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The urgent recognition of phosphate resource scarcity and pollution

Thibaut L. M. Martinon 

DOI: 10.1039/d3su90040a

rsc.li/rscsus

To quote Isaac Asimov (1974),¹ “We may be able to substitute nuclear power for coal, and plastics for wood, and yeast for meat, and friendliness for isolation—but for phosphorus, there is neither substitute nor replacement”. In fact, phosphorus only accounts for 0.14% of the Earth’s crust, one of the rarest elements essential to organic life and what is considered “life’s bottleneck”. The chronic and inefficient use of nonrenewable synthetic fertilizers in the last century has led to vast amounts of phosphate eroding into surface-level water bodies all over the planet.² The consequences manifest in a multifaceted threat to human society,³ where the most basic necessities of food and water hang in the balance.

An often-overlooked pollutant, the signs of phosphate contamination appear after a build-up of nutrient pollution fixed by cyanobacteria; these species will reproduce uncontrollably in a “bloom”, their decay resulting in hypoxic dead zones composed of excessive cyanotoxins. This eutrophication renders freshwater sources inaccessible and devastates the local biota’s biodiversity. The uncontrolled cycle of growth and death repeats until the excess nutrient is removed from the environment. A prominent example of this effect can be witnessed in Lake Erie, of which the western basin experienced blooms of *Microcystis* in 2014, which actively compromised the drinking water of half a million people in Toledo, Ohio.⁴ While this nutrient pollution was originally discovered in 1995, mitigation strategies and cooperation agreements between the United States and Canada have proved ineffective in

preventing the re-eutrophication of this vital source of fresh water.⁵

The irony of the growing abundance of polluting phosphates is that sustained population growth has inadvertently led to the compounding dependence on scarce, nonrenewable phosphate rock following the Green Revolution of the 1960s.⁶ This resource is primarily monopolized between a dozen countries, with Morocco and occupied Western Sahara holding 70% of the world’s overland reserves (Fig. 1).⁷ Although scientists generally agree humans have over 50–100 years before synthetic phosphate reserves are depleted,⁸ the geopolitical ramifications of monopolized export of such a critical resource are likely to shape worldwide political/economic instability in the coming decades in a manner akin to the 1973 Organization of Petroleum Exporting Countries (OPEC) oil embargo.⁹ Some premonitions regarding

Department of Chemistry, University of Minnesota Twin-Cities, Minneapolis, MN 55455, USA. E-mail: mart5691@umn.edu



Thibaut Martinon originates from France and now lives in the United States. He received his MSc from the University of Minnesota in 2022 and is currently pursuing a PhD in supramolecular chemistry. His work includes a focus on the achieving allosteric control of supramolecular anion reception and establishing new selective lanthanide receptors for essential oxyanions such as bicarbonate and orthophosphate in water as part of the Pierre Group. Following his PhD, he would like to expand the supramolecular approaches to lanthanide separation to mitigate dependence on environmentally disastrous methods such as industrial liquid–liquid extraction required to obtain sufficient lanthanide purity.



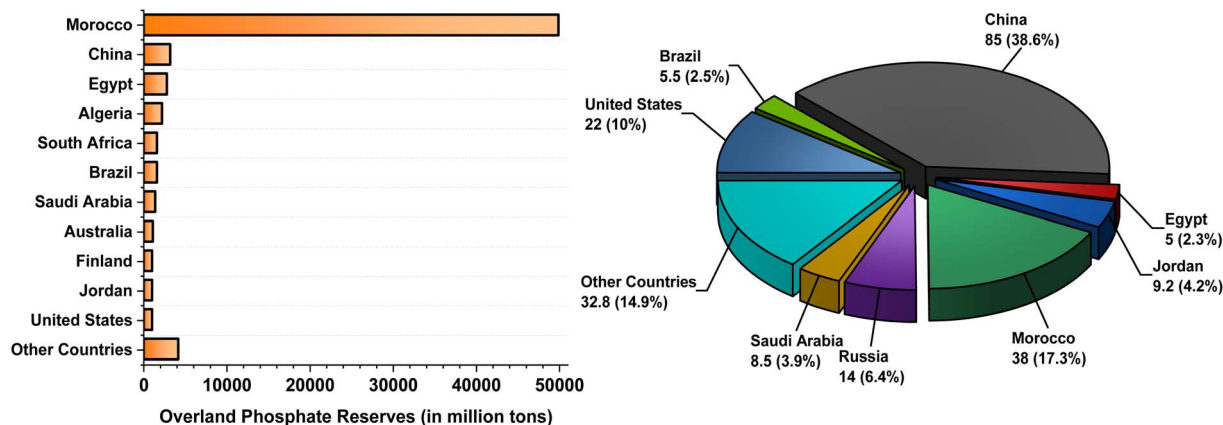


Fig. 1 2021 total world phosphate rock reserves (left) and yearly production (right) by country in million tons. Data obtained from Jasinski⁷ and figures constructed with Origin 2021.

