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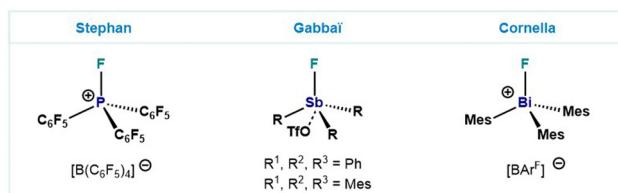
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**This work reports the synthesis and structural study of a family of chlorotriarylbismuthonium salts. The abstraction of a chlorine atom with  $\text{NaBAr}^F$  from triarylbismuth dichloride species leads to monomeric and dimeric chlorotriarylbismuthonium species, which show a distinct behavior in solution and solid-state in comparison to their fluorotriarylbismuthonium analogues.**

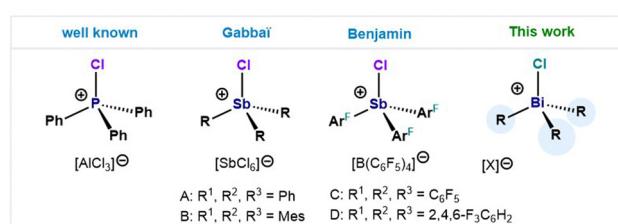
The synthesis and design of novel Lewis acids based on cheap and more abundant main group elements opens the door to sustainable catalysts that can have an impact in organic synthesis and applications thereof. Of particular importance, are those derived from organopnictonium ions. In this regard, numerous fluorophosphonium salts have been synthesized, demonstrating their Lewis acidity in a wide range of organic transformations (Fig. 1A).<sup>1–10</sup> The heightened Lewis acidity observed in cationic haloorganopnictonium salts is ascribed to the low-lying  $\sigma^*$ -orbital positioned *trans* to the halogen atom.<sup>11–13</sup> In contrast to its lighter congeners, heavier analogues in the group 15 have received comparable less attention despite their well-documented electrophilicity.<sup>14–17</sup> Remarkable examples in this front are the monomeric fluorotriarylstibonium ions developed by Gabbaï and co-workers where interaction between the cationic Sb(v) and a OTf<sup>–</sup> anion can be observed in the solid state (Fig. 1A).<sup>12</sup> Based on our interest in the study of organometallic compounds of bismuth,<sup>18–20</sup> our group has recently reported a structural study on a series of fluorotriarylbismuthonium salts,<sup>21</sup> which have shown a distinctive behaviour in comparison with its lighter analogues. In contrast to the monomeric cationic salts based on P and Sb,<sup>12</sup> we observed the formation of di- and trinuclear fluorobismuthonium compounds. This distinct behaviour was attributed to the highly electropositive Bi(v) center in combination

with the strong donor–acceptor interactions between the fluorine and Bi atom. Evaluation of the steric constraints on the aryl moieties eventually led to the isolation of the first monomeric fluorotriarylbismuthonium salt (Fig. 1A). Following this rationale, we realized that in contrast to fluoropnictonium salts, few examples of chlorotriarylpnictonium salts have been reported,<sup>12,17,22–28</sup> (Fig. 1B). Gabbaï demonstrated a remarkable example of a monomeric chlorotriarylstibonium hexachloroantimonate salt, and its application in the polymerization of THF and dimerization of 1,1-diphenylethylene (Fig. 1B).<sup>12</sup> Recently, Benjamin *et al.* enhanced the Lewis acidity by the introduction of electron-withdrawing aryl ligands,

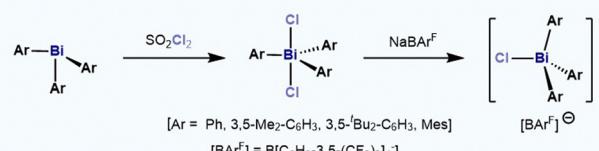
**A. Selected examples of monomeric fluoropnictonium salts**



**B. Selected examples of monomeric chloropnictonium salts**



**C. This work: Synthesis of high-valent chlorotriarylbismuthonium salts**



**Fig. 1** (A) Representative examples of mononuclear fluorotriarylpnictonium salts. (B) Examples of mononuclear chlorotriarylpnictonium salts. (C) This work: synthesis of high-valent chlorotriarylbismuthonium salts.

