


 Cite this: *RSC Adv.*, 2023, **13**, 15606

# Theoretical investigation of CO<sub>2</sub> capture in the MIL-88 series: effects of organic linker modification†

 Nguyen Thi Xuan Huynh,<sup>a</sup> Ong Kim Le,<sup>bc</sup> Tran Phuong Dung,<sup>cd</sup> Viorel Chihai<sup>e</sup> and Do Ngoc Son <sup>\*bc</sup>

CO<sub>2</sub> capture is a crucial strategy to mitigate global warming and protect a sustainable environment. Metal–organic frameworks with large surface area, high flexibility, and reversible adsorption and desorption of gases are good candidates for CO<sub>2</sub> capture. Among the synthesized metal–organic frameworks, the MIL-88 series has attracted our attention due to their excellent stability. However, a systematic investigation of CO<sub>2</sub> capture in the MIL-88 series with different organic linkers is not available. Therefore, we clarified the topic *via* two sections: (1) elucidate physical insights into the CO<sub>2</sub>@MIL-88 interaction by van der Waals–dispersion correction density functional theory calculations, and (2) quantitatively study the CO<sub>2</sub> capture capacity by grand canonical Monte Carlo simulations. We found that the 1π<sub>g</sub>, 2σ<sub>u</sub>/1π<sub>u</sub>, and 2σ<sub>g</sub> peaks of the CO<sub>2</sub> molecule and the C and O p orbitals of the MIL-88 series are the predominant contributors to the CO<sub>2</sub>@MIL-88 interaction. The MIL-88 series, *i.e.*, MIL-88A, B, C, and D, has the same metal oxide node but different organic linkers: fumarate (MIL-88A), 1,4-benzene-dicarboxylate (MIL-88B), 2,6-naphthalene-dicarboxylate (MIL-88C), and 4,4'-biphenyl-dicarboxylate (MIL-88D). The results exhibited that fumarate should be the best replacement for both the gravimetric and volumetric CO<sub>2</sub> uptakes. We also pointed out a proportional relationship between the capture capacities with electronic properties and other parameters.

Received 10th March 2023

Accepted 13th May 2023

DOI: 10.1039/d3ra01588b

[rsc.li/rsc-advances](https://rsc.li/rsc-advances)

## 1 Introduction

The energy consumption activities of industries, vehicles, and daily demands intensively depend on fossil fuels and natural gases. The by-products of such processes contain harmful emissions, especially CO<sub>2</sub>, a greenhouse gas related to the global warming phenomenon.<sup>1</sup> Also, there is a huge amount of CO<sub>2</sub> emitted by agricultural activities and forest exploitation and reclamation.<sup>2</sup> For sustainable environment and energy usage, CO<sub>2</sub> capture and reduction are indispensable. Until now, post-combustion (mainly used in coal-gasification plants), pre-combustion, and oxyfuel combustion (primarily used in gas and coal-fired plants) have been crucial technologies for carbon capture.<sup>3</sup> However, these technologies face many challenges

such as CO<sub>2</sub> separation and purification from flue gas streams and gas mixtures.<sup>4</sup> The solution to these problems is searching for solid materials that meet the requirements,<sup>5</sup> *i.e.*, high adsorption capacity, high selectivity, fast adsorption and desorption, and durability for CO<sub>2</sub> capture based on adsorption effects.<sup>6–8</sup> Solid sorbents that have been used to adsorb CO<sub>2</sub> include rigid classical adsorbents (activated carbon and zeolites) and novel porous adsorbents (covalent organic frameworks, zeolitic imidazolate frameworks, and metal–organic frameworks). Nowadays, metal–organic frameworks (MOFs)<sup>9,10</sup> have become the most popular CO<sub>2</sub> capture material due to their many outstanding features: ultra-large surface area, high porosity, and remarkable chemical, thermal, and mechanical stabilities.<sup>11,12</sup>

So far, MOF-200 and MOF-210 recorded the highest capacity for the gravimetric CO<sub>2</sub> uptake of 2400 mg g<sup>−1</sup> (74.2 wt%) and 2396 mg g<sup>−1</sup> (73.9 wt%) at 298 K and 50 bar.<sup>13</sup> At the same condition of the temperature and pressure, MOF-117 was the best sorbent among the considered MOFs with the gravimetric and volumetric CO<sub>2</sub> uptakes of 33.5 mmol g<sup>−1</sup> and 320 cm<sup>3</sup> (STP) per cm<sup>3</sup>, respectively.<sup>14</sup> Besides, MIL-100(Cr) and MIL-101c(Cr) also achieved 18 mmol g<sup>−1</sup> and 40 mmol g<sup>−1</sup> (280 cm<sup>3</sup> (STP) per cm<sup>3</sup> and 390 cm<sup>3</sup> (STP) per cm<sup>3</sup> for volumetric loading) at 304 K and 50 bar; MIL-47(V) with 11 mmol g<sup>−1</sup> and 250 cm<sup>3</sup> (STP) per cm<sup>3</sup> at 302 K, 20 bar, IRMOF-1 with 21.7 mmol g<sup>−1</sup> and 290 cm<sup>3</sup> (STP) per cm<sup>3</sup> at 298 K, 35 bar.<sup>15</sup>

<sup>a</sup>Faculty of Natural Sciences, Quy Nhon University, 170 An Duong Vuong, Quy Nhon City, Binh Dinh Province, Vietnam

<sup>b</sup>Ho Chi Minh City University of Technology (HCMUT), 268 Ly Thuong Kiet Street, District 10, Ho Chi Minh City, Vietnam. E-mail: dnson@hcmut.edu.vn

<sup>c</sup>Vietnam National University Ho Chi Minh City, Linh Trung Ward, Ho Chi Minh City, Vietnam

<sup>d</sup>Department of Chemistry, University of Science, Ho Chi Minh City, Vietnam

<sup>e</sup>Institute of Physical Chemistry “Ilie Murgulescu” of the Romanian Academy, Splaiul Independentei 202, Sector 6, 060021 Bucharest, Romania

 † Electronic supplementary information (ESI) available: Theoretical investigation of CO<sub>2</sub> capture in the MIL-88 series: effects of organic linker modification. See DOI: <https://doi.org/10.1039/d3ra01588b>


Other MOFs were also investigated for CO<sub>2</sub> capture as NU-100 (69.8 wt%, at 298 K and 40 bar), Mg-MOF-74 (68.9 wt%, at 278 K and 36 bar), MOF-5 (58 wt%, at 273 K and 10 bar), and HKUST-1 (19.8 wt%, 298 K and 1 bar).<sup>16</sup>

Among the successfully synthesized MOFs, the MIL-88 series have attracted our attention due to their exceptionally high flexibility and stability that MIL-88 series could swell upon immersion in various liquids with reversible variations in unit cell volume from 85% (MIL-88A) to 240% (MIL-88D) depending on the nature and length of the organic spacer without breaking the bonds, and fully retains its open framework topology.<sup>17</sup> Therefore, the MIL-88 series became a good candidate for CO<sub>2</sub> capture. In reality, MIL-88A(Fe) has been studied for catalysis,<sup>18,19</sup> NO adsorption,<sup>20</sup> and drug transportation.<sup>21</sup> Besides, MIL-88s (s = A, B, C, D) were investigated for hydrogen storage.<sup>22–24</sup> In 2016, the experiment measured CO<sub>2</sub> sorption capacity of MIL-88A as a function of temperatures, and the value achieved was 4.95 mmol g<sup>-1</sup> at 30 °C and the pressure of 1 bar.<sup>25</sup> MIL-88 family has been synthesized by experiment,<sup>26,27</sup> where the structure of MIL-88A, B, C, and D has the same metal site but different organic linkers, *i.e.*, fumarate, 1,4-benzene-dicarboxylate, 2,6-naphthalene-dicarboxylate, and 4,4'-biphenyl-dicarboxylate, respectively. However, no research is available to systematically investigate the CO<sub>2</sub> capture capacity in the MIL-88 series and the physical meanings of the CO<sub>2</sub>-MOF interaction under the influences of replacing organic linkers. Therefore, we elucidated these topics in the present work using the van der Waals dispersive correction density functional theory calculations for the earlier *via* electronic structure properties and the grand canonical Monte Carlo simulations for the later *via* adsorption isotherms, and isosteric heat of adsorption for the pressures below 100 bar. Elucidating the effects of defects and structural flexibility is out of the scope of the current work.

