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Functionalization of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ nanoparticles with Cu(I)-thiosemicarbazone complex as a robust and efficient heterogeneous nanocatalyst for *N*-arylation of N-heterocycles with aryl halides†

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In this research, the functionalized silica-coated magnetite nanoparticles with Cu(I)-thiosemicarbazone complex ($\text{Fe}_3\text{O}_4@\text{SiO}_2-\text{[CuL]}$) has been designed and synthesized as a magnetically retrievable nanocatalyst. Different techniques were employed to characterize the structure of $\text{Fe}_3\text{O}_4@\text{SiO}_2-\text{[CuL]}$ comprising FT-IR, FE-SEM, TEM, DLS, XRD, EDX, TGA, AAS, and VSM analysis. The catalytic performance of $\text{Fe}_3\text{O}_4@\text{SiO}_2-\text{[CuL]}$ was perused in Ullmann-type *N*-arylation of nucleobases, xanthines, and other N-heterocycles with diverse aryl halides which acquired the desired *N*-aryl products in good to excellent yields. $\text{Fe}_3\text{O}_4@\text{SiO}_2-\text{[CuL]}$ is a thermal and chemical stable, easy to prepare and recyclable, inexpensive, and ecofriendly catalyst that needs no additional additive or ligand as promoters. This catalyst could be separated without difficulty by a simple magnet and reused for at least seven sequential runs without a significant decline in its catalytic performance.

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1. Introduction

Indeed, the C–N bond formation is one of the most significant reactions in organic synthesis since it provides numerous attractive and important substrates which have plenty of utilizations in different scientific areas and human life. Amongst, the N-heterocyclic compounds have found immense research interest especially due to their key roles in chemistry, medicine, and material sciences.^{1,2} Therefore, the synthesis of N-heterocycles and modification of the N-heterocyclic cores are of massive research interest. The *N*-arylation of the N-heterocycles is an important approach to modify the parent N-heterocyclic unit since it not only increases the lipophilicity nature of the N-heterocyclic unit but also can mask the nucleophilic nitrogen. The *N*-arylated N-heterocycles have found diverse applications in agriculture, organic, and medicinal chemistry. Many herbicides and approved drugs are involved in the *N*-aryl N-heterocycle residues in their structures (Fig. 1). The *N*-aryl nucleobases are used as agrochemicals,³ molecular tools, and probes to study biological systems.⁴ They exhibit some biological profiles such as anticancer,⁵ antibacterial,⁶ and

antiviral.⁷ Some *N*-aryl xanthine derivatives inhibit proliferation in T47D tumor cells and phosphoinositide 3-kinases (PI3Ks).⁸ The *N*-aryl azoles and *N*-aryl indoles are useful precursors for the synthesis of drugs, natural products, and bioactive agents.^{2,9–11} Furthermore, they display various biological activities such as anti-HIV1,¹² antiallergic,¹³ melatonin receptor MT1 agonists,¹⁴ COX-2 inhibitors,¹⁵ and antipsychotic agents.¹⁶

Owing to the unique properties and applications of *N*-aryl N-heterocycles, different synthetic methodologies have been established to afford these important compounds.^{17–19} These general approaches are involved (i) the nucleophilic aromatic substitution ($S_N\text{Ar}$) reactions of electron-deficient aryl halides with N-heterocycles,²⁰ (ii) the treatment of N-heterocycles with aryl iodonium salts,²¹ (iii) the heterocyclization of appropriate precursors,²² (iv) the *N*-arylation of N-heterocycles with aryl boronic acids,^{8,23} (v) Ullmann-type C–N cross coupling reaction between aryl halides and heterocycles.^{24,25} However, these protocols are often associated with numerous shortcomings such as the use of the stoichiometric amount of Cu catalyst, the relative instability and high cost of boronic acids, the use of toxic and expensive ligands, harsh reaction conditions, unavailability of some precursors, and limited scope. In this connection, the Ullmann-type reaction is traditionally preferred since more starting materials are available compared to other methods. However, the classical Ullmann-type reaction was often carried out at very high temperatures and in the presence of the stoichiometric amount of copper catalyst which is not adequate for the scale-up synthesis.²⁶ Buchwald *et al.*²⁷ and

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† Electronic supplementary information (ESI) available: Data and copies of NMR spectra (^1H and ^{13}C NMR) for synthesized compounds. See DOI: <https://doi.org/10.1039/d3ra06327e>



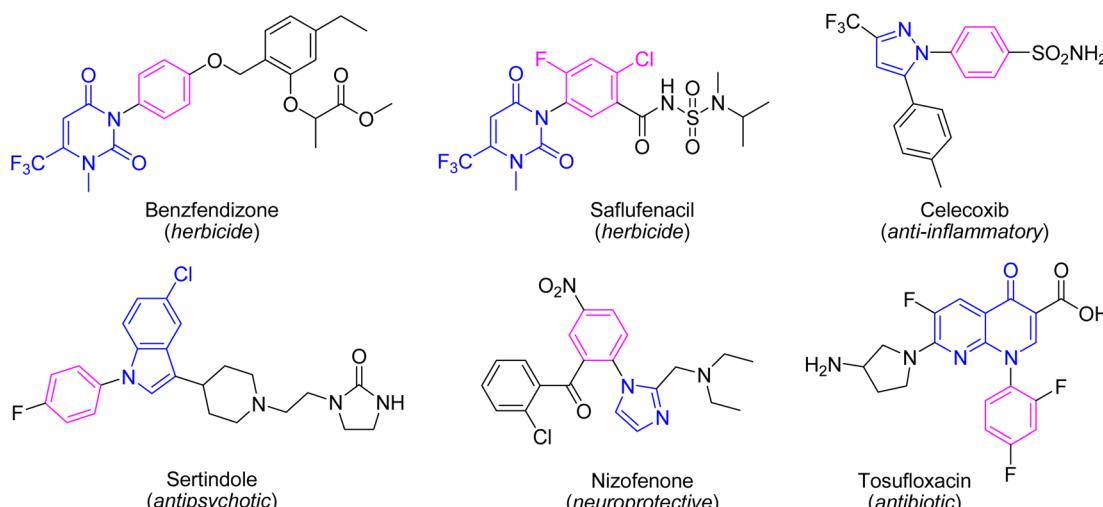
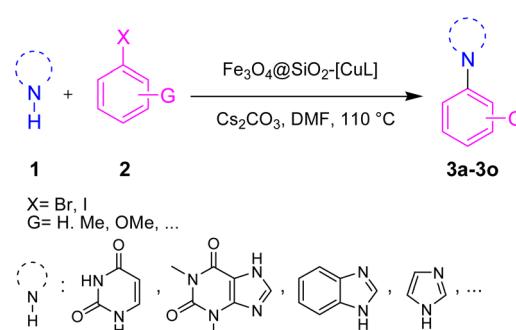


Fig. 1 The structure of some approved drugs and herbicides containing the *N*-aryl *N*-heterocycle residue.

