

Cite this: *Chem. Sci.*, 2023, 14, 9016

All publication charges for this article have been paid for by the Royal Society of Chemistry

Metal-free reductive desulfurization of C-sp³-substituted thiols using phosphite catalysis†

Rana M. I. Morsy,^a Ganesh Samala,^a Ankur Jalan,^b Michael E. Kopach,^b Naresh M. Venneti^{*,a} and Jennifer L. Stockdill^{*,a}

Phosphines and phosphites are critical tools for non-metal desulfurative methodologies due to the strength of the P=S bond. An overarching premise in these methods has been that stoichiometric (or excess) P(III) reagent is required for reactivity. Despite decades of research, a desulfurative process that is catalytic in phosphine/phosphite has not been reported. Here, we report the successful merging of two thermal radical processes: the desulfurization of unactivated and activated alkyl thiols and the reduction of P(V) = S to P(III) by reaction with a silyl radical species. We employ catalytic trimethyl phosphite, catalytic azo-bis(cyclohexyl)nitrile, and two equivalents of tris(trimethylsilyl)silane as the stoichiometric reductant and sulfur atom scavenger. This method is tolerant of common organic functional groups and affords products in good to excellent yields.

Received 5th January 2023
Accepted 3rd July 2023

DOI: 10.1039/d3sc00045a

rsc.li/chemical-science

Introduction

Desulfurization is an important process¹ for diverse applications from organic synthesis² and chemical protein synthesis² to crude oil processing and refinement.³ Numerous reports describe transition metal-based reduction reactions, including Ni,⁴ Co,⁵ W,⁶ Mo,⁷ and others.⁸ Unfortunately, critical limitations remain, which include pyrophoricity,^{4,5} toxic H₂S gas evolution,⁴ reactivity with other S-functional groups,^{4,6} and the need for specialized reaction setups.⁵ Additionally, when the desulfurized products are deployed in living systems, trace metal contamination can be problematic.⁹ Photoredox strategies^{10a,b} avoid functional group incompatibility^{10a} and may enable further functionalization of the substrate; however, they do require the use of expensive rare earth transition metal complexes, which may limit industrial scale applications. With only two exceptions,^{10d,e} modern and non-metal strategies for thiol desulfurization^{10,11} employ excess phosphine/phosphite reagents, whose aquatic toxicity⁹ has limited their utility in >kg-scale processes. The last several years have seen a surge of interest in the potential of P(III)/P(V) catalysis,¹² largely focused on strategies to facilitate polar P(III)/P(V) = O cycles.¹³ We were intrigued by the potential of employing P(III)/P(V) = S catalysis for desulfurative radical transformations. Here, we present the first step along this journey, a metal-free desulfurization of unactivated thiols by P(III)/P(V) = S catalysis. We achieve this

reactivity by leveraging the affinity of silyl radicals toward sulfur¹⁴ to achieve *in situ* regeneration of phosphites by reduction of the P=S bond while sequestering the S atom to avoid H₂S formation.

Results and discussion

Reports as recent as 2021 have asserted that stoichiometric phosphine is a requirement for radical desulfurization reactions.^{12b} Both phosphines and phosphites are well established reagents for desulfurization reactions,^{11,15} with the electronics of the R group on PR₃ being key to reactivity.¹⁶ Phosphine sulfides react with super silyl hydride (TTMSS)¹⁷ under free radical conditions to produce phosphines in good yields.¹⁴ Unlike P(III)/P(V) = O redox cycles, which occur *via* polar pathways, the P(III)/P(V) = S reduction occurs *via* a radical mechanism.¹⁴ Thus, we envisioned that this cycle would merge seamlessly with a radical desulfurization mechanism. We employed Ac-N-Cys-OMe¹⁸ (**1a**) as a model substrate to investigate the feasibility of the catalytic transformation in the presence of various phosphines/phosphites with TTMSS as a phosphine sulfide reductant¹⁴ and/or terminal hydrogen atom donor¹⁹ (Table 1). We selected azo-bis(cyclohexyl)nitrile (ACHN) as the initiator because of its 10 h half-life²⁰ and to avoid the formation of oxygenated by-products commonly observed in the presence of organic peroxide initiators.²¹ Because of the electrophilic nature of thiyl radicals,²² we began our studies with 10 mol% of the electron-rich tris(dimethylanion)phosphine ((Me₂N)₃P).²³ Only a single turnover was observed, with the desired product (**1b**) formed in 21% conversion (entry 1). We limited reaction time to 24 h for the purposes of this screen. The less basic tri-*tert*-butylphosphine was markedly better with 65%

^aDepartment of Chemistry, Wayne State University, Detroit, MI 48202, USA. E-mail: gu0132@wayne.edu; stockdill@wayne.edu

^bEli Lilly and Company, Indianapolis, IN 46285, USA

† Electronic supplementary information (ESI) available: Experimental details, tabulated spectroscopic data, and images of spectra. See DOI: <https://doi.org/10.1039/d3sc00045a>

Table 1 Establishing PR₃ reactivity under catalytic conditions^a

Entry	PR ₃ (equiv.)	TTMSS (equiv.)	ACHN (equiv.)	Temp. (°C)	Time (h)	Conversion ^b (%)
1	(Me ₂ N) ₃ P (0.1)	1	0.1	80	24	21
2	^t Bu ₃ P (0.1)	1	0.1	80	24	65
3	ⁿ Bu ₃ P (0.1)	1	0.1	80	24	54
4	P(Cy) ₃ (0.1)	1	0.1	80	24	38
5	P(OMe) ₃ (0.1)	1	0.1	80	24	80
6	P(OMe)₃ (0.2)	2	0.1	88	16	>99
7	P(OMe) ₃ (0.2)	2	—	88	16	0
8	—	2	0.1	88	16	0
9	P(OMe) ₃ (0.2)	—	0.1	88	16	7
10	P(OMe) ₃ (0.2)	2	0.1	25	16	0

