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Ag₂NCN anchored on Ti₃C₂T_x MXene as a Schottky heterojunction: enhanced visible light photocatalytic efficiency of rhodamine B degradation†

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The quick charge recombination of light-generated electrons and holes severely restricts the photocatalytic applications of single semiconductors. Here, a straightforward electrostatically driven self-assembly technique was used to construct an Ag₂NCN/Ti₃C₂T_x Schottky heterojunction, which was then used to degrade Rhodamine B (RhB) in the illumination of visible light. The findings from the experiments revealed that as a cocatalyst, Ti₃C₂T_x significantly suppresses the recombination rate and broadens visible absorptivity to improve Ag₂NCN photocatalytic efficiency. The optimized Ag₂NCN/Ti₃C₂T_x (AT2) composite exhibited an outstanding photocatalytic rate in 96 min, with the highest RhB degradation rate ($k = 0.029 \text{ min}^{-1}$), which was around fifteen times that of pure Ag₂NCN ($k = 0.002 \text{ min}^{-1}$). Furthermore, the trapping-agent experiment showed photogenerated superoxide radicals and holes were the principal active agents inside the photodegradation of RhB. Compared with Ag-based semiconductors, the composite exhibited outstanding photostability, highlighting its excellent potential for application in visible-light photocatalysis.

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Introduction

Nowadays, the kinds and amounts of chemicals entering water bodies through various routes have drastically increased as a result of the quick development of modern industry. Additionally, organic effluent includes a number of toxic contaminants that are challenging to degrade. Among them, RhB is highly refractory to degradation and is easily accumulated; as a result, the environment and public health are seriously threatened by RhB contamination.^{1,2} The main techniques employed for RhB removal include adsorption, reverse osmosis membrane filtration, and advanced oxidation processes.^{3–5} Notably, photocatalysis has attracted considerable attention as a solution to the environmental issues caused by organic pollutants, such as dyes, antibiotics, and persistent organic compounds.^{6–9} It is an eco-friendly and highly efficient technology for dealing with organic pollutant wastewater, owing to its easy operation, lower energy consumption, and minimal secondary pollution,¹⁰ and it could be an efficient strategy for

RhB removal. To date, a number of semiconductor photocatalysts for environmental cleanup have been developed, including TiO₂, Ag₃PO₄, and g-C₃N₄.^{11–13} Although these photocatalysts exhibit remarkable performance in photocatalytic reactions, it is still challenging to apply them. This is because solo catalytic photoinduced carrier recombination is relatively fast. In addition, photocatalysts have a low specific surface area, weak photostability, short light absorption range, and limited active sites.¹⁴ Hence, highly efficient and stable photocatalysts are required. Structure improvement and composition improvement are two essential methods used for enhancing photocatalyst performance,^{15,16} where the creation of heterojunction interfaces by joining various semiconductors is one of the most efficient ways to separate or utilize charges and increase active sites.¹⁷

According to theory, silver cyanamide (Ag₂NCN) has a strong photocatalytic activity and an excellent visible-light response due to its substantial nitrogen components and naturally occurring conductive carbon (C-sp²) content. The bandgap for Ag₂NCN is smaller compared to the silver-oxide instance. Due to the strong dispersion influence of the energy band, it may simultaneously increase electron mobility, decrease electron-hole recombination, and improve photocatalytic conversion efficiency.^{18–20} Nevertheless, the application of Ag₂NCN as a photocatalyst is restricted due to the rapid recombination of photoinduced carriers, low light absorption, and photocatalytic instability caused by the self-photoreduction of Ag⁺.²¹

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Constructing a semiconductor heterojunction is considered one of the most effective strategies for solving the aforementioned problems.^{22,23} MXene ($\text{Ti}_3\text{C}_2\text{T}_x$), a novel 2D transition-metal carbide nanosheet, is similar to graphene.²⁴ $\text{Ti}_3\text{C}_2\text{T}_x$ has recently attracted significant attention as a cocatalyst in the field of photocatalysis. This is due to its huge surface area, which includes plenty of uncovered active sites. Meanwhile, the efficient migration of charge carriers on its surface forms a close connection with other semi-conductors to construct the heterojunction.²⁵ In addition, $\text{Ti}_3\text{C}_2\text{T}_x$ has been confirmed to occupy a suitable Fermi-level (EF) position to construct the Schottky heterojunction at the interface of composites.²⁶ For example, Cai *et al.* prepared $\text{Ag}_3\text{PO}_4/\text{Ti}_3\text{C}_2$, which displayed noticeably improved photocatalytic degradation activity and stability for organic pollutants in comparison to Ag_3PO_4 .²⁷ This renders it an effective cocatalyst to depress the recombination of photoproduced carries, thereby improving its photocatalytic performance.²⁸

Here, the Schottky heterojunction between the interfaces of Ag_2NCN and $\text{Ti}_3\text{C}_2\text{T}_x$ was constructed by electrostatically driving the self-assembly to prepare $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ composites. As compared to pure Ag_2NCN , these composites demonstrated very effective photocatalytic performance as well as stability for the degradation of RhB. The optimized $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ (AT2) composite demonstrated an extraordinary photocatalytic rate in 96 min, with the greatest RhB degradation efficiencies of 98.8%. A large special surface area, prominent metallic conductivity, and favorable EF position of the 2D $\text{Ti}_3\text{C}_2\text{T}_x$ played a key role in the enhancement of photocatalysis. These advantages render $\text{Ti}_3\text{C}_2\text{T}_x$ an excellent cocatalyst to construct a Schottky heterojunction, which can efficiently depress the recombination of carries and widen the light absorption during the photodegradation, resulting in $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ nanocomposites with improved efficacy and stability.

Experimental

Synthesis of Ag_2NCN

In a typical synthesis,²⁹ 0.85 g (5 mmol) of AgNO_3 was first dissolved in 20 mL of deionized water in a conventional synthesis. Following that, 125 mL of $\text{NH}_3 \cdot \text{H}_2\text{O}$ (1.5 M) aqueous solution was swiftly added to make a clear solution while magnetic stirring was performed. The translucent solution became golden yellow after 5 mL of an aqueous solution containing 0.21 g (5 mmol) of cyanamide (H_2NCN) was added dropwise. Subsequently, the precipitate was rinsed three times with deionized water and centrifuged at 3500 rpm for 5 minutes. The resulting product was dried in a vacuum (60 °C) overnight and designated as AT0 for short.