Max-Planck-Institut für Kohlenforschung, Kaiser-Wilhelm-Platz 1,  
45470 Mülheim an der Ruhr, Germany. E-mail: cornella@kofo.mpg.de

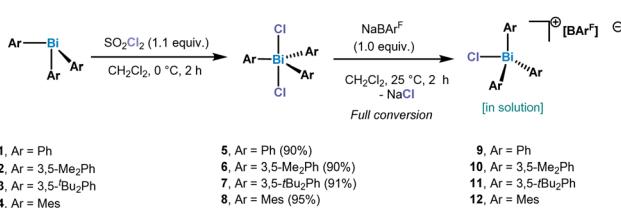
† Electronic supplementary information (ESI) available. CCDC 2356069–2356071, 2289340, 2296787, 2289343 and 2296786. For ESI and crystallographic data in CIF or other electronic format see DOI: <https://doi.org/10.1039/d4cc03364g>



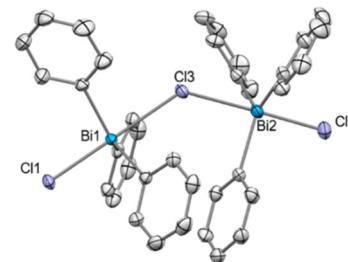
showing also the potential for formation of C–C bonds in a Friedel–Crafts alkylation reactions (Fig. 1B).<sup>28</sup> Moreover, the same group has demonstrated the formation of a dinuclear chlorotriarylstibonium salt by the use of 0.4 equiv. of  $[\text{Et}_3\text{Si}(\text{C}_7\text{H}_8)][\text{B}(\text{C}_6\text{F}_5)_4]$  as chloride abstracting agent. Based on our previous work,<sup>21</sup> we present the synthesis and structural study of chlorotriarylbismuthonium salts, including the first example of a monomeric chlorotriarylbismuthonium salt (Fig. 1C).

Our studies of chlorotriarylbismuthonium salts started with the synthesis of triarylbismuth dichlorides. In order to be able to compare the chlorobismuthonium cations with the fluorinated analogues, unsubstituted and substituted aryl ligands bearing  $^t\text{Bu}$  and Me groups have been selected. Oxidation of triarylbismuth complexes with 1.3 equiv. of  $\text{SO}_2\text{Cl}_2$  led to the isolation of the corresponding triarylbismuth dichlorides in high yields, which have been characterized by NMR, HRMS and SC-XRD (single crystal X-ray diffraction) (see ESI† for more details) (Scheme 1). Treating  $\text{Ph}_3\text{BiCl}_2$  with 1.0 equiv. of  $\text{NaBAr}^F$  ( $\text{BAr}^F = \text{B}[\text{C}_6\text{H}_3\text{-}3,5\text{-}(\text{CF}_3)_2]_4^-$ ) in  $\text{CH}_2\text{Cl}_2$  at 25 °C indicated the presence of a formally monomeric  $[\text{Ph}_3\text{BiCl}][\text{BAr}^F]$  (9) species in solution as suggested by the integration of the phenyl signals with respect to the  $\text{BAr}^F$  anion in the  $^1\text{H}$  NMR spectrum.