^a Reactions were performed with 1 equiv. **1a** (0.05 M) in degassed PhMe. ^b Conversion = consumption of **1a**; based on ¹H-NMR integration.

conversion (entry 2), while the less hindered and less electron-rich tri(*n*-butyl)phosphine led to 54% conversion (entry 3). Interestingly, tri-cyclohexylphosphine, which bridges these steric and electronic properties, was even less effective (38% conversion, entry 4). Trimethyl phosphite balanced these effects, improving the conversion to 80% (entry 5). Increasing the phosphite loading to 20 mol% led to complete conversion by 16 h (entry 6).²⁴ Control experiments were consistent with the radical pathway we envisioned for this chemistry. Omission of the radical initiator (entry 7) or the phosphite (entry 8) abolished reactivity. In the absence of TTMSS, slow conversion was observed and did not reach 20% because the substrate thiol must serve as the H-atom donor for C–H bond formation (entry 9). Finally, heating is required for the reduction of the P=S bond by TTMSS, as evidenced by entry 10.

We proceeded with trimethyl phosphite at 88 °C for 16 h while screening different reaction parameters such as initiator,

reductant, and solvent (Table 2). Using different initiators like VA-044 or Luperox A98 led to large decreases in conversion (entries 1–2). Dicumyl peroxide led to a 3-fold increase in conversion percentage relative to VA-044 at the same temperature (entry 3), but it was still inferior to ACHN.

Alternative hydrogen atom/silyl radical donors had a deleterious effect on conversion. When tri-isopropylsilane (TIPS) was used, the reaction takes place, but the phosphite is not reduced (Table 2, entry 4). This apparent limited efficiency of TIPS compared to TTMSS is logical given the difference in their Si–H bond-dissociation energies (BDEs) with trialkylsilanes ranging from (90–96 kcal mol^{−1})²⁵ compared to the Si–H in TTMSS at 79 kcal mol^{−1}.²⁶ The instability of the TIPS[•] interferes with the turnover of the phosphine sulfide, presumably because H-atom abstraction from the substrate or solvent occurs more rapidly than Si–S bond formation. Alternatively, the *i*-Pr group is slow to migrate after α -scission. Consistent with this reasoning, and with the similar properties of Ge and Si,²⁷ Et₃GeH (86 kcal mol^{−1}) was able to effect ~1 phosphine sulfide turnover in 16 h, proceeding with 43% conversion (entry 5). The best alternate P=S reductant was *n*-Bu₃SnH (74 kcal mol^{−1}), which still reduced the phosphine sulfide more slowly than TTMSS (entry 6).²⁸ The respective BDEs of TTMSS and *n*-Bu₃SnH indicate that the Si[•] and Sn[•] species are similarly stable and the rate difference likely arises from differences in affinity to sulfur or the ability of TTMSS to irreversibly sequester the S-atom *via* an additional α -fragmentation and 1,2-group shift (see Scheme 2).

In terms of solvent, THF and MeCN, each of which were performed at 88 °C using a sealed microwave tube, were less competent than toluene (entries 7–8). Interestingly, the reaction proceeded to complete conversion with 1,4-dioxane as solvent (entry 9). Indeed, the success of 1,4-dioxane increases the scope of this method to include polar substrates.

With two sets of established conditions in hand, we evaluated the scope of our efficient reductive desulfurization reaction on different relevant substrates, including short peptides and small molecule substrates that would form intermediate C-

Table 2 Probing the impact of radical initiator, reductant, and solvent^a

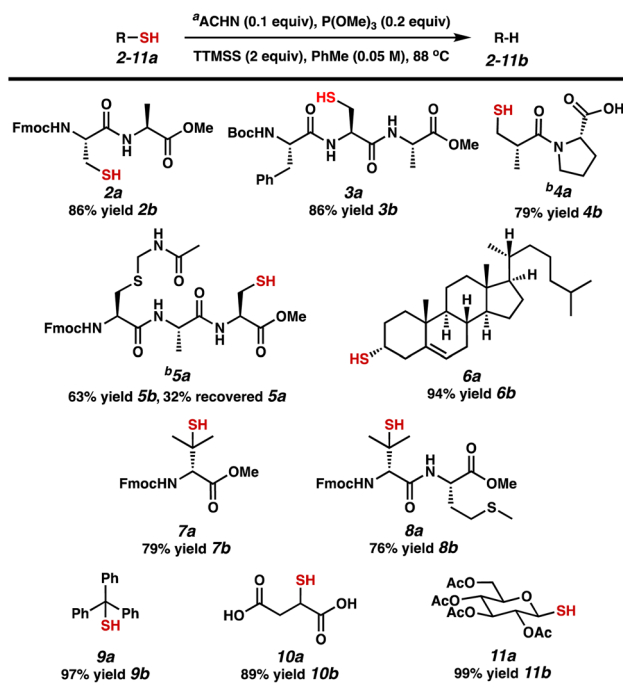
Entry	Initiator	Reductant	Solvent	Conversion ^b (%)
1	VA-044	TTMSS	H ₂ O/MeCN	44
2	Luperox A98	TTMSS	PhMe	22
3	Dicumyl peroxide	TTMSS	PhMe	65
4	ACHN	TIPS	PhMe	7
5	ACHN	Et ₃ GeH	PhMe	43
6	ACHN	ⁿ Bu ₃ SnH	PhMe	63
7	ACHN	TTMSS	MeCN	62
8	ACHN	TTMSS	THF	82
9	ACHN	TTMSS	1,4-Dioxane	>99

^a Reactions were performed with 1 equiv. **1a** (0.05 M) in degassed PhMe.

^b Conversion = consumption of **1a**; based on ¹H-NMR integration.



