. These results prove the formation of cyclophane dimer **(*m*-ICz-CN)$_2$** at ambient conditions (Fig. S3, ESI†), which tends to dissociate into diradical monomers at high temperatures (Fig. S2, ESI†). Mass spectrometry is an useful analytical tool for measuring the mass-to-charge ratio (*m*/*z*) of one or more molecules present in a sample which allows the assignment of possible structures to the detected fragmentation patterns.[37] The mass spectrum of ***m*-ICz-CN** displays two main peaks (Fig. S4, ESI†); one peak is associated with the monomer structure (*m*/*z* 830) and the other one corresponds to double the molecular weight of the monomer (*m*/*z* 1660), in consonance with the co-facial coupling between the two diradical monomers giving rise to the dimeric cyclophane, **(*m*-ICz-CN)$_2$**.

2.2 DFT-calculations

2.2.1 Cyclophane structural characterization. In order to predict the most favourable intermolecular arrangement in

(*m*-ICz-CN)$_2$, two different dimer models are proposed: (i) an antiparallel orientation of the two ICz units and (ii) a parallel arrangement with co-facially superimposed units. The calculated formation free energies revealed that the antiparallel structure is the more favourable (−84.4 kcal mol^{-1} for the antiparallel disposition *vs.* −45.9 kcal mol^{-1} for the parallel structure, see Fig. S6 in ESI†), as also found for **(*p*-ICz-CN)$_2$**.[15] When comparing the aggregates with different DCM substitutions, the formation of **(*m*-ICz-CN)$_2$** is found to be more favourable than **(*p*-ICz-CN)$_2$** with the potential energy curve of the dimerization showing a slightly more pronounced well-defined minimum at ∼1.65 Å in the former (Fig. 2a). This can be explained by the presence of more favourable π–π interactions in **(*m*-ICz-CN)$_2$**, as evidenced by the stronger π–π interactions (3.64 Å *vs.* 3.74 Å) and shorter C–C σ-bonds between the monomers (1.63 Å *vs.* 1.65 Å) when compared to **(*p*-ICz-CN)$_2$**, which might confer a higher stability of the *meta*-substituted cyclophane (Fig. 2b).

2.2.2 Diradical character elucidation. As shown in Fig. 3a, focusing on the HOMO–LUMO gap (HL gap), several important points are worth noting when comparing ***p*-ICz-CN** with ***m*-ICz-CN**: (i) the substitution of the DCM groups at the *meta*-position causes a strong HOMO destabilization together with a moderate LUMO stabilization, which results in HL gap decreased by ∼1 eV (from 2.28 eV in ***p*-ICz-CN** to 1.21 eV in ***m*-ICz-CN**). (ii) Whereas the HOMO is mainly distributed over the whole π-conjugated core in ***p*-ICz-CN**, the HOMO electron density in ***m*-ICz-CN** is mainly localized at the external DCM groups; this might be ascribed to a less favourable conjugation when the DCM groups are substituted at the *meta*-position.[32]

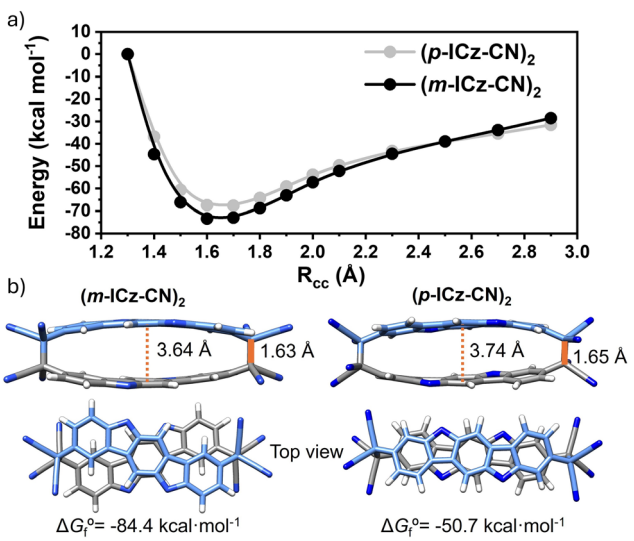

Fig. 2 (a) Potential energy curve (PES) scan of **(*p*-ICz-CN)$_2$** and **(*m*-ICz-CN)$_2$** in gas phase computed at the (U)M06-2X/6-31G** level. Here R_{CC} is the length of the C–C σ-bonds formed between two ICz units. (b) Lateral and top view of the DFT-calculated global minimum structure for **(*m*-ICz-CN)$_2$** (left) and **(*p*-ICz-CN)$_2$** (right) displaying the distances of the closest C–C σ-bonds and π–π interactions between the co-facially oriented monomers. The *N*-substituted alkyl groups have been omitted for clarity. Free energies of formation (ΔG_f°) of the **(*m*-ICz-CN)$_2$** and **(*p*-ICz-CN)$_2$** are also shown.

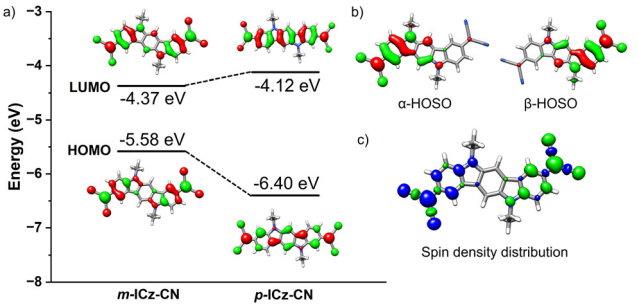

Fig. 3 (a) DFT-calculated (M06-2X/6-31G**) molecular orbital diagram of ***m*-ICz-CN** and ***p*-ICz-CN**. (b) Calculated highest-occupied single orbitals (HOSO) of α and β electrons for ***m*-ICz-CN**. (c) Spin density distribution in the diradical singlet ground state of ***m*-ICz-CN** at the (U)M06-2X/6-31G** level of theory with an isosurface value of 0.03 a.u. The blue and green surfaces represent α and β spin densities, respectively.

The investigation of the molecular orbitals of the open-shell (OS) singlet state is useful to get additional insight into the nature of the diradicals.[38,39] As shown in Fig. 3b, the highest-occupied single orbital (HOSO) of α and β electrons in the ***m*-ICz-CN** system can be classified as singlet disjoint diradicals where the two unpaired electrons reside on different parts of the molecule with almost no overlap between them (see Fig. S7, ESI,† for ***p*-ICz-CN**). Interestingly, the resulting spin density distribution in the ***m*-ICz-CN** system is located exclusively on the external parts of the conjugated backbone. Consequently, the two unpaired electrons in ***m*-ICz-CN** are less electronically coupled because of the disruption of the effective π-conjugation.[32,36] This coincides with the trend of the calculated diradical character shown below.

