

these transformations suffer from certain limitations, such as the need for excessive amounts of waste generating agents, harsh reaction conditions, poor functional group tolerance, and tedious synthetic procedures to access substrates. Hence, there is a high demand for efficient and environmentally benign new methods for the synthesis of cyanamide from readily available starting materials under mild conditions. As part of our continuing studies on green chemistry,¹¹ we envisage that the sulfur in thiourea could transfer to useful compounds in elegant reactions. We herein report an efficient method for the synthesis of cyanamides and disulfanes from aryl thiourea and halides in one pot procedure.

Initially, 1-phenylthiourea (**1a**) and (2-chloroethyl)benzene (**2a**) were selected as the model substrates to screen the optimal reaction conditions, and the results are summarized in Table 1. To our delight, under the combination of Na₂CO₃ and Cu(OH)₂ in DMSO in the open air at 120 °C for 1 h, the desired product *N*-phenethyl-*N*-phenylcyanamide (**3aa**) and disulfane (**4a**) were obtained in 95% and 79% yield, respectively (Table 1, entry 1). The yields of compounds **3** and **4** were calculated based on arylthiourea (**1**). In addition, thioether (**5a**) was also detected in trace amount by GC-MS. The reaction fails without base (entry 2), and the starting materials **1a** and **2a** were all recovered in yield >95%, showing that base is critical to the reaction. In this reaction, inorganic bases may be acid scavenger, promoting the nucleophilic addition reaction between arylthiourea and halide. The yield of cyanamide (**3aa**) was increased as the amount of base used in the reaction, however, the yield of disulfane was low without Cu(OH)₂ (entries 3–6). Various kinds of bases were then

screened, such as K₂CO₃, Cs₂CO₃, KHCO₃, NaHCO₃, and Et₃N; among them, K salts gave the similar results to Na salts, but no desired products were determined under organic bases (entries 7–11). These results indicated that Na₂CO₃ is the best choice to promote the reaction. With the increase of Cu(OH)₂, the yield of disulfane decreased, while the yield of cyanamide remained excellent (entries 12; see ESI† for details). These results may be due to the chelation between copper and mercaptan. Subsequently, different solvents, including water, MeCN, DMF, DMA, acetone, and toluene were estimated, and DMSO was superior to the others (entries 15–18). Furthermore, the reaction temperature and time were also optimized. It was found that increasing the temperature higher than 80 °C can obviously shorten the reaction time, and the reaction finished under 120 °C within 1 h (for more details, see ESI†).

With the optimal reaction conditions in hand, we then tried to explore the generality and limitations of the protocol. First, various halides were investigated and the results are summarized in Table 2. Long-chain chloroalkanes are good substrates for the preparation of corresponding cyanamides (**3ad–3ae**) in 84% to 95% yields and disulfanes (**4d–4e**) in 73% to 95% yields. Phenol ether is tolerant in the reaction, and the corresponding cyanamide (**3af**) and disulfane (**4f**) were obtained in high yields. Simple primary chloroalkane and steric-hindrance halides can also offer corresponding products (**3ag, 3ah**) in 70% and 88% yields, respectively. Benzyl chloride and its alkyl, alkoxy, halo, and cyan substituted derivatives will react smoothly and produce cyanamides (**3ai–3aq**) in moderate to good yields. To our surprise, except for 3-methylbenzyl chloride affording

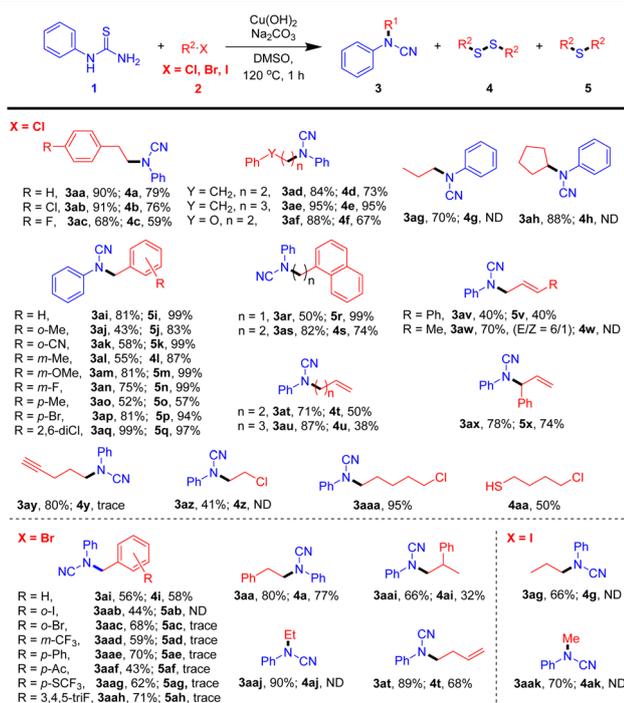
Table 1 Optimization of reaction conditions^a



