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Highly stable Fe/CeO₂ catalyst for the reverse water gas shift reaction in the presence of H₂S†

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This study focused on evaluating the catalytic properties for the reverse water gas shift reaction (RWGS: CO₂ + H₂ → CO + H₂O ΔH⁰ = 42.1 kJ mol⁻¹) in the presence of hydrogen sulfide (H₂S) over a Fe/CeO₂ catalyst, commercial Cu–Zn catalyst for the WGS reaction (MDC-7), and Co–Mo catalyst for hydrocarbon desulfurization. The Fe/CeO₂ catalyst exhibited a relatively high catalytic activity to RWGS, compared to the commercial MDC-7 and Co–Mo catalysts. In addition, the Fe/CeO₂ catalyst showed stable performance in the RWGS environment that contained high concentrations of H₂S. The role of co-feeding H₂S was investigated over the Fe/CeO₂ catalyst by the temperature programmed reaction (TPR) of CO₂ and H₂ in the presence of H₂S. The result of TPR indicated that the co-feeding H₂S might enhance RWGS performance due to H₂S acting as the hydrogen source to reduce CO₂.

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1. Introduction

Most of the world's natural gas resources are in the form of sour gas, which contains high concentrations of hydrogen sulfide (H₂S) and carbon dioxide (CO₂), with typical compositions of 0–30 vol% H₂S and 0–80 vol% CO₂.^{1–3} Sour gas is widely distributed worldwide, including in the Middle East, Canada, Northern Europe, and China. High concentrations of H₂S (over 30 vol%) is contained in large gas fields in Saudi Arabia, Abu Dhabi, and offshore areas in Iran.⁴ Moreover, biogas, a renewable energy source that has attracted much attention in recent years, contains a relatively high concentration of H₂S.^{5–7} In order to use the hydrocarbons in natural gas or methane in biogas as fuel, pollutants (H₂S and CO₂) that exist beyond allowable limits must be removed. For example, in natural gas, the heavy hydrocarbons are condensed and then passed through a scrubbing unit typically composed of liquid amine-based adsorbents to remove the sour gas components (H₂S and CO₂), in a process called “sweetening”.⁸ However, the liquid amine sorbents

cannot control the H₂S/CO₂ selectivity because the concentration of these contaminants varies depending on the source of the natural gas.^{9,10} In this regard, acid gas remediation is costly, and it is often uneconomical to extract and use natural gas resources that contain high concentrations of acid gases. In many cases, most of the recovered CO₂ and H₂S is buried in landfills.^{11,12} This has necessitated the exploration of new strategies for the active use of C and S-containing substances (CO₂ and H₂S) to address the critical issues of global climate change and sustainability. H₂S, one of the sour gas components, is known to cause catalytic deactivation; hence, its academic value and economic benefits are expected to be significant if S-resistant catalysts can be used to directly convert the CO₂ in sour natural gas into value-added products.

There has been extensive research on high-performance catalysts for CO₂ conversion in the presence of H₂S. Sharma *et al.* reported that MoS₂ was effective for CO₂ hydrogenation under relatively high concentrations (<2500 ppm) of H₂S.¹³ They suggested that H₂S acts as a “conduit” for the active hydrogen required for the hydrogenation of CO₂, contributing to the enhancement of activity. Guilera *et al.* investigated the S resistance of Ni/Al₂O₃ doped with rare-earth oxides of CeO₂ and La₂O₃ for the low-temperature hydrogenation of CO₂.^{14,15} They reported that even in an atmosphere containing low concentrations (0.4 ppm) of H₂S, the Ni-based catalysts were not deactivated by sintering or S poisoning. To the best of our knowledge, there have only been a few reports on S-resistant CO₂ conversion catalysts. Our group studied the dehydrogenation of alkanes in the presence of high concentrations of H₂S (several tens of vol%), and found that Fe-based catalysts showed high activity and high selectivity with a stable performance.^{16,17}