## 2 Computational techniques

### 2.1. Density functional theory calculations

We used the van der Waals dispersive correction density functional theory (vdW-DF) calculations with the revPBE-based exchange–correlation energy<sup>28,29</sup> *via* the Vienna Ab initio Simulation Package (VASP).<sup>30,31</sup> The projector augmented-wave technique was used to describe the electron–ion interactions and a cut-off energy of 700 eV for the plane-wave basis set.<sup>32,33</sup> The first Brillouin zone was sampled with a special *k*-point mesh of 3 × 3 × 3 for MIL-88A and B, 3 × 3 × 2 for MIL-88C and D, centred at the Gamma point following the Monkhorst–Pack method. The Methfessel–Paxton smearing of order 1 with the smearing width of 0.1 eV was used to support the convergence speed of geometry relaxation and the electronic density of states (DOS).<sup>34</sup> Besides, the tetrahedron method with the Bloch corrections has been selected for the total energy calculations.<sup>35</sup> Because the magnetic moments of the iron atom of the MIL-88 series are inconsistent with the spin-polarized calculations due to the structural flexibility, we, therefore, performed the non-spin polarized scheme in the present study. The structure and unit cell of the MIL-88 series are shown in Fig. 1. MIL-88A, B, C, and D share the same metal oxide node (Fe<sub>3</sub>O) but different

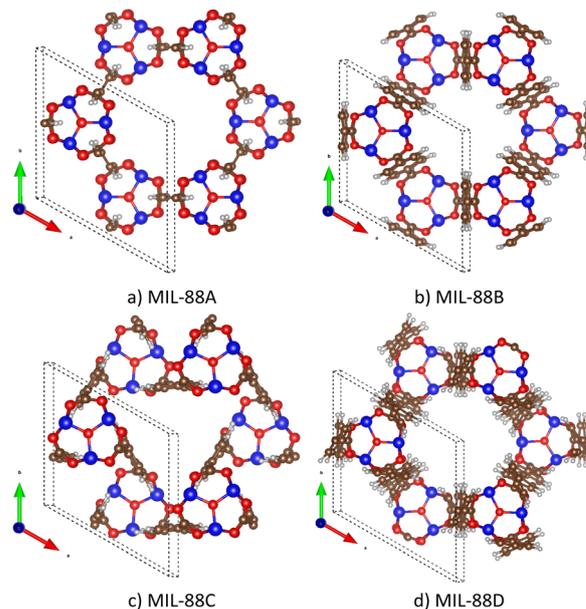


Fig. 1 The structure of MIL-88 series. The dotted frame indicates the unit cell. Fe (blue), O (red), and C (brown), and H (light-grey).

ligands, *i.e.*, fumarate or FMA (OOC–C<sub>2</sub>H<sub>2</sub>–COO), 1,4-benzene-dicarboxylate or BDC (OOC–C<sub>6</sub>H<sub>4</sub>–COO), 2,6-naphthalene-dicarboxylate or NDC (OOC–C<sub>10</sub>H<sub>6</sub>–COO), and 4,4'-biphenyl-dicarboxylate or BPDC (OOC–C<sub>12</sub>H<sub>8</sub>–COO), respectively.

Before exploring the CO<sub>2</sub> adsorption configurations, we have to fully optimize the atomic positions and the unit cell volume of the MIL-88 series, which followed the procedure of our previous publications.<sup>22,23</sup> We then carried out the geometric optimization to search for the CO<sub>2</sub> adsorption configurations and favorable sites. The adsorption energy of the CO<sub>2</sub> molecule in the MIL-88 series was calculated by

$$\Delta E_{\text{ads}} = E_{\text{MIL}+\text{CO}_2} - (E_{\text{MIL}} + E_{\text{CO}_2}), \quad (1)$$

where  $E_{\text{MIL}+\text{CO}_2}$ ,  $E_{\text{MIL}}$ , and  $E_{\text{CO}_2}$  are the total energy of the [MIL-88 + CO<sub>2</sub>] system, the pristine MIL-88, and the isolated CO<sub>2</sub> molecule, respectively. Besides, we calculated the charge density difference of the [MIL-88 + CO<sub>2</sub>] system *via* the expression

$$\Delta\rho = \rho_{\text{MIL}+\text{CO}_2} - (\rho_{\text{MIL}} + \rho_{\text{CO}_2}), \quad (2)$$

here  $\rho_{\text{MIL}+\text{CO}_2}$ ,  $\rho_{\text{MIL}}$ , and  $\rho_{\text{CO}_2}$  are the charge density of the [MIL-88 + CO<sub>2</sub>] system, the pristine MIL-88, and the isolated CO<sub>2</sub> molecule, respectively.

### 2.2. Grand canonical Monte Carlo simulations

After optimizing the volume and structure of the MIL-88 series, we obtained the charge distribution in the unit cell. We then used the density-derived electrostatic and chemical charge (DDEC6) method<sup>36–39</sup> to calculate the partial point charge for the atoms of the MIL-88 series. This charge assignment was used to parameterize the force fields as follows.



The interaction between the CO<sub>2</sub> molecule and the MIL-88 series was modelled by

$$U(r_{ij}) = U_{\text{vdW}}(r_{ij}) + U_{\text{elec}}(r_{ij}), \quad (3)$$

where  $r_{ij}$  is the distance between the  $i^{\text{th}}$  and  $j^{\text{th}}$  atoms. The above interaction consists of two parts:

(i) The van der Waals (vdW) interaction was described by the Lennard-Jones (LJ) potential,

$$U_{\text{vdW}}(r_{ij}) = 4\epsilon_{ij} \left[ \left( \frac{\sigma_{ij}}{r_{ij}} \right)^{12} - \left( \frac{\sigma_{ij}}{r_{ij}} \right)^6 \right], \quad (4)$$

where, the parameters  $\epsilon_{ij}$  and  $\sigma_{ij}$  are the Lennard-Jones (LJ) potential well depth and diameter, respectively. These parameters were determined by the Lorentz–Berthelot mixing rule for a pair of unlike atoms:

$$\epsilon_{ij} = \sqrt{\epsilon_i \epsilon_j} \quad \text{and} \quad \sigma_{ij} = \frac{\sigma_i + \sigma_j}{2}, \quad (5)$$

here,  $\sigma_i$  and  $\epsilon_i$  ( $i$  = the Fe, H, C, O atoms of MIL-88A, B, C, and D) were taken from the generic force fields for MOFs.<sup>40</sup> Notice that the cutoff radius of 22 Å has been checked for the convergence of the vdW interaction.

(ii) The electrostatic interaction with a cutoff radius of 13 Å,

$$U_{\text{elec}}(r_{ij}) = \frac{1}{4\pi\epsilon_0} \frac{q_i q_j}{r_{ij}}. \quad (6)$$

Here,  $\epsilon_0$  is the dielectric constant of the vacuum, and  $q_i$  is the partial charge of the  $i^{\text{th}}$  atom, which was obtained by the DDEC6 method.<sup>36–39</sup> The atomic charge of the atoms in the MIL-88 series was listed in Table 1.<sup>23</sup> Where, Fe, H, C1, C2, C3, C4, C5, O1, and O2 were also displayed in Fig. 2. The force field parameters for the CO<sub>2</sub> molecule were described as a rigid three-site molecule using the EPM2 model (elementary physical model).<sup>41,42</sup>

With the obtained force fields, we carried out the GCMC simulations using the RASPA code<sup>40</sup> to calculate the CO<sub>2</sub> adsorption isotherms in the MIL-88 series with  $\mu$ VT ensembles (*i.e.*, fixed the temperature, volume, and chemical potential) for various pressures up to 60 bar and 298 K (room temperature). In