Hartwig *et al.*²⁸ have developed the Pd-catalyzed C–N coupling reactions using aryl halides. However, the use of toxic and expensive Pd catalysts and ligands has restricted its applicability in scale up synthesis and/or industrial processes. Thus, the use of less toxic, safe, inexpensive, and environmentally benign metals instead of Pd is extensively preferred. A striking breakthrough was made by Taillefer²⁹ and Buchwald³⁰ in the Cu-catalyzed C–N cross coupling reactions using aryl halides under mild reaction conditions. In this context, numerous Cu-catalyzed Ullmann-type C–N cross coupling reactions have been described over the years.^{18,31} Nevertheless, several problems were associated with these methods such as less activity because of the low surface area of the catalyst, long reaction time, the use of toxic and sensitive ligands, by-product formation, and the low yields causing a remarkable need for further development of more efficient catalytic systems to achieve the desired *N*-aryl *N*-heterocycles.

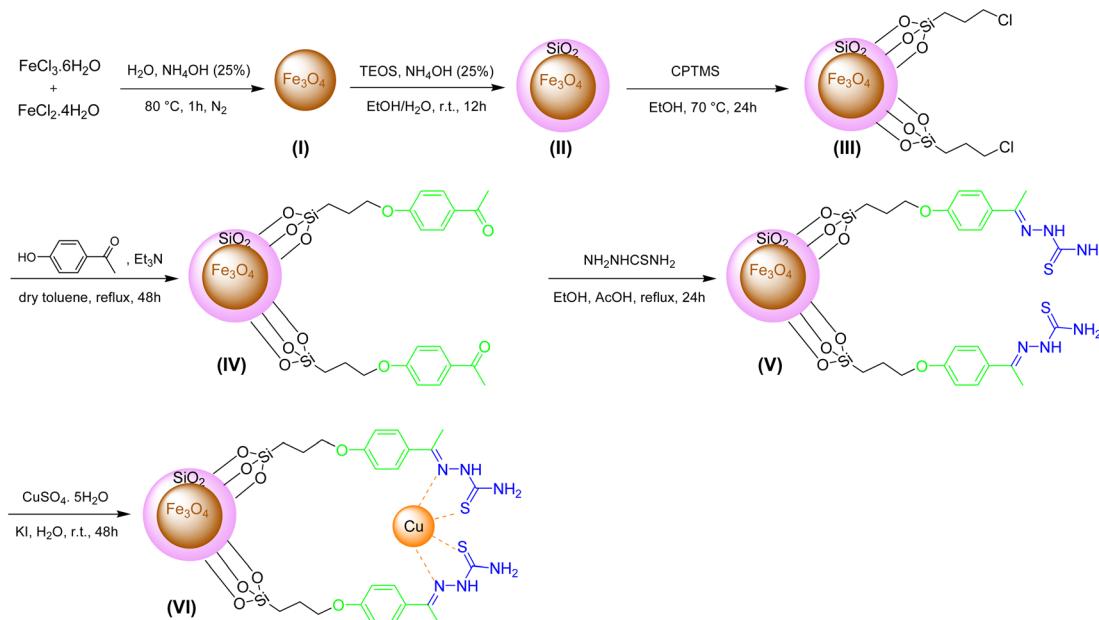
The heterogeneous catalysts exhibit many environmental and economic benefits such as recoverability, reusability, and stability of the catalyst, waste minimization and ease of handling.³² Practically, in some cases, the main products are insoluble in the reaction media and hence the catalyst cannot be recovered by simple centrifugation or filtration. On the other hand, the purification of the product and the work-up process will be difficult without initial separation of the catalyst. To avoid these shortcomings, using magnetic nanoparticle (MNPs) based catalysts would be an ideal strategy since the catalyst could be removed from the reaction media using an external magnet field.³³ To prevent the agglomeration and coagulation of MNPs and conserve their activity, they usually coated with a protecting layer such as carbon, polymer, metal oxide, and silica.³⁴ Among various Fe_3O_4 core-shell composites, silica-coated Fe_3O_4 MNPs ($\text{Fe}_3\text{O}_4@\text{SiO}_2$) are popular, in particular for preparing sustainable and powerful heterogeneous catalysts.^{33,35} It is well known and fully established that silica has large pore volumes, high surface area, and biocompatibility as well as high chemical and thermal

stability. Additionally, the modification of the silica surface could be simply achieved with different functional groups to afford the desired core-shell structures. Moreover, applying an external magnetic field is solely enough to separate the modified $\text{Fe}_3\text{O}_4@\text{SiO}_2$ nanoparticles from the reaction media. In this regard, some Cu complex-functionalized $\text{Fe}_3\text{O}_4@\text{SiO}_2$ nanoparticles were developed for the Ullmann-type *N*-arylation of *N*-heterocycles such as imidazole, benzimidazole, and indole.^{36–38} However, to the best of our knowledge, there has been no report yet on the Ullmann-type *N*-arylation of nucleobase and xanthine derivatives in the presence of Cu complex-functionalized $\text{Fe}_3\text{O}_4@\text{SiO}_2$ nanoparticles. Hence, there is still ongoing interest in preparing and applying such catalysts having broader applicability for Ullmann-type *N*-arylation of *N*-heterocycles. To this end and in pursuit of our research interest on copper-catalyzed C–N bond coupling reactions,^{39–42} herein we introduce the functionalized silica-coated magnetite nanoparticles with Cu(i)-thiosemicarbazone complex ($\text{Fe}_3\text{O}_4@\text{SiO}_2[\text{CuL}]$) as a new magnetically recoverable nano catalyst for Ullmann-type *N*-arylation of nucleobases, xanthines, and other *N*-heterocycles with aryl halides in DMF at 110 °C (Scheme 1).



Scheme 1 *N*-Arylation of *N*-heterocycles with aryl halides using $\text{Fe}_3\text{O}_4@\text{SiO}_2[\text{CuL}]$.





Scheme 2 The procedures for synthesis of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL].

2. Results and discussion

2.1. Preparation of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL]

The preparation of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] as a new magnetically recoverable nano catalyst was achieved through 6 steps as indicated in Scheme 2. The nanoparticles **I–III** were synthesized according to the described procedures in the literature.^{43,44} In brief, the co-precipitation of ferrous and ferric ions under the nitrogen gas afforded Fe_3O_4 nanoparticles (**I**). The reaction was carried out under the nitrogen gas since nanoparticles **I** are susceptible to oxidation.⁴⁵ Hence, magnetically separable Fe_3O_4 nanoparticles (**I**) were acquired *via* the treatment of $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ and $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$ under N_2 gas in NH_4OH (25 wt%) and deionized water. The black solid **I** was attained after washing the solid with water and ethanol and drying it in a vacuum oven (80°C). Tetraethyl orthosilicate (TEOS) was employed for improving the chemical stability of **I** with silica shell according to the Stöber process.^{43,44} The solid **I** and TEOS were reacted together in the presence of NH_4OH (25 wt%) in a deionized $\text{H}_2\text{O}/\text{EtOH}$ solution. After 12 h, the solid was washed several times with deionized water and EtOH . Then, it was dried at 70°C to attain the silica-coated Fe_3O_4 nanoparticles **II** ($\text{Fe}_3\text{O}_4@\text{SiO}_2$). Next, the functionalization of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ (**II**) was done in EtOH using (3-chloropropyl)trimethoxysilane (CPTMS). After 24 h, the solid was collected using an external magnet field, washed with EtOH , and dried in an oven which affords chloropropyl-functionalized $\text{Fe}_3\text{O}_4@\text{SiO}_2$ nanoparticles (**III**). CPTMS was selected for this step since the electrophilic sites were needed for the nucleophilic substitution reactions ($\text{S}_{\text{N}}2$) with 4-hydroxyacetophenone as the chosen nucleophile. In the fourth step, $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -hydroxyacetophenone intermediate (**IV**) was prepared through the $\text{S}_{\text{N}}2$ reaction between 4-hydroxyacetophenone and adduct **III** in refluxing dry toluene.