Entry	Base	Solvent	Yield of 3aa ^b /%	Yield of 4a ^b /%	Yield of 5a ^b /%
1	Na ₂ CO ₃	DMSO	95 (93)	79 (70)	Trace
2	—	DMSO	0	0	0
3 ^{c,d}	Na ₂ CO ₃	DMSO	33	16	20
4 ^{c,e}	Na ₂ CO ₃	DMSO	65	30	32
5 ^{c,f}	Na ₂ CO ₃	DMSO	74	26	30
6 ^{c,g}	Na ₂ CO ₃	DMSO	81	26	40
7	K ₂ CO ₃	DMSO	97	70	Trace
8	Cs ₂ CO ₃	DMSO	70	50	Trace
9	NaHCO ₃	DMSO	90	75	Trace
10	KHCO ₃	DMSO	92	70	Trace
11	Et ₃ N	DMSO	0	0	0
12 ^h	Na ₂ CO ₃	DMSO	97	0	0
13 ⁱ	Na ₂ CO ₃	DMSO	81	57	Trace
14 ^j	Na ₂ CO ₃	DMSO	92	70	Trace
15 ^k	Na ₂ CO ₃	H ₂ O	0	0	0
16 ^l	Na ₂ CO ₃	CH ₃ CN	76	30	Trace
17	Na ₂ CO ₃	DMAc	79	62	Trace
18	Na ₂ CO ₃	DMF	91	71	Trace

^a Unless otherwise note, all reactions were performed with **1a** (0.20 mmol), **2a** (0.48 mmol, 2.4 equiv.), base (2.5 equiv.), Cu(OH)₂ (10 mol%), solvent (0.5 mL) at 120 °C for 1 h. ^b Yields were based on ¹H NMR analysis of the crude product using CH₂Br₂ as an internal standard. The value in parentheses is the isolated yield of product. ^c Without Cu(OH)₂. ^d Base (0.5 equiv.). ^e Base (1.0 equiv.). ^f Base (1.5 equiv.). ^g Base (2.0 equiv.). ^h Cu(OH)₂ (1.0 equiv.). ⁱ **2a** (2.0 equiv.). ^j **2a** (2.2 equiv.). ^k At 100 °C for 12 h. ^l At 80 °C for 12 h. DMAc = *N,N*-dimethylacetamide.

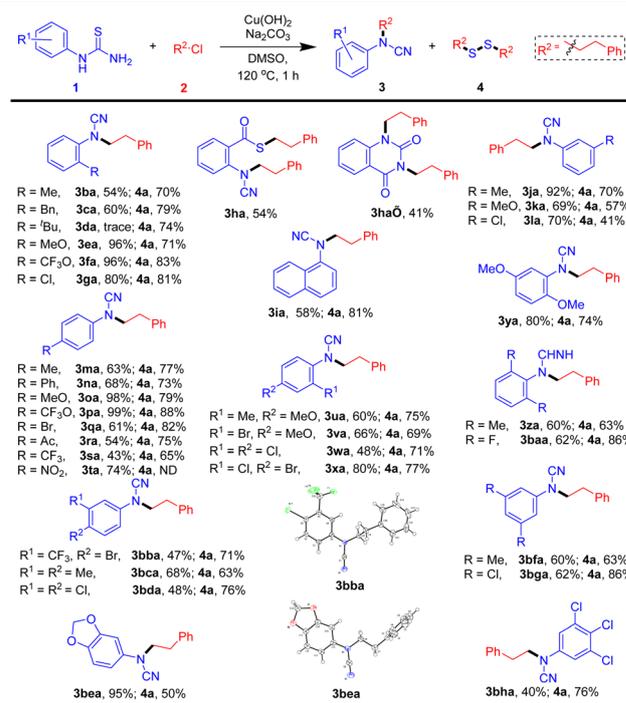


Table 2 Substrate scope of various halides 2^a

^a All reactions were performed with **1a** (0.2 mmol), **2** (2.4 equiv.), Cu(OH)₂ (10 mol%), Na₂CO₃ (2.5 equiv.), DMSO (0.5 mL) at 120 °C for 1 h. Yields referred to isolated yield.