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Therefore, we focused on Fe as a component as it is effective in water gas shift (WGS) reactions, and CeO₂ as the support because of its excellent functionality for CO₂ conversion.^{18–22} By employing an Fe-supported CeO₂ (Fe/CeO₂) catalyst in the reverse water gas shift (RWGS) reaction, a highly stable performance is expected even when the reaction atmosphere contains a high concentration of H₂S. In this study, the characteristics of the RWGS reaction using the Fe/CeO₂ catalyst with/without H₂S co-feeding were compared with those of the reaction using a commercial Cu–ZnO catalyst (MDC-7) and Co–Mo-based desulfurization catalyst, which are known to exhibit catalytic performance even in a high concentration of H₂S in the atmosphere. In addition, the stability of the Fe/CeO₂ catalyst was investigated over a reaction time of 12 h while co-feeding H₂S. The structure of the catalyst after the reaction was qualitatively investigated by X-ray diffraction (XRD) and STEM-EDX mapping to estimate the structure of the Fe and Ce component phases in the RWGS atmosphere in the presence of H₂S. Furthermore, to identify the effect of H₂S on the RWGS, the reaction mechanism was estimated by the temperature programmed reaction (TPR) with CO₂ and H₂ in the presence of H₂S.

2. Experimental

2.1. Preparation and characterization of the catalyst

The Fe/CeO₂ catalyst was prepared by an impregnation method. The CeO₂ (JRC-CEO-2) catalyst support was provided by the Catalysis Society of Japan. First, 2.0 g of CeO₂ was immersed in 30 mL of distilled water for 6 h, following which 1.624 g of Fe(NO₃)₃·9H₂O was added and stirred for an additional 2 h. The Fe-based catalyst was then calcined at 500 °C for 1 h. The loading amount of Fe was 10 wt%. MDC-7 (Clariant Catalysts K.K.), with 34 wt% Cu, was used for comparison with the Fe/CeO₂ catalyst. The reference Co–Mo-based catalyst was prepared by a previously reported method.²³ Some promoters of Zn and P components were included in the reference Co–Mo-based catalyst. The crystal structures of the Fe species and Ce species in the Fe/CeO₂ catalyst were determined *via* XRD (Ultima IV, Rigaku) using a Cu K α radiation source. The elemental mapping of the Fe/CeO₂ catalyst after the RWGS with H₂S co-feeding was performed by STEM-EDX (JEM-2100F, JEOL, Japan). The element maps of the Fe, Ce, S, and O components of the Fe/CeO₂ catalyst were collected for investigating the dispersion of the Fe components and coverage of the S atoms on the catalyst.

2.2. Evaluation of catalytic performance

Catalytic reaction tests were conducted at ambient pressure in a fixed bed reactor. After placing the catalyst in the center of the reaction tube, H₂ reduction (100 vol%) was carried out at 500 °C for 1 h. After purging the reduction gas, the reaction temperature was set at 400 to 600 °C, the catalyst amount was 75 mg, and the reaction gas with the composition CO₂/H₂/H₂S/He = 40/40/*x*/20 – *x* mL min^{–1} (*x* = 0–10). The gas was quantified by an FID gas chromatograph (GC-2014, Shimadzu,

Japan) and FPD gas chromatograph (GC-2030, Shimadzu, Japan).

2.3. Temperature programmed reaction method

To clarify the reaction mechanism of the RWGS, TPR analysis was carried out with the following steps. First, the Fe/CeO₂ catalyst of 100 mg was pretreated with the gas composition of H₂/Ar = 25/5 mL min^{–1} at 500 °C. After 1 h of the pretreatment, the reduced gas was purged by He, following the catalyst temperature was decreased to 40 °C. Then, CO₂ with co-feeding H₂S (gas composition: CO₂/H₂S/Ar/He = 1/1/5/93 mL min^{–1}) was supplied to the catalyst with increasing temperature from 40 °C to 800 °C. After that, the catalyst temperature was decreased to 40 °C, again. Subsequently, H₂ with co-feeding H₂S (gas composition: H₂/H₂S/Ar/He = 1/1/5/93 mL min^{–1}) was supplied with increasing temperature from 40 °C to 800 °C. The produced gases in these operations were monitored by a quadrupole mass spectrometer (QMS, HIDEN ANALYTICAL, UK).