After 48 h, the solid **IV** was separated by a magnet field, washed with dry toluene, and dried in an oven (50°C).

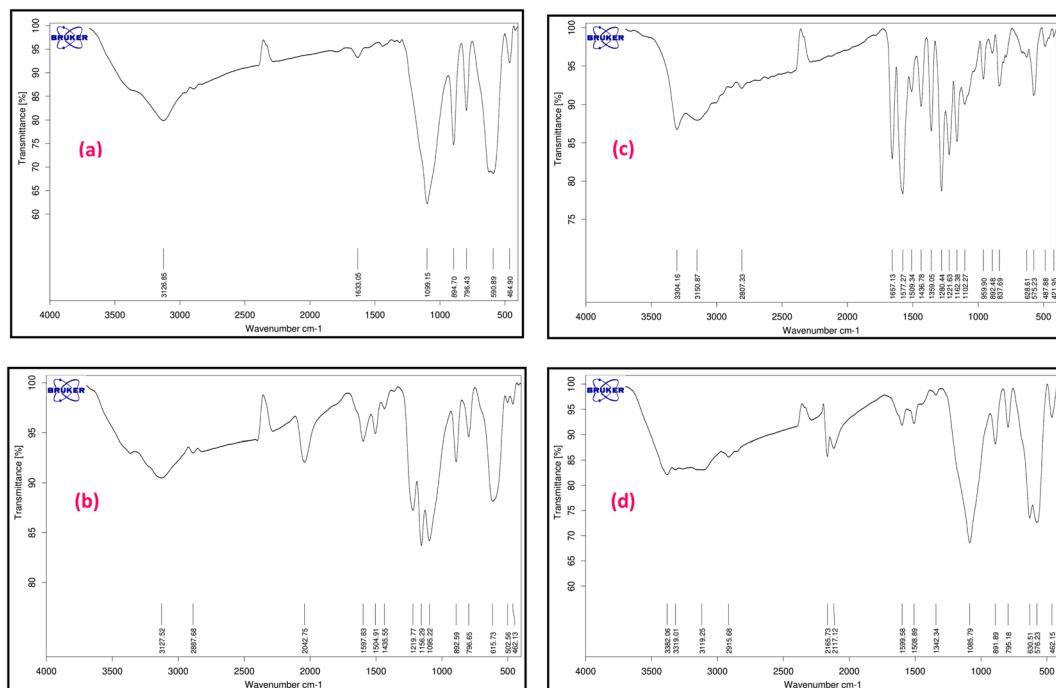
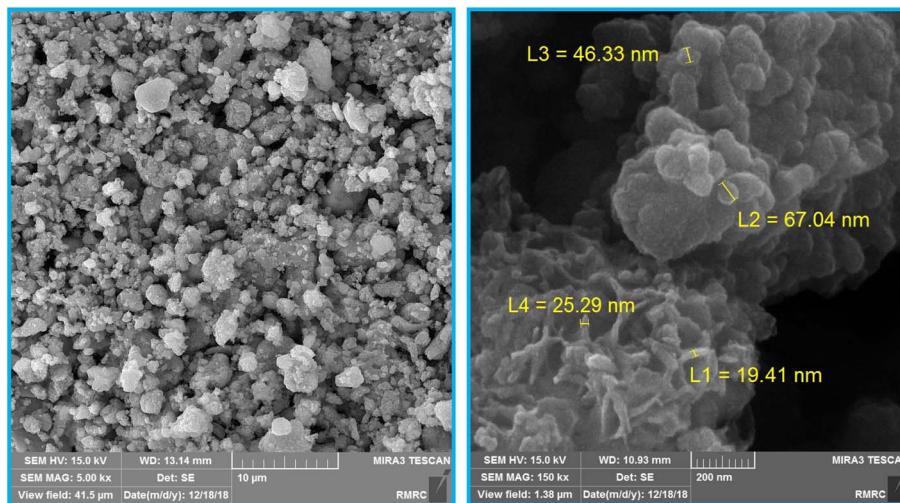
In the subsequent step, a mixture containing thiosemicarbazide, adduct **IV**, and acetic acid (a few drops) was refluxed in EtOH for 24 h. The thiosemicarbazone adduct **V** (creamy solid) was acquired after separating the solid by an external magnet, washing (deionized water and EtOH), and drying in a vacuum oven (50°C). In the last step, thiosemicarbazone adduct **V** was reacted with copper sulfate pentahydrate and potassium iodide in deionized water. After 2 days, the solid was separated using an external magnet, washed with deionized water, and dried in an oven (70°C) to afford $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] (**VI**) as a new magnetically separable nano catalyst.

2.2. Characterization of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL]

The characterization of magnetic $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] was achieved using various techniques including Fourier-transform infrared spectroscopy (FT-IR), field emission scanning electron microscopy (FE-SEM), transmission electron microscopy (TEM), dynamic laser scattering (DLS), X-ray diffraction spectroscopy (XRD), energy dispersive X-ray spectroscopy (EDX), thermogravimetric analysis (TGA), atomic absorption spectroscopy (AAS), and vibrating sample magnetometer (VSM).

Fig. 2 represents the FT-IR spectra of adducts **III–VI**. Apparently, the formation of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] was confirmed by the FT-IR spectra. In Fig. 2a, the absorption bands at 590 and 618 cm^{-1} correspond to the vibration of $\text{Fe}-\text{O}$ (stretching). The peak that appeared at 796 cm^{-1} is ascribed to $\text{C}-\text{Cl}$ vibration (stretching). The peaks in the region of 1099 , 894 , and 464 cm^{-1} are related to the vibration of $\text{Si}-\text{O}-\text{Si}$ (asymmetric and symmetric stretching), and the vibration of $\text{Fe}-\text{O}-\text{Si}$ (stretching), respectively. The peak in the region of 1633 cm^{-1} is



Fig. 2 FT-IR spectra of adducts III (a), IV (b), V (c), and $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] (d).Fig. 3 FE-SEM images of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL].

