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Thermo-switchable hydrophobic deep eutectic solvent for CuAAC^{†‡}

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The use of a deep eutectic solvent with thermo-switchable hydrophobicity properties (lidocaine: decanoic acid) is presented. Upon a gentle thermal trigger, the solvent separated into aqueous and organic layers. This system is well adapted for the CuAAC reaction using an N-heterocyclic carbene copper catalyst. An alkyne–azide cycloaddition of hydrophilic compounds, such as carbohydrates, could be conducted with good yields in the aqueous layer after the thermal separation. The method presented herein causes low contamination of the products by metals and the solvent can be reused at least five times.

Sustainability spotlight

A decrease in the impact of emissions from organic solvents during chemical synthesis is important. Therefore, the substitution of a traditional organic solvent by a renewable one offers a simple solution to improve sustainability. In this context, non-volatile solvents are of great importance, especially if smart properties are needed. We envisioned developing a protocol to improve the efficiency of the extraction phases in which a thermomorphic multiphase system could separate the product from the reaction mixture, including the metals of the catalyst. The unique properties of oleic or decanoic acid/lidocaine mixture toward water rendered them good candidates for such a separation. Remarkably, we achieved a sequence based on reaction, extraction, and separation during a one-pot process for the copper(i)-catalyzed alkyne–azide cycloaddition upon a small thermal trigger. Our work aligns with the following UN sustainable development goals: industry, innovation, and infrastructure (SDG 9) and sustainable consumption and production patterns (SDG 12).

Introduction

Classical organic processes often involve the use of large solvents quantities (reaction, extraction, purification) for fine chemical and pharmaceutical sector applications.¹ The UN has presented a sustainability plan entitled “Transforming Our World: The 2030 Agenda for Sustainable Development” which targets the need for greener and more sustainable chemistry.² Environmental directives and legislation of the European Union focus on reducing solvent emissions and regulating the usage of potentially harmful or environmentally damaging chemical substances.³ When possible, the substitution of a traditional harmful organic solvent by a renewable and benign one offers a simple solution to improve the sustainability of the overall process. Alternative chemical procedures have appeared to decrease the *E*-factor. Among these procedures, mechanochemistry is perfect because it eliminates the usage of solvents, but it does not solve the issue of the purification step.⁴ One issue in using volatile solvents is the eventuality of releasing

some into the atmosphere and in the environment. To solve this issue, it is possible to use ionic liquids that are practically non-volatile.⁵ More recently, deep eutectic solvents (DES) that are formed from mixtures of Brønsted or Lewis acids and bases have emerged as new solvents.⁶ The melting points of DES are lower than their individual components, and they are generally simply prepared by mixing two chemicals to form a liquid mixture. Different types of DES have been described, and most of them contain hydrogen-bond acceptors (HBAs) and hydrogen-bond donors (HBDs), such as choline chloride/urea. Like ionic liquids, they present very high enthalpies of vaporisation, and therefore, could be considered as non-volatile for most applications.^{6b} Indeed, some of the ionic liquids have found applications in catalysis and extraction due to the thermomorphic multiphase systems (TMS) nature of their mixtures with organic solvents.⁷ By changing the temperature, the system can switch from one to two phases opening the possibility to perform a chemical reaction followed by a fast separation of the reacting materials (and the catalysts) upon a thermal trigger. However, most of these systems still contain volatile organic solvents. Recently, we reported the first example of a DES with switchable smart properties when using temperature as a trigger: a mixture of oleic acid/lidocaine (1/1 molar ratio) and water (50 wt%), which is homogeneous at room temperature, spontaneously separates into two phases (water and DES) at temperature higher than its lower critical solution temperature (LCST, 25 °C in this case).⁸ Although DES are mainly used for

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[‡] Indeed, the *pK_a* of lidocaine is ~8 at room temperature, but it decreases rapidly with temperature. As the ionic form of lidocaine is much more soluble than the molecular one, the increase in temperature reduces the solubility of organics in the aqueous phase, until separation of the two liquid phases occurs.



extraction, they have found applications in organic synthesis.⁹ Meyer *et al.* took advantage of this smart property to conduct biocatalysis in a homogeneous aqueous DES phase.^{9f} After the reaction, the temperature increases, and the enzyme stays in the aqueous phase while the product is found in the DES. This separation of the two phases triggered by temperature perfectly fits to avoid the extraction step during the organic synthesis of aqueous soluble compounds.

Therefore, we hypothesised that:

1- The separation of the two phases triggered by temperature perfectly replaces the extraction step during the organic synthesis of aqueous soluble compounds.

2- As DES are recognised to efficiently extract transition metals,¹⁰ we expect that the metal from a catalyst could easily separate from the reaction product as well.

We suspect that linking two partially hydrophilic compounds would produce an easier water-extractable compound, whereas the rest of the starting material would predominantly stay in the DES (Fig. 1). This approach could be of great importance in the synthesis of radiopharmaceuticals bearing metals or pharmaceutical compounds that are obtained through metal catalysis. In this context, the copper alkyne-azide cycloaddition (CuAAC) was chosen as a case study.¹¹ This reaction, catalysed by Cu¹⁺, furnishes exclusively 1,5 disubstituted 1,2,3-triazoles from the [3 + 2] cycloaddition of organic azides with terminal alkynes, while maintaining a perfect atom economy. Indeed, CuAAC is compatible with a broad range of solvents (from organic to aqueous) and accommodates many functional groups. In fact, a few publications reporting the application of CuAAC on different DES can already be found in the literature.¹² Among these, the choline chloride/glycerol (1 : 2) system is well-adapted for hydrophobic compounds that precipitate after the addition of a large volume of water into the reaction vessel. We have already reported the properties of lidocaine/oleic acid or decanoic acid (1/1 mol ratio, **DES I** and **DES II**, respectively) mixtures that separate in two aqueous and organic phases just above room temperature.^{8d} This feature is due to the decrease in ion concentrations with an increase in the temperature ref. 8d. For the CuAAC reaction, a large set of

