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Empowering boronic acids as hydroxyl synthons for aryne induced three-component coupling reactions†

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Boronic acids have become one of the most prevalent classes of reagents in modern organic synthesis, displaying various reactivity profiles via C–B bond cleavage. Herein, we describe the utilization of a readily available boronic acid as an efficient surrogate of hydroxide upon activation via fluoride complexation. The hitherto unknown aryne induced ring-opening reaction of cyclic sulfides and three-component coupling of fluoro-azaarenes are developed to exemplify the application value. Different from metal hydroxides or water, this novel hydroxy source displays mild activation conditions, great functionality tolerance and structural tunability, which shall engender a new synthetic paradigm and in a broad context offer new blueprints for organoboron chemistry. Detailed computational studies also recognize the fluoride activation mode, provide in-depth insights into the unprecedented mechanistic pathway and elucidate the reactivity difference of $\text{ArB}(\text{OH})_x\text{F}_y$ complexes, which fully support the experimental data.

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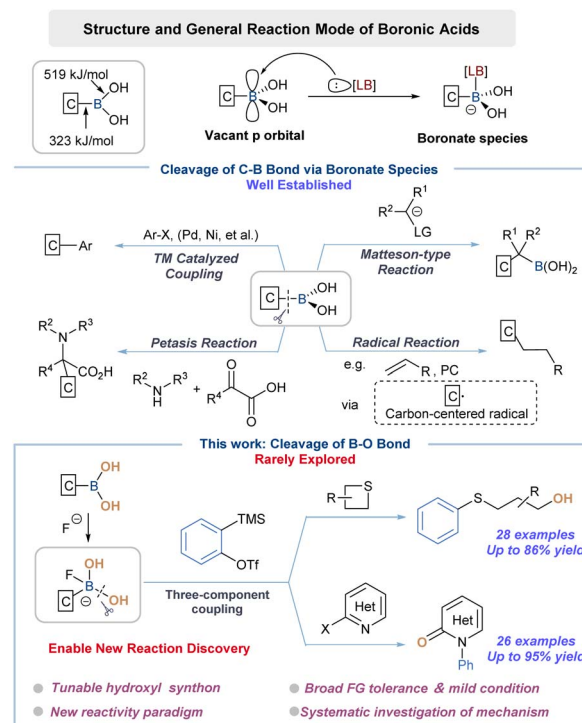
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Introduction

Boronic acids (BAs) have found myriad applications in modern organic synthesis as they can be readily converted to various functionalities.¹ Aside from versatile transformations, their synthetic merit is further emphasized by bench stability, low toxicity and operational simplicity. Specifically, the relatively weak C–B bonds (323 kJ mol^{-1}) in BAs have played a major synthetic role as carbon nucleophiles (Scheme 1).² For instance, the Suzuki–Miyaura coupling reaction, a “household” name in both academic and industrial settings, employed aryl halides and BAs to forge C–C bonds with high efficiency and great robustness.³ Other transformations including Chan–Lam,⁴ Matteson type⁵ and Petasis reaction⁶ have also been extensively investigated and advanced to applications. The recent resurgence of visible-light photochemistry was also fueled by the progress that BAs and their derivatives could serve as radical precursors through C–B bond homolysis.⁷ In particular, these remarkable achievements relied on the addition of a suitable Lewis base (LB) to activate BAs

toward C–B bond cleavage through the formation of anionic boron-centered “ate” complexes.⁸

Despite the aforementioned achievements, the strategic use of BA through B–O bond cleavage has been very rare, which might



Scheme 1 Reactivity of boronic acids: past and present.

