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Te⋯I secondary-bonding interactions in crystals containing tellurium(II), tellurium(IV) and iodide atoms: supramolecular aggregation patterns, nature of the non-covalent interactions and energy considerations†

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This study delves into the intriguing realm of chalcogen-bonding interactions, specifically focusing on tellurium(II) and/or tellurium(IV) interactions with iodide species within six crystals. Utilising comprehensive computational chemistry calculations and underpinned by crystallographic data retrieved from the Cambridge Structural Database, this research elucidates the supramolecular aggregation, bonding nature and interaction energetics of Te⋯I non-covalent bonds in a series of crystals containing each of tellurium(II), tellurium(IV) and iodide. The investigation encompasses a variety of molecular assemblies, ranging from zero-dimensional to complex three-dimensional architectures. Tellurium(II) atoms formed Te⋯I interactions in six of the crystals but tellurium(IV) participated in Te⋯I interactions in only three crystals despite there being equal numbers of tellurium(II) and tellurium(IV) atoms in the zero- and one-dimensional assemblies. Several computational tools, including Molecular Electrostatic Potential (MEP) surfaces, Quantum Theory of Atoms in Molecules (QTAIM) and Natural Bond Orbital (NBO) analyses, provide a nuanced understanding of the electron density topology and charge transfer mechanisms. In some assemblies, where the Te⋯I interactions are ~ 3.2 Å, the interaction energies are very large, *i.e.* 25.6 to 38.8 kcal mol^{−1}, suggesting partial covalent character, an observation corroborated by the NBO analysis; longer Te⋯I separations correlate with reduced energies, *i.e.* 3.1 to 7.0 kcal mol^{−1}. Crucially, it is noteworthy that the Te⋯I interactions identified in this study are distinctly characterised as chalcogen bonds, rather than halogen bonds.

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1. Introduction

In molecular crystals, tellurium has long been known to expand its coordination sphere to accommodate proximate atoms at separations longer than the sum of the covalent radii but shorter than the sum of the van der Waals radii; usually these intermolecular contacts introduce distortions in the central atom's coordination geometry.^{1–9} The term “secondary-bonding” emerged as the all-encompassing description of these types of interactions operating in crystals of the main group elements.^{10–12} While the atoms forming intermolecular interactions with tellurium are usually the lighter chalcogens, *i.e.* oxygen and sulphur, and halides, more recently, tellurium(lone-pair)⋯ π (aryl) interactions have been

recognised as important in providing stabilising interactions in crystals where such interactions can form.^{13–15} There are three readily accessed oxidation states for tellurium, namely +II, +IV and +VI, having two, one and zero stereochemically active lone-pair of electrons, respectively. The interaction of electron-rich tellurium(II)/(IV) species with electron-rich donors, such as sulphur and iodide, is often rationalised in terms of the σ -hole concept.^{16–19} This explanation for the interaction between these electron-rich species relates to an electron-deficient region at the extension of a covalent bond capable of interacting with nucleophiles such as lone-pairs of electrons and π -systems. The electrophile is normally attached to an electron-withdrawing group and increasing polarizability of the bond correlates with a greater σ -hole – thus, of the group 16 elements, the σ -hole will be at a maximum for tellurium.

When the electrophile is a chalcogen atom, the σ -hole model is given the specific term chalcogen-bonding²⁰ with these non-covalent connections capable of providing structure-directing interactions in crystals.²¹ However,

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systematic reviews of crystals featuring $\text{Se}\cdots\text{O}$,²² $\text{Se}\cdots\text{N}$,^{23,24} $\text{Se}\cdots\pi(\text{aryl})$,²⁵ $\text{Te}\cdots\text{N}$,⁹ $\text{Te}\cdots\text{I}$ (ref. 26) and $\text{Te}\cdots\pi(\text{aryl})$ ¹⁴ interactions confirm these contacts do not always arise from the interaction of a chalcogen-bound σ -hole with a nucleophile. Indeed, the situation is rather more complicated, especially in systems where an interacting atom forms multiple secondary-bonding interactions. Other interactions are sometimes observed, such as π -hole interactions, where the electron deficient region is located perpendicular to a framework sustained by covalent bonds.^{27–29} In crystals of chalcogen-containing molecules, π -hole interactions can cooperate with σ -hole interactions^{30–32} or even operate independently.^{33,34} Instances of bifurcated^{35–37} and charge-assisted^{38,39} chalcogen bonds are also known. Computational chemistry calculations on the aforementioned neutral systems^{31–37} and other non-charged systems^{26,40} confirm that the energy imparted by the different interactions involving the tellurium are like or exceed those imparted by conventional hydrogen bonding.⁴¹

While the focus of the present study is upon the crystals, secondary-bonding interactions are known to persist in solution,^{42,43} with recent studies highlighting the presence of intramolecular $\text{Te}\cdots\text{F}$ (ref. 44) contacts in solution, and the gas-phase $(\text{Te}\cdots\text{N})$.⁴⁵ With such persistence, it is not surprising that chalcogen-bonding finds applications⁴⁶ beyond the solid-state⁴⁷ such as in coordination chemistry,⁴⁸ catalysis,^{49,50} molecular recognition⁵¹ and materials science.⁵² While less numerous, chalcogen-bonding has many parallels with the ubiquitous halogen-bonding.⁵³

The importance of strong and directional halogen-bonding in many spheres of endeavour⁵⁴ have been summarised in many reviews focussing upon catalysis,^{55–57} liquid crystals,⁵⁸ polymer science,⁵⁹ phosphorescent materials,⁶⁰ molecular recognition,⁶¹ *etc.* in addition to crystal engineering considerations.^{62–64}

Of the halogens, the largest, iodide, is expected have the greatest σ -hole.¹⁶ This was demonstrated in a bibliographic review of $\text{X}\cdots\pi(\text{aryl})$, for $\text{X} = \text{I}, \text{Br}, \text{Cl}$ and F , interactions in crystals.⁶⁵ Thus, delocalised⁶⁶ $\text{I}\cdots\pi(\text{aryl})$ interactions in crystals did not necessarily persist in crystals of the lighter congeners. However, other factors must come into play as demonstrated in a survey of competing $\text{I}\cdots\text{Br}$ interactions,⁶⁷ which saw a 1:2 ratio of type I to type II^{68,69} $\text{I}\cdots\text{Br}$ interactions, showing the σ -hole on iodide is not all pervasive and subject to other considerations in crystals. In the present work, a survey and computational chemistry investigation of $\text{Te}\cdots\text{I}$ interactions formed in heterovalent^{70,71} crystals containing both tellurium(II) and tellurium(IV) as well as iodide is described. The aim of the work is to determine the propensity of formation, competition, *i.e.* $\text{Te}^{\text{II}}\cdots\text{I}$ vs. $\text{Te}^{\text{IV}}\cdots\text{I}$, nature and energies of stabilisation provided by the identified $\text{Te}\cdots\text{I}$ interactions.