As seen in Fig. 4a the diradical character has been calculated using two methods (diradical character, y_0, and fractional orbital density, N^{FOD}). In all cases, the FOD plots show that the spatial distribution of the unpaired electrons is highly delocalized over the whole molecule with a strong contribution from the central carbon atoms of the DCM groups. However, the N^{FOD} values can be categorized as (i) mild for ***p*-Cz-CN**

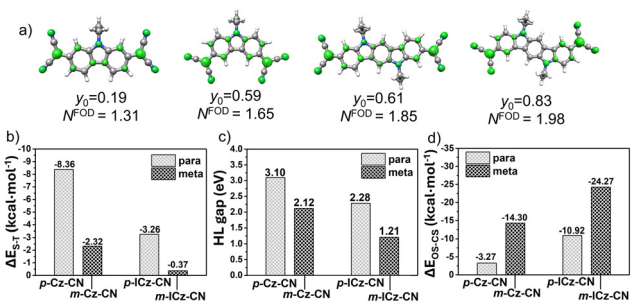

Fig. 4 (a) Isocontour plot of the FOD density (σ = 0.005 e bohr^{-3}) and predicted N^{FOD} values together with the diradical character y_0 values of four DCM-substituted diradicals (***p*-Cz-CN**,[10,14] ***m*-Cz-CN**,[7,14] ***p*-ICz-CN**[15] and ***m*-ICz-CN**). (b) Comparison of the singlet–triplet gap (ΔE_{S-T}), (c) the HL gap and (d) the energy difference between OS and CS states (ΔE_{OS-CS}) of ***p*-Cz-CN**,[10,14] ***m*-Cz-CN**,[7,14] ***p*-ICz-CN**[15] and ***m*-ICz-CN** compounds at (U)M06-2X/6-31G** level of theory.

(N^{FOD} ~ 1.3), (ii) moderate for ***m*-Cz-CN** (N^{FOD} ~ 1.6) and (iii) pronounced for ***p*-ICz-CN** and ***m*-ICz-CN** with the longest conjugated cores (N^{FOD} ~ 1.9 and 2, respectively). Thus, the diradical character can remarkably be tuned upon increasing the chain length (*i.e.*, y_0 = 0.59 in ***m*-Cz-CN** *vs.* y_0 = 0.83 in ***m*-ICz-CN**) and when going from the *para*- to *meta*-substitution (*i.e.*, y_0 = 0.61 in ***p*-ICz-CN** *vs.* y_0 = 0.83 in ***m*-ICz-CN**).

As seen in Fig. 4b–d, DFT-calculated diradical character values are in consonance with the trend predicted for the physical parameters within the series (*i.e.*, the larger diradical character, the larger ΔE_{OS-CS}, the smaller ΔE_{S-T} and smaller HL gap values). Overall, these data reveal that the extension of the molecular size and the *meta*-substitution promote the diradical state. Thus, the calculations reveal a stronger diradical character of ***m*-ICz-CN** and more reactive radical sites compared to ***p*-ICz-CN**. Note that the OS singlet ground state of ICz-based systems is supported by negative J_{ab} values (see Table S1, ESI†) revealing a lower interaction between the unpaired electrons in line with the large diradical character.

2.3 Cyclophane/monomer interconversion in solution

2.3.1 Optical properties. Aiming to explore the possibility of C–C bond cleavage of cyclophanes upon external stimuli, we monitored the optical properties of the goal samples by using UV-Vis-NIR absorption spectroscopy as a function of temperature. As seen in Fig. 5, remarkable absorption spectral changes were observed for **(*m*-ICz-CN)$_2$** in *o*-dichlorobenzene (*o*-DCB) upon heating to 410 K. A broad absorption band at ~700 nm appeared together with two shoulders at 615 and 888 nm. At the same time, the composed band around 450 nm, ascribed to the dimeric cyclophane, gradually decreased and the pale yellow solution colour gradually turned green. The colour change is fully reversed by cooling back to 300 K. The spectral profile of these new bands is characteristic of open-shell systems with unpaired electrons because of the mixing of the doubly excited electronic configuration (HOMO, HOMO → LUMO, LUMO) into the ground state.[40] This is in line with the behaviour observed for **(*p*-ICz-CN)$_2$** whose electronic absorption spectrum shows upon heating a significant increase in the intensity of the ***p*-ICz-CN** monomer band at ~700 nm.[15] The weaker absorption for ***m*-ICz-CN** monomer indicates a less efficient π-conjugation system due to its zwitterionic resonance structure.[36]

TD-DFT calculations support (i) the experimental UV-Vis-NIR spectral evolution when going from the dimeric cyclophanes to the isolated monomers (Fig. S9, ESI†), and (ii) the red-shifting of the maximum absorption band on going from the *para*- to the *meta*-substituted compound. These results are in good agreement with the HOMO–LUMO gap decreasing on going from **(*m*-ICz-CN)$_2$** to ***m*-ICz-CN** (Fig. S12, ESI†). Interestingly, the UV-Vis absorption spectrum of the brominated precursor ***m*-ICz-Br** resembles that of the aggregate **(*m*-ICz-CN)$_2$** (Fig. S9, ESI†). This is in line with the similar energy levels and topologies of the frontier molecular orbitals of ***m*-ICz-Br** and **(*m*-ICz-CN)$_2$** (Fig. S10–S12, ESI†), which in both cases exhibit an aromatic ICz backbone.

2.3.2 EPR spectroscopy. When comparing the ^{1}H NMR spectra of **(*m*-ICz-CN)$_2$** recorded in tetrachloroethane-d_2 at room temperature (RT) and high temperature up to 403 K (Fig. S2, ESI†), several proton signals broaden gradually upon heating and their hyperfine structure becomes unresolved. These observations indicate paramagnetic broadening due to the gradual development of the dominant OS diradical form of the ***m*-ICz-CN** monomer upon breaking the fairly strong C–C σ-bonds in the dimeric cyclophane. In the preceding section, the monomers ***p*-ICz-CN** and ***m*-ICz-CN** could be distinguished by electronic absorption in the visible-light spectral region reflecting differences in their specific π-conjugation.

The magnetic properties of the dissolved monomers were investigated in *o*-dichlorobenzene (Fig. 6) and toluene (Fig. S16, ESI†) by variable-temperature (VT) EPR spectroscopy in the

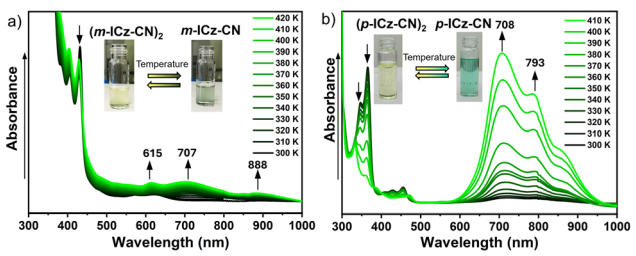

Fig. 5 (a) UV-Vis-NIR absorption of **(*m*-ICz-CN)$_2$** in *o*-DCB as a function of temperature. (b) UV-Vis-NIR spectral changes accompanying conversion of dimeric cyclophane **(*p*-ICz-CN)$_2$** to ***p*-ICz-CN** upon heating in *o*-DCB from 300 to 410 K.