3. Results and discussion

3.1. Effect of H₂S on RWGS performance

The performance of the Fe/CeO₂ catalyst was evaluated and compared with those of the reference MDC-7 and Co–Mo catalysts. The Co–Mo catalyst has been reported to be highly active in WGS reactions involving H₂S. Hence, this catalyst was used as a desulfurization catalyst, and its RWGS performance was evaluated. Fig. 1 shows the effect of the H₂S concentration on the CO₂ conversion of the Fe/CeO₂, and commercial MDC-7 and Co–Mo catalysts at various reaction temperatures. The RWGS reaction selectively proceeded since the selectivity to CO was *ca.* 100% under all conditions. For the Fe/CeO₂ catalyst, the CO₂ conversion increased with increasing reaction temperature, and reached 29.6% at 600 °C. Compared to its RWGS performance with no H₂S, there was no significant decrease in the catalytic performance at H₂S concentrations of 1, 5 and 10%. In other words, the performance of the Fe/CeO₂ catalyst was not adversely affected by H₂S. There has been no report on the catalyst for RWGS almost not degrading under a high

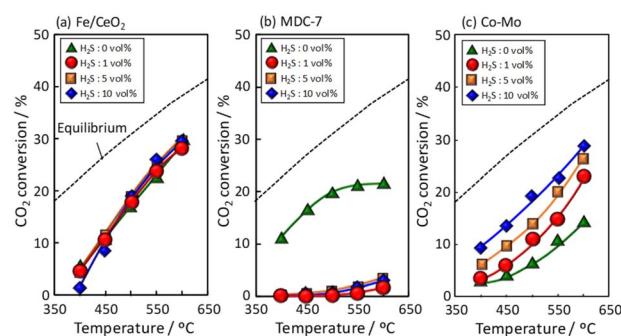


Fig. 1 Effect of H₂S co-feeding on the RWGS performance of (a) Fe/CeO₂, (b) MDC-7, and (c) Co–Mo catalysts. The reaction conditions were as follows: the catalyst amount was 75 mg, gas composition was CO₂/H₂/H₂S/He = 40/40/*x*/20 – *x* mL min^{–1} (*x* = 0–10), and reaction temp. ranged from 400 to 600 °C. Note that sel. of CO was 100%.



concentration of H₂S. The performance of the MDC-7 catalyst was superior to that of the Fe/CeO₂ catalyst. However, there was no significant increase in activity with increasing temperature, which is possibly attributed to the sintering and oxidation of the active Cu component by the water vapor produced during the progress of the RWGS reaction.²⁴ In addition, the co-feeding of H₂S greatly decreased the RWGS activity on the MDC-7 catalyst, where the CO₂ conversion was approximately 1.7% at 600 °C under 1 vol% H₂S. Interestingly, the degradation of the Cu-based catalyst progressed with the coexistence of H₂S, while that of the Fe/CeO₂ catalyst was almost negligible. For the Co-Mo catalyst, in a H₂S-free atmosphere, the WGS performance was low, but with increasing H₂S concentration, improved. In particular, the highest WGS performance was observed under 10 vol% H₂S. Although the Co-Mo catalyst was effective under a high concentration of H₂S, the Fe/CeO₂ catalyst showed better performance under a low concentration of H₂S, which indicated that the Fe/CeO₂ catalyst could be applicable to the biogas.

To evaluate the durability of the Fe/CeO₂ and Co-Mo catalysts, long-term RWGS reaction was tested under 1 vol% of H₂S. Moreover, to determine the role of the Fe component and CeO₂ support, RWGS tests were carried out over the Fe/SiO₂ and bare CeO₂ catalysts. Fig. 2 shows the CO₂ conversion over the Fe/CeO₂ and Co-Mo catalysts with the reaction time. When H₂S was supplied, the CO₂ conversion was high for 12 h over the Fe/CeO₂ catalyst. Since the products mostly contained CO, the RWGS selectively proceeded over the Fe/CeO₂ catalyst. Compared with the RWGS performance of the Co-Mo catalyst, that of the Fe/CeO₂ catalyst was superior, with high stability. The Fe/SiO₂ and bare CeO₂ catalysts were less active than the Fe/CeO₂ catalysts, indicating that both, Fe and CeO₂, are required for the RWGS when co-feeding H₂S. Since CeO₂ had a low activity and Fe/SiO₂ exhibited almost no activity, it was assumed that there might be an interaction between the Fe and CeO₂ components.

