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## Film-forming, stable, conductive composites of polyhistidine/graphene oxide for electrochemical quantification of trace $\text{Pb}^{2+}$

 Zhe-Han Yang, \* Xin Lei, Guangming Jiang and Xianming Zhang\*  

Nanomaterials with unique properties, such as good film-formation and plentiful active atoms, play a vital role in the construction of electrochemical sensors. In this work, an *in situ* electrochemical synthesis of conductive polyhistidine (PHIS)/graphene oxide (GO) composite film (PHIS/GO) was designed to construct an electrochemical sensor for the sensitive detection of  $\text{Pb}^{2+}$ . Herein, GO as an active material can directly form homogeneous and stable thin films on the electrode surface because of its excellent film-forming property. Then GO film was further functionalized by *in situ* electrochemical polymerization of histidine to obtain plentiful active atoms (N). Due to strong van der Waals forces between GO and PHIS, PHIS/GO film exhibited high stability. Furthermore, the electrical conductivity of PHIS/GO films was greatly improved by *in situ* electrochemical reduction technology and the plentiful active atoms (N) in PHIS are profitable for adsorbing  $\text{Pb}^{2+}$  from solution, tremendously enhancing the assay sensitivity. With the above unique property, the proposed electrochemical sensor showed high stability, a low detection limit ( $0.045 \mu\text{g L}^{-1}$ ) and a wide linear range ( $0.1$ – $300 \mu\text{g L}^{-1}$ ) for the quantification of  $\text{Pb}^{2+}$ . The method can also be extended to the synthesis of other film-forming nanomaterials to functionalize themselves and widen their potential applications, avoiding the addition of non-conductive film-forming substances.

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## Introduction

The United States Environmental Protection Agency (U.S. EPA) reported that 10–20% of adults and 40–60% of infants are exposed to Pb *via* drinking water and food.<sup>1,2</sup> Even at trace concentrations, heavy metals pose a major threat to human health because of their inherent properties, including bio-accumulation and non-biodegradability as well as toxicity.<sup>3,4</sup> The World Health Organization (WHO) has established a guideline that limits the total Pb concentration in drinking water to 10 ppb.<sup>5</sup> Therefore, it is essential to accurately assay the total amount of Pb in drinking water and food. Commonly used assay methods include flame atomic absorption spectrometry,<sup>6</sup> inductively coupled plasma (ICP)-atomic emission spectrometry,<sup>7</sup> ICP-mass spectrometry,<sup>8</sup> and potentiometric ion selective electrodes.<sup>9</sup> Among these, the electrochemical method has aroused a lot of attention due to various advantages, such as simplicity, high sensitivity, low cost and the choice of functional materials.<sup>10,11</sup> Because the stable attachment of functional materials onto the electrode plays a key role in guaranteeing the excellent performance of an electrochemical assay for  $\text{Pb}^{2+}$ , the

functional materials need to mix with film-forming substances such as Nafion to enhance stability.<sup>12,13</sup> However, film-forming substances are organic polymers which dramatically hinder electron transport, decreasing the performance of the electrochemical assay for  $\text{Pb}^{2+}$ .<sup>14</sup> Therefore, a method that can attach functional materials onto an electrode surface free of film-forming substance is urgently being sought.

Graphene oxide (GO) as a dimensional layered material has been widely applied in various fields, such as electronics, energy, composite materials and bio-applications for its excellent electrical, thermal, and mechanical properties.<sup>15–18</sup> Moreover, unlike other large  $\text{sp}^2$ -conjugated structures, GO can easily form a homogeneous dispersion because it has rich  $-\text{OH}$  and  $-\text{COOH}$  groups on the GO surface.<sup>19</sup> More importantly, GO can be solution-processed into homogeneous and stable thin films, which endow it with the ability to modify electrode materials free of film-forming substances.<sup>15,20,21</sup> However, oxygen-containing functional groups (especially  $\text{C}=\text{O}$ ,  $-\text{COOH}$ ) have a strong electron absorption ability, leading to negative effects on electrical conductivity, decreasing the performance of the modified electrode.<sup>22</sup> Therefore, many attempts have been made to enhance the electrical conductivity and improve the performance of GO-modified electrodes. For example, L-cysteine that contains  $-\text{SH}$  and  $-\text{NH}_2$  was chosen as a functional molecule to modify GO by forming an amide bond between  $-\text{NH}_2$  in L-cysteine and  $-\text{COOH}$  in GO for the detection of  $\text{Pb}^{2+}$ .<sup>23</sup> The polypyrrole was functionalized on the GO surface to introduce N

