


 Cite this: *RSC Adv.*, 2023, **13**, 32589

Mechanism of synergistic removal of NO and SO₂ by sodium bicarbonate

 Song Shu,^{ab} Yiqi Huang,^c Longhua Zou,^d Xinyi Zhang^c and Jianjun Li^{id}*^{ab}

Sodium bicarbonate (NaHCO₃) is considered to be an effective alkaline adsorbent for SO₂ removal and surprisingly, the concentration of NO is significantly reduced along with the generation of NO₂ during its desulfurization. Unfortunately, the mechanism of NO interaction with NaHCO₃, SO₂ and O₂ is ambiguous. In this work, the effects of absorption gas and absorber composition on SO₂/NO absorption performance were explored, the absorption products were characterized using XPS and SEM, and the Gibbs free energy of the inferred reaction path was calculated based on density functional theory (DFT). The results showed that SO₂ and O₂ synergistically promoted the absorption and removal of NO by NaHCO₃, which could completely remove SO₂ and absorb 90% of NO at 160 °C. Sodium metabisulfite (Na₂S₂O₅) and sodium dithionate (Na₂S₂O₆) were identified as the active substances responsible for efficient NO absorption, and the oxidation of Na₂S₂O₅ to Na₂S₂O₆ is the controlling step of the NO removal reaction. Specifically, Na₂S₂O₅ is an intermediate produced by the reaction of NaHCO₃ with SO₂, and subsequently reacts with O₂ to produce Na₂S₂O₆, which releases reactive oxygen species to oxidize NO to NO₂. In addition, when the S/N ratio is greater than 1 and the O₂ content is greater than 5%, both SO₂ and NO can maintain removal efficiency higher than 90%, indicating that the absorption reaction of SO₂ and NO by NaHCO₃ is highly adaptable to the flue gas composition.

 Received 12th July 2023
 Accepted 4th October 2023

DOI: 10.1039/d3ra04672a

rsc.li/rsc-advances

1. Introduction

As the world's population and energy demand increase, continued industrialization will undoubtedly raise the levels of atmospheric pollutants, such as SO₂ and NO, which are harmful to the environment and human health. Flue gas desulfurization (FGD) and ammonia-based selective catalytic reduction (NH₃-SCR) denitrification are simultaneously applied for controlling pollutant emissions in power plants.^{1–3} However, desulfurization performed using wet sorbents requires large installation space, large amounts of water and high capital as well as operating costs,^{4–6} and research has shifted toward dry methods of SO₂ removal. Various types of solid sorbents/catalysts are being used for dry FGD, like calcium based,² sodium based,^{7,8} activated carbon,^{9,10} and metal oxide.¹¹ Among the dry sorption, sodium bicarbonate (NaHCO₃) has attracted particular interest because of its ability to produce valuable solid compounds with SO₂, such as sodium sulfite and sulfate, which can be used as alkali agent, anti-caking agent and neutralizer. In addition, NaHCO₃ as a SO₂ dry-sorbent can couple cost-effectiveness and

environmental compatibility in practical engineering applications, and it has been implemented in a great number of plants in Europe.^{12–15}

Surprisingly, the concentration of NO_x was simultaneously reduced during the desulfurization with NaHCO₃,¹⁶ providing the possibility to realize integrated removal of multiple pollutants in a single system. Barbara Walawska *et al.* stated that the efficiency of NaHCO₃ for NO_x removal was about 21% at 300 °C,¹⁷ but failed in the absence of SO₂. It was presumed that the intermediate of Na/SO_x generated by the desulfurization reaction plays a key role in NO removal, especially sodium metabisulfite (Na₂S₂O₅). Unfortunately, the efficient Na/SO_x intermediates are still ambiguous and the synergistic removal mechanisms of SO₂ and NO_x are currently lacking in research.

In this work, the effect of gas composition on the absorption of SO₂ and NO by NaHCO₃ was investigated using a fixed-bed reactor. The Na/SO_x intermediates generated by the FGD process were characterized by scanning electron microscopy (SEM) and X-ray photoelectron spectroscopy (XPS), and then the adsorption properties of a series of possible Na/SO_x on SO₂ and NO were examined. Furthermore, the effects of O₂ in NO oxidation and Na/SO_x formation was evaluated. Most interestingly, the reaction paths were inferred from density functional theory (DFT) calculations of Gibbs freedom to elucidate the mechanism of NO and SO₂ removal by Na/SO_x. Overall, we pave the way for the development of NaHCO₃ for simultaneous desulfurization and denitrification.

^aCollege of Carbon Neutrality Future Technology, Sichuan University, Chengdu 610065, China. E-mail: jjli@scu.edu.cn
^bNational Engineering Research Center for Flue Gas Desulfurization, Chengdu 610065, China

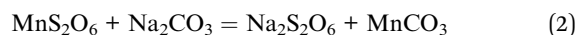
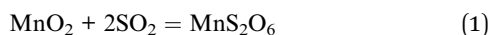
^cCollege of Architecture and Environment, Sichuan University, Chengdu 610065, China

^dCollege of Food and Biological Engineering, Chengdu University, Chengdu 610106, China


2. Materials and methods

2.1 Materials

The analytical grade of NaHCO₃ with mean diameter of 50–100 μm, sodium carbonate (Na₂CO₃), sodium bisulfite (NaHSO₃), sodium sulfite (Na₂SO₃) and Na₂S₂O₅ were all purchased by Kelong Company. Na₂S₂O₆, which is in science labs, was synthesized by the following chemical method:



2.2 Experimental sections

The experiments were carried out in the fixed-bed reactor with a diameter of 2 cm and a height of 1.6 cm at ambient, depicted schematically in Fig. 1. The simulated flue gas consisted of SO₂, NO, O₂ and N₂, supplied by a compressed cylinder and metered by a mass flow controller. The concentration of the gas compositions in the inlet and outlet flows was monitored by a flue gas analyzer (Testo 350, Germany). The temperature inside the bed was controlled by thermocouples.