Crystals of compound 9 suitable for SC-XRD analysis could not be obtained and therefore its structure in the solid state remains elusive. However, when  $\text{Ph}_3\text{BiCl}_2$  was treated with 0.5 equiv. of  $\text{NaBAr}^F$  in  $\text{CH}_2\text{Cl}_2$  at 25 °C, the formation of a dinuclear chlorotriphenylbismuthonium salt  $[(\text{Ph}_3\text{BiCl})_2\text{Cl}][\text{BAr}^F]$  (9a) was observed in solution and solid-state as analyzed by NMR spectroscopy and SC-XRD (Fig. 2). The different formation of 9 and 9a in solution is in stark contrast to our previous observations with the parent difluorotriphenylbismuth, where the amount of  $\text{NaBAr}^F$  did not affect the reactivity of the halogen abstraction step.<sup>21</sup> The formation of 9a resembles rather its lighter element homologue, the chlorotriphenylstibonium cation, which was obtained after treating  $\text{Ph}_3\text{SbCl}_2$  with 0.4 equiv. of  $[\text{Et}_3\text{Si}(\text{C}_7\text{H}_8)][\text{B}(\text{C}_6\text{F}_5)_4]$ .<sup>28</sup> Solid-state analysis of 9a reveals no additional intermolecular  $\text{Bi}\cdots\text{Cl}$  interactions between the individual units. As illustrated in Fig. 2, both Bi centres adopt a trigonal bipyramidal geometry with chlorine ligands in apical and the phenyl groups in equatorial positions. The chlorobismuthonium cation (9a) exhibits a bent  $\text{Bi}1\text{-Cl}3\text{-Bi}2$  geometry ( $\text{Bi}1\text{-Cl}3\text{-Bi}2: 128.10(2)^\circ$ ), attributed to packing effects during crystallization and the larger atomic size of chlorine compared to fluorine. Whereas the terminal  $\text{Bi}1\text{-Cl}1$  and  $\text{Bi}2\text{-Cl}2$  bond



**Scheme 1** Synthesis of triarylbismuth dichlorides 5–8 and chlorotriarylbismuthonium salts.

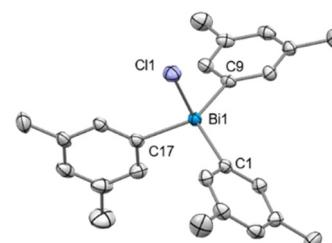


**Fig. 2** Solid state structure of 9a.<sup>29</sup> Ellipsoids are drawn at the 50% probability level. H atoms, disordered parts and  $\text{BAr}^F$  are omitted for clarity. Selected bond lengths (Å) and angles (°):  $\text{Bi}1\text{-Cl}1: 2.4856(7)$ ,  $\text{Bi}1\text{-Cl}3: 2.7995(7)$ ,  $\text{Bi}2\text{-Cl}2: 2.4996(8)$ ,  $\text{Bi}2\text{-Cl}3: 2.8274(7)$ ,  $\text{Bi}1\text{-Cl}3\text{-Bi}2: 128.10(2)$ ,  $\text{Cl}1\text{-Bi}1\text{-Cl}3: 179.25(2)$ ,  $\text{Cl}2\text{-Bi}2\text{-Cl}3: 177.70(2)$ .

lengths were determined to be 2.4856(7) Å and 2.4996(8) Å, the distances between the bridging chloride and the Bi centres are characteristically longer ( $\text{Bi}1\text{-Cl}3: 2.7995(7)$ ,  $\text{Bi}2\text{-Cl}3: 2.8274(7)$  Å).<sup>21</sup>

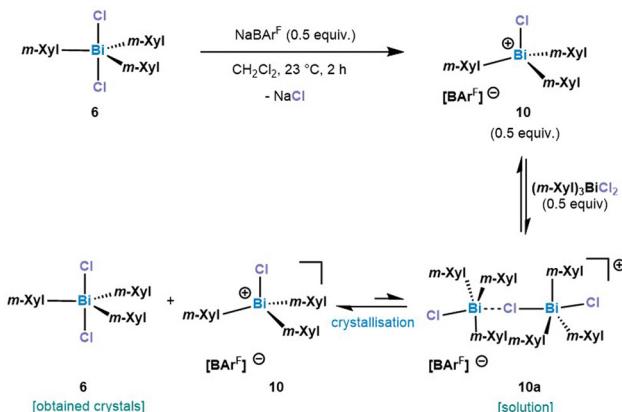
The addition of 1.0 equiv.  $\text{NaBAr}^F$  to  $(m\text{-Xyl})_3\text{BiCl}_2$  (6) ( $m\text{-Xyl} = \text{meta-xylene}$ ) in  $\text{CH}_2\text{Cl}_2$  for 2 h at 25 °C resulted in the formation of mononuclear  $[(m\text{-Xyl})_3\text{BiCl}][\text{BAr}^F]$  (10) in solution, as judged by the integration in the  $^1\text{H}$  NMR spectrum. Crystals suitable for X-ray diffraction confirmed the isolation of a mononuclear bismuthonium salt as depicted in Fig. 3. The solid state structure reveals that the Bi atoms adopts a distorted tetrahedral geometry ( $\tau_4 = 0.89$ ).<sup>30</sup> The Bi–Cl bond length of 2.4028(8) Å is slightly shorter in complex 10 compared to the terminal Bi–Cl length of complex 9a. Interestingly, the presence of the less electronegative chlorine atom prevents dimerization or trimerization processes, and stands in contrast to the observations for its fluorinated analogues, where sterics controlled these processes.<sup>21</sup>