this crisis have already manifested in surging fertilizer prices and countries such as China prioritizing domestic supply over exports critically desired by non-phosphate-producing countries.¹⁰

As with all climate disasters, developing countries will bear the effect most heavily. As demonstrated by a 2007 study sponsored by the International Fertilizer and Development Center (IFDC), prices for fertilizers mount quickly due to compounding transport costs, leaving many impoverished farmers without access to crucial synthetic fertilizers.¹¹ As phosphate prices continue to fluctuate, food security in Sub-Saharan Africa is expected to decrease should inter-state cooperation collapse in the scramble to address mounting scarcity.¹² Recent events in Sri Lanka also reveal a foreboding omen for countries already dependent on synthetic fertilizers. The sudden transition from synthetic fertilizers resulted in varying catastrophes, including an estimated 33% reduction in rice and a 35% reduction in tea yields. These approximate damages alone totaled \$425 million USD, contributing to the toppling of the ruling government within a year.¹³

Developed countries that manage these costs will instead find themselves footing an ever-increasing bill for pollution treatment and fertilizer costs. The United States public and private sectors alone have spent \$1.9 trillion USD as of 2014 on abating surface-water pollution since the 1960s.¹⁴ These costs would accompany increased mass emigration as food insecurity displaces growing population centers, requiring major shifts in

asylum laws along already congested borders.¹⁵ While one could not be blamed for assuming mitigating this disaster is inevitable, some progress has been made in the past decades. The Indian state of Sikkim, for example, was able to adopt an organic fertilizer economy through a gradual transition from 2003–2014, strongly supported *via* major public funding and education programs.¹⁶ Each region of the planet will need to adopt new strategies to tackle the unique aspects of its phosphate footprint, and the chemical sciences are excellently placed in leading these innovations.

The chemical sciences have been at the forefront of revolutionary innovations, as noted by Asimov. These same innovations have also equally been the architects of the crises that haunt our fragile planet. We have the duty to develop new strategies for mitigating the impending crisis by directing public policy. Our first step should be at a local level. Farmers are keenly aware of their contribution to this disaster and have access to many modern tools for collecting studies in cooperation with the scientific community, from sensors to monitor soil conditions to drones that analyze field conditions from above.¹⁷ The United States Department of Agriculture (USDA), for example, organizes critical Natural Resources Conservation Service (NRCS) conservation practices in tandem with farmers, saving \$927 million USD's worth of anhydrous ammonia in fields annually and, in some cases, reducing agricultural phosphate run-off by 75 percent.¹⁸ Taking advantage

of these tools to mitigate phosphate run-off and overuse of fertilizers could be a key step in stopping phosphate contamination at its primary source.

We should also prioritize messaging around these economic impacts, because this provides a huge incentive for change on a political and individual level. Education is also crucial, spreading awareness of safeguarding water resources and food waste prevention to local community leaders with the trust and best interests of their communities. My own lab group at the University of Minnesota has led outreach efforts in our community to educate young students about phosphate pollution, and we strongly encourage our colleagues to do the same to combat the declining interest and trust in science.¹⁹

Immediate actions can also be taken to address nutrient pollution, particularly achieving international recognition in removing phosphates from detergents and personal care products. The role of detergents in nutrient pollution is well documented,²⁰ and many countries have pursued legislation to limit or ban their use. Companies such as Proctor and Gamble, which composed 25% of the North American market, transitioned to phosphate-free alternatives in 2014.²¹ While this change will continue organically as phosphate prices increase, banning phosphate detergents prematurely should be pursued globally. In 2007, detergents alone made up 61% of all surface water pollution from industrial uses in Western Europe.²²