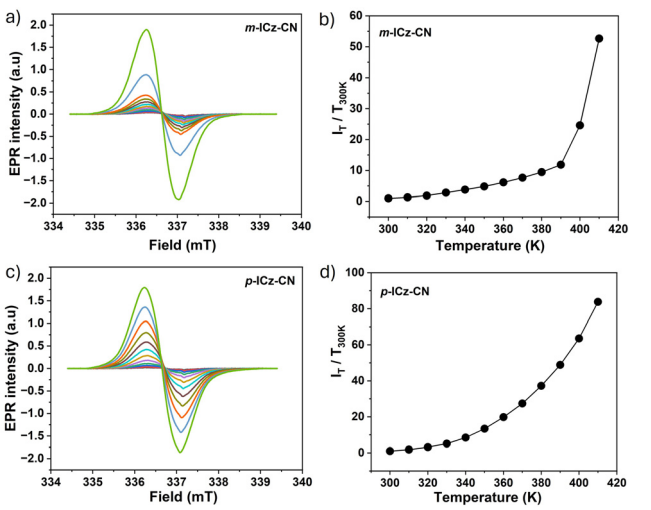

Fig. 6 (Left) Variable-temperature, cw-EPR spectra of diradical (a) ***m*-ICz-CN** and (c) ***p*-ICz-CN** generated from dimer cyclophane **(*m*-ICz-CN)$_2$** and **(*p*-ICz-CN)$_2$**, respectively, in *o*-dichlorobenzene upon gradual temperature elevation from 300 K to 410 K. Experimental conditions: microwave power 23 dB (1 mW), modulation amplitude 1 G, time constant 82 ms, conversion time ~12 ms, sweep time 30 s, receiver gain 30 dB and an average microwave frequency of 9.441 GHz. (Right) *T*-Dependent relative intensities of EPR signals of diradicals (b) ***m*-ICz-CN** and (d) ***p*-ICz-CN** measured as fluid solution in *o*-dichlorobenzene.

temperature range of 300–410 K and 300–380 K, respectively. The diradical resonance forms of both monomers exhibit in *o*-dichlorobenzene an unresolved singlet signal at g = 2.0028. The signals are very weak at ambient temperature, in line with the dominant diamagnetic dimeric cyclophane structure in the solution revealed by VT NMR and UV-Vis-NIR spectroscopies. The intensity of the EPR signal of ***p*-ICz-CN** grows much faster with increasing temperature compared to ***m*-ICz-CN** until 390 K (Fig. 6). This difference reflects the stronger C–C σ-bonds in **(*m*-ICz-CN)$_2$** (Fig. 2). Above 390 K the EPR signal of ***m*-ICz-CN** starts to grow abruptly faster. In less polar toluene, the broad EPR signal of ***p*-ICz-CN** shows a partly resolved hyperfine structure. This signal was growing in intensity with temperature less steeply compared to when heating in *o*-dichlorobenzene. For ***m*-ICz-CN** the difference became even larger and its EPR signal remained very weak up to 380 K (Fig. 6).

These observations comply qualitatively well with the outcomes of the corresponding VT UV-Vis-NIR absorption measurements. The dissolved cyclophane-type dimers **(*p*-ICz-CN)$_2$** and **(*m*-ICz-CN)$_2$** remain stable at ambient temperature. Stronger thermal activation is required for **(*m*-ICz-CN)$_2$** to break the C–C σ-bonds and the π–π interaction between the monomer chains. The fluid solution spectra of ***p*-ICz-CN** and ***m*-ICz-CN** in *o*-dichlorobenzene and in toluene were both simulated with the same isotropic g = 2.0028 within experimental error and are shown in Fig. S16 and 17 in ESI.† The spectrum of ***p*-ICz-CN** in toluene shows poorly resolved hyperfine couplings due to the unknown number of ^{14}N and ^{1}H nuclei, which are not included in the simulation to avoid over-parameterisation.

2.3.3 Redox properties. Reference ***m*-ICz-Br** undergoes in dichloromethane at ambient temperature reversible oxidation to the corresponding monocation (Fig. 7). Compared with ***p*-ICz-Br**, the oxidation potential is shifted slightly less positively (Table 1), which agrees with the higher-lying HOMO of ***m*-ICz-Br** (−6.25 eV (*meta*) *vs.* −6.37 eV (*para*)). In contrast, **(*m*-ICz-CN)$_2$** oxidizes more positively than **(*p*-ICz-CN)$_2$**, which is also reflected in their different HOMO energies (−6.93 eV (*meta*) *vs.* −6.76 (*para*)). This comparison of the HOMO energies also explains the larger difference between the oxidation potentials of **(*m*-ICz-CN)$_2$** and ***m*-ICz-Br** than reported[15] for **(*p*-ICz-CN)$_2$** and ***p*-ICz-Br** (Table 1). The lower anodic peak current of **(*m*-ICz-CN)$_2$** compared to ***m*-ICz-Br** in their equimolar solutions nicely corresponds with the diffusion coefficients determined[15] for **(*p*-ICz-CN)$_2$** (D = 4.05 × 10^{-10} m^2 s^{-1}) and ***m*-ICz-Br** (D = 1.76 × 10^{-10} m^2 s^{-1}). In contrast to stable singly oxidized [**(*p*-ICz-CN)$_2$**]$^+$ (Fig. 7 in ref. 15), the cyclic voltammogram of **(*m*-ICz-CN)$_2$** (Fig. 7 and 8a) reveals that the oxidized cyclophane undergoes rapid dissociation reflected in the absence of the corresponding reduction counterwave (in contrast to ***m*-ICz-Br**). The significantly shifted reduction wave at 0.57 V recorded on the reverse potential scan belongs to monomeric [**(*m*-ICz-CN)**]$^+$, as confirmed by cyclic voltammetry of ***m*-ICz-CN** at T = 393 K in *o*-dichlorobenzene and the corresponding UV-Vis spectroelectrochemical study in the following section. The less positive oxidation potential of ***m*-ICz-CN** compared to **(*m*-ICz-CN)$_2$** (Table 1) again complies with the calculated HOMO energies (−5.58 eV for the monomer and −6.93 V for the dimer, *cf.* Fig. S12 in ESI†).

The electrochemical reduction of **(*m*-ICz-CN)$_2$** at ambient temperature in the negative potential region is poorly defined, similar to the behaviour reported[15] for **(*p*-ICz-CN)$_2$**. The reason

Table 1 Redox properties of ***m*,*p*-ICz-Br**, isolated monomer ***m*,*p*-ICz-CN** and cyclophane-type dimer **(*m*,*p*-ICz-CN)$_2$** determined by cyclic voltammetry at a Pt microdisc electrode

Compound	E(ox.)/V	E(red.)/V	Ref.
***m*-ICz-Br**	0.38[ab]	—[c]	This work
***p*-ICz-Br**	0.45[ab]	—[c]	15
(*m*-ICz-CN)$_2$	0.84[ae]	[d]	This work
(*p*-ICz-CN)$_2$	0.63[ab]	[d]	15
***m*-ICz-CN**	0.63[e]	−0.75[ef]	This work
***p*-ICz-CN**	0.70[e]	−0.46/−0.67[beg]	15
[***m*-ICz-CN**]$^+$	—[c]	0.57[a]	This work
[***m*-ICz-CN**]$^{2-}$	−0.45[af]	—[c]	This work
	−0.61[ef]	—[c]	This work

[a] In dichloromethane/Bu$_4$NPF$_6$ at T = 298 K, E_p values *vs.* $E_{1/2}$(Fc$^+$/Fc), unless stated otherwise. [b] $E_{1/2}$ value, reversible redox reaction; ΔE_p = 80–100 mV. [c] Beyond the electrode potential limit of the electrolyte. [d] Poorly resolved broad waves. [e] In *o*-dichlorobenzene/Bu$_4$NPF$_6$ at T = 393 K. [f] Unresolved 2e$^-$ step. [g] Closely spaced 1e$^-$ steps.

