



Cite this: *RSC Adv.*, 2023, 13, 17869

Received 17th April 2023

Accepted 7th June 2023

DOI: 10.1039/d3ra02556j

rsc.li/rsc-advances

Sb(III)/Gum Arabic composite as a new natural-based environmentally green catalyst for the one-pot pseudo-four-component synthesis of 2*H*-indazolo[2,1-*b*] phthalazinetriones†

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Sb(III)–Gum Arabic composite as a unique natural-based and nontoxic catalyst was synthesized and characterized by FT-IR, XRD, TGA, ICP, BET, EDX and mapping. A four-component reaction of phthalic anhydride, hydrazinium hydroxide, aldehyde, and dimedone has been carried out in the presence of Sb(III)/Gum Arabic composite to synthesise 2*H*-indazolo[2,1-*b*] phthalazine triones. The advantages of the present protocol are the appropriate reaction times, eco-friendly nature and high yields.

Introduction

Recently, science and technology are shifting emphasis toward green and environmental friendly organic processes.¹ Catalysis is one of the pillars of green chemistry. Indeed, one of the main strategies for the minimization of the amount of unwanted by-products is the replacement of stoichiometric chemical procedures by efficient catalytic protocols. Organic reactions under solvent-free conditions have attracted much interest from chemists, particularly from the viewpoint of green chemistry.² In this regard, bio-supported and solid heterogeneous catalysts for the synthesis of fine chemicals have attracted considerable interest for both environmental and economical reasons.³ Gum Arabic (GA) is a natural polymer that consists mainly of high molecular weight polysaccharides and is widely used as a stabilizer, thickening agent and hydrocolloid emulsifier. It is used in the food, textile, pottery, lithography, cosmetics and pharmaceutical industries.⁴ The existence of OH groups in Gum Arabic are suitable moieties for bonding of elemental cations to form supported Lewis acids. Heterocyclic compounds containing phthalazine moiety are of interest because they show some biological and pharmacological activities. Moreover, these compounds can be used as new fluorescence probes or luminescent materials.⁵ In the past decade, several methods for the synthesis of 2*H*-indazolo[2,1-*b*] phthalazinetriones using different catalysts such as silica sulfuric acid,⁶ phosphosulfonic acid,⁷ *N*-halosulfonamides,⁸ cyanuric chloride,⁹ H₂SO₄,¹⁰

dodecylphosphonic acid,¹¹ and [PVPH]ClO₄ (ref. 12) have been reported.

The above mentioned catalysts have high yields,^{6–8,11} short reaction times,^{7,10,11} easy work-up procedure,^{7,10,11} solvent-free condition^{6,11} and reusability of catalyst.¹²

Some procedures have disadvantages such as strong acid condition¹⁰ and no reusability catalyst.^{8–10}

Sb(III) is a good Lewis acid for activation of carbonyl compound.

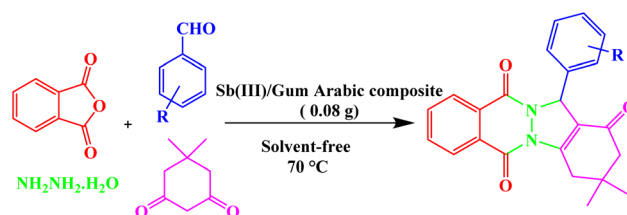
Herein, we have supported Sb(III) on Gum Arabic to preparation of Sb(III)/Gum Arabic as a heterogenous acid catalyst. Thus, we report an eco-friendly protocol for the synthesis of 2*H*-indazolo[2,1-*b*] phthalazinetriones in the presence of Sb(III)/Gum Arabic composite as a new natural-based green reusable catalyst *via* the reaction of dimedone, phthalic anhydride, hydrazinium hydroxide and aldehydes (Scheme 1).

Experimental

Materials

Chemicals were purchased from Merck, Fluka and Aldrich Chemical Companies.