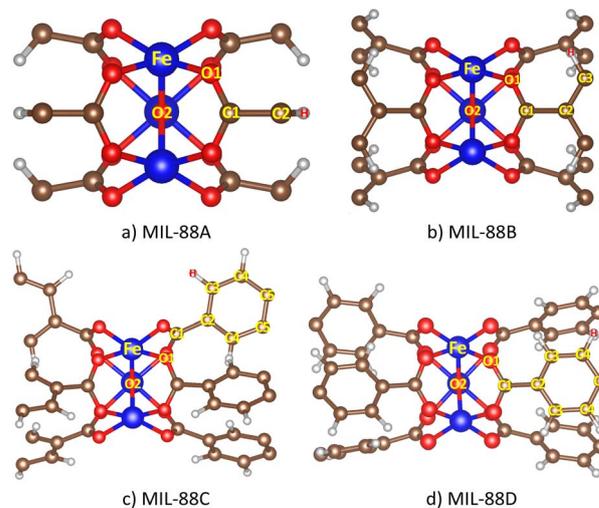


Fig. 2 The structure of MIL-88 series. Fe (blue), O (red), and C (brown), H (light-grey).

the GCMC simulation process, the MIL-88 series was kept fixed while the CO<sub>2</sub> molecule could move by inserting, translating, rotating, and deleting operations. We obtained the total ( $n_{\text{tot}}$ ) and excess ( $n_{\text{exc}}$ ) uptakes of CO<sub>2</sub>, satisfying the relationship

$$n_{\text{exc}} = n_{\text{tot}} - V_{\text{g}} \rho_{\text{g}}. \quad (7)$$

Here,  $V_{\text{g}}$  and  $\rho_{\text{g}}$  are the pore volume of the MIL-88 and the molar density of the bulk gas phase, respectively.

To evaluate the strength of the adsorbate–adsorbent interaction within the GCMC simulations, the total heat of adsorption (or the isosteric heat of adsorption) was also calculated from the average energies of the [MIL-88 + CO<sub>2</sub>] system:

$$Q_{\text{st}} = [\langle U_{\text{gh}} \rangle - \langle U_{\text{h}} \rangle - \langle U_{\text{g}} \rangle] - RT. \quad (8)$$

where  $\langle U_{\text{gh}} \rangle$ ,  $\langle U_{\text{h}} \rangle$ , and  $\langle U_{\text{g}} \rangle$  are the average potential energy of the guest–host (CO<sub>2</sub> + MIL-88), the host (MIL-88), and the gas (CO<sub>2</sub>), respectively.

Table 1 The Lennard-Jones parameters, the atomic charges of the CO<sub>2</sub> molecule and the MIL-88 series<sup>a</sup>

Ions	$\epsilon/k_{\text{B}}$ (K)	$\sigma$ (Å)	Partial charges (e <sup>-</sup> )
Fe_MIL	6.54	2.59	1.222 (A), 1.212 (B), 1.208 (C), 1.306 (D)
H_MIL	7.65	2.85	0.118 (A), 0.099 (B), 0.096 (C, D)
C1_MIL	47.86	3.47	0.734 (A), 0.689 (B), 0.700 (C), 0.718 (D)
C2_MIL			-0.178 (A), -0.095 (B), -0.071 (C), -0.135 (D)
C3_MIL			-0.059 (B), -0.062 (C), -0.026 (D)
C4_MIL			-0.114 (C), -0.139 (D)
C5_MIL			0.058 (C), 0.073 (D)
O1_MIL	48.16	3.03	-0.570 (A, B), -0.574 (C), -0.603 (D)
O2_MIL ( $\mu_3$ -O)			-0.875 (A), -0.844 (B), -0.849 (C), -0.939 (D)
C_CO <sub>2</sub> (ref. 41 and 42)	28.129	2.757	0.6512
O_CO <sub>2</sub> (ref. 41 and 42)	80.507	3.033	-0.3256

<sup>a</sup>  $\mu_3$ -O: the oxygen atom at the center of the trimer of the MIL-88 series. The A, B, C, and D letters indicate MIL-88A, MIL-88B, MIL-88C, and MIL-88D, respectively.



### 3 Results and discussion

#### 3.1. Adsorption configurations and energies

After optimizing the atomic positions of MIL-88s using the vdW-DF calculations, we loaded a CO<sub>2</sub> molecule into the MIL-88s structure to create the CO<sub>2</sub>@MIL-88s systems, then continued fully relaxing their atomic positions. With many possibilities of the CO<sub>2</sub> adsorption sites on the Fe metal, the organic linker, and iron–oxygen hollow sites, see Fig. 3, we obtained the stable configurations of the CO<sub>2</sub> molecule and the corresponding adsorption energies. With three stable adsorption sites and two configurations (end-on and side-on) for each MOF, see Fig. 4 and Table 2, we found that there is a similarity among MIL-88A, B, and D that the side-on configuration is more stable than the end-on configuration on the hollow and linker, while the reverse tendency holds for the CO<sub>2</sub> molecule on the Fe metal. MIL-88C does not adsorb the CO<sub>2</sub> molecule on the linker. Even though the most favourable adsorption configuration of CO<sub>2</sub> is end-on on the Fe metal for MIL-88A, B, and C, while it is side-on on the linker for MIL-88D. We also observed that the CO<sub>2</sub> molecule adsorbed in the pore position among the metal site and the iron–oxygen hollow (see Fig. S2 in the ESI†). Therefore, the CO<sub>2</sub> molecule belongs to both the metal end-on and hollow end-on or metal side-on and hollow side-on configurations in MIL-88C. The bond distance from the CO<sub>2</sub> molecule to the nearest atoms of the MIL-88s is about 3 Å, Table 2. Besides, we found that the average adsorption strength of CO<sub>2</sub> at the favourable sites is in the order of MIL-88A (−37.52 kJ mol<sup>−1</sup>) > MIL-88D (−19.33 kJ mol<sup>−1</sup>) > MIL-88B (−14.25 kJ mol<sup>−1</sup>) > MIL-88C (−2.9 kJ mol<sup>−1</sup>). Notably, the adsorption of the CO<sub>2</sub> molecule is strongest and weakest on MIL-88A and C, respectively. The number of the favourable adsorption sites of MIL-88C is fewer than that of the others because its organic linker does not offer negative adsorption energy. A more detailed analysis revealed that the adsorption strength is MIL-88A (FMA linker) > MIL-88B (BDC linker) > MIL-88D (BPDC linker) > MIL-88C (NDC linker) for the metal, but it is MIL-88A > MIL-88D > MIL-88B > MIL-88C for the linker and hollow sites. The latter has the same

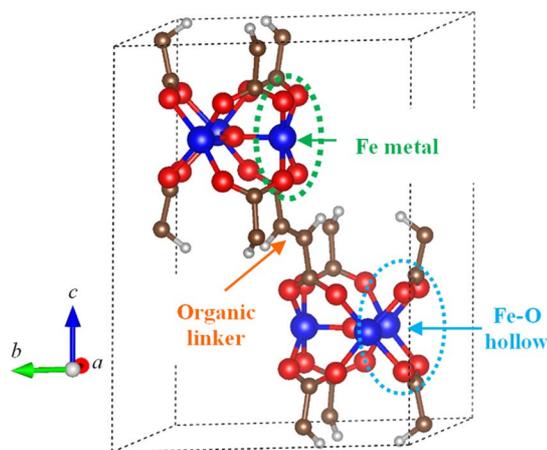


Fig. 3 The possible adsorption regions in a unit cell of MIL-88 series. Fe (blue), O (red), C (brown), and H (light-grey).