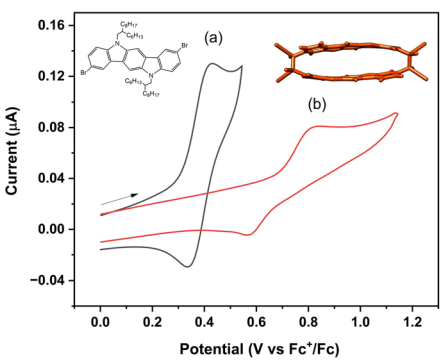

Fig. 7 Cyclic voltammogram of (a) 0.2 mM ***m*-ICz-Br** and (b) 0.2 mM **(*m*-ICz-CN)$_2$** in CH_2Cl_2/Bu_4NPF_6 at 298 K; v = 100 mV s^{-1}, Pt disk microelectrode.

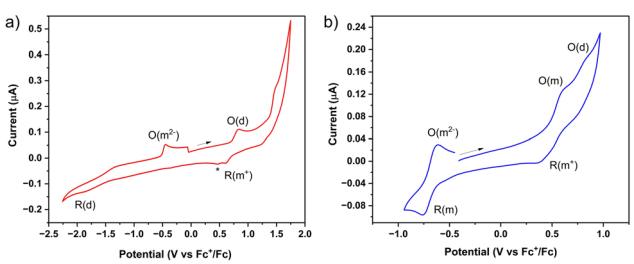

Fig. 8 Cyclic voltammograms of (a) 0.2 mM **(*m*-ICz-CN)$_2$** in CH_2Cl_2/Bu_4NPF_6 at 293 K; v = 100 mV s^{-1}, Pt disk microelectrode, and (b) *ca.* 0.2 mM ***m*-ICz-CN** in *o*-dichlorobenzene/Bu_4NPF_6 at 393 K; v = 100 mV s^{-1}, Pt disk microelectrode. Abbreviations: d = dimer **(*m*-ICz-CN)$_2$**, m = monomer ***m*-ICz-CN**, O = oxidation, R = reduction.

behind this observation is the thermal equilibrium between **(*m*-ICz-CN)$_2$** and ***m*-ICz-CN** that still exists in dichloromethane and in *o*-dichlorobenzene at ambient temperature despite being strongly shifted towards the dimer. The cyclic voltammogram of **(*m*-ICz-CN)$_2$** (Fig. 8a) shows a broad reduction wave (R(d)) around −1.9 V. Passing this wave and scanning back to the starting point leads to the appearance of a new oxidation wave at −0.45 V that is ascribed to the oxidation of the fully reduced monomer [***m*-ICz-CN**]$^{2-}$. This assignment, which complies with the reduction behaviour of **(*p*-ICz-CN)$_2$**,[15] receives support hereinafter from UV-Vis spectroelectrochemistry.

Cyclic voltammetry in *o*-dichlorobenzene at T = 393 K provides an important insight into the redox behaviour of monomer ***m*-ICz-CN** and supports the assignments made above for the cyclovoltammetric responses at ambient temperature where both the oxidation and reduction of **(*m*-ICz-CN)$_2$** triggers dissociation of the cyclophane C–C σ-bonds. Based on the EPR study under these conditions, the major component present in the electrolyte is monomer ***m*-ICz-CN**. This is also clearly reflected in the cyclic voltammogram in Fig. 8b (recorded in with a CV cell thermostatted at 393 K), which shows **(*m*-ICz-CN)$_2$** only as a minor component (the anodic wave O(d) at 0.84 V). The dominant oxidation wave O(m) at 0.63 V belongs to ***m*-ICz-CN**, being less positively shifted compared to that of ***p*-ICz-CN** under the same conditions (Table 1) due to the higher-lying HOMO of ***m*-ICz-CN** (−5.58 eV *vs.* −6.40 V calculated for ***p*-ICz-CN**). A very important new feature in the high-temperature cyclic voltammogram of dominant ***m*-ICz-CN** is its reduction wave R(m) at −0.75 V, which is negatively shifted in comparison to the reduction of ***p*-ICz-CN** (Table 1) in line with the higher-lying LUMO (−4.37 eV (*meta*) *vs.* −4.12 eV (*para*)). The scan-reversal oxidation wave O(m^{2-}) at −0.61 V belongs to [***m*-ICz-CN**]$^{2-}$ converting directly back to neutral ***m*-ICz-CN**, as confirmed by UV-Vis spectroelectrochemistry in the following section. The striking difference between the unresolved two-electron reduction of the radical centres in ***m*-ICz-CN** to [***m*-ICz-CN**]$^{2-}$ and the stepwise one-electron reduction of ***p*-ICz-CN** to [***p*-ICz-CN**]$^{-}$ and [***p*-ICz-CN**]$^{2-}$ (Table 1) can be explained by considering the very limited electronic communication between the radical halves of ***m*-ICz-CN** through the central benzene ring, visualized by the LUMO characters in Fig. 3.

2.3.4 UV-Vis-NIR spectroelectrochemistry. The one-electron oxidation of reference ***m*-ICz-Br** to corresponding [***m*-ICz-Br**]$^{+}$ in dichloromethane at ambient temperature results in UV-Vis spectral changes shown in Fig. 9. The UV-Vis absorption spectrum of [***m*-ICz-Br**]$^{+}$ (Fig. 9) has appreciably been reproduced by TD-DFT calculations (Fig. S14 (top) and Table S3, ESI†). The UV-Vis-NIR absorption spectrum of [***p*-ICz-Br**]$^{+}$ exhibits a very similar pattern,[15] with the corresponding bands being slightly shifted to shorter wavelengths.

Changes in the UV-Vis absorption spectra accompanying the oxidation and reduction of **(*m*-ICz-CN)$_2$** were recorded in parallel with the thin-layer cyclic voltammogram at a Pt mesh (Fig. S13, ESI†). The latter reproduced all key features in the conventional cyclic voltammogram of the dimer cyclophane at a Pt microelectrode (Fig. 10). The oxidation of **(*m*-ICz-CN)$_2$** may

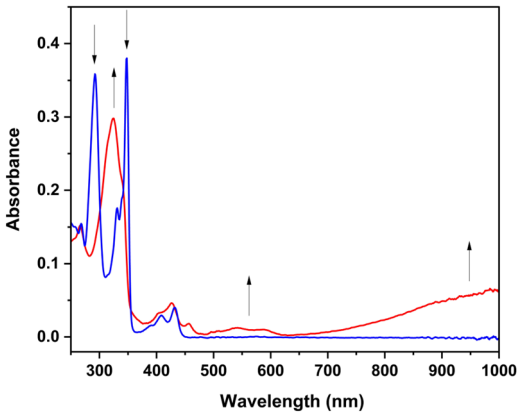

Fig. 9 UV-Vis absorption spectral changes resulting from oxidation of ***m*-ICz-Br** (blue) to [***m*-ICz-Br**]$^{+}$ (red.). Conditions: in CH_2Cl_2/Bu_4NPF_6 at T = 298 K; Pt minigrid in an OTTLE cell.