To investigate the synergy effects on the adsorption of SO₂ and NO by NaHCO₃, individual and simultaneous sorption process were designed. The uniform reaction conditions set for SO₂ and NO adsorption by NaHCO₃ included $T = 160$ °C, GHSV = 10 000 h⁻¹, flow rate = 1 L min⁻¹, 5 vol% O₂ and adsorption residence time of 60 min. The composition of the adsorption gas was adjusted as a variable and its specific names are listed in Table 1.

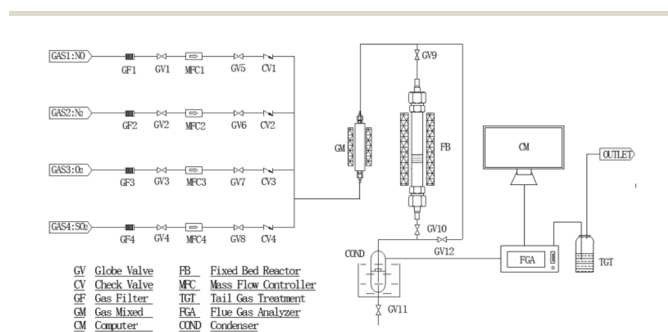


Fig. 1 Schematic diagram of the fixed-bed laboratory apparatus.

Table 1 The named samples corresponding to the specific gas compositions

Sample	Gas compositions	
	0–30 min	31–60 min
S550N0	550 ppm SO ₂	
N500S0	500 ppm NO	
S550/N500	550 ppm SO ₂	500 ppm NO
S550-N500	550 ppm SO ₂	550 ppm SO ₂ + 500 ppm NO
N500-S550	500 ppm NO	500 ppm NO + 550 ppm SO ₂
S550N500	550 ppm SO ₂ + 500 ppm NO	

The desulfurization and denitration properties of the possibility Na/SO_x included Na₂CO₃, NaHSO₃, Na₂SO₃, Na₂S₂O₅, and Na₂S₂O₆ were then examined, respectively. Since NaHSO₃ and Na₂S₂O₅ would release SO₂ at 160 °C according to the results of pre-experiments, NaHSO₃ and Na₂S₂O₅ as additives were set in the ratios of 1 : 4 with NaHCO₃ to measure their effects on NO removal. In addition, the role of O₂ in the oxidation of NO was investigated by coexisting/not coexisting O₂ in the sorption gas.

2.3 Characterization methods

The thermal decomposition properties were characterized using an SDTQ600 instrument (TA Instruments, USA) at a heating rate of 5 °C min⁻¹ to 200 °C and a N₂ flow rate of 100 mL min⁻¹. The specific surface areas were calculated by the Brunauer–Emmett–Teller (BET) method, and the pore sizes and pore capacities were calculated by the Barrett–Joyner–Halenda (BJH) method. SEM images of sorbent particle were performed using a JSM-7500F scanning electron microscope (JEOL Japan) at an acceleration voltage of 5 kV. X-ray photoelectron spectra (XPS) with Al Kα source carried out on a Thermo ESCALAB250Xi instrument was used to characterize and semi-quantify the chemical compositions of reacted Na/SO_x.

2.4 Computational section

To identify the differences in Gibbs free energy (ΔG) of the reactions involved, Density Functional Theory (DFT) calculations were performed with VASP 5.3.5 code.^{18,19} Generalized gradient approximation with Perdew–Burke–Ernzerhof exchange and correlation functional were used.²⁰ A plane-wave basis set of 400 eV cut off energy was employed in the framework of projector-augmented wave method.²¹ The Gaussian smearing with a width of 0.2 eV was used. The convergence criteria for the energy and force were set to 10⁻⁵ eV and 0.01 eV Å⁻¹. ΔG is calculated as the following:

$$\Delta G = G_{\text{product}} - G_{\text{reactant}} \quad (3)$$

where the G_{product} and G_{reactant} represent Gibbs free energies of products and reactants, respectively. The Gibbs free energies of gas phase can be calculated as the eqn (5):²²

$$G_{\text{g}}(T, P) = E_{\text{DFT}}(K^0, P^0) + \Delta H_{\text{g}}(K^0 \rightarrow T, P^0) - TS_{\text{g}}(T, P^0) \quad (4)$$

where the first term is the energy calculated by DFT, and the second and the third terms are the contribution of gas enthalpy and entropy under standard state pressure ($P^0 = 0.1$ MPa), respectively.