Engineering Research Center for Waste Oil Recovery, Technology and Equipment of Ministry of Education, Chongqing Key Laboratory of Catalysis and Functional Organic Molecules, College of Environment and Resources, Chongqing Technology and Business University, Chongqing 400067, China. E-mail: [yzhzn89@ctbu.edu.cn](mailto:yzhzn89@ctbu.edu.cn); Tel: +86-023-62768056



atoms for adsorbing target metal ions.<sup>24</sup> Although the performance of the above-modified electrode was improved to some extent compared with the GO-modified electrode, its conductivity was not obviously enhanced due to the still existing amount of electron-withdrawing group (C=O).<sup>25</sup> Therefore, a method that can enhance electrical conductivity as well as improve the performance of the GO-modified electrode is still needed.

In this work, an *in situ* electrochemical synthesis of polyhistidine (**PHIS**)/GO film coupled with an electrochemical reduction strategy was designed to construct an electrochemical sensor for a Pb<sup>2+</sup> assay. Firstly, a GO dispersion solution was dropped onto the electrode surface and dried at room temperatures to obtain a GO film-modified electrode (GO/GCE). Then, *in situ* electrochemical polymerization was conducted to obtain **PHIS**-functionalized GO/GCE (PHIS/GO/GCE).<sup>26</sup> Next, the -C=O group in the PHIS/GO film was removed under the action of constant potential electrolyzation, achieving the reduction of PHIS/GO (r-PHIS/GO). Three outstanding properties of r-PHIS/GO endow the proposed electrochemical chemical sensor with excellent performance for the monitoring of Pb<sup>2+</sup>: (1) With its film-forming property, GO can directly form homogeneous and stable thin films at room temperature, avoiding the addition of a non-conductive film-forming substance. Moreover, due to the large sp<sup>2</sup>-conjugated structures, GO can help stabilize the **PHIS** film on the electrode for strong van der Waals forces between GO and **PHIS**. (2) Because of successive electrochemical reduction, the removal of the -C=O group in PHIS/GO film would increase the electrical conductivity. (3) **PHIS** is a polyaminoacid-bearing imidazole group ( $pK_a = 6.0$ ) that can provide plentiful active atoms (N) to adsorb Pb<sup>2+</sup>, enhancing the sensitivity of the assay. As a result, the proposed electrochemical sensor showed

high stability, a low detection limit (0.045  $\mu\text{g L}^{-1}$ ) and a wide linear range (0.1–300  $\mu\text{g L}^{-1}$ ) for the quantification of Pb<sup>2+</sup> (Scheme 1).

## Experimental section

### Reagents and instruments

Graphite oxide (GO) was obtained from Nanjing Xianfeng Nano Co. (Nanjing, China). Pb(NO<sub>3</sub>)<sub>2</sub>, Na<sub>2</sub>HPO<sub>4</sub>, NaH<sub>2</sub>PO<sub>4</sub>, KCl, sodium acetate (NaAc) and acetic acid (HAc) were purchased from Chengdu KeLong Chemical Reagents (Chengdu, China). **HIS** was obtained from Aladdin (Shanghai, China).