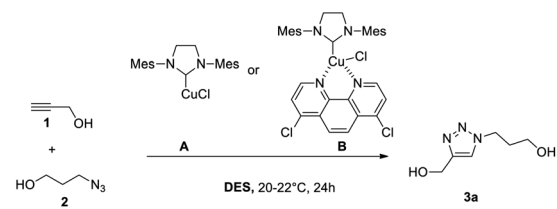
ligands that strongly accelerate the process are available, ranging from highly hydrophilic (*i.e.*, THPTA, bathophenanthroline disulfonic acid) to hydrophobic (*i.e.*, TBTA, 1,10-phenanthroline, *N*-heterocyclic carbenes (NHCs)). Among these, the systems based on NHC ligands are particularly interesting because they do not require the addition of a reducing reagent (*e.g.*, ascorbic acid) to keep copper at its +1 oxidation level. Thus, we selected the well-explored reagents, SIMesCuCl (**A**) and [SIMes(Cl₂-Phen)CuCl] (**B**, generated *in situ* by the addition of 4,7-dichloro-1,10-phenanthroline to **A**) as catalysts.¹³

Results and discussion

We first examined the reaction of propargyl alcohol (**1**) with 1-azido propan-3-ol (**2**) furnishing triazole **3a** (Table 1). Reactions were conducted at room temperature (typically 20 °C) for 24 h with 0.5 M solutions of **1** and **2**. Conversions were measured by ¹H NMR (10–15 μL of the DES/water mixture dissolved in dms_o-d₆) by comparing the triazole signal (8 ppm) and the propyne proton (2.4 ppm).

No reaction occurs between **DES I** (Table 1, entry 1) and catalyst **A**, but a complete conversion takes place after the addition of the *N*-donor (entry 2, catalyst **B**). The addition of an equal volume of water has a deleterious effect on both catalysts (entries 3–4), probably due to the precipitation of the catalyst. However, decreasing the ratio to 2/1 (v/v) allows the reaction to proceed with catalyst **B** (entry 6), whereas the catalyst **A** stays inactive. The same tendency was observed with **DES II**: **A** is

Table 1 DES and catalyst screening



Entry	Solvent ^a	Cat. (mol%)	Conversion ^b
1	DES I (1/0)	A (2.5)	0%
2	DES I (1/0)	B (2.5)	100%
3	DES I (1/1)	A (2.5)	0%
4	DES I (1/1)	B (2.5)	0%
5	DES I (2/1)	A (2.5)	0%
6	DES I (2/1)	B (2.5)	100%
7	DES II (1/0)	A (2.5)	0%
8	DES II (1/0)	B (2.5)	100% (93%) ^c
9	DES II (1/1)	A (2.5)	0%
10	DES II (1/1)	B (2.5)	0%
11	DES II (2/1)	A (2.5)	0%
12	DES II (2/1)	B (2.5)	100%
13	DES II (2/1)	B (1.0)	100%
14	DES II (2/1)	B (0.5)	100%
15	DES II (2/1)	B (0.25)	24%
16	Water	B (2.5)	0%

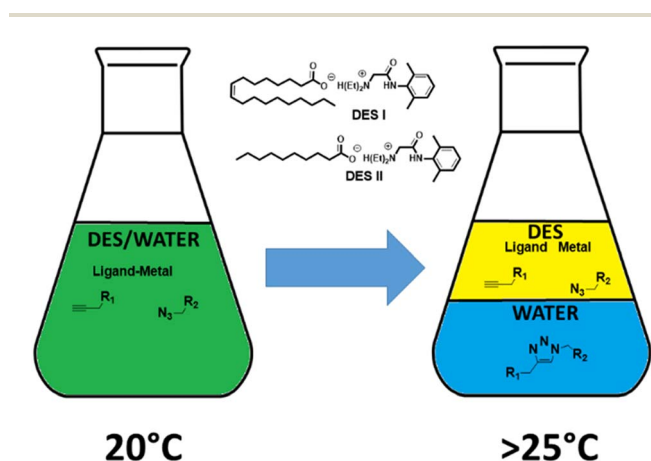


Fig. 1 Scheme of the reaction and separation steps using thermal switchable aqueous DES.

^a DES/water (v/v). ^b Taken by ¹H NMR. ^c Isolated yield.



inactive (entries 7, 9, and 11), which definitively rules out the possibility of using this catalyst, whereas **B** performs well in the pure solvent or in the 2/1 (v/v) mixture (entries 8, 12 *versus* entry 10). The triazole **3a** was extracted well (see general procedures), furnishing the product with 93% yield (entry 8). With the 2/1 volumetric ratio, the loading of **B** could be decreased to 0.5 mol% (entries 13, 14 *versus* 15) without altering the conversion.

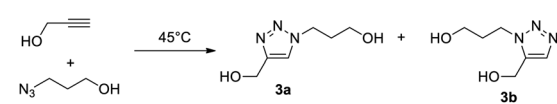
The blank experiment (entry 16) shows that no reaction takes place in pure water. Although it has been reported that Cu-NHCs can catalyse 'on-water' cycloadditions even at ppm-scale loadings,¹⁴ our result can be explained by the total solubility of the starting material in water, which helps avoid the formation of an organic layer on the water in which the catalyst dissolves. Our results outline the necessary presence of the N-donor heterocycle (catalyst **B**) for the reaction to proceed. This is explained by the breaking of the copper-chloride bond to free the active catalyst. Therefore, pure DES (**I** and **II**) do not seem to have strong hydrolytic properties. It is important to mention that the presence of a double bond in DES **I** does not interact with the catalyst. However, it could potentially create troubles with other transition metal catalysis (Pt²⁺, Pd²⁺), therefore, DES **II** was preferred for the rest of this study. To assess that the reaction is catalysed and does not occur spontaneously in the DES, especially during the extraction process during which the temperature increases, the azide and alkyne were mixed at 45 °C for 24 h in different DES **II**/water mixtures (Table 2). Entries 1–3 indicate that uncatalysed cycloaddition occurs to a lesser extent at 45 °C in DES **II**, in water or in the 2/1 mixture (two phases in this case as the reaction temperature is above 25 °C) after 24 h. As expected, a substantial proportion of 1,5 regioisomer **3b** is observed in each reaction. This is in sharp contrast with the high yield and the total selectivity for the 1,4-regioisomer **3a** observed in the presence of catalyst **B** (entry 4).