disulfane **4l**, the other benzyl chloride derivatives produce thioethers (**5i–5q**) in excellent yields under standard conditions. Halides with a terminal or internal C=C double bond are suitable partners and give the desired products (**3av–3ax**, **4v–4x**) in moderate yields. Particularly, 5-chloropent-1-yne is also perfectly compatible under this reaction condition, and the corresponding cyanamide (**3ay**) was obtained in an 80% yield. To our delight, dichloroalkanes are acceptable for the preparation of halo-substituted cyanamides (**3az–3aaa**) and halo-substituted disulfane products (**4aa**). Benzyl bromide and its electron-withdrawing or donating derivatives can offer moderate yields of cyanamide products (**3ai**, and **3aab–3aah**). To our confusion, sulfane products were not easily obtained in these cases, and oily dibenzyl disulfane (**4i**) was obtained in 58% isolated yield. Moreover, bromoethane can produce cyanamide in excellent yield (90%) under standard reaction conditions. Furthermore, 1-iodopropane and methyl iodide successfully provide corresponding products (**3ag** and **3aak**) in yields of 66% and 70%, respectively. These results indicated that chloride, bromide, and iodo hydrocarbons are all available materials for the corresponding cyanamide products.

To explore further the generality of this protocol, various structurally diverse arylthioureas were then evaluated, and the representative results are summarized in Table 3. *Ortho*-substituted phenyl thioureas are tested first. Satisfactorily, both electron-rich (Me, Bn, MeO, and CF₃O) and electron-withdrawing (Cl) substituents arylthioureas can be

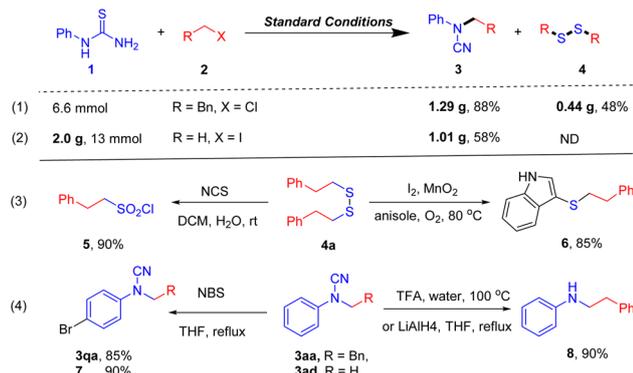
Table 3 Substrate scope of various arylthioureas 1^a

^a All reactions were performed with **1** (0.2 mmol), **2a** (2.4 equiv.), Cu(OH)₂ (10 mol%), Na₂CO₃ (2.5 equiv.), DMSO (0.5 mL) at 120 °C for 1 h. Yields referred to isolated yield.

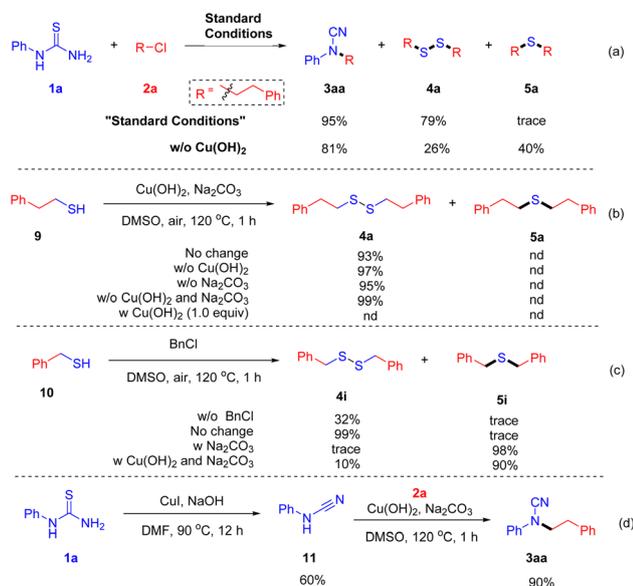
successfully coupled with **2a**, and the corresponding cyanamide products **3ba–3ga** and disulfane **4a** were furnished in yields ranging from 54% to 96%. However, the high-steric-hindrance substrate 1-(2-(*tert*-butyl)phenyl)thiourea cannot obtain a cyanamide product. In addition, ethyl 2-thioeidobenzoate reacting with **2a** will produce **3ha** and **3ha'** in moderate yields. In this case, **3ha** was generated *via* thiolysis between ethyl ester and *in situ* disulfane **4a**. Similarly, naphthyl thiourea substrate gives the corresponding cyanamide **3ia** in 58% yield and disulfane **4a** in 81% yield. *Meta*-substituted phenyl thioureas were well tolerated under the current reaction conditions, and the cyanamide products **3ja–3la** were obtained in good to excellent yields. Either electron-donating or electron-withdrawing substituents on the *para* site, aryl thioureas are successful in achieving the desired products. To our delight, multi-substituted alkyl, alkoxy, and/or halogen on the phenyl ring were well accommodated, and the corresponding products were generated in yields ranging from 48 to 95%. The configurations of **3bba** (ref. 12) and **3bea** (ref. 13) were unequivocally confirmed by X-ray crystallographic analysis.