3. Results and discussion

3.1 Thermal decomposition properties of NaHCO₃

The Fig. 2A shows that NaHCO₃ decomposes at 80–160 °C, with the fastest rate of decomposition at 145 °C, while the weight loss at 80 °C corresponds to the separation of bound water.²³ The thermal decomposition process of NaHCO₃ is an active development, which is shown in the SEM images (Fig. 2B). As seen in



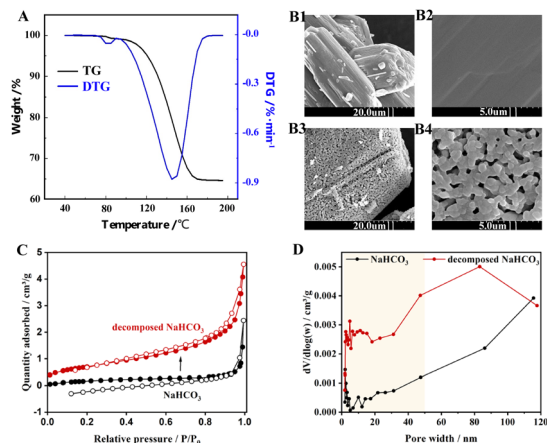


Fig. 2 (A) TG/DTG curves of NaHCO_3 at a heating rate of $5\text{ }^\circ\text{C min}^{-1}$. SEM images of (B1 and B2) NaHCO_3 , and (B3 and B4) NaHCO_3 thermally decomposed at $160\text{ }^\circ\text{C}$. (C) Nitrogen adsorption–desorption isotherm and (D) pore size distribution of pristine and $160\text{ }^\circ\text{C}$ decomposed NaHCO_3 .

Fig. 2B1 and B2, the physical structure of the initial NaHCO_3 particles is nonporous, while the decomposed NaHCO_3 in Fig. 2B3 and B4 produces some visible micro-grade pores of about $1\text{ }\mu\text{m}$. BET measurements were further carried out to characterize the pore structure of NaHCO_3 . The N_2 adsorption–desorption isotherm results of are shown in Fig. 2C, where the specific surface area of NaHCO_3 increases from 0.7 to $2.4\text{ m}^2\text{ g}^{-1}$ after thermal decomposition at $160\text{ }^\circ\text{C}$, which is consistent with the shrinkage nucleation model.²⁴ The results of the pore size distribution in Fig. 2D indicate that the thermal decomposition of NaHCO_3 favours the formation of multistage pores, and these well-developed pore structures are expected to provide channels and sites for the physical and chemisorption of SO_2/NO .

3.2 Performance of NaHCO_3 in absorbing SO_2/NO

To investigate the effect of SO_2 on the absorption of NO by NaHCO_3 , a series of single-component absorption experiments were first carried out and the results correspond to Fig. 3. In Fig. 3A1, the concentration of SO_2 drops sharply to 0 at the beginning of the reaction for 10 min, and remains complete absorption until the end of the reaction, indicating the dominance of NaHCO_3 in desulfurization. However, the concentration of NO in Fig. 3B1 decreases from 500 to 400 ppm at most and then rises to 500 ppm, implying that the individual uptake of NO by NaHCO_3 is weak and occurs mainly through physical adsorption and the spontaneous reaction of NO oxidation to NO_2 . On this basis, a performance test (Fig. 3C1) was conducted with pre-sorption of SO_2 followed by NO adsorption, resulting in no decrease in NO concentration after stopping SO_2 injection. The above results show that NaHCO_3 has good SO_2 absorption performance, but it is ineffective for single NO absorption, and pre-absorption of SO_2 has no effect on NO absorption by NaHCO_3 .

The sorption performance of the S550-N500, N500-S550 and S550N500 samples with SO_2 and NO coexisting was shown in

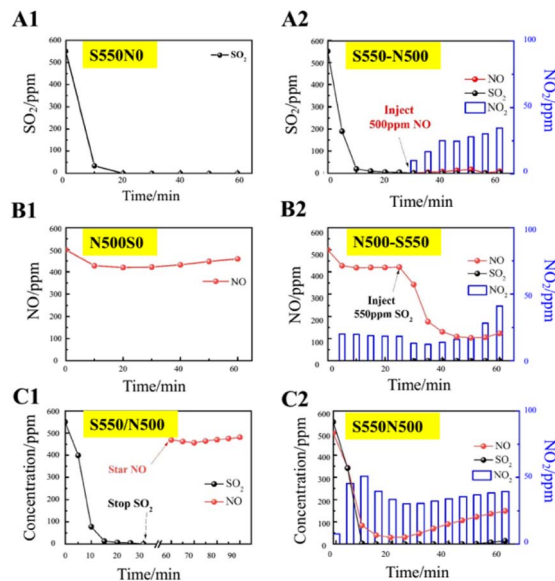


Fig. 3 SO_2/NO adsorption properties. $T = 160\text{ }^\circ\text{C}$, $\text{GHSV} = 10\,000\text{ h}^{-1}$, flow rate = 1 L min^{-1} and the total sorption time is 60 min. Single-component adsorption experiments: (A1) S550N0, (B1) S0N500 and (C1) S550/N500. Simultaneous adsorption of SO_2 and NO : (A2) S550-N500, (B2) N500-S550 and (C2) S550N500.