Synthesis of $\text{Ti}_3\text{C}_2\text{T}_x$ pseudo-solution

$\text{Ti}_3\text{C}_2\text{T}_x$ was fabricated by etching Ti_3AlC_2 . LiF (3.2 g) was mixed with 40 mL of hydrochloric acid (9 M) to make the etching solution. To begin, 1.0 g of Ti_3AlC_2 was gently added to 20 mL of etching solution and then held at 40 °C for 48 hours with constant stirring. The product was washed with hydrochloric acid and deionized water until the pH reached a value greater than 6.

The power was then diffused in 100 mL of deionized water and delaminated for 1 hour under N_2 flow using ice bath sonication. The product was then centrifuged for 1 hour at 3500 rpm, and the supernatant was collected. The concentration of $\text{Ti}_3\text{C}_2\text{T}_x$ pseudo-solution was calculated by filtering a portion of the supernatant and weighing the dried film mass. Finally, the colloidal solution of $\text{Ti}_3\text{C}_2\text{T}_x$ was diluted to obtain a concentration of 10 mg mL^{-1} .³⁰

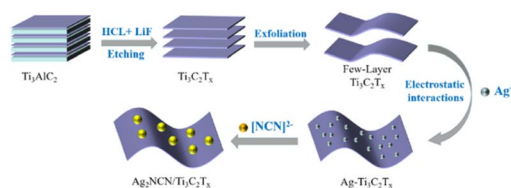
Synthesis of $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ nanocomposites

$\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ was prepared through an electrostatically driven self-assembly process. Firstly, 0.85 g (5 mmol) of AgNO_3 was dissolved in 20 mL of deionized water. Subsequently, 125 mL of $\text{NH}_3 \cdot \text{H}_2\text{O}$ (1.5 M) aqueous solution was quickly added to form a transparent solution under magnetic stirring. A certain amount of $\text{Ti}_3\text{C}_2\text{T}_x$ colloid (0.5 mL, 1 mL, 2 mL, and 3 mL) was then dropwise added into the above transparent solution and stirred at room temperature for 20 min. After 5 mL of an aqueous solution containing 0.21 g (5 mmol) of cyanamide (H_2NCN) was added dropwise with stirring for 30 min, and the precipitate obtained was dried in an oven at 60 °C to obtain $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ (AT0.5, AT1, AT2, and AT3 for short). The detailed synthetic procedure for the preparation of $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ composites was shown in Scheme 1.

Material characterizations

X-ray powder diffraction (XRD, Bruker D8 advance) was used to analyze the crystal structures of the samples, with Cu K radiation ($\lambda = 0.15328$ nm) as the radiation source and a scanning rate of 5° per min ranging from 5 to 90°. Transmittance electron microscopy was used to examine the surface morphology of the samples (TEM, FEI Tecnai G2 F20). X-ray photoelectron spectroscopy was used to determine the elemental makeup of the produced materials (XPS, Thermo Scientific K-Alpha). The light absorption was measured using diffuse ultraviolet-visible reflectance (UV-vis DRS, Hitachi U-3900). Photoluminescence (PL) study was performed on the performance of photo-generated carrier transfer using a transient fluorescence spectrophotometer (Hitachi F-7000) with an excitation wavelength of 325 nm, scanning speed of 240 nm min^{-1} , and PMT voltage of 650 V.

Photoelectrochemical experiments were carried out using a CHI660E electrochemical analyzer, which was utilized to record the electrochemical response using a typical three-electrode setup (A working electrode, an Ag/AgCl reference electrode, and a Pt wire counter electrode). The AT0 and AT2



Scheme 1 Schematics illustration for the preparation of $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$.



were coated on fluorine-doped tin oxide (FTO) conductive glass and naturally dried during the transient photocurrent measurement. The light source was a 300 W xenon lamp with a bias voltage of 0.5 V. Additionally, as the electrode solution, Na_2SO_4 (0.5 M) solution was employed.

Photocatalytic performance tests

The photocatalytic activity of the catalysts as prepared was evaluated through experiments involving the photocatalytic degradation of RhB under visible light exposure. The light source used was a 300 W Xe lamp with a cut-off filter ($\lambda > 420$ nm). To reach an adsorption–desorption balance, photocatalysts (100 mg) were distributed in 200 mL of RhB (20 mg L^{-1}) reaction solution, which was stirred in a lightless environment for 30 minutes. At regular intervals, 3 mL of the suspension was collected and centrifuged (8000 rpm, 10 min) to separate the catalysts for further analysis. A UV-vis spectrophotometer was used to measure the concentration of Rhodamine B (MAPADA, UV-1800PC).

To assess the photocatalytic stability of the AT2 nanocomposite, cycle photocatalytic studies were carried out using the aforementioned procedures. The catalyst was gathered after every cycle of the photocatalytic procedure, cleaned with deionized water, and then re-dispersed into a new RhB solution for the following cycle. Trapping agent studies were used to investigate the photocatalytic mechanism. Except that a small quantity of scavenger is introduced to the reaction system, the experimental procedure is quite identical to the photo-degradation experiment.

Results and discussion

Characterization of structures

The crystal structures of as-fabricated Ag_2NCN , $\text{Ti}_3\text{C}_2\text{T}_x$, and $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ were initially examined using XRD. According to Fig. 1a, after the etching treatment by LiF and hydrochloric acid, the peak at 39.2° (104) in the Ti_3AlC_2 pattern disappeared. The (002) and (004) diffraction peaks at 9.6° and 19.3° , respectively, in the XRD pattern of $\text{Ti}_3\text{C}_2\text{T}_x$ shifted to a lower angle, indicating that the Al layer in Ti_3AlC_2 was removed and that Ti_3AlC_2 was mostly converted into $\text{Ti}_3\text{C}_2\text{T}_x$.³¹ When H_2NCN solution was dropped into a solution containing silver ammonia solution and a little quantity of $\text{Ti}_3\text{C}_2\text{T}_x$, diffraction peaks were observed for Ag_2NCN (Fig. 1b). According to Jia *et al.*, the characteristic 2θ peaks of 19° , 32° , and 37.9° can be indexed to the (110), (102), and (300) crystal planes, respectively (JCPDS 70-5232).¹⁹ This proved that pure Ag_2NCN could be synthesized and that the structure of Ag_2NCN was unaltered by the synthetic procedure. It is worth noting that a trace quantity of metallic Ag emerged in the XRD pattern of $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$, which might be owing to Ag self-reduction in $\text{Ti}_3\text{C}_2\text{T}_x$ MXene solution produced by low-valence Ti.^{32,33} According to reports in the literature,³⁴ Ag's surface plasmon resonance effect boosts the photoelectric field close to the metal particles and creates resonance absorption in the ultraviolet-visible region, particularly in the visible region. This can increase the noble metal's electron energy density and boost the photocatalytic activity. Due to the