Inspection of the advancement shows that the reaction progresses slowly in pure DES **II** (24 h), whereas the presence of water strongly increases the reaction rate (1 h 30 min to 2 h – it should be noted that the addition of water to the DES **II** induces a decrease in the viscosity that could be involved in the improvement of the kinetics of the reaction). Interestingly, the strongly exothermic reaction causes an increase in the

temperature (up to 40 °C with 10 mol% of **B**) in DES **II** that causes phase separation (temperature superior to the LCST, 25 °C). This interesting result indicates that the solvent itself is responsive to triggers from the reaction and paves the way to smart solvents that would change their properties in case of run-away reactions (therefore, in the current study, the vials were placed at 20 °C with a water bath). Another concern was the fate of the catalyst **B** after the reaction. Due to the sensitivity of Cu⁺¹ toward oxygen, the reaction usually ends with unliganded Cu²⁺ and the protonated ligand (1,3-dimesitylimidazolium). ICP-OES indicates that the mass contamination of the product by copper represents 0.05 wt% (copper *vs.* product). ¹H quantitative NMR was used to determine the partition of the imidazolium and 4,7-dichloro-1,10-phenanthroline. The compounds stay essentially in the DES **II** layer (>99%). In all cases, after copper catalysis of the test reaction, we observed a small amount of lidocaine in the aqueous phase after extraction. This is consistent with the results reported by Kroon *et al.*¹⁰ which studied the extraction of transition metals (including copper) from water with the hydrophobic mixture of decanoic acid and lidocaine (2/1 molar ratio) forming a biphasic liquid system with water, unlike our (1/1) mixture that separated upon heating. The transition metal stays bound to the organic acid in the DES phase, thereby forming Cu(Deca)₂ and releasing two protonated lidocaine ions in the aqueous layer. We therefore determined the amount of lidocaine in the water (using ¹H q-NMR) and found that it represents twice the amount of catalyst added which agrees well with Kroon's observations. § Before exemplifying our system, we examined the extractability of the compounds in the model reaction (the triazole was synthesised separately, purified, and re-engaged in an extraction protocol). The extraction consists of the dissolution of the compound at 0.5 M in a 2/1 (v/v) DES **II**/water mixture at room temperature (<25 °C), followed by an increase in temperature to 40–45 °C for 20 minutes to separate the two phases. The number of compounds was determined by quantitative ¹H NMR (3-(trimethylsilyl)propane-1-sulfonic acid sodium salt for the aqueous phase and 1,4-dimethyl benzene for the organic layer as internal standards). We found that propargyl alcohol has a strong propensity to stay in the DES **II** layer with a partition ratio of 9/1. Azido propanol splits quasi-equally between the two phases (partition of 6/4). However, the extraction of triazole **3a** occurs more efficiently albeit with the introduction of a hydrophobic heterocycle (partition 8/2). A second extraction (same volume ratio, 1/2 DES/water) affords a total transfer in the aqueous layer. Fig. 2 shows a typical example of the process. As can be seen, the reaction starts with a reddish colour, which is typical of the catalyst **B** (copper-to-ligand charge transfer). At the end of the reaction, oxidation by molecular oxygen furnishes a greenish colour from a copper²⁺ carboxylate. Upon a thermal trigger, a clean-phase separation occurs leaving the copper in the DES **II** layer (top phase).

Therefore, we suppose that the "diol effect" could be applied to other compounds, such as carbohydrates. This protocol was

Table 2 Thermal *versus* catalysed reaction



Entry	Ratio (DES II /water)	Catalyst B ^a	Yield	3a/3b
1	1/0	0	<10% ^c	1/1
2	0/1	0	<10% ^c	2/1
3	2/1 ^b	0	<10% ^c	1/1
4	2/1 ^b	0.7	93% ^d	1/0

^a Mol%. ^b Reaction conducted in a 2-phase system because reaction temperature is superior to the LCST (25 °C). ^c Determined by ¹H qNMR in the crude mixture. ^d Isolated yield.

§ See ESI for lidocaine determination.



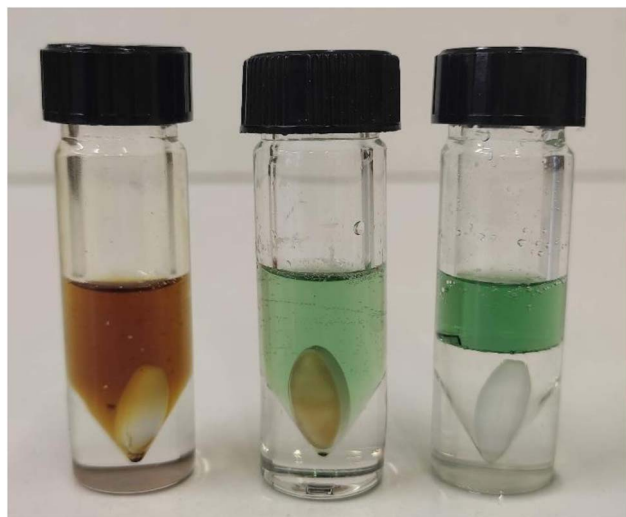


Fig. 2 Reaction process. Left: reaction at $t = 0$, $T = 20$ °C. Middle: end of the reaction, $T = 20$ °C. Right: phase separation immediately after a thermal trigger, $T = 45$ °C.

then applied to exemplify the system; Table 3 summarises the isolated yields and conversions obtained for 11 polyols.