result either in stable [**(*m*-ICz-CN)$_2$**]$^{+}$ (as observed[15] for **(*p*-ICz-CN)$_2$**), or monomer radical cation [***m*-ICz-CN**]$^{+}$. The experimental spectra recorded at the oxidation wave O(d) are shown in Fig. 10a, revealing smooth conversion to a single product. TDDFT calculations (Fig. S14 (top) and Table S3, ESI†) allow its assignment to stable [***m*-ICz-CN**]$^{+}$, in particular by the presence of medium-intensity absorption between 400–500 nm, seen in the experimental spectrum as shoulders at the intense structured band at 380 nm. Only very weak absorption in this region is predicted for [**(*m*-ICz-CN)$_2$**]$^{+}$ (Fig. S13, ESI†). similar to that reported[15] for stable [**(*p*-ICz-CN)$_2$**]$^{+}$.

The electrochemical reduction of **(*m*-ICz-CN)$_2$** within the OTTLE cell at the wave R(d) is shifted to a less negative electrode potential compared to the conventional CV response (Fig. 10a). This has been ascribed[15] to a dynamic behaviour of the dimer cyclophane that co-exists with the corresponding diradical monomer, ***m*-ICz-CN**. No observable response of ***m*-ICz-CN** in the conventional CV at a less negative reduction potential (much below −1 V) complies with the solution VT EPR data and stronger C–C σ bonds in **(*m*-ICz-CN)$_2$**. UV-Vis spectral changes accompanying the single-step reduction of **(*m*-ICz-CN)$_2$** at R(d) are presented in Fig. 10b. The single product exhibits characteristic medium-intensity absorption in the visible region between 400–550 nm and two strong structured bands in the UV region. This pattern complies with TDDFT

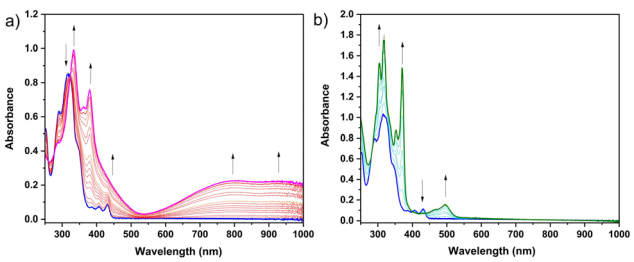

Fig. 10 UV-Vis absorption spectral changes resulting from (a) oxidation of **(*m*-ICz-CN)$_2$** (blue) to [***m*-ICz-CN**]$^{+}$ (purple), and (b) single-step reduction of **(*m*-ICz-CN)$_2$** (blue) to [***m*-ICz-CN**]$^{2-}$ (green). Conditions: in CH_2Cl_2/Bu_4NPF_6 at T = 298 K; Pt minigrid in an OTTLE cell.

data calculated for the doubly reduced monomer $[\boldsymbol{m}\text{-ICz-CN}]^{2-}$ (Table S3, middle, ESI†). The UV-Vis absorption spectrum of $[\boldsymbol{m}\text{-ICz-CN}]^{2-}$ also closely resembles that of $[\boldsymbol{p}\text{-ICz-CN}]^{2-}$ published elsewhere.[15] Reverse oxidation of $[\boldsymbol{m}\text{-ICz-CN}]^{2-}$ at O(m^{2-}) (Fig. S13, ESI†) fully restored the electronic absorption of parent $(\boldsymbol{m}\text{-ICz-CN})_2$ in a single step, alike the stepwise reoxidation of $[\boldsymbol{p}\text{-ICz-CN}]^{2-}$. As discussed above, the stronger diradical nature of ***m*-ICz-CN** not only results in the higher stability of $(\boldsymbol{m}\text{-ICz-CN})_2$ (alike the other redox states of the cyclophane dimer) but also imposes the instability of the singly reduced ('mixed-valence') radical anions $[\boldsymbol{m}\text{-ICz-CN}]^{-}$ that rapidly disproportionate to $[\boldsymbol{m}\text{-ICz-CN}]^{2-}$ and reducible parent ***m*-ICz-CN**. The likely zwitterionic resonance structure of ***m*-ICz-CN**, in contrast to the more π-conjugated quinoidal resonance form of ***p*-ICz-CN**,[15] prevents sufficient electronic communication between the radical termini in ***m*-ICz-CN** and separation of the two one-electron reduction steps, as observed[15] for the series $[\boldsymbol{m}\text{-ICz-CN}]^{n-}$ (n = 0–2).

2.4 Cyclophane/monomer interconversion in solid-state

2.4.1 IR spectroscopy. We now turn to evaluate here the C–C σ-bond dissociation of the dimeric structures in the solid state by IR spectroscopy, on applying external stimuli such as temperature and pressure. As seen in Fig. 11, the IR spectra of both systems upon grinding (as KBr pellets) display an intense ν(CN) band at ~2250 cm^{-1} associated with the cyclophane structure, which is very similar to that observed in nonconjugated nitriles (around 2270 cm^{-1} in malononitrile), and small band at ~2220 cm^{-1} ascribed to the monomer. This indicates that the application of mechanical grinding (*i.e.*, pressures below 0.1 GPa) promotes the cyclophane/monomer transformation in both systems resulting in increased conjugation of the CN groups with the central ICz core. Furthermore, increasing the temperature promotes an increase in band intensity at ~2220 cm^{-1} in the *meta*-substituted compound revealing a more pronounced transformation to the monomer, in line with the experimental EPR results shown in the following section. We also observed a similar effect in Cz-based cyclophanes[14] with the *meta*-substituted systems exhibiting a lower tendency towards the C–C bond dissociation in solution than in solid state. It is therefore unquestionable that intermolecular interactions, which are very different in solution and in the solid state, must play a fundamental role in this cyclophane/monomer transformation.