3.2. Structure characterization

To investigate the structure of the Fe/CeO₂ catalyst, XRD and STEM-EDX analyses were performed on the catalysts after the RWGS with H₂S co-feeding (Fig. 3(a) and (b), respectively). From Fig. 3(a), regardless of the change in the H₂S concentration, no

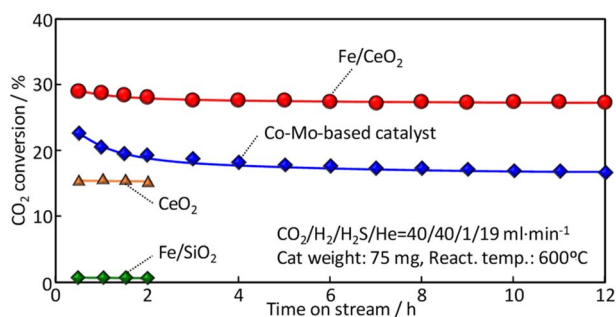


Fig. 2 Performances of Fe/CeO₂, Co-Mo, Fe/SiO₂, and CeO₂ catalysts for the RWGS with 1 vol% H₂S at 600 °C. Note that sel. of CO was 100%.

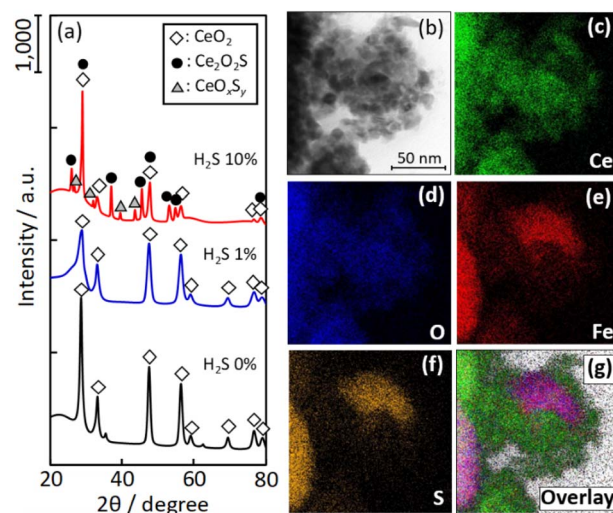


Fig. 3 (a) XRD pattern and (b)–(g) STEM-EDX mapping of the Fe/CeO₂ catalyst after RWGS with H₂S co-feeding.

diffraction peaks of the Fe component are observed and only the Ce component is detected. Further, only CeO₂ is detected when the H₂S concentrations are 0 vol% and 1 vol%, and the peak intensity of CeO₂ decreases with increasing H₂S supply. When the H₂S concentration is further increased from 1 vol% to 10 vol%, the CeO₂ phase transforms into oxysulfates of Ce₂O₂S and Ce₂O_xS_y. From the atomic mapping of Fig. 3(b), the Fe species are dispersed on the CeO₂ support; however, there are some areas where the Fe species are heavily segregated. The Fe mapping is visually similar to the S mapping, which suggests the absorption of S species on the Fe species in the catalyst. ESI 1† shows the thermodynamically stable phases of the Fe component from 25 °C to 700 °C in the reaction simulated gas atmosphere of CO and H₂, and H₂S. Since Fe_{0.877}S (major phase) and FeS (minor phase) are the stable phases of Fe species at the reaction temperature from 25 °C to 700 °C, the S identified from the STEM-EDX mapping is assumed to be incorporated in a sulfide phase.

Fig. 4 shows the Raman spectra of the Fe/CeO₂ catalyst after RWGS with different hydrogen sulfide concentrations (1 vol% and 10 vol%). Some peaks attributed to pyrrhotite (FeS_x) were observed at 213 cm⁻¹, 277 cm⁻¹, and 391 cm⁻¹ for the Fe/CeO₂ catalyst after RWGS containing 1 vol% and 10 vol% of H₂S.²⁶ The results are roughly in agreement with those inferred from thermodynamic calculations, as shown in ESI 1†.