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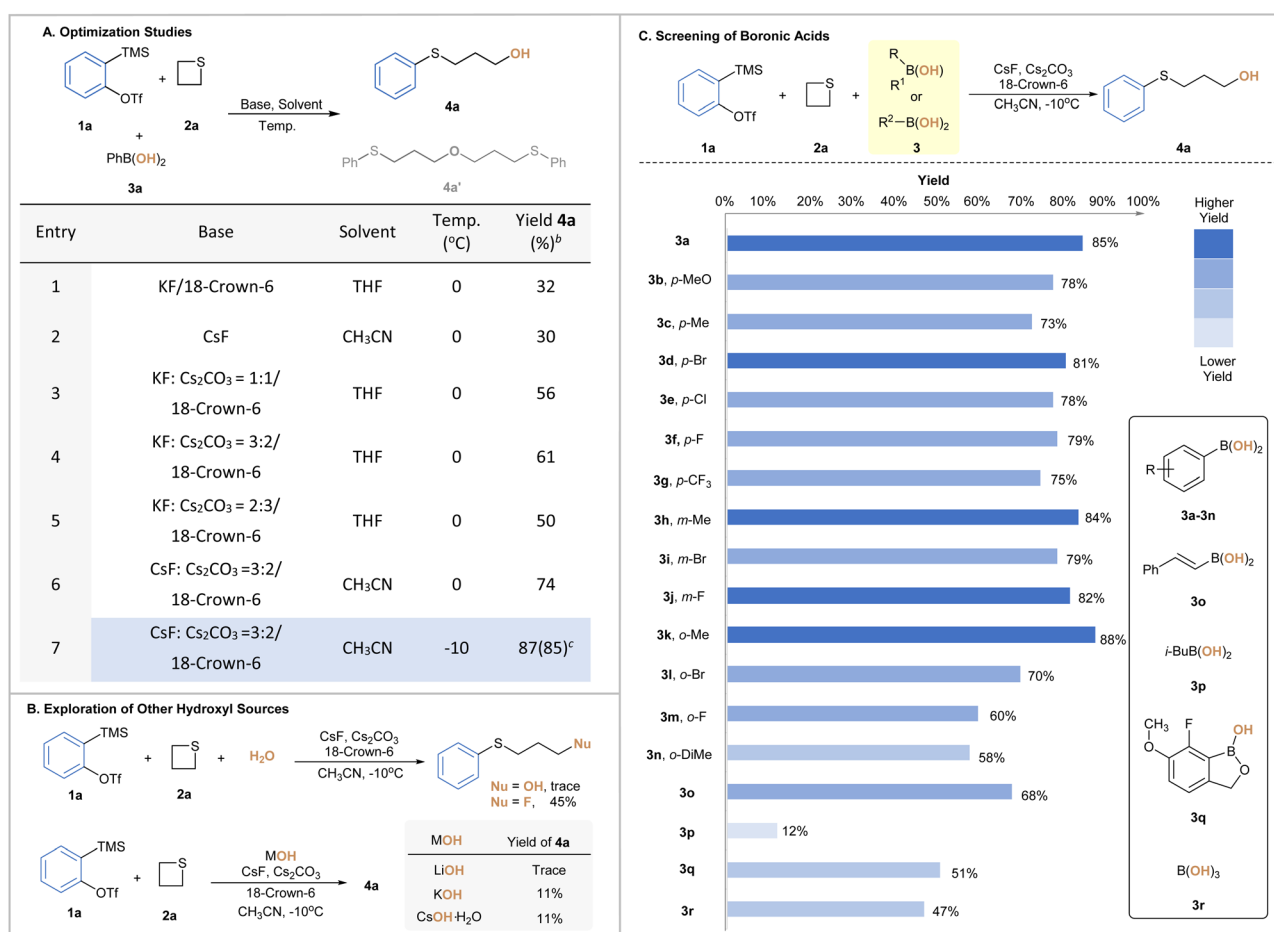
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be attributed to the much higher B–O bond energy than that of the C–B bond (519 vs. 323 kJ mol^{−1}).^{2a} In consideration of the synthetic advantage of BAs, a systematic investigation on employing BAs as a –OH source would be therefore of particular interest, which ought to expand the reactivity profile as well as the application value of such reagents, and even advance the fundamental knowledge of organoboron chemistry. In an attempt to fill the apparent synthetic gap, we recognized that identifying a suitable activator was crucial to switch the reactivity trend of BAs. From a conceptual standpoint, the leveraged activator ought to form a stronger bond with the boron center to elicit target behavior and not compete with desired hydroxylation reactions while displaying good functional group tolerance. Indeed, our research plan was initially guided by the recognition that aromatic BAs occasionally served as the affinity handle of fluorescent probes for fluoride ions.⁹ Naturally, we then questioned if the strong complexation of fluoride with boron (BDE B–F 646 kJ mol^{−1}) would necessitate our design, ensuring activation while forcing the hydroxylation reactivity from the “ate” complex. The only available example, nonetheless, focused solely on looking for similarities in reactivities between BA and

water rather than differences, letting alone detailed mechanistic studies.¹⁰ Hence, another ultimately significant task associated with our proposition was to fully understand the chemical reactivity character, mechanism of action and structure–activity relationship (SAR) of such boronate species.

In addition to the aforementioned considerations, our group's systematic research on exploring zwitterion chemistry to engage in a multicomponent reaction eventually inspired us to choose such a platform to thoroughly investigate our working hypothesis.¹¹ Note that we have established the nucleophilicity scale on these transformations, which ought to serve as a precise tool to map the whole reactivity profile of a B–F “ate” complex through comparative studies with other counterparts.¹² Besides being the “tailor-made” activator for BAs, fluorides fulfilled an additional task in facilitating the generation of aryl species from Kobayashi precursors.^{13,14} Despite recent efforts, we herein would like to report the successful execution of these ideas, and present a proof-of-concept use of BAs as a tunable hydroxy source in unprecedented aryl-based multicomponent coupling reactions. DFT calculations fully support the experimental data and shed light on the activation mode along with distinct mechanistic

Table 1 Reaction condition optimization^a



^a Reaction conditions: **1** (0.1 mmol), **2a** (0.2 mmol), **3** (0.15 mmol), base (0.5 mmol), and solvent (1 mL) for 12 hours at −10 °C. ^b Yields from ¹H NMR analysis. ^c Isolated yield.

underpinning involving a B–F ate complex. In addition to classic C–B bond cleavage manifolds, a variety of new discovery can be envisaged through the expansion of BA chemistry and zwitterion synthesis toolkits.

