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# Cubane-1,3-dicarboxamides as structural isosteres for isophthalamides in hydrogen bond templated interlocked molecules†

Sean R. Barlow,<sup>a</sup> Nathan R. Halcovitch,<sup>a</sup> Geoffrey R. Akien,<sup>a</sup>  
Susannah C. Coote<sup>b</sup> and Nicholas H. Evans<sup>b\*</sup>

**The synthesis and characterization of the first examples of cubane containing interlocked molecules are reported. Catenanes and rotaxanes have been prepared by hydrogen bond templation with cubane-1,3-dicarboxamides replacing isophthalamide motifs.**

Due to the geometric correlation between a cubane and a benzene ring, there has been considerable interest amongst organic and medicinal chemists into whether cubanes may act as bioisosteres for benzene rings in biologically active compounds.<sup>1</sup> With the notable exception of a cubane analogue of a 1,4-benzene-dicarboxylic acid containing metal organic framework,<sup>2</sup> cubanes have yet to be incorporated into self-assembled supramolecular structures—despite the potential to exploit the three-dimensional vectors of the cubic structure. Very recently, MacMillan<sup>3</sup> and Coote<sup>4</sup> have disclosed syntheses of cubane-1,3-dicarboxylic acid derivatives, which may be considered structural isosteres of isophthalic acid derivatives (Fig. 1). In the field of mechanically interlocked molecules (MIMs), following upon the work of Hunter,<sup>5</sup> Vögtle,<sup>6</sup> Leigh<sup>7</sup> and Philp,<sup>8</sup> Evans and co-workers have prepared hydrogen bond templated catenanes and rotaxanes that include isophthalamide macrocycles.<sup>9,10</sup> Here we report that it is possible to synthesize and spectrally characterize hydrogen bond templated MIMs where 1,3-substituted cubanes are replacing isophthalic groups. These are the first reported examples of cubane containing catenanes and rotaxanes.

We chose to investigate a range of hydrogen bond templated catenane and rotaxane syntheses as reported by Leigh<sup>7a,c</sup> and Evans<sup>9c,g</sup> (Scheme 1). Our initial target was to prepare a cubane analogue of Leigh's serendipitous [2]catenane (Scheme 1a).<sup>7a</sup> To allow for confident yield comparisons, Leigh's original synthesis was repeated; in our hands we obtained the previously reported

all-isophthalamide [2]catenane in 17% yield (20% lit. yield, see ESI† pp. S6 and S36). To prepare the cubane analogue, solutions of cubane-1,3-diacyl chloride **1**<sup>11</sup> and *p*-xylylenediamine in dry CHCl<sub>3</sub> were simultaneously added dropwise to a solution of Et<sub>3</sub>N in CHCl<sub>3</sub> over three hours and then the reaction was stirred for a further 16 hours. After work-up, rather than being able to isolate a pure sample of the target [2]catenane **2**, a crude mixture was obtained, which by analytical silica TLC and <sup>1</sup>H NMR analysis indicated the presence of at least two cubane containing species. Silica gel column chromatography and preparative TLC allowed for separation of an analytical sample (isolated 4% yield) of [2]catenane **2** for spectral characterization. In addition to detection of the molecular ion peak by electrospray mass spectrometry (see ESI† p. S21), isolation of [2]catenane **2** (rather than non-interlocked macrocycle) is evidenced by the <sup>1</sup>H NMR spectrum (Fig. 2). In particular, the splitting of aromatic protons **8** (as well as protons **1** and **6**) indicates the presence of one macrocyclic ring passing through the cavity of its interlocked partner.