Table 3 Strategic substrate scope for catalytic desulfurization



^a Reactions were performed with 1 equiv. **2a–11a** in degassed toluene. Yields are of isolated material following purification *via* silica gel flash chromatography or HPLC. ^b Reaction was performed in 1,4-dioxane.

radicals with varying stability (Table 3). Both protected dipeptide **2a** and tripeptide **3a** – each containing a 1° alkyl thiol – were desulfurized in 86% yield. Carboxylic acid-containing captropil

(**4a**) was desulfurized in 79% yield. Tripeptide **5a** performed more sluggishly, producing **5b** in 63% yield with 32% recovered **5a**. We next investigated 2° and 3° alkyl thiol substrates. Cholesterol derivative **6a** was converted to cholest-5-ene **6b** in 94% yield. The B-ring olefin was unaffected. Penicillamine derivatives **7a** and **8a** proceed through tertiary C-radicals to afford valine-containing peptides **7b** (79% yield) and **8b** (76% yield). Notably, the thioether of methionine was stable to these conditions. Finally, we tested substrates that proceed through α -stabilized radicals. Trityl thiol **9a** was desulfurized to form triphenylmethane (**9b**) in 97% yield. The thiols in diacid **10a** and glucose derivative **11a** were readily converted to the corresponding C–H bonds to provide **10b** in 89% yield and a chiral pyran derivative (**11b**) in 99% yield.

Efforts to extend these results to longer and unprotected peptides were not fruitful due to lack of solubility of the starting materials. A brief co-solvent screen of a dipeptide acid (Fmoc-Cys-Ala-OH, SI-13†) was conducted to probe the conversion efficiency in the presence of various amounts of water, which would be needed to increase the efficiency of the desulfurization process for peptide substrates. Water is tolerated in the reaction, with the caveat that the reactivity is significantly slowed and by-product formation occurs (disulfide, thioether, and C–C bond formation were observed).²³ The maximum conversion obtained for dipeptide SI-13† was 62%, requiring 2 additions of ACHN and P(OMe)₃. We suspect that the hydrophobicity of TTMSS is limiting in these reactions. Presently, a water-soluble analog of TTMSS is not available, and this chemistry is therefore best suited to organic-soluble substrates. Importantly, our interest was not in achieving an ideal peptide desulfurization,^{10e} but rather in demonstrating P(III)/P(V) = S catalysis and learning about its behavior.

Table 4 Functional group compatibility screen

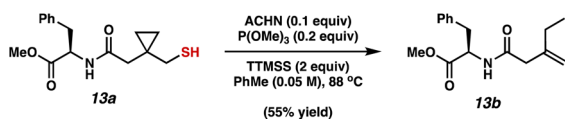
<div style="text-align: center;"> </div>							
Entry ^a	Additive	Unreacted 12a ^b (%)	GC yield 12b ^b (%)	Entry ^a	Additive	Unreacted 12a ^b (%)	GC yield 12a ^b (%)
1	None	0	87	6		100	0
2		0	87	7		0	57
3		7	91	8		95	4
4		0	93	9		0	81
5		0	59	10		0	55

^a Reaction conditions: (4-methoxyphenyl)methanethiol (0.2 mmol, 1 equiv.), P(OMe)₃ (0.04 mmol, 0.2 equiv.), ACHN (0.02 mmol, 0.1 equiv.), TTMSS (0.4 mmol, 2 equiv.), toluene (0.05 M), 88 °C, 3 h. ^b Calculated *via* GCMS using 1,3,5-trimethoxybenzene as internal standard.

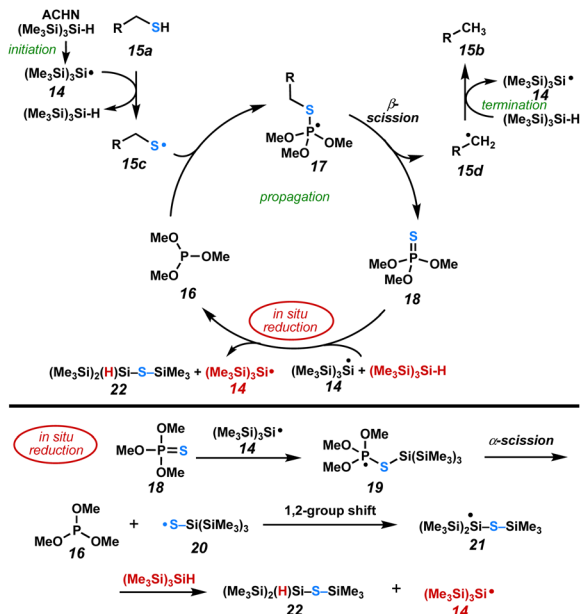


To probe the amenability of the desulfurization to large scale conditions, we performed the reaction on 4-methoxybenzyl mercaptan (**12a**). On 6.84 mmol scale in 1,4-dioxane, 100% conversion was observed in 3 h, with clean reactivity and no unexpected by-product formation (Scheme SI-01†). To probe the compatibility of various functional groups with the reaction conditions, we again employed 4-methoxybenzyl mercaptan (**12a**) in toluene along with different additives. The additives shown in Table 4 were compatible with the reaction conditions, showing no impact on the additive. In most cases, good to excellent reactivity of **12a** was maintained. In entries 6 and 8, the desulfurization reaction did not occur. Alcohol additives had no negative effect on the desulfurization but were silylated. Aldehyde and ketone additives led to an unknown by-product and interfered somewhat with the desulfurization (see Table SI-02†).

A radical clock experiment (Scheme 1) was performed with cyclopropane-containing thiol, **13a**, which underwent ring opening to afford olefin **13b** in 55% yield. This indicates that a carbon-centered radical is an intermediate in this reaction. Based on our observations, we hypothesize that reduction of the phosphine sulfide is rate-limiting and proceeds *via* the radical pathway described by Chatgililoglu.¹⁴ The postulated overall radical chain mechanism for the catalytic reductive desulfurization is illustrated in Scheme 2.



Scheme 1 Cyclopropylcarbinyl radical rearrangement product supports a radical pathway.



Scheme 2 Mechanism of catalytic desulfurization.