2.4.2 EPR spectroscopy. The EPR spectrum of the powder sample of $(\boldsymbol{p}\text{-ICz-CN})_2$ shows a strong, partly resolved heptet signal of monomer diradical ***p*-ICz-CN** at ambient temperature (Fig. 12). The observed hyperfine splitting is tentatively assigned to the two ^{14}N nuclei of the DCM group and two ^{1}H nuclei in the *ortho* positions at the adjacent ring. This is in line with the calculated[13] distribution of the unpaired spin density. The relatively high intensity and better resolution of the signal compared to the previously published data reflect the different mode of mechanoscopic sample preparation (filling of a small glass capillary *vs.* adhering a tape with a thin layer of the powder on the surface of a 5-mm quartz tube). When the temperature is continuously increased from 300 K, the signal intensity is nearly doubled (Fig. 12) at 410 K. On the other hand, the microcrystalline powder of yellow $(\boldsymbol{m}\text{-ICz-CN})_2$ contained at ambient temperature only a small amount of paramagnetic ***m*-ICz-CN** (Fig. 12) signal. Gradual elevation of the temperature to 410 K increased the signal intensity *ca.* 15 times (Fig. 12; relative to that of the signal observed at 300 K); however, the signal still remained of lower intensity than the signal ***p*-ICz-CN** at ambient temperature. These observations confirm the higher stability of $(\boldsymbol{m}\text{-ICz-CN})_2$ at ambient conditions also in the solid state compared to $(\boldsymbol{p}\text{-ICz-CN})_2$. At the same time, the effect of the thermal activation is stronger for solid $(\boldsymbol{m}\text{-ICz-CN})_2$, in

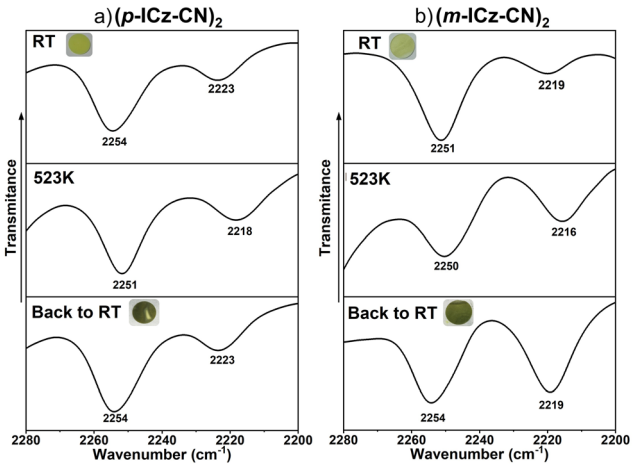

Fig. 11 IR spectrum as a yellow KBr pellet at RT (top), after heating (middle) and back to RT (down) for (a) $(\boldsymbol{p}\text{-ICz-CN})_2$ and (b) $(\boldsymbol{m}\text{-ICz-CN})_2$ compounds.

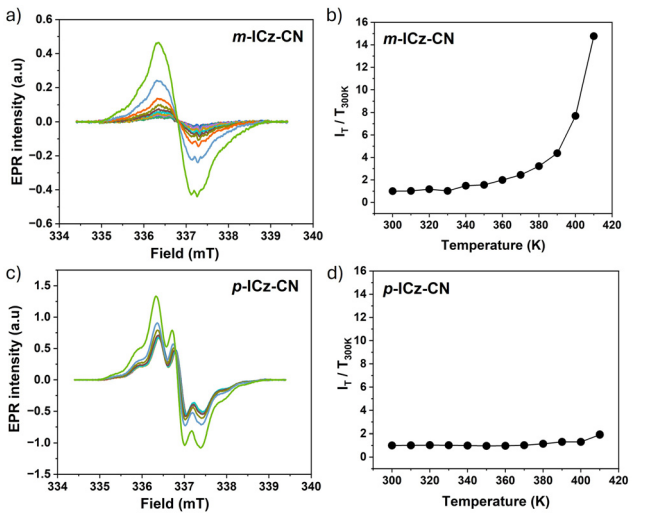

Fig. 12 Variable temperature, cw-EPR spectra of diradical (a) ***m*-ICz-CN** and (c) ***p*-ICz-CN** generated from dimer cyclophane $(\boldsymbol{m}\text{-ICz-CN})_2$ and $(\boldsymbol{p}\text{-ICz-CN})_2$, respectively, in solid state upon gradual temperature elevation from 300 K to 410 K. Experimental conditions: see Fig. 6. *T*-Dependent relative intensities of EPR signals of diradicals (b) ***m*-ICz-CN** and (d) ***p*-ICz-CN** measured as polycrystalline powder.

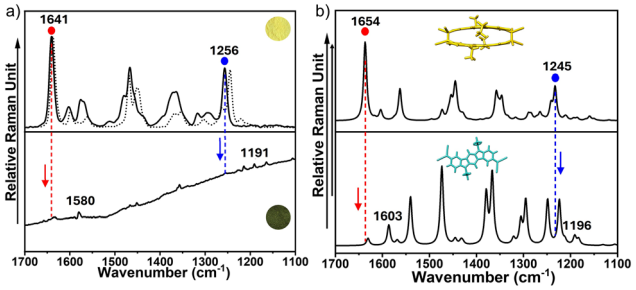

Fig. 13 (a) Top: FT-Raman spectra at λ_{exc} = 1064 nm of the yellow powder of **(*m*-ICz-CN)$_2$** (black line) and its brominated precursor ***m*-ICz-Br** (black dots). Bottom: Raman spectrum at λ_{exc} = 785 nm of isolated diradical ***m*-ICz-CN** in the solid state at 410 K. (b) Theoretical Raman spectra of **(*m*-ICz-CN)$_2$** (top) and ***m*-ICz-CN** (down) at the M06-2X/6-31G** level.

agreement with the VT IR data in the preceding section (Fig. 11). The EPR spectrum of polycrystalline ***p*-ICz-CN** is simulated using two equivalent ^{14}N and ^{1}H nuclei with the hyperfine splitting of $a_{iso}(^{14}N)$ = 12.8 and $a_{iso}(^{1}H)$ = 13.1 MHz with g = 2.0027 consistent with the unpaired spin density distribution of ***p*-ICz-CN** isomer obtained previously in DFT calculations.[15] The modelled spectrum reasonably reproduced all the features in the experimental spectrum. The heptet feature is also observed in the ***m*-ICz-CN** isomer, but very poorly resolved and not included in the simulation.

2.4.3 Raman spectroscopy. Raman spectroscopy is a sensitive approach to evaluate the open-shell diradical systems and to probe structural changes causing any mechanochromic alteration.[41–47] Fig. 13a-(top) exhibits the FT-Raman spectrum of the yellow powder of **(*m*-ICz-CN)$_2$** together with the Raman spectrum of the monomer ***m*-ICz-Br** at ambient conditions. Note that the brominated ***m*-ICz-Br** precursor exhibits an almost coincident spectrum to that of **(*m*-ICz-CN)$_2$** in line with their similar aromatic ICz structural backbones. The Raman spectrum of the diradical monomer ***m*-ICz-CN** in the solid state at 410 K is also shown in Fig. 13a-(bottom). The following spectral changes are found and compared to those seen for **(*m*-ICz-CN)$_2$**: (i) the most intense band of the dimeric structure appears at 1640 cm^{-1}, assigned to a ν(CC) mode of the ICz rings, whilst this mode downshifts to 1580 cm^{-1} in the isolated monomer, suggesting a transformation of the aromatic ICz backbone in the dimer to a CS zwitterionic resonance structure in the monomer. (ii) The band at 1252 cm^{-1} related to CC stretching vibrations and CH bending of the external phenyl rings in the dimer, is replaced by a new band at ~1190 cm^{-1} that is attributed to the CH bending of the central phenylene ring in the monomer. This experimental spectral evolution is in excellent agreement with the DFT-calculated Raman spectra on going from the dimer to the monomer (Fig. 13b).

