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Metal-free reductive desulfurization of C-sp³-substituted thiols using phosphite catalysis†

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Phosphines and phosphites are critical tools for non-metal desulfurative methodologies due to the strength of the P=S bond. An overarching premise in these methods has been that stoichiometric (or excess) P(III) reagent is required for reactivity. Despite decades of research, a desulfurative process that is catalytic in phosphine/phosphite has not been reported. Here, we report the successful merging of two thermal radical processes: the desulfurization of unactivated and activated alkyl thiols and the reduction of P(V) = S to P(III) by reaction with a silyl radical species. We employ catalytic trimethyl phosphite, catalytic azo-bis(cyclohexyl)nitrile, and two equivalents of tris(trimethylsilyl)silane as the stoichiometric reductant and sulfur atom scavenger. This method is tolerant of common organic functional groups and affords products in good to excellent yields.

Introduction

Desulfurization is an important process¹ for diverse applications from organic synthesis² and chemical protein synthesis² to crude oil processing and refinement.³ Numerous reports describe transition metal-based reduction reactions, including Ni,⁴ Co,⁵ W,⁶ Mo,⁷ and others.⁸ Unfortunately, critical limitations remain, which include pyrophoricity,^{4,5} toxic H₂S gas evolution,⁴ reactivity with other S-functional groups,^{4,6} and the need for specialized reaction setups.⁵ Additionally, when the desulfurized products are deployed in living systems, trace metal contamination can be problematic.⁹ Photoredox strategies^{10a,b} avoid functional group incompatibility^{10a} and may enable further functionalization of the substrate; however, they do require the use of expensive rare earth transition metal complexes, which may limit industrial scale applications. With only two exceptions,^{10d,e} modern and non-metal strategies for thiol desulfurization^{10,11} employ excess phosphine/phosphite reagents, whose aquatic toxicity⁹ has limited their utility in >kg-scale processes. The last several years have seen a surge of interest in the potential of P(III)/P(V) catalysis,¹² largely focused on strategies to facilitate polar P(III)/P(V) = O cycles.¹³ We were intrigued by the potential of employing P(III)/P(V) = S catalysis for desulfurative radical transformations. Here, we present the first step along this journey, a metal-free desulfurization of unactivated thiols by P(III)/P(V) = S catalysis. We achieve this

reactivity by leveraging the affinity of silyl radicals toward sulfur¹⁴ to achieve *in situ* regeneration of phosphites by reduction of the P=S bond while sequestering the S atom to avoid H₂S formation.

Results and discussion

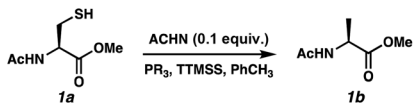
Reports as recent as 2021 have asserted that stoichiometric phosphine is a requirement for radical desulfurization reactions.^{12b} Both phosphines and phosphites are well established reagents for desulfurization reactions,^{11,15} with the electronics of the R group on PR₃ being key to reactivity.¹⁶ Phosphine sulfides react with super silyl hydride (TTMSS)¹⁷ under free radical conditions to produce phosphines in good yields.¹⁴ Unlike P(III)/P(V) = O redox cycles, which occur *via* polar pathways, the P(III)/P(V) = S reduction occurs *via* a radical mechanism.¹⁴ Thus, we envisioned that this cycle would merge seamlessly with a radical desulfurization mechanism. We employed Ac-N-Cys-OMe¹⁸ (**1a**) as a model substrate to investigate the feasibility of the catalytic transformation in the presence of various phosphines/phosphites with TTMSS as a phosphine sulfide reductant¹⁴ and/or terminal hydrogen atom donor¹⁹ (Table 1). We selected azo-bis(cyclohexyl)nitrile (ACHN) as the initiator because of its 10 h half-life²⁰ and to avoid the formation of oxygenated by-products commonly observed in the presence of organic peroxide initiators.²¹ Because of the electrophilic nature of thiyl radicals,²² we began our studies with 10 mol% of the electron-rich tris(dimethylanion)phosphine ((Me₂N)₃P).²³ Only a single turnover was observed, with the desired product (**1b**) formed in 21% conversion (entry 1). We limited reaction time to 24 h for the purposes of this screen. The less basic tri-*tert*-butylphosphine was markedly better with 65%

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Table 1 Establishing PR₃ reactivity under catalytic conditions^a


Entry	PR ₃ (equiv.)	TTMSS (equiv.)	ACHN (equiv.)	Temp. (°C)	Time (h)	Conversion ^b (%)
1	(Me ₂ N) ₃ P (0.1)	1	0.1	80	24	21
2	^t Bu ₃ P (0.1)	1	0.1	80	24	65
3	ⁿ Bu ₃ P (0.1)	1	0.1	80	24	54
4	P(Cy) ₃ (0.1)	1	0.1	80	24	38
5	P(OMe) ₃ (0.1)	1	0.1	80	24	80
6	P(OMe)₃ (0.2)	2	0.1	88	16	>99
7	P(OMe) ₃ (0.2)	2	—	88	16	0
8	—	2	0.1	88	16	0
9	P(OMe) ₃ (0.2)	—	0.1	88	16	7
10	P(OMe) ₃ (0.2)	2	0.1	25	16	0