The second target was to prepare a cubane analogue of Leigh's furamide [2]rotaxane (Scheme 1b).<sup>7c</sup> As for the catenane, Leigh's original synthesis was repeated; in our hands we synthesized the all-isophthalamide [2]rotaxane in 75% yield (calculated from the <sup>1</sup>H NMR spectrum of the purified rotaxane contaminated with free axle, *c.f.* lit. yield: 97%). To prepare the cubane analogue **4**, solutions of cubane-1,3-diacyl chloride **1** and *p*-xylylenediamine in dry CHCl<sub>3</sub> were simultaneously added dropwise to a solution of axle **3** and Et<sub>3</sub>N in a 9 : 1 mixture of dry

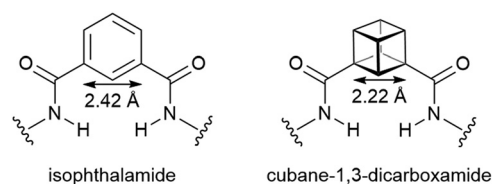


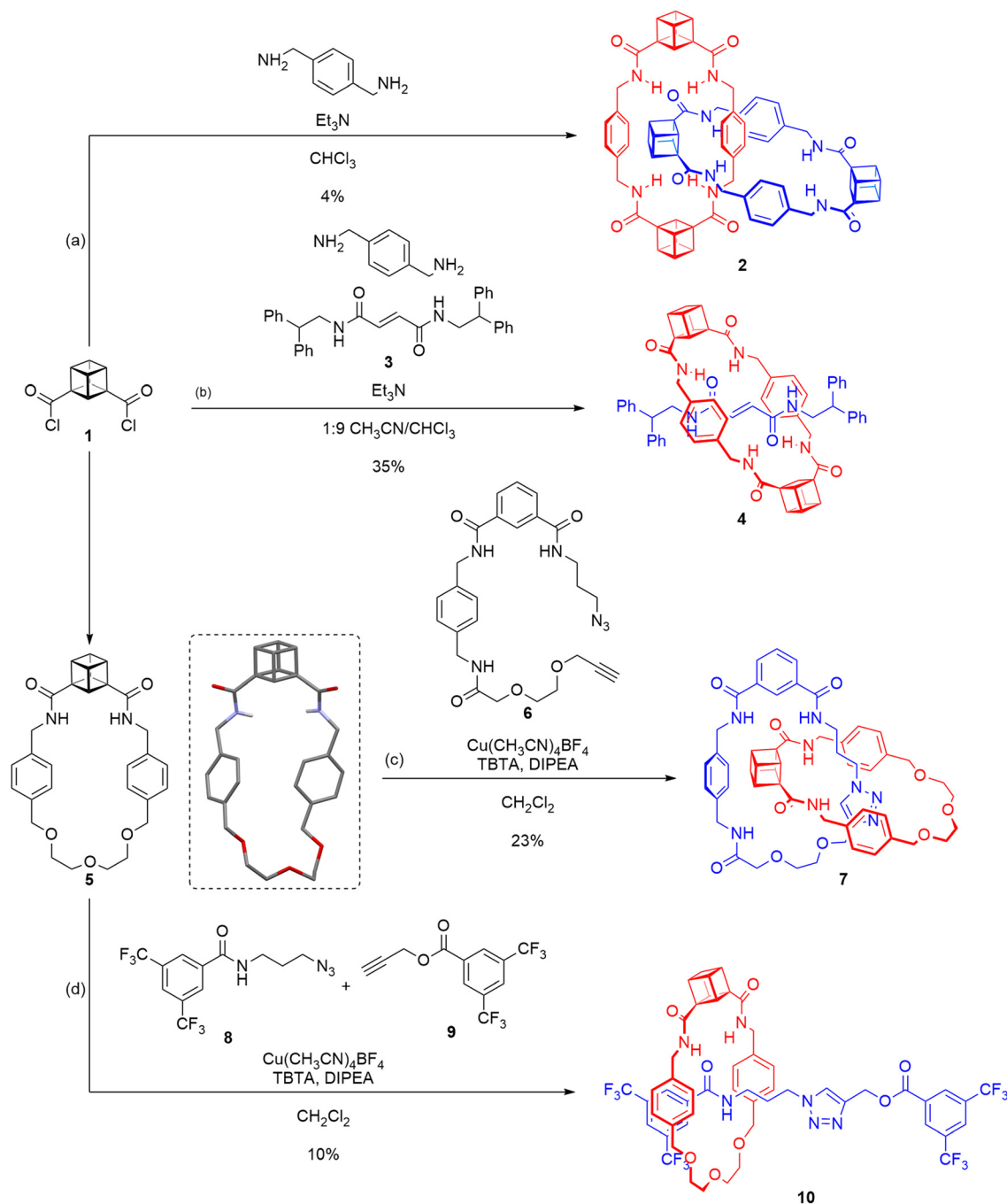
Fig. 1 Illustration of isosteric isophthalamide and cubane-1,3-dicarboxamide.

