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Bowl-shaped, nitrogen-doped polycyclic aromatic hydrocarbons were examined for the first time as molecular receptors for the recognition of metal cations. Potentiometric and spectrofluorimetric assays, supported with density functional theory computations, revealed that the title compounds recognise metal cations with a special focus on caesium (Cs^+) cations.

Bowl-shaped compounds are an intriguing group of molecules, whose synthesis, properties, and functions have been of significant interest in recent years. The curvature in the structure of such compounds influences their unique properties and functions, commonly superior in comparison to their flat congeners.^{1–3} The curvature not only influences the improved photophysical properties of bowl-shaped molecules, but also enables the creation of interesting molecular arrangements and supramolecular systems. The nonplanar polycyclic aromatic hydrocarbons (PAHs)³ constitute a unique class of buckybowl arrangements^{4,5} accompanied by pronounced redox and optical properties modulated by heteroatom dopant(s), such as mainly nitrogen (N) but also sulphur,⁶ phosphorus, selenium,⁷ and tellurium^{8,9} atoms, installed at both their central or peripheral parts. Hence, a heteropine analogue of azacorannulene was demonstrated to exhibit unique p-type semiconducting characteristics with high charge carrier mobility for use in single-crystal field-effect transistors.^{5b} While considering the possible utility of a hetera-buckybowl scaffold as an organic light-emitting diode emitter, our group has introduced a new paradigm for bowl-shaped dibenzoazepine-based PAHs attributed to electron-accepting phenazine species, demonstrating distinctive thermally-activated delayed fluorescence (TADF) properties.¹⁰ Building on that, subsequent derivatization with

Metal cations recognition by bowl-shaped *N*-pyrrolic polycyclic aromatic hydrocarbons[†]

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twisted electron-rich moieties towards D–A–D emitters was accomplished.¹¹ This led to a significant performance leap, achieving a PLQY as high as 96% and an EQE of 21.9% for orange OLED devices.

On the other hand, electron-rich and π -expanded architectures were anticipated to show remarkable binding features, forming supramolecular complexes between electron-deficient C_{60} and twizzer-like azacorannulene.¹² Recently, the possibility of applying the C_3 -symmetric sumanene buckybowl in the construction of caesium (Cs^+) cation selective sensors was demonstrated.^{13,14} The importance of the bowl shape of sumanene towards providing Cs^+ recognition ability was elucidated. Moreover, tellurium- and phosphorus-doped sumanenes turned out to be potent coordinating partners for silver(I) cations.¹⁵ Seeking detection possibilities towards more challenging cations, it could be reasoned that our parent bipolar N-doped PAH could be a perfectly suited system for fluorescent probing of metal cations, not only because of the well-defined emissive properties but even more importantly because of the affinity of the pyramidal centrally positioned N-atom encircled by π -electron units to attract a specific cation. In the light of these considerations, herein we report on the first application of bowl-shaped N-doped PAHs as molecular receptors of metal cations. Their binding abilities were examined using potentiometric methods, spectrofluorimetry, and theoretical calculations (DFT) showing Cs^+ selectivity, together with satisfactory apparent binding constant (K_{app}) and limit of detection (LOD) values for the formed systems at the level of 10^5 M^{-1} and 10^{-6} M , respectively.

The structures of bowl-shaped and N-doped PAHs **1–4** selected for our receptor studies are presented in Fig. 1. Among our library of such derivatives,¹⁰ molecules **1–4** were exclusively chosen due to their satisfactory solubility profiles and different electronic nature of the substituents at the outer benzene ring.