Fourier transform infrared (FT-IR) spectra recorded by ATR method on a Bruker (EQUINOX 55) spectrometer. The nuclear



Scheme 1 Synthesis of 2*H*-indazolo[2,1-*b*] phthalazinetriones.

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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3ra02556j>



magnetic resonance (NMR) spectra were recorded in DMSO- d_6 on Bruker (DRX-400, Avance) NMR 400 MHz. Melting points were determined by a Büchi B-540 instrument. Field emission scanning electron microscopy (FESEM) apparatus (Mira 3-XMU) was used for recording of FESEM images. The XRD graph of catalyst was obtained by X-ray diffractometer (XRD, Bruker-binary V3) using a Cu $k\alpha$ anode ($k = 1.54 \text{ \AA}$, radiation at 36 kV and 36 mA) in the 2θ range from 10° to 80° . Energy-dispersive X-ray spectrometer (EDS) and maps of catalyst were recorded by Phenom pro X. Thermal gravimetric analysis (TGA) was done using "BÄHR-(model: STA 503)" instrument. BELSORP MINI II nitrogen adsorption apparatus (Japan) was used for recording of Brunauer–Emmett–Teller (BET) specific surface area of catalyst at 77 K. Inductively coupled plasma mass spectrometry (ICP-MS) was recorded by AGILENT 7500 apparatus. All IR and ^1H NMR spectra data are available in ESI (Fig. S1–S20).†

Catalytic test

General procedure for the synthesis of 2*H*-indazolo[1,2-*b*]phthalazinetriones. Firstly, a mixture of phthalic anhydride (1 mmol) and hydrazinium hydroxide (1.2 mmol) was mixed at 70°C for 20 min to form phthalhydrazide. Then, aldehyde (1 mmol), dimedone (1 mmol) and Sb(III)/Gum Arabic composite (0.08 g) were added to above mentioned mixture and stirred under solvent-free conditions at 70°C . After the completion of the reaction, monitored by TLC, (*n*-hexane : ethyl acetate (8 : 2)), 5 mL of hot ethanol was added to the mixture of reaction and the catalyst was separated by filtration. By adding H_2O (5 mL) to filtrate, the solid product was appeared and then recrystallized from ethanol to give the pure product.

Results and discussion

Synthesis and characterization of Sb(III)/Gum Arabic composite

For the synthesis of Sb(III)/Gum Arabic, a mixture of Gum Arabic (0.5 g), $\text{Sb}(\text{OAc})_3$ (0.5 g) and dichloromethane (10 mL) was refluxed for 2 hours. Then solid catalyst was filtered, washed with dichloromethane and dried at room temperature.

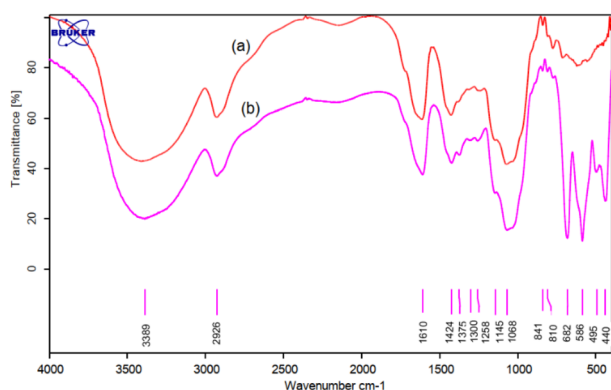


Fig. 1 FT-IR spectra of (a) Gum Arabic, (b) Sb(III)/Gum Arabic composite.

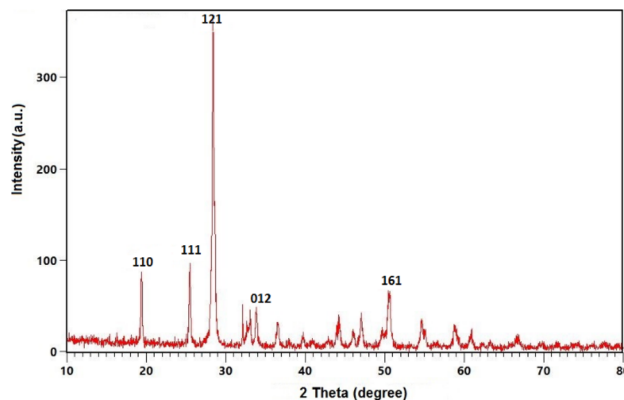


Fig. 2 XRD pattern of Sb(III)/Gum Arabic composite.