<sup>a</sup> Department of Chemistry, Lancaster University, Lancaster, LA1 4YB, UK.  
E-mail: n.h.evans@lancaster.ac.uk

<sup>b</sup> Department of Chemistry, University of Bath, Claverton Down, Bath, BA2 7AY, UK

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**Scheme 1** Synthesis of cubane containing interlocked molecules. Inset: X-ray structure of macrocycle **5**.

$\text{CHCl}_3/\text{CH}_3\text{CN}$  and then the reaction was stirred for a further 16 hours. After aqueous work-up, the crude material was then subjected to careful silica gel column chromatography, which allowed for the isolation of the target [2]rotaxane **4** in 35% yield.

Evidence for rotaxane formation is provided by comparison of the  $^1\text{H}$  NMR spectra of axle **3** and [2]rotaxane **4** (see ESI† p. S37); for example, there is a dramatic upfield shift in the  $\text{sp}^2$  C–H furamide axle protons consistent with shielding due to the presence of the macrocyclic ring.<sup>12</sup> Irrefutable demonstration of interlocked structure formation was secured by the growth of

single crystals of sufficient quality for X-ray diffraction from chloroform/methanol solution. The solved crystal structure (Fig. 3) reveals that the cubane-1,3-dicarboxamides are *syn-anti* with hydrogen bonding of the axle amide oxygen atoms to a cubane carboxamide N–H, while the oxygen atom of the other cubane carboxamides are hydrogen bonding to the axle amide N–Hs. This contrasts with Leigh's report of the all-isophthalamide analogue, where all four N–Hs of the *syn-syn* isophthalamides hydrogen bond to axle amide oxygen atoms.<sup>7c</sup> This variation in conformation is attributed to the differences

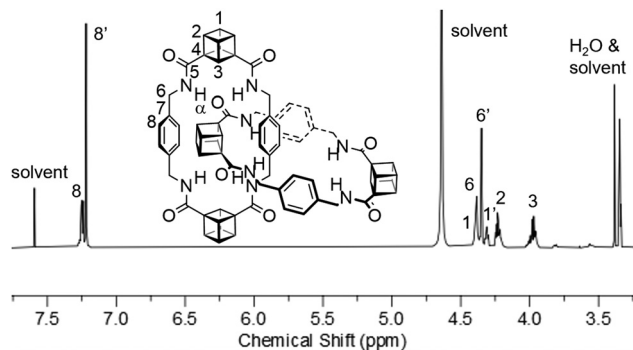


Fig. 2  $^1\text{H}$  NMR spectrum of [2]catenane **2** (1 : 1  $\text{CDCl}_3/\text{CD}_3\text{OD}$ , 400 MHz, 298 K).