Fig. 3A2, B2 and C2. It is easily found that the coexistence of SO_2 significantly promotes the absorption and conversion of NO and increase the production of NO_2 , inferring that the Na/SO_x intermediates produced by the reaction of SO_2 with NaHCO_3 as active substances stimulate the absorption of NO . When NO is exposed to 550 ppm of SO_2 (Fig. 3A2), the NO concentrations immediately decrease to 0 ppm, while a maximum of about 35 ppm of NO_2 is detected in 30 min. The data in Fig. 3B2 show that when SO_2 is introduced into the NO being absorbed, the concentration of SO_2 rapidly drops to 0 ppm and the concentration of NO followed down to a minimum of 100 ppm accompanied by an increase in the concentration of NO_2 from 20 to 40 ppm. Similarly, as shown in Fig. 3C2, the concentration of SO_2 decreases faster than that of NO when SO_2 and NO are in contact with NaHCO_3 at the same time, and the absorption of SO_2 is significantly higher than that of NO , which indicates that NaHCO_3 preferentially absorbs SO_2 relative to NO .

Another finding is that the amount of NO_2 present is much smaller than the amount of NO absorbed, implying that a large number of nitrogenous species are stored in the absorber. The distribution of nitrogen species given in Fig. 4A for other nitrogenous species exceeds 60% of the NO feeding, indicating that the S550N500 absorber removes NO mainly through the formation of other nitrogenous species. However, it is uncertain whether NO_2 is an active intermediate species for NO removal, so experiments of mixed NO_2 and SO_2 removal were conducted. The results in Fig. 4b show that the NO_2 concentration is reduced from 360 to 140 ppm with 60% removal when SO_2 is coexisted and this implies that NaHCO_3 has good absorption properties for NO_2 . Therefore, NO_2 is an important intermediate for NO removal, and it is the unabsorbed NO_2 that is detected in the gas.



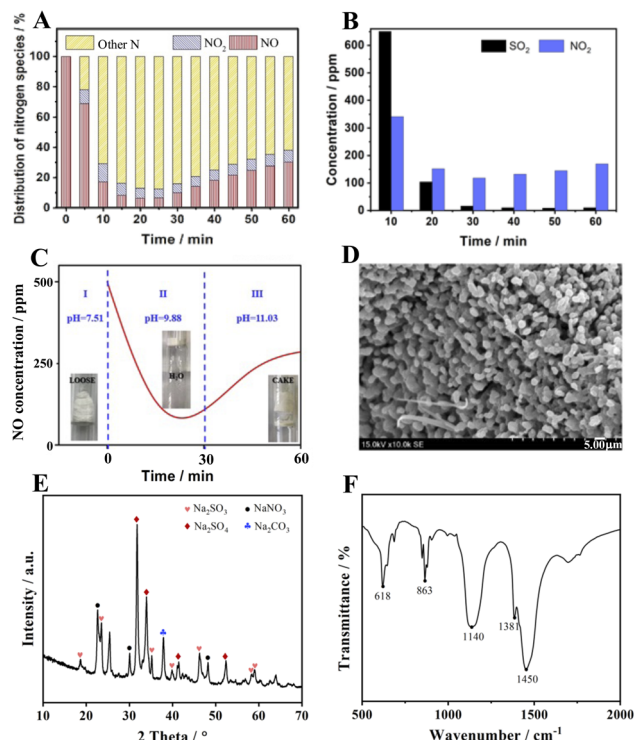


Fig. 4 (A) Nitrogen species distribution of NO absorbed by S550N500, (B) absorption properties of NO₂ in the coexistence of SO₂, (C) pH evolution of the absorption process of S550N500, (D) SEM image, (E) XRD spectra and (F) FTIR spectra of the absorber S550N500 after reaction with SO₂ and NO.

In addition, the changes in absorber morphology and pH were also recorded, and the results are shown in Fig. 4C. The unreacted NaHCO₃ is a loose, weakly basic salt with a pH of 7.51. When the reaction proceeds for 30 min, a large amount of water mist appears in the tube due to the release of water molecules from the thermal decomposition of NaHCO₃, at which time the pH is 9.88. It is noteworthy that the best NO removal efficiency is achieved at this time. As the reaction is prolonged, the absorber gradually condenses into lumps as the pH increases to 11.0, leading to a decrease in NO absorption performance. The SEM image of S550N500 after reaction with SO₂ and NO is shown in Fig. 4D. Unlike the morphology of NaHCO₃ after thermal decomposition, crystalline products cover the surface of S550N500, confirming the occurrence of the chemisorption reaction.

The XRD data of the absorption products are shown in Fig. 4E, and not surprisingly, Na₂CO₃, the thermal decomposition product of NaHCO₃, is detected at 37.8°. The characteristic peaks of Na₂SO₃ (JCPD 37-1488) appear at 18.7°, 23.5°, 35.2°, 40.0°, 46.4°, 58.3° and 59.1°, which are easily oxidized to Na₂SO₄ with $2\theta = 31.8^\circ, 33.9^\circ, 41.4^\circ,$ and 52.3° (JCPD 37-1475). In addition, diffraction peaks located at 22.6°, 30.1° and 48.2° are attributed to NaNO₃ according to JCPD 36-1474, confirming the chemical reaction between NO and the absorbent.