Upon heating, ACHN releases N_2 to initiate the reaction by forming the C-centered cyclohexyl nitrile radical. H-atom abstraction from TTMSS leads to super silyl radical **14**. H-atom abstraction from the sulfhydryl group in **15a** forms thiyl radical **15c** and regenerates TTMSS.²² Thiyl radical **15c** reacts with phosphite²⁹ **16** to generate phosphoranyl radical intermediate **17**.³⁰ Tetravalent P-centered radical **17** undergoes β -scission to produce trimethyl thiophosphate **18** and the carbon centered radical **15d**. H-atom transfer from TTMSS to **15d** produces the desulfurized product **15b** and regenerates TTMSS radical **14**. Trimethyl thiophosphate **18** reacts with silyl radical **14** to form phosphoranyl radical intermediate **19**. Subsequent α -scission regenerates the catalytically active phosphite (**16**) to be reused in the propagation step. Notably, this α -scission is facilitated by the thiophilic nature of the silane (Si-S 148.4 *versus* P-S 105.6 kcal mol⁻¹).²⁵ The resulting S-centered radical **20** induces a 1,2-group shift³¹ of a TMS group from silicon to sulfur, furnishing silyl radical **21** that reacts with a second equivalent of TTMSS to afford TTMSS radical **14** and silyl sulfide **22**.³² This completes the *in situ* reduction process and permits propagation of the radical cascade. GCMS analysis of the crude reaction mixture indicates the presence of a small amount of phosphine sulfide **18** and multiple TTMSS-related peaks during the reaction.

Conclusions

In summary, we have established reaction conditions for P(III)/P(V) = S catalysis that proceeds through a tetravalent P-centered radical rather than through a pentavalent P(V) intermediate as is observed for P=O reduction. The reductive desulfurization reaction sequesters the S atom, avoids the use of rare earth and transition metals, and affords desulfurized products in good yields. The method has been tested across a range of substrate types including thiols that lead to benzylic, primary, secondary, and tertiary alkyl C-centered radicals. Importantly, good functional group compatibility is observed, including free carboxylic acid, ester, carbamate, amide, nitrile, aryl halide, amine, isocyanate, thioether functionalities. Steroid, carbohydrate, amino acid, pyridine, and benzene scaffolds are compatible with the reaction. No special setup is required for this catalytic method, so it can be readily employed and adapted. This study expands the scope of metal-free and thiol-additive-free desulfurization methods while establishing the viability of P(III)/P(V) = S catalysis under radical conditions. Efforts to harness this catalysis strategy to effect other transformations and develop new phosphorus-based catalysts are ongoing.

Data availability

The datasets supporting this article have been uploaded as part of the ESI.†

Author contributions

All authors have contributed to this research.



Conflicts of interest

There are no conflicts to declare.

Acknowledgements

In Fondlest Memory of Prof. Robert H. Grubbs, a world leader in chemistry, humility, and kindness. These studies were supported by the National Science Foundation (CAREER Award to JLS; CHE-1554752) and the Lilly Research Award Program (LRAP) at Eli Lilly and Co. We thank A. Paul and Carol C. Schaap and Wayne State University for a Rumble-Schaap Fellowship to (RMIM).