The reactions of propargyl and homopropargyl alcohols with azido-3-propanol (compounds **3a** and **4** entries 1–2) occur with a total conversion; after two aqueous extractions, isolated yields are excellent. Here again, contamination by copper from the aqueous extract was determined by ICP-OES and represented 0.05% of the mass product. A lower yield is observed when switching to pentyne-1-ol (**5**, entry 3), albeit a total conversion is observed. ^1H NMR inspection of the DES layer reveals that a large amount of the product stays inside, highlighting the poor efficiency of the extraction in this case. The reaction of the PEG azidoethoxy-ethoxy-ethanol (**6**, entry 4) allows an averaged isolated yield. ^1H NMR of the crude aqueous extract shows that it is contaminated by 10–20% of the **DES II** mixture, probably because of the amphiphilic properties of the product. Stirring the aqueous extract with *tert*-butyl methyl or *di*-isopropyl ether allows the removal of the DES (the very pale green colour of the ether layer indicates the presence of a trace amount of copper). As an alternative, a simple purification through a C18 column affords the pure compound. The reaction of benzyl azide (**7**, entry 5) affords a low isolated yield although the reaction reaches completion, indicating that several hydrophilic functions are needed to conduct the extraction step efficiently. Therefore, the reactions of highly hydrophilic carbohydrates were explored. The reaction of α -D-galactopyranosyl bearing a propargyl function at the anomeric centre with azido propanol occurs efficiently (**8**, entry 6). Similarly, the reaction with benzyl azide presented total completion (**9**, entry 7, 85% isolated yield), although it was still contaminated by 10–20% of DES. This successful extraction proves that the lipophilic moiety is well counterbalanced by the carbohydrate. Using the same strategy, the reaction of the azido-PEG affords an isolated yield of 84% after further purification (**10**, entry 8, the crude extract is still contaminated by the DES). We also installed the azide function

Table 3 Examples of CuAAC in aqueous **DES II** (1.6 mol% of catalyst)

Entry	Compound	Yield ^a
1	3a 	93% (100%) ^b
2	4 	92% (100%) ^b
3	5 	38% (100%) ^b
4	6 	60% (100%) ^b
5	7 	<10% (100%) ^b
6	8 	90% (100%) ^b
7	9 	85% (100%) ^b
8	10 	84% (100%) ^b
9	11 	86% (100%) ^b
10	12 	0% (0%) ^b
11	13 	85% (100%) ^b

^a Isolated pure compound. ^b Completion.

on the carbohydrate (β -D-galactopyranosyl), which was reacted with propargyl alcohol in 86% isolated yield (**11**, entry 9). Notably, the reaction is totally inhibited for phenylacetylene (entry 10). Finally, the two carbohydrates were efficiently linked affording 85% of isolated yield (**13**, entry 11). Finally, we examined the reusability of the solvent after extraction. ¶ Table 4 indicates the yield of triazole **3a** and copper contamination after each reaction cycle (**DES II** was not purified between the two experiments). We also examined the effects of generating the catalyst by adding **A** and 4,7-dichloro-1,10-phenanthroline (Table 4, upper results); as the phenanthroline additive is left

¶ Alternatively, it has been reported that sodium oxalate could be used to extract the metal from similar DES, see ref. 9.



Table 4 Reusability of the DES

Experiment	1	2	3	4	5
Yield (catalyst) ^a	95% (B)	99% (B)	95% (B)	95% (B)	95% (B)
Copper ^b	0.03%	0.04%	0.02%	0.05%	0.02%
Yield (catalyst) ^a	97% (B)	95% (A)	97% (A)	80% (A)	59% (A)
Copper ^b	0.03%	0.2%	0.07%	0.18%	0.37%

^a Measured by ¹H NMR with 3-trimethylsilyl-1-propanesulfonic acid sodium salt as internal standard. ^b Contamination: mass Cu/total mass.

untouched after the reaction, the addition of only A takes place only after the first run (results at the bottom).

The iterative experiments show that the solvent could be used at least 5 times with catalyst B without a decrease in the yield of the reaction or copper contamination of the product. The second method shows that the addition of 4,7-dichloro-1,10-phenanthroline is necessary as the yield decreases and the metal contamination increases after the third cycle.

The process presented here allows a reaction/extraction sequence in a 'one pot' process with good separation of the metal and the catalyst's ligand; the main contamination occurred due to the presence of lidocaine salt in small amounts (2%). Regarding the *E*-factor, the overall process comprises the DES that could be used 5 times and a small amount of *di*-isopropyl ether. With this smart solvent that switches from hydrophilic to hydrophobic, it is possible to obtain an *E*-factor below 5.

Conclusion

In conclusion, we have reported that the azide-alkyne cycloaddition catalyzed by NHC-Cu^I takes place effectively in the aqueous deep eutectic solvents made of lidocaine and fatty acid. Using the thermal trigger, hydrophilic compounds are extracted efficiently in water, permitting an efficient separation of the product. However, for compounds possessing amphiphilic properties, a substantial amount of DES is produced within the aqueous layer, and an additional extraction step with ether solvent is necessary because of the relatively high concentrations of reactants used. Notably, the DES acts as a metal trap providing products with low copper contamination. The solvent can be reused at least 5 times, still providing a good yield and low metal content. Investigations of the properties of other DES for catalysis are under evaluation in our laboratory.