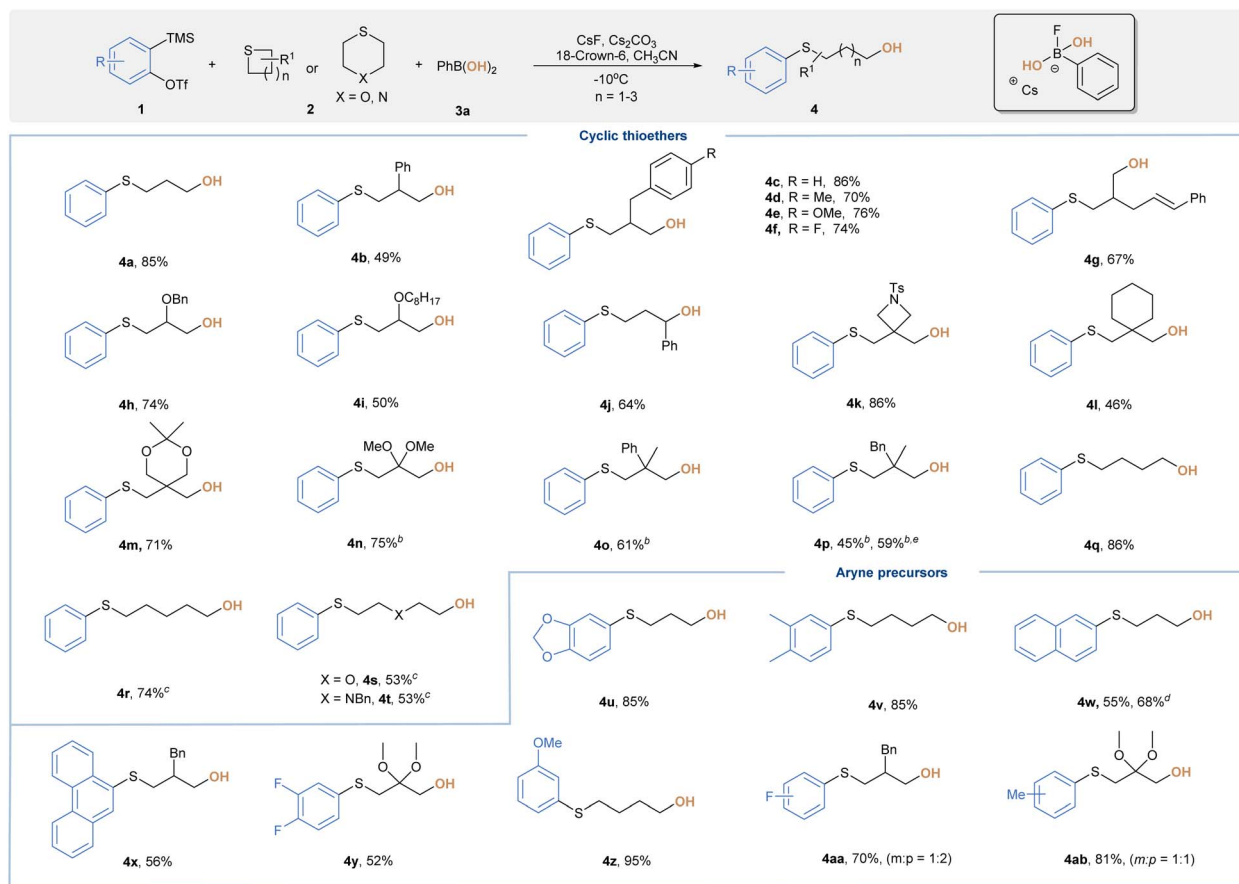
Results and discussion

To test the hypothesis, our investigation commenced with the model reaction of *o*-(trimethylsilyl)phenyl triflates (**1a**) and thietane (**2a**) with phenyl boronic acid (**3a**). To our delight, the ring-opening hydroxylation product **4a** was obtained in moderate yield without engaging any extra reagents to assist the B–O bond cleavage (Table 1A, entries 1–2). Meanwhile, ether (**4a'**) was observed as the side product albeit in a small amount. We speculated that it was probably formed *via* the ring opening of electrophilic aryne activated **4a** with the *in situ* generated alkoxide.¹⁵ To our delight, the addition of cesium carbonate (Cs₂CO₃) and varying the temperature could significantly inhibit such side products. Various mixtures of fluoride sources and cesium carbonate were then screened, among which a 3 : 2 ratio of CsF and Cs₂CO₃ improved the overall yield of **4a** to 74% (Table 1A, entries 3–6). Lowering the temperature to –10 °C

completely suppressed the formation of side product **4a'**, and the yield of **4a** was improved to 87% yield (Table 1A, entry 7). Finally, other known hydroxy sources were examined, and the obtained results demonstrated the distinct reactivity of BAs (Table 1B). In the presence of water as the 3rd component, a ring-opening fluorination product^{11d} was obtained as the major one through sulfonium intermediates under otherwise identical conditions.^{16,17} Clearly, this platform reaction served as an evaluation criterion of the nucleophilic scale to aid our understanding of the nature of such boronate complexes. When employing potassium hydroxide, lithium hydroxide or cesium hydroxide monohydrate as another option, we observed a significant loss of yield, which could be possibly attributed to the decomposition caused by their strong basicity.

To unveil the SAR, a diverse array of BAs was then screened. As shown in Table 1C, aromatic BAs bearing either electron-donating or -withdrawing groups with different substitution patterns were all viable, among which 2-methylphenylboronic acids (**3k**) offered optimal results.¹⁸ BAs with either alkenyl or aliphatic backbones proved to be less suitable under similar reaction conditions, which was presumably due to the poor affinity of the boron center at different levels. Finally, the use of

Table 2 Substrate scope^a



^a Reaction conditions: **1** (0.2 mmol), **2** (0.4 mmol), **3a** (0.3 mmol), CsF (0.6 mmol), Cs₂CO₃ (0.4 mmol), 18-crown-6 (0.4 mmol), and CH₃CN (2 mL) for 12 hours at –10 °C, and isolated yield. ^b CsF (0.15 mmol), Cs₂CO₃ (0.6 mmol), 18-crown-6 (0.6 mmol), and 40 °C. ^c 40 °C. ^d 2-MePhB(OH)₂ used instead. ^e 3-MePhB(OH)₂ used instead.

benzoxaborole (**3q**) and boric acid (**3r**) also gave **4a** in 51% and 47% yield, respectively. Clearly, varying electronic and steric parameters around the boron center could effectively tune the reactivity of BAs, providing further benefits to our strategy. In consideration of the compromise between the reaction outcome and cost effectiveness, **3a** was eventually chosen as the hydroxy-transfer precursors in the subsequent investigations.