2. Methods

The data in the Cambridge Structural Database (CSD; version 5.44, April 2023)⁷² was searched using ConQuest

(version 2023.1.0).⁷³ The analysis sought all crystal structures having an intermolecular $\text{Te}\cdots\text{I}$ interaction less than or equal to the sum of the van der Waals radii of tellurium (2.06 Å) and iodide (1.98 Å), *i.e.* 4.04 Å,⁷⁴ in crystals containing tellurium and iodide. Additional criteria were applied in that only molecular single-crystal structures with three-dimensional coordinates and $R \leq 0.075$ were sought and having no errors nor disorder and were charge neutral. Manual inspection of the 269 “hits” retrieved nine mixed tellurium(II)/tellurium(IV) structures and six of these had $\text{Te}\cdots\text{I}$ interactions between non-charged species and operating independently of alternate, directional intermolecular interactions. The analysis of the six crystals was conducted with PLATON⁷⁵ and DIAMOND.⁷⁶ All crystallographic diagrams in section 3.1 and ESI† Fig. S1–S5 are original, being generated using DIAMOND.⁷⁶

The computational chemistry calculations were carried out with the Turbomole 7.7 program,⁷⁷ using the crystallographic coordinates available in the respective CIFs as input for the evaluation of the supramolecular assemblies. The positions of the hydrogen atoms were optimised in dimers where hydrogen bonds are established. The level of theory used for the calculations was PBE0 (ref. 78)-D4 (ref. 79)/def2-TZVP.^{80,81} For heavier elements, such as tellurium and iodide, this basis set includes effective core potentials (ECPs) and takes into consideration relativistic effects for the inner electrons.⁸¹ This combination of method and basis set has been previously used by us⁸² and others⁸³ to analyse σ -hole interactions, including chalcogen bonds (ChBs) involving tellurium.²⁶ The MEP surface plots were generated using the wavefunction obtained at the same level of theory and 0.001 a.u. isosurfaces. The topological analysis of the electron density was carried out according to the quantum theory of atoms in molecules (QTAIM) method proposed by Bader⁸⁴ and represented using the VMD program.⁸⁵ The NBO analysis⁸⁶ was performed using the same level of theory and the NBO7.0 program.⁸⁷ The NBOs and charge density plots were represented using the VMD software.⁸⁵

3. Results and discussion

The chemical diagrams for the six systems investigated herein by a range of theoretical procedures are presented in Fig. 1. Based on the standard van der Waals radii assumed in PLATON,⁷⁵ two, three and one zero-, one- and three-dimensional aggregation patterns, respectively, featuring $\text{Te}\cdots\text{I}$ interactions functioning independently of alternative non-covalent interactions, such as conventional hydrogen-bonding and $\text{I}\cdots\text{I}$ halogen-bonding, have been identified.

3.1 Supramolecular aggregation

Table 1 presents a summary of the coordination number (CN) and donor sets for the tellurium atoms participating in $\text{Te}\cdots\text{I}$ interactions as well as salient geometric parameters characterising these.



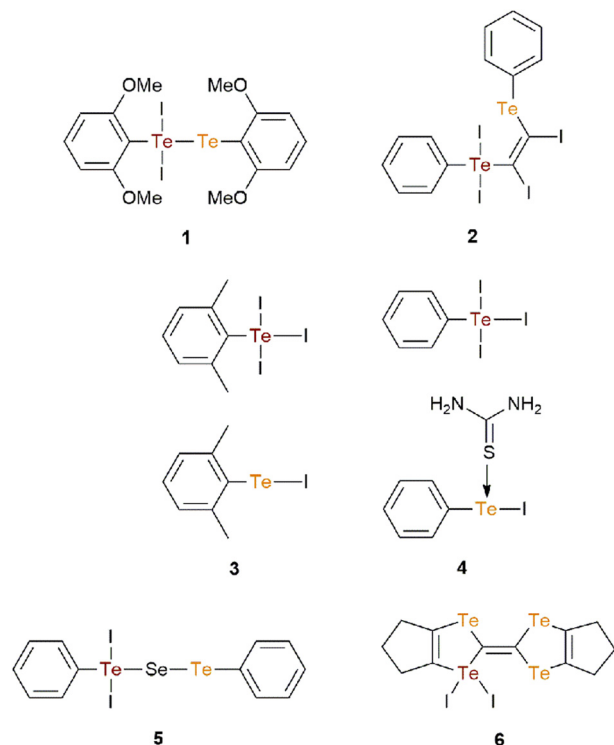


Fig. 1 Chemical diagrams for the mixed tellurium(II)/tellurium(IV) iodide species investigated herein, 1–6. Compounds 3 and 4 are 1:1 co-crystals with the cofomers being connected by Te⋯I interactions in their respective crystals.

The first and simplest aggregate to be described is formed in the crystal of ditellane, $L^1\text{Te}^{\text{II}}\text{Te}^{\text{IV}}\text{I}_2L^1$ for $L^1 = 2,6\text{-dimethoxyphenyl}$ (1).⁸⁸ Here, the tellurium(II) atom forms a Te⋯I interaction with one of the tellurium(IV)-bound iodide atoms over a centre of inversion to form a dimer. As shown in ESI† Fig. S1(a), the aggregation results in the formation of a flat, six-membered $\{\cdots\text{Te}_2\text{I}\}_2$ synthon. Within the aggregate, intramolecular Te⋯O interactions are noted, with the shortest Te⋯O separation being 2.905(4) Å. In the crystal, the dimeric aggregates assemble into a chain *via* relatively long I⋯I halogen-bonding interactions; see ESI† Fig. S1(b). A very

similar mode of aggregation as for 1, is found in the crystal of $\text{PhTe}^{\text{II}}\text{C}(\text{I})=\text{C}(\text{I})\text{Te}^{\text{IV}}\text{PhI}_2$ (2),⁸⁹ but results in a twisted, centrosymmetric 10-membered synthon $\{\cdots\text{Te}_2\text{TeI}\}_2$, ESI† Fig. S2(a). An intramolecular Te⋯Te interaction [3.4205(4) Å] is noted. The aggregates are connected into a three-dimensional architecture *via* I⋯I halogen-bonding; see ESI† Fig. S2(b) for details.

There are three examples of one-dimensional aggregation patterns. The first of these, formulated as $(2,6\text{-dimethoxyphenyl})\text{ITe}^{\text{II}}(\mu\text{-I})\text{Te}^{\text{IV}}(2,6\text{-dimethoxyphenyl})\text{I}_2$,⁹⁰ 3, resembles those of 1 and 2 in that it is only the tellurium(II) atom that forms Te⋯I interactions within the aggregation pattern with the notable difference being that the tellurium(II) atom engages in two Te⋯I interactions, Fig. 2(a), with one separation being, (i), being significantly shorter than the second, (ii). The tellurium(II) atom is tri-coordinated within a Cl_2 donor set as one of the tellurium(IV)-bound iodide atoms links the two tellurium atoms; the bridge is close to symmetric with the Te^{IV}–I bond length [2.9130(3) Å] being marginally shorter than the Te^{II}–I bond [3.0267(3) Å]; the tellurium(IV) atom is coordinated within a C_3I donor set. The molecules assemble through centrosymmetric, four-membered $\{\cdots\text{Te}\cdots\text{I}\}_2$ synthons which are flat and are linked *via* larger, eight-membered $\{\cdots\text{TeTeI}\}_2$ synthons. These latter synthons are also centrosymmetric and have the form of an elongated chair with the iodide atoms engaged in the Te⋯I interactions lying above and below the plane defined by the remaining atoms. Only one of the iodide atoms participates in the formation of Te⋯I interactions. The Te⋯I interactions feature within a supramolecular ribbon, Fig. 2(a). The presence of I⋯I halogen-bonding in the crystal of 3 lead to the formation of two-dimensional arrays; see ESI† Fig. S3(b).