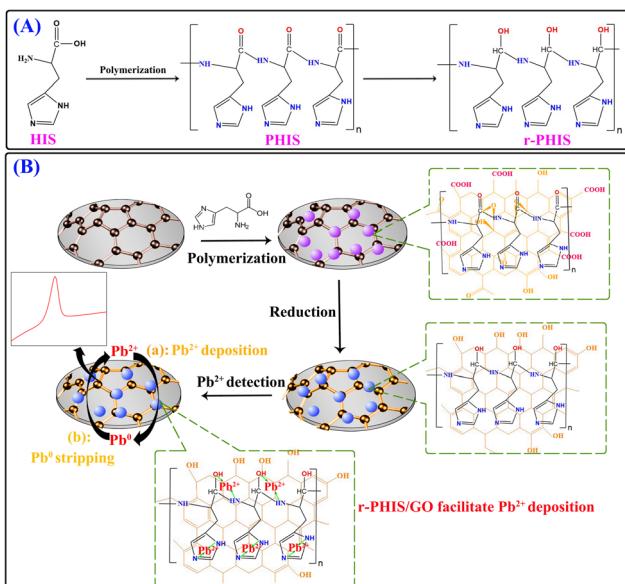
Electrochemical measurements were carried out on a CHI 660E electrochemical workstation (Shanghai Chenhua Instrument Co., Ltd, China) using a three-electrode system: a platinum wire auxiliary electrode, a saturated calomel reference electrode and a glassy carbon electrode (GCE,  $\Phi = 3\text{ mm}$ ) as the working electrode. X-ray photoelectron spectroscopy (XPS) data was obtained on the surface of the samples using an electron spectrometer (Thermo Escalab 250Xi, Thermo Scientific, USA) and fitted with XPS PEAK software. The morphology of the nanomaterials was characterized by a high-resolution scanning electron microscope (SEM, JSM-7001F INCA X-MAX, Japan). Cyclic voltammetry (CV) was performed in 5.0 mM [Fe(CN)<sub>6</sub>]<sup>3-/4-</sup> with a scanning potential from -0.2 to 0.6 V at a scan rate of 50 mV s<sup>-1</sup>. Electrochemical impedance spectroscopy (EIS) was performed at 5.0 mM [Fe(CN)<sub>6</sub>]<sup>3-/4-</sup> with parameters including a 50 mV s<sup>-1</sup> sweep rate, 0.22 V initial electric potential, 0.1 Hz low frequency, 100 000 Hz high frequency and a 2 s pulse period.

### Preparation of r-PHIS/GO modified electrode

Firstly, GCEs were pre-treated according to the reported protocol.<sup>22</sup> 10 mg of GO was dispersed into 10 mL of double-distilled water and ultrasonicated to form a homogeneous dispersion solution. After that, 5  $\mu\text{L}$  of the prepared GO solution was dropped onto the cleaned GCE and dried at room temperature to obtain the GO layer (GO/GCE). The GO/GCE was placed in a 0.1 M phosphoric acid buffer solution (PBS, pH 9.0) containing 0.02 M histidine (monomer) and CV was conducted with six cycles between -0.8 and +2.0 V (vs. Hg/Hg<sub>2</sub>Cl<sub>2</sub>) at a scan rate of 100 mV s<sup>-1</sup>, obtaining polyhistidine modified GO/GCE (PHIS/GO/GCE). Following that, the PHIS/GO/GCE was immersed in 0.5 M KCl solution, and successive electrochemical reduction was performed by chronoamperometry at a constant potential of -1.3 V for 1 h to achieve a reduced PHIS/GO layer (r-PHIS/GO).

### Pb<sup>2+</sup> quantification

Square wave anodic stripping voltammetry (SWASV) was used to analyze the concentrations of Pb<sup>2+</sup> in HAc buffer (pH 5.0) with various concentrations of Pb<sup>2+</sup>, where Pb<sup>2+</sup> was electrodeposited at 0.8 V for 240 s, and then stripped by SWV from -1.1 to -0.1 V with a step size of 5 mV, pulse width of 0.2 s and amplitude of 50 mV. At the end of each detection test, a 1 V potential was applied on the working electrode for 100 s in order to remove the deposited residual species from its surface.



**Scheme 1** (A) Schematic representation of molecular structure before and after polymerization and electrochemical reduction. (B) Schematic illustration of the modification process of an electrochemical sensor and its application in the detection of Pb<sup>2+</sup>.