With the optimal conditions in hand, we then assessed the substrate scope of this transformation (Table 2). Three-component reactions of 2 or 3-monosubstituted thietanes with aryne and **3a** readily afforded ring-opening hydroxylation products (**4b–4j**) in good yields. A wide range of functional groups were discovered to be suitable, including phenyl, benzyl, cinnamyl and alkoxy substitutions. Thietanes bearing all quaternary centers at the 3-position were also well-tolerated, leading to **4k–4p** in 45–86% yields. It is noteworthy that the transformation of 6-tosyl-2-thia-6-azaspiro[3.3]heptane (**2k**) occurred chemoselectively at the more nucleophilic sulfur center to provide **4k** in 86% yield, which could be further derived through ring opening of azetidines. For substrates bearing ketal moieties, the reaction also proceeded well to deliver **4m** and **4n** in 71% and 75% yield, respectively. We were also delighted to find that 5 or 6-membered cyclic sulfides with less ring strain were amenable to provide **4q–4t** in moderate yields, although elevated reaction temperature was required. Encouraged by these results, we further examined versatile aryne precursors. Symmetrical and unsymmetrical arynes bearing various functional groups all underwent the title reaction smoothly to deliver target products (**4u–4ab**) in good yields. When a 3-methoxybenzyl precursor was subjected to standard conditions, we obtained **4z** as a single regioisomer, which was consistent with the model proposed by Garg and Houk.¹⁹ As a further benefit, we were able to improve the overall reaction outcome for certain substrates by simply screening aryl BAs with different scaffolds (**4p** and **4w**). Of note, neither ring-opening fluorination products nor direct C–O coupling side reactions were observed during our investigations.

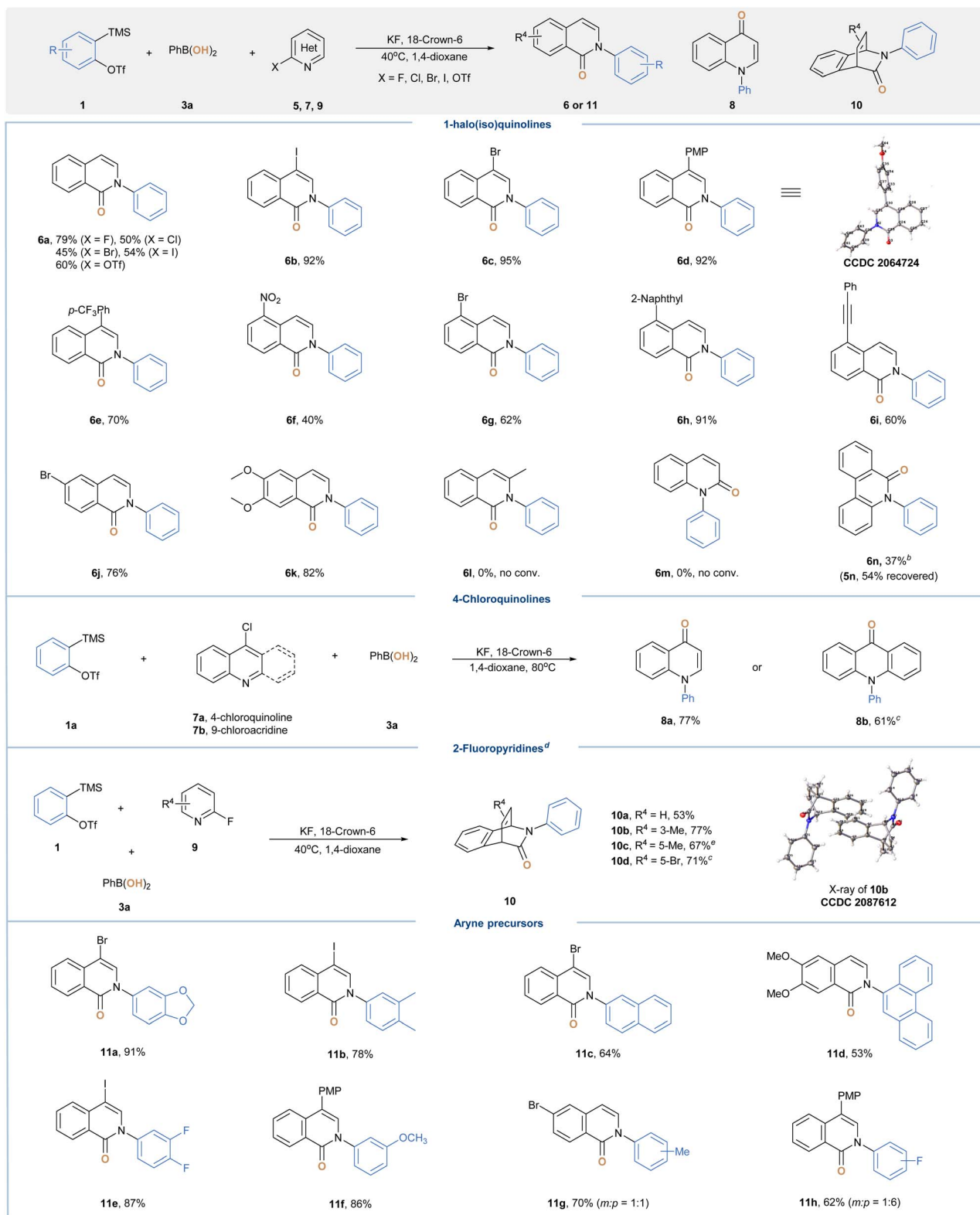
Having demonstrated the ability of aryl BAs as efficient nucleophilic hydroxy sources, we next ventured to explore other multicomponent coupling cascades in order to construct structurally complex and medicinally relevant heterocycles.²⁰ To this end, we sought to incorporate insights gained *vide supra* to investigate the aryne-induced electrophilic activation-nucleophilic substitution of halo-azaarenes with phenylboronic acids (Table 3). Once optimal conditions were identified,¹⁵ a series of 1-haloisoquinolines and their pseudo-halo counterparts were explored. Among them, 1-fluoroisoquinoline (**5a**) resulted in a better yield of 2-phenylisoquinolin-1-one (**6a**), and the observed reactivity trend was consistent with the leaving group effect for S_NAr .²¹ Next, substrates bearing various substitution groups were examined, and the desired 2-arylisquinolin-1-one products (**6b–6k**) were all obtained in good yields. Notably, bromo and iodo groups as potential synthetic handles for further derivatization were compatible with our conditions.²² Unfortunately, the transformations of 3-methyl-1-fluoroisoquinoline (**5l**) and 2-fluoroquinoline (**5m**) failed to deliver the corresponding products, whereas most of