Experimental section

General information

Commercial grade reagents and solvents (*tert*-butyl and *di*-isopropyl ether, 4,7-dichloro-1,10-phenanthroline) were used as received from Sigma-Aldrich. Lidocaine, oleic, and decanoic acids were purchased from TCI. Mono distilled water was used. Purifications for analytical purposes were performed on a spot II Flash chromatography ARMEN instrument with a C18 Büchi column (40 μm irregular) using H₂O/MeOH/TFA (v/v = 10/0 to 6/

4, 0.1% TFA). The nuclear magnetic resonance (NMR) spectra were recorded with a 400 MHz spectrometer (Bruker Advance III). ¹H NMR and ¹³C{¹H} NMR chemical shifts (δ, relative to tetramethylsilane for ¹H and relative to the solvent for ¹³C{¹H}) are reported in ppm (part per million). Coupling constants (*J*) are reported in hertz (Hz). Infrared spectra were measured using a PerkinElmer (spectrum two, ATR mode) and are reported in cm⁻¹. High-resolution mass spectra (HRMS) were obtained using time-of-flight (TOF) and orbitrap MS. Copper contamination was determined using an ICP-OES Agilent 5800 equipment. 1-Azido-propane-3-ol,¹⁵ 2,3,4,6-tetra-*O*-acetyl-β-*D*-galactopyranosyl azide,¹⁶ 1-*O*-propargyl-*D*-mannose,¹⁷ and SiMe₃CuCl¹⁸ were synthesised from procedures reported in literature.

DES preparation

The two components of the DES (1 to 1 molar ratio) were added together and then heated at 80 °C in a closed vial; this temperature was above the melting temperatures of both HBD and HBA. DES mixtures became perfectly transparent liquid after a few minutes. DES was then left overnight to cool down at room temperature before further use.

General procedure

In a conical reaction vial (surrounded by a water bath at room temperature, typically 20–22 °C) that contained 1.3 mL of DES, 6.5 mg of SiMe₃CuCl (0.016 mmol, 1.6 mol%) and 4.7 mg of 4,7-dichloro-1,10-phenanthroline (0.016 mmol, 1.6 mol%) was introduced. The mixture was heated at 45 °C, while stirring until all the solids were dissolved and a red colour was developed. After cooling to room temperature, a mixture of 101 mg of 1-azido propanol (1 mmol, 1.0 equivalent) and 56 mg of propargyl alcohol (1 mmol, 1.0 equivalent) in 0.7 mL of water was added. After 24 h, the solution was heated to 45 °C for 20 minutes to produce a phase separation (in case of a poor separation, 1 mL of water was added). The lower aqueous layer was separated using a syringe. Two millilitres of water were added and the reaction mixture was cooled to room temperature under stirring until it turned limpid. The liquid-liquid separation was repeated at 45 °C for 20 minutes, and the two aqueous phases were mixed. The rest of DES II (in compounds 3a, 4, 5, and especially 6) was removed from the aqueous layers by extraction with 0.5–0.7 mL of *di*-isopropyl ether. For analytical purposes, the products containing sugars were purified on C18 using 0 to 20% methanol gradient (0.1% TFA).

Quantitative NMR

Trimethylsilyl propane-1-sulfonic acid sodium (1/9 of the expected number of mol) was added to the aqueous extracts and 20% (v/v) of D₂O. An aliquot of the aqueous layer was analysed by quantitative ¹H NMR.¹⁹ Dosage was performed by comparing signals at –0.2 ppm (TMS) and 8.0 ppm (triazole).

2-(2-(2-Azidoethoxy)ethoxy)ethanol. This was synthesised from a modified procedure.²⁰ 2-(2-(2-Chloroethoxy)ethoxy) ethanol (16 g, 100 mmol, 1.0 equivalent) was added to 100 mL of water containing 19.5 g of sodium azide (300 mmol, 3.0



equivalent). The reaction mixture was heated at 65 °C for 48 h, extracted with dichloromethane (3 × 250 mL), dried over MgSO₄, and evaporated to furnish 16.5 g of a pale-yellow liquid (94%). ¹H NMR (D₂O, 400 MHz): δ (ppm) = 3.7–3.3 (m, 8H), 3.55 (t, *J* = 5.0 Hz, 2H), 3.45 (t, *J* = 5.0 Hz, 2H). ¹³C NMR (CDCl₃, 100 MHz) δ (ppm) = 72.1, 70.1, 69.8, 69.4, 61.0, and 50.1.

1-β-D-Galactopyranosyl azide. 2,3,4,6-Tetra-*O*-acetyl-β-D-galactopyranosyl azide was deprotected using the Zemplin procedure: 3.7 g of the chemical was dissolved in 50 mL of dry methanol containing 100 mg of dissolved sodium. The reaction was stirred for 6 h and quenched with amberlist. It was then evaporated to produce a white solid (quantitative yield). ¹H NMR (D₂O, 400 MHz): δ (ppm) = 4.57 (d, *J* = 8.7 Hz, 1H), 3.87 (d, *J* = 3.3 Hz, 1H), 3.72–3.65 (m, 3H), 3.67 (dd, *J* = 9.9, 3.4 Hz, 1H), 3.40 (t, *J* = 8.8 Hz, 1H). ¹³C NMR (dms_o-d₆, 400 MHz): δ (ppm) = 90.5, 77.4, 73.1, 70.1, 68.0, 60.3. IR (Neat) ν (cm⁻¹) = 3308, 3453, 3169, 2890, 2125 (st), 1501, 1548, 1372, 1307, 1050, 891, 775, 706. HRMS (ESI) *m/z*: [M + Na]⁺ calculated for C₆H₁₁O₅N₃Na 228.05909 and found to be 228.0591.

1-O-Propargyl-D-mannose.¹⁷ ¹H NMR (D₂O, 400 MHz): δ (ppm) = 4.95 (s, 1H), 4.27 (dd, *J* = 12 Hz, 3 Hz, 1H), 4.21 (dd, *J* = 12, 3 Hz, 1H), 3.9–3.5 (m, 6H), 2.84 (t, *J* = 3 Hz, 1H). ¹³C NMR (dms_o-d₆, 400 MHz): δ (ppm) = 98.2, 79.8, 77.3, 74.5, 71.0, 70.2, 67.0, 61.2, 53.1. IR (Neat) ν (cm⁻¹) = 3353, 3275, 2936, 2911, 2118, 1446, 1367, 1334, 1217, 1134, 1059, 1003, 963, 908, 916.