## Results and discussion

### Characterization of the as-prepared r-PHIS/GO

The morphologies of GO and r-PHIS/GO were characterized using SEM. As can be seen in Fig. 1A, GO showed restacked layers and wrinkles in some regions, which may be ascribed to electronic repulsion among the soft layers. Fig. 1B shows the morphologies of **PHIS** prepared by direct deposition onto the GCE surface. It can be seen that the **PHIS** is composed of large irregular lumps. When **PHIS** was deposited on the GO surface, **PHIS** showed smaller nano-scale irregular particles, dispersed uniformly on the GO surface, indicating that GO helps to form uniformly distributed **PHIS**.

### Elemental analysis of r-PHIS/GO

Photo electron spectra were performed to investigate the change in the elements during the preparation process of r-PHIS/GO. The detailed N 1s peaks of pure **HIS**, **PHIS/GO** and r-PHIS/GO are shown in Fig. 2A. The spectrum of pure **HIS** can be split into three peaks at around 398.3, 399.7 and 400.5 eV. The peaks at 398.3 eV and 399.7 eV were attributed to the imidazole group ( $N=C-N$  and  $N=C-N$ ). The peak at 400.5 eV was assigned to amino functions. When the electrochemical polymerization of **HIS** onto the GO/GCE surface was conducted, only two peaks at 399.1 eV and 400.1 eV could be observed. This result was due to the fact that the reaction between amino functions and carboxyl was driven at  $-1.3$  V with the formation of  $O=C-NH$ , and electrochemical polymerization caused the imidazole group peaks ( $N=C-N$  and  $N=C-N$ ) to shift to higher binding energy. After electrochemical reduction, no obvious change in imidazole group peaks was observed. Similarly, for a better comparison, XPS analysis of C 1s was further investigated (Fig. 2B). The four XPS peaks of C 1s for pure **HIS** could be observed at 284.4 eV, 284.9 eV, 285.4 eV and 287.4 eV, which were attributed to  $C-C/C=C$ ,  $C=O/N-C=N$ ,  $C-NH_2$  and  $COOH$ , respectively. However, the peaks corresponding to  $COOH$  disappeared after electrochemical polymerization, suggesting that  $COOH$  was transferred to other groups ( $O=C-NH$ ).

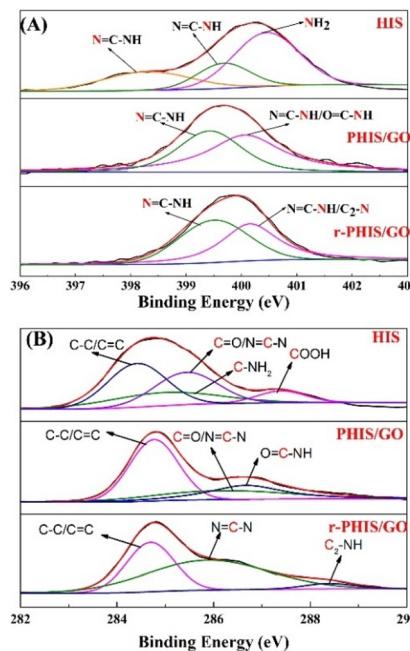


Fig. 2 (A) N 1s XPS spectra of pure **HIS**, **PHIS/GO**, r-PHIS/GO; (B) C 1s XPS spectra of pure **HIS**, **PHIS/GO**, r-PHIS/GO.

Furthermore, a new peak at 288.3 eV could be observed after electrochemical reduction, which implies the formation of  $H-N-C_2$  bonds.