After treating 6 with 0.5 equiv. of  $\text{NaBAr}^F$  a new set of  $^1\text{H}$  NMR signals was observed that is in line with a formally formation of a dinuclear bismuthonium  $[(m\text{-Xyl})_3\text{BiCl}]_2\text{Cl}][\text{BAr}^F]$  species (10a) in solution as judged by  $^1\text{H}$  NMR. However, contrary to our expectations, crystallization of this compound led to an SC-XRD of  $(m\text{-Xyl})_3\text{BiCl}_2$  (6). We hypothesize that the chloride abstraction with 0.5 equiv.  $\text{NaBAr}^F$  leads to the formation of a monomeric chlorobismuthonium cation 10, which reacts with a neutral triarylbismuth dichloride (6) resulting in a



**Fig. 3** Solid state structure of 10. Ellipsoids are drawn at the 50% probability level. H atoms and  $\text{BAr}^F$  are omitted for clarity. Selected bond lengths (Å) and angles (°):  $\text{Bi}1\text{-Cl}1: 2.4028(8)$ ,  $\text{Bi}1\text{-C}1: 2.185(3)$ ,  $\text{Bi}1\text{-C}9: 2.180(3)$ ,  $\text{Bi}1\text{-C}17: 2.189(3)$ ,  $\text{C}1\text{-Bi}1\text{-C}9: 121.63(11)$ ,  $\text{C}1\text{-Bi}2\text{-C}17: 111.32(10)$ ,  $\text{C}9\text{-Bi}2\text{-C}17: 112.55(11)$ .





Scheme 2 Proposed behaviour of **10a** in solution and solid state. *m*-Xyl = 3,5-Me<sub>2</sub>-C<sub>6</sub>H<sub>3</sub>.

reversible equilibrium with dinuclear chlorobismuthonium salt (**10a**) in solution (Scheme 2). Further crystallization leads to the precipitation of complex **6**. A similar equilibrium was observed with the trimeric fluorobismuthonium salt bearing phenyl as ligands using 1.0 equiv. of BAR<sup>F</sup>.<sup>21</sup>

Increasing the steric bulk by replacing the Me groups by <sup>t</sup>Bu, a different behaviour in the formation of the bismuthonium salt could be observed. The addition of 1.0 equiv. of NaBAR<sup>F</sup> to complex **7** leads to a new set of <sup>1</sup>H NMR signals that are in line with a formally monomeric chlorotriarylbismuthonium salt [<sup>t</sup>BuAr<sub>3</sub>BiCl][BAR<sup>F</sup>] (**11**). However, crystallization of complex **11** revealed the formation of a dimeric chlorobismuthonium salt (**11a**) as illustrated in Fig. 4, resembling the product and structure of reported fluorine analogue.<sup>21</sup> The solid-state structure of **11** reveals that both Bi atoms adopt a trigonal bipyramidal geometry with two chloride ligands in apical positions and aromatic rings in equatorial positions (Fig. 4A). Both Bi atoms are united by a bridging chloride ligand, generating a Bi1-Cl2-Bi1 angle of 180.0°. Whereas the terminal Bi1-Cl1 bond lengths are shorter compared to the neutral parent complex **7** [**11**, Bi1-Cl1: 2.4857(7) Å; **7**, Bi1-Cl1: 2.5954(5) Å (see ESI† for more details), the bond length between the