Interestingly, in the range between 2300 cm^{-1} and 2180 cm^{-1} (related to the CN stretching modes), the band at 2250 cm^{-1} undergoes a downshift to 2205 cm^{-1} upon heating, corresponding to the formation of the isolated diradical ***m*-ICz-CN** (Fig. S15, ESI†); this is in line with the IR spectroscopic data. Overall, the Raman spectral evolution reveals that the cyclophane/monomer transformation involves a loss of aromaticity of the ICz units and the appearance of a more conjugated scenario of the CN groups.

3. Conclusions

In this work, the different structural and electronic parameters that control the cyclophane formation from ICz-based diradical systems are investigated in depth. Subtle modifications of the chemical structure (*i.e.*, DCM substitution position, core elongation) are found to significantly affect the dynamic covalent chemistry behaviour. Theoretical data reveal that: (i) the elongation of the conjugated core and the *meta*-substitution provoke less interaction between the unpaired electrons resulting in a large diradical character, and (ii) *meta*-substituted ICz-based systems can form more stable cyclophanes when compared to their *para*-substituted analogues. All of these factors are shown to influence the cyclophane/monomer transformation. Interestingly, the dynamic ICz-based cyclophane/monomer transformation is accompanied by a colour change from light yellow to deep green, although a less reversible dissociation is found in the ICz-based diradicals when compared to the shorter Cz-based analogues. On the other hand, the effect of the linkage position is key to understand how the external stimuli affect the C–C bond dissociation. The *meta* substitution pattern favours the presence of a zwitterionic resonance structure in ***m*-ICz-CN**, resulting in a less favoured dissociation towards the monomer in solution compared to the *para*-substituted system. However, a more feasible C–C bond dissociation is found in the solid state when comparing *meta vs. para*, since a more restricted conformational space is created in solid state and intermolecular interactions might play a significant role. This is in line with a recent theoretical study that reveals that resonance structure and the diradical character can be modulated even in regions dominated by the dispersion forces.[48] This cyclophane/monomer conversion has successfully been followed by NMR, vibrational spectroscopy and EPR techniques. In addition, the redox properties of the isomeric cyclophane dimers (at ambient temperature) and diradical monomers (at elevated temperature) confirm the differences in electronic structures and frontier orbital energies. The dissociation of the C–C σ-bonds in **(*m*-ICz-CN)$_2$** upon both oxidation and reduction was also revealed by UV-Vis spectroelectrochemistry. To summarize, this comprehensive study helps to evaluate the stimuli-responsive properties of diradical-based systems, providing new strategies for the design of novel dynamic covalent chemistry materials.

4. Methodology

4.1 Experimental details

Commercial reagents were used without further purification and anhydrous solvents were purified by standard distillation methods prior to use. NMR spectra were recorded on an Agilent 400 spectrometer. Mass spectra were recorded on a Orbitrap

Q-Exactive high-resolution mass spectrometer and atmospheric-pressure ionization system using heated electrospray H-ESI (Thermo Scientific). Elemental analyses were performed on an Elementar Vario EL III elemental analyzer. Gel permeation chromatography results were recorded on a PolymerLaboratories GPC 220 instrument, using standard polystyrene (PS) as the reference. High-performance liquid chromatography measurements were carried out on an Agilent ZORBAX Eclipse XDB-C18 apparatus. Nitrogen adsorption–desorption isotherm was measured at 77.3 K using a micromeritics ASAP2020 HD88 surface area and porosity analyzer. The samples were degassed at 393 K for 12 h under vacuum (10^{-5} bar) before analysis. The surface areas were calculated in the relative pressure (P/P_0) range from 0.05 to 0.30.

The characteristics of the analyses performed by high-performance liquid chromatography (HPLC) are as follows. Chromatographic column: cortex C18, 2.7 μm, 4.6 × 100 mm mobile phase: water (A), acetonitrile (B) gradient elution: 0 min 50% B, 10 min 100% B, 15 min 100% B. Flow rate: 0.6 mL min^{-1} temperature: ambient temperature. Sample preparation: we take an appropriate amount of the sample, dissolve it in 200 μL of chloroform, and add 800 μL of acetonitrile. We filter the organic phase through a membrane filter and proceed to HPLC analysis.

UV-Vis-NIR spectra were measured on an Agilent 8453 diode-array spectrophotometer between 190–1100 nm, with a resolution of 1 nm. A Varian Cary 5000 spectrophotometer was used in the UV-Vis-NIR range (170–3300 nm), with a PbSmart detector and 0.4 nm resolution. Variable-temperature UV-Vis-NIR absorption spectra were recorded using an optical cryostat Optistat DN (Oxford Instruments). IR spectra were carried out in the spectral range 4000–400 cm^{-1} using a Bruker Vertex 70 FT-IR spectrometer, which consists of a KBr beam splitter and a DLaGTS detector. A spectral resolution of 4 cm^{-1} and 64 scans were used to record the spectra. The compound was studied dispersed in KBr pellets (transmission). For variable-temperature FT-IR measurements, the Specac P/N 2100 accessory was used, which allows the recording of both solid and liquid samples. It has a Graseby Specac automatic temperature control system and can work in the range of 103 to 523 K.

Raman measurements were performed in an FT-Raman Bruker spectrometer which makes use of a continuous-wave Nd-YAG laser working at 1064 nm for excitation and a germanium detector operating at liquid nitrogen temperature. Raman radiation performs the detection in a back-scattering geometry, forming an angle of 180° between the incident radiation and the scattered radiation with a standard spectral resolution of 4 cm^{-1}. Variable-temperature Raman spectra were registered by using a Bruker Senterra dispersive Raman microscope which makes use of two integrated lasers, with a wavelength of 785 nm. The detection system is a charge couple device (CCD) that is thermoelectrically cooled to 208 K. It also consists of a Neon lamp, used for laser calibration. For thermo-spectroscopic measurements, the Linkam FTIR600 accessory was used, which allows to access a thermal range between 93 and 873 K. Cyclic voltammetry (CV) was performed under an inert argon atmosphere with a heart-shaped single-compartment glass cell containing a coiled Ag wire pseudo-reference electrode, a 0.4-mm diameter Pt microdisc working electrode, and a coiled Pt wire counter electrode. The cell was held inside an earthed Faraday cage and connected to a Metrohm-Autolab PGSTAT302N potentiostat. The internal standard ferrocenium/ferrocene (Fc^+/Fc) was added prior to the final scans. The sample solutions were 0.2 mM in analyte dissolved in dichloromethane containing 10^{-1} M tetrabutylammonium hexafluorophosphate (TBAH). UV-Vis spectroelectrochemistry was performed using an optically transparent thin-layer electrochemical (OTTLE) cell[49] equipped with a Pt mini-grid working electrode. The sample solution contained 2 × 10^{-1} M TBAH as the supporting electrolyte and 1 mM redox active compounds in dichloromethane. The course of the spectroelectrochemical experiment was monitored by thin-layer cyclic voltammetry (v = 2 mV s^{-1}), with a potential control by a PalmSens EmStat3 potentiostat operated with the PSTrace5 software. UV-Vis spectra were recorded with a Scinco S-3100 diode-array spectrophotometer covering 200–1100 nm.