3.3. Role of H₂S

To clarify the role of H₂S in the Fe/CeO₂ catalyst on the RWGS reaction, a TPR analyses were performed. Fig. 5(a) and (b) respectively show the profile of product during the CO₂-TPR and H₂-TPR analyses. The products were detected by a quadrupole mass spectrometer, where the temperature was increased by 10 °C min⁻¹, along with supplying 1 vol%-CO₂ or 1 vol%-H₂ in the presence of 1 vol%-H₂S to the Fe/CeO₂ catalyst. The shaded area indicated the amount change of each gas; positive area



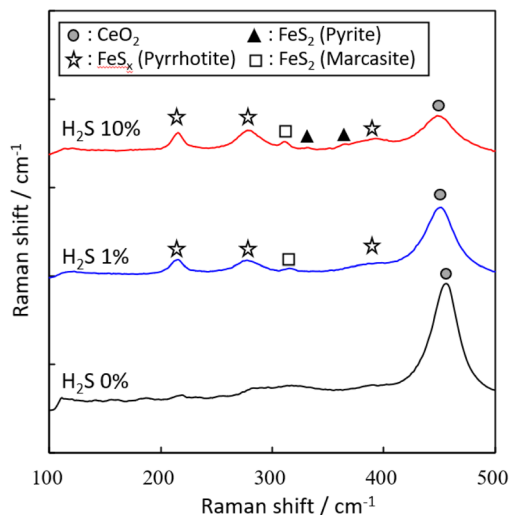


Fig. 4 Raman spectra of the Fe/CeO₂ after RWGS with H₂S co-feeding.

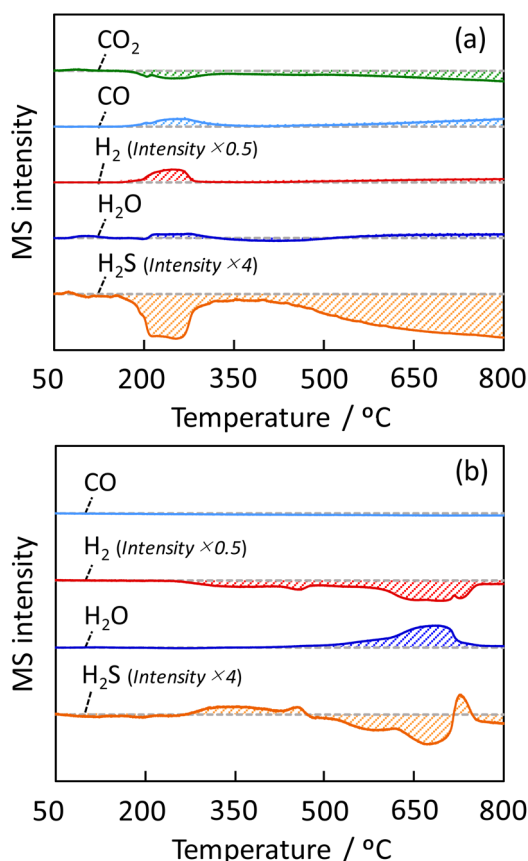


Fig. 5 Products behavior in TPR with (a) CO₂ and (b) H₂ in the presence of H₂S over the Fe/CeO₂ catalyst. The H₂ and H₂S intensities were adjusted because the sensitivity was quite different from the other products.

represents generation, and negative area represents consumption. By flowing CO₂ to the Fe/CeO₂ catalyst, CO is observed to form at from 180 °C to 350 °C. At the same temperature of the

formation of CO, the production of H₂ and H₂O and the consumption of H₂S were observed. The above phenomena can be understood as follows; H₂ generated by sulfurization of the Fe/CeO₂ catalyst is thought to have reduced CO₂ to form CO. Similarly, at temperatures above 350 °C, H₂S was consumed to generate H₂, and CO₂ was reduced to CO. The reason for the consumption of H₂S and CO₂ in both low (180–350 °C) and high temperature ($T > 350$ °C) regions is thought to be due to the formation of the lattice sulfide ion (S²⁻) by feeding H₂S on both the surface and the bulk of the catalyst. As confirmed by Fig. 5(b), H₂ was consumed at around 250 °C by the reaction with lattice S²⁻ ($H_2 + S^{2-} \rightarrow H_2S + V_S$). In this experiment, H₂-TPR is performed after CO₂-TPR; the regeneration of lattice S²⁻ from H₂S during CO₂-TPR was considered to contribute the formation of H₂S at low temperature from 250 to 400 °C. Furthermore, the formation of H₂O can be observed at temperatures above 500 °C. The result meant that a lattice oxygen was generated by catalyst oxidation ($CO_2 + V_{Ox} \rightarrow CO + O^{2-}$) in CO₂-TPR, and followed by the consumption of generated lattice oxygen by H₂ to produce H₂O.