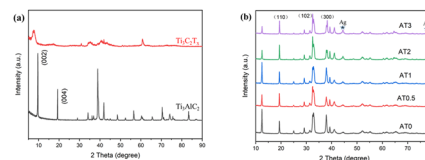


Fig. 1 (a) XRD patterns of Ti_3AlC_2 and $\text{Ti}_3\text{C}_2\text{T}_x$, (b) XRD patterns of Ag_2NCN and $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ hybrids.

low $\text{Ti}_3\text{C}_2\text{T}_x$ concentration in the $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ composition, the characteristic peaks of $\text{Ti}_3\text{C}_2\text{T}_x$ were unable to be found in the XRD patterns.

Morphology analysis

The microstructure of the AT2 photocatalyst was investigated using transmission electron microscopy (TEM). It demonstrated the existence of $\text{Ti}_3\text{C}_2\text{T}_x$. According to Fig. 2a, large lamellar $\text{Ti}_3\text{C}_2\text{T}_x$ sheets and Ag_2NCN nanoparticles were observed in the AT2 composite. Fig. 2b shows that the Ag_2NCN nanoparticles were assembled on the surface of $\text{Ti}_3\text{C}_2\text{T}_x$, indicating the intimate coupling between $\text{Ti}_3\text{C}_2\text{T}_x$ and Ag_2NCN . Moreover, the electrostatically induced self-assembly of Ag^+ and negative charge $\text{Ti}_3\text{C}_2\text{T}_x$ resulted in a very uniform dot-like nanostructure in Ag_2NCN . Notably, the inclusion of $\text{Ti}_3\text{C}_2\text{T}_x$ resulted in a modest reduction in the diameters of the Ag_2NCN particles in the AT2 composite when compared to the diameters of pure Ag_2NCN particles (0.1–1.6 μm). The findings suggested that $\text{Ti}_3\text{C}_2\text{T}_x$ could effectively control the shape and size of Ag_2NCN particles, which facilitated the photocatalytic process.²⁷ The (0110) and (300) crystal planes of $\text{Ti}_3\text{C}_2\text{T}_x$ and Ag_2NCN were represented by the lattice fringes with spacings of 0.266 and 0.238 nm (Fig. 2c and d).³⁵ The TEM results showed that ultrathin, few-layered 2D $\text{Ti}_3\text{C}_2\text{T}_x$ was synthesized with Ag_2NCN with a large area and high quality.

XPS analysis

The XPS analysis of AT2 is shown in Fig. 3. Fig. 3a shows Ti, Ag, O, and C in AT2, related to the presence of Ag_2NCN and $\text{Ti}_3\text{C}_2\text{T}_x$. As shown in Fig. 3b, two characteristic peaks at 373.9 and 367.9 eV were ascribed to the $\text{Ag } 3d_{5/2}$ and $\text{Ag } 3d_{3/2}$ of Ag_2NCN ,³⁶ respectively. Three peaks at 288.9, 286.7, and 284.8 eV were observed in the C 1s pattern (Fig. 3c), ascribable to the C–F, C–O, and C–C bonds.³⁷ The Ti 2p spectra of AT2 included three peaks (Fig. 3d), and the binding energies at 453.4, 458.2, and 463.9 eV corresponded to the Ti–C $2p_{3/2}$, Ti–O $2p_{3/2}$, and Ti–O $2p_{1/2}$, respectively.³⁸ The peak of the C–Ti bond (281.3 eV) was not observed in the C 1s spectra. The distinctive peaks of Ti–C mostly vanished in the Ti 2p spectra, suggesting that several low-valence Ti bonds were broken during the synthesis of AT2.²⁷ The coexistence of Ag_2NCN and $\text{Ti}_3\text{C}_2\text{T}_x$ in the AT2 composite was verified by the XPS findings, which correspond to the previous part of the TEM analysis.

Photocatalytic performance

$\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ composites of various weight proportions were tested for their photocatalytic capabilities using RhB