The molecule in the 1:1 co-crystal 4 has the formula $\text{PhTe}^{\text{II}}\text{I}[\text{S}=\text{C}(\text{NH}_2)_2]\cdot\text{PhTe}^{\text{IV}}\text{I}_3$.⁹¹ The tellurium(II) centre is tri-coordinated by the *ipso*-carbon atom of a phenyl group, an iodide and a loosely bound thione-sulphur atom of a thiourea molecule, while the tellurium(IV) atom is tetra-coordinated by the *ipso*-carbon atom of a phenyl group and three iodide atoms. Each of the independent tellurium(II) and tellurium(IV) atom participates in the formation of Te⋯I interactions

Table 1 Listing of coordination numbers (CN), donor sets, Te⋯I separations and selected angles for 1–5

Aggregate/interaction ^a	Te ^{II} – CN/donor set	Te ^{IV} – CN/donor set	<i>d</i> (Te⋯I) (Å)	Y–Te⋯I (°)
1/(i)/Te ^{II}	2/Te ^{IV} C	4/Te ^{II} I ₂ C	4.0169(5)	Y = Te ^{IV} , 157.377(15)
2/(i)/Te ^{II}	2/C ₂	4/I ₂ C ₂	3.7708(4)	Y = ethylene-C, 171.97(11)
3/(i)/Te ^{II}	2/Cl ₂	4/Cl ₃	3.2175(4)	Y = I, 173.288(11)
3/(ii)/Te ^{II}	3/Cl ₂	4/C ₃ I	3.9902(4)	Y = Te ^{IV} , 153.679(10)
4/(i)/Te ^{IV}	3/Cl ₂	4/Cl ₃	3.2651(7)	Y = I, 174.501(19)
4/(ii)/Te ^{IV}			3.7601(14)	Y = phenyl-C, 174.08(16)
4/(iii)/Te ^{II}			3.5063(8)	Y = Te ^{IV} , 170.20(2)
4/(iv)/Te ^{II}			3.9034(14)	Y = phenyl-C, 157.14(16)
5/(i)/Te ^{IV}	2/CSe	4/Cl ₂ Se	3.9424(10)	Y = Se, 171.40(3)
5/(ii)/Te ^{IV}			3.9712(10)	Y = phenyl-C, 170.3(2)
5/(iii)/Te ^{II}			3.9925(10)	Y = phenyl-C, 166.4(2)
5 ^b /(iv)/Te ^{II}			3.8910(10)	Y = Te ^{IV} , 73.40(2); Y = Se, 93.36(3); Y = phenyl-C, 92.9(2)

^a The interaction is as in Fig. 2. ^b This interaction is intramolecular.



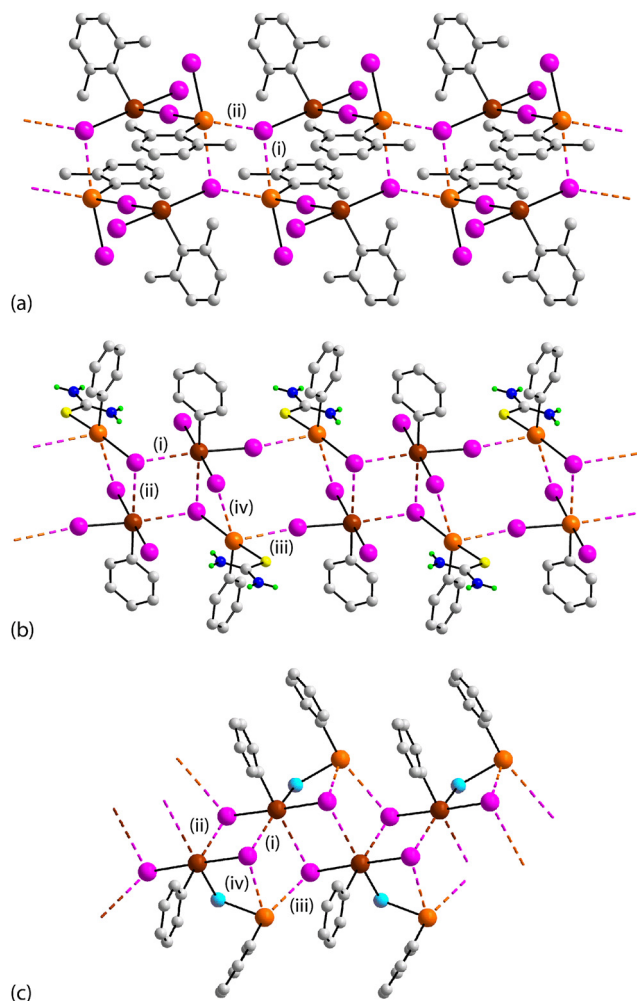


Fig. 2 Views of the supramolecular ribbons featuring Te...I interactions in the crystals of (a) **3**, (b) **4** and (c) **5**.

leading to three distinct synthons. The tellurium(II)-bound iodide forms two Te...I interactions while two of the tellurium(IV)-bound iodide atoms form Te...I interactions. Referring to Fig. 2(b), centrosymmetric, four-membered $\{\cdots\text{Te}^{\text{IV}}\cdots\text{I}\}_2$ synthons, resembling those seen in **3**, involve tellurium(IV) atoms exclusively. Here, each of the tellurium(IV) and tellurium(II)-bound iodide atoms forms two Te...I interactions. Larger, centrosymmetric and eight-membered $\{\cdots\text{Te}^{\text{II}}\cdots\text{I}\cdots\text{Te}^{\text{IV}}\cdots\text{I}\}_2$ synthons are constructed from the tellurium(II) atom forming two Te...I interactions with two independent tellurium(IV)-bound iodide atoms. The third synthon, is non-symmetric and four-membered $\{\cdots\text{Te}^{\text{II}}\cdots\text{I}\cdots\text{Te}^{\text{IV}}\cdots\text{I}\}$. This is the first synthon to be described herein involving tellurium in both the +II and +IV oxidation states. The Te...I separations in this latter synthon are the longest of the four independent Te...I interactions. In the crystal, I...I halogen-bonding combine with N-H...I hydrogen-bonds to form a double-layer; see ESI† Fig. S4.