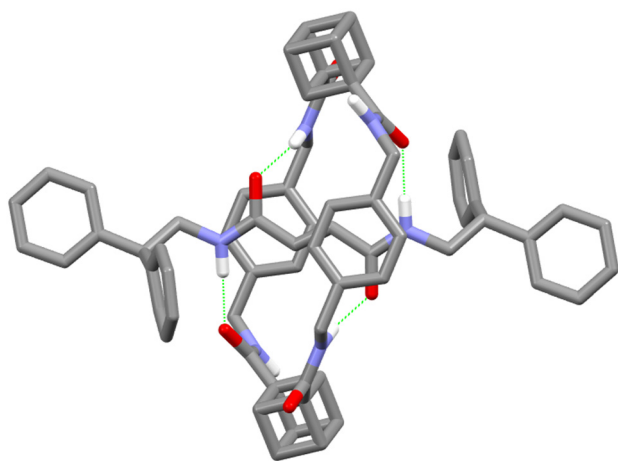


Fig. 3 X-ray crystal structure of [2]rotaxane **4** (C–H bonds omitted for clarity).

in strength of the possible inter-component hydrogen bonds in the [2]rotaxanes.

Our attention then turned to studying the use of cubane-1,3-dicarboxamides in examples of Evans style hydrogen bond templated synthesis. We first sought to trial novel macrocycle **5**<sup>13</sup> in the synthesis of a [2]catenane using our recently published methodology, specifically seeking to generate the cubane containing analogue of the highest yielding reported example (afforded in a yield of 70%).<sup>9g</sup> Following the published methodology, cubane macrocycle **5** and alkyne-azide **6**<sup>9g</sup> were dissolved in dry  $\text{CH}_2\text{Cl}_2$ , and stirred for 15 minutes to allow for the formation of the pseudorotaxane complex. Then,  $\text{Cu}(\text{CH}_3\text{CN})_4\text{BF}_4$ , TBTA and DIPEA were added, and the reaction stirred for 16 hours under an inert atmosphere. After aqueous work-up, the crude material was purified by silica gel column chromatography and preparative TLC to afford the cubane containing [2]catenane **7** in 23% yield.

The formation of cubane containing [2]catenane **7** is evident by comparison of the  $^1\text{H}$  NMR spectra of alkyne-azide **6**, [2]catenane **7** and macrocycle **5** (Fig. 4). For example, the upfield shifts of both sets of aromatic protons *h*/*g* and *j*/*k* are consistent with intercalation of aromatic rings within the interlocked molecule. In addition, cubane proton environments **1** and **3** split as the two faces of cubane macrocycle **5** become inequivalent upon formation of the mechanical bond, due to the second macrocycle being rotationally asymmetric.<sup>12</sup>

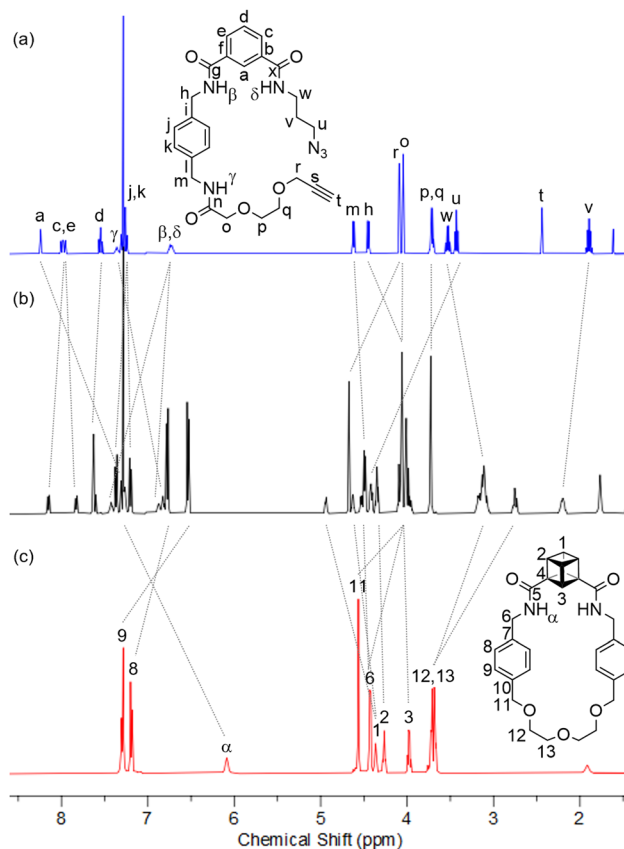


Fig. 4  $^1\text{H}$  NMR spectra of (a) alkyne-azide **6**, (b) [2]catenane **7** and (c) macrocycle **5** ( $\text{CDCl}_3$ , 400 MHz, 298 K).