attributed to the O-H vibration (stretching) absorbed on the surface of Fe nanostructures. The band at 2952 cm^{-1} is assigned to the vibration of aliphatic C-H groups (stretching) which is followed by a broad band related to the O-H vibration in the surface of nanoparticles (stretching). Accordingly, CPTMS was grafted on the surface of magnetic $\text{Fe}_3\text{O}_4@\text{SiO}_2$ nanoparticles. Treatment of 4-hydroxyacetophenone with adduct **III** led to the synthesis of adduct **IV** which shows the new peaks in the FT-IR spectra (Fig. 2b). Hence, the newly appeared band at 1435 cm^{-1} is attributed to the vibration of methyl moiety in 4-hydroxyacetophenone (bending). The bands that appeared at 1156 and 1219 cm^{-1} are related to the presence of C-O stretching. The

bands at 1504 and 1597 cm^{-1} correspond to the vibration of C=C and C=O (stretching), respectively. The aromatic C-H vibration (stretching) is expected to be around 3000 – 3100 cm^{-1} which is covered by the vibration of O-H (stretching) on the surface of nanoparticles. Fig. 2c displays the FT-IR spectra of thiosemicarbazone intermediate **V**. The peaks at 1280 , 1359 , and 1657 cm^{-1} are assigned to the vibration of C=S, C-NH₂, and C=N bonds, respectively. The peaks around 3304 cm^{-1} correspond to the amine groups. A comparison between the FT-IR spectra of intermediate **V** and $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] confirmed the preparation of the desired catalyst (Fig. 2c and d).

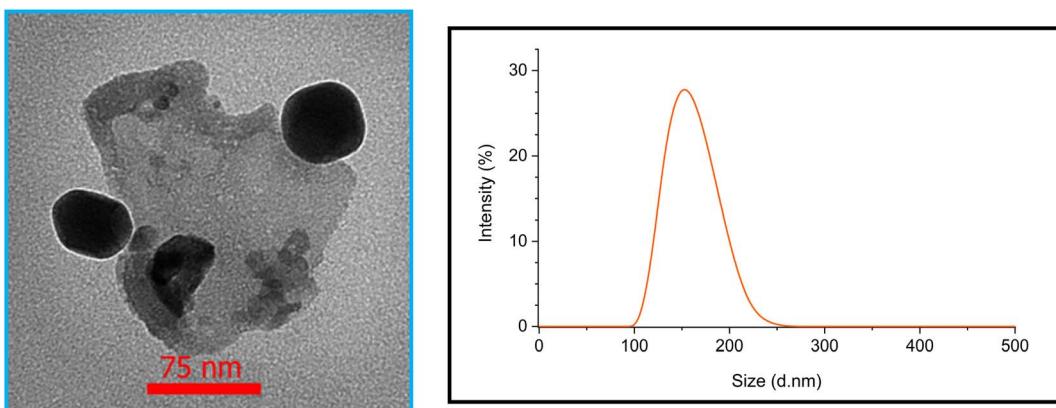


Fig. 4 TEM image (left) and DLS analysis (right) of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL].

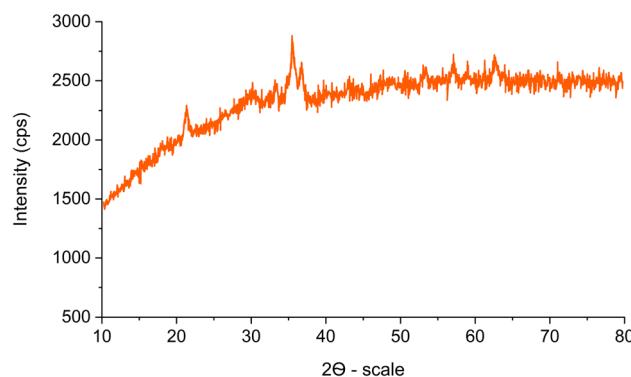


Fig. 5 The XRD pattern of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL].

The size and surface morphology of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] were explored by the FE-SEM images as shown in Fig. 3. Consequently, the fabricated nanoparticles have almost a sphere-shaped morphology with a slight agglomeration. The synthesized particles of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] were also in a nanoscale size. To have better insight into the morphology and size of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL], the TEM image was also assessed (Fig. 4, left). Apparently, the obtained nanoparticles have a spherical-like shape with a core-shell structure. This structure involves the light amorphous silica shell and the ligand that coated and wrapped the dark magnetic Fe_3O_4 .

Dynamic light scattering (DLS) technique was employed to study the hydrodynamic diameter of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] and the distribution of its particle size. The results are depicted in Fig. 4 (right). The size distribution with a mean diameter of 151 nm was observed for the particles. It is established that the hydrodynamic diameter by the DLS technique is larger than the obtained diameter by other techniques such as TEM and SEM. The rational reason for this observation could be the inability of DLS to differentiate between constituent and agglomerate particles.⁴⁶

The XRD pattern of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] is illustrated in Fig. 5. The characteristic diffraction peaks at $2\theta = 30.13^\circ$, 35.45° , 43.17° , 53.41° , 57.17° , and 62.65° were related to the (220), (311), (400), (422), (511) and (440) indices, respectively.

These results correspond to the standard XRD pattern for Fe_3O_4 (JCPDS card no. 19-0629) having a cubic spinel structure of the magnetite. Therefore, the crystal structure of the synthesized magnetic Fe_3O_4 nanoparticles remained unaffected during the synthesis of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL]. The broad peaks below 30° are ascribed to the amorphous silica wrapping the Fe_3O_4 nanoparticles. The presence of broad peaks in the XRD pattern and the scattering of Cu ions in the catalyst structure caused some difficulties in the detection of the copper species. In general, an amorphous shape could be suggested for the copper species.

The energy-dispersive X-ray spectroscopy (EDX) was employed to determine the chemical composition of magnetic $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] nanoparticles (Fig. 6). The resulting spectrum from EDX analysis endorses the presence of Fe, O, Si, C, N, S, and Cu in the structure of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL]. Consequently, the functionalization of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ with [CuL] was confirmed.

The thermal stability of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] and the weight loss of its material upon heating were investigated by TGA analysis (Fig. 7). As shown in Fig. 7, around 0.3% weight loss was detected below 200 °C which is ascribed to the removal of adsorbed water from the catalyst pores. Hence, $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] is thermally stable up to 200 °C. A gentle weight loss (~4.7%) was observed in a range of 200–800 °C which