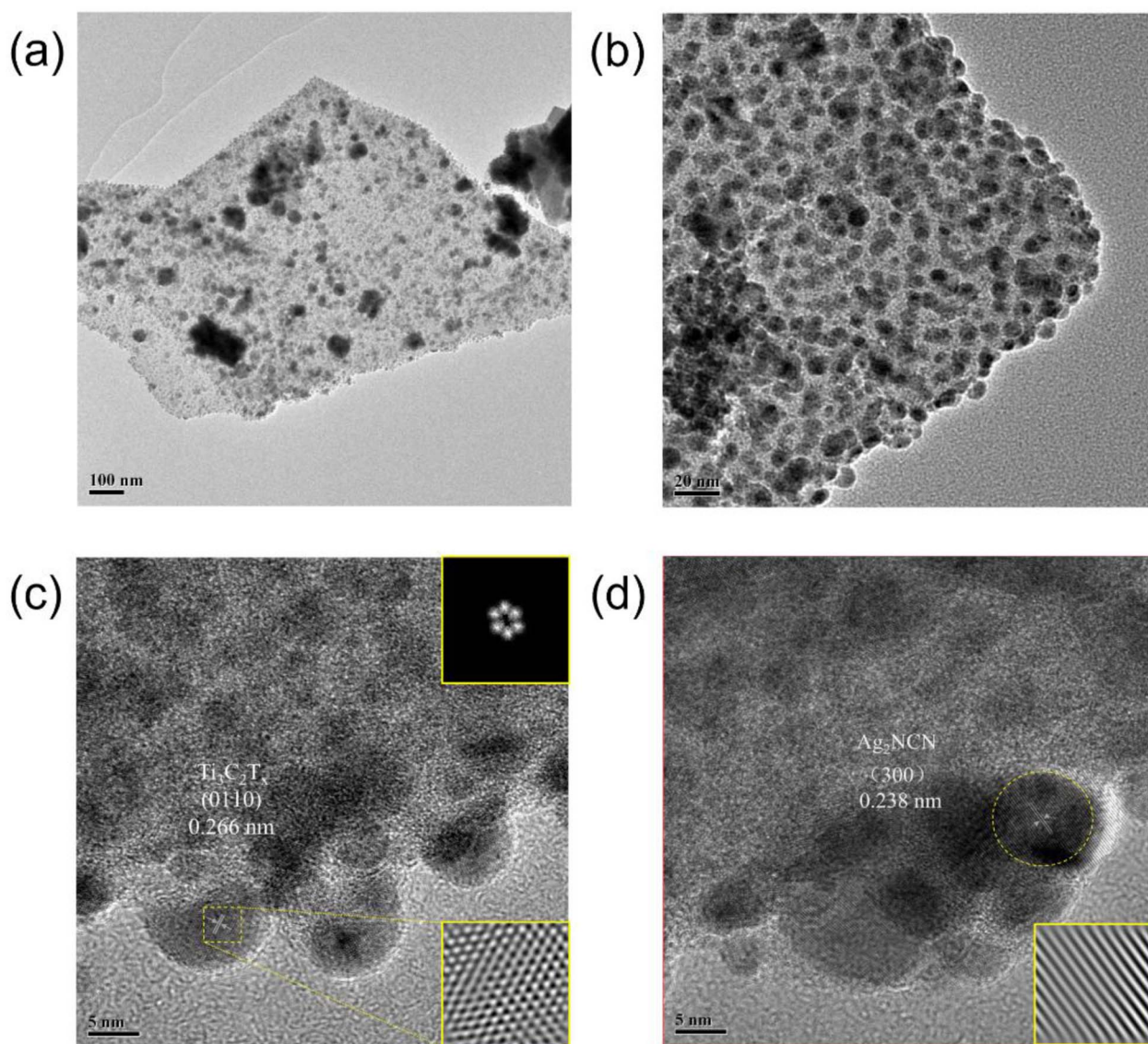


Fig. 2 (a and b) TEM images of AT2, (c and d) high-resolution TEM (HRTEM) images of AT2 composite.

degradation under visible light. According to Fig. 4a, the $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ nanocomposites own better adsorb ability during the adsorption–desorption process than pure Ag_2NCN and nanocomposites outperformed pure Ag_2NCN in terms of photocatalytic activity. $\text{Ti}_3\text{C}_2\text{T}_x$ did not exhibit photoactivity because of its metallic properties.³⁹ Among the catalysts with various weight ratios of $\text{Ti}_3\text{C}_2\text{T}_x$, AT2 exhibited the best photocatalytic activity; its removal rate reached 98.8% in 96 min, nearly two times that of pure Ag_2NCN (45.6%). However, with an increase in the $\text{Ti}_3\text{C}_2\text{T}_x$ content from 2 to 3 mL, a decreased photocatalytic performance was recorded. On the one hand, excess $\text{Ti}_3\text{C}_2\text{T}_x$ decreased the Ag_2NCN content, leading to insufficient photogenerated carriers.⁴⁰ On the other hand, the large weight concentration of $\text{Ti}_3\text{C}_2\text{T}_x$ may have occluded visible light to diminish Ag_2NCN light absorption and light intensity *via* the reaction solution, decreasing photocatalytic performance.⁴¹ The RhB photodegradation over $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ with different $\text{Ti}_3\text{C}_2\text{T}_x$ weight ratios is shown in Fig. 4b and is

matched with a pseudo-first-order kinetics model. For RhB, AT2 demonstrated the highest photocatalytic efficiency in relation to other catalysts.

The initial pH value is an important element in Fenton or Fenton-like reactions, hence the rate of RhB elimination was examined at various pH values (ranging from 3.0 to 5.0 to 6.9 to 9.0 to 11.0). According to Fig. S1(a),† the removal rate of RhB fell from 99.1% to 76.3% in 100 minutes, with the baseline pH increasing from 3.0 to 11.0. Because the pH of the solution changes, the surface charge of $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ changes, affecting the rate of RhB removal. The impact of the $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ dose (0.2, 0.3, 0.4, 0.5, 0.6 g L^{-1}) on the removal rate was examined. The elimination rate of RhB increased with increasing catalyst dose in 96 min, as shown in Fig. S1(b).† The increase of 0.5 g L^{-1} of $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ (AT2) results in the greatest removal rate (95.6%) compared to other dosages, suggesting that increasing the catalyst dosage might give more active sites for RhB adsorption. Fig. S1(c) and (d)† are UV-vis