the starting materials were recovered. We speculated that the steric hindrance around the nitrogen center suppressed the initial nucleophilic addition step. In line with such rationalization, the subjection of 6-fluorophenanthridine (**5n**) to standard conditions gave 5-phenylphenanthridin-6-one (**6n**) in 37% yield along with 54% of **5n** recovered. Subsequently, we assessed the applicability of the current method with respect to other electron-deficient azaarenes. To our delight, both 4-chloroquinoline and 9-chloroacridine were amenable to give 1-phenylquinolin-4-one (**8a**) and 10-phenylacridin-9-one (**8b**) in 77% and 61% yield, respectively. A unique reaction cascade comprising such three-component coupling and [4 + 2] cycloaddition was observed for 2-fluoropyridine substrates. In the presence of excess arynes, three-dimensional, bicyclic products (**10a–10d**) were readily acquired in 53–77% yields from planar heteroaromatics, which clearly demonstrated the potential “escape-from-flatland” value of our approach.²³ Finally, we extend this protocol to a similar set of aryne components and in all cases obtained desired heterocyclic products (**11a–11h**) in moderate to good yields.

To demonstrate the synthetic utility, scale-up experimentation was conducted under standard conditions, affording **4u** and **6c** at the same level of isolated yields (Scheme 2A). Next, we sought to present exemplary downstream derivatization of the obtained three-component coupling products (Scheme 2B). Sulfide **4u** could be easily converted to sulfoxide **12** and sulfone **13** in 92% and 98% yield respectively, depending on the reaction time and amount of *m*-CPBA used. An $FeCl_3$ -promoted intramolecular Friedel–Crafts reaction of **4j** proceeded smoothly to give the annulated product 4-phenylthiochromane (**14**) in 81% isolated yield. The treatment of **6g** with Lawesson's reagent furnished 5-bromoisoquinoline-1(2*H*)-thione **15** in 89% yield, while the palladium-catalyzed C–N coupling gave **16** in 83% yield. In addition, product **8a** was subjected to both copper catalyzed conjugate addition with *p*-tolylmagnesium bromide and *N*-iodosuccinimide (NIS) promoted iodination reactions, leading to **17** and **18** in 82% and 98% yield, respectively. Finally, our protocol proved to be applicable for late-stage modification of the active pharmaceutical ingredient, giving drug-like fasudil derivative (**6o**) in 50% yield (Scheme 2C). The showcased reactivity spectrum further signified the value of our protocol for facile preparation of highly versatile synthetic building blocks.