bridging chlorine and both Bi atoms are increased to 2.80068(13) Å. While the Bi1-Cl2-Bi1 axis displays an angle of 180.0°, the angle of a single <sup>t</sup>BuAr<sub>3</sub>BiCl<sub>2</sub> unit is slightly deviated from linearity [Cl1-Bi1-Cl2: 173.722(17)°]. Moreover, the presence of the chlorine atoms in the bismuthonium salt **11** allows the aromatic ligands a greater flexibility due to the longer Bi1-Cl2 bond lengths [Bi1-Cl2: 2.80068(13) Å] compared to the fluorinated analogue [Bi1-F1: 2.282(3) Å],<sup>21</sup> leading to a propeller-like conformation of the aromatic ligands (Fig. 4B). The addition of 0.5 equiv. of NaBAR<sup>F</sup> to **7** resulted in a dimeric bismuthonium salt **11a** [((*m*-Xyl)<sub>3</sub>BiCl)<sub>2</sub>Cl][BAR<sup>F</sup>] in solution and solid state (see ESI† for more details).

When comparing the formation of complexes **10** and **11a** in the solid state, it can be observed that the *meta* *tert*-butyl groups might exert an attractive effect as a consequence of the London-dispersion forces. Although the *m*-<sup>t</sup>Bu substituents are more sterically demanding in comparison to the *m*-Me groups, they serve as dispersion energy donors,<sup>31</sup> leading to the formation of a thermodynamically more stable dimeric species **11a** in solid state. The influence on the stability by *m*-<sup>t</sup>Bu groups through attractive dispersion interactions has been previously observed in the literature.<sup>31</sup>

Increasing the sterics at the *ortho*-position by adding Me substituents, resulted in the formation of a monomeric chlorotrimesylbismuthonium salt **12** in solution. Despite several crystallization attempts, suitable crystals could not be obtained. In order to get more insight into the potential structure, we reacted **12** with 1.0 equiv. of pyridine-*N*-oxide for 2 h at 25 °C in CH<sub>2</sub>Cl<sub>2</sub>. This reaction led to the isolation of complex **13** which was confirmed by NMR and SC-XRD (Scheme 3).

A broadening of the <sup>13</sup>C NMR signals of the pyridine ring and the <sup>15</sup>N NMR shift ( $\delta = -108.0$  ppm) indicated a weak Bi-O bond which was also supported by the results of the HRMS data, as only the chlorobismuthonium salt has been detected. The solid-state structure of **13** displays a coordination of the pyridine-*N*-oxide *via* the O atom to the Bi centre, displaying a Bi-O bond length of 2.404(6) Å (Fig. 5). Moreover, the Bi center adopts a trigonal bipyramidal geometry with the chlorine and oxygen atom in apical positions as well as Mes ligands in equatorial positions. Complex **13** exhibits a Cl-Bi-O angle of 173.5(2)°.

In conclusion, we have reported a series of chlorotriarylbismuthonium salts with BAR<sup>F</sup> as a weakly coordinating anion and extended the library of halotriarylbismuthonium salts. Contrary to our previous work on fluorobismuthonium salts, where the formation of either mono-, di- and trinuclear formation were

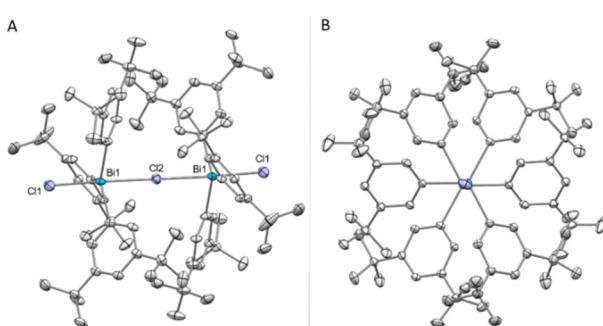
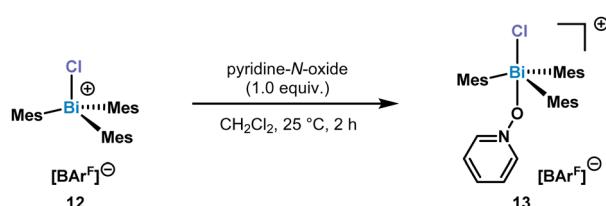
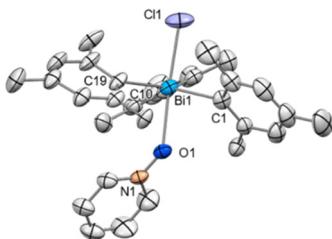