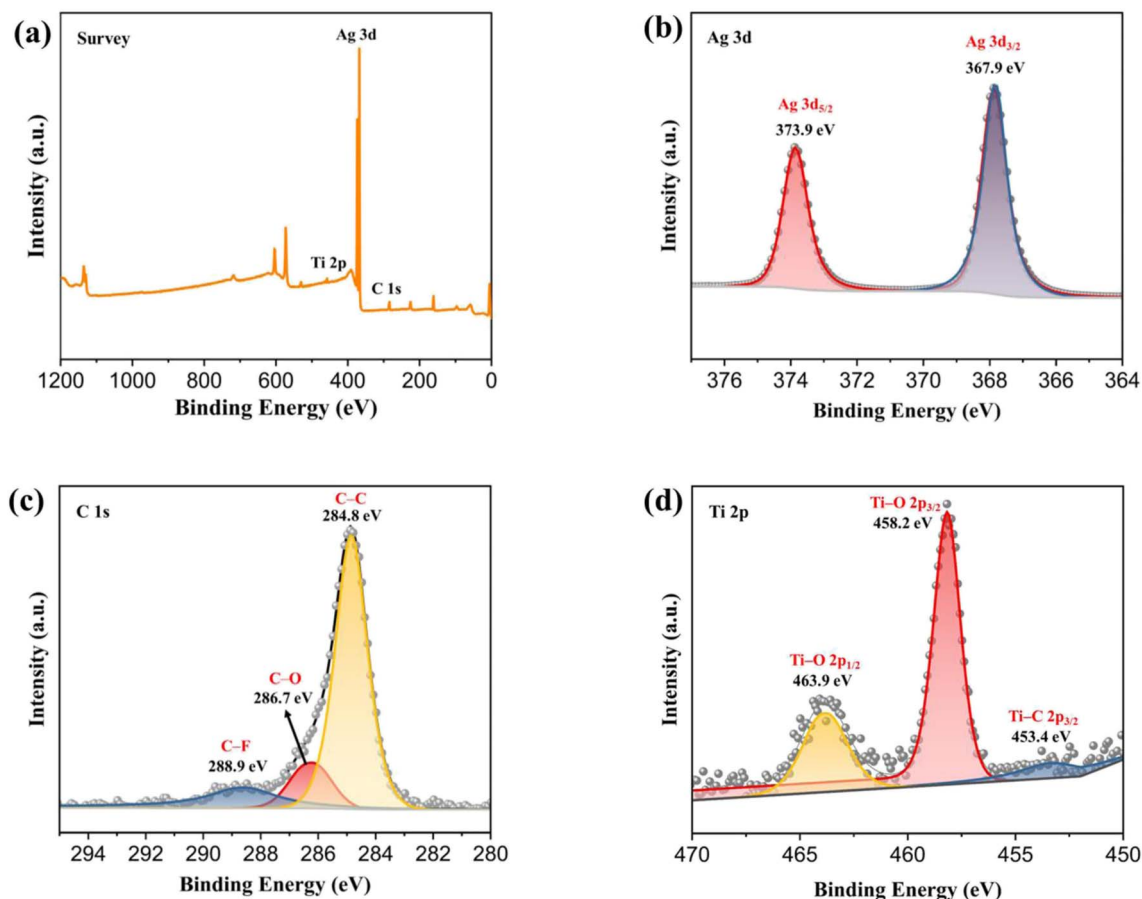


Fig. 3 (a) XPS survey spectrum of AT2 composite; (b–d) High-resolution XPS spectra of Ag 3d, Ti 2p and C 1s of AT2 composite.

spectra of RhB concentration changed with time and the degradation kinetics, under the condition of RhB = 20 mg L⁻¹ 200 mL, Ag₂NCN/Ti₃C₂T_x (AT2) = 0.5 g L⁻¹, pH = 3.0. As the reaction time changes, 99.2% RhB was removed from the solution within 60 min, this reaction belongs to the first order reaction kinetics.

The recyclable stability of a catalyst is critical in assessing its photocatalyst efficacy. A stability test was performed for the AT2 catalyst (Fig. 5). The photodegradation rate of RhB in AT2 reduced significantly (5%) during the second photocatalytic cycle, as illustrated in Fig. 5a. Nevertheless, during the following four cycles, AT2 did not show a decline in the photodegradation efficiency of RhB, showing that the AT2 composite catalyst exhibited high reusability. The XRD pattern of AT2 prior to and after the experiment is shown in Fig. 5b. The results from the FT-IR spectra and the XRD patterns were in agreement (Fig. S2†). The AT2 sample's distinctive peaks didn't seem to have changed much. After five photocatalytic cycles, no changes were observed in the structure of AT2, indicating that the composite catalyst was stable in the photocatalytic degradation reaction. It is usually believed that Ag-based semiconductors are prone to corrosion under light irradiation; however, the photocorrosion of Ag₂NCN was effectively avoided by introducing 2D Ti₃C₂T_x.⁴²

Photoelectrochemical characteristics

To examine the light absorption properties, Fig. 6a shows the UV-vis absorption spectra for Ag₂NCN and AT2. These two samples showed impressive absorption in the scope about 400 nm or less. The absorbance of AT2 photocatalyst increased noticeably between 400 and 800 nm, compared with that of pure Ag₂NCN, indicating that Ti₃C₂T_x successfully coupled with Ag₂NCN. Additionally, photoluminescence emission spectra were used to assess electron and hole coupling (Fig. 6b). The emission peak of AT2 intensity was lower compared to pure Ag₂NCN, demonstrating that an addition of Ti₃C₂T_x reduced the recombination of photogenerated carriers. In order to understand the charge transport process, the transient photocurrent response was investigated. The photocurrent density curves of the Ag₂NCN and AT2 composite exposed to visible light (light on/off cycles: 10 s) are displayed in Fig. 6c. Similar and reproducible photocurrent responses were observed for pure Ag₂NCN and AT2, demonstrating that the photocurrent was reversible. Compared to a single Ag₂NCN, the AT2 composite had a greater photocurrent density. The slightly higher photocurrent density clearly demonstrated that Ag₂NCN combined with Ti₃C₂T_x efficiently boosted the production and segregation of charge carriers. Also, the results of electrochemical impedance spectroscopy (EIS) showed how quickly photogenerated carriers



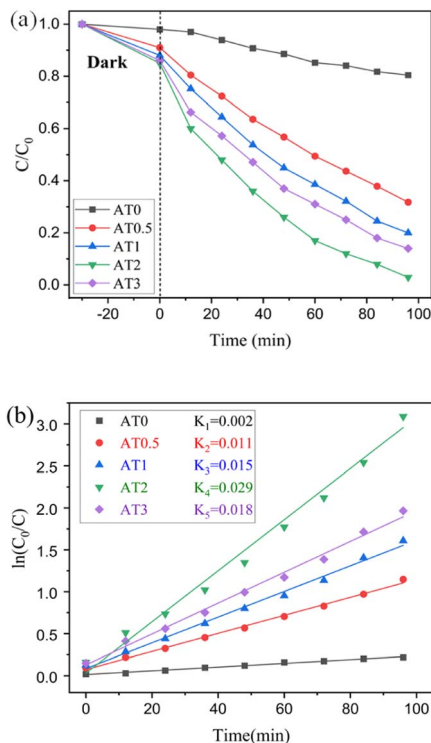


Fig. 4 (a) The degradation rate of dyes with different catalysts; (b) plots of $\ln(C_0/C)$ against reaction time.

transferred. As seen in Fig. 6d, the AT2 composite's semicircle diameter was smaller than that of a single pure Ag_2NCN , indicating that the AT2 carriers had a lower charge transfer resistance and a faster transfer rate. With $\text{Ti}_3\text{C}_2\text{T}_x$ present, the charge transfer resistance was reduced, improving visible light absorption and resulting in a high separation efficiency of photogenerated carriers.⁴³ As a result, AT2 performed better in terms of RhB degradation.