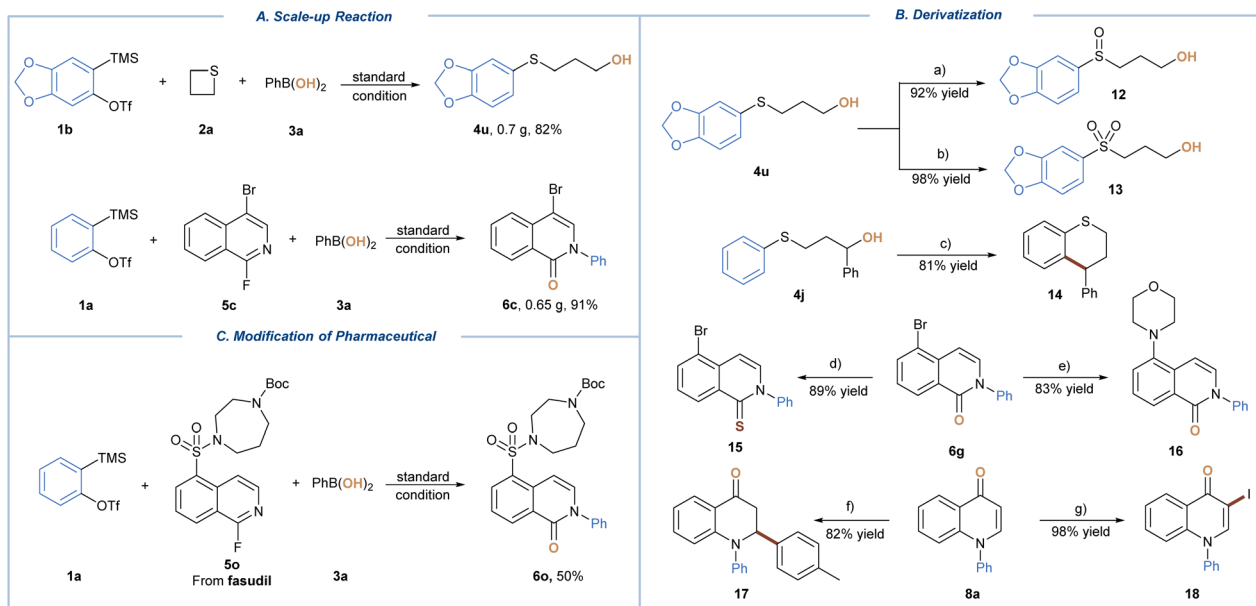
To shed light on the mechanistic underpinnings of these methods, both experimental and theoretical studies have been performed. First, a deuterium-labeling study was conducted employing $PhB(OD)_2$ (75% D-incorporation) under standard conditions, which gave an *ortho*-deuterated sulfide product (**4a-D**) in 86% yield and 68% deuterium incorporation, which not only indicated that the proton came from aryl BAs, but also ruled out the possibility of involving cyclic sulfonium ylide type intermediates (Scheme 3A).²⁴ In the absence of arynes, neither isoquinolin-1-one or 1-hydroxyisoquinoline was observed under otherwise identical conditions, thereby excluding the possible pathway involving hydroxylation (S_NAr) and *N*-arylation cascades.¹⁵ With the subjection of **3a** to CsF or the combination of KF and 18-crown-6, a new upfield peak at 5 ppm could be observed in the ¹¹B spectrum whereas the peak of **3a**



Table 3 Substrate scope^a

^a Reaction conditions: 1 (0.6 mmol), 5/7/9 (0.2 mmol), 3 (0.3 mmol), KF (1 mmol), 18-crown-6 (1 mmol), and dioxane (2 mL) for 24 hours at 40 °C, and isolated yield. ^b 80 °C. ^c 100 °C. ^d 1a (5 equiv). ^e 60 °C.

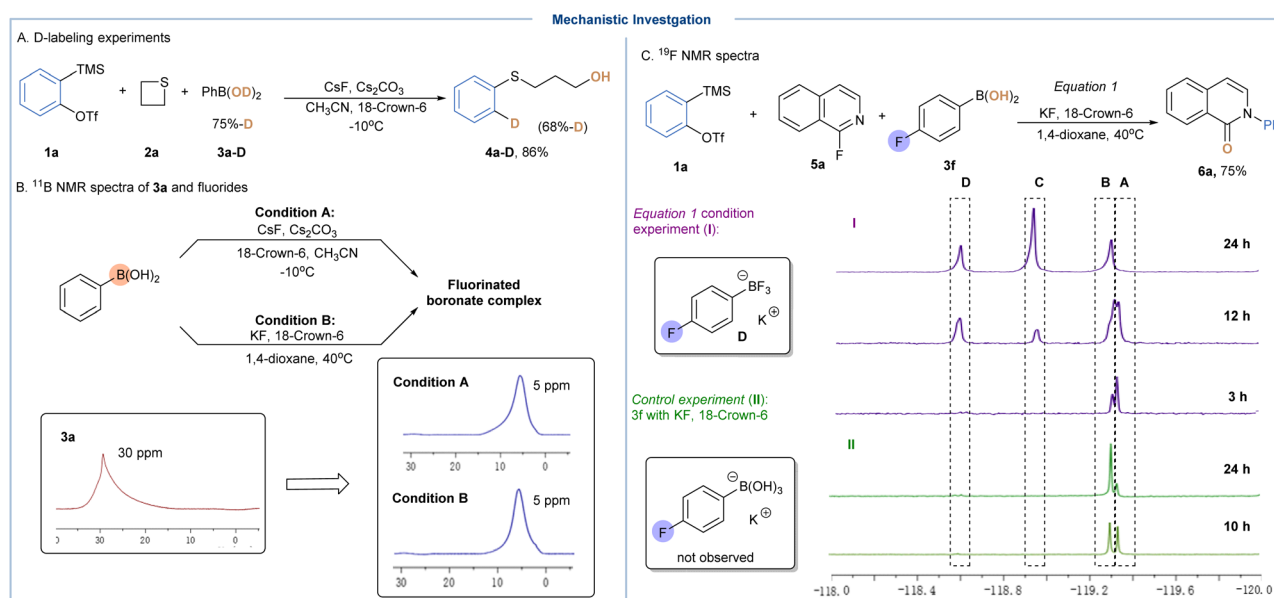




Scheme 2 Synthetic application. Reaction conditions: (a) *m*-CPBA (1.1 equiv), DCM, and 0 °C. (b) *m*-CPBA (3.2 equiv), DCM, and 0 °C to r.t. (c) FeCl_3 (1 equiv), DCE, and 70 °C. (d) Lawesson's reagent (4 equiv), toluene, and 120 °C. (e) Morpholine (2 equiv), $\text{Pd}_2(\text{dba})_3$ (10 mol%), BINAP (10 mol%), *t*-BuOK (3 equiv), toluene, and 110 °C. (f) CuI (2 equiv), TMSCl (3 equiv), *p*-TolMgBr (2 equiv), THF/DCM, and -78 °C. (g) NIS (1.5 equiv), AcOH, and r.t.