### Characterization of the stepwise fabrication of the chemical sensor

Fig. 3A shows the cyclic voltammograms recorded in  $0.1\text{ mol L}^{-1}$  PBS (pH 9.0) containing  $0.02\text{ mol L}^{-1}$  **HIS** (monomer) using GO/GCE. An anodic peak at 1.25 V could be observed due to oxidation and deposition of the **HIS** monomers, indicating that **HIS** could be successfully polymerized. Fig. 3B shows the  $i-t$  curve of **PHIS/GO/GCE** under 1.3 V for 1 h in 0.5 M KCl. When  $t > 0$ , a large current can be observed, and then the current decreases rapidly. With an increase in electrolysis time, the current gradually decreases and finally remains relatively stable, indicating that the electrochemical properties of the electrode materials tend to be stable. To characterize the stepwise fabrication of r-PHIS/GO/GCE, electrochemical measurements (CV and EIS) were conducted. As shown in Fig. 3C, an obvious decrease in CV redox peaks in GO/GCE could be observed in comparison with GCE, indicating that GO was successfully attached onto the GCE surface because amounts of O-containing groups on GO surface hindered electron transfer. After *in situ* polycondensation in **HIS** solution, a further decrease in peak current was obtained, indicating that **PHIS** was successfully deposited onto the GO/GCE surface. To improve the conductivity of **PHIS/GO/GCE**, constant voltage scanning was performed and the peak currents of r-PHIS/GO/GCE were apparently increased, which indicated that g-constant voltage scanning was beneficial for improving the conductive property of **PHIS/GO/GCE**. In addition, EIS was conducted to further verify the CV results. As shown in Fig. 3D, GO/GCE exhibited

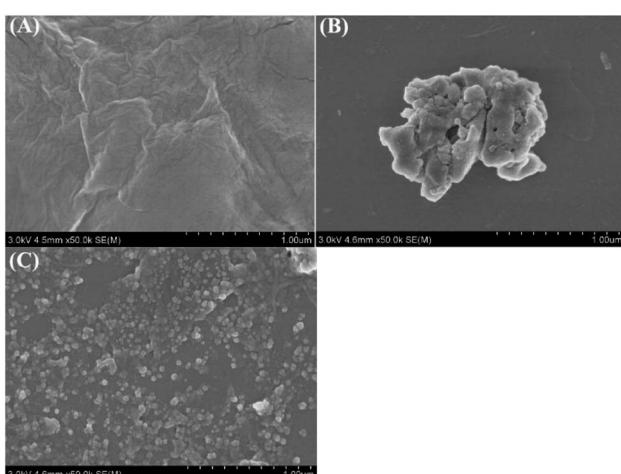


Fig. 1 SEM images of GO (A), **PHIS** (B) and r-PHIS/GO (C).



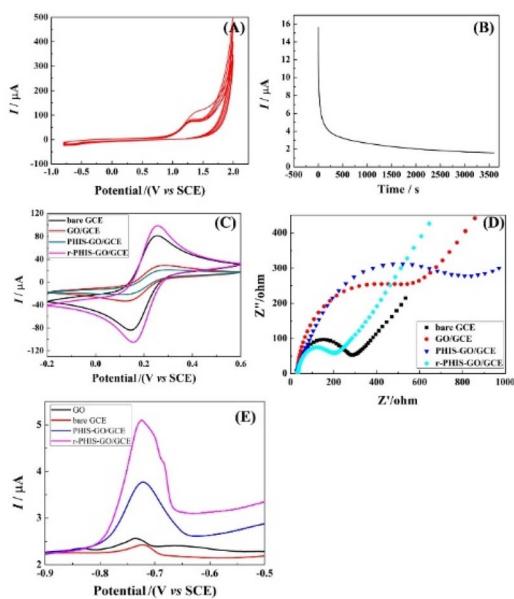


Fig. 3 (A) The cyclic voltammograms of GO/GCE in 0.1 M PBS (pH 9.0) containing 0.02 M HIS (monomer). (B)  $i$ - $t$  curve of PHIS/GO/GCE under 1.3 V for 1 h in 0.5 M KCl. Electrochemical characterizations of different modified electrodes: (C) CV and (D) EIS in the  $[\text{Fe}(\text{CN})_6]^{3-/4-}$  system. (E) SWASV curves of bare GCE, GO/GCE, PHIS-GO/GCE and r-PHIS-GO/GCE in 0.1 M acetate buffer solution (pH 5.0) containing 100  $\mu\text{g L}^{-1}$   $\text{Pb}^{2+}$ .

bigger impedance than GCE. An increase in impedance was obtained after **HIS** polymerization. A dramatic decrease in EIS value was achieved after electrochemical reduction. The EIS characterization result is consistent with the CV results, suggesting the successful preparation of r-PHIS/GO/GCE.