3-(4-(Hydroxymethyl)-1H-1,2,3-triazol-1-yl)propan-1-ol, 3a. ¹H NMR (D₂O, 400 MHz): δ (ppm) = 7.94 (s, 1H), 4.66 (s, 2H), 4.46 (t, *J* = 7.0 Hz, 2H), 3.52 (q, *J* = 6.5 Hz, 2H), 2.07 (q, *J* = 6.5 Hz, 2H). ¹³C NMR (D₂O, 100 MHz): δ (ppm) = 146.6, 124.2, 58.1, 54.6, 47.1, 31.8. According to literature.²¹

3-(4-(2-Hydroxyethyl)-1H-1,2,3-triazol-1-yl)propan-1-ol, 4. ¹H NMR (D₂O, 400 MHz): δ (ppm) = 7.74 (s, 1H), 4.38 (t, *J* = 7.0 Hz, 2H), 3.75 (t, *J* = 6.5 Hz, 2H), 4.47 (t, *J* = 6.2 Hz, 2H), 2.83 (t, *J* = 6.5 Hz, 2H), 2.01 (q, *J* = 6.5 Hz, 2H). ¹³C NMR (D₂O, 100 MHz): δ (ppm) = 145.1, 123.9, 60.5, 58.1, 47.0, 31.8, 27.7. IR (Neat) ν (cm⁻¹) = 3307, 2937, 2877, 1340, 1359, 1313, 1217, 1134, 1046, 804. HRMS (ESI) *m/z*: [M + H]⁺ was calculated for C₇H₁₄O₂N₃: 172.10805 and was found to be 172.1080.

3-(4-(Hydroxypropyl)-1H-1,2,3-triazol-1-yl)propan-1-ol, 5. ¹H NMR (D₂O, 400 MHz): δ (ppm) = 7.68 (s, 1H), 4.35 (t, *J* = 7.0 Hz, 2H), 3.50 (t, *J* = 6.6 Hz, 2H), 3.45 (t, *J* = 7.0 Hz, 2H), 2.63 (t, *J* = 7.6 Hz, 2H), 2.00 (q, *J* = 6.0 Hz, 2H), 1.77 (q, *J* = 6.6 Hz, 2H). ¹³C NMR (D₂O, 100 MHz): δ (ppm) = 147.7, 123.2, 60.7, 58.1, 47.0, 31.8, 31.0, 21.0. IR (Neat) ν (cm⁻¹) = 3277, 2920, 2850, 1435, 1217, 1187, 1055, 799. HRMS (ESI) *m/z*: [M + H]⁺ was calculated for C₈H₁₆O₂N₃: 186.1237 and was found to be 186.137.

2-(2-(2-(4-(Hydroxymethyl)-1H-1,2,3-triazol-1-yl)ethoxy)ethoxy)ethanol, 6. ¹H NMR (D₂O, 400 MHz): δ (ppm) = 7.98 (s, 1H), 4.66 (s, 2H), 4.57 (t, *J* = 5.1 Hz, 2H), 3.92 (t, *J* = 5.1 Hz, 2H), 3.7–3.5 (m, 6H), 3.49 (t, *J* = 5.0 Hz). ¹³C NMR (D₂O, 100 MHz): δ (ppm) = 146.8, 124.4, 71.6, 69.6, 69.3, 68.7, 60.3, 54.6, 50. IR (Neat) ν (cm⁻¹) = 3343, 2869, 1455, 1351, 1224, 1112, 1058, 1012, 929, 823. HRMS (ESI) *m/z*: [M + H]⁺ was calculated for C₉H₁₇N₃O₄: 232.1219 and was found to be 232.1290.

6-((1-(3-Hydroxypropyl)-1H-1,2,3-triazol-4-yl))-α-manno pyranose, 8. Analytical sample purified on C18 column to remove 2–3% of lidocaine salt contamination (±2%). ¹H NMR (MeOH-d₄,

400 MHz): δ (ppm) = 8.00 (s, 1H, H_{triaz}), 4.85 (d, *J* = 1.4 Hz, 1H, H_{ano}), 4.80 (d, *J* = 12.5 Hz, 1H, CH₂), 4.65 (d, *J* = 12.5 Hz, 1H, CH₂), 4.50 (t, *J* = 6.9 Hz, 2H, CH₂OH), 3.9–3.5 (m, 6H), 2.10 (q, *J* = 6.8 Hz, 2H, CH₂CH₂OH). ¹³C (MeOH-d₄, 100 MHz): δ (ppm) = 125.5, 100.6, 74.8, 72.3, 71.8, 68.5, 62.8, 60.5, 59.1, 48.3, 33.8 (Cq_{triaz}, Missing). IR (Neat) ν (cm⁻¹) = 3307, 2927, 1674, 1429, 1227, 1130, 1049, 960, 808. HRMS (ESI) *m/z*: [M + H]⁺ was calculated for C₁₂H₂₁N₃O₇: 320.14523 and was found to be 320.1450.

6-((1-(3-Benzyl)-1H-1,2,3-triazol-4-yl))-α-mannopyranose.

1TFA, 9. The analytical sample was purified on a C18 column. ¹H NMR (MeOH-d₄, 400 MHz): δ (ppm) = 8.00 (s, 1H, H_{triaz}), 7.4–7.3 (m, 5H, H_{phenyl}), 5.59 (s, 2H, CH₂-benzyl), 4.83 (d, *J* = 1.4 Hz, 1H, H_{ano}), 4.80 (d, *J* = 12.4 Hz, CH₂), 4.63 (d, *J* = 12.4 Hz, CH₂), 3.9–3.5 (m, 6H). ¹³C (MeOH-d₄, 100 MHz): δ (ppm) = 161.7 (q, *J* = 35 Hz, TFA), 144.2, 135.3, 128.6, 128.3, 127.8, 124.0, 116.5, (q, *J* = 293 Hz, TFA), 99.4, 73.5, 71.0, 70.6, 67.0, 61.2, 59.3, 53.6. (Cq_{triaz}, missing). IR (Neat) ν (cm⁻¹) = 3199, 3039, 2890, 1654 (st), 1435, 1183, 1128, 1057, 1028, 839, 799, and 723. HRMS (ESI) *m/z*: [M + H]⁺ was calculated for C₁₆H₂₂O₆N₃: 352.1503 and was found to be 352.1499.