Notes and references

- 1 J. Rentner, M. Kljajic, L. Offner and R. Breinbauer, Recent advances and applications of reductive desulfurization in organic synthesis, *Tetrahedron*, 2014, **70**, 8983–9027, DOI: [10.1016/j.tet.2014.06.104](#).
- 2 (a) J. J. Eisch, L. E. Hallenbeck and K. I. Han, Organic chemistry of subvalent transition metal complexes. 12. Hydrodesulfurization of organosulfur heterocycles by metal hydride-nickel (0) complexes: accelerated single-electron transfer in carbon-sulfur bond cleavage, *J. Am. Chem. Soc.*, 1986, **108**, 7763–7767, DOI: [10.1021/ja00284a045](#); (b) H. Rohde and O. Seitz, Invited review ligation—desulfurization: a powerful combination in the synthesis of peptides and glycopeptides, *Pept. Sci.*, 2010, **94**, 551–559, DOI: [10.1002/bip.21442](#); (c) K. Jin and X. Li, Advances in native chemical ligation—desulfurization: a powerful strategy for peptide and protein synthesis, *Chem.-Eur. J.*, 2018, **24**, 17397–17404, DOI: [10.1002/chem.201802067](#).
- 3 J. Zhao and R. Wang, Research Progress of Phase Transfer Catalysts Used in Oxidative Desulfurization of Fuel Oil, *Mini Rev. Org. Chem.*, 2021, **18**, 626–648, DOI: [10.2174/1570193X17999200820160851](#).
- 4 (a) K. M. Ho, C. H. Lam and T. Y. Luh, Transition metal promoted reactions. Part 27. Nickel (II)-lithium aluminum hydride mediated reduction of carbon-sulfur bonds, *J. Org. Chem.*, 1989, **54**, 4474–4476, DOI: [10.1021/jo00279a046](#); (b) M. C. Chan, K. M. Cheng, K. M. Ho, C. T. Ng, T. M. Yam, B. S. Wang and T. Y. Luh, Transition-metal-promoted reactions. 19. Nickelocene-lithium aluminum hydride: a versatile desulfurization reagent, *J. Org. Chem.*, 1988, **53**, 4466–4471, DOI: [10.1021/jo00254a009](#); (c) M. C. Chan, K. M. Cheng, M. K. Li and T. Y. Luh, Nickelocene-lithium aluminium hydride, a new effective desulphurization reagent, *J. Chem. Soc., Chem. Commun.*, 1985, 1610–1611, DOI: [10.1039/C39850001610](#).
- 5 (a) H. Alper, F. Sibtain and J. Heveling, Phase transfer catalyzed desulfurization reactions, *Tetrahedron Lett.*, 1983, **24**, 5329–5332, DOI: [10.1016/S0040-4039\(00\)87860-0](#); (b) S. C. Shim, S. Antebi and H. Alper, Desulfurization of mercaptans to hydrocarbons by carbon monoxide and water in the presence of cobalt carbonyl, *Tetrahedron Lett.*, 1985, **26**, 1935–1938, DOI: [10.1016/S0040-4039\(00\)98345-X](#).
- 6 C. T. Ng, X. Wang and T. Y. Luh, Transition metal-promoted reactions. Part 21. Tungsten hexacarbonyl-mediated carbon-sulfur bond cleavage reactions, *J. Org. Chem.*, 1988, **53**, 2536–2539, DOI: [10.1021/jo00246a024](#).
- 7 (a) H. Alper and C. Blais, Modified molybdenum carbonyl species; excellent reagents for the desulphurization of thiols, *J. Chem. Soc., Chem. Commun.*, 1980, 169–170, DOI: [10.1039/C39800000169](#); (b) T. Y. Luh and C. S. Wong, Transition-metal promoted reactions. 12. Molybdenum hexacarbonyl-promoted reductive cleavage of the carbon-sulfur bond, *J. Org. Chem.*, 1985, **50**, 5413–5415, DOI: [10.1021/jo00225a091](#); (c) Z. Wang, Y. Kuninobu and M. Kanai, Molybdenum-mediated desulfurization of thiols and disulfides, *Synlett*, 2014, **25**, 1869–1872, DOI: [10.1055/s-0034-1378315](#).
- 8 (a) H. Alper, S. Ripley and T. L. Prince, Desulfurization of thiols and thioketones by sodium triethylborohydride and iron (II) chloride on alumina, *J. Org. Chem.*, 1983, **48**, 250–252, DOI: [10.1021/jo00150a020](#); (b) Z. Yu and J. G. Verkade, Mild and efficient desulfurization of alkyl sulfides with sodium, *Tetrahedron Lett.*, 1998, **39**, 2671–2674, DOI: [10.1016/S0040-4039\(98\)00368-2](#); (c) H. Alper and T. L. Prince, Exceedingly Mild Desulfurization of Thiols by Sodium Triethylhydroborate and Transition Metal Chlorides, *Angew. Chem., Int. Ed. Engl.*, 1980, **19**, 315–316, DOI: [10.1002/anie.198003151](#); (d) T. Ganguly, A. Das and A. Majumdar, Iron (II) Mediated Desulfurization of Organosulfur Substrates Produces Nonheme Diiron (II)-hydrosulfides, *Inorg. Chem.*, 2019, **58**, 9998–10011, DOI: [10.1021/acs.inorgchem.9b01144](#); (e) E. Gretz, W. M. Vetter, H. A. Stecher and A. Sen, Homogeneous Desulfurization of Thiols and Sulfides by Yb [N (SiMe₃)₂]₂·2 THF. Relevance to Heterogeneous and Biological Desulfurizations and the Synthesis of Lanthanide (II) Sulfides, *Isr. J. chem.*, 1990, **30**, 327–329, DOI: [10.1002/jch.199000034](#).
- 9 (a) A. M. Sciuto, B. J. Wong, M. E. Martens, H. Hoard-Fruchey and M. W. Perkins, Phosphine toxicity: a story of disrupted mitochondrial metabolism, *Ann. N. Y. Acad. Sci.*, 2016, **1374**, 41–51, DOI: [10.1111/nyas.13081](#); (b) J. Roels and W. Verstraete, Biological formation of volatile phosphorus compounds, *Bioresour. Technol.*, 2001, **79**, 243–250, DOI: [10.1016/S0960-8524\(01\)00032-3](#); (c) B. Wang, Q. Shen, C. Han, Y. Zheng, Z. Wang, C. Liu, L. Zhang and J. Ren, New insights into the growth response of the macrophyte *Vallisneria spiralis* exposed to phosphite, *Sci. Total Environ.*, 2022, **851**, 158189, DOI: [10.1016/j.scitotenv.2022.158189](#).
- 10 (a) X. F. Gao, J. J. Du, Z. Liu and J. Guo, Visible-light-induced specific desulfurization of cysteinyl peptide and glycopeptide in aqueous solution, *Org. Lett.*, 2016, **18**, 1166–1169, DOI: [10.1021/acs.orglett.6b00292](#); (b) M. Lee, S. Neukirchen, C. Cabrele and O. Reiser, Visible-light photoredox-catalyzed desulfurization of thiol- and disulfide-containing amino acids and small peptides, *J. Pept. Sci.*, 2017, **23**, 556–562, DOI: [10.1002/psc.3016](#); (c) E. H. Discekici, S. L. Shankel, A. Anastasaki, B. Oschmann, I. H. Lee, J. Niu, A. J. McGrath, P. G. Clark, D. S. Laitar, J. Read de Alaniz, C. J. Hawker and D. J. Lunn, Dual-