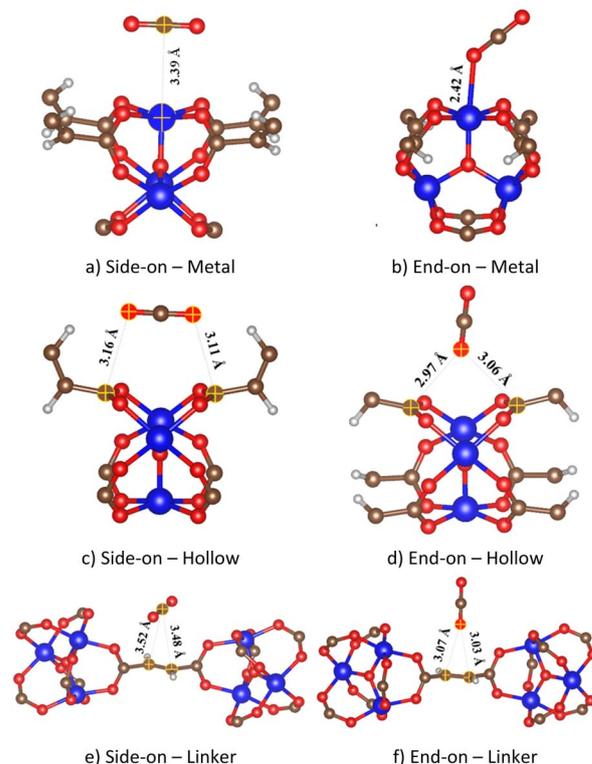


Fig. 4 The favorable CO<sub>2</sub> adsorption configurations in MIL-88A. Similarly, the CO<sub>2</sub> adsorption configuration in MIL-88B, C, and D was presented in the ESI (Fig. S1–S3).†

tendency as the average adsorption energy. Also, the adsorption energy is significantly large on the linkers compared to the metal and hollow sites of MIL-88A, B, and D, which implies a significant role of the linkers.

#### 3.2. Electronic properties

The nature of the CO<sub>2</sub>@MIL-88s interaction can be understood *via* the analysis of point charges, which were calculated by the Bader partition technique<sup>43,44</sup> and then excluded the neutral charge value of the corresponding atoms.<sup>45,46</sup> Table 3 shows that the C and O atoms of the CO<sub>2</sub> molecule always donate and accumulate the negative charge for all cases, respectively. However, the CO<sub>2</sub> molecule can gain (in most cases) or lose the charge. For MIL-88s, the Fe and C atoms always lose, while the O atoms always gain the negative charge. The charge of the H atoms of MIL-88s can vary depending on case by case, resulting in the total charge of MIL-88s, which can compensate for the charge losses or gains of the CO<sub>2</sub> molecule. Based on the obtained charge exchange of the CO<sub>2</sub> molecule and the MIL-88s, we found that the maximum value of the charge exchange of 11% e<sup>−</sup> was achieved for the CO<sub>2</sub> adsorption with the side-on configuration at the Fe metal site in the MIL-88D. Therefore, the nature of the CO<sub>2</sub>@MIL-88s interaction is the weak charge exchange.

The charge density difference in Fig. 5 shows that the charge accumulation and donation clouds of the CO<sub>2</sub> molecule have a relationship with that of the Fe and O atoms of the MIL-88



Table 2 Adsorption energy  $\Delta E_{\text{ads}}$  (kJ mol<sup>-1</sup>), the distance between the nearest atoms of CO<sub>2</sub> and MOF,  $d_{\text{MIL-CO}_2}$  (Å)

Site	Configuration	MIL-88A		MIL-88B		MIL-88C		MIL-88D	
		$\Delta E_{\text{ads}}$	$d_{\text{MIL-CO}_2}$	$\Delta E_{\text{ads}}$	$d_{\text{MIL-CO}_2}$	$\Delta E_{\text{ads}}$	$d_{\text{MIL-CO}_2}$	$\Delta E_{\text{ads}}$	$d_{\text{MIL-CO}_2}$
Metal	Side-on	-30.65	3.39	-0.46	3.37	-0.870 <sup>a</sup>	3.36	-12.73	3.42
	End-on	<b>-57.88</b>	2.42	<b>-37.23</b>	2.38	<b>-4.962<sup>b</sup></b>	2.33	-24.96	2.44
Hollow	Side-on	-28.59	3.16, 3.11	-6.87	3.08, 3.09	-0.870 <sup>a</sup>	3.58, 3.36	-20.84	3.17, 3.19
	End-on	-20.62	2.97, 3.06	-6.26	3.49, 3.67	<b>-4.962<sup>b</sup></b>	3.22, 3.17	-11.01	2.99, 3.01
Linker	Side-on	-45.16	3.48, 3.52	-15.55	3.57, 3.58	—	—	<b>-26.88</b>	3.39, 3.45
	End-on	-42.23	3.03, 3.07	-5.36	3.44, 3.57	—	—	-19.57	3.27, 3.27
Average adsorption energy		-37.52		-11.96		-2.916		-19.33	

<sup>a</sup> These adsorption configurations are identical. <sup>b</sup> These adsorption configurations are identical.

Table 3 The Bader charge transfer (e<sup>-</sup>) between the CO<sub>2</sub> molecule and the MIL-88 series. Positive and negative values implied the negative charge (e<sup>-</sup>) accumulation and donation, respectively

Site/configuration		Metal		Hollow		Linker		
		Side-on	End-on	Side-on	End-on	Side-on	End-on	
CO <sub>2</sub> @MIL-88A	1C	-1.949	-2.071	-1.995	-2.013	-2.074	-2.057	
	2O	1.954	2.097	2.059	2.065	2.100	2.068	
	CO <sub>2</sub>	<b>0.006</b>	<b>0.025</b>	<b>0.065</b>	<b>0.052</b>	<b>0.026</b>	<b>0.011</b>	
	12H	0.132	0.068	-0.013	0.073	0.146	0.129	
	24C	-19.016	-18.923	-18.892	-18.996	-19.069	-19.029	
	26O	27.221	27.239	27.360	27.269	27.340	27.315	
	6Fe	-8.343	-8.409	-8.382	-8.398	-8.443	-8.426	
	<b>MIL-88A</b>	<b>-0.006</b>	<b>-0.025</b>	<b>-0.065</b>	<b>-0.052</b>	<b>-0.026</b>	<b>-0.011</b>	
	CO <sub>2</sub> @MIL-88B	1C	-2.072	-2.064	-1.966	-1.983	-2.055	-2.070
		2O	2.080	2.053	1.973	1.991	2.085	2.098
CO <sub>2</sub>		<b>0.008</b>	<b>-0.011</b>	<b>0.007</b>	<b>0.008</b>	<b>0.030</b>	<b>0.028</b>	
24H		-0.186	0.093	-0.101	-0.110	-0.172	-0.239	
48C		-18.771	-18.945	-18.910	-18.894	-18.727	-18.785	
26O		27.356	27.296	27.424	27.369	27.300	27.431	
6Fe		-8.407	-8.434	-8.420	-8.374	-8.431	-8.436	
<b>MIL-88B</b>		<b>-0.008</b>	<b>0.011</b>	<b>-0.007</b>	<b>-0.008</b>	<b>-0.030</b>	<b>-0.028</b>	
CO <sub>2</sub> @MIL-88C		1C	-2.021			-2.090		
		2O	2.039			2.069		
	CO <sub>2</sub>	<b>0.018</b>			<b>-0.021</b>			
	36H	0.316			0.263			
	72C	-19.275			-19.286			
	26O	27.166			27.288			
	6Fe	-8.225			-8.245			
	<b>MIL-88C</b>	<b>-0.018</b>			<b>0.021</b>			
	CO <sub>2</sub> @MIL-88D	1C	-1.935	-2.029	-1.943	-1.999	-1.974	-2.050
		2O	2.045	2.014	1.951	2.006	1.994	2.076
CO <sub>2</sub>		<b>0.110</b>	<b>-0.016</b>	<b>0.008</b>	<b>0.007</b>	<b>0.020</b>	<b>0.026</b>	
48H		0.256	0.249	0.207	0.228	0.117	0.220	
85C		-19.288	-19.248	-19.331	-19.385	-19.160	-19.248	
26O		27.340	27.427	27.556	27.549	27.440	27.416	
6Fe		-8.418	-8.413	-8.440	-8.399	-8.417	-8.414	
<b>MIL-88D</b>		<b>-0.110</b>	<b>0.016</b>	<b>-0.008</b>	<b>-0.007</b>	<b>-0.020</b>	<b>-0.026</b>	

series, while the C and H atoms of the MIL-88 series exhibit an ignorable role in the interaction with the CO<sub>2</sub> molecule as no significant charge clouds have been found on these atoms. Besides, we explored the nature of the CO<sub>2</sub>@MIL-88s interaction at the electronic orbital level through the analysis of the orbital-projected density of states (PDOS), which was presented in Fig. 6. This figure revealed that the  $1\pi_g$ ,  $2\sigma_u/1\pi_u$ , and  $2\sigma_g$

peaks of the CO<sub>2</sub> molecule locate near -4.5, -8.0, and -9.5 eV, respectively.<sup>47</sup> The  $p_x$ ,  $p_y$ ,  $p_z$ , and s orbitals of the MIL-88 series distribute below -1.0 eV, while the d orbitals of the MIL-88s position around the Fermi level. Generally, the interaction between adsorbate and adsorbent stems from the attraction between occupied states of the isolated adsorbate with unoccupied states of the isolated adsorbent and *vice versa*. When the