All EPR samples were prepared either as a polycrystalline powder or as fluid solutions in *o*-dichlorobenzene/toluene solvents. Microwave power saturation experiments were performed with reference ***p*-ICz-CN** (a solid-state sample) and TEMPO (2,2,6,6-tetramethyl-1-piperidinyloxy) in toluene (data not shown). From the analysis, non-saturating experimental conditions were chosen for all EPR experiments, which are: microwave power of 23 dB (1.0 mW) and modulation amplitude (MA) of 1 G for both fluid solutions and polycrystalline powders at room/high temperature. For quantitative/qualitative analysis of the polycrystalline powders, both ***p*-ICz-CN** (*ca.* 0.34 mg) and ***m*-ICz-CN** (*ca.* 0.4 mg) were loaded into a 1.3-mm (OD)/1-mm (ID) glass-capillary tube at *ca.* 21-mm height. This glass capillary tube was inserted into a 4.0-mm (OD)/3-mm (ID) quartz EPR tube (Wilmad LabGlass) for subsequent measurements. The fluid solutions in *o*-dichlorobenzene contained *ca.* 0.4 mg of ***p*-ICz-CN** in 20 μL and *ca.* 1.1 mg of ***m*-ICz-CN** in 50 μL. The fluid solutions in toluene contained an undetermined amount of ***p*-ICz-CN** soluble in 70 μL and *ca.* 0.82 mg of ***m*-ICz-CN** in 40 μL. Also, the solutions were loaded into a 1.3-mm (OD)/1-mm (ID) glass-capillary tube at *ca.* 21-mm height. The optimum position of the samples (both powders and solutions) inside the sample cavity (resonator) and the temperature at the sample were calibrated before the actual EPR measurements. Note that in addition to the diradical resonance forms of both monomers which exhibit in *o*-dichlorobenzene an unresolved singlet signal at g = 2.0028, an extremely weak, sharp radical signal is also observed at 3371 G with g = 2.0005 and its contribution to the overall spectrum is estimated to be less than one percent (see Fig. S17, ESI†). This weak sharp EPR signal is likely to belong to a trace paramagnetic impurity in the quartz tube(s) used for the measurements.

The RT and elevated temperatures X-band measurements were achieved on an EMXmicro EPR spectrometer using the Bruker ER414VT variable temperature accessory, connected to an ER 4112SHQ X-band resonator equipped with cooling side

plates (ER4112SHQE-1004). The optimum conditions used for recording the spectra are: microwave power 23 dB (1 mW), modulation amplitude 1 G, time constant 82 ms, conversion time ~12 ms, sweep time 30 s, receiver gain 30 dB and an average microwave frequency of 9.441 GHz. All experimental parameters were kept constant, except sample concentration and/or weight. All spectra were multi-scanned and averaged (*ca.* 30 scans) to obtain good S/N ratio. The analysis of the continuous wave EPR spectra and simulations were performed using EasySpin toolbox (5.2.36) for the Matlab program package.[50] The observed polycrystalline powder and fluid solutions spectra looks similar to that of the previously reported open-shell monomer, ***p*-ICz-CN** of EPR-silent cyclophane dimer, **(*p*-ICz-CN)$_2$**.[15]

4.2 Theoretical details

All geometry optimizations were carried out in the gas-phase, using the Gaussian 16 package[51] based on the density functional theory (DFT). In all calculations, the long alkyl *N*-substituents were replaced by ethyl groups. Herein, we have made use of the exchange–correlation functional M06-2X[52] and the 6-31G (d,p)[53,54] basis set. This functional has been parameterized so that the electronic correlation for medium-range distances (5 Å or less) is enough to describe the non-covalent interactions within many complexes.[55] This fact allows that the functional M06-2X has widely been used to solve problems of non-covalent interactions, and so far, it has been recognized as one of the best in the description of the interactions of π–π stacking and the study of molecular aggregates (dimeric species).[56] The broken-symmetry solution was simulated for OS singlet species and triplet state form using the keyword guess = mix and unrestricted wavefunction. To evaluate the importance of the solvent on the pancake dimer structural properties, the geometries optimized were obtained in toluene, chloroform and *o*-DCB by using the PCM method[57] (see Table S2, ESI†). The effective electron exchange interaction (J_{ab}) which reflects the overlap integral between the two nearly energetically degenerate molecular orbitals was calculated by using eqn (1):

$$J_{ab} = \frac{(E_{OS} - E_T)}{\langle S^2\rangle_T - \langle S^2\rangle_{OS}} \quad (1)$$

where E_{OS} and E_T correspond to the energies of the OS singlet and triplet state, with spin contamination values $\langle S^2\rangle_T$ and $\langle S^2\rangle_{OS}$ of the triplet and singlet state, respectively.[58–61]

The spin correction for ΔE_{S-T} was computed by eqn (2).

$$\Delta E_{S-T} = (E_{OS} - E_T)\frac{\langle S^2\rangle_T}{\langle S^2\rangle_T - \langle S^2\rangle_{OS}} \quad (2)$$

In addition, the diradical-character calculations were measured using two different ways. First, the diradical character (y_0) is defined as twice the weight of the double excitation configuration in a multiconfigurational calculation.[62] Nevertheless, although multi-reference methods should preferably be employed for addressing diradical systems calculations,[63–65] Nakano and others demonstrated that spin-unrestricted mono-reference methods, such as DFT, can also reproduce well the diradical character calculated by multiconfigurational methods.[66–69] Therefore, the diradical character y_0 (the closed-shell structure has y_0 = 0 and the pure diradical state has y_0 = 1) was calculated from the occupation numbers of the highest occupied natural orbital (HONO) and the lowest unoccupied natural orbital (LUNO) using the spin-projected formalism.[70] This parameter is obtained from eqn (3),

$$y_0 = 1 - \frac{2T_i}{1 + T_i^2} \quad (3)$$

where

$$T_i = \frac{n_{HONO-i} - n_{LUNO+i}}{2} \quad (4)$$

and $n_{HONO-i,LUNO+i}$ are the occupation numbers of the occupied and unoccupied natural orbitals (NO). The diradical character corresponds to i = 0.

Second, the fractional orbital density (FOD) method incorporating a strong correlation effect has also been used to calculate the N^{FOD} values by the ORCA 4.0 program.[71] An analysis based on FOD was performed, whose main characteristics is to provide the localization of ''hot'' electrons (strongly correlated and chemically active) in a molecule.[72,73] This method gives a quantitative description of the OS diradical form in the ground state. We used the FT-TPSS[74]/def2-TZVP[75] level at a default electronic temperature (T_{el}) of 5000 K in order to calculate the number of ''hot'' electrons, N_{FOD}. The isocontour value of the ρ^{FOD} plots was fixed to 0.005 e bohr^{-3}.

Author contributions

V. H. J., I. B. D. and M. C. R. D: DFT-calculations together with UV-Vis absorption and Raman/IR spectra at RT and *T*-dependent. R. N. and F. H.: redox properties and spectroelectrochemistry experiments. D. C. and M. S.: *T*-dependent EPR experiments. H. L. and D. W.: synthesis and characterization of the sample. All authors have read and agreed to the published version of the manuscript.

Data availability

The data that support the findings of this study are available within the article and the ESI.† Details about methodology, experimental procedures, characterization data, theoretical calculations, EPR studies and NMR spectra are included.

Conflicts of interest

The authors declare no conflict of interest.

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