- pathway chain-end modification of RAFT polymers using visible light and metal-free conditions, *Chem. Commun.*, 2017, **53**, 1888–1891, DOI: [10.1039/C6CC08370F](#); (d) S. Kaur, D. P. Luciano, X. Fan, G. Zhao, S. Messier, M. M. Walker, Q. Zhang and T. Wang, Radical functionalization of thioglycosides in aqueous medium, *Tetrahedron Lett.*, 2021, **86**, 153499, DOI: [10.1016/j.tetlet.2021.153499](#); (e) N. M. Venneti, G. Samala, R. M. I. Morsy, L. G. Mendoza, A. Isidro-Llobet, J. K. Tom, S. Mukherjee, M. E. Kopach and J. L. Stockdill, Phosphine-dependent photoinitiation of alkyl thiols under near-UV light facilitates user-friendly peptide desulfurization, *J. Am. Chem. Soc.*, 2023, **145**, 1053–1061.
- 11 For selected publications see (a) F. W. Hoffmann, R. J. Ess, T. C. Simmons and R. S. Hanzel, The Desulfurization of Mercaptans with Trialkyl Phosphites, *J. Am. Chem. Soc.*, 1956, **78**, 6414, DOI: [10.1021/ja01605a033](#); (b) C. Walling and R. Rabinowitz, The Reaction of Thiyl Radicals with Trialkyl Phosphites, *J. Am. Chem. Soc.*, 1957, **79**, 5326, DOI: [10.1021/ja01576a077](#); (c) C. Walling and R. Rabinowitz, The reaction of trialkyl phosphites with thiyl and alkoxy radicals, *J. Am. Chem. Soc.*, 1959, **81**, 1243–1249, DOI: [10.1021/ja01514a057](#); (d) C. Walling, O. H. Basedow and E. S. Savas, *J. Am. Chem. Soc.*, 1960, **82**, 2181–2184, DOI: [10.1021/ja01494a023](#); (e) C. Walling and M. S. Pearson, Some Radical Reactions of Trivalent Phosphorus Derivatives with Mercaptans, Peroxides, and Olefins. A New Radical Cyclization, *J. Am. Chem. Soc.*, 1964, **86**, 2262–2266, DOI: [10.1021/ja01065a032](#); (f) W. G. Bentrude, Phosphoranyl radicals-their structure, formation, and reactions, *Acc. Chem. Res.*, 1982, **15**, 117–125, DOI: [10.1021/ar00076a004](#); (g) W. G. Bentrude, E. R. Hansen, W. A. Khan and P. E. Rogers, α vs. β Scission in Reactions of Alkoxy and Thiyl Radicals with Diethyl Alkylphosphonites, *J. Am. Chem. Soc.*, 1972, **94**, 2867–2868, DOI: [10.1021/ja00763a059](#); (h) W. G. Bentrude, E. R. Hansen, W. A. Khan, T. B. Min and P. E. Rogers, Free-Radical Chemistry of Organophosphorus Compounds. III. α vs. β Scission in Reactions of Alkoxy and Thiyl Radicals with Trivalent Organophosphorus Derivatives, *J. Am. Chem. Soc.*, 1973, **95**, 2286–2293, DOI: [10.1021/ja00788a031](#); (i) W. G. Bentrude, J. H. Hargis, N. A. Johnson, T. B. Min, P. E. Rusek Jr, H. W. Tan and R. A. Wielesek, Free radical studies of organophosphorus compounds. 7. The stereochemistries of alkoxy and thiyl radical oxidations of phosphines and cyclic phosphites. Possible permutational isomerization modes for phosphoranyl radicals, *J. Am. Chem. Soc.*, 1976, **98**, 5348–5357, DOI: [10.1021/ja00433a049](#); (j) W. G. Bentrude, J. J. L. Fu and P. E. Rogers, Free-radical chemistry of organophosphorus compounds. IV. Polar, bond strength, and resonance effects on phosphoranyl radical formation, *J. Am. Chem. Soc.*, 1973, **95**, 3625–3635, DOI: [10.1021/ja00792a027](#); (k) W. G. Bentrude and P. E. Rogers, Free-radical chemistry of organophosphorus compounds. V. Reactions of thiyl radicals with alkylphosphonites. Effects of activation energies and A values on the α vs. β scission competition, *J. Am. Chem. Soc.*, 1976, **98**, 1674–1677, DOI: [10.1021/ja00423a006](#); (l) W. G. Bentrude and F. R. Hartley, *The Chemistry of Organophosphorus Compounds*, Wiley, Chichester, 1990, vol. 1, pp. 531–566, DOI: [10.1002/9780470034439m](#); (m) S. Marque and P. Tordo, Reactivity of Phosphorus Centered Radicals, *Top. Curr. Chem.*, 2005, **250**, 43–76, DOI: [10.1007/b100981](#).
 - 12 (a) H. Guo, Y. C. Fan, Z. Sun, Y. Wu and O. Kwon, Phosphine organocatalysis, *Chem. Rev.*, 2018, **118**, 10049–10293, DOI: [10.1021/acs.chemrev.8b00081](#); (b) C. Xie, A. J. Smaligo, X. R. Song and O. Kwon, Phosphorus-based catalysis, *ACS Cent. Sci.*, 2021, **7**, 536–558, DOI: [10.1021/acscentsci.0c01493](#); (c) M. C. Fu, J. X. Wang and R. Shang, Triphenylphosphine-Catalyzed Alkylative Iododecarboxylation with Lithium Iodide under Visible Light, *Org. Lett.*, 2020, **22**, 8572–8577, DOI: [10.1021/acs.orglett.0c03173](#).
 - 13 (a) N. L. Dunn, M. Ha and A. T. Radosevich, Main group redox catalysis: reversible PIII/PV redox cycling at a phosphorus platform, *J. Am. Chem. Soc.*, 2012, **134**, 11330–11333, DOI: [10.1021/ja302963p](#); (b) K. D. Reichl, N. L. Dunn, N. J. Fastuca and A. T. Radosevich, Biphilic organophosphorus catalysis: Regioselective reductive transposition of allylic bromides via PIII/PV redox cycling, *J. Am. Chem. Soc.*, 2015, **137**, 5292–5295, DOI: [10.1021/jacs.5b01899](#); (c) K. Lee, A. V. Blake, A. Tanushi, S. M. McCarthy, D. Kim, S. M. Loria, C. M. Donahue, K. D. Spielvogel, J. M. Keith, S. R. Daly and A. T. Radosevich, Validating the biphilic hypothesis of nontrigonal phosphorus (III) compounds, *Angew. Chem., Int. Ed.*, 2019, **58**, 6993, DOI: [10.1002/anie.201901779](#); (d) G. Li, T. V. Nykaza, J. C. Cooper, A. Ramirez, M. R. Luzung and A. T. Radosevich, An Improved PIII/PV = O-Catalyzed Reductive C–N Coupling of Nitroaromatics and Boronic Acids by Mechanistic Differentiation of Rate- and Product-Determining Steps, *J. Am. Chem. Soc.*, 2020, **142**, 6786–6799, DOI: [10.1021/jacs.0c01666](#); (e) J. A. Rossi-Ashton, A. K. Clarke, W. P. Unsworth and R. J. Taylor, Phosphoranyl radical fragmentation reactions driven by photoredox catalysis, *ACS Catal.*, 2020, **10**, 7250–7261, DOI: [10.1021/acscatal.0c01923](#); (f) G. Zeng, S. Maeda, T. Taketsugu and S. Sakaki, Theoretical study of hydrogenation catalysis of phosphorus compound and prediction of catalyst with high activity and wide application scope, *ACS Catal.*, 2016, **6**, 4859–4870, DOI: [10.1021/acscatal.5b02968](#).
 - 14 R. Romeo, L. A. Wozniak and C. Chatgililoglu, Radical-based reduction of phosphine sulfides and phosphine selenides by $(\text{Me}_3\text{Si})_3\text{SiH}$, *Tetrahedron Lett.*, 2000, **41**, 9899–9902, DOI: [10.1016/S0040-4039\(00\)01759-7](#).
 - 15 J. M. Lopp and V. A. Schmidt, Intermolecular Phosphite-Mediated Radical Desulfurative Alkene Alkylation Using Thiols, *Org. Lett.*, 2019, **21**, 8031–8036, DOI: [10.1021/acs.orglett.9b03018](#).
 - 16 J. M. Lopp and V. A. Schmidt, Experimental and Computational Investigations into Trivalent Phosphorous-Mediated Radical Thiol Desulfurization, *ChemRxiv*,