The third one-dimensional aggregation is seen in the crystal of **5** which comprises binuclear molecules formulated as $\text{PhTe}^{\text{II}}(\mu\text{-Se})\text{Te}^{\text{IV}}\text{PhI}_2$.⁹² In terms of Te...I interactions, each

independent tellurium atom forms two such interactions as does each of the two independent, tellurium(IV)-bound iodide atoms; all separations are long, Table 1. A non-symmetric, four-membered $\{\cdots\text{Te}^{\text{IV}}\cdots\text{I}\}_2$ synthon is formed where tellurium(IV) atoms arranged in a zigzag fashion along the glide plane are bridged by the two independent iodide atoms. The synthons are corner-shared through the tellurium(IV) atoms to form a ribbon and represented by entries (i) and (ii) in Fig. 2(c). A second, non-symmetric and four-membered $\{\cdots\text{Te}^{\text{II}}\cdots\text{I}\cdots\text{Te}^{\text{IV}}\cdots\text{I}\}$ synthon, like that seen in the crystal of **4**, provides additional links between the $\{\cdots\text{Te}^{\text{IV}}\cdots\text{I}\}_2$ synthons, *i.e.* entries (iii) and (iv); these additional Te...I interactions involve the tellurium(II) atom. The overall assembly features concatenated and alternating four-membered synthons. The ribbons are connected into a two-dimensional array *via* centrosymmetric $\{\cdots\text{Te}^{\text{II}}\text{Se}\}_2$ synthons as illustrated in ESI† Fig. S5.

There is a sole example where molecules are assembled within a three-dimensional architecture with Te...I interactions between the molecules operating in three dimensions. The tetra-tellurium molecule, **6**, hexamethylenetetratellurafulvalene di-iodide,⁹³ comprises a single tellurium(IV) site with three positions occupied by tellurium(II) atoms in the inner five-membered rings, Fig. 1. Owing to its complexity, computational chemistry was not conducted on this crystal so the description of the molecular packing along with diagrams is included in ESI† Fig. S6.

3.2 Theoretical study

A detailed theoretical investigation of the specific assemblies inclusive of Te...I interactions in **1–5** has been conducted with a focus upon an exploration of the nature of the non-covalent bonding interactions through various computational techniques. As highlighted in section 3.1, these interactions play a crucial role in determining the dimensionality of the assemblies and crystal structure organisation. The analysis commences with the zero-dimensional aggregates identified in the crystals of **1** and **2**.

The molecular electrostatic potential (MEP) surface for the molecule in dimeric **1**, depicted in Fig. 3(a), reveals that the MEP maximum is positioned at the hydrogen atoms of the methoxy groups, registering at +26.3 kcal mol^{−1}. The anticipated σ -holes on the tellurium(IV) and tellurium(II) atoms are mostly masked by the intramolecular interactions with the lone-pairs (LPs) on the oxygen atoms of the methoxy groups, which are orientated towards the tellurium atoms. Notably, a σ -hole on the tellurium(II) atom remains partially accessible, attributed to the rotation of the dimethoxyphenyl ring, and measures +7.5 kcal mol^{−1}. The MEP minimum, indicating nucleophilic regions, is found at the iodide atom, *i.e.* −32.7 kcal mol^{−1}. The larger MEP value at the hydrogen atoms compared to the tellurium atoms suggests that this molecule is more predisposed to forming hydrogen bonds, *e.g.* CH...I, rather than chalcogen bonds.

The Quantum Theory of Atoms in Molecules (QTAIM) analysis of the self-assembled dimer, illustrated in Fig. 3(b),



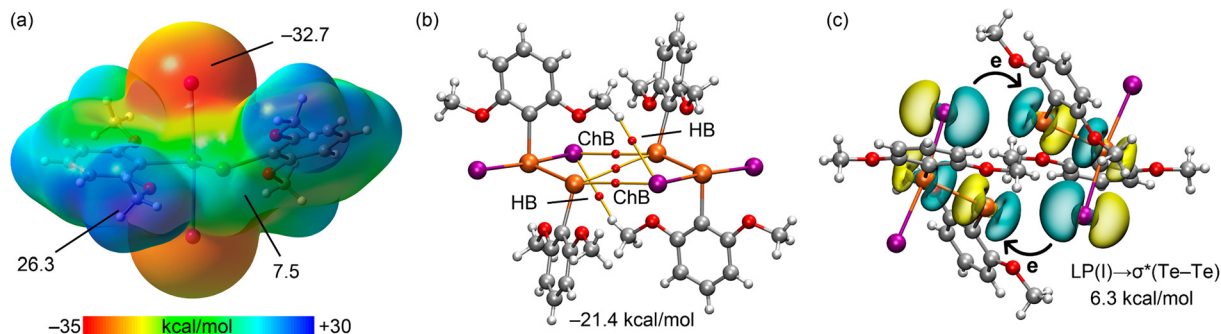


Fig. 3 (a) MEP surface of the molecule in **1** with isosurface 0.001 a.u., (b) QTAIM analysis (BCPs are represented as red spheres and bond paths as orange lines; only intermolecular contacts are indicated for clarity) and (c) plot of the NBOs involved in the chalcogen bonds (ChB). Stabilisation energies are indicated in (b) and (c). HB = hydrogen bond.

validates the presence of Te \cdots I interactions through bond critical points (BCPs) and bond paths linking the tellurium and iodide atoms. Additionally, this analysis identifies two symmetric CH \cdots I contacts, corroborating the MEP findings. The QTAIM analysis also reveals the existence of a BCP and a bond path interconnecting the tellurium(II) atoms suggesting a van der Waals interaction between them. The dimerization energy is significantly great at -21.4 kcal mol $^{-1}$, due to the cooperation of both interactions. The density (ρ), the Lagrangian kinetic energy density (G), the potential energy density (V) the total energy density (H) and the Laplacian ($\nabla^2\rho$) of the electron density values at the BCPs represented

in Fig. 3(b) are collated in Table 2. The positive values of H and $\nabla^2\rho$ along with low values of ρ , confirm the non-covalent nature of the interactions. Higher values of ρ and V at the BCPs associated with chalcogen bonds indicate that these are the predominant forces driving the formation of the dimer.

Charge transfer effects were examined next using Natural Bond Orbital (NBO) analysis, with the relevant NBOs involved in the chalcogen bonds show cased in Fig. 3(c). This analysis illustrates electron donation from the LP at the iodide atom to the anti-bonding orbital, $\sigma^*(\text{Te}^{\text{II}}-\text{Te}^{\text{IV}})$, resulting in a total stabilisation energy of 6.3 kcal mol $^{-1}$ for the two LP(I) \rightarrow $\sigma^*(\text{Te}^{\text{II}}-\text{Te}^{\text{IV}})$ contributions. This relatively modest orbital contribution is consistent with the elongated Te \cdots I experimental distance of $4.0169(5)$ Å.