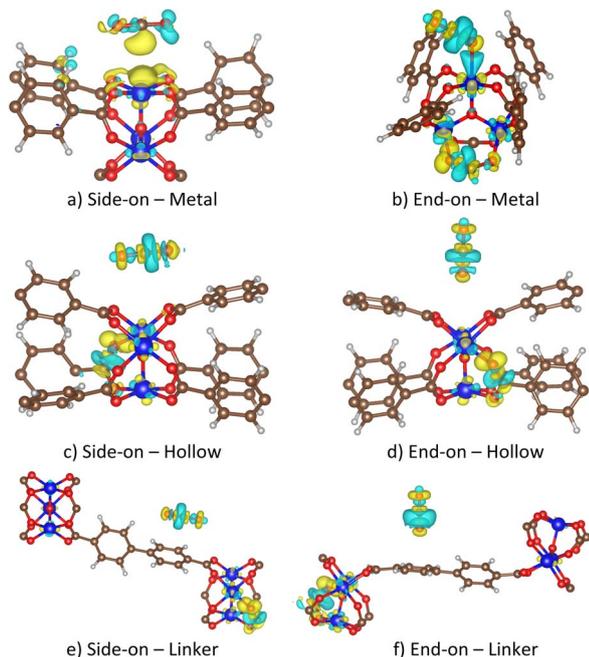


Fig. 5 The charge density difference of the  $\text{CO}_2$ @MIL-88D system with the different adsorption configurations and sites. Negative charge accumulation (yellow) and donation (cyan). Similar features of the charge clouds have been shown for MIL-88A, B, and C in Fig. S4–S6 in the ESI.† Isosurface values ( $e^- \text{ bohr}^{-3}$ ) for the charge density difference of the MIL-88D@ $\text{CO}_2$  system at different sites are listed in Table S1.†

combined system is formed, the electronic states of the adsorbate and the adsorbent become completely overlapping to make the peak resonance. Therefore, the overlapping between the electronic states of the adsorbate and adsorbent becomes an indication of their interaction. We found in Fig. 6 that the peak resonance forms between the DOS peaks of the  $\text{CO}_2$  molecule with mainly the C and O  $p_x$ ,  $p_y$ , and  $p_z$  orbitals and weakly the Fe  $s$  and  $d_{z^2}$  orbitals of the MIL-88 series for the energy below  $-1.0$  eV. This observation and the charge density difference implied that the C and O  $p_x$ ,  $p_y$ , and  $p_z$  orbitals of the organic linkers should be the dominant factors contributing to the  $\text{CO}_2$ @MIL-88s interaction.

### 3.3. The $\text{CO}_2$ capture capacity of MIL-88 series

Using the force fields presented in the Computational techniques section, we calculated the  $\text{CO}_2$  capture gravimetric and volumetric capacities of the MIL-88 series *via* the isotherms as a function of pressures at the temperature of 298 K. Fig. 7a displays the obtained isotherms in gravimetric uptake, *i.e.*, the adsorbed amount per unit mass of adsorbent (wt%), and Fig. 7b is in volumetric uptakes, *i.e.*, the adsorbed amount per unit volume ( $\text{cm}^3$  (STP) per  $\text{cm}^3$ ). Besides, the number of  $\text{CO}_2$  molecules loading per unit cell of MIL-88A, MIL-88B, MIL-88C, and MIL-88D were also displayed in Fig. S8 in ESI.† We found that the total and excess isotherms rapidly increase at the low-pressure range and gradually enhance with the increase in pressure. Afterward, the excess isotherms approach their maximum value<sup>48,49</sup> and slowly decrease after passing by the

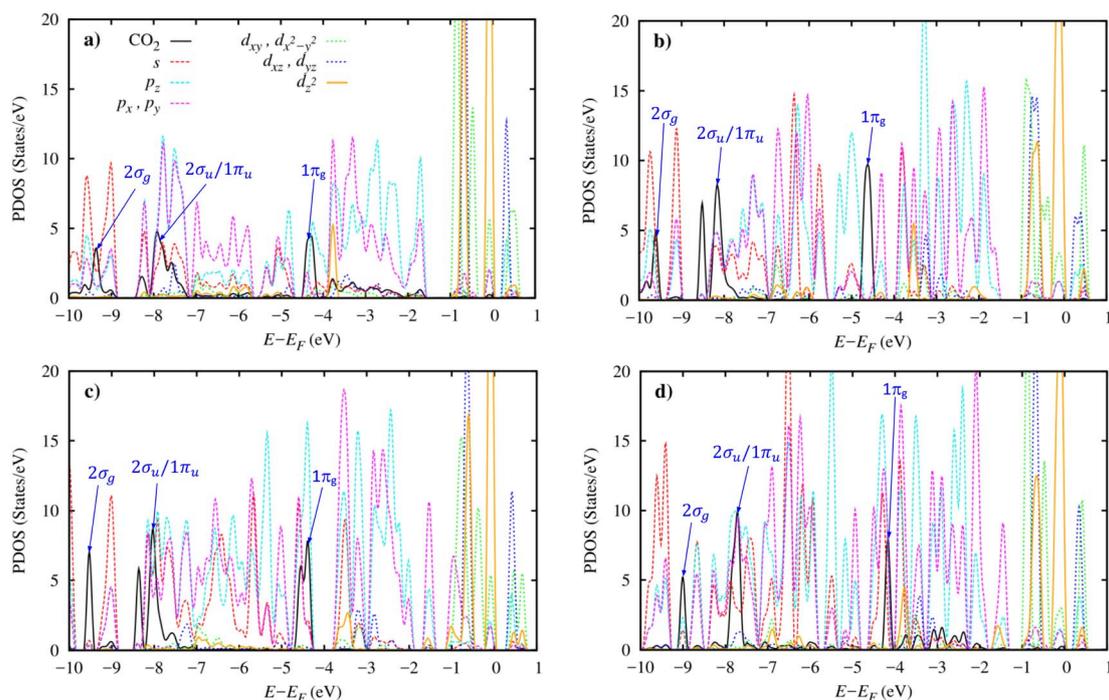


Fig. 6 The orbital-projected density of states of the  $\text{CO}_2$ @MIL-88 systems in the most favourable adsorption state: (a) end-on at the Fe metal of MIL-88A, (b) end-on at the Fe metal of MIL-88B, (c) end-on at the Fe metal of MIL-88C, and (d) side-on at the organic linker of MIL-88D. The Fermi level was set to 0 eV.



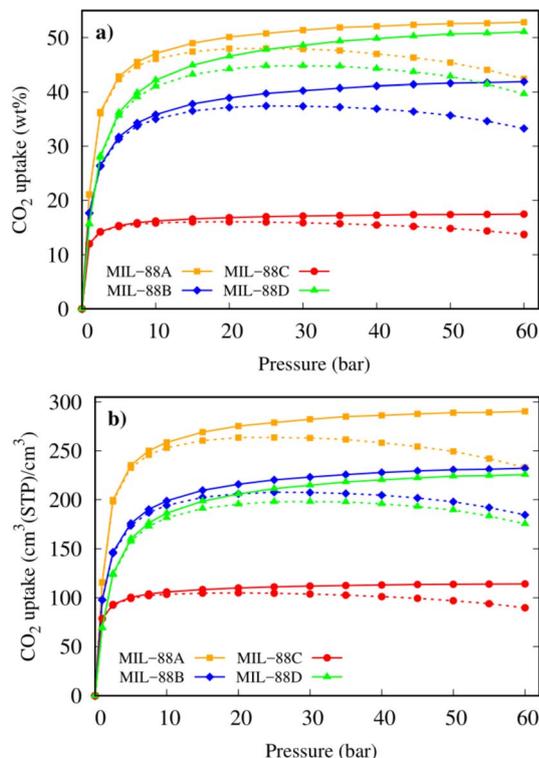


Fig. 7 The excess (dashed line) and total (solid line) CO<sub>2</sub> adsorption isotherms of MIL-88A, B, C, and D at 298 K: (a) gravimetric capacity and (b) volumetric capacity.