### Electrochemical response of different modified electrodes

To assess the superiority of r-PHIS/GO/GCE for  $\text{Pb}^{2+}$  assay, four modified electrodes, bare GCE, GO/GCE, PHIS/GO/GCE, and r-PHIS/GO/GCE, were prepared for a comparison of the response to the same concentration of  $\text{Pb}^{2+}$ . Fig. 3E shows the stripping current curves of the above four electrodes in 0.1 M HAc buffer solution (pH 5.0) containing 100  $\mu\text{g L}^{-1}$   $\text{Pb}^{2+}$ . The red line and black line show the stripping peak current at bare GCE and GO/GCE, respectively. It can be seen that both electrodes show weak current responses. Moreover, a remarkable increase in the stripping peak current for  $\text{Pb}^{2+}$  ions was obtained at PHIS/GO/GCE (blue line), indicating that **PHIS** was beneficial for absorbing more metal ions. Furthermore, compared with PHIS/GO/GCE, the current response of r-PHIS/GO/GCE was obviously increased (red line), which may be ascribed to the excellent electrical conductivity of r-PHIS/GO/GCE. Therefore, r-PHIS/GO/GCE is superior for a  $\text{Pb}^{2+}$  assay.

### Optimization of conditions

To achieve optimal performance for  $\text{Pb}^{2+}$  detection, the assay conditions, including the pH of the detection solution, polymerization time of **HIS** and deposition time of  $\text{Pb}^{2+}$  were

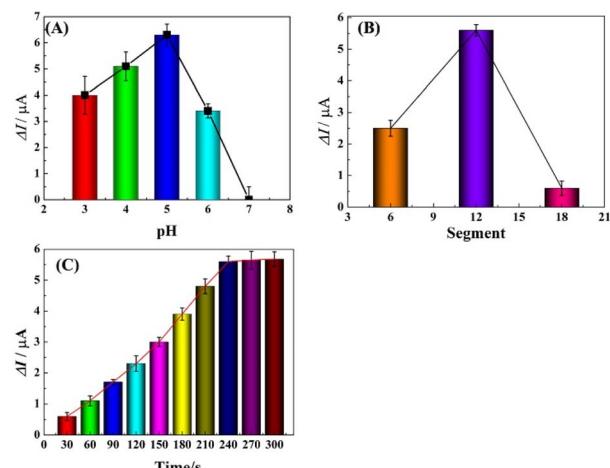


Fig. 4 Influence of pH value (A), deposition potential (B) and accumulation time (C) on the stripping peak currents.

optimized. The stripping responses of  $\text{Pb}^{2+}$  in HAc buffer solution of various pH (ranging from 3.0 to 7.0) were investigated using SWASV. As shown in Fig. 4A, the peak current response was enhanced from pH 3.0 to 5.0 and decreased from 6.0 to 7.0, which was ascribed to the competitive binding between proton ions and metal ions to the donating atoms in lower pH solution and the hydrolysis of  $\text{Pb}^{2+}$  in high pH solution. Therefore, pH 5.0 was used as the optimized value and applied in the following studies.

Electrochemical characterization verifies that **PHIS** is non-conductive; thus the amount of **PHIS** should be optimized by controlling the polymerization time (CV measurement used segments to mean polymerization time) and comparing the stripping current responses. As shown in Fig. 4B, the effect of polymerization time in the range of 6 to 18 segments was investigated. It can be seen that the current response of 100  $\mu\text{g L}^{-1}$   $\text{Pb}^{2+}$  was enhanced from 6 segments to 12 segments and then decreased from 12 segments to 18 segments. Thus, the optimal polymerization time was 12 segments.

The stripping current response was influenced by the deposition time of  $\text{Pb}^{2+}$ ; thus the effect of deposition time on the stripping current response at r-PHIS/GO/GCE was further investigated using SWASV. From Fig. 4C, we can see that the current response of 100  $\mu\text{g L}^{-1}$   $\text{Pb}^{2+}$  was enhanced with increasing deposition time from 30 s to 240 s and then tended to plateau from 240 s to 300 s. Thus, we chose 240 s as the optimal deposition time.