A similar analysis was conducted for dimeric **2**, where the MEP surface distinctly illustrates the σ -holes on both tellurium(II) and tellurium(IV) atoms, Fig. 4(a), exhibiting comparable energies of $+24.8$ and $+24.3$ kcal mol $^{-1}$, respectively. The MEP minimum is found at the iodide atom's negative belt, marked by -16.9 kcal mol $^{-1}$. Moreover, the MEP highlights negative potential, *i.e.* -8.8 kcal mol $^{-1}$, over the π -system of the aromatic ring, suggesting regions prone to electron donation.

The QTAIM analysis, as shown in Fig. 4(b), confirms the engagement of both σ -holes in chalcogen-bonding interactions. Each Te $^{\text{II}}\cdots$ I interaction is delineated by a BCP and a bond path that connects the tellurium and iodide atoms. Intriguingly, the QTAIM analysis further reveals a BCP and bond path linking the tellurium(IV) atom to a carbon atom within the aromatic ring, demonstrating the formation of two additional chalcogen bonds within the aggregate with the π -systems serving as electron donors as highlighted in the inset plot of Fig. 4(b). A BCP and bond path linking the two tellurium(II) atoms were also observed, indicating an interaction between these atoms. The QTAIM data at the BCPs (Table 2) is typical of non-covalent interactions, with positive values of H and $\nabla^2\rho$.

The interaction energy for the aggregate in **2** is significantly negative (-28.2 kcal mol $^{-1}$) and is predominantly attributed to the chalcogen bonds, as corroborated by the QTAIM data of Table 2. Moreover, the QTAIM values are

Table 2 QTAIM parameters in a.u. measured at the bond critical points shown in Fig. 3–7. The atom labels correspond to those of the X-ray structures (retrieved from the respective CIFs). See also Fig. 2 for the labelling of some interactions

	ρ	$G(r)$	$V(r)$	$H(r)$	$\nabla^2\rho$
1					
Te1–Te2	0.0071	0.0028	−0.0025	0.0004	0.0128
Te2 \cdots I1 (i)	0.0073	0.0033	−0.0027	0.0006	0.0155
C13–H7 \cdots I1	0.0037	0.0020	−0.0015	0.0005	0.0099
2					
Te2 \cdots I2 (i)	0.0108	0.0050	−0.0045	0.0005	0.0219
Te2 \cdots Te2	0.0049	0.0019	−0.0016	0.0003	0.0090
Te1 \cdots C10	0.0098	0.0056	−0.0048	0.0008	0.0257
3					
C2–H3 \cdots I1	0.0052	0.0030	−0.0023	0.0007	0.0147
Te2 \cdots I2 (i)	0.0301	0.0117	−0.0149	−0.0032	0.0341
Te1 \cdots C7	0.0094	0.0053	−0.0045	0.0008	0.0247
C9–H8 \cdots I2	0.0072	0.0041	−0.0034	0.0008	0.0198
C15–H15 \cdots C5	0.0069	0.0044	−0.0035	0.0009	0.0210
4					
Te2 \cdots I1 (iv)	0.0096	0.0044	−0.0039	0.0005	0.0193
Te2 \cdots I3 (ii)	0.0119	0.0054	−0.0051	0.0004	0.0232
Te2 \cdots I2	0.0060	0.0025	−0.0021	0.0004	0.0117
C4–H3 \cdots C12	0.0043	0.0028	−0.0021	0.0008	0.0145
C6–H5 \cdots C9	0.0050	0.0032	−0.0024	0.0008	0.0159
Te1 \cdots I3 (i)	0.0293	0.0108	−0.0136	−0.0028	0.0322
5					
Te1 \cdots I2 (i)	0.0093	0.0038	−0.0034	0.0004	0.0170
C4–H3 \cdots I2	0.0056	0.0031	−0.0024	0.0007	0.0150
Te2 \cdots I2 (iii)	0.0075	0.0035	−0.0029	0.0006	0.0167
Te2 \cdots C8	0.0100	0.0054	−0.0047	0.0008	0.0248
Te1 \cdots I1 (ii)	0.0080	0.0037	−0.0031	0.0005	0.0168



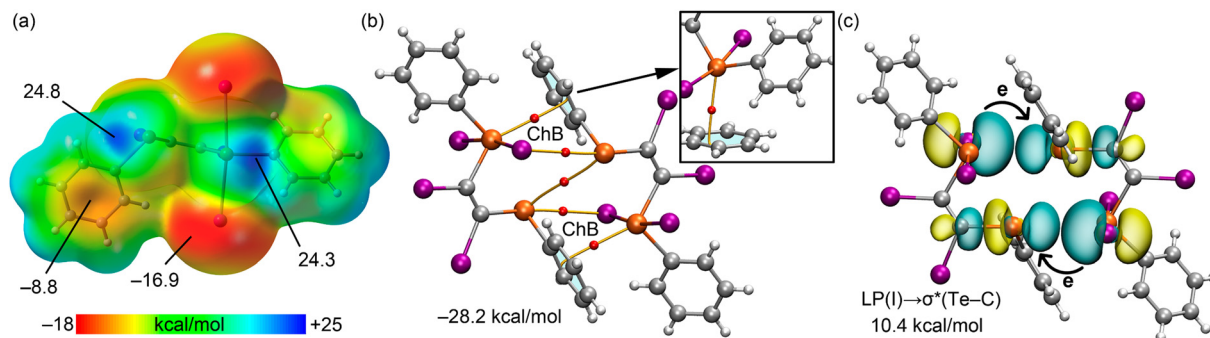


Fig. 4 (a) MEP surface of molecule **2** with isosurface 0.001 a.u., (b) QTAIM analysis (BCPs are represented as red spheres and bond paths as orange lines; only intermolecular contacts are indicated for clarity) and (c) plot of the NBOs involved in the chalcogen bonds (ChB). Stabilisation energies are indicated in (b) and (c).

larger in **2** with respect to **1** at the $\text{Te}\cdots\text{I}$ BCPs, consistent with the larger binding energy of the former. The NBO analysis elucidates the lone-pairs (LPs) at the iodide atoms are precisely aligned with the anti-bonding $\sigma^*(\text{Te}^{\text{II}}-\text{C})$ orbitals, Fig. 4(c). This alignment results in a total stabilisation energy ($10.4 \text{ kcal mol}^{-1}$) surpassing that observed for the aggregate in **1**. This greater stabilisation energy in **2** correlates with the shorter $\text{Te}\cdots\text{I}$ distance of $3.7708(4) \text{ \AA}$ (*cf.* $4.0169(5) \text{ \AA}$ in **1**) and enhanced directionality (the $\text{I}\cdots\text{Te}-\text{C}$ angle is $171.97(11)^\circ$ *cf.* $\text{I}\cdots\text{Te}^{\text{II}}-\text{Te}^{\text{IV}} = 157.377(15)^\circ$ in **1**), indicating a stronger and more directed chalcogen-bonding interaction in **2** compared to **1**.