disappeared, offering direct evidence on the formation of tetrahedral boronate complexes (Scheme 3B). Next, the reaction progress of 4-fluorophenylboronic acid (3f) with 1a was monitored by the ^{19}F NMR technique. Upon treatment with KF and 18-crown-6, we immediately observed two signals at -119.34 ppm (peak A) and -119.36 ppm (peak B), which should refer to the formation of a B-F ate complex as supported by the control experiment (Scheme 3C).²⁵ As the reaction proceeded, peak C (-118.97 ppm) and peak D (-118.63 ppm) gradually

appeared, and the former was successfully identified as 4-fluorophenyltrifluoroborate by control experiments. Note that this observation suggested that our “fluoride activation-nucleophilic hydroxylation” strategy could drive the equilibrium on B-F coordination of BAs through the consumption of hydroxyl groups. Additionally, high-resolution mass spectrometry (HRMS) was employed to support the direct evidence on the *in situ* formed B-F complex.¹⁵



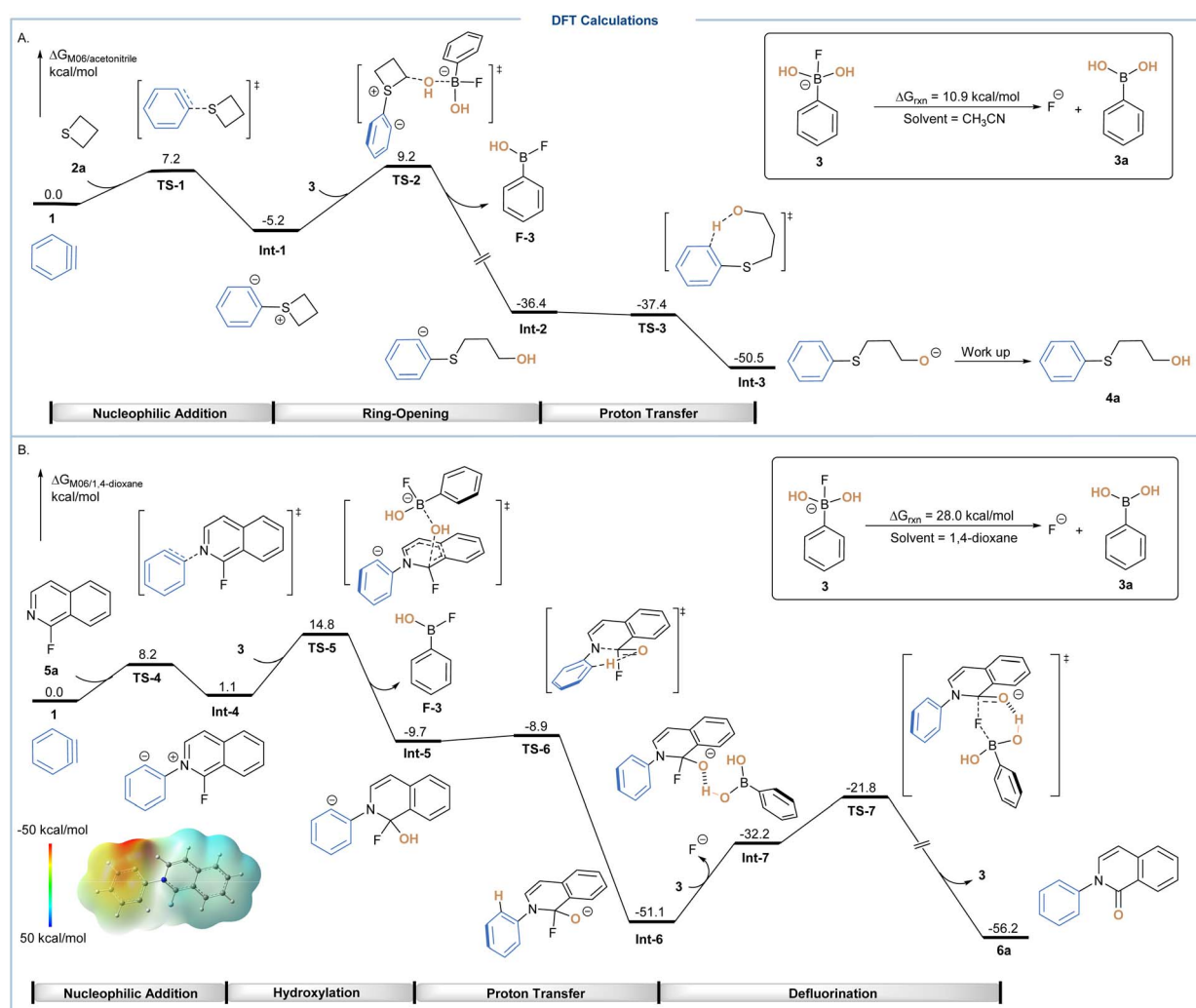
Scheme 3 Mechanistic study.