### Control of electron transfer process

In order to further investigate the electron transfer process on r-PHIS/GO/GCE, the CV current at different scan rates was investigated. As shown in Fig. 5A, the redox peak currents increased with an increment in scan rates, accompanied by an increase in potential gap. Moreover, the oxidation and reduction peak currents were linearly proportional ( $R^2 = 0.9989, 0.9992$ ) to the square root of the scan rate ( $v^{1/2}$ ) in the



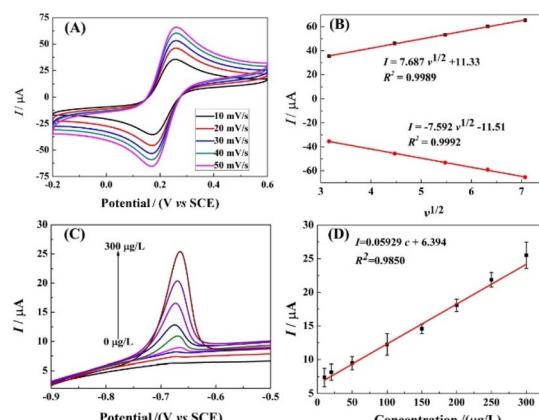


Fig. 5 (A) CV response in PBS buffer containing 5.0 mM  $[\text{Fe}(\text{CN})_6]^{3-/4-}$  at different scan rates: 10, 20, 30, 40, and 50  $\text{mV s}^{-1}$ ; (B) calibration of linear relationship between the CV currents and scan rates. (C) SWASV response of various  $\text{Pb}^{2+}$  concentrations at  $\text{g-C}_3\text{N}_4/\text{r-GO/GCE}$  in HAC buffer (pH 5.0); (D) calibration of linear relationship between the SWASV currents and the concentrations of  $\text{Pb}^{2+}$ .

range from 10 to 50  $\text{mV s}^{-1}$ . The linear equations are  $I_{\text{pa}} (\mu\text{A}) = 8.678v^{1/2} + 11.33$  and  $I_{\text{pc}} (\mu\text{A}) = -7.592v^{1/2} - 11.51$  (Fig. 5B). The above results suggest that the electron transfer process on r-PHIS/GO/GCE was diffusion-controlled. Therefore, the stripping peak current value is relevant to the concentration of  $\text{Pb}^{2+}$ .

### Performance of r-PHIS/GO/GCE for $\text{Pb}^{2+}$ detection

Under the optimal conditions, the performance of r-PHIS/GO/GCE for  $\text{Pb}^{2+}$  detection was assessed in HAc buffer (pH 5) containing various concentrations of  $\text{Pb}^{2+}$  by SWASV. As shown in Fig. 5C, increased oxide currents were observed with the increase in  $\text{Pb}^{2+}$  concentration and were shown to be linear with concentration in the range of 1.0–300  $\mu\text{g L}^{-1}$ . The equation of the calibration curve was  $I (\mu\text{A}) = 0.05929c (\mu\text{g L}^{-1}) + 6.394$  with correlation coefficient ( $R^2$ ) 0.9850 (Fig. 5D). The limit (3S/N) was 0.045  $\mu\text{g L}^{-1}$ . Furthermore, Table 1 shows a comparison of the performance of r-PHIS/GO/GCE for  $\text{Pb}^{2+}$  with some previous reported work. The results suggested that the performance of r-PHIS/GO/GCE for  $\text{Pb}^{2+}$  was acceptable and competitive.

Table 1 Comparison of r-PHIS/GO/GCE for  $\text{Pb}^{2+}$  assay with previous reported work

Electrode materials	Method	Detection limit ( $\mu\text{g L}^{-1}$ )	Detection range ( $\mu\text{g L}^{-1}$ )	Ref.
BiMOF/CP	SWASV	2.7	8.5–34	27
$\text{Au}@\text{Nb}_4\text{C}_3\text{T}_x$	DPASV	0.83	2.1–1035	28
$\text{Au}@\text{SiO}_2@\text{Fe}_3\text{O}_4/\text{NG}$	DPASV	0.60	5–80	29
ZnO-RGO	DPASV	0.090	0.5–10	30
$\text{g-C}_3\text{N}_4/\text{r-GO}$	SWASV	0.15	1–300	25
$\text{NH}_2\text{-MIL-53}(\text{Cr})$	SWASV	$6.0 \times 10^{-3}$	$8.3 \times 10^{-2}$ to $1.6 \times 10^3$	31
2D MoS <sub>2</sub> nanofilm	SWASV	0.30	0–20	1
$\text{UiO-66-NH}_2$	DPASV	0.030	$1.0 \times 10^{-3}$ to $1.0 \times 10^3$	32
Boron nitride	LSASV	0.15	$4.0\text{--}1.5 \times 10^3$	33
r-PHIS/GO	SWASV	0.045	0.1–300	This work