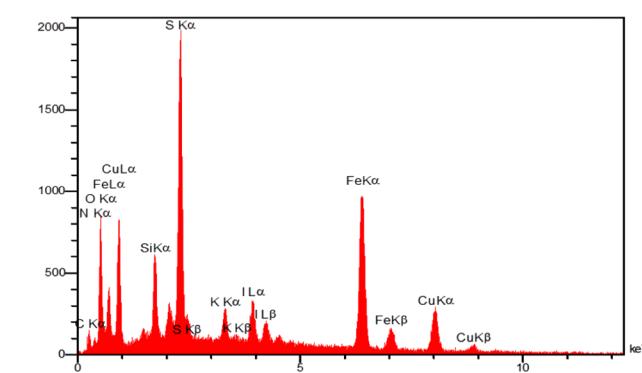
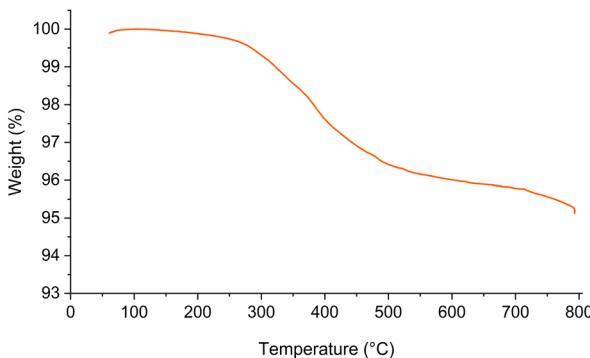
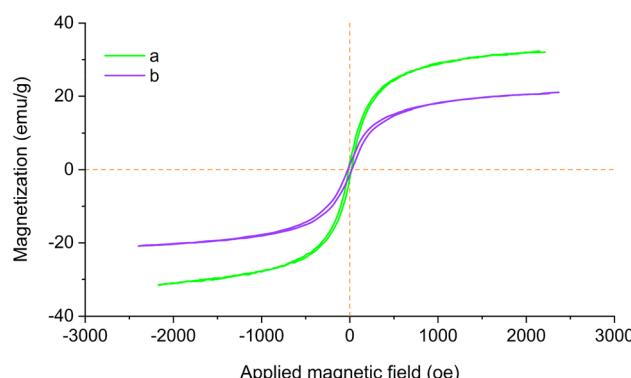


Fig. 6 EDX analysis of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL].



Fig. 7 TGA of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL].Fig. 8 Magnetization curves of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ (a) and $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] (b).

corresponds to the thermal degradation of organic residue tethered to the surface of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ nanoparticles and decomposition of the copper-thiosemicarbazone complex. Consequently, grafting the organic residue on the $\text{Fe}_3\text{O}_4@\text{SiO}_2$ surface was confirmed by TGA analysis.

Atomic absorption spectroscopy (AAS) was conducted to investigate the amount of Cu content supported on $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -Ligand (adduct V). According to the AAS analysis, the loading amount of copper in $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] was determined to be 1.04 mmol g⁻¹. Thus, the synthesized nanoparticles V have a high capacity to form chelates with the Cu species which could be attributed to the presence of thiosemicarbazone residues.

The magnetic properties of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ and $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] were assessed with magnetic hysteresis measurements. The magnetization curves of both studied nanoparticles are shown in Fig. 8. According to the obtained results, the saturation magnetization (M_s) values of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] and $\text{Fe}_3\text{O}_4@\text{SiO}_2$ proved to be 20.12 and 30.96 emu g⁻¹, respectively. The hysteresis loops characteristic displays the magnetic property of the studied substrates. As a result, the magnetic $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] nanoparticles may simply be collected from the reaction mixture using a magnet field. Clearly, the M_s value of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] is lower than that of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ since the organic moiety and the silica shell on the Fe_3O_4 surface are non-magnetic materials.

Table 1 Influence of diverse reaction parameters on Ullmann-type synthesis of 3m^a

Entry	Solvent	Base	Catalyst amount (mol%)	Catalyst amount		
				T (°C)	Time (h)	Yield ^b (%)
1	H_2O	Cs_2CO_3	1	Reflux	24	NR ^c
2	EtOH	Cs_2CO_3	1	Reflux	10	30
3	MeCN	Cs_2CO_3	1	Reflux	10	48
4	THF	Cs_2CO_3	1	Reflux	12	26
5	Toluene	Cs_2CO_3	1	Reflux	10	54
6	DMF	Cs_2CO_3	1	110	5	93
7	DMF	K_2CO_3	1	110	8	79
8	DMF	Et_3N	1	110	12	36
9	DMF	NaOH	1	110	10	68
10	DMF	NaOAc	1	110	12	40
11	DMF	MgO	1	110	18	Trace
12	DMF	DMAP	1	110	16	45
13	DMF	Cs_2CO_3	1	r.t.	24	Trace
14	DMF	Cs_2CO_3	1	80	8	71
15	DMF	Cs_2CO_3	1	100	6	83
16	DMF	Cs_2CO_3	1	120	5	93
17	DMF	Cs_2CO_3	—	110	24	NR
18	DMF	Cs_2CO_3	0.5	110	12	47
19	DMF	Cs_2CO_3	0.7	110	8	76
20	DMF	Cs_2CO_3	0.9	110	5	90
21	DMF	Cs_2CO_3	1.1	110	5	93

^a Imidazole (1.0 mmol), iodobenzene (1.2 mmol), base (2.0 mmol), $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] (X mol%), and solvent (4 mL). ^b Isolated yield. ^c No reaction.

2.3. Ullmann-type N-arylation of N-heterocycles with aryl halides using $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL]

After the preparation and characterization of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL], its efficiency was explored by its application in the Ullmann-type N-arylation of N-heterocycles with diverse aryl halides (3a–3o). Our initial attempt was conducted to find out the optimized reaction conditions. Therefore, the reaction of imidazole with iodobenzene in the presence of 0.01 g (1 mol%) of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] was chosen as the sample reaction to attain 1-phenyl-1H-imidazole (3m). In this regard, the influence of key factors including solvent, base, catalyst amount, and temperature was assessed on the progress of the sample reaction (Table 1).

Primarily, it was attempted to determine the best solvent for the progress of the reaction. In this connection, several protic and aprotic solvents were employed to perform the sample reaction (Table 1, entries 1–6). When the reaction was carried out in water as a green and protic solvent, 3m was not prepared even after prolonging the reaction time up to 24 h (entry 1). The use of EtOH as another protic solvent afforded 3m in 30% yield after 10 h (entry 2). Low to moderate yields of 3m were obtained when MeCN, THF, and toluene were applied as the reaction media (entries 3–5). Practically, the best result was achieved when DMF was used as the reaction solvent (entry 6). The next



Table 2 Structure of the synthesized *N*-aryl *N*-heterocycles 3a–3o^a

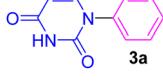
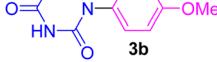
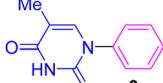
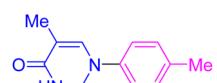
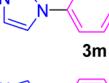
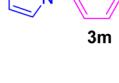
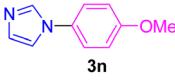
Entry	Aryl halide	N-Heterocycle	Product	Time (h)	Yield ^b (%)
1	C ₆ H ₅ I	Uracil		6	67
2	4-MeO-C ₆ H ₄ I	Uracil		7	65
3	C ₆ H ₅ I	Thymine		7	59
4	4-Me-C ₆ H ₄ Br	Thymine		9	54
5	C ₆ H ₅ I	Theophylline		5	77
6	4-MeO-C ₆ H ₄ I	Theophylline		5	74
7	C ₆ H ₅ I	Theobromine		5.5	78
8	4-MeO-C ₆ H ₄ I	Theobromine		5.5	76
9	4-MeO-C ₆ H ₄ I	Indole		6	88
10	2-MeO-C ₆ H ₄ I	Indole		9	73
11	C ₆ H ₅ I	Benzimidazole		4.5	92
12	1-Br-C ₁₀ H ₇	Imidazole		7	82
13	C ₆ H ₅ I	Imidazole		5	93
14	C ₆ H ₅ Br	Imidazole		7	88



Table 2 (Contd.)