Fig. 4F shows the FTIR spectra of the samples after the absorption reaction. In addition to the peaks of Na₂CO₃ at 863

and 1450 cm⁻¹, the characteristic peaks corresponding to Na₂SO₄ appear at 618 and 1140 cm⁻¹, while that of NaNO₃ is at 1381 cm⁻¹, implying that the reaction products of NaHCO₃ with SO₂ and NO mainly include Na₂CO₃, Na₂SO₄ and NaNO₃.^{25–27}

The reaction products were then characterized by XPS, and the high-resolution spectra of S 2p and N 1s were illustrated in Fig. 5. For S 2p, the peaks with binding energy located at 166.7 ± 0.2 eV can be attributed to Na₂SO₃ and NaHSO₃, while the peaks at 168.8 eV correspond to the mixed Na₂S₂O₅ and Na₂S₂O₆,^{28–30} which are generated by the following eqn (6)–(13).³¹ In addition, two characteristic peaks corresponding to the product Na₂SO₄ (Na₂SO₄ (1) and Na₂SO₄ (2)) are found at 168.4 eV and 169.6 eV, respectively.³² This is consistent with Tim C. Keener that NaHCO₃ can react directly with SO₂ or with SO₂ after thermal decomposition, which gives rise to a multi-path reaction scheme between NaHCO₃ and SO₂ with temperature-sensitive properties.^{24,33} When the reaction temperature reaches 413 K, the formation of Na₂SO₄ (1) by SO₂ uptake occurs mainly in the Na₂CO₃ micropores produced by the thermal decomposition of NaHCO₃,⁸ and the amount of this Na₂SO₄ production is huge, accounting for 61% of the total S of S550N0. Moreover, Na₂SO₄ (2) located at 169.6 eV can be reasonably attributed to the direct reaction generation at the S550N0 surface with a percentage of about 4.4%.

For S550-N500, N500-S550 and S550N500, four forms of sulfur species existed after simultaneously absorption of SO₂ and NO, including two types of Na₂SO₄, Na₂S₂O₅ and Na₂S₂O₆, with Na₂S₂O₆ and Na₂SO₄ (2) predominating. However, the introduction of NO affects the distribution of sulfur species. Notably, when NO is present in the absorbed gas, Na₂SO₃/NaHSO₃ disappears while Na₂SO₄ (2) is enriched, indicating that the conversion of SO₃²⁻ to SO₄²⁻ is promoted.

The N 1s XPS spectra given in Fig. 5B can distinguish two typical components with peaks located at 407.3 eV and 403.9 eV, which belong to sodium nitrate (NaNO₃) and sodium nitrite (NaNO₂), respectively.^{34,35} Combined with the absorption performance of NO in the coexistence of SO₂ in Fig. 3, it is confirmed that the main products of NO conversion are nitrate nitrogen (NO₂⁻ and NO₃⁻).

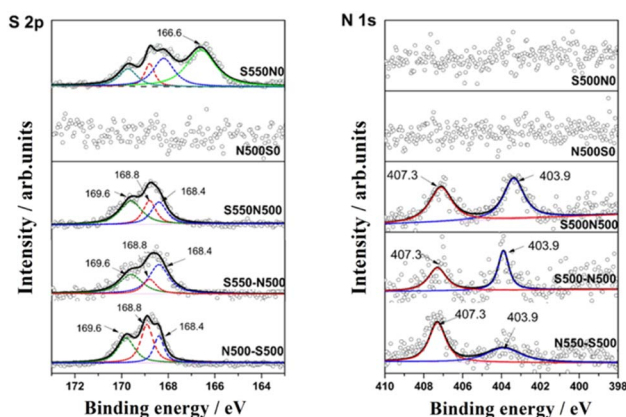


Fig. 5 Fitted S 2p and N 1s spectra of S550N0, N500S0, S550-N500, N500-S550 and S550N500.



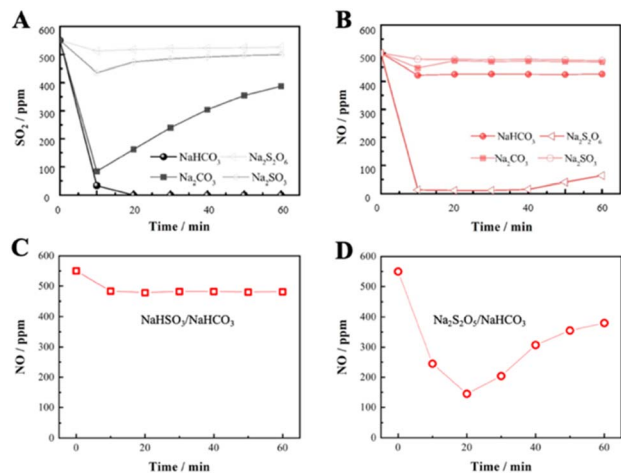


Fig. 6 (A) Desulfurization performance of NaHCO_3 , Na_2CO_3 , Na_2SO_3 , and $\text{Na}_2\text{S}_2\text{O}_6$. (B) Denitrification performance NaHCO_3 , Na_2CO_3 , Na_2SO_3 and $\text{Na}_2\text{S}_2\text{O}_6$. (C) Denitrification performance $\text{NaHSO}_3/\text{NaHCO}_3$. (D) Denitrification performance $\text{Na}_2\text{S}_2\text{O}_5/\text{NaHCO}_3$.