### Specificity and stability

The specificity of r-PHIS/GO/GCE was assessed by challenging it against other usual metal ions, including  $\text{Al}^{3+}$ ,  $\text{Fe}^{3+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Mg}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Ca}^{2+}$ ,  $\text{Hg}^{2+}$ ,  $\text{K}^+$ , and  $\text{Na}^+$ . The concentration of the interfering ions was 5  $\mu\text{g L}^{-1}$  and the concentration of target  $\text{Pb}^{2+}$  was 50  $\mu\text{g L}^{-1}$ . As shown in Fig. 6, the results demonstrated that a 100-fold high concentration of interfering metal ions showed minimal current response, but a large current response was exhibited at 50  $\mu\text{g L}^{-1}$   $\text{Pb}^{2+}$  (peak current change <7%), suggesting r-PHIS/GO/GCE has excellent selectivity for a  $\text{Pb}^{2+}$  assay. Moreover, the stability was assessed by storing r-PHIS/GO/GCE with physical protection and measuring the same concentration of  $\text{Pb}^{2+}$  every four days. After storage for 20 days, the current response was 89.1% of the initial response, indicating that r-PHIS/GO/GCE showed relatively excellent stability for a  $\text{Pb}^{2+}$  assay.

### Analysis of real samples

The prepared r-PHIS/GO/GCE was further challenged to detect  $\text{Pb}^{2+}$  from industrial wastewater. Prior to detection, all of the samples were filtered through a 0.45 mm membrane. The four industrial wastewater samples (pH 5.0) were detected by the r-PHIS/GO/GCE and ICP methods, respectively. The percentage variance between the measured results of the above two methods for the four samples was within 9.5%, revealing that the prepared r-PHIS/GO/GCE was suitable for the practical detection of real samples with acceptable accuracy (Table 2).

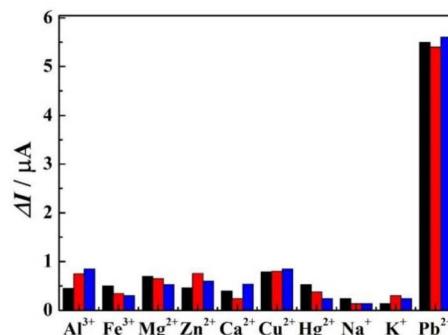


Fig. 6 The specificity of r-PHIS/GO/GCE for  $\text{Pb}^{2+}$  detection.

**Table 2** Analysis of real samples using the electrochemical method and ICP

No.	Electrochemical method	ICP	Percentage variance (%)
1	2.57	2.7	9.5
2	10.5	11	4.5
3	89.6	83.9	6.7
4	112.3	122.9	8.6

## Conclusions

In conclusion, the *in situ* electrochemical synthesis of r-PHIS/GO/GCE was designed to construct a chemical sensor for a Pb<sup>2+</sup> assay. With the film-forming property of GO and *in situ* polymerization of **PHIS**, r-PHIS/GO/GCE avoids using an extra film-forming substance to stabilize nanomaterials on the electrode surface. GO also helps stabilize **PHIS** film on an electrode for strong van der Waals forces between GO and **PHIS**. Because of the removal of the -C=O group in PHIS/GO film by electrochemical reduction and the plentiful active atoms (N) in **PHIS**, r-PHIS/GO/GCE ensures high sensitivity for Pb<sup>2+</sup> detection. Furthermore, the method can also be extended to the synthesis of other film-forming nanomaterials to functionalize themselves and widen their potential applications.

## Conflicts of interest

The authors declare no competing financial interest.

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