Entry	Aryl halide	N-Heterocycle	Product	Time (h)	Yield ^b (%)
15	C ₆ H ₅ Cl	Imidazole		24	Trace
16	4-MeO-C ₆ H ₄ I	Imidazole		5	91
17	C ₆ H ₅ I	2-Phenylimidazole		6	79

^a Reaction conditions: aryl halide (1.2 mmol), N-heterocycle (1 mmol), Cs₂CO₃ (2 mmol), Fe₃O₄@SiO₂-[CuL] (0.01 g, 1 mol%), DMF (4 mL), 110 °C.

^b Isolated yield.

endeavor was carried out to find an efficient base to perform the reaction (Table 1, entries 6–12). As indicated in Table 1, Et₃N, NaOAc, MgO, and DMAP didn't influence adequately on the reaction time and yield (entries 8 and 10–12). A considerable improvement in yield of **3m** was acquired when K₂CO₃ and NaOH were employed as the reaction base (entries 7 and 9). The Cs₂CO₃ proved to be an adequate base for enhancing the yield and the completion of the reaction (entry 6), hence it was selected base for further optimization experiments.

Temperature is undeniably important factor in determining the optimized condition. To this end, the synthesis of **3m** was carried out at different temperatures (Table 1, entries 6 and 13–16). Practically, enhancing the reaction temperature from room temperature up to 100 °C remarkably increased the reaction efficacy (entries 13–15). The highest yield of **3m** in the shortest reaction time was obtained when the sample reaction was conducted at 110 °C (entry 6). However, more elevation in the reaction temperature did not lead to further improvement and progress of the reaction (entry 16). The last attempt was focused on the assessment of loading the catalyst amount on the reaction progress (Table 1, entries 6 and 17–21). Firstly, the sample

reaction was achieved in the absence of Fe₃O₄@SiO₂-[CuL]. However, as expected no reaction happened and **3m** was not acquired even in trace amount (entry 17). As depicted in Table 1, the reaction efficacy was improved by enhancing the amount of Fe₃O₄@SiO₂-[CuL]. The best result was observed when 1 mol% Fe₃O₄@SiO₂-[CuL] was employed for the preparation of **3m** (entry 6). Loading more catalyst amounts didn't improve the reaction progress. Overall, the optimized conditions for this current Ullmann-type *N*-arylation of N-heterocycles proved to be the use of 1 mol% Fe₃O₄@SiO₂-[CuL] in the presence of Cs₂CO₃ and DMF at 110 °C. The optimized stoichiometric ratios of N-heterocycle/aryl halide/Cs₂CO₃ for preparation of **3m** were determined to be 1:1.2:2. It is also worth mentioning that while a higher amount of aryl halides was employed, the formation of byproduct like biaryl compounds was observed.

Having determined the optimized condition, the generality and limitations of this procedure were explored by its application to Ullmann-type *N*-arylation of some N-heterocycle compounds comprising pyrimidine nucleobases (uracil, thymine), xanthines (theophylline, theobromine), azoles (imidazole, 2-phenylimidazole), and the other analogs

Table 3 Comparing the efficiency of Fe₃O₄@SiO₂-[CuL] with some reported catalysts for synthesis of **3m** from the reaction between imidazole and iodobenzene

Entry (ref.)	Catalyst	Conditions	Time (h)	Yield (%)
1 (ref. 47)	Cu ₂ O (10 mol%)	DMSO, KOH, 110 °C	24	90
2 (ref. 48)	Cu-Y zeolite (10.8 mol%)	DMF, K ₂ CO ₃ , 120 °C	24	99
3 (ref. 49)	CuFAP (12.5 mol%)	DMSO, K ₂ CO ₃ , 110 °C	6	92
4 (ref. 50)	Nano-CuO (2.5 mol%)	DMSO, KOH, 110 °C	24	91
5 (ref. 51)	SiO ₂ -Py-Cu(OAc) ₂ (5 mol%)	Toluene, Cs ₂ CO ₃ , 100 °C	8	92
6 (ref. 37)	Fe ₃ O ₄ @SiO ₂ -PVA-Cu(II) (0.6 mol%)	DMF, <i>t</i> -BuONa, 100 °C	5	94
7 (this work)	Fe ₃ O ₄ @SiO ₂ -[CuL] (1 mol%)	DMF, Cs ₂ CO ₃ , 110 °C	5	93



(benzimidazole, indole) with structurally diverse aryl halides. In this regard, the different aryl halides including iodobenzene, bromobenzene, chlorobenzene, 4-iodoanisole, 2-iodoanisole, 4-bromotoluene, and 1-bromonaphthalene were employed. The starting materials, yields, and structure of the synthesized compounds **3a–3o** are depicted in Table 2. The structures of **3a–3o** were confirmed by ^1H -, ^{13}C -NMR, and IR spectroscopy methods. The synthesized *N*-aryl *N*-heterocycles **3a–3o** were obtained in good to excellent yields depending on the nature of examined aryl halides and *N*-heterocycles. In the case of uracil and thymine, products **3a–3d** were obtained in 54–67% yields (Table 2, entries 1–4). For nucleobases, the lower yield of products is attributed to the weaker nucleophilic power of nucleobases compared to other screened *N*-heterocycles. It is worth mentioning that good regioselectivity was observed at the site of nucleobase *N*-arylation using the present protocol. As a consequence, the N^1 -aryl adducts **3a–3d** were majorly prepared while the N^1,N^3 -diaryl products were attained in trace amounts. Theophylline and theobromine were also examined under the optimized conditions to afford products **3e–3h** in 74–78% yields (Table 2, entries 5–8). Indole, benzimidazole, and imidazole derivatives were successfully performed in the Ullmann-type *N*-arylation reaction with aryl halides. It seems that the reaction is sensitive to the steric hindrance on both aryl halide and *N*-heterocycle. Practically, a lower yield of products was attained when 2-iodoanisole, 1-bromonaphthalene, and 2-phenylimidazole were used as the sterically hindered starting materials (Table 2, entries 10, 12, and 17). The influence of the steric factor is well recognized by comparing the reaction efficiency for 4-iodoanisole and 2-iodoanisole as the starting materials (Table 2, entries 9 and 10). In addition, the type of halogen atom in aryl halide influenced the reactivity in the Ullmann-type *N*-arylation reaction. In this connection, the iodobenzene, bromobenzene, and chlorobenzene were reacted with imidazole under the optimized conditions to acquire product **3m** (entries 13–15). The reactivity of bromobenzene is lower than that of iodobenzene. While bromobenzene affords a comparable yield of **3m**, however, the reaction took a longer time to be completed. Among pointed out halobenzenes, chlorobenzene has the least reactivity due to the higher bond dissociation energy. Consequently, **3m** was attained in trace amount when chlorobenzene was employed as the starting aryl halide.