Fig. 6 exhibits the ESR signal of oxygen and/or sulfur vacancy after TPR with CO₂ and H₂. The vacancy was confirmed at $g = 2.03$ before the reaction, due formation of lattice vacancy in CeO₂.²⁵ After TPR with CO₂ and H₂ with H₂S, the peak at $g = 2.03$ was not also confirmed, indicating that the vacancy was replenished as S²⁻/O²⁻ by H₂S/CO₂. Based on these data, H₂S might act as a sulfidizing-agent in the Fe/CeO₂ catalyst, meaning that generated lattice vacancy was replenished rapidly by S²⁻. To confirm this assumption, H₂S was supplied to the Fe/CeO₂ catalyst before the reaction. It is clear from this figure that the peak at $g = 2.03$ originating from defects has almost disappeared. The result suggests that the following reaction ($H_2S + V \rightarrow S^{2-} + H_2$) proceeds and restores lattice S²⁻ to the vacancy.

The role of H₂S was discussed from the viewpoint of reaction mechanism. The reaction pathway envisioned is that CO₂ is activated on the surface of the CeO₂ support²⁷ and reduced to CO by reacting with hydrogen activated by FeS_x. The coexisting H₂S is also expected to regenerate lattice S²⁻ to produce activated H atoms, and this generated hydrogen species further activates CO₂. In other words, the generated H₂ from H₂S on the

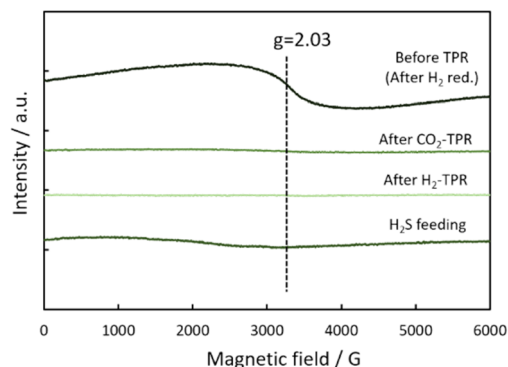


Fig. 6 ESR signal of the Fe/CeO₂ catalyst before TPR (after H₂ reduction), after H₂S feeding, after CO₂-TPR and H₂-TPR.



catalyst could work as hydrogen source to CO₂ reduction. In the case of Fe/CeO₂, the activity does not change much when the gas-phase H₂S concentration is changed because the bulk diffusion of lattice S²⁻ is the rate-limiting factor. In the Co-Mo system, on the other hand, the performance is greatly enhanced by the H₂S concentration, so we expect that H₂S → H₂ (active H species) + S²⁻ at the catalyst surface is the rate-limiting factor.

4. Conclusions

The catalytic properties of Fe/CeO₂ and MDC-7 catalysts for the RWGS reaction with H₂S co-feeding were evaluated. MDC-7 showed a relatively high activity in the low-temperature range, but underwent significant deactivation during the RWGS with H₂S co-feeding. The Fe/CeO₂ catalyst similarly showed a relatively high activity for the RWGS, and showed a stable activity for 12 h in the reaction atmosphere containing H₂S. The result TPR with CO₂ and H₂ indicated that co-feeding H₂S might act as the hydrogen source for RWGS over the Fe/CeO₂ catalyst.

Conflicts of interest

There are no conflicts to declare.