3.3 Performance of Na/SO_x in absorbing NO

Based on the above, the desulfurization/denitrification performance of specific intermediates was further investigated, including Na_2CO_3 , Na_2SO_3 and $\text{Na}_2\text{S}_2\text{O}_6$. From Fig. 6A, NaHCO_3 is the best absorbent for desulfurization, and Na_2CO_3 is easily saturated by SO_2 adsorption, while the desulfurization efficiency of Na_2SO_3 and $\text{Na}_2\text{S}_2\text{O}_6$ is poor. During denitrification, as shown in Fig. 6B, $\text{Na}_2\text{S}_2\text{O}_6$ is the only effective absorber that can immediately react with NO and maintain prolonged uptake of NO. This indicates that $\text{Na}_2\text{S}_2\text{O}_6$ generated by the desulfurization is mainly responsible for the absorption reaction of NO. Further, an absorber with a 1 : 4 ratio of NaHSO_3 or $\text{Na}_2\text{S}_2\text{O}_5$ to NaHCO_3 was configured for probing its absorption activity of NO. As shown in Fig. 6C and D, NaHSO_3 has almost no denitrification activity, while $\text{Na}_2\text{S}_2\text{O}_5$ performs well in absorbing NO, reducing the concentration from 550 to 150 ppm. According to eqn (7)–(12), $\text{Na}_2\text{S}_2\text{O}_5$ as a reactant can produce $\text{Na}_2\text{S}_2\text{O}_6$, which promotes the absorption and conversion of NO.

3.4 Effect of O₂ on the absorption of SO₂ and NO

The role of O_2 in NO and SO_2 adsorption and removal was further investigated. The results in Fig. 7A show that without O_2 ,

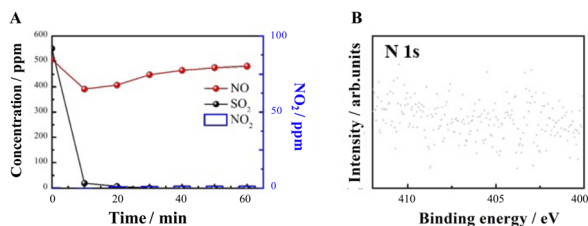


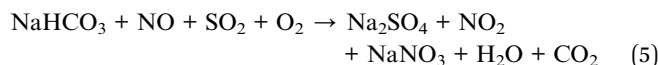
Fig. 7 SO_2 and NO uptake on NaHCO_3 in the absence of O_2 . (A) $T = 160\text{ }^\circ\text{C}$, $\text{GHSV} = 10\,000\text{ h}^{-1}$, flow rate = 1 L min^{-1} , $\text{SO}_2 = 550\text{ ppm}$, $\text{NO} = 500\text{ ppm}$, $\text{O}_2 = 0\text{ vol}\%$ and the total sorption time is 60 min. (B) Fitted N 1s spectra.

SO_2 sorption is not affected, but NO removal and NO_2 generation are significantly inhibited. XPS results (Fig. 7B) show that both NO_2^- and NO_3^- are not generated in the absence of O_2 , speculating O_2 is necessary for the synergistic removal of NO by SO_2 and NaHCO_3 .

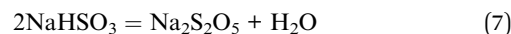
3.5 Gibbs free energy of reactions

For the possible intermediate species and reaction pathways involved in the process of SO_2 and NO uptake by NaHCO_3 , the following schematic diagram 8 and eqn (5)–(14) are described, and the relevant Gibbs free energy (ΔG) is calculated.

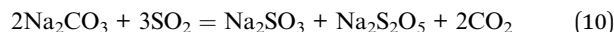
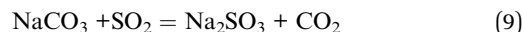
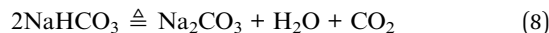
Overall reaction:



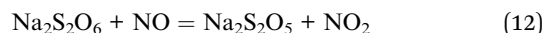
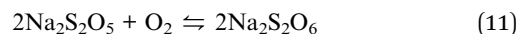
Step A:



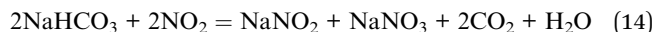
Step B:



Step C:



Step D:



The generation of $\text{Na}_2\text{S}_2\text{O}_5$ from NaHCO_3 during desulfurization is thermodynamically favorable, and steps A and B are considered as two possible reactions with ΔG of -3276 and -5105 kJ mol^{-1} , respectively. Step B shows that the generation of $\text{Na}_2\text{S}_2\text{O}_5$ from Na_2CO_3 is also thermodynamically feasible, which ΔG is about -1281 kJ mol^{-1} . This indicates that NaHCO_3 can be directly or thermally decomposed into Na_2CO_3 to react with SO_2 to form $\text{Na}_2\text{S}_2\text{O}_5$. Steps C and D, as key steps in NO removal, describe the oxidation of NO to NO_2 by O_2 and $\text{Na}_2\text{S}_2\text{O}_5$ and further conversion of NO_2 to NO_x^- with ΔG of 675 kJ mol^{-1} (11), -158 kJ mol^{-1} (12), -858 kJ mol^{-1} (13) and -1721 kJ mol^{-1} (14), respectively. This means that the oxidation of $\text{Na}_2\text{S}_2\text{O}_5$ to $\text{Na}_2\text{S}_2\text{O}_6$ is the controlling step of the whole reaction (Table 2).