The metal-cation recognition properties of the bowl-shaped *N*-pyrrolic PAHs were first studied using the potentiometric method. Derivatives **1** and **2**, possessing electron-donating (methyl groups) or electron-withdrawing (halogen) substituents

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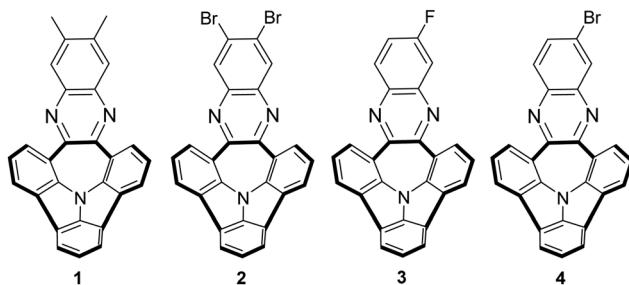


Fig. 1 Structures of investigated N-doped PAHs.

in the outer benzene rings, have been selected as model receptors and introduced into polymeric membranes of ion-selective electrodes.¹⁶

The ion binding ability of these receptors was evaluated on the basis of potentiometric selectivity coefficients¹⁷ (see Fig. 2 and values of selectivity coefficients in Table S1, ESI†), expressing quantitatively the influence of a given cation on the signal of a potentiometric sensor. It should be emphasized that the values of the selectivity coefficients were determined in a heterogeneous system (polymeric membrane/water solution), and therefore did not strictly reflect the supramolecular interactions taking place in solution. Moreover, an anion-exchanger – potassium tetrakis[3,5-bis(trifluoromethyl)phenyl]borate (KTFPB) – has been added to PVC membranes to provide proper operation of the sensors containing receptors **1** or **2**, assumed to function as neutral carriers.¹⁸ The presence of anionic additive KTFPB favours ion-selectivity of the membranes towards more lipophilic cations, *i.e.*, characterized by less negative values of standard molar Gibbs free energy of hydration, like for example caesium (Cs^+) or tetramethylammonium cations.¹⁹

Despite the presence of KTFPB, the examined membranes exhibited increased affinity towards rubidium (Rb^+) and especially Cs^+ cations with respect to other metals ions, even in comparison with the so-called blank membranes, containing only the ion-exchanger. For instance, the membranes based on

receptor **1** were 350 times more selective for Cs^+ against sodium cations (Na^+) and at least 1000 times more selective against lithium cations (Li^+) as well as divalent metal ions. Although a similar selectivity pattern was observed for membranes based on both receptors, the derivative **1** provided higher Cs^+ selectivity (*i.e.*, a higher value of $\log K_{\text{Na},\text{Cs}}$ was determined). According to our expectations, the membranes exhibited significantly worse selectivity towards transition metal cations (negative values of $\log K_{\text{Na},\text{X}}$, comparable with those calculated for the blank membrane), which can be explained by their high hydrophilicity, *i.e.*, very low value of molar Gibbs free energy of hydration (compare for example: $-1915 \text{ kJ mol}^{-1}$ and -250 kJ mol^{-1} for cobalt (Co^{2+}) and Cs^+ ions, respectively¹⁹).

The tested ion-selective electrodes were selective but also sensitive towards Cs^+ cations, enabling their fast and selective quantitative analysis in environmental samples. Potentiometric responses of the sensors based on membranes doped with receptor **1** or **2**, determined in 0.01 M NaNO_3 solution, exhibited a wide linear range with a detection limit of 25 μM Cs^+ (refer to Fig. S1, ESI†). Comparable sensitivity, *i.e.*, slopes of calibration curves close to 50 mV decade⁻¹, as well as short response times in the range of 120–180 s were noticed for both sensors.

Supramolecular interactions between the metal cations and investigated bowl-shaped N-doped PAHs were further examined in solution with fluorescence spectroscopy. Since exclusive Cs^+ recognition abilities were observed during potentiometric experiments, the interactions between Cs^+ and receptor **1**, featuring the best solubility profile among the compounds tested, were first investigated. Due to the insolubility of the tested receptors in water, the addition of organic solvent was necessary to enable running spectroscopic investigations at reasonable receptor concentration. To our delight, receptor **1** was found to be soluble in the $\text{THF}:\text{H}_2\text{O}$ 1:1 *vol/vol* solvent system, which was selected for spectrofluorimetric assays. Cs^+ cations were introduced to the solution in the form of the corresponding hexafluorophosphate (PF_6^-) salt.