Analysis of the photocatalytic mechanism

Scavengers were utilized to study the photocatalytic reaction process of the AT2 hybrid. According to Fig. 7, during the photocatalytic degradation, disodium

ethylenediaminetetraacetic acid (EDTA-2Na), 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO), and benzoquinone (BQ) were selected as capture agents for h^+ , $\cdot\text{OH}$, and $\cdot\text{O}_2^-$, respectively.^{19,44,45} After the addition of BQ, the rate of RhB degradation drastically reduced, suggesting that $\cdot\text{O}_2^-$ was essential to the photocatalytic reaction. After adding EDTA-2Na, a reduction in the rate of RhB degradation was seen, indicating h^+ was also an important active species. Adding TEMPO had no significant impact on the photocatalytic activity. These findings revealed that $\cdot\text{O}_2^-$ and h^+ were essential for RhB degradation, but $\cdot\text{OH}$ had no effect on the process.

Scheme 2 illustrates a probable photocatalytic reaction mechanism. When exposed to visible light, the light-generated electron of the Ag_2NCN nanoparticles was transported from the VB to the CB (0.12 eV).¹⁹ The EF of $\text{Ti}_3\text{C}_2\text{T}_x$ was approximately -0.36 eV (vs. a typical hydrogen electrode) prior to contact.⁴⁶ After contact, owing to the different work functions between $\text{Ti}_3\text{C}_2\text{T}_x$ and Ag_2NCN , the electrons in $\text{Ti}_3\text{C}_2\text{T}_x$ migrated to Ag_2NCN , which had a higher EF than that of $\text{Ti}_3\text{C}_2\text{T}_x$. As a consequence, the surface of Ag_2NCN would acquire a negative charge, whereas $\text{Ti}_3\text{C}_2\text{T}_x$ would accumulate a positive charge. The electrostatic force is generated internally at the junction of Ag_2NCN and $\text{Ti}_3\text{C}_2\text{T}_x$.⁴⁷ By successfully balancing the two EFs, it could be possible to reduce the recombination of photo-generated carriers and considerably boost the rate at which photoinduced carriers are involved in the photocatalytic activity.⁴⁸ A Schottky junction was also created at the interface after contact with $\text{Ti}_3\text{C}_2\text{T}_x$ caused the CB and VB of Ag_2NCN to flex upward. The generating electrons precluded diffusion return to Ag_2NCN because the Schottky barrier was present.⁴⁹ According to the above analysis, the main possible reactions for the degradation of RhB were shown as the following equations.

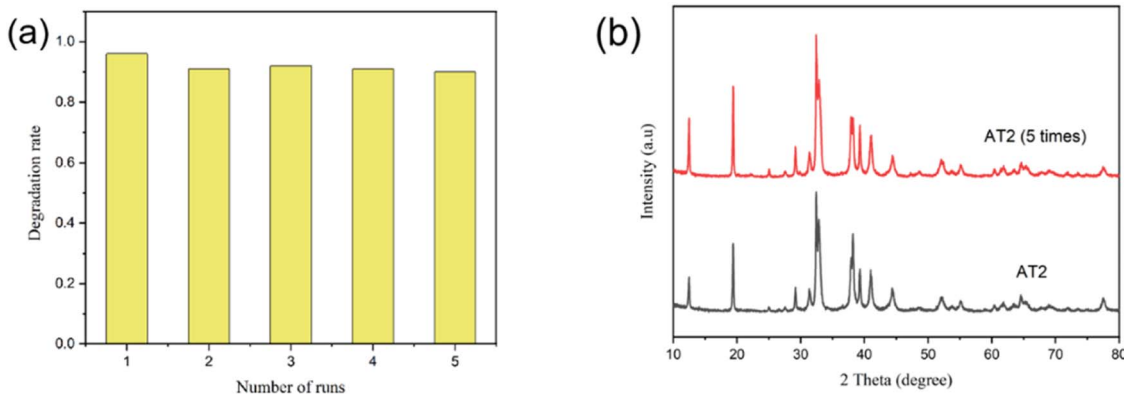
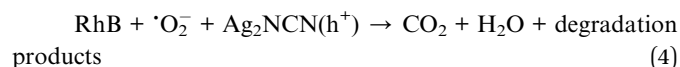
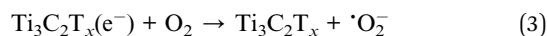
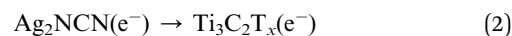
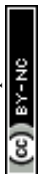


Fig. 5 (a) Cycling utilization of AT2 in the photocatalytic degradation of RhB and (b) XRD patterns of the AT2 after cycling five times.