^a Reactions were performed with 1 equiv. **1a** (0.05 M) in degassed PhMe. ^b Conversion = consumption of **1a**; based on ¹H-NMR integration.

conversion (entry 2), while the less hindered and less electron-rich tri(*n*-butyl)phosphine led to 54% conversion (entry 3). Interestingly, tri-cyclohexylphosphine, which bridges these steric and electronic properties, was even less effective (38% conversion, entry 4). Trimethyl phosphite balanced these effects, improving the conversion to 80% (entry 5). Increasing the phosphite loading to 20 mol% led to complete conversion by 16 h (entry 6).²⁴ Control experiments were consistent with the radical pathway we envisioned for this chemistry. Omission of the radical initiator (entry 7) or the phosphite (entry 8) abolished reactivity. In the absence of TTMSS, slow conversion was observed and did not reach 20% because the substrate thiol must serve as the H-atom donor for C–H bond formation (entry 9). Finally, heating is required for the reduction of the P=S bond by TTMSS, as evidenced by entry 10.

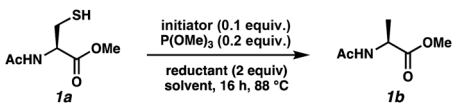
We proceeded with trimethyl phosphite at 88 °C for 16 h while screening different reaction parameters such as initiator,

reductant, and solvent (Table 2). Using different initiators like VA-044 or Luperox A98 led to large decreases in conversion (entries 1–2). Dicumyl peroxide led to a 3-fold increase in conversion percentage relative to VA-044 at the same temperature (entry 3), but it was still inferior to ACHN.

Alternative hydrogen atom/silyl radical donors had a deleterious effect on conversion. When tri-isopropylsilane (TIPS) was used, the reaction takes place, but the phosphite is not reduced (Table 2, entry 4). This apparent limited efficiency of TIPS compared to TTMSS is logical given the difference in their Si–H bond-dissociation energies (BDEs) with trialkylsilanes ranging from (90–96 kcal mol^{−1})²⁵ compared to the Si–H in TTMSS at 79 kcal mol^{−1}.²⁶ The instability of the TIPS[•] interferes with the turnover of the phosphine sulfide, presumably because H-atom abstraction from the substrate or solvent occurs more rapidly than Si–S bond formation. Alternatively, the *i*-Pr group is slow to migrate after α -scission. Consistent with this reasoning, and with the similar properties of Ge and Si,²⁷ Et₃GeH (86 kcal mol^{−1}) was able to effect ~1 phosphine sulfide turnover in 16 h, proceeding with 43% conversion (entry 5). The best alternate P=S reductant was *n*-Bu₃SnH (74 kcal mol^{−1}), which still reduced the phosphine sulfide more slowly than TTMSS (entry 6).²⁸ The respective BDEs of TTMSS and *n*-Bu₃SnH indicate that the Si[•] and Sn[•] species are similarly stable and the rate difference likely arises from differences in affinity to sulfur or the ability of TTMSS to irreversibly sequester the S-atom *via* an additional α -fragmentation and 1,2-group shift (see Scheme 2).

In terms of solvent, THF and MeCN, each of which were performed at 88 °C using a sealed microwave tube, were less competent than toluene (entries 7–8). Interestingly, the reaction proceeded to complete conversion with 1,4-dioxane as solvent (entry 9). Indeed, the success of 1,4-dioxane increases the scope of this method to include polar substrates.

With two sets of established conditions in hand, we evaluated the scope of our efficient reductive desulfurization reaction on different relevant substrates, including short peptides and small molecule substrates that would form intermediate C-

Table 2 Probing the impact of radical initiator, reductant, and solvent^a


Entry	Initiator	Reductant	Solvent	Conversion ^b (%)
1	VA-044	TTMSS	H ₂ O/MeCN	44
2	Luperox A98	TTMSS	PhMe	22
3	Dicumyl peroxide	TTMSS	PhMe	65
4	ACHN	TIPS	PhMe	7
5	ACHN	Et ₃ GeH	PhMe	43
6	ACHN	ⁿ Bu ₃ SnH	PhMe	63
7	ACHN	TTMSS	MeCN	62
8	ACHN	TTMSS	THF	82
9	ACHN	TTMSS	1,4-Dioxane	>99

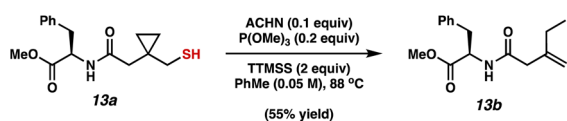
^a Reactions were performed with 1 equiv. **1a** (0.05 M) in degassed PhMe.

^b Conversion = consumption of **1a**; based on ¹H-NMR integration.