The structure of prepared Sb(III)/Gum Arabic composite was studied using FTIR, XRD, BET, FE-SEM, TGA, EDX(EDS), elemental mapping and ICP. Fourier transform infrared (FT-IR) spectroscopy spectra of Gum Arabic and Sb(III)/Gum Arabic composite were shown in Fig. 1. In Gum Arabic spectrum, the characteristic absorption bands at 1068 and 1145 cm^{-1} (C–O, stretch), 1610 (O–H, bending), 2926 cm^{-1} (C–H, stretch), and

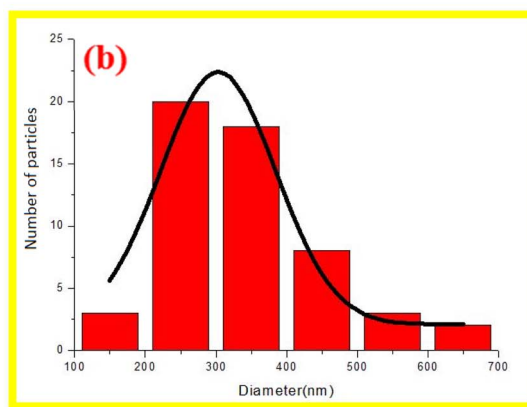
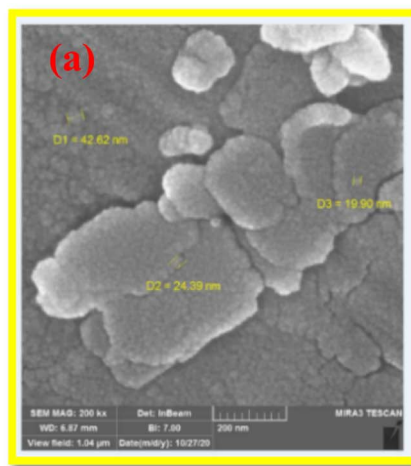


Fig. 3 FESEM image (a) and particle size distribution histogram of Sb(III)/Gum Arabic composite (b).



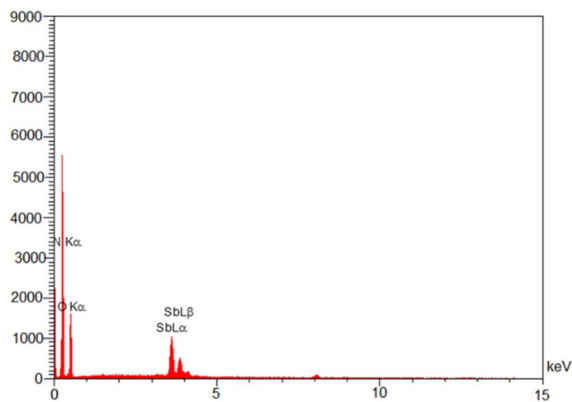


Fig. 4 EDX patterns of Sb(III)/Gum Arabic composite.

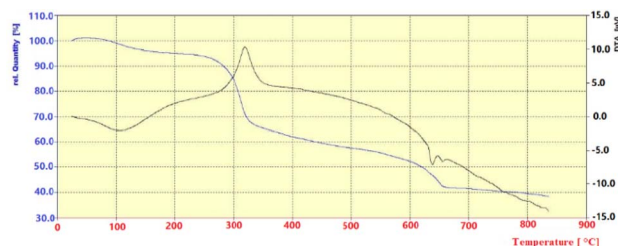


Fig. 5 TGA diagram of Sb(III)/Gum Arabic composite.