Finally we targeted the synthesis of an Evans style [2]rotaxane, specifically a cubane analogue of a single amide containing axle [2]rotaxane (reported 47% yield).<sup>9c</sup> Once again following the published methodology, a pseudorotaxane was formed by stirring a solution of cubane macrocycle **5** and azide **8**<sup>9c</sup> in dry  $\text{CH}_2\text{Cl}_2$  for 15 minutes. Then, alkyne **9**,<sup>9c</sup>  $\text{Cu}(\text{CH}_3\text{CN})_4\text{BF}_4$ , TBTA and DIPEA were added, and the reaction stirred for 16 hours under an inert atmosphere. The reaction was then submitted to aqueous work-up and the crude material purified by column chromatography and preparative TLC to afford the cubane containing [2]rotaxane **10** in 10% yield.<sup>12</sup>

In all cases, the yields were lower for the cubane-containing MIMs in comparison with their all-isophthalamide analogues.<sup>14,15</sup> This is tentatively attributed to a combination of (i) the amides being less optimally arranged spatially to support the hydrogen bonding interactions required to create the necessary self-assembled complexes prior to covalent capture of the interlocked structure and (ii) the cubane-1,3-dicarboxamide N–Hs being less acidic than the isophthalamide N–Hs.<sup>16</sup>

Evidence for weaker intercomponent hydrogen bonds in at least one of the cubane-containing MIMs is provided by NMR spectroscopy. At room temperature, the all-isophthalamide analogue of the Evans style cubane [2]catenane **7** has a broad  $^1\text{H}$  NMR spectrum in  $\text{CDCl}_3$  indicating ring rotation occurs at a rate between that of the fast and slow exchange regimes on the NMR timescale. In comparison, the cubane-containing [2]catenane **7** has a sharp



$^1\text{H}$  NMR spectrum in  $\text{CDCl}_3$  at room temperature, consistent with ring rotation being in the fast exchange regime and weaker inter-ring hydrogen bonding (see ESI† pp. S40 and S41).

While generated in modest yields, the successful preparation of cubane containing MIMs is nevertheless exciting. It has been demonstrated that Coote's synthesis<sup>4</sup> does allow for production of sufficient quantities of 1,3-disubstituted cubanes for useful synthetic studies. Aided by further work optimizing the synthesis of cubane-containing MIMs, we envisage incorporating cubanes into more sophisticated MIMs for host-guest<sup>17</sup> or catalytic<sup>18</sup> applications that exploit the three-dimensional vector potential of the cubane structure to augment the inherent three-dimensional architectures of MIMs. In addition, an ability to alter the hydrogen bonding interactions between interlocked components by substituting isophthalamide(s) for cubane dicarboxamide(s), provides supramolecular chemists with another strategy to adjust the dynamic co-conformational behaviour of future MIM-based molecular machines.<sup>19</sup>

NHE and SCC proposed the study. SRB and SCC conducted the synthesis and characterization of materials (with contributions from NRH, GRA and NHE). NHE supervised SRB. SRB, SCC and NHE wrote the manuscript. All authors discussed and commented on the manuscript.

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## Data availability

Underlying data for this paper are provided in the ESI†. Electronic copies of NMR spectra (including fid files) will be available upon publication from: DOI: <https://doi.org/10.17635/lancaster/researchdata/662> and DOI: <https://doi.org/10.17635/lancaster/researchdata/682>. Crystallographic data for **4** and **5** have been deposited at CCDC (<https://ccdc.cam.ac.uk>) under structure numbers 2335982 and 2335981.