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References

- W. F. J. Burgers, P. S. Northrop, H. S. Khashgi and J. A. Valencia, *Energy Procedia*, 2011, **4**, 2178–2184.
- W. Taifan and J. Baltrusaitis, *Catal. Sci. Technol.*, 2017, **7**, 2919–2929.
- N. W. Chakroun and A. F. Ghoniem, *Int. J. Greenhouse Gas Control*, 2015, **41**, 163–173.
- A. Hofmann, W. van Strien, and R. Malekzadeh, in *Abu Dhabi International Petroleum Exhibition & Conference*, OnePetro, 2017.
- Y. Belmabkhout, G. De Weireld and A. Sayari, *Langmuir*, 2009, **25**(23), 13275–13278.
- J. I. Huertas, N. Giraldo and S. Izquierdo, *Mass transfer in chemical engineering processes*, 2011.

- P. Iovane, F. Nanna, Y. Ding, B. Bikson and A. Molino, *Fuel*, 2014, **135**, 352–358.
- W. Y. Lee, S. Y. Park, K. B. Lee and S. C. Nam, *Energy Fuels*, 2020, **34**(2), 1992–2000.
- K. Huang, X. M. Zhang, X. B. Hu and Y. T. Wu, *AIChE J.*, 2016, **62**, 4480–4490.
- K. Huang, D. N. Cai, Y. L. Chen, Y. T. Wu, X. B. Hu and Z. B. Zhang, *ChemPlusChem*, 2014, **79**(2), 241–249.
- C. Khan, R. Amin and G. Madden, *J. Pet. Explor. Prod. Technol.*, 2013, **3**, 55–60.
- A. Y. Kaita, O. Ogolo, X. Wu, I. Mohammed and E. A. Akpan, *J. Pet. Explor. Prod. Technol.*, 2020, **10**, 1575–1589.
- L. Sharma, R. Upadhyay, S. Rangarajan and J. Baltrusaitis, *ACS Catal.*, 2019, **9**, 10044–10059.
- J. Guilera, J. D. Valle, A. Alarcón, J. A. Díaz and T. Andreu, *J. CO₂ Util.*, 2019, **30**, 11–17.
- A. Alarcón, J. Guilera, R. Soto and T. Andreu, *Appl. Catal., B*, 2020, **263**, 118346.
- R. Watanabe, N. Hirata, Y. Yoda, Y. Fushimi and C. Fukuhara, *J. Jpn. Pet. Inst.*, 2020, **63**(4), 228–237.
- R. Watanabe, N. Hirata, K. Miura, Y. Yoda, Y. Fushimi and C. Fukuhara, *Appl. Catal., A*, 2019, **587**, 117238.
- Y. Sekine, T. Chihara, R. Watanabe, Y. Sakamoto, M. Matsukata and E. Kikuchi, *Catal. Lett.*, 2010, **140**(3), 184–188.
- R. Watanabe, Y. Sakamoto, K. Yamamuro, S. Tamura, E. Kikuchi and Y. Sekine, *Appl. Catal., A*, 2013, **457**, 1–11.
- C. Fukuhara, K. Hayakawa, Y. Suzuki, W. Kawasaki and R. Watanabe, *Appl. Catal., A*, 2017, **532**, 12–18.
- S. Ratchahat, M. Sudoh, Y. Suzuki, W. Kawasaki, R. Watanabe and C. Fukuhara, *J. CO₂ Util.*, 2018, **24**, 210–219.
- C. Fukuhara, A. Kamiyama, M. Itoh, N. Hirata, S. Ratchahat, M. Sudoh and R. Watanabe, *Chem. Eng. Sci.*, 2020, **219**, 115589.
- N. Nakajima, Y. Seriguchi, M. Kato, M. Yoshinari, T. Yoshida and K. Watanabe, *J. Jpn. Pet. Inst.*, 2013, **56**(6), 388–394.
- C.-S. Chen, W.-H. Chen and S.-S. Lin, *Appl. Catal., A*, 2004, **257**, 97–106.
- X. Chen, S. Zhan, D. Chen, C. Hea, S. Tian and Y. Xiong, *Appl. Catal., B*, 2021, **286**(5), 119928.
- Y. El Mendili, A. Abdelouas, H. El Hajja and J.-F. Bardeau, *RSC Adv.*, 2013, **3**, 26343.
- F. Cao, Y. Xiao, Z. Zhang, J. Li, Z. Xia, X. Hu, Y. Ma and Y. Qu, *J. Catal.*, 2022, **414**, 25–32.