Beyond the zero-dimensional assemblies of **1** and **2**, the analysis extends to the chalcogen bonds in the crystal featuring supramolecular tapes, starting with **3**, with a particular focus on the shorter $\text{Te}\cdots\text{I}$ interaction (labelled as (i) in Fig. 2(a), and distance = $3.2175(4) \text{ \AA}$). The longest chalcogen bond interaction ($3.9902(4) \text{ \AA}$, labelled as (ii) in Fig. 2(a)), is not discussed in detail because it is similar to those already analysed in **1** and **2**. The MEP surface of the molecule in **3**, Fig. 5(a), reveals a pronounced σ -hole on the tellurium(II) atom opposite the $\text{Te}^{\text{II}}-\text{I}$ bond, showing a high potential of $+40.8 \text{ kcal mol}^{-1}$, and a less intense σ -hole opposite the $\text{Te}^{\text{II}}-\text{C}$ bond, with a potential of $+21.3 \text{ kcal mol}^{-1}$. Additionally, the MEP is positive around the hydrogen

atom of the methyl groups ($+19.5 \text{ kcal mol}^{-1}$), with the MEP minimum once again located at one of the iodide atoms ($-31.4 \text{ kcal mol}^{-1}$), echoing the findings for **1**. The significant MEP values at both the iodide and tellurium atoms suggest the likelihood of strong chalcogen-bonding interactions.

The QTAIM analysis of the dimer reveals a complex arrangement of BCPs and bond paths, Fig. 5(b). Besides the BCPs and bond paths indicative of chalcogen bonds, the presence of methyl groups promotes the formation of additional $\text{CH}\cdots\text{I}$ and $\text{CH}\cdots\pi$ interactions, characterised by BCPs and bond paths connecting the hydrogen atoms to iodide and aromatic-C atoms, respectively. This complex network of interactions contributes to a notably strong dimerization energy, *i.e.* $-59.5 \text{ kcal mol}^{-1}$. The QTAIM parameters in Table 2 for **3** corroborate the chalcogen bonds are the strongest interactions. Moreover, one chalcogen bond presents a negative value of H , suggesting a partial covalent interaction. Furthermore, the NBO analysis presented in Fig. 5(c) reveals charge transfer from the iodide-bound lone-pair (LP) to the anti-bonding $\sigma^*(\text{Te}^{\text{II}}-\text{I})$ orbital is exceptionally substantial ($51.2 \text{ kcal mol}^{-1}$ in total), nearly mirroring the total interaction energy. This finding strongly implies partial covalent nature for these chalcogen bonds in line with the QTAIM analysis and underscores the significant role of electron donation in stabilising these interactions.

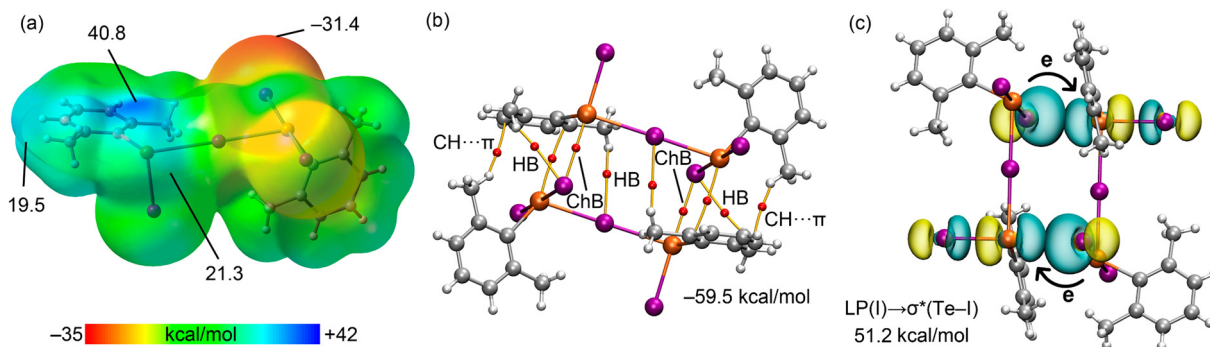


Fig. 5 (a) MEP surface of **3** with isosurface 0.001 a.u., (b) QTAIM analysis (BCPs are represented as red spheres and bond paths as orange lines; only intermolecular contacts are indicated for clarity) and (c) plot of the NBOs involved in the chalcogen bonds. Stabilisation energies are indicated in (b) and (c).



The chalcogen bonds within the 1:1 co-crystal **4** have also been analysed theoretically with an emphasis on the distinct interactions involving the tellurium(IV) centre as the $\text{Te}^{\text{IV}}\cdots\text{I}$ interactions resemble those already analysed in detail in **3**. The MEP surfaces for both coformers are illustrated in Fig. 6(a). For the tellurium(IV)-containing coformer, the MEP values at the σ -holes are 40.2 and 26.3 kcal mol⁻¹, and correspond to the regions at the extension of the Te–I and Te–C bonds, respectively; these values mirror the values seen in **3**. In the case of the second coformer, the attachment of thiourea to the tellurium(II) atom significantly diminishes the σ -hole opposite the $\text{Te}^{\text{II}}\text{--C}$ bond, *i.e.* to 9.4 kcal mol⁻¹. As expected for this molecule, the MEP maximum is located at the thiourea-bound hydrogen atoms (90.0 kcal mol⁻¹). The MEP minima at the iodide atoms are –13.8 kcal mol⁻¹ for $\text{PhTe}^{\text{IV}}\text{I}_3$ and –42.7 kcal mol⁻¹ for $\text{PhTe}^{\text{II}}\text{I}[\text{S}=\text{C}(\text{NH}_2)_2]$, indicating a strong $\text{Te}^{\text{IV}}\cdots\text{I}$ interaction, supported by the short experimental distance of 3.2651(7) Å (interaction labelled as (i) in Fig. 2(b) *cf.* 3.7601(14) Å for interaction (ii) in Fig. 2(b)).

The distribution of BCPs and bond paths in the tetrameric assembly, shown in Fig. 6(b), verifies a complex network comprising up to eight $\text{Te}\cdots\text{I}$ interactions as well as $\text{CH}\cdots\pi$ contacts. As noted above, both tellurium(II) and tellurium(IV) function as σ -hole donors in establishing chalcogen bonds. The

formation energy of this assembly is significantly great (–86.5 kcal mol⁻¹), again hinting at a partial covalent nature for the shorter chalcogen bonds. This covalent character is further confirmed by the NBO analysis, which examined both types of chalcogen bonds. For the $\text{Te}^{\text{IV}}\cdots\text{I}$ interactions, the stabilisation energy from the $\text{LP}(\text{I}) \rightarrow \sigma^*(\text{Te}^{\text{IV}}\text{--I})$ charge transfer is 77.5 kcal mol⁻¹ (Fig. 6(c), left) underscoring the strong covalent character. This is further corroborated by the QTAIM parameters, Table 2, revealing a negative total energy density value ($H = -0.0028$ a.u.) and a significant charge density (0.0293 a.u.) compared to the remaining BCPs. Conversely, the stabilisation energy from the $\text{LP}(\text{I}) \rightarrow \sigma^*(\text{Te}^{\text{IV}}\text{--C})$ charge transfer (Fig. 6(c), right) is considerably lower (14.0 kcal mol⁻¹), consistent with the longer $\text{Te}^{\text{IV}}\cdots\text{I}$ distance, illustrating the varying strength and nature of these $\text{Te}^{\text{IV}}\cdots\text{I}$ interactions within the co-crystal.