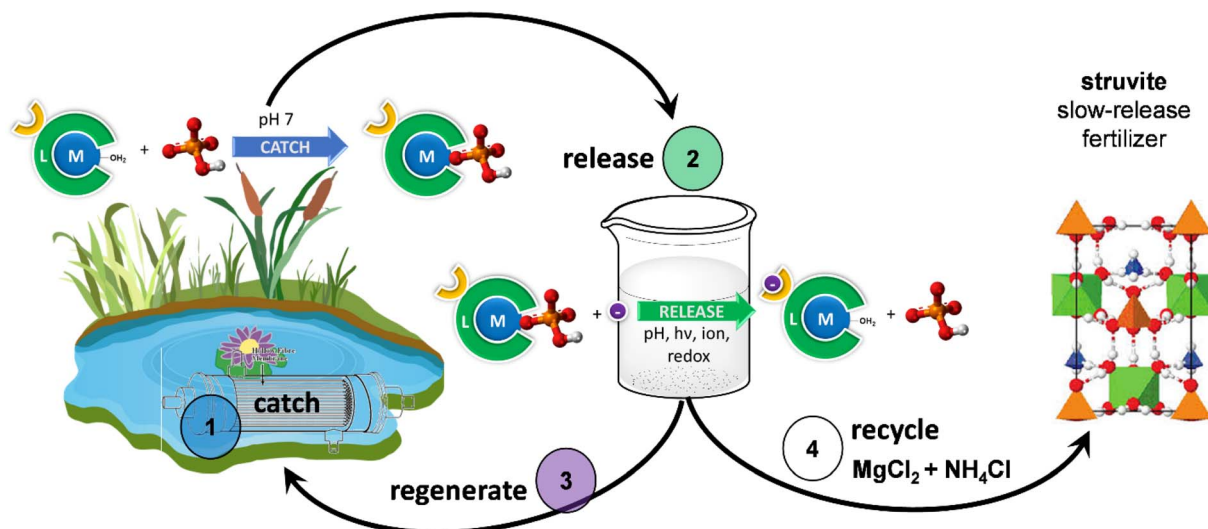


Fig. 2 A proposed cycle of catch-and-release with regenerable lanthanide receptors to remediate polluting phosphate into struvite fertilizer. Affinity for phosphate is controlled via reversible instillation of peripheral negative charge on the receptor. Image reproduced courtesy of Dr. Valérie Pierre.

Of course, activism will not be able to solve this issue alone. The re-eutrophication of Lake Erie serves as a dire warning that simple prevention methods alone will not suffice.⁵ Farmers need to be given viable alternatives to common fertilizers; this will come in the synthesis of cheaper, long-release fertilizers that will cushion the transition to a sustainable farming economy. Developing novel agents such as struvite has shown favorable advantages regarding slow-release properties, reduced heavy-metal contamination, and a salt index ideal for plant growth.²³ The low water solubility, however, has raised questions about struvite's agronomic utility,²⁴ and hence further exploration into other slow-release fertilizers with either mineral or organic polymer coatings may be necessary.²⁵

For phosphate already contaminating surface-level water bodies, several recent approaches have been studied in close context in the past decades, regarding phosphate removal and recovery.^{6,26} Phosphate removal is principally performed via chemical precipitation of bioavailable phosphate with cations of either calcium, iron, or aluminum. While great in the short term in preventing eutrophication witnessed in China,²⁷ there are concerns the phosphate salts could contaminate waterbeds and lead to long-term elevated phosphate/cation

levels. Phosphate recovery instead solves both the scarcity and pollution issues at once, and is principally utilized in the treatment of phosphate-contaminated wastewater. Enhanced biological phosphorous removal of activated sludge with polyphosphate-accumulating organisms is one of the primary methods to refine phosphate into an aqueous state.⁶ Phosphates harvested in this manner are unfortunately well-known to bioaccumulate heavy metals through stabilization,²⁸ unideal for phosphate capture and leaving over 60% of recovered phosphate to languish in landfills as heavy-metal-contaminated waste.² An answer is likely found within supramolecular phosphate receptors and membrane technologies.²⁹ Lanthanide-based receptors have proved to be of particular interest due to their unique properties in forming labile, selective bonds with phosphates, surpassing the hydration energy of phosphate in water (Fig. 2).^{30–32}

These innovations will require vertical integration to facilitate novel discoveries; lanthanides used in phosphate binders could be more readily accessible. However, limitations in the purification process have led to higher prices, despite them being more abundant than traditional metals such as gold.³³ While this sprawling field is already gaining traction, due to the increased requirement of rare earth elements (REEs) in electronics,

novel lanthanide ligands³⁴ and separation methods³⁵ should be explored to incur fewer environmental costs and increase accessibility. It should also be noted that some alternatives to solve climate change should also be re-evaluated, particularly biofuels. First-generation biofuels require heavy uses of phosphate-based fertilizers, and their impact in depleting inorganic phosphorous (Pi) sources outweighs the benefit of combating climate change.³⁶ Our fight against climate change should not jeopardize the security of the most basic resources required for human life. Neither should we idle as the 50th anniversary of Asimov's warning comes into view, and should continue to build a path forward for the coming generations.

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