Fig. 4 (A) Solid state structure of **11a**.<sup>29</sup> Ellipsoids are drawn at the 50% probability level. H atoms and BAR<sup>F</sup> are omitted for clarity. Selected bond lengths (Å) and angles (°): Bi1-Cl1: 2.4857(7), Bi1-Cl2: 2.80068(13), Bi1-Cl2-Bi1: 180.0, Cl1-Bi1-Cl2: 173.722(17). (B) View along the Cl-Bi-Cl axis.



Scheme 3 Synthesis of complex **13**.





**Fig. 5** Solid state structure of **13**. Ellipsoids are drawn at the 50% probability level. H atoms and  $\text{BAr}^{\text{F}}$  are omitted for clarity. Selected bond lengths (Å) and angles (°):  $\text{Bi1}-\text{Cl1}$ : 2.507(7),  $\text{Bi1}-\text{O1}$ : 2.404(6),  $\text{Bi1}-\text{C1}$ : 2.229(6),  $\text{Bi1}-\text{C10}$ : 2.226(5),  $\text{Bi1}-\text{C19}$ : 2.236(6),  $\text{N1}-\text{O1}$ : 1.334(8),  $\text{C1}-\text{Bi1}-\text{C10}$ : 117.9(2),  $\text{C1}-\text{Bi1}-\text{C19}$ : 121.7(2),  $\text{C10}-\text{Bi1}-\text{C19}$ : 119.3(2),  $\text{Cl1}-\text{Bi1}-\text{O1}$ : 173.5(2),  $\text{N1}-\text{O1}-\text{Bi}$ : 136.7(4).

controlled by steric factors, the use of a chloride ligand has displayed a distinctive behaviour, leading to fast dynamic exchanges in solution. The less electronegative chlorine ligands generate more labile bismuthonium units, allowing easier access to monomeric halobismuthonium salts. However, London dispersion forces assist in the formation of dimeric species in solid-state.

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## Data availability

The data supporting this article have been included as part of the ESI.† Crystallographic data for **6-8**, **9a-11** and **13** have been deposited at the CCDC under CCDC numbers 2356069-2356071 (**6**, **10**, **13**) and 2289340 (**7**), 2296787 (**8**), 2289343 (**9a**), 2296786 (**11a**).

## Conflicts of interest

There are no conflicts to declare.

## Notes and references

<sup>1</sup> C. B. Caputo, L. J. Hounjet, R. Dobrovetsky and D. W. Stephan, *Science*, 2013, **341**, 1374.