## Conflicts of interest

There are no conflicts to declare.

## Notes and references

- (a) J. Wlochal, R. D. M. Davies and J. Burton, *Org. Lett.*, 2014, **16**, 4094–4097; (b) B. A. Chalmers, H. Xing, S. Houston, C. Clark, S. Ghassabian, A. Kuo, B. Cao, A. Reitsma, C.-E. P. Murray, J. E. Stok, G. M. Boyle, C. J. Pierce, S. W. Littler, D. A. Winkler, P. V. Bernhardt, C. Pasay, J. J. De Voss, J. McCarthy, P. G. Parsons, G. H. Walter, M. T. Smith, H. M. Cooper, S. K. Nilsson, J. Tsanaktisidis, G. P. Savage and C. M. Williams, *Angew. Chem., Int. Ed.*, 2016, **55**, 3580–3585; (c) T. A. Reekie, C. M. Williams, L. M. Redina and M. Kassiou, *J. Med. Chem.*, 2019, **62**, 1078–1095.
- L. K. Macreadie, E. J. Mensforth, R. Babarao, K. Konstas, S. G. Telfer, C. M. Doherty, J. Tsanaktisidis, S. R. Batten and M. R. Hill, *J. Am. Chem. Soc.*, 2019, **141**, 3828–3832.
- M. P. Wiesenfeldt, J. A. Rossi-Ashton, I. B. Perry, J. Diesel, O. L. Garry, F. Bertels, S. C. Coote, X. Ma, C. S. Yeung, D. J. Bennett and D. W. C. MacMillan, *Nature*, 2023, **618**, 513–518.
- N. Kazi, M. C. Aublette, S. L. Allinson and S. C. Coote, *Chem. Commun.*, 2023, **59**, 7971–7973.
- C. A. Hunter, *J. Am. Chem. Soc.*, 1992, **114**, 5303–5311.
- (a) F. Vögtle, S. Meier and R. Hoss, *Angew. Chem., Int. Ed. Engl.*, 1992, **31**, 1619–1622; (b) F. Vögtle, M. Händel, S. Meier, S. Ottens-Hildebrandt, F. Ott and T. Schmidt, *Liebigs Ann.*, 1995, 739–743.
- (a) A. G. Johnston, D. A. Leigh, R. J. Pritchard and M. D. Deegan, *Angew. Chem., Int. Ed. Engl.*, 1995, **34**, 1209–1212; (b) D. A. Leigh, A. Murphy, J. P. Smart and A. M. Z. Slawin, *Angew. Chem., Int. Ed. Engl.*, 1997, **36**, 728–732; (c) F. G. Gatti, D. A. Leigh, S. A. Nepogodiev, A. M. Z. Slawin, S. J. Teat and J. K. Y. Wong, *J. Am. Chem. Soc.*, 2001, **123**, 5983–5989; (d) A. Altieri, G. Bottari, F. Dehez, D. A. Leigh, J. K. Y. Wong and F. Zerbetto, *Angew. Chem., Int. Ed.*, 2003, **42**, 2296–2300.
- (a) A. Vidonne and D. Philp, *Tetrahedron*, 2008, **64**, 8464–8475; (b) N. I. Hassan, V. del Amo, E. Calder and D. Philp, *Org. Lett.*, 2011, **13**, 458–461; (c) T. Kosikova, N. I. Hassan, D. B. Cordes, A. M. Z. Slawin and D. Philp, *J. Am. Chem. Soc.*, 2015, **137**, 16074–16083; (d) A. Vidonne, T. Kosikova and D. Philp, *Chem. Sci.*, 2016, **7**, 2592–2603.
- (a) C. N. Marrs and N. H. Evans, *Org. Biomol. Chem.*, 2015, **13**, 11021–11025; (b) N. H. Evans, C. E. Gell and M. J. G. Peach, *Org. Biomol. Chem.*, 2016, **14**, 7972–7981; (c) B. E. Fletcher, M. J. G. Peach and N. H. Evans, *Org. Biomol. Chem.*, 2017, **15**, 2797–2803; (d) N. H. Evans and G. R. Akien, *Supramol. Chem.*, 2018, **30**, 758–764; (e) C. E. Gell, T. A. Mcardle-Ismagulov and N. H. Evans, *Chem. Commun.*, 2019, **55**, 1576–1579; (f) M. J. Young, G. R. Akien and N. H. Evans, *Org. Biomol. Chem.*, 2020, **18**, 5203–5209; (g) S. R. Barlow, G. R. Akien and N. H. Evans, *Org. Biomol. Chem.*, 2023, **21**, 402–414; (h) R. L. Spicer, C. C. Shearman and N. H. Evans, *Chem. – Eur. J.*, 2023, **29**, e2022030502; (i) S. R. Barlow, N. R. Halcovitch and N. H. Evans, *Org. Biomol. Chem.*, 2024, **22**, 3001–3008.