Table 1 Obtained parameters from porosity analysis

BET plot	
V _{micro} , (cm ³ g ⁻¹)	0.5
S(BET), (m ² g ⁻¹)	2.2
V (Σ), (cm ³ g ⁻¹)	0.006
D _{pore} , (nm)	10.5

3389 cm⁻¹ (O–H, stretch) can be observed. In Sb(III)/Gum Arabic composite spectrum, the additional bands at 586 and 682 are assigned to the Sb–O stretching vibrations.

The XRD pattern of Sb(III)/Gum Arabic composite in 2θ = 10–80° is shown in Fig. 2. The broad peak at 2θ = 19.43 relates to the amorphous Gum Arabic and 2θ = 25.46, 28.41 attributed to Sb(III).

The surface morphology of Sb(III)/Gum Arabic composite was characterized using FESEM analysis (Fig. 3). According to the FESEM images, the particles are spherical with the size in the range of nanomaterial. The histogram of the size distribution detected by MIRA 3 TESCAN software in FESEM.

Energy-dispersive X-ray spectroscopy EDS (EDX) analysis was applied for the identification of elemental composition in Sb(III)/Gum Arabic composite (Fig. 4). The EDX data confirmed the existence of N, O and Sb elements in the catalyst.

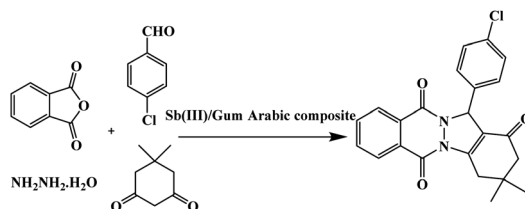
The weight percentages of Sb(III) in catalyst was studied by inductively coupled plasma (ICP). According to obtained data, the weight percentage of Sb(III) is 43% (Fig. S21, ESI†).

The elemental mapping of Sb(III)/Gum Arabic composite was shown in (Fig. S22, ESI†) which confirmed homogenous distribution of C, N, O, Sb in catalyst.

Fig. 5 shows the thermal gravimetric analysis (TG-DTA) of the Sb(III)/Gum Arabic composite in 20–900 °C. The obtained results indicated that there were three weight loss steps. The first observed weight loss of sample occurred at 100–200 °C due to removal of the catalyst moisture. The second step of weight loss, in 280–320 °C, was a result of dehydration and decarboxylation of Gum Arabic. The third step of weight loss occurs in the 343–663 °C range, most likely due to complete decomposition of the glycoprotein of Gum Arabic. Finally, the high weight of ash can be improved the presence of Sb(III) ion in catalyst.

The nitrogen adsorption–desorption graph, BET plot, Langmuir plot, *t* plot and BJH plot were shown in (Fig. S23, ESI†),

Table 2 Synthesis of 3,4-dihydro-3,3-dimethyl-13-(4-chlorophenyl)-2*H*-indazolo[2,1-*b*] phthalazine-1,6,11(13*H*)-trione in the presence of Sb(III)/Gum Arabic composite under different conditions^a



Entry	Solvent	Conditions	Sb(III)/Gum Arabic composite (g)	Time (min)	Yield ^b (%)
1	—	R.T	0.08	120	28
2	—	70 °C	0.08	30	97
3	—	Grinding/60 °C	0.08	80	60
4	—	50 °C	0.08	85	40
5	—	65 °C	0.08	25	75
6	EtOH	Reflux	0.08	50	62
7	H ₂ O	Reflux	0.08	120	54
8	—	70 °C	0.07	40	80
9	—	70 °C	0.06	45	74

^a Phthalic anhydride (1 mmol), hydrazinium hydroxide (1.2 mmol), 4-chlorobenzaldehyde (1 mmol) and dimedone (1 mmol). ^b Isolated yield.



which were used to determine the specific surface area, volume and distribution of pores. Nitrogen adsorption was measured on gas-depleted samples at 77 K. The adsorption-desorption isotherm is the type of III while indicating the presence of broader pore size distributions with H3 type hysteresis loop for Sb(III)/Gum Arabic composite, which is consistent with the IUPAC classification. The Brunauer–Emmett–Teller surface area (SBET), total pore volume and mean pore diameter of Sb(III)/Gum Arabic composite were determined as $2.2 \text{ m}^2 \text{ g}^{-1}$, $0.006 \text{ cm}^3 \text{ g}^{-1}$ and 10.5 nm , respectively (Table 1).