Fig. 3 shows the fluorescence spectra results of titration of receptor **1** with Cs^+ (refer to ESI† Section S3 for all the spectra and data on fluorescence spectra assays). Upon the addition of further portions (molar equivalents) of Cs^+ , turn-on fluorescence behaviour was observed. This feature was ascribed to the dynamic formation of cation– π systems,^{20,21} revealing the recognition of Cs^+ by the tested bowl-shaped PAH also in solution. The stoichiometry of the formed complexes between **1** and Cs^+ , analysed with Job's plot method,²² was equal to 1:1 (see plots for **1** in Section S3, ESI†). The K_{app} value for the system, estimated with the Benesi–Hildebrand method,²³ was on the level of 10^4 M^{-1} , while the LOD value was on the micromolar concentration level (Table 1). Thus, satisfactory binding parameters for the formed systems can be concluded. Respective Cs^+ recognition studies were also performed for bowl-shaped receptors **3–4** (receptor **2** was excluded from fluorescence spectroscopy studies due to solubility limitations). As compared with receptor **1**, interactions between Cs^+ and receptors **3–4** were characterised by *ca.* 10-fold higher binding constants (see data in Table 1). This result might suggest that

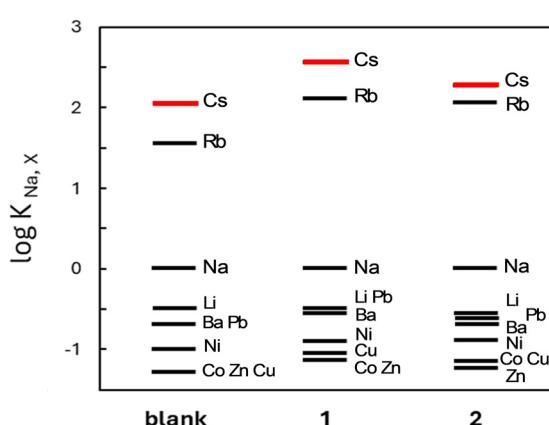


Fig. 2 Graphical presentation of the selectivity of PVC/o-NPOE membranes formulated with receptor **1** or **2** (50 mol% KTFPB) and without receptor (blank membrane); mean values were calculated for 3 electrode specimens.

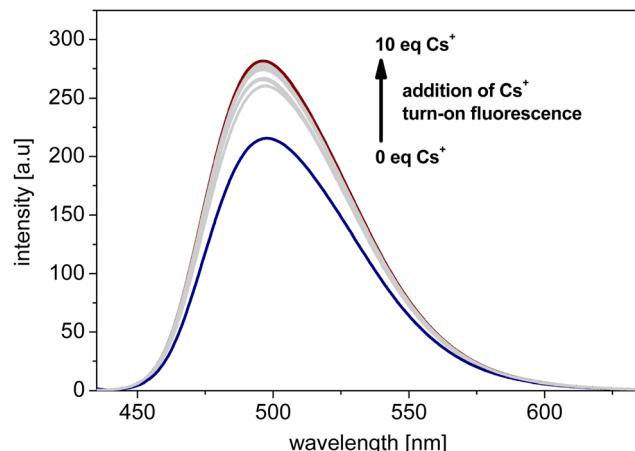


Fig. 3 Fluorescence Cs^+ titration of representative bowl-shaped and N-doped PAH receptor **1** ($\lambda_{\text{ex}} = 320 \text{ nm}$).

Table 1 Summary of binding parameters for non-covalent systems comprising Cs^+ and receptors **1**, **3** or **4**

Receptor	$K_{\text{app}} [\text{M}^{-1}]$	LOD [μM]
1	3.8×10^4	3.7
3	3.0×10^5	4.8
4	2.9×10^5	23.6

the presence of electron-withdrawing groups in the outer phenyl rings of the tested N-doped PAHs should have a beneficial effect on K_{app} . The LOD for the system containing **3** was similar to that of **1**, whereas recognition of Cs^+ by **4** was characterised by *ca.* 10-fold higher LOD.