Table 4 The maximum excess and absolute capture capacity of CO<sub>2</sub> in MIL-88A, B, C, and D at 298 K

Adsorbent	mmol g <sup>-1</sup>		wt%		cm <sup>3</sup> (STP) per cm <sup>3</sup>	
	Total	Excess	Total	Excess	Total	Excess
MIL-88A	12.11	10.90	53.26	47.98	292.65	263.61
MIL-88B	9.61	8.51	42.29	37.45	234.44	207.60
MIL-88C	4.00	3.64	17.58	16.04	115.11	104.98
MIL-88D	11.74	10.18	51.66	44.79	228.35	197.97

maximum while the total isotherms continue developing up to 60 bar. Table 4 lists the maximal CO<sub>2</sub> uptakes for gravimetric and volumetric capacities, where maximum excess loading was achieved at 25 bar for MIL-88A and B, 20 bar for MIL-88C, and 30 bar for MIL-88D. The highest value of total uptake was achieved at 60 bar for MIL-88A, B, and D, while it reached the earlier pressure value, about 20 bar for MIL-88C because the isotherms become flat for pressures greater than 20 bar. Table 4 shows that both excess and total isotherms are in the order MIL-88A (FMA linker) > MIL-88D (BPDC linker) > MIL-88B (BDC linker) > MIL-88C (NDC linker) for the gravimetric CO<sub>2</sub> uptake (proportional to the adsorption energy tendency at the linker and hollow sites) and MIL-88A > MIL-88B > MIL-88D > MIL-88C (proportional to the adsorption energy tendency at the metal site) for the volumetric uptake. This finding implies that MIL-

88A and MIL-88C would be the best and worst candidates for CO<sub>2</sub> capture, respectively. Also, the tendency of the gravimetric uptakes is in good agreement with that of the average adsorption energy obtained above, *i.e.*, the stronger the adsorption energy, the higher the capture capacity.

Next, we calculated the isosteric heat of the CO<sub>2</sub> adsorption,  $Q_{st}$ , in MIL-88s, as shown in Fig. 8. The average value of the curves exhibited  $Q_{st}$  in the order: MIL-88A (34.9 kJ mol<sup>-1</sup>) > MIL-88B (32.5 kJ mol<sup>-1</sup>) > MIL-88D (32.1 kJ mol<sup>-1</sup>) > MIL-88C (31.8 kJ mol<sup>-1</sup>). This tendency is consistent with the volumetric uptakes and the adsorption energy at the metal site, as mentioned above. Moreover, the value of  $Q_{st}$  represents the CO<sub>2</sub>@MIL-88s interaction. The obtained values of  $Q_{st}$  are rather high compared to those of many other MOFs.<sup>50</sup>

To elucidate the effect of moisture on the CO<sub>2</sub> adsorption capacity in the MIL-88 series, we calculated both gravimetric and volumetric amounts of CO<sub>2</sub> adsorption in the presence of H<sub>2</sub>O. The results were described in Fig. S7 and Table S2 in ESI.† Compared to those without H<sub>2</sub>O, the CO<sub>2</sub> adsorption capacities in MIL-88s insignificantly decreased in the presence of H<sub>2</sub>O. The decrease was about 4–8% for the absolute uptakes at 40 bar and 2–6% for the maximum excess.

It is necessary to consider the influences of macroscopic parameters on the CO<sub>2</sub> capture capacity. Fig. 9 presents the dependence of the gravimetric uptakes of the CO<sub>2</sub> capture on the specific surface area and pore volume of the MIL-88 series. We found a simple relationship, *i.e.*, adsorption capacity proportionally depends on the specific surface area (Fig. 9a) and the pore volume (Fig. 9b). Here, the specific surface area and pore volume data of the MIL-88 series were calculated theoretically by grand canonical Monte Carlo simulations with nitrogen gas utilized to obtain them.

Substituting the organic linkers caused the effects on adsorption energy and uptakes, which were perhaps due to the following reasons: (1) geometric modification, *i.e.*, the organic linkers led to the different SSA and pore volume of MIL-88s. The larger SSA and pore volume lead to higher CO<sub>2</sub> capacity. (2) Electronic property modification *via* the overlapping of wavefunctions<sup>22</sup> between CO<sub>2</sub> and the organic linkers, even for the metal and hollow adsorption sites, see Fig. S9 in ESI.† We found that the wavefunctions of CO<sub>2</sub> and that of the C and O atoms of

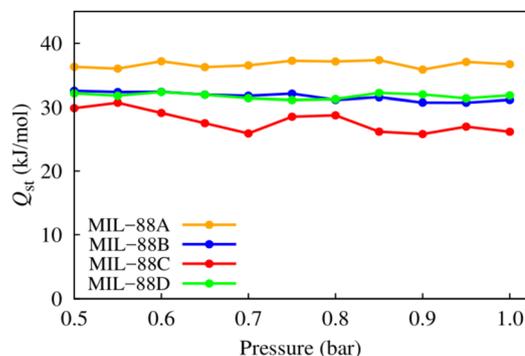


Fig. 8 The isosteric heat of the CO<sub>2</sub> adsorption in MIL-88 series at the low-pressure range and 298 K.



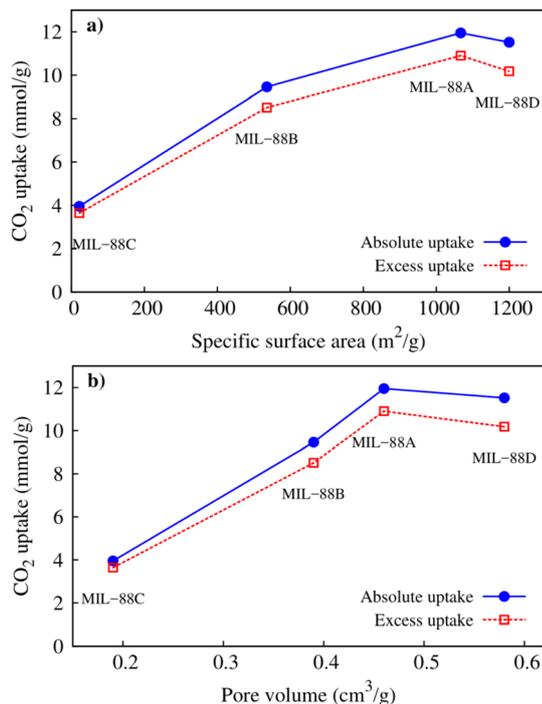


Fig. 9 The relationship between the  $\text{CO}_2$  uptakes with (a) specific surface area and (b) pore volume.

the organic linkers can overlap. The same or opposite direction oscillations of the wavefunctions cause constructive and destructive interferences, respectively. We can see that the wavefunction interaction of  $\text{CO}_2$  with the organic linker shows more clearly for MIL-88A and D than for MIL-88B and C, which somehow shows a relationship to the obtained  $\text{CO}_2$  uptakes.