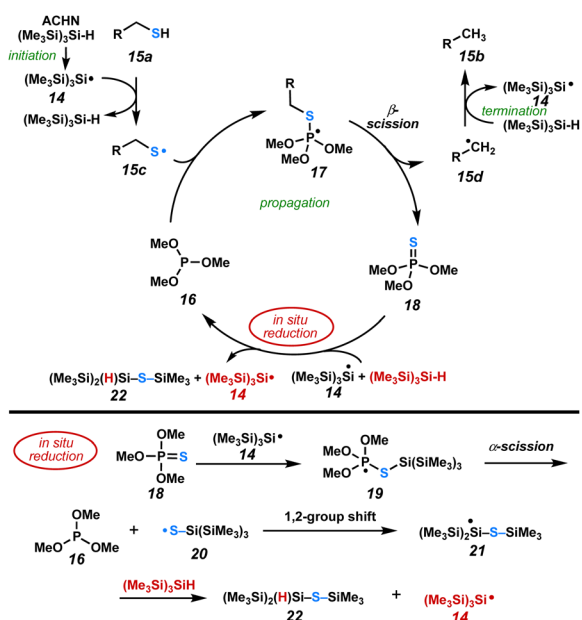


To probe the amenability of the desulfurization to large scale conditions, we performed the reaction on 4-methoxybenzyl mercaptan (**12a**). On 6.84 mmol scale in 1,4-dioxane, 100% conversion was observed in 3 h, with clean reactivity and no unexpected by-product formation (Scheme SI-01†). To probe the compatibility of various functional groups with the reaction conditions, we again employed 4-methoxybenzyl mercaptan (**12a**) in toluene along with different additives. The additives shown in Table 4 were compatible with the reaction conditions, showing no impact on the additive. In most cases, good to excellent reactivity of **12a** was maintained. In entries 6 and 8, the desulfurization reaction did not occur. Alcohol additives had no negative effect on the desulfurization but were silylated. Aldehyde and ketone additives led to an unknown by-product and interfered somewhat with the desulfurization (see Table SI-02†).

A radical clock experiment (Scheme 1) was performed with cyclopropane-containing thiol, **13a**, which underwent ring opening to afford olefin **13b** in 55% yield. This indicates that a carbon-centered radical is an intermediate in this reaction. Based on our observations, we hypothesize that reduction of the phosphine sulfide is rate-limiting and proceeds *via* the radical pathway described by Chatgililoglu.¹⁴ The postulated overall radical chain mechanism for the catalytic reductive desulfurization is illustrated in Scheme 2.



Scheme 1 Cyclopropylcarbinyl radical rearrangement product supports a radical pathway.



Scheme 2 Mechanism of catalytic desulfurization.

Upon heating, ACHN releases N_2 to initiate the reaction by forming the C-centered cyclohexyl nitrile radical. H-atom abstraction from TTMSS leads to super silyl radical **14**. H-atom abstraction from the sulfhydryl group in **15a** forms thiyl radical **15c** and regenerates TTMSS.²² Thiyl radical **15c** reacts with phosphite²⁹ **16** to generate phosphoranyl radical intermediate **17**.³⁰ Tetravalent P-centered radical **17** undergoes β -scission to produce trimethyl thiophosphate **18** and the carbon centered radical **15d**. H-atom transfer from TTMSS to **15d** produces the desulfurized product **15b** and regenerates TTMSS radical **14**. Trimethyl thiophosphate **18** reacts with silyl radical **14** to form phosphoranyl radical intermediate **19**. Subsequent α -scission regenerates the catalytically active phosphite (**16**) to be reused in the propagation step. Notably, this α -scission is facilitated by the thiophilic nature of the silane (Si-S 148.4 *versus* P-S 105.6 kcal mol⁻¹).²⁵ The resulting S-centered radical **20** induces a 1,2-group shift³¹ of a TMS group from silicon to sulfur, furnishing silyl radical **21** that reacts with a second equivalent of TTMSS to afford TTMSS radical **14** and silyl sulfide **22**.³² This completes the *in situ* reduction process and permits propagation of the radical cascade. GCMS analysis of the crude reaction mixture indicates the presence of a small amount of phosphine sulfide **18** and multiple TTMSS-related peaks during the reaction.

Conclusions

In summary, we have established reaction conditions for P(III)/P(V) = S catalysis that proceeds through a tetravalent P-centered radical rather than through a pentavalent P(V) intermediate as is observed for P=O reduction. The reductive desulfurization reaction sequesters the S atom, avoids the use of rare earth and transition metals, and affords desulfurized products in good yields. The method has been tested across a range of substrate types including thiols that lead to benzylic, primary, secondary, and tertiary alkyl C-centered radicals. Importantly, good functional group compatibility is observed, including free carboxylic acid, ester, carbamate, amide, nitrile, aryl halide, amine, isocyanate, thioether functionalities. Steroid, carbohydrate, amino acid, pyridine, and benzene scaffolds are compatible with the reaction. No special setup is required for this catalytic method, so it can be readily employed and adapted. This study expands the scope of metal-free and thiol-additive-free desulfurization methods while establishing the viability of P(III)/P(V) = S catalysis under radical conditions. Efforts to harness this catalysis strategy to effect other transformations and develop new phosphorus-based catalysts are ongoing.

Data availability

The datasets supporting this article have been uploaded as part of the ESI.†

Author contributions

All authors have contributed to this research.



Conflicts of interest

There are no conflicts to declare.

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