Building on the above-described procedure, receptor **1** was also titrated with Na^+ , Co^{2+} or Rb^+ cations, as a representative small s-block cation, d-block cation and cation featuring similar van der Waals radius to Cs^+ , respectively. Receptor **1** featured the

property of recognising these metal cations in solution with similar turn-on fluorescence behaviour as in Cs^+ titration assay (refer to ESI,† Section S3 for these spectra). The K_{app} values for the systems were slightly higher than for Cs^+ , equalling $7.9 \times 10^4 \text{ M}^{-1}$, $6.6 \times 10^4 \text{ M}^{-1}$ and $5.4 \times 10^4 \text{ M}^{-1}$, for Na^+ , Co^{2+} and Rb^+ , respectively. The LOD values for Rb^+ ($1.9 \mu\text{M}$) and Co^{2+} ($3.4 \mu\text{M}$) were slightly lower (slightly higher for Na^+ – $9.1 \mu\text{M}$) in comparison to the respective value for Cs^+ ($3.7 \mu\text{M}$). While enhanced selectivity of receptor **1** towards recognising Rb^+ was also observed in potentiometric experiments, we believe that Na^+ and Co^{2+} binding observed during fluorescence assays resulted from the differences between the used techniques. In potentiometric experiments, the water-immiscible phase was doped with the receptor and formed a heterogeneous system contacting the solution, limiting the system dynamics in the light of the recognition process. On the other hand, analyses in solution (fluorescence titration assays) provided higher system dynamics which enabled the formation of complexes between **1** and such cations as small s-block cation Na^+ and d-block cation Co^{2+} .

To further study the investigated non-covalent systems, DFT calculations with pseudopotential plane wave (PSPW)²⁴ were conducted (simulation cell was 100 \AA). The computational works were performed with molecule **1** as a representative N-doped PAH receptor. Complexes containing one Cs^+ cation per receptor molecule (1:1 complex stoichiometry) were subjected to calculation experiments, according to the above-discussed spectrofluorimetric results. Eight different complex arrangements were subjected to DFT calculations. These included Cs^+ concave or convex bound to different structural parts of the receptor, namely six- or seven-membered rings, mostly N-doped ones, see Fig. 4. One considered arrangement (labelled as **1_B**) included Cs^+ bound to the non-bowl-shaped part of the receptor, while other arrangements (labelled as **1_A**, **1_C**, **1_D**) assumed binding Cs^+ by bowl-shaped parts of the receptor.

Fig. 4 presents the DFT calculated interaction energies for each complex. The interaction energies for all considered