Density functional theory (DFT) calculations at the M06 level of theory were then conducted to explore the role of the B-F complex in both unprecedented three-component protocols. As shown in Scheme 4A, *in situ* generated aryne **1** was chosen as the relative zero in the calculated free energy profiles, which could react with thietane **2a** via nucleophilic addition transition state **TS-1** to afford a sulfonium zwitterion intermediate **Int-1**. The calculated activation free energy for this step proved to be only 7.2 kcal mol⁻¹, indicating an easily reachable process. After that, a ring-opening process could occur through the nucleophilic attack of the hydroxyl group on boronate via transition state **TS-2** with an energy barrier of 14.4 kcal mol⁻¹, leading to the cleavage of the B-O bond and formation of aryl anion intermediate **Int-2**. Subsequently, an intramolecular proton transfer via transition state **TS-3** through a barrier-less process led to the formation of alkoxide **Int-3**, which could be subsequently protonated during work up to yield the target product

4a. It is noteworthy that this boronate based hydroxylation system showed a unique pathway where instead of protonation to form sulfonium ions, the direct hydroxylation of sulfonium zwitterions occurred at the first place.

For the three-component coupling of 1-fluoroisoquinoline, *in situ* generated aryne **1** was again chosen as the relative zero in the calculated free energy profiles, which was nucleophilically attacked by 1-fluoroisoquinoline (**5a**) to form the zwitterionic isoquinolinium **Int-4** (the energy of activation is 8.2 kcal mol⁻¹) (Scheme 4B). For **Int-4**, the electrostatic potential surface clearly revealed that the isoquinoline moiety was still electron-deficient. Therefore, rather than protonation to form isoquinolinium salt, it could directly react with boronate **3** through an electrophilic substitution via transition state **TS-5** with an energy barrier of 13.7 kcal mol⁻¹. Afterwards, a dearomatized intermediate **Int-5** was formed, which then underwent a rapid intramolecular 1,5-proton transfer to form the stable



Scheme 4 DFT calculations. (A) Free energy profile for the ring-opening hydroxylation of aryne **1** and thietane **2a** with phenyl boronic acid **3a**. The energy values are in kcal mol⁻¹ and represent the relative free energies calculated at the M06/def2-TZVP/SMD(acetonitrile)//B3LYP/def2-SVP/SMD(acetonitrile) level of theory in acetonitrile solvent. (B) Free energy profile for the three-component reaction of aryne **1** and 1-fluoroisoquinoline **5a** with phenyl boronic acid **3a**. The energy values are in kcal mol⁻¹ and represent the relative free energies calculated at the M06/def2-TZVP/SMD(1,4-dioxane)//B3LYP/def2-SVP/SMD(1,4-dioxane) level of theory in 1,4-dioxane solvent.



isoquinolinolate **Int-6**. As the generation of **3a** was endergonic by 28.0 kcal mol⁻¹ from the corresponding boronate **3**, the combination of such a process and conversion of intermediate **Int-6** to **Int-7** was endergonic by 18.9 kcal mol⁻¹. Therefore, the overall activation free energy for the BA-assisted defluorination to yield target product **6a** was 29.3 kcal mol⁻¹, proceeding *via* a six-member ring transition state **TS-7**.

Alternatively, we considered the possible participation of phenyl(difluoro)(hydroxy)borate **19** as a hydroxy source *via* transition state **TS-8**, the relative free energy of which was only 3.0 kcal mol⁻¹ higher than its counterpart **TS-5** (see Fig. S6†).¹⁵ Accordingly, potassium phenyltrifluoroborate could be formed after the consumption of the hydroxyl group, which was consistent with the ¹⁹F NMR study (Scheme 3C). In the case of the fluoride dissociation step, other pathways involving either phenyl(hydroxy)(difluoro)borate **19** or phenyltrifluoroborate **21** were also taken into account, both of which displayed much higher reaction energies, making the corresponding pathway less favorable (see Fig. S7†).¹⁵

Conclusions

In summary, we have demonstrated that BAs can be employed as novel tunable hydroxide surrogates. The synthetic value was presented by hitherto unknown aryne triggered ring opening of cyclic sulfides and three-component synthesis of *N*-aryl (iso) quinolinones. Fluorides played a “tailor made” role in generating aryne species, activating BAs *via* complexation and facilitating B–O bond cleavage. The B–F “ate” complex displayed a unique nucleophilic paradigm that water or a metal hydroxide base couldn't convey. Furthermore, DFT calculation suggested that the unique hydroxylation of sulfur/nitrogen-containing zwitterions with boronate complexes occurred in favor of the formation of sulfonium/isoquinolinium intermediates *via* protonation. Experimental and theoretical studies also shed light on both the activation mode and reactivity trend of ArB(OH)₂F₂. Overall, we are strongly confident that the systematic knowledge gained in this investigation can be broadly translated to other reaction manifolds, thereby providing new synthetic blueprints for organoboron chemistry.

Data availability

The data supporting this study is available within the main text and the associate ESI.† The X-ray crystal structures reported in this paper have been deposited at the Cambridge Crystallographic Data Centre (CCDC) under the deposition numbers CCDC 2064724 (**6d**) and 2087612 (**10b**).

Author contributions

R. F. and J. T. conceived the project; R. F. performed the experiments; Q. Y. and J. W. prepared some starting materials; Y. W. discussed the results; S. L. performed the DFT calculation; Y. L. supervised the DFT calculations. R. F., S. L., Y. L. and Y. W. prepared this manuscript.

Conflicts of interest

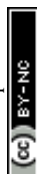
There are no conflicts to declare.

Acknowledgements

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