Fig. 10 shows the excess gravimetric capacity of  $\text{CO}_2$  versus temperature at a low pressure of 1 bar. The uptake decreases as the temperature increases, following the straight line for MIL-88C; however, the curves for MIL-88A, B, and D. The isotherms approach a convergent value as the temperature is high enough from the room temperature. Particularly, MIL-88A achieved the excess uptake of  $4.00 \text{ mmol g}^{-1}$  at 303 K, which is

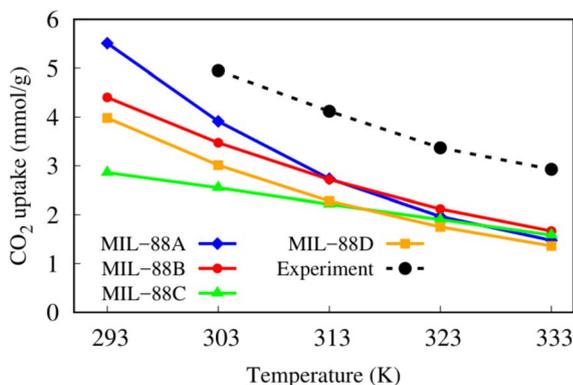


Fig. 10 The excess uptakes versus temperatures of MIL-88 series at 1 bar. The experimental data were obtained from ref. 25.

in good agreement with the experimental value of  $4.95 \text{ mmol g}^{-1}$  at  $30^\circ\text{C}$  (303 K).<sup>25</sup> Moreover, the shape of the  $\text{CO}_2$  adsorption curve on MIL-88A is consistent with that of the experiment in the temperature range from 293 to 333 K.<sup>25</sup>

## 4 Conclusions

This work investigated the  $\text{CO}_2$ @MIL-88s interaction and the  $\text{CO}_2$  capture capacity in the MIL-88 series using the vdW-DF density functional theory calculations and grand canonical Monte Carlo simulations. We found a proportional relationship between the adsorption energy with the gravimetric uptake FMA linker (MIL-88A) > BPDC linker (MIL-88D) > BDC linker (MIL-88B) > NDC linker (MIL-88C), proportional to the adsorption strength of  $\text{CO}_2$  at the linker and hollow sites and the isosteric heat of  $\text{CO}_2$  adsorption with the volumetric uptake FMA linker (MIL-88A) > BDC linker (MIL-88B) > BPDC linker (MIL-88D) > NDC linker (MIL-88C), proportional to the adsorption strength of  $\text{CO}_2$  at the metal site). We have to emphasize that the replacement of the organic linkers is the only difference in the geometric structure of MIL-88A, B, C, and D. Therefore, the substitution of the linkers also influences the adsorption strength of  $\text{CO}_2$  not only at the organic linkers but also at the metal and hollow sites. The gravimetric and volumetric uptakes indicated that MIL-88A (with fumarate linker) should be the best candidate among the MIL-88 series for  $\text{CO}_2$  capture. Furthermore, the physical insights have been elucidated that the contributions of the MIL-88 series to the interaction with  $\text{CO}_2$  are as follows: the C and O p orbitals (dominant) and the Fe s and  $d_{z^2}$  orbitals (minor). The H atoms of the MIL-88 series played an ignorable role in the  $\text{CO}_2$ @MIL-88 interaction.

## Author contributions

Conceptualization (DNS, VC), formal analysis (NTXH, DNS), investigation (NTXH, DNS), resources (DNS, VC), supervision (DNS), validation (NTXH, VC, DNS), visualization (NTXH, OKL, TPD), writing of original draft (DNS), reviewing and editing (VC, DNS).

## Conflicts of interest

There are no conflicts of interest to declare.

## Acknowledgements

This research is funded by Vietnam National University Ho Chi Minh City (VNU-HCM) under grant number VL2022-20-01.

## References

- 1 K. Sumida, D. L. Rogow, J. A. Mason, T. M. McDonald, E. D. Bloch, Z. R. Herm, T. H. Bae and J. R. Long, A short review of recent advances in  $\text{CO}_2$  hydrogenation to hydrocarbons over heterogeneous catalysts, *Chem. Rev.*, 2012, **112**, 724–781.



- 2 R. Quadrelli and S. Peterson, The energy-climate challenge: recent trends in CO<sub>2</sub> emissions from fuel combustion, *Energy Policy*, 2007, **35**, 5938–5952.
- 3 S. E. M. Elhenawy, M. Khraisheh, F. Almomani and G. Walker, Metal-organic frameworks as a platform for CO<sub>2</sub> capture and chemical processes: adsorption, membrane separation, catalytic-conversion, and electrochemical reduction of CO<sub>2</sub>, *Catalysts*, 2020, **10**, 1–33.
- 4 M. Kanniche, R. Gros-Bonnivard, P. Jaud, J. Valle-Marcos, J. M. Amann and C. Bouallou, Pre-combustion, post-combustion and oxy-combustion in thermal power plant for CO<sub>2</sub> capture, *Appl. Therm. Eng.*, 2010, **30**, 53–62.
- 5 A. Rehman, S. Farrukh, A. Hussain and E. Pervaiz, Synthesis and effect of metal-organic frameworks on CO<sub>2</sub> adsorption capacity at various pressures: a contemplating review, *Energy Environ.*, 2020, **31**, 367–388.
- 6 D. Y. C. Leung, G. Caramanna and M. M. Maroto-valer, An overview of current status of carbon dioxide capture and storage technologies, *Renew. Sust. Energ. Rev.*, 2014, **39**, 426–443.
- 7 J. Y. Jung, F. Karadas, S. Zulfiqar, E. Deniz, S. Aparicio, M. Atilhan, C. T. Yavuz and S. M. Han, Limitations and high pressure behavior of MOF-5 for CO<sub>2</sub> capture, *Phys. Chem. Chem. Phys.*, 2013, **15**, 14319–14327.
- 8 W. S. Drisdell, R. Poloni, T. M. McDonald, T. A. Pascal, L. F. Wan, C. Das Pemmaraju, B. Vlasisavljevich, S. O. Odoh, J. B. Neaton, J. R. Long, D. Prendergast and J. B. Kortright, Probing the mechanism of CO<sub>2</sub> capture in diamine-appended metal-organic frameworks using measured and simulated X-ray spectroscopy, *Phys. Chem. Chem. Phys.*, 2015, **17**, 21448–21457.
- 9 Y. Li, H. Xu, S. Ouyang and J. Ye, Metal-organic frameworks for photocatalysis, *Phys. Chem. Chem. Phys.*, 2016, **18**, 7563–7572.
- 10 S. Kang and J. Yu, Electronic structure and magnetic properties of transition metal Kagome metal-organic frameworks, *Phys. Chem. Chem. Phys.*, 2022, **24**, 22168–22180.
- 11 A. J. Howarth, Y. Liu, P. Li, Z. Li, T. C. Wang, J. T. Hupp and O. K. Farha, Chemical, thermal and mechanical stabilities of metal-organic frameworks, *Nat. Rev. Mater.*, 2016, **1**, 1–15.
- 12 T. T. T. Huong, P. N. Thanh, N. T. X. Huynh and D. N. Son, Metal-organic frameworks: state-of-the-art material for gas capture and storage, *VNU J. Sci. Math. –Phys.*, 2016, **32**, 67–85.
- 13 H. Furukawa, N. Ko, Y. B. Go, N. Aratani, S. B. Choi, E. Choi, A. O. Yazaydin, R. Q. Snurr, M. O’Keeffe, J. Kim and O. M. Yaghi, Ultrahigh porosity in metal-organic frameworks, *Science*, 2010, **329**, 424–428.
- 14 A. R. Millward and O. M. Yaghi, Metal-organic frameworks with exceptionally high capacity for storage of carbon dioxide at room temperature, *J. Am. Chem. Soc.*, 2005, **127**, 17998–17999.
- 15 P. L. Llewellyn, S. Bourrelly, C. Serre, A. Vimont, M. Daturi, L. Hamon, G. De Weireld, J. Chang, D. Hong, Y. K. Hwang and S. H. Jung, High uptakes of CO<sub>2</sub> and CH<sub>4</sub> in mesoporous metal-organic frameworks MIL-100 and MIL-101, *Langmuir*, 2008, **24**, 7245–7250.
- 16 C. Dey, *et al.*, Crystalline metal-organic frameworks (MOFs): synthesis, structure and function, *Acta Crystallogr., Sect. B: Struct. Sci., Cryst. Eng. Mater.*, 2014, **70**, 3–10.
- 17 P. Horcajada, F. Salles, S. Wuttke, T. Devic, D. Heurtaux, A. Vimont, M. Daturi, O. David, E. Magnier, N. Stock, Y. Filinchuk, D. Y. Popov, C. Riekkel, G. Férey, C. Serre, G. Maurin and D. Popov, How linker’s modification controls swelling properties of highly flexible iron (III) dicarboxylates MIL-88, *J. Am. Chem. Soc.*, 2011, **133**, 17839–17847.
- 18 J. Wang, J. Wan, Y. Ma, Y. Wang, M. Pu and Z. Guan, Metal-organic frameworks MIL-88A with suitable synthesis conditions and optimal dosage for effective catalytic degradation of Orange G through persulfate activation, *RSC Adv.*, 2016, **6**, 112502–112511.
- 19 W.-T. Xu, L. Ma, F. Ke, F.-M. Peng, G.-S. Xu, Y.-H. Shen, J.-F. Zhu, L.-G. Qiu and Y.-P. Yuan, Metal-organic frameworks MIL-88A hexagonal microrods as a new photocatalyst for efficient decolorization of methylene blue dye, *Dalton Trans.*, 2014, **43**, 3792–3798.
- 20 A. C. McKinlay, J. F. Eubank, S. Wuttke, P. S. Wheatley, P. Bazin, J.-C. Lavalley, M. Daturi, A. Vimont, G. De Weireld, P. Horcajada, C. Serre and R. E. Morris, Nitric oxide adsorption and delivery in flexible MIL-88(Fe) metal-organic frameworks, *Chem. Mater.*, 2013, **25**, 1592–1599.
- 21 S. N. Kim, C. G. Park, B. K. Huh, S. H. Lee, C. H. Min, Y. Y. Lee, Y. K. Kim, K. H. Park and Y. B. Choy, Metal-organic frameworks, NH<sub>2</sub>-MIL-88(Fe), as carriers for ophthalmic delivery of brimonidine, *Acta Biomater.*, 2018, **79**, 344–353.
- 22 N. T. X. Huynh, O. M. Na, V. Chihaiia and D. N. Son, A computational approach towards understanding hydrogen gas adsorption in Co-MIL-88A, *RSC Adv.*, 2017, **7**, 39583–39593.
- 23 N. T. X. Huynh, V. Chihaiia and D. N. Son, Hydrogen storage in MIL-88 series, *J. Mater. Sci.*, 2019, **54**, 3994–4010.
- 24 N. T. X. Huynh, V. Chihaiia and D. N. Son, Enhancing hydrogen storage by metal substitution in MIL-88A metal-organic framework, *Adsorption*, 2020, **26**, 509–519.
- 25 S. Wongsakulphasatch, W. Kiatkittipong, J. Saupsor, J. Chaiwisesphol, P. Piroonlerkgul, V. Parasuk and S. Assabumrungrat, Effect of Fe open metal site in metal-organic frameworks on post-combustion CO<sub>2</sub> capture performance, *Greenhouse Gases: Sci. Technol.*, 2016, **7**, 383–394.
- 26 C. Serre, F. Millange, S. Surblé and G. Férey, A route to the synthesis of trivalent transition-metal porous carboxylates with trimeric secondary building units, *Angew. Chem., Int. Ed.*, 2004, **43**, 6286–6289.
- 27 S. Surblé, C. Serre, C. Mellot-Draznieks, F. Millange and G. Férey, A new isoreticular class of metal-organic-frameworks with the MIL-88 topology, *Chem. Commun.*, 2006, 284–286.