Catalytic properties of Sb(III)/Gum Arabic composite

In order to optimize reaction conditions for synthesis of 2*H*-indazolo[2,1-*b*] phthalazinetrione in the presence of Sb(III)/Gum Arabic composite, the condensation of dimedone, 4-chlorobenzaldehyde, phthalic anhydride and hydrazinium hydroxide was selected as a model reaction. The obtained results of this reaction such as temperature, solvent and amount of catalyst were summarized in Table 2.

Efficiency of Sb(III)/Gum Arabic composite

According to obtained optimal conditions, the synthesis of 2*H*-indazolo[1,2-*b*] phthalazinetriones were done in the presence of Sb(III)/Gum Arabic composite at 70°C and solvent-free conditions, which the results are summarized in Table 3.

According to Hammet correlation, the presence of electron-withdrawing group in *para* position of aromatic aldehyde (4-COOH $\sigma_{para} = 0.45$, 4-Cl $\sigma_{para} = 0.23$, 4-NO₂ $\sigma_{para} = 0.78$) caused increasing of the yield of reaction and electron-donating group

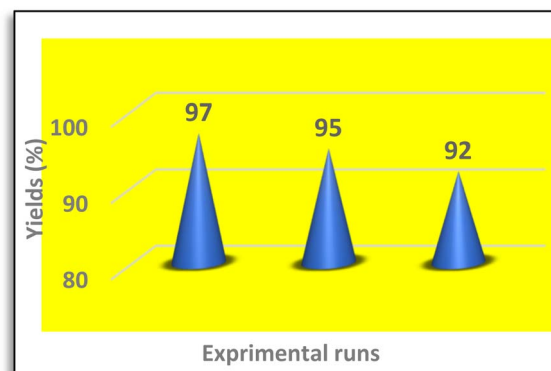


Fig. 6 Reusability of Sb(III)/Gum Arabic composite.

in *para* position (4-OH $\sigma_{para} = -0.37$, 4-CH₃ $\sigma_{para} = -0.17$, 4-OCH₃ $\sigma_{para} = -0.27$) reduced the yield of reaction.

Reusability of Sb(III)/Gum Arabic composite

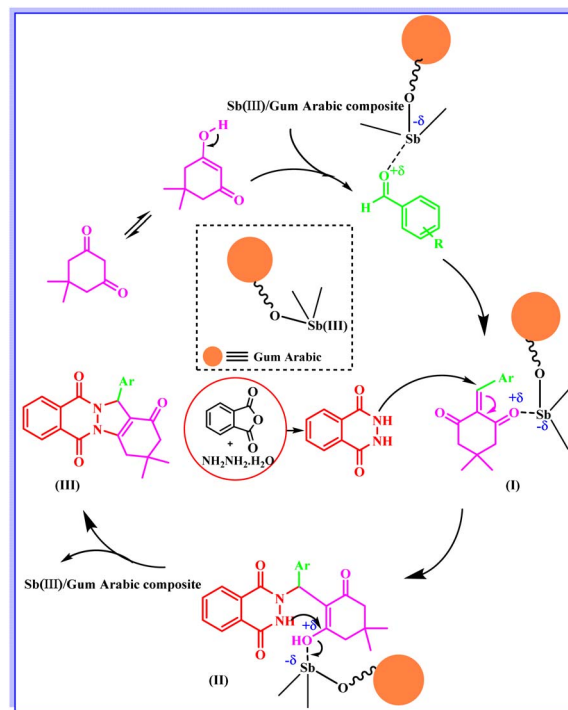
The reusability of the catalyst was examined in the model reaction for three times. The result showed low decrease of catalytic activity (Fig. 6).