In summary, pathways A and B describe the chemical interaction between NaHCO_3 and SO_2 to produce the intermediate $\text{Na}_2\text{S}_2\text{O}_5$. According to path C, the reaction properties of



Table 2 The standard molar Gibbs energy ($\Delta_f G^\circ$) of solid phase species,³⁶ and the standard entropy and enthalpy data are referred to thermochemical tables at $p^0 = 0.1$ Mpa, $T = 500$ K in website <http://kinetics.nist.gov/janaf/>. The energetics and Gibbs free energy of gas phase species (E_{DFT}) are obtained by DFT calculations

Structure	$\Delta_f G^\circ$ kJ mol ⁻¹	Structure	E_{DFT} eV	$\Delta H_g(0 \text{ K} \rightarrow 500 \text{ K}, p^0)$ kJ mol ⁻¹	$S_g(T, p^0)$ J mol ⁻¹ K ⁻¹	$G_g(500 \text{ K}, p^0)$ kJ mol ⁻¹
NaHCO ₃	-852	O ₂	-9.8	4.34	172.20	-1031.2
Na ₂ CO ₃	-1048	SO ₂	-11.0	8.76	270.49	-1200
Na ₂ SO ₃	-1012	NO	-12.2	6.06	226.26	-1290.4
Na ₂ SO ₄	-1270	NO ₂	-15.5	8.1	260.64	-1626.4
NaNO ₂	-285	H ₂ O	-13.6	6.92	206.53	-1415.8
NaNO ₃	-367	CO ₂	-22.7	8.30	234.90	-2304.9
Na ₂ S ₂ O ₅	-1354	N ₂	-16.4	5.91	206.74	-1683.7
Na ₂ S ₂ O ₆	-1532					

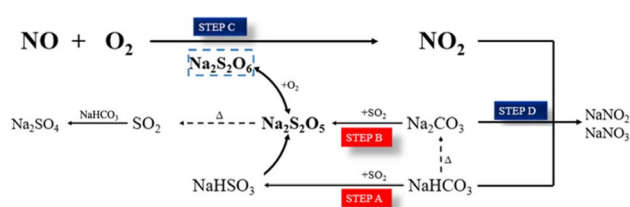


Fig. 8 A proposed reaction mechanism.

NaHCO₃ with NO depend on the formation of Na₂S₂O₆ from Na₂S₂O₅. The NO molecule acts as a Lewis acid and reacts on the surface of Na₂S₂O₆ to form NO₂.³⁷ Gradually, the generated NO₂ achieves a dynamic equilibrium of adsorption-desorption, and then reacts with NaHCO₃ to form NO_x⁻. The removal mechanism and reaction pathway are shown in Fig. 8.

Considering that SO₂ and O₂ are the key factors that synergistically promote NO absorption, it is necessary to optimize the content of coexisting SO₂ and O₂ for improving the NO absorption performance. The ration of SO₂ and NO (S/N) of the absorbed gases are set to 0.5, 1, 2, 3, and 4, respectively, where the NO concentration is fixed at 500 ppm and the SO₂ concentration corresponds to 250, 500, 1000, 1500, and 2000 ppm. The results in Fig. 9A show that S/N has a small effect on the SO₂ removal, which is maintained above 95%. However, the efficiency of NO removal is only 34% when S/N is 0.5, which is considered that the oxidative absorption of NO is limited by the small number of reactive intermediates generated on the absorber surface. With S/N higher than 1, the absorption efficiency of NO is stabilized at about 90%. As the S/N increases to

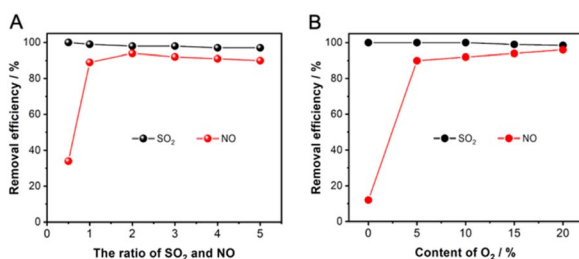


Fig. 9 Effect of (A) the ration of SO₂ and NO and (B) the content of O₂ on the adsorption of SO₂ and NO.

1, the NO removal stabilizes to about 90%, emphasizing the necessity of SO₂ for NO absorption. As shown in Fig. 9B, the O₂ concentration does not affect the SO₂ absorption efficiency, which is always maintained at about 100%. However, when the O₂ content is increased from 0 to 5%, the NO removal efficiency dramatically increases from 11% to about 91%, and increasing the content of O₂ to 20% has no significant effect on the NO removal. Therefore, the NO removal by NaHCO₃ has the advantage of being highly adaptable to SO₂ and O₂ content of flue gas.

4. Conclusions

The NaHCO₃ simultaneous removal of SO₂ and NO perform a good activity. NaHCO₃ can completely adsorb SO₂ and absorb 90% of NO at 160 °C. The NO oxidation observed after the reaction with SO₂ and NaHCO₃ is due to the formation of Na/SO_x species. The most likely reactive Na/SO_x demonstrated theoretically and experimentally include Na₂S₂O₆ and Na₂S₂O₅. In addition, coexisting SO₂ and O₂ are the main factors for forming Na/SO_x. The Na₂S₂O₅ formed by the reaction of SO₂ with NaHCO₃ further reacts with O₂ to form Na₂S₂O₆, which could significantly alter the overall rate of the NO oxidation. Mechanistic studies have shown that the oxidation of NO to NO₂ is the key step for its efficient uptake and requires additional energy. This study clearly demonstrates the removal of NO by NaHCO₃ in the presence of SO₂ and O₂, which will be instructive for the synergistic removal of SO₂ and NO.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by Sichuan Science and Technology Program (2023YFS0354).

Notes and references

- 1 Y. Zhao, T. Guo, Z. Chen and Y. Du, *Chem. Eng. J.*, 2010, **160**, 42–47.