In the final example to be analysed, the chalcogen bonds present in the crystal of **5** were evaluated. The MEP surface, depicted in Fig. 7(a), reveals two distinct σ -holes on each type of tellurium atom. Notably, the σ -holes opposite the Te–Se bonds on both tellurium(II) and tellurium(IV) atoms are equivalent, each valued at 25.0 kcal mol⁻¹, and those opposite the Te–C bonds are comparable, with values of 15.7 and 18.8 kcal mol⁻¹ for the tellurium(II) and tellurium(IV) atoms, respectively. The MEP minimum is located on an

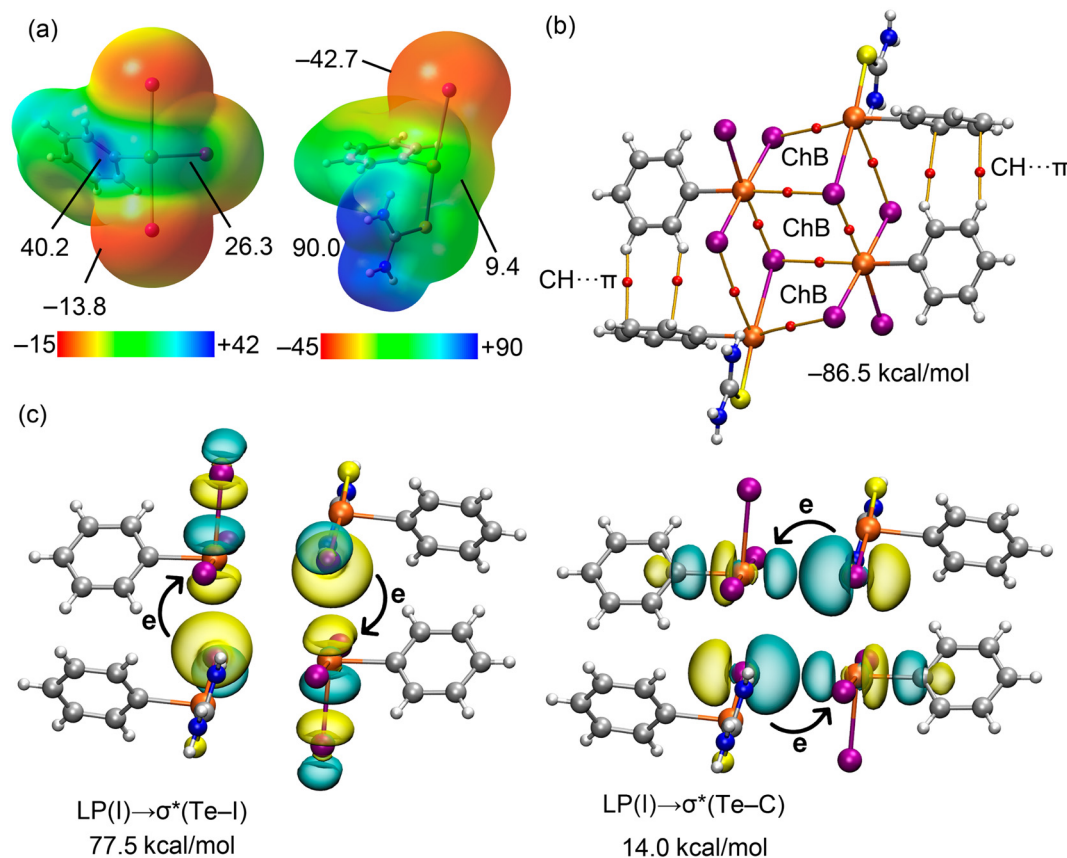


Fig. 6 (a) MEP surfaces of the two coformers of co-crystal **4** with isosurface 0.001 a.u., (b) QTAIM analysis (BCPs are represented as red spheres and bond paths as orange lines; only intermolecular contacts are indicated for clarity) and (c) plot of the NBOs involved in both types of chalcogen bonds. Stabilisation energies are indicated in (b) and (c).



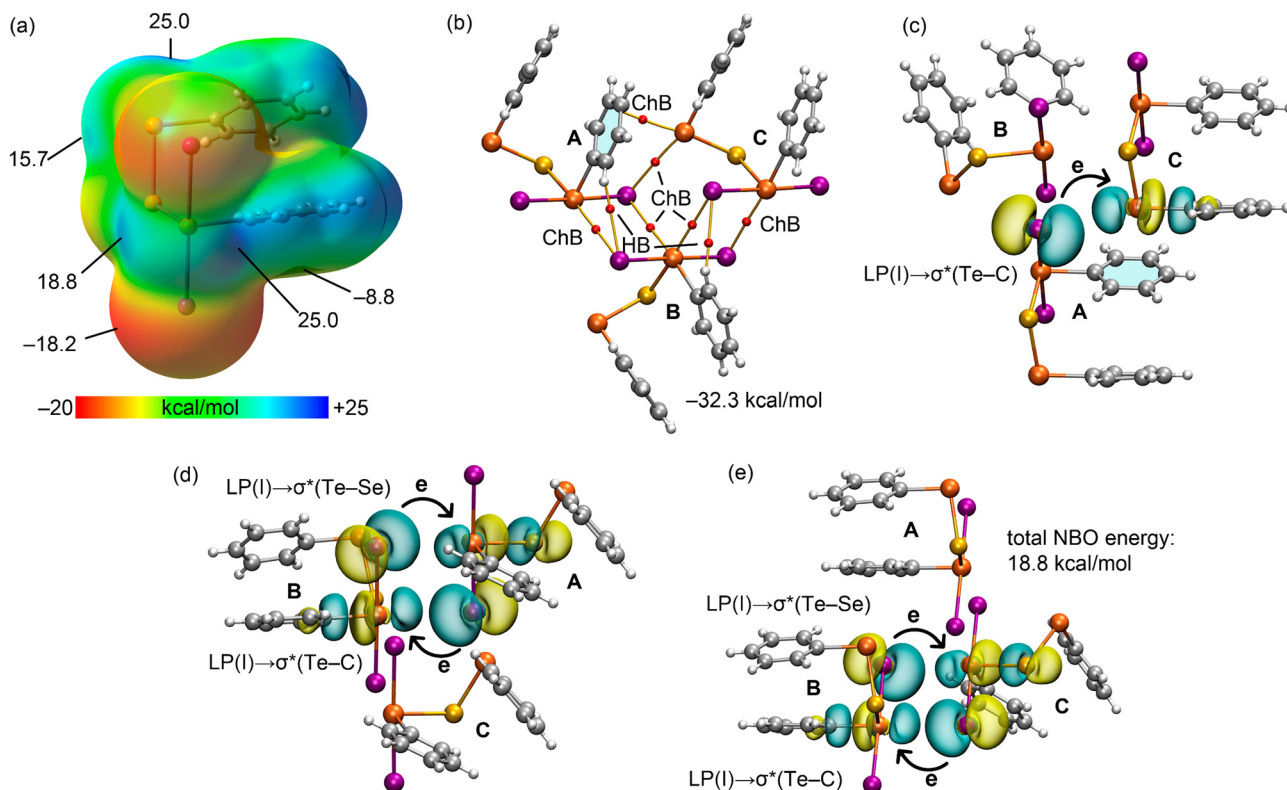


Fig. 7 (a) MEP surface of the molecule in **5** with isosurface 0.001 a.u., (b) QTAIM analysis (BCPs are represented as red spheres and bond paths as orange lines; only intermolecular contacts are indicated for clarity) and (c–e) plots of the NBOs involved in both types of chalcogen bonds. Stabilisation energies are indicated in (b) and (e).