To investigate the practicality of this method, gram-scale reactions and further transformations were carried out (Scheme 2). On a 6.6 mmol scale, the desired cyanamide **3aa** and disulfane **4a** were isolated in 88% and 48% yield, respectively. It is noteworthy that this approach can perform at a 2 g scale with iodomethane for the synthesis of the desired methyl cyanamide product (**3aak**) in 58% yield. Subsequently, disulfane

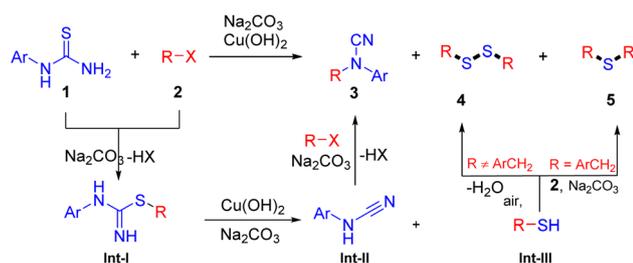




Scheme 2 Scale-up reactions and application transformations of products.



Scheme 3 Control experiments (a–d).



Scheme 4 Plausible mechanism.

4a can be readily converted to the corresponding sulfonyl chloride **5** and sulfenylindole **6** (ref. 14) in high yield. These results show that disulfane can replace mercaptan as an odorless reagent in organic synthetics. Phenyl cyanamides were easily brominated in the para site of the phenyl ring when treated with NBS in THF under reflux. It is a pity that the cyan group was eliminated when treating with TFA or LiAlH₄.

In order to further understand the reaction mechanism, some control experiments (Scheme 3) were conducted. First, the model reaction is performed under standard conditions (eqn (a)). Cyanamide (**3a**) and disulfane (**4a**) were found in the yields of 95% and 79%, respectively, and only a trace amount of thioether (**5a**) was detected by GC-MS. In the absence of Cu(OH)₂, the yield of thioether (**5a**) increased dramatically. Meanwhile, trace amounts of thiol and *N*-phenyl cyanamide were detected in GC-MS. These results showed that Cu(OH)₂ may promote the formation of disulfane (**4a**). Subsequently, 2-phenylethanethiol (**9**) was tested under various conditions (eqn (b)), and only disulfane (**4a**) was obtained. In the presence of 1.0 equiv. Cu(OH)₂, thiol (**9**) is consumed, and no sulfane products are detected by GC-MS. Indicating that excess copper salt may chelate with mercaptan. Moreover, 1,2-dibenzyl disulfane (**4i**) is produced in high yield when benzyl mercaptan (**10**) is performed under air in the presence or absence of benzyl chloride (eqn (c)). However, thioether (**5i**) becomes the main product in the presence of a base. Indicating that base facilitates the S_N2 reaction between thiol and halide. Furthermore, substituted cyanamide (**3aa**) was obtained successfully when treating *N*-phenyl cyanamide (**11**) with **2a** under standard conditions. These results indicated that mercaptan and *N*-phenyl cyanamide might be the intermediates in the reaction (Scheme 4).

Based on the above experimental results and previous reports,¹⁵ a tentative mechanistic interpretation of the preceding observations is proposed in Scheme 4. First, base-promoted nucleoaddition of halide to arylthiourea forms isothiurea **Int-I**, which undergoes further intramolecular elimination in the presence of base and copper to generate cyanamide **Int-II** and thiol intermediate **Int-III**. Then halide **2** and **Int-II** undergo S_N2 progress to afford the substituted cyanamide product (**3**). Meanwhile, thiol **Int-III** is oxidized by air to form a sulfane compound.

Conclusions

In summary, a simple method for the preparation of cyanamides and disulfanes has been established under mild reaction conditions in one pot. Many diverse substituted cyanamides and functional sulfanes are successfully prepared in moderate to excellent yields with this strategy. Importantly, the present methodology demonstrates excellent functional group compatibility, high atom- and step-economy, and mild reaction conditions. Further studies on the synthetic application of this strategy are underway in our lab.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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