(1-(2-(2-(2-Hydroxyethoxy)ethoxy)ethyl)-1H-1,2,3-triazol-4-yl))-α-mannopyranose, 1 TFA, 10. The analytical sample was purified on a C18 column to remove 2–3% of lidocaine salt contamination (±2%). ¹H NMR (MeOH-d₄, 400 MHz): δ (ppm) = 8.08 (s, 1H), 4.80 (d, *J* = 12.3 Hz, 1H), 4.62 (d, *J* = 12.3 Hz, 1H), 4.59 (t, *J* = 4.9 Hz, 2H), 3.9 (t, *J* = 5.1 Hz, 2H), 3.85 (dd, *J* = 2.2, 11.2 Hz, 1H), 3.8 (m, 1H), 3.75–3.5 (m, 12H). ¹³C NMR (D₂O, 100 MHz): δ (ppm) = 162.1 (TFA), 142.2, 126.7, 117.6 (TFA), 99.5, 72.9, 71.5, 70.3, 69.8, 69.5, 69.2, 68.3, 66.51, 60.7, 60.2, 59.0, 50.7. IR (Neat) ν (cm⁻¹) = 3349, 2878, 1351, 1228, 1126, 1103, 1054, 883, 812. HRMS (ESI) *m/z*: [M + H]⁺ was calculated for C₁₅H₂₈O₉N₃: 394.1826 and was found to be 394.1817.

6-(4-(Hydroxymethyl)-1H-1,2,3-triazol-1-yl)-β-galactopyranose, 11. The analytical sample was purified on a C18 column to remove 2–3% of lidocaine salt contamination (±2%). ¹H NMR (MeOH-d₄, 400 MHz): δ (ppm) = 8.19 (s, 1H, H_{triaz}), 5.58 (d, *J* = 9.2 Hz, 1H, H_{ano}), 4.71 (s, 2H), 4.15 (t, *J* = 9.29, 1H), 3.98 (d, *J* = 2.8 Hz, 1H), 3.9–3.6 (m, 4H). ¹³C (MeOH-d₄, 100 MHz): δ (ppm) = 90.4, 80.1, 75.4, 71.6, 70.5, 62.5, 565.5 (2C triazole missing). IR (Neat) ν (cm⁻¹) = 3251, 2932, 1672, 1422, 1125, 1089, 1051, 1010, 888. HRMS (ESI) *m/z*: [M + H]⁺ was calculated for C₉H₁₆N₃O₆: 262.1039 and was found to be 262.1030.

((β-D-Galactopyranosyl)-1H-1,2,3-triazol-4-yl) methylene α-D-mannopyranose, 13. The analytical sample was purified on a C18 column to remove 2–3% of lidocaine salt contamination (±2%). ¹H (400 MHz, dms_o-d₆) δ (ppm) = 8.23 (s, 1H), 5.47 (d, *J* = 9.2 Hz, 1H), 4.73 (d, *J* = 1.5 Hz, 1H), 4.70 (d, *J* = 8.3 Hz, 1H), 4.52 (d, *J* = 12 Hz, 1H), 4.02 (t, *J* = 9.3 Hz, 1H), 3.8–3.35 (m, 11H). ¹³C (dms_o-d₆, 100 MHz) δ (ppm) = 143.9, 12.34, 99.3, 88.6, 78.9, 74.7, 74.1, 71.4, 70.6, 69.8, 67.5, 61.8, 60.9, and 59.3. IR (Neat) ν (cm⁻¹) = 3277, 2927, 1672, 1407, 1326, 1128, 1045, 884, 810, and 700. HRMS (ESI) *m/z*: [M + H]⁺ was calculated for C₁₅H₂₆N₃O₁₁: 424.15618 and was found to be 424.1558.

Conflicts of interest

There are no conflicts to declare.