iodide atom ($-18.2 \text{ kcal mol}^{-1}$), with the MEP also showing nucleophilicity over the π -system of the aromatic ring.

A trimeric assembly of this compound, as presented in Fig. 7(b), was examined using the QTAIM method, which identifies five $\text{Te}\cdots\text{I}$ chalcogen bonds and one $\text{Te}\cdots\pi(\text{aryl})$ interaction. In addition, the QTAIM analysis uncovers the presence of two $\text{CH}\cdots\text{I}$ hydrogen bonds. The trimerization energy, calculated at $-32.3 \text{ kcal mol}^{-1}$, aligns well with the dimerization energies observed for the aggregates in **1** and **2**. All identified chalcogen bonds were further scrutinised using NBO analysis, which consistently observed lone-pair (LP) to anti-bonding σ^* charge transfer across the interactions, *i.e.* $\text{LP}(\text{I}) \rightarrow \sigma^*(\text{Te}-\text{C})$ and $\text{LP}(\text{I}) \rightarrow \sigma^*(\text{Te}-\text{Se})$. Here, the cumulative stabilisation energy was $18.8 \text{ kcal mol}^{-1}$, suggesting the predominantly non-covalent character of these interactions. This aligns well with the QTAIM values (Table 2) with positive total energy density (H) values and ρ values that range from 0.0056 to 0.0100 a.u. An interesting point to note is the significance of all four σ -holes in the formation of the supramolecular tape illustrated in Fig. 2(c) and revealed by the QTAIM analysis, highlighting their critical role in the assembly of the molecules in the crystal.

4. Overview

Six crystals containing at least one each of tellurium(II), tellurium(IV) and iodide have been evaluated for their

propensity to form $\text{Te}\cdots\text{I}$ interactions. In zero-dimensional **1** and **2**, and one-dimensional **3–5** there are equal numbers of tellurium(II) and tellurium(IV) atoms. In **1–3** the $\text{Te}\cdots\text{I}$ interactions involve tellurium(II) only, but in each of **4** and **5**, equal numbers of $\text{Te}^{\text{II}}\cdots\text{I}$ and $\text{Te}^{\text{IV}}\cdots\text{I}$ interactions are apparent; in three-dimensional **6**, there is a 3:1 ratio between tellurium(II) and tellurium(IV) atoms, and of the five $\text{Te}\cdots\text{I}$ interactions identified, four involve tellurium(II). The dominance of tellurium(II) over tellurium(IV) in forming $\text{Te}\cdots\text{I}$ interactions in this rather small data set may simply relate to steric crowding given the tellurium(II) atoms are usually two-coordinate whereas the tellurium(IV) centres are four-coordinate.

Prior to summarising the results of the computational chemistry investigations, it is salient to place the observed number of mixed tellurium(II), tellurium(IV) and iodide crystals featuring $\text{Te}\cdots\text{I}$ interactions in context. In the CSD⁷² there are 207 independent, non-charged crystals containing both tellurium and iodide but excluding other metals. Following these criteria, there are seven examples of mixed tellurium(II), tellurium(IV) and iodide crystals, while each of these featured $\text{Te}\cdots\text{I}$ interactions one of the crystals saw these operating in concert with $\text{I}\cdots\text{I}$ interactions.⁹⁴ There are 71 examples of crystals containing tellurium(II) and iodide of which 15 (21%) contain $\text{Te}\cdots\text{I}$ interactions operating independently of other, structure-directing non-covalent interactions while there are 129 crystals containing



tellurium(IV) and iodide, 69 (54%) exhibit Te \cdots I interactions operating independently of other obvious directional non-covalent interactions. These data suggest that tellurium(II) compounds are less likely to form Te \cdots I interactions than tellurium(IV) compounds.

Comprehensive theoretical analyses, employing Molecular Electrostatic Potential (MEP), Quantum Theory of Atoms in Molecules (QTAIM) and Natural Bond Orbital (NBO) techniques, have elucidated the intricate chalcogen-bonding interactions across the series of compounds, 1–5. These studies reveal the significant influence of σ -holes on the formation of chalcogen bonds and their subsequent impact on the structural dimensionality and stability of the observed molecular assemblies. From zero-dimensional aggregates to complex trimeric and tetrameric arrangements within one-dimensional tapes, the findings underscore the variability in interaction strengths, highlighting the covalent to non-covalent nature of these bonds. The discernible σ -holes, critical to the formation of supramolecular tapes, demonstrate the nuanced interplay between molecular geometry, electron density and interaction energy. It is also noteworthy to emphasise that the Te \cdots I interactions identified in this study are characterized as chalcogen bonds rather than halogen bonds. This distinction is based on the directionality of the interactions and the observed positive and negative MEP values at the interacting atoms.

In terms of energies of stabilisation imparted by the identified Te \cdots I interactions, these span a considerable range. Thus, while most of the interaction energies are generally small, *i.e.* ranging from 3.1 to 7.0 kcal mol^{−1} in 1, 2 and 5, these increase markedly in two examples, *i.e.* 25.6 to 38.8 kcal mol^{−1} in 3 and 4, involving tellurium(II) and tellurium(IV) centres, respectively, where the Te \cdots I separations are consistent with some covalent character.

5. Conclusions

When both tellurium(II) and tellurium(IV) atoms are present in iodide-containing crystals, a tellurium(II) atom will always form a Te \cdots I interaction but not necessarily the tellurium(IV) atom. The Te \cdots I interactions are always classified as chalcogen bonds being distinguished by the directionality of the interactions. This work not only advances the understanding of chalcogen bonds in supramolecular assembly/crystal engineering but also sets the groundwork for future exploration into the design and synthesis of novel materials guided by chalcogen-bonding principles.

Author contributions

Rosa M. Gomila: computational chemistry, investigation, visualisation, writing – original draft. Antonio Frontera: computational chemistry, investigation, visualisation, writing – original draft. Edward R. T. Tiekink: conceptualisation, CSD analysis, investigation, visualisation, writing – original draft.

Conflicts of interest

The authors declare no competing financial interest.

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