- Cambridge Open Engage, Cambridge, 2021, preprint, DOI: [10.26434/chemrxiv-2021-fgq3c](https://doi.org/10.26434/chemrxiv-2021-fgq3c).
- 17 (a) C. Chatgililoglu, *Organosilanes in radical chemistry*, John Wiley & Sons, 2004, DOI: [10.1002/0470024755](https://doi.org/10.1002/0470024755); (b) C. Chatgililoglu, Organosilanes as radical-based reducing agents in synthesis, *Acc. Chem. Res.*, 1992, **25**, 188–194, DOI: [10.1021/ar00016a003](https://doi.org/10.1021/ar00016a003); (c) C. Chatgililoglu and V. I. Timokhin, Silanes as reducing reagents in radical chemistry, in *Encyclopedia of Radicals in Chemistry, Biology and Materials*, 2012, DOI: [10.1002/9781119953678.rad074](https://doi.org/10.1002/9781119953678.rad074); (d) H. Janbazi, C. Schulz, I. Wlokas, H. Wang and S. Peukert, Thermochemistry of organosilane compounds and organosilyl radicals, *Proc. Combust. Inst.*, 2021, **38**, 1259–1267, DOI: [10.1016/j.proci.2020.06.114](https://doi.org/10.1016/j.proci.2020.06.114); (e) B. Giese, B. Kopping and C. Chatgililoglu, Tris(trimethylsilyl) silane as mediator in organic synthesis *via* radicals, *Tetrahedron Lett.*, 1989, **30**, 681–684, DOI: [10.1016/S0040-4039\(01\)80281-1](https://doi.org/10.1016/S0040-4039(01)80281-1); (f) M. Lesage, C. Chatgililoglu and D. Griller, Tris(trimethylsilyl) silane: A catalyst for radical mediated reduction reactions, *Tetrahedron Lett.*, 1989, **30**, 2733–2734, DOI: [10.1016/S0040-4039\(00\)99111-1](https://doi.org/10.1016/S0040-4039(00)99111-1); (g) M. Ballestri, C. Chatgililoglu, K. B. Clark, D. Griller, B. Giese and B. Kopping, Tris(trimethylsilyl) silane as a radical-based reducing agent in synthesis, *J. Org. Chem.*, 1991, **56**, 678–683, DOI: [10.1021/jo00002a035](https://doi.org/10.1021/jo00002a035); (h) C. Chatgililoglu, A. Guerrini and M. Lucarini, The trimethylsilyl substituent effect on the reactivity of silanes. Structural correlations between silyl radicals and their parent silanes, *J. Org. Chem.*, 1992, **57**, 3405–3409, DOI: [10.1021/jo00038a031](https://doi.org/10.1021/jo00038a031); (i) C. Chatgililoglu, The Tris(trimethylsilyl) silane/Thiol Reducing System: A Tool for Measuring Rate Constants for Reactions of Carbon-Centered Radicals with Thiols, *Helv. Chim. Acta*, 2006, **89**, 2387–2398, DOI: [10.1002/hlca.200690221](https://doi.org/10.1002/hlca.200690221); (j) C. Chatgililoglu, (Me₃Si)₃SiH: Twenty Years After Its Discovery as a Radical-Based Reducing Agent, *Chem.-Eur. J.*, 2008, **14**, 2310–2320, DOI: [10.1002/chem.200701415](https://doi.org/10.1002/chem.200701415); (k) C. Chatgililoglu, C. Ferreri, Y. Landais and V. I. Timokhin, Thirty years of (TMS)₃SiH: A milestone in radical-based synthetic chemistry, *Chem. Rev.*, 2018, **118**, 6516–6572, DOI: [10.1021/acs.chemrev.8b00109](https://doi.org/10.1021/acs.chemrev.8b00109); (l) C. Chatgililoglu and J. Lalevée, Recent applications of the (TMS)₃SiH radical-based reagent, *Molecules*, 2012, **17**, 527–555, DOI: [10.3390/molecules17010527](https://doi.org/10.3390/molecules17010527); (m) C. Chatgililoglu and V. I. Timokhin, 2 Silyl Radicals in Chemical Synthesis, *Adv. Organomet. Chem.*, 2008, **57**, 117–181.
 - 18 Ac-N-Cys-OMe was prepared according to reported procedure: Y. A. Lin, J. M. Chalker, N. Floyd, G. J. Bernardes and B. G. Davis, Allyl sulfides are privileged substrates in aqueous cross-metathesis: application to site-selective protein modification, *J. Am. Chem. Soc.*, 2008, **130**, 9642–9643, DOI: [10.1021/ja8026168](https://doi.org/10.1021/ja8026168).
 - 19 C. Chatgililoglu, D. Griller and M. Lesage, Tris(trimethylsilyl) silane. A new reducing agent, *J. Org. Chem.*, 1988, **53**, 3641–3642, DOI: [10.1021/jo00250a051](https://doi.org/10.1021/jo00250a051).
 - 20 https://www.sigmaaldrich.com/deepweb/assets/sigmaaldrich/marketing/global/documents/411/888/thermal_initiators.pdf, accessed 28 November 2022.
 - 21 M. B. Cole Jr, An “azo” substitute for benzoyl peroxide in polymerizing glycol methacrylate (GMA), *Micron*, 1979, **10**, 135–137, DOI: [10.1016/0047-7206\(79\)90005-0](https://doi.org/10.1016/0047-7206(79)90005-0).