- 2 C.-F. Liu and S.-M. Shih, *Ind. Eng. Chem. Res.*, 2006, **45**, 8765–8769.
- 3 Y. Liu, W. Xu, L. Zhao, Y. Wang and J. Zhang, *Energy Fuels*, 2017, **31**, 12364–12375.
- 4 A. G. Chmielewski, E. Zwolińska, J. Licki, Y. Sun, Z. Zimek and S. Bułka, *Radiat. Phys. Chem.*, 2018, **144**, 1–7.
- 5 R.-B. Lin, S.-M. Shih and C.-F. Liu, *Chem. Eng. Sci.*, 2003, **58**, 3659–3668.
- 6 J.-H. Park, J.-W. Ahn, K.-H. Kim and Y.-S. Son, *Chem. Eng. J.*, 2019, **355**, 351–366.
- 7 I. Bahrabadi-Jovein, S. Seddighi and J. Bashtani, *Energy Fuels*, 2017, **31**, 14007–14017.
- 8 C. Wu, S.-J. Khang, T. C. Keener and S.-K. Lee, *Adv. Environ. Res.*, 2004, **8**, 655–666.
- 9 S. Wang, S. Xu, S. Gao, P. Xiao, M. Jiang, H. Zhao, B. Huang, L. Liu, H. Niu, J. Wang and D. Guo, *Sci. Rep.*, 2021, **11**, 11003.
- 10 L. Xu, J. Guo, F. Jin and H. Zeng, *Chemosphere*, 2006, **62**, 823–826.
- 11 R. Sundararaman and C. Song, *Energy Fuels*, 2013, **27**, 6372–6376.
- 12 A. Dal Pozzo, D. Guglielmi, G. Antonioni and A. Tugnoli, *J. Cleaner Prod.*, 2017, **162**, 1061–1074.
- 13 J. Vehlow, *Waste Manage.*, 2015, **37**, 58–74.
- 14 M. A. Paisley and M. Millspaugh, in *19th Annual North American Waste-to-Energy Conference*, ASMEDC, Lancaster, Pennsylvania, USA, 2011, pp. 1–8.
- 15 S. Kalisz, R. Wejkowski, I. Maj and P. Garbacz, *Energy*, 2023, **279**, 128046.
- 16 National Energy Technology Laboratory, *Integrated Dry NO_x/SO₂ Emissions Control System, A DOE Assessment*, 2001.
- 17 B. Walawska, A. Szymanek and A. Pajdak, *Chemik*, 2013, **67**, 903–912.
- 18 G. Kresse and J. Furthmüller, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **54**, 11169–11186.
- 19 G. Kresse and J. Furthmüller, *Comput. Mater. Sci.*, 1996, **6**, 15–50.
- 20 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 21 P. E. Blöchl, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **50**, 17953–17979.
- 22 Y. Liu, W. Cen, Z. Wu, X. Weng and H. Wang, *J. Phys. Chem. C*, 2012, **116**, 22930–22937.
- 23 S. Ma, X. Bie, C. Gong, B. Qu and D. Liu, *RSC Adv.*, 2021, **11**, 8846–8856.
- 24 T. C. Keener and S.-J. Khang, *Chem. Eng. Sci.*, 1993, **48**, 2859–2865.
- 25 X. Zhang, S. Wang, J. Li, X. Xiao and S. Shu, *Chem. Eng. Sci.*, 2023, **281**, 119157.
- 26 T. Zhou, Y. Zhao, X. Xiao, Y. Liu, H. Bai, X. Chen, J. Dou and J. Yu, *ACS Omega*, 2022, **7**, 29171–29180.
- 27 T. Ishizuka, H. Kabashima, T. Yamaguchi, K. Tanabe and H. Hattori, *Environ. Sci. Technol.*, 2000, **34**, 2799–2803.
- 28 D. Littlejohn and S.-G. Chang, *J. Electron Spectrosc. Relat. Phenom.*, 1995, **71**, 47–50.
- 29 K. Stejskalová, I. Spirovová, E. Lippert, K. Mocek and Z. Bastl, *Appl. Surf. Sci.*, 1996, **103**, 509–516.
- 30 B. J. Lindberg, K. Hamrin, G. Johansson, U. Gelius, A. Fahlman, C. Nordling and K. Siegbahn, *Phys. Scr.*, 1970, **1**, 286–298.
- 31 K. He, S. Su, S. Ding and W. Sun, *React. Kinet., Mech. Catal.*, 2018, **123**, 757–770.
- 32 H. Peisert, T. Chassé, P. Streubel, A. Meisel and R. Szargan, *J. Electron Spectrosc. Relat. Phenom.*, 1994, **68**, 321–328.
- 33 S. Enami, C. D. Vecitis, J. Cheng, M. R. Hoffmann and A. J. Colussi, *J. Phys. Chem. A*, 2007, **111**, 13032–13037.
- 34 K. Burger, F. Tschisnarov and H. Ebel, *J. Electron Spectrosc. Relat. Phenom.*, 1977, **10**, 461–465.
- 35 M. Datta, H. J. Mathieu and D. Landolt, *Appl. Surf. Sci.*, 1984, **18**, 299–314.
- 36 Á. Vegas, J. F. Liebman and H. D. B. Jenkins, *Acta Crystallogr., Sect. B: Struct. Sci.*, 2012, **68**, 511–527.
- 37 M. C. Paganini, M. Chiesa, P. Martino, E. Giamello and E. Garrone, *J. Phys. Chem. B*, 2003, **107**, 2575–2580.