- Lewis has also recently reported upon similar MIMs: (a) J. E. M. Lewis, *Org. Biomol. Chem.*, 2019, **17**, 2442–2447; (b) N. Hoyas Pérez, P. S. Sherin, V. Posligua, J. L. Greenfield, M. J. Fuchter, K. E. Jelfs, M. K. Kuimova and J. E. M. Lewis, *Chem. Sci.*, 2022, **13**, 11368–11375.
- Successful activation of cubane-1,3-dicarboxylic acid by use of oxalyl chloride was confirmed by preparation of a simple model system ESI-1 see ESI† pp. S5, S16 and S17.
- In addition to comparison of  $^1\text{H}$  NMR spectra, spectral evidence of interlocked structure formation is provided by numerous intercomponent through-space correlations in the  $^1\text{H}$ - $^1\text{H}$  ROESY NMR spectrum and detection of the molecular ion peak by electrospray mass spectroscopy (see ESI†).
- The synthesis and characterization of macrocycle **5** is included in ESI† pp. S10, S26, S27, S42 & S43. The crystal structure of **5** (omitting C–H bonds for clarity) is depicted in Scheme 1.
- Similar trends in yield have been observed by Martínez-Cuezva and Berna in preparation of adamantane-containing analogues of furamides and succinimide rotaxanes: (a) A. Martínez-Cuezva, F. Morales, G. R. Marley, A. Lopez-Lopez, J. C. Martínez-Costa, D. Bautista, M. Alajarin and J. Berna, *Eur. J. Org. Chem.*, 2019, 3480–3488; (b) A. Martínez-Cuezva, A. Pastor, M. Marin-Luna, C. Diaz-Marin, D. Bautista, M. Alajarin and J. Berna, *Chem. Sci.*, 2021, **12**, 747–756; (c) J. de Maria Perez, M. Alajarin, A. Martínez-Cuezva and J. Berna, *Org. Chem. Front.*, 2022, **9**, 2690–2696.
- Highlighted by one of the reviewers, a recent report of modifying previously reported MIM synthesis by motif substitution: A. Sarwa, A. Białońska, M. Sobieraj, J. P. Martínez, B. Trzaskowski and B. Szyszko, *Angew. Chem., Int. Ed.*, 2024, **62**, e202316489.
- We believe the lower yields for the cubane-containing MIMs cannot be attributed solely to losses in purification, even though most notably for the Leigh style [2]catenane 2-challenging chromatographic separation may play a role.
- J. T. Wilmore and P. D. Beer, *Adv. Mater.*, 2024, 2309098.
- R. W. Heard, J. M. Suárez and S. M. Goldup, *Nat. Rev. Chem.*, 2022, **6**, 182–196.
- (a) J.-P. Sauvage, *Angew. Chem., Int. Ed.*, 2017, **56**, 11080–11093; (b) J. F. Stoddart, *Angew. Chem., Int. Ed.*, 2017, **56**, 11094–11125.