 - 22 (a) H. Subramanian, R. Moorthy and M. P. Sibi, Thiyl Radicals: From Simple Radical Additions to Asymmetric Catalysis, *Angew. Chem., Int. Ed.*, 2014, **53**, 13660–13662, DOI: [10.1002/anie.201408781](https://doi.org/10.1002/anie.201408781); (b) F. Denes, M. Pichowicz, G. Povie and P. Renaud, Thiyl radicals in organic synthesis, *Chem. Rev.*, 2014, **114**, 2587–2693, DOI: [10.1021/cr400441m](https://doi.org/10.1021/cr400441m).
 - 23 See the ESI† for details.
 - 24 Use of an aromatic phosphine (PPh₃) under the optimized conditions afforded 98% (toluene) and 92% (1,4-dioxane) conversions.
 - 25 Y.-R. Luo, *Comprehensive Handbook of Chemical Bond Energies*, CRC Press, 1st edn, 2007, DOI: [10.1201/9781420007282](https://doi.org/10.1201/9781420007282), ISBN 9780429128684.
 - 26 (a) J. F. Brazeau, Tris(trimethylsilyl) silane (TTMSS), *Synlett*, 2007, 1972–1973, DOI: [10.1055/s-2007-984873](https://doi.org/10.1055/s-2007-984873); (b) L. J. Laarhoven, P. Mulder and D. D. Wayner, Determination of bond dissociation enthalpies in solution by photoacoustic calorimetry, *Acc. Chem. Res.*, 1999, **32**, 342–349, DOI: [10.1021/ar9703443](https://doi.org/10.1021/ar9703443).
 - 27 J. Lalevée, N. Blanchard, B. Graff, X. Allonas and J. P. Fouassier, Tris(trimethylsilyl) silyl *versus* tris(trimethylsilyl) germyl: radical reactivity and oxidation ability, *J. Organomet. Chem.*, 2008, **693**, 3643–3649, DOI: [10.1016/j.jorganchem.2008.08.039](https://doi.org/10.1016/j.jorganchem.2008.08.039).
 - 28 G. Avar and W. P. Neumann, *J. Organomet. Chem.*, 1977, **131**, 207–214, DOI: [10.1016/S0022-328X\(00\)93403-X](https://doi.org/10.1016/S0022-328X(00)93403-X).
 - 29 (a) C. Walling and R. Rabinowitz, The reaction of thiyl radicals with trialkyl phosphites¹, *J. Am. Chem. Soc.*, 1957, **79**, 5326, DOI: [10.1021/ja01576a077](https://doi.org/10.1021/ja01576a077); (b) F. W. Hoffmann, R. J. Ess, T. C. Simmons and R. S. Hanzel, The desulfurization of mercaptans with trialkyl phosphites, *J. Am. Chem. Soc.*, 1956, **78**, 6414, DOI: [10.1021/ja01605a034](https://doi.org/10.1021/ja01605a034); (c) C. Walling and M. S. Pearson, Some radical reactions of trivalent phosphorus derivatives with mercaptans, peroxides, and olefins. A new radical cyclization, *J. Am. Chem. Soc.*, 1964, **86**, 2262–2266, DOI: [10.1021/ja01065a032](https://doi.org/10.1021/ja01065a032).
 - 30 (a) D. Leca, L. Fensterbank, E. Lacôte and M. Malacria, Recent advances in the use of phosphorus-centered radicals in organic chemistry, *Chem. Soc. Rev.*, 2005, **34**, 858–865, DOI: [10.1039/B500511F](https://doi.org/10.1039/B500511F); (b) W. G. Bentrude, Phosphoranyl radicals-their structure, formation, and reactions, *Acc. Chem. Res.*, 1982, **15**, 117–125, DOI: [10.1021/ar00076a004](https://doi.org/10.1021/ar00076a004); (c) K. Luo, W. C. Yang and L. Wu, Photoredox catalysis in organophosphorus chemistry, *Asian J. Org. Chem.*, 2017, **6**, 350–367, DOI: [10.1002/ajoc.201600512](https://doi.org/10.1002/ajoc.201600512); (d) C. Lopin, G. Gouhier, A. Gautier and S. R. Piettre, Phosphonyl, Phosphonothioyl, Phosphonodithioyl, and Phosphonotrithioyl Radicals: Generation and Study of Their Addition onto Alkenes, *J. Org. Chem.*, 2003, **68**, 9916–9923, DOI: [10.1021/jo0348064](https://doi.org/10.1021/jo0348064).



- 31 (a) M. Ballestri, C. Chatgililoglu and G. Seconi, (Me₃Si)₃SiS. H.: a new radical-based reducing agent, *J. Organomet. Chem.*, 1991, **408**, C1–C4, DOI: [10.1016/0022-328X\(91\)80107-U](https://doi.org/10.1016/0022-328X(91)80107-U); (b) C. Chatgililoglu, Structural and chemical properties of silyl radicals, *Chem. Rev.*, 1995, **95**, 1229–1251, DOI: [10.1021/cr00037a005](https://doi.org/10.1021/cr00037a005).
- 32 C. Chatgililoglu and K. D. Asmus, *Sulfur-centered reactive intermediates in chemistry and biology*, Springer Science & Business Media, 2013, vol. 197, DOI: [10.1007/978-1-4684-5874-9](https://doi.org/10.1007/978-1-4684-5874-9).