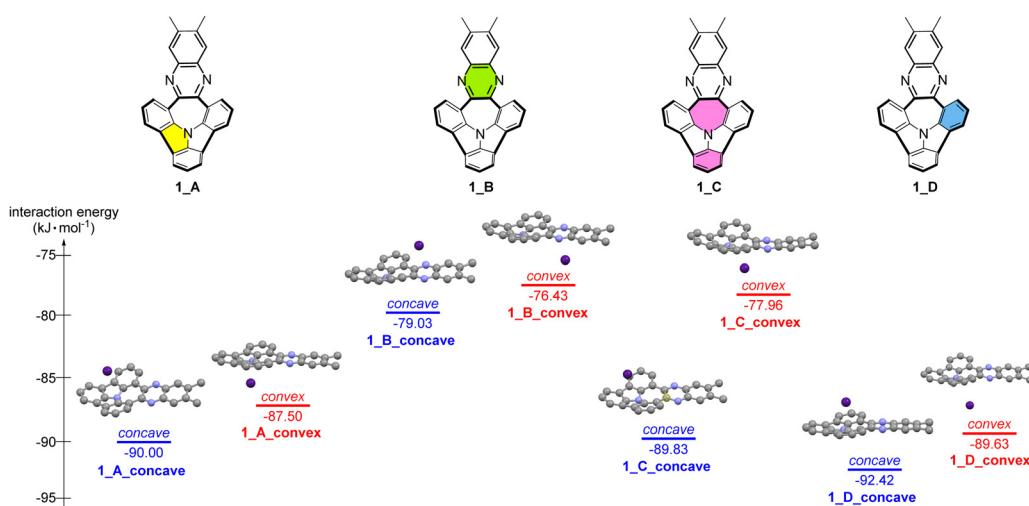


Fig. 4 Graphical presentation of the interaction energies for the systems comprising Cs^+ and **1**, together with the computed complexes' arrangements (hydrogen atoms are omitted for the clarity of the images). Considered parts of receptor **1** included in the Cs^+ binding process are marked with colours.



systems featured negative values, confirming the formation of energetically favourable complexes. According to our computational results, binding of Cs^+ at the bowl parts of the receptor (arrangements **1_A**, **1_C**, **1_D**) was energetically more favoured than binding at the flat part (arrangement **1_B**). This conclusion was supported by further DFT calculations on electrostatic surface potential (ESP), which indicated a non-uniform charge distribution within the receptor molecule (see Fig. S24, ESI[†]). It emphasized the importance of curvature in the considered N-doped PAHs towards providing metal cation recognition features. Furthermore, the formation of both concave- but also convex-bound 1 : 1 complexes was expected, since in most cases no significant differences in the concave- *versus* convex-oriented interaction energies of Cs^+ binding were found, with the exception of one arrangement **1_C** (refer also to the calculated geometry of the HOMOs, Fig. S25, ESI[†]). Interestingly, our DFT analyses suggest that the most energetically favourable Cs^+ complex might be the concave-bound system including the six-membered ring located between N-doped rings of the considered PAH (**1_D** **concave**, interaction energy $-92.42 \text{ kJ mol}^{-1}$).

In conclusion, this work revealed that bowl-shaped and N-doped PAHs constitute a new class of molecular receptors dedicated to the recognition of metal cations. According to our potentiometric experiments, a preference towards Cs^+ binding could be concluded, whereas spectrofluorimetric analyses in solution pointed out satisfactory K_{app} and LOD values for the formed systems at the level of 10^5 M^{-1} and 10^{-6} M , respectively. The importance of the bowl shape towards providing Cs^+ recognition properties of the studied receptors was supported with DFT calculations. We believe that this work opens up new avenues in yet-unexplored applications of bowl-shaped PAHs.

M. L., W. W. and A. K. conceived the project, whereas W. W. and A. K. supervised the experiments; D. S. performed most potentiometric and spectrofluorimetric analyses; J. S. C. performed part of the spectrofluorimetric analyses; J. A. performed DFT calculations; J. W. synthesised the compounds; A. K. conceptualized the manuscript and corresponded with the Editor and Reviewers; M. L., W. W. and A. K. contributed to writing of the manuscript, including preparation of the revised version.

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Data availability

The data supporting this article have been included as part of the ESI.[†]

Conflicts of interest

There are no conflicts to declare.