- Pérez, L. J. Hounjet, C. B. Caputo, R. Dobrovetsky and D. W. Stephan, *J. Am. Chem. Soc.*, 2013, **135**, 18308.
- C. B. Caputo, D. Winkelhaus, R. Dobrovetsky, L. J. Hounjet and D. W. Stephan, *Dalton Trans.*, 2015, **44**, 12256.
- M. Mehta, M. H. Holthausen, I. Mallov, M. Pérez, Z.-W. Qu, S. Grimme and D. W. Stephan, *Angew. Chem., Int. Ed.*, 2015, **54**, 8250.
- M. Pérez, T. Mahdi, L. J. Hounjet and D. W. Stephan, *Chem. Commun.*, 2015, **51**, 11301.
- M. H. Holthausen, M. Mehta and D. W. Stephan, *Angew. Chem., Int. Ed.*, 2014, **53**, 6538 (*Angew. Chem.*, 2014, **126**, 6656).
- M. Mehta, I. García de la Arada, M. Pérez, D. Porwal, M. Oestreich and D. W. Stephan, *Organometallics*, 2016, **35**, 1030.
- I. Mallov and D. W. Stephan, *Dalton Trans.*, 2016, **45**, 5568.
- S. Postle, V. Podgorny and D. W. Stephan, *Dalton Trans.*, 2016, **45**, 14651.
- M. Vogler, L. Süss, J. H. W. LaFortune, D. W. Stephan and M. Oestreich, *Organometallics*, 2018, **37**, 3303.
- S. Benz, A. I. Poblador-Bahamonde, N. Low-Ders and S. Matile, *Angew. Chem. Int. Ed.*, 2018, **57**, 5408 (*Angew. Chem.*, 2018, **130**, 5506).
- M. Yang and F. P. Gabbaï, *Inorg. Chem.*, 2017, **56**, 8644.
- T. Thorwart and L. Greb, in *Encyclopedia of Inorganic and Bioinorganic Chemistry*, ed. R. A. Scott, Wiley, 2021, pp. 1-26.
- J. S. Murray, P. Lane and P. A. Politzer, *Int. J. Quantum Chem.*, 2007, **107**, 2286.
- A. Bauzá, T. J. Mooibroek and A. Frontera, *Chem. Phys. Chem.*, 2015, **16**, 2496.
- J. S. Murray, P. Lane, T. Clark, K. E. Riley and P. Politzer, *J. Mol. Model.*, 2012, **18**, 541.
- M. A. García-Monforte, M. Baya, D. Joven-Sancho, I. Ara, A. Martín and B. Menjón, *J. Organomet. Chem.*, 2019, **897**, 185.
- Y. Pang, N. Nöthling, M. Leutzsch, L. Kang, E. Bill, M. Van Gastel, E. Reijerse, R. Goddard, L. Wagner, D. SantaLucia, S. DeBeer, F. Neese and J. Cornell, *Science*, 2023, **380**, 1043.
- X. Yang, E. J. Reijerse, N. Nöthling, D. J. SantaLucia, M. Leutzsch, A. Schnegg and J. Cornell, *J. Am. Chem. Soc.*, 2023, **145**, 5618.
- M. Magre, J. Kuziola, N. Nöthling and J. Cornell, *Org. Biomol. Chem.*, 2021, **19**, 4922.
- J. Kuziola, M. Magre, N. Nöthling and J. Cornell, *Organometallics*, 2022, **41**, 1754.
- D. B. Sowerby, *J. Chem. Soc., Dalton Trans.*, 1983, 5.
- S. M. Godfrey, C. A. McAuliffe, R. G. Pritchard and J. M. Sheffield, *Chem. Commun.*, 1996, 2521.
- K. Nikitin, E. V. Jennings, S. Al Sulaimi, Y. Ortin and D. G. Gilheany, *Angew. Chem.*, 2018, **130**, 1496.
- L. Kapuśnian, P. N. Plessow, D. Trzybiński, K. Woźniak, P. Hofmann and P. I. Jolly, *Organometallics*, 2021, **40**, 693.
- S. Pal, L. Hong, R. V. M. Freire, S. Farooq, S. Salentinig and A. F. M. Kilbinger, *Macromolecules*, 2023, **56**, 7984.
- K. Akiba, K. Okada and K. Ohkata, *Tetrahedron Lett.*, 1986, **27**, 5221.
- O. Coughlin, T. Krämer and S. L. Benjamin, *Organometallics*, 2023, **42**, 339.
- J. Kuziola Dissertation: Synthesis and Characterization of Mono- and Bimetallic Organobismuth(v) Compounds, Ruhr-Universität Bochum, Max-Planck Institut für Kohlenforschung, 2024.
- L. Yang, D. R. Powell and R. P. Houser, *Dalton Trans.*, 2007, 955.
- (a) S. Grimme and P. R. Schreiner, *Angew. Chem., Int. Ed.*, 2011, **50**, 12639; (b) J. P. Wagner and P. R. Schreiner, *Angew. Chem., Int. Ed.*, 2015, **54**, 12274; (c) S. Rösel, J. Becker, W. D. Allen and P. R. Schreiner, *J. Am. Chem. Soc.*, 2018, **140**, 14421; (d) E. Solel, M. Ruth and P. R. Schreiner, *J. Am. Chem. Soc.*, 2021, **143**, 20837; (e) J. M. Schümann, L. Ochmann, J. Becker, A. Altun, I. Harden, G. Bistoni and P. R. Schreiner, *J. Am. Chem. Soc.*, 2023, **145**, 2093.