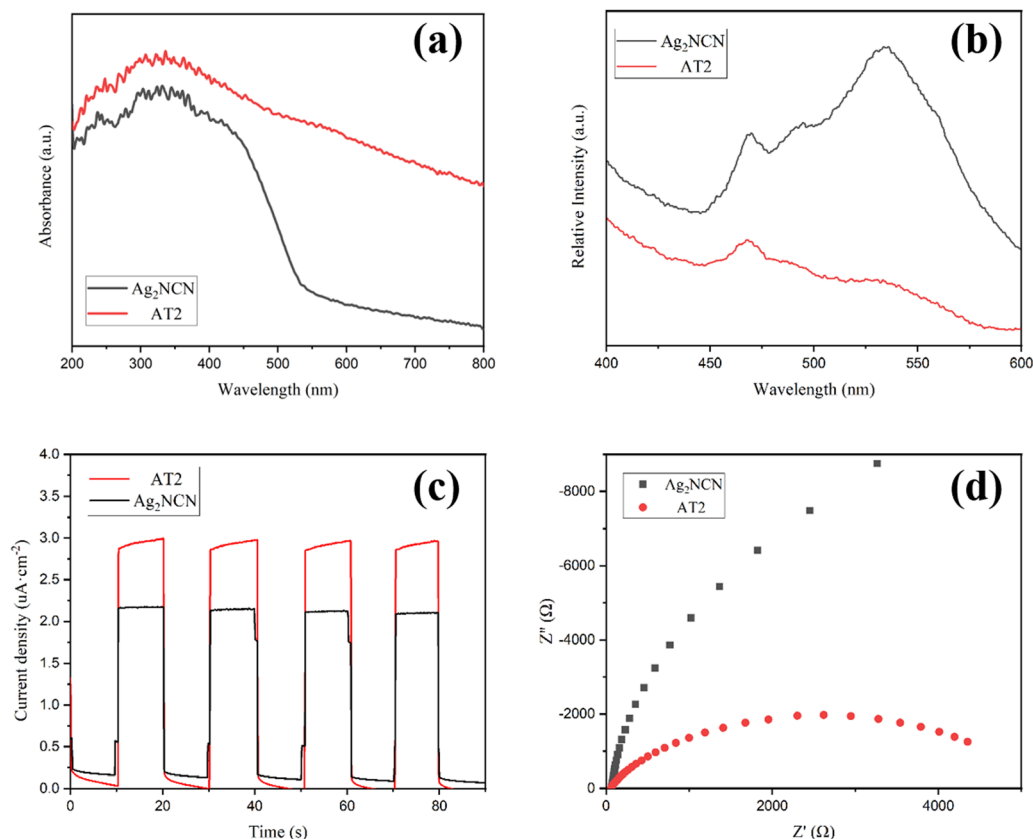


Fig. 6 (a) UV-vis absorption spectra, (b) photoluminescence (PL) emission spectra for pure Ag_2NCN and AT2, (c) transient photocurrent responses, (d) electrochemical impedance spectroscopy (EIS) of pure Ag_2NCN and AT2.

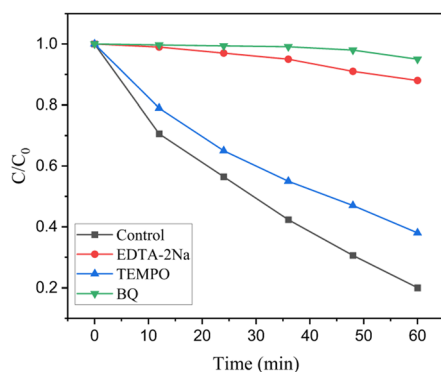
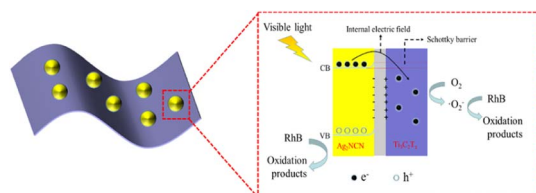


Fig. 7 Photocatalytic performance with scavengers.



Scheme 2 Schematic photocatalytic reaction mechanism of $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ nanocomposite.