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Notes and references

- D. J. C. Constable, C. Jimenez-Gonzalez and R. K. Henderson, Perspective on Solvent Use in the Pharmaceutical Industry, *Org. Process Res. Dev.*, 2007, **11**(1), 133–137, DOI: [10.1021/op060170h](https://doi.org/10.1021/op060170h).
- United Nation Home Page, 'United Nations General Assembly. Transforming Our World: The 2030 Agenda for Sustainable Development; 2015', <https://sdgs.un.org/2030agenda>, accessed 2023–2024.
- (a) EU Home Page, 'On the Limitation of Emissions of Volatile Organic Compounds Due to the Use of Organic Solvents in Certain Activities and Installations', <https://eur-lex.europa.eu/legal-content/EN/TXT/?uri=celex%3A31999L0013>, accessed 2023–2024; (b) F. Kerton and R. Marriott, Chapter 2. Green Solvents – Legislation and Certification, *RSC Green Chemistry*, The Royal Society of Chemistry, UK, 2013; pp. 31–50.
- J. L. Do and T. Frišćić, *ACS Cent. Sci.*, 2017, **3**(1), 13–19.
- M. J. Earle, J. M. S. S. Esperança, M. A. Gilea, J. N. Canongia Lopes, L. P. N. Rebelo, J. W. Magee, K. R. Seddon and J. A. Widegre, *Nature*, 2006, **439**, 831–834.
- (a) B. B. Hansen, S. Spittle, B. Chen, D. Poe, Y. Zhang, J. M. Klein, A. Horton, L. Adhikari, T. Zelovich, B. W. Doherty, B. Gurkan, E. J. Maginn, A. Ragauskas, M. Dadmun, T. A. Zawodzinski, G. A. Baker, M. E. Tuckerman, R. F. Savinell and J. R. Sangoro, *Chem. Rev.*, 2021, **121**, 1232–1285; (b) K. Xin, I. Roghair, F. Gallucci and M. van Sint Annaland, *J. Mol. Liq.*, 2021, **325**, 115227–115235.
- Y. Qiao, W. Ma, N. Theysen, C. Chen and Z. Hou, *Chem. Rev.*, 2017, **117**, 6881–6928; Y. Kohno and H. Ohno, *Chem. Commun.*, 2012, **48**, 7119–7130; J. Luo, T. Xin and Y. Wang, *New J. Chem.*, 2013, **37**, 269–273.
- (a) A smart solvent can have his properties changed reversibly depending of the application of a trigger, P. J. Jessop, D. J. Heldebrant, X. Li, C. A. Eckert and C. L. Liotta, *Nature*, 2005, **436**, 1102; (b) As example, the addition/removal of CO₂ into some liquid mixtures can switch from hydrophobic to hydrophilic, J. J. Jessop, L. Phan, A. Carrier, S. Robinson, C. J. Dürra and J. R. A Harjania, *Green Chem.*, 2010, **12**, 809–814; (c) Temperature can also be use as trigger, few aqueous amine mixtures have a Lower Critical Solution Temperature (LCST), M. Goral, D. G. Shaw, A. Maczynski, B. Wiśniewska-Gocłowska and P. Oracz, *J. Phys. Chem. Ref. Data*, 2012, **41**, 043107; (d) O. Longeras, A. Gautier, K. Ballerat-Busserolles and J.-M. Andanson, *ACS Sustainable Chem. Eng.*, 2020, **8**(33), 12516–12520; (e) J. Castaneda Corzo, K. Ballerat-Busserolles, J. Y. Coxam, A. Gautier and J. M. Andanson, *J. Mol. Liq.*, 2023, **377**, 121468; (f) L.-E. Meyer, M. B. Andersen and S. Kara, *Angew. Chem., Int. Ed.*, 2022, **61**, e2022038.
- Selected examples: (a) P. H. Tran, H. T. Nguyen, P. E. Hansen and T. E. Le, *RSC Adv.*, 2016, **6**, 37031–37038; (b) R. Martínez, L. Berbegal, G. Guillena and D. J. Ramón, *Green Chem.*, 2016, **18**, 1724–1730; (c) S. A. Padvi and S. P. Dalal, *Synth. Commun.*, 2017, **47**, 779–787; (d) X. Maset, J. M. Pérez and D. J. Ramón, *Green Chem.*, 2016, **18**, 826–833; (e) C. Vidal, J. García-Álvarez, A. Hernán-Gómez, A. R. Kennedy and E. Hevia, *Angew. Chem., Int. Ed.*, 2016, **55**, 16145–16148.
- D. J. G. P. van Osch, D. Parmentier, C. H. J. T. Dietz, A. van den Bruinhorst, R. Tuinier and M. C. Kroon, *Chem. Commun.*, 2016, **52**, 11987–11990.
- (a) C. W. Tornøe, C. Christensen and M. Meldal, *J. Org. Chem.*, 2002, **67**, 3057–3064; (b) H. C. Kolb, M. G. Finn and K. B. Sharpless, *Angew. Chem., Int. Ed.*, 2001, **40**, 2004–2021.
- (a) S. Sethi, N. C. Jana, S. Behera, R. R. Behera and B. Bag, *ACS Omega*, 2023, **8**, 868–878; (b) S. V. Giofrè, M. Tiecco, A. Ferlazzo, R. Romeo, G. Ciancaleoni, R. Germani and D. Iannazzo, *Eur. J. Org. Chem.*, 2021, **34**, 4777–4789; (c) A. Kafle and S. T. Handy, *Tetrahedron*, 2017, **73**, 7024–7029.
- (a) S. Díez-González, A. Correa, L. Cavallo and S. P. Nolan, *Chem.–Eur. J.*, 2006, **12**, 7558–7564; (b) M.-L. Teyssot, A. Chevry, M. Traïkia, M. El-Ghozzi, D. Avignat and A. Gautier, *Chem.–Eur. J.*, 2009, **15**, 6322–6326.
- S. Díez-González and S. P. Nolan, *Angew. Chem., Int. Ed.*, 2008, **47**, 8881–8884.
- S. Kalhor-Monfared, C. Beauvineau, D. Scherman and C. Gurard, *Eur. J. Med. Chem.*, 2016, **122**, 436–441.
- S. Wang, J. A. Cuesta-Seijo, D. Lafont, M. M. Palcic and S. Vidal, *Chem.–Eur. J.*, 2013, **19**, 15346–15357.
- I. Krabicová, B. Dolenský and M. Rezanka, *Molecule*, 2022, **27**, 1483–1493.
- C. Gibard, H. Ibrahim, A. Gautier and F. Cisnetti, *Organometallics*, 2013, **32**(15), 4279–4283.
- (a) G. F. Pauli, S.-N. Chen, C. Simmler, D. C. Lankin, T. Gödecke, B. U. Jaki, J. B. Friesen, J. B. McAlpine and J. G. Napolitano, *J. Med. Chem.*, 2014, **57**, 9220–9231; (b) Y. Nishizaki, D. C. Lankin, S.-N. Chen and G. F. Pauli, *Anal. Chem.*, 2021, **93**, 2733–2741; (c) A. Akhdar, J. M. Andanson, S. Faure, M. Traïkia and A. Gautier, *J. Organomet. Chem.*, 2021, **950**, 121991–212995.
- A. Palmioli, A. Aliprandi, D. Septiadi, M. Mauro, A. Bernardi, L. De Colab and M. Panigati, *Org. Biomol. Chem.*, 2017, **15**, 1686–1699.
- C. Girard, E. Önen, M. Aufort, S. Beauvière, E. Samson and J. Herscovici, *Org. Lett.*, 2006, **8**(8), 1689–1692.

