


 Cite this: *RSC Adv.*, 2023, **13**, 7834

# N-Doped carbon nanoparticles on highly porous carbon nanofiber electrodes for sodium ion batteries†

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Nitrogen doped carbon nanoparticles on highly porous carbon nanofiber electrodes were successfully synthesized *via* combining centrifugal spinning, chemical polymerization of pyrrole and a two-step heat treatment. Nanoparticle-on-nanofiber morphology with highly porous carbon nanotube like channels were observed from SEM and TEM images. Nitrogen doped carbon nanoparticles on highly porous carbon nanofiber (N-PCNF) electrodes exhibited excellent cycling and C-rate performance with a high reversible capacity of around 280 mA h g<sup>-1</sup> in sodium ion batteries. Moreover, at 1000 mA g<sup>-1</sup>, a high reversible capacity of 172 mA h g<sup>-1</sup> was observed after 300 cycles. The superior electrochemical properties were attributed to a highly porous structure with enlarged *d*-spacings, enriched defects and active sites due to nitrogen doping. The electrochemical results prove that N-PCNF electrodes are promising electrode materials for high performance sodium ion batteries.

Received 30th January 2023

Accepted 3rd March 2023

DOI: 10.1039/d3ra00635b

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## 1. Introduction

Lithium ion batteries (LIBs) have been the most widely used rechargeable batteries in portable electronics. However, limited sources and the high price of lithium metal limit the application of LIBs in power grids and electric vehicles. Considering the abundant sodium sources, sodium ion batteries (SIBs) have been presented as an alternative to LIBs. However, low energy density and the poor cycle life of SIBs need to be addressed for further development of SIBs.<sup>1–8</sup>

Carbon materials have been reported as electrodes for SIBs owing to their good electrical conductivity, cost effectiveness, and excellent chemical and physical stability, as well as their high porosity and good cycling performance.<sup>9–12</sup> Moreover, disordered carbons with larger interlayer space are beneficial for Na<sup>+</sup> insertion and extraction. Fabricating porous carbon as well as N-doping have been studied for improving the capacity and cycling performance of carbon anodes. Porous carbon

nanofibers have the merits of high conductivity, excellent flexibility, a short ion diffusion pathway, good stress tolerance and a large surface area to volume ratio.<sup>13,14</sup> N doping creates defects and improves electronic conductivity, thus electrochemical performance is improved by providing more sites for Na<sup>+</sup> storage.<sup>15–17</sup>

Polyacrylonitrile (PAN) has been commonly used as carbon precursor owing to its high carbon yield and good mechanical properties.<sup>18–22</sup> Production of high performance electrode materials *via* low cost and facile approach is vital to use sodium ion batteries in practical applications.<sup>3,9,23</sup> Morphology, porosity and pore structure, affect the electrochemical properties of carbon materials and nanofiber production technique influence the porosity and pore structure. Centrifugal spinning allows fast and cost-effective production of nanofibers and centrifugally spun nanofibers show higher porosity compared to electrospun nanofibers. Furthermore, porosity and pore size could be easily adjusted in this technique.<sup>19,24</sup> Polymethylmethacrylate (PMMA) and polystyrene (PS) are the most commonly used polymers to create pore structure in carbon fibers because of their low carbon yield. These polymers decompose and create pores during high temperature carbonization processes.<sup>3,25,26</sup> The novel morphology of N-doped carbon nanoparticles on porous carbon nanofibers obtained from polypyrrole (PPy) coated centrifugally spun PAN/PMMA and PAN/PS nanofibers have not been studied yet. Herein, two types of blend nanofibers, PAN/PMMA and PAN/PS were centrifugally spun as precursors to prepare highly porous carbon nanofiber electrodes for sodium ion batteries for the first time. The effect of low carbon yield polymer type (PMMA and PS) and content of these polymers on

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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3ra00635b>



the morphology of carbon nanofibers were investigated. Furthermore, N-doped carbon nanoparticles on highly porous carbon nanofiber electrodes were synthesized *via* combining centrifugal spinning, chemical polymerization of pyrrole and two step heat treatment. N-PCNF electrodes delivered high reversible capacity with excellent cycling performance owing to N-doping and highly porous structure with enlarged interlayer spacings. Considering low cost and facile production approach and excellent electrochemical properties, nitrogen doped carbon nanoparticles on highly porous carbon nanofiber electrodes are promising electrode materials for sodium ion batteries.

## 2. Experimental

### 2.1 Chemicals

Polyacrylonitrile (PAN,  $M_w = 150\ 000$ ), polymethylmethacrylate (PMMA,  $M_w = 120\ 000$ ), polystyrene (PS,  $M_w = 192\ 000$ ), pyrrole, ammonium per sulfate (APS), hydrochloric acid, ethanol and *N,N*-dimethyl formamide (DMF) were supplied from Sigma-Aldrich.

### 2.2 Preparation of N-doped carbon nanoparticles on highly porous carbon nanofibers

PAN/PMMA blend solutions with two different PAN/PMMA weight ratios (1/0.2, 1/0.4) and PAN/PS blend solution with the ratio of 1/0.4 (PAN/PS) were prepared in DMF and labeled as PAN/PMMA1, PAN/PMMA2 and PAN/PS, respectively. These blend solutions were fed into the centrifugal spinning device by a syringe pump. The rotational speed of 4000 rpm, the feeding rate of 60 ml h<sup>-1</sup>, the collector distance of 20 cm, and the nozzle diameter of 0.5 mm was applied in the centrifugal spinning. 10 wt% polymer blend solutions were used in the centrifugal spinning and approximately 5.7 g of nanofibers produced after an hour due to the losses of the solution during production. PAN, PAN/PMMA1, PAN/PMMA2 and PAN/PS were stabilized at 250 °C for 2.5 h in air and carbonized at 800 °C for 2 h in N<sub>2</sub> and named as CNFs, PCNF1, PCNF2 and PCNF3, respectively. The carbon yield was around 40 wt%. In order to synthesize nitrogen doped carbon nanoparticles on highly porous carbon nanofibers, pyrrole was polymerized on PAN/PMMA2 and PAN/PS blend nanofibers by using APS as an initiator. In order to further improved the electrochemical performance of PCNFs, PAN/PMMA2 and PAN/PS blend nanofiber precursors were selected for PPy coating because of the higher porosity and enhanced capacity of PCNFs prepared from PAN/PMMA2 and PAN/PS blend nanofibers compared to those of CNFs prepared by using PAN and PAN/PMMA1. Pyrrole and APS were dissolved in aqueous hydrochloric acid solution separately. PAN/PMMA2 and PAN/PS blend nanofibers were immersed in pyrrole solution at 0 °C and then APS solution was added dropwise to pyrrole solution. The molar ratio of pyrrole to APS was 1 : 1 and the polymerization was ended in 12 hours. Polypyrrole (PPy) coated PAN/PMMA2 and PAN/PS blend nanofibers were washed and dried and then PPy coated PAN/PMMA2 and PAN/PS blend nanofibers were stabilized at 250 °C for 2.5 h in air and

carbonized at 800 °C for 2 h in N<sub>2</sub> and then named as N-PCNF4 and N-PCNF5 respectively.

### 2.3 Structure characterization

The morphology of nanofibers was investigated by FEI Quanta FEG 250 scanning electron microscope and Hitachi HighTech HT7700 tunneling electron microscope (TEM). The structure of all studied sample was analyzed by using XRD and Raman spectroscopy.

### 2.4 Electrochemical characterization

A slurry of N