Conclusions

In summary, $\text{Ag}_2\text{NCN}/\text{Ti}_3\text{C}_2\text{T}_x$ was constructed by an electrostatically driven self-assembly process and used as the photocatalyst in the degradation of RhB. As compared to pure Ag_2NCN , the AT2 composites exhibited improved photocatalytic effectiveness due to the synergistic effect of Ag_2NCN and $\text{Ti}_3\text{C}_2\text{T}_x$. A superb degradation rate of RhB was achieved for the AT2 composite, which was two times that of pure Ag_2NCN . The AT2 catalyst demonstrated superb photocatalytic stability. The superior visible light absorption and Schottky junction, which promoted charge transport and decreased the combination of carriers due to the close contact between Ag_2NCN and $\text{Ti}_3\text{C}_2\text{T}_x$, are responsible for the improved photocatalytic performance. This work stimulates new ideas for creating sophisticated and effective heterostructure photocatalysts that possess outstanding photocatalytic stability and performance for pollutants removal using Ag-based semiconductors and 2D MXenes materials.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- W. Cheng, Y. Wang, S. Ge, X. Ding, Z. Cui and Q. Shao, *Adv. Compos. Hybrid Mater.*, 2021, **4**, 150–161.
- S. Balasurya, M. K. Okla, A. Mohebaldin, A. A. Al-Ghamdi, M. A. Abdel-Maksoud, B. Almunqedhi, H. AbdElgawad, A. M. Thomas, L. L. Raju and S. S. Khan, *J. Environ. Manage.*, 2022, **308**, 114614.
- C. Berberidou, S. Avlonitis and I. Poulios, *Desalination*, 2009, **249**, 1099–1106.
- L. Huang, X. Huang, J. Yan, Y. Liu, H. Jiang, H. Zhang, J. Tang and Q. Liu, *J. Hazard. Mater.*, 2023, **442**, 130024.
- K. Sharma, D. Vaya, G. Prasad and P. K. Suroliya, *Int. J. Environ. Sci. Technol.*, 2022, **20**, 4615–4634.
- B. Shao, X. Liu, Z. Liu, G. Zeng, Q. Liang, C. Liang, Y. Cheng, W. Zhang, Y. Liu and S. Gong, *Chem. Eng. J.*, 2019, **368**, 730–745.
- Y. Liu, M. Cheng, Z. Liu, G. Zeng, H. Zhong, M. Chen, C. Zhou, W. Xiong, B. Shao and B. Song, *Chemosphere*, 2019, **236**, 124387.
- S. Zhao, C. Chen, J. Ding, S. Yang, Y. Zang and N. Ren, *Front. Environ. Sci. Eng.*, 2021, **16**, 36.
- H. Yu, H. Jiang, S. Zhang, X. Feng, S. Yin and W. Zhao, *Processes*, 2023, **11**.
- N. Boonprakob, N. Wetchakun, S. Phanichphant, D. Waxler, P. Sherrell, A. Nattestad, J. Chen and B. Inceesungvorn, *J. Colloid Interface Sci.*, 2014, **417**, 402–409.
- D. Yu, Y. Zhang, F. Wang and J. Dai, *RSC Adv.*, 2021, **11**, 2307–2314.
- L. Ma, H. Han, L. Pan, M. Tahir, L. Wang, X. Zhang and J.-J. Zou, *RSC Adv.*, 2016, **6**, 63984–63990.
- P. Li, W. Zhang, Y. Zhang, Y. Sun and F. Dong, *RSC Adv.*, 2016, **6**, 96334–96338.
- X. Wang, F. Liang, C. Huang, Y. Li and B. Chen, *Catal. Sci. Technol.*, 2016, **6**, 6551–6560.
- X. Liu, X. Duan, W. Wei, S. Wang and B.-J. Ni, *Green Chem.*, 2019, **21**, 4266–4289.
- R. Zheng, Z. Wang, C. Zhang, B. Chen, H. San and H. Yu, *Composites, Part B*, 2021, **222**.
- Y. Ding, Y. Cao, D. Chen, J. Li, H. Wu, Y. Meng, J. Huang, J. Yuan, Y. Su, J. Wang and H. Li, *Chem. Eng. J.*, 2023, **452**.
- Y. Li, C. Cao, Q. Zhang, Y. Lu, Y. Zhao, Q. Li, X. Li and T. Huang, *RSC Adv.*, 2021, **11**, 10235–10242.
- B. Jia, W. Zhao, L. Fan, G. Yin, Y. Cheng and F. Huang, *Catal. Sci. Technol.*, 2018, **8**, 1447–1453.
- W. Zhao, J. Pan and F. Huang, *Chem. Commun.*, 2018, **54**, 1575–1578.
- C. Li, S. Yu, Y. Wang, J. Han, H. Dong and G. Chen, *J. Taiwan Inst. Chem. Eng.*, 2018, **87**, 272–280.
- C. Jin, S. Rao, J. Xie, Z. Sun, J. Gao, Y. Li, B. Li, S. Liu, L. Liu, Q. Liu and J. Yang, *Chem. Eng. J.*, 2022, **447**.
- C. Yang, Q. Tan, Q. Li, J. Zhou, J. Fan, B. Li, J. Sun and K. Lv, *Appl. Catal., B*, 2020, **268**.
- A. Shahzad, K. Rasool, W. Miran, M. Nawaz, J. Jang, K. A. Mahmoud and D. S. Lee, *ACS Sustainable Chem. Eng.*, 2017, **5**, 11481–11488.
- M. Tahir and B. Tahir, *Chem. Eng. J.*, 2020, **400**.
- L.-f. Hong, R.-t. Guo, Y. Yuan, X.-y. Ji, Z.-s. Li, Z.-d. Lin and W.-g. Pan, *Mater. Today Energy*, 2020, **18**.
- T. Cai, L. L. Wang, Y. T. Liu, S. Q. Zhang, W. Y. Dong, H. Chen, X. Y. Yi, J. L. Yuan, X. N. Xia, C. B. Liu and S. L. Luo, *Appl. Catal., B*, 2018, **239**, 545–554.
- Q. Zhong, Y. Li and G. Zhang, *Chem. Eng. J.*, 2021, **409**.
- X. Meng, P. Yao, Y. Xu, H. Meng and X. Zhang, *RSC Adv.*, 2016, **6**, 61920–61926.
- X. Xie, N. Zhang, Z.-R. Tang, M. Anpo and Y.-J. Xu, *Appl. Catal., B*, 2018, **237**, 43–49.
- Z. Zhuge, X. Liu, T. Chen, Y. Gong, C. Li, L. Niu, S. Xu, X. Xu, Z. A. Allothman, C. Q. Sun, J. G. Shapter and Y. Yamauchi, *Chem. Eng. J.*, 2021, **421**.
- Z. Zhang, H. Li, G. Zou, C. Fernandez, B. Liu, Q. Zhang, J. Hu and Q. Peng, *ACS Sustainable Chem. Eng.*, 2016, **4**, 6763–6771.
- G. Zou, Z. Zhang, J. Guo, B. Liu, Q. Zhang, C. Fernandez and Q. Peng, *ACS Appl. Mater. Interfaces*, 2016, **8**, 22280–22286.
- M. Zayats, S. P. Pogorelova, A. B. Kharitonov, O. Lioubashevski, E. Katz and I. Willner, *Chem.-Eur. J.*, 2003, **9**, 6108–6114.
- B. Cao, S. Wan, Y. Wang, H. Guo, M. Ou and Q. Zhong, *J. Colloid Interface Sci.*, 2022, **605**, 311–319.
- P. Xu, J. Yang, Y. Chen, Y. Li, X. Jia and H. Song, *Mater. Des.*, 2019, **183**.
- S. Cao, B. Shen, T. Tong, J. Fu and J. Yu, *Adv. Funct. Mater.*, 2018, **28**, 1800136.
- W. Yuan, L. Cheng, Y. Zhang, H. Wu, S. Lv, L. Chai, X. Guo and L. Zheng, *Adv. Mater. Interfaces*, 2017, **4**, 20.
- K. Rajavel, S. Shen, T. Ke and D. Lin, *2d Materials*, 2019, **6**, 3.
- X. Chen, Y. Guo, R. Bian, Y. Ji, X. Wang, X. Zhang, H. Cui and J. Tian, *J. Colloid Interface Sci.*, 2022, **613**, 644–651.
- H. Wang, P. Hu, J. Zhou, B. Maarten, J. Roeflaers, B. Weng, Y. Wang and H. Ji, *J. Mater. Chem. A*, 2021, 19984–19993, DOI: [10.1039/d1ta03573h](https://doi.org/10.1039/d1ta03573h).
- T. Cai, L. Wang, Y. Liu, S. Zhang, W. Dong, H. Chen, X. Yi, J. Yuan, X. Xia, C. Liu and S. Luo, *Appl. Catal., B*, 2018, **239**, 545–554.
- J.-X. Yang, W.-B. Yu, C.-F. Li, W.-D. Dong, L.-Q. Jiang, N. Zhou, Z.-P. Zhuang, J. Liu, Z.-Y. Hu, H. Zhao, Y. Li, L. Chen, J. Hu and B.-L. Su, *Chem. Eng. J.*, 2021, **420**.
- W. Li, F. Wang, Y. Shi and L. Yu, *Chin. Chem. Lett.*, 2023, **34**, 1.
- Y. Li, J. Wang, H. Yao, L. Dang and Z. Li, *J. Mol. Catal. A: Chem.*, 2011, **334**, 116–122.
- B. Shao, Z. Liu, G. Zeng, Y. Liu, Q. Liang, Q. He, T. Wu, Y. Pan, J. Huang, Z. Peng, S. Luo, C. Liang, X. Liu, S. Tong and J. Liang, *Appl. Catal., B*, 2021, **286**.
- Y. Liu, Y. H. Li, X. Li, Q. Zhang, H. Yu, X. Peng and F. Peng, *ACS Nano*, 2020, **14**, 14181–14189.
- Y. Chen, T. Shi, P. Liu, X. Ma, L. Shui, C. Shang, Z. Chen, X. Wang, K. Kempa and G. Zhou, *J. Mater. Chem. A*, 2018, **6**, 19167–19175.
- B. Li, S. Liu, C. Lai, G. Zeng, M. Zhang, M. Zhou, D. Huang, L. Qin, X. Liu, Z. Li, N. An, F. Xu, H. Yi, Y. Zhang and L. Chen, *Appl. Catal., B*, 2020, **266**.