The proposed mechanism for the synthesis of 2*H*-indazolo[2,1-*b*] phthalazinetrione is shown in Scheme 2. It was anticipated that, at first, the Sb(III) Lewis acid in Sb(III)/Gum Arabic composite activates the carbonyl group of aldehyde to condensation with dimedone to form adduct (I). In continue, phthalhydrazide react with adduct (I) *via* a Michael-type addition to

Table 3 Synthesis of 2*H*-indazolo[2,1-*b*] phthalazinetrions derivatives in the presence of Sb(III)/Gum Arabic composite at 70°C and solvent-free conditions^a

Entry	R	Time (min)	Yield ^b (%)	m.p. ($^\circ\text{C}$)
1	4-HCOO	120	94	265–267
2	4-Cl	30	97	263–265
3	3-Cl	35	95	185–187
4	4-OH	35	90	262–264
5	3-NO ₂	32	91	270
6	4-CH ₃	40	89	232–234
7	4-OCH ₃	43	90	214–216
8	2-OH-5-Br	37	92	265
9	4-NO ₂	30	93	228–230
10	4-Isopropyl	60	87	258–260

^a Reaction conditions: aldehyde (1 mmol), phthalic anhydride (1 mmol), hydrazinium hydroxide (1.2 mmol) and dimedone (1 mmol), Sb(III)/Gum Arabic composite (0.08 g) solvent-free and 70°C . ^b Isolated yield.



Scheme 2 A plausible mechanism for the formation of 2*H*-indazolo[2,1-*b*] phthalazinetrione in the presence of Sb(III)/Gum Arabic composite.



Table 4 Comparison of Sb(III)/Gum Arabic composite with other reported catalysts in model reaction^a

Ent.	Catal.	Cond.	Time (min)	Yield (%) [ref.]	Product (mmol)/ catalyst (g)	Catalyst (g)/ reaction mixture (g)
1	CAN (5 mol%)	PEG 400, 50 °C	120	94 (ref. 13)	0.94/0.027	0.027/0.44
2	H ₂ SO ₄ (0.15 mmol)	[bmim]BF ₄ , 80 °C	30	92 (ref. 10)	0.92/0.015	0.015/0.44
3	Iodine (0.1 g)	Sonic bath	10	92 (ref. 14)	0.92/0.1	0.1/0.44
4	A (5 mol%)	S.F. 100 °C	6	93 (ref. 15)	0.93/0.1	0.1/0.44
5	B (0.02 g)	S.F. 110 °C	10	95 (ref. 16)	0.95/0.02	0.02/0.44
6	C (0.08 g)	S.F. 70 °C	30	97 [this work]	0.097/0.08	0.08/0.44

^a A: poly phosphoric acid–SiO₂, B: Fe₃O₄@SiO₂–ZrCl₂–MNPs, C: Sb(III)/Gum Arabic composite.

produce the intermediate (II). Finally, the OH group in intermediate (II) was activated by catalyst to production of product (III) via an intramolecular cyclization.

The comparison between our reported protocol with others was done and the results were tabulated in Table 4.

Conclusions

In conclusion, an easy and green method for the synthesis of 2*H*-indazolo[2,1-*b*] phthalazintriones using an eco-friendly and reusable catalyst was described. Some advantages of this protocol is mild reaction condition, solvent free and high yields of products.

Data availability

All data generated or analysed during this study are included in this published article.

Author contributions

MK, BFM and AB designed and performed the research, analysed the data, interpreted the results, and prepared the manuscript. MK performed the assay and conducted the optimization, and purification of compounds. All authors read and approved the final manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This study was financially supported by Yazd University. The funding bodies played no role in the design of the study and collection, analysis, and interpretation of data and in writing the manuscript. The authors thank the Research Council of Yazd University for the support of this research.

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