To realize the efficacy and potency of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL], the present catalytic system was compared with several previously reported methods for Ullmann-type *N*-arylation of imidazole with iodobenzene to afford **3m** (Table 3). As can be seen in Table 3, the use of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] is associated with several benefits such as short reaction time, excellent yield, low amount of catalyst, simple recovery, and separability from the reaction media.

The recovery and recycling of a catalyst are essential issues from both economic and green chemistry aspects. In this context, the application of heterogeneous catalysts is highly preferred in comparison with homogeneous catalysts. In this regard, the consecutive synthesis of 1-phenyl-1*H*-imidazole (**3m**) via the Ullmann-type *N*-arylation of imidazole with iodobenzene

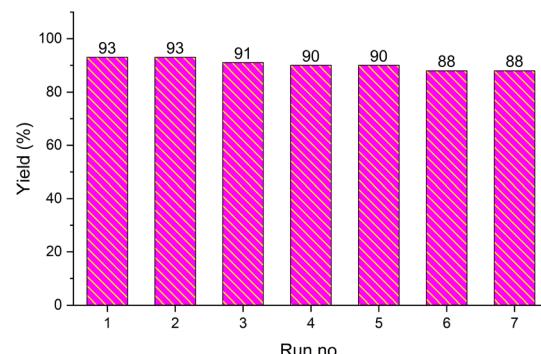


Fig. 9 Catalytic potency of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] in 7 repeated runs for preparation of **3m**.

using 0.01 g (1 mol%) $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] was considered. The heterogeneity and recoverability of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] was evaluated for seven sequential reactions in a model reaction. After the achievement of each reaction run, $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] was separated and retrieved from the reaction by a magnet. After washing the separated catalyst with deionized water and ethanol, it was dried in an oven at 70 °C. The dried $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] was directly used for achieving the subsequent reaction in the absence of any fresh catalyst. All successive reaction runs were quenched after 5 h (Fig. 9). As illustrated in Fig. 9, the performance and efficacy of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] are remarkable for at least up to seven consecutive runs with no significant deterioration in its reactivity. Among the synthesized compounds shown in Table 2, the lowest yield was related to 5-methyl-1-*p*-tolylpyrimidine-2,4(1*H*,3*H*)-dione (**3d**). The recycling test was also performed for the synthesis of **3d** via the same procedure that was used for **3m**. Accordingly, product **3d** was obtained in 54% to 47% yields after the first and seventh consecutive reaction runs, respectively. These experiments have verified the heterogeneous nature, stability, and recyclability of $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL].

The structure of recovered $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] was studied by the FT-IR spectrum after the seventh reaction run which is shown in Fig. 10. As can be understood from Fig. 10, there is no significant difference between the FT-IR spectrum of the recycled $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL] with that of the fresh catalyst. This

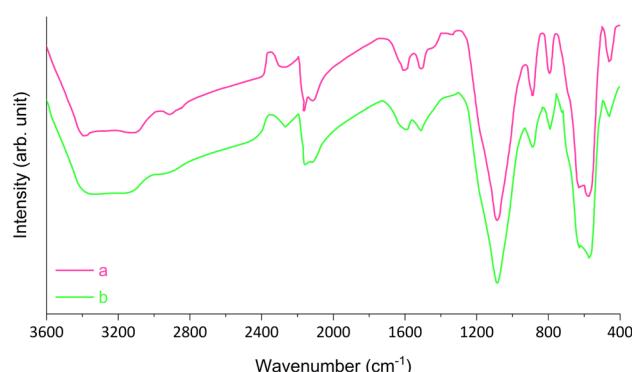
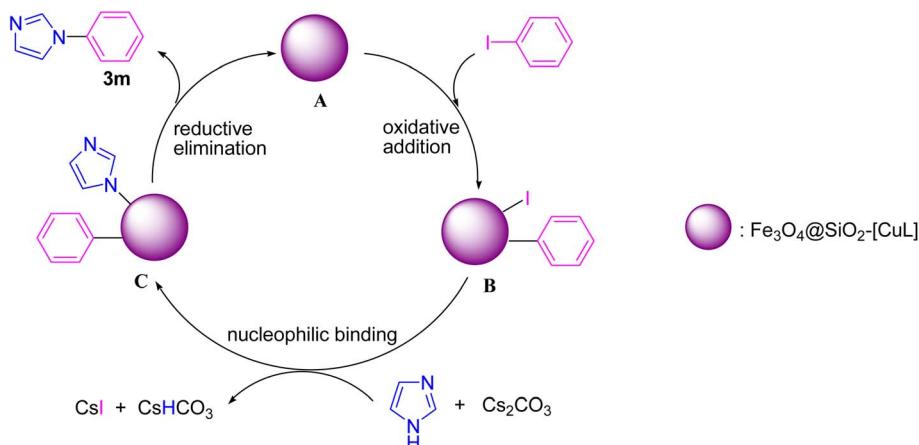


Fig. 10 FT-IR spectra of fresh (a) and recovered (b) $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -[CuL].





Scheme 3 A plausible mechanism for synthesis of **3m** using $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$.

evidence proves that the chemical structure of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ is remaining intact after seven runs. Besides, the hot-filtration test (split test) was performed on the sample reaction to explore the heterogeneous nature and stability of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$.⁵² In this context, the sample reaction was achieved for synthesis of **3m** and after 2.5 h the catalyst was retrieved from the reaction media using a magnet. The remaining residue was then allowed to continue the reaction for a further 2.5 h. As expected, further conversion and improvement of the sample reaction was paused which is good evidence for negligible leaching of copper species from the surface of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$. Overall, the heterogeneity, strong binding of Cu specios on the catalyst surface, and remarkable stability of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ were endorsed by these experiments.

Based on the previously reported mechanism for Ullmann-type *N*-arylation of *N*-heterocycles with aryl halides,^{53,54} we proposed a plausible mechanism for the synthesis of **3m** in the presence of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ (Scheme 3). Initially, the oxidative addition of iodobenzene to $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ occurs to produce a transient Cu(III) species (**B**). Then, the treatment of adduct **B** with the *in situ* base-activated imidazole generates adduct **C**. Eventually, the C–N coupling reaction and synthesis of **3m** are conducted *via* the reductive elimination from adduct **C**. Simultaneously, the active $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ is regenerated to proceed to the next catalytic cycle.