- 28 J. P. Perdew, J. Chevary, S. Vosko, K. Jackson, M. Pederson, D. Singh and C. Fiolhais, Atoms, molecules, solids, and surfaces: applications of the generalized gradient approximation for exchange and correlation, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1992, **46**, 6671–6687.
- 29 J. P. Perdew, K. Burke and M. Ernzerhof, Generalized gradient approximation made simple, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 30 G. Kresse and J. Furthmüller, Efficient iterative schemes for ab initio total-energy calculations using a plane-wave basis set, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **54**, 11169–11186.
- 31 G. Kresse and J. Furthmüller, Efficiency of ab initio total energy calculations for metals and semiconductors using a plane-wave basis set, *Comput. Mater. Sci.*, 1996, **6**, 15–50.
- 32 P. E. Blöchl, Projector augmented-wave method, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **50**, 17953–17979.
- 33 G. Kresse and D. Joubert, From ultrasoft pseudopotentials to the projector augmented-wave method, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1999, **59**, 1758–1775.
- 34 M. Methfessel and A. T. Paxton, High-precision sampling for Brillouin-zone integration in metals, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1989, **40**, 3616–3621.
- 35 P. E. Blochl, O. Jepsen and O. K. Andersen, Improved tetrahedron method for Brillouin-zone integrations, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **49**, 16223–16233.
- 36 T. A. Manz and N. G. Limas, Introducing DDEC6 atomic population analysis: Part 1. Charge partitioning theory and methodology, *RSC Adv.*, 2016, **6**, 47771–47801.
- 37 N. G. Limas and T. A. Manz, Introducing DDEC6 atomic population analysis: Part 2. Computed results for a wide range of periodic and nonperiodic materials, *RSC Adv.*, 2016, **6**, 45727–45747.
- 38 T. A. Manz, Introducing DDEC6 atomic population analysis: part 3. Comprehensive method to compute bond orders, *RSC Adv.*, 2017, **7**, 45552–45581.
- 39 N. G. Limas and T. A. Manz, Introducing DDEC6 atomic population analysis: part 4. Efficient parallel computation of net atomic charges, atomic spin moments, bond orders, and more, *RSC Adv.*, 2018, **8**, 2678–2707.
- 40 D. Dubbeldam, S. Calero, D. E. Ellis and R. Q. Snurr, RASPA: molecular simulation software for adsorption and diffusion in flexible nanoporous materials, *Mol. Simul.*, 2016, **42**, 81–101.
- 41 J. G. Harris and K. H. Yungt, Carbon dioxide's liquid-vapor coexistence curve and critical properties as predicted by a simple molecular model, *J. Phys. Chem.*, 1995, **99**, 12021–12024.
- 42 K. Makrodimitris, G. K. Papadopoulos and D. N. Theodorou, Prediction of permeation properties of CO<sub>2</sub> and N<sub>2</sub> through silicalite via molecular simulations, *J. Phys. Chem. B*, 2001, **105**, 777–788.
- 43 G. Henkelman, A. Arnaldsson and H. Jónsson, A fast and robust algorithm for Bader decomposition of charge density, *Comput. Mater. Sci.*, 2006, **36**, 354–360.
- 44 W. Tang, E. Sanville and G. Henkelman, A grid-based Bader analysis algorithm without lattice bias, *J. Phys.: Condens. Matter*, 2009, **21**(7), 084204.
- 45 O. K. Le, V. Chihaiia, P. T. H. Hoa, P. T. Hai and D. N. Son, Physical insights into the Au growth on the surface of a LaAlO<sub>3</sub>/SrTiO<sub>3</sub> heterointerface, *RSC Adv.*, 2022, **12**, 24146–24155.
- 46 O. K. Le, V. Chihaiia, V. V. On and D. N. Son, N-type and p-type molecular doping on monolayer MoS<sub>2</sub>, *RSC Adv.*, 2021, **11**, 8033–8041.
- 47 D. N. Son, T. T. T. Huong and V. Chihaiia, Simultaneous adsorption of SO<sub>2</sub> and CO<sub>2</sub> in an Ni(bdc)(ted)<sub>0.5</sub> metal-organic framework, *RSC Adv.*, 2018, **8**, 38648–38655.
- 48 K. S. Walton, A. R. Millward, D. Dubbeldam, H. Frost, J. J. Low, O. M. Yaghi and R. Q. Snurr, Understanding inflections and steps in carbon dioxide adsorption isotherms in metal-organic frameworks, *J. Am. Chem. Soc.*, 2008, **130**, 406–407.
- 49 M. Gheytaizadeh, A. Baghban, S. Habibzadeh, A. Esmaili, O. Abida, A. Mohaddespour and M. T. Munir, Towards estimation of CO<sub>2</sub> adsorption on highly porous MOF-based adsorbents using Gaussian process regression approach, *Sci. Rep.*, 2021, **11**, 1–13.
- 50 Z. Li, P. Liu, C. Ou and X. Dong, Porous metal-organic frameworks for carbon dioxide adsorption and separation at low pressure, *ACS Sustainable Chem. Eng.*, 2020, **8**, 15378–15404.