Notes and references

- 1 M. Saito, H. Shinokubo and H. Sakurai, *Mater. Chem. Front.*, 2018, **2**, 635–661.
- 2 E. Nestoros and M. C. Stuparu, *Chem. Commun.*, 2018, **54**, 6503–6519.
- 3 W.-S. Wong and M. Stępień, *Trends Chem.*, 2022, **4**, 573–576.
- 4 A. Borissov, Y. K. Maurya, L. Moshniaha, W.-S. Wong, M. Żyła-Karwowska and M. Stępień, *Chem. Rev.*, 2022, **122**, 565–788.
- 5 (a) W. Wang and X. Shao, *Org. Biomol. Chem.*, 2021, **19**, 101–122; (b) W. Wang, F. Hanindita, Y. Tanaka, K. Ochiai, H. Sato, Y. Li and T. Yasuda, *Angew. Chem., Int. Ed.*, 2023, **62**, e202218176.
- 6 K. Imamura, K. Takimiya, T. Otsubo and Y. Aso, *Chem. Commun.*, 1999, 1859–1860.
- 7 X. Li, Y. Zhu, J. Shao, B. Wang, S. Zhang, Y. Shao, X. Jin, X. Yao, R. Fang and X. Shao, *Angew. Chem., Int. Ed.*, 2014, **53**, 535–538.
- 8 S. Wang, J. Shang, C. Yan, W. Wang, C. Yuan, H.-L. Zhang and X. Shao, *Org. Chem. Front.*, 2019, **6**, 263–272.
- 9 W. Wang, F. Hanindita, Y. Tanaka, K. Ochiai, H. Sato, Y. Li, T. Yasuda and S. Ito, *Angew. Chem., Int. Ed.*, 2023, **62**, e202218176.
- 10 J. Wagner, P. Zimmermann Crocomo, M. A. Kochman, A. Kubas, P. Data and M. Lindner, *Angew. Chem., Int. Ed.*, 2022, **61**, e202202232.
- 11 J. Wagner, D. Kumar, M. A. Kochman, T. Gryber, M. Grzelak, A. Kubas, P. Data and M. Lindner, *ACS Appl. Mater. Interfaces*, 2023, **15**, 37728–37740.
- 12 Y. Tokimaru, S. Ito and K. Nozaki, *Angew. Chem., Int. Ed.*, 2018, **130**, 9966–9970.
- 13 J. S. Cyniak, L. Kocobolska, N. Bojdecka, A. Gajda-Walczak, A. Kowalczyk, B. Wagner, A. M. Nowicka, H. Sakurai and A. Kasprzak, *Dalton Trans.*, 2023, **52**, 3137–3147.
- 14 J. Ażgiń, M. Wesoły, K. Durka, H. Sakurai, W. Wróblewski and A. Kasprzak, *Dalton Trans.*, 2024, **53**, 2964–2972.
- 15 S. Wang, C. Yan, J. Shang, W. Wang, C. Yuan, H. Zhang and X. Shao, *Angew. Chem., Int. Ed.*, 2019, **58**, 3819–3823.
- 16 Membranes contained: 1% wt receptor 1–2, 65% wt plasticizer (o-NPOE) and 33% wt PVC. Assuming the 1:1 stoichiometry of the formed complexes, 50% mol vs receptor of KTFPB was added to the membranes to gain their highest selectivity.
- 17 According to the Nikolski-Eisenman equation, the values of the selectivity coefficients ($\log K_{\text{Na},\text{X}}$) express quantitatively the influence of a given interfering ion X (in our case $\text{X} = \text{Cs}^+, \text{Rb}^+, \text{Li}^+, \text{Ba}^{2+}, \text{Pb}^{2+}$ and transition metal cations) on the potential of the ion-selective electrode sensitive towards arbitrary chosen primary ion (in our case: Na^+); the higher the value of the selectivity coefficient the higher selectivity towards given interfering cation X.
- 18 P. C. Meier, W. E. Morf, M. Läubli and W. Simon, *Anal. Chim. Acta*, 1984, **156**, 1–8.
- 19 Y. Marcus, *J. Chem. Soc., Faraday Trans.*, 1991, **87**, 2995–2999.
- 20 A. Kasprzak, *Angew. Chem., Int. Ed.*, 2024, **63**, e202318437.
- 21 K. Sikligar, S. P. Kelley, G. A. Baker and J. L. Atwood, *Cryst. Growth Des.*, 2022, **22**, 2806–2811.
- 22 J. S. Renny, L. L. Tomasevich, E. H. Tallmadge and D. B. Collum, *Angew. Chem., Int. Ed.*, 2013, **52**, 11998–12013.
- 23 H. A. Benesi and J. H. Hildebrand, *J. Am. Chem. Soc.*, 1949, **71**, 2703–2707.
- 24 E. J. Bylaska, *Annu. Rep. Comput. Chem.*, 2017, **13**, 185–228.