3. Conclusion

In this study, we have described the synthesis of the functionalized silica-coated magnetite nanoparticles with Cu(I)-thiosemicarbazone complex ($\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$) as a new magnetically retrievable nano catalyst. The $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ was efficiently employed for the synthesis of *N*-aryl *N*-heterocycles *via* Ullmann-type *N*-arylation of nucleobases, xanthines, and other *N*-heterocycles with aryl halides. The structure of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ was characterized by several techniques including FT-IR, FE-SEM, TEM, DLS, XRD, EDX, TGA, AAS, and VSM analysis. The significant benefits associated with this catalytic system are: (i) the stability and the facile

synthesis of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$, (ii) the cheapness and the heterogeneous nature of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$, (iii) the ease of catalyst separation and collection by an external magnet field, (iv) the reusability of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ for at least 7 sequential reaction runs, (v) the ease of handling, operation and simple work-up process, (vi) good to excellent yields of the *N*-aryl adducts, and (vii) no need to use external promoters such as ligands or additives. Altogether, $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ could be considered for the synthesis of industrially and medicinally important compounds as well as other copper-catalyzed Ullmann-type coupling reactions.

4. Experimental

4.1. General

The solvents and substrates were purchased from Merck and other chemical vendors. The chemicals were directly applied without further purification. The following instruments were employed for characterizing the structure and morphology of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ as well as products. XRD: Brüker D8 Advance X-ray diffractometer; EDX: SAMX; FE-SEM: MIRA3 TESCAN-XMU; TEM: 912 AB (120 kV) microscope; DLS: Horiba SZ-100; TGA: Shimadzu (TG-50), heating rate = 14 °C min⁻¹; VSM: Lake Shore VSM-7400; Atomic absorption spectrometer (AAS): GBC-932 AA; FT-IR: Thermo Nicolet Avatar 370; Melting points: Electrothermal IA 9000, uncorrected data; ¹H- and ¹³C-NMR: Brüker Avance (DPX-300 MHz and DPX-250 MHz), internal standard: tetramethylsilane, abbreviations: coupling constant (*J* in Hz), singlet (*s*), doublet (*d*), triplet (*t*), quartet (*q*), quintet (*quint*), multiplet (*m*), broad (*br*).

4.2. General procedure for synthesis of $\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$

$\text{Fe}_3\text{O}_4@\text{SiO}_2\text{-}[\text{CuL}]$ was synthesized *via* 6 steps as shown in Scheme 2. The nanoparticles **I–III** were synthesized through the procedures reported in the literature.^{43,44} Initially, the co-precipitation method was employed to prepare the Fe_3O_4 nanoparticles. Thus, a mixture of $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$ (20 mmol, 3.976 g) and $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ (40 mmol, 10.812 g) was added to the deionized water (140 mL) which was preheated to 80 °C in an oil



bath. A solution of NH_4OH (0.64 mol, 25.4 mL, 25 wt%) was added dropwise to vigorously stirred the reaction mixture under the N_2 (g) atmosphere. The Fe_3O_4 nanoparticles were prepared after 1 h which was magnetically separated and collected from the reaction container. The black solid was then washed with deionized water (4×15 mL) and EtOH (4×15 mL) and dried in a vacuum oven at 80°C to attain Fe_3O_4 (**I**) as a black powder. Afterward, the Fe_3O_4 nanoparticles (**I**, 2.0 g) were added to a solution of deionized water and EtOH ($1:3$, 80 mL). After the sonication of the reaction mixture for half an hour, NH_4OH (6.0 mL, 25 wt%) was added. The mixture was stirred at room temperature while tetraethyl orthosilicate (TEOS, 10 mL) was added portionwise. After 12 h, the obtained solid was separated by an external magnet field, washed with deionized water (4×15 mL) and EtOH (4×15 mL), and dried in a vacuum oven at 60°C to afford the silica-coated Fe_3O_4 nanoparticles ($\text{Fe}_3\text{O}_4@\text{SiO}_2$, **II**). In the third step, (3-chloropropyl)trimethoxysilane (CPTMS) was used to functionalize the silica surface of **II**. Therefore, CPTMS (3.64 mL, 20 mmol) and **II** (2.0 g) were reacted together in EtOH (70 mL) at 70°C . The magnetic solid **III** was salvaged from the reaction media after 24 h using a magnet field, washed with EtOH (4×15 mL), and kept in a vacuum oven at 60°C to be dried. In step 4, a mixture of Et_3N (4.04 g, 40 mmol), 4-hydroxyacetophenone (4.08 g, 30 mmol), and freshly prepared **III** (2.0 g) was added in a double-necked round bottom flask containing dry toluene (70 mL). The reaction was conducted in refluxing toluene using an oil bath. After 48 h, the solid was magnetically collected and separated from the reaction flask and washed with dry toluene (4×15 mL). Then, it was placed and dried in a vacuum oven at 50°C to give the functionalized $\text{Fe}_3\text{O}_4@\text{SiO}_2$ -hydroxyacetophenone intermediate (**IV**). To prepare the thiosemicarbazone adduct **V** in step 5, thiosemicarbazide (2.42 g, 26 mmol), adduct **IV** (2.0 g), and AcOH (few drops) was added in a double-necked round bottom flask containing EtOH (40 mL) and the flask was heated to reflux in an oil bath for 24 h. The thiosemicarbazone adduct **V** was obtained as a creamy solid after magnetic separation of the solid, washing with deionized water (4×15 mL) and EtOH (4×15 mL), and drying in a vacuum oven at 50°C . Finally, a mixture of KI (1.5 g, 9 mmol), $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ (2.5 g, 10 mmol), and thiosemicarbazone adduct **V** (2.0 g) was dispersed in a flask containing deionized water (30 mL). After stirring the mixture at room temperature for 48 h, the dark-brown solid was separated from the flask by a magnet, washed with deionized water (4×15 mL), and kept in a vacuum oven at 70°C to be dried. The obtained solid which is $\text{Fe}_3\text{O}_4@\text{SiO}_2[\text{CuL}]$ (**VI**) was corked and stored in a desiccator at room temperature.

4.3. General procedure for the Ullmann-type N-arylation of N-heterocycles

In a round bottom flask containing DMF (4 mL), it was charged with the desired N-heterocycle (1 mmol), aryl halide (1.2 mmol), Cs_2CO_3 (2 mmol), and $\text{Fe}_3\text{O}_4@\text{SiO}_2[\text{CuL}]$ (0.01 g, 1 mol%). The mixture was heated to 110°C and the reaction progress was checked by TLC. When the reaction was completed (the indicated times in Table 2), the crude mixture was cooled to room

temperature. Then, $\text{Fe}_3\text{O}_4@\text{SiO}_2[\text{CuL}]$ was separated by a magnet from the reaction mixture and DMF was evaporated under reduced pressure. The remaining residue was dissolved in EtOAC (10 mL), washed twice with water (20 mL), and dried using sodium sulfate. Afterward, the extracted EtOAC was evaporated using a rotary evaporator under a vacuum to afford the crude product. The desired N-aryl N-heterocycle was obtained after purification *via* column chromatography on silica gel eluted with the proper solvent.

Note: The characterization data of **3a–3o** can be found in the ESI.†

4.4. Recycling of $\text{Fe}_3\text{O}_4@\text{SiO}_2[\text{CuL}]$

The separated solid by a magnet was washed with deionized water (2×30 mL) and ethanol (2×30 mL). Afterward, the catalyst was kept in a vacuum oven to be dried at 70°C and reused for subsequent reactions.

Conflicts of interest

The authors declare that they have no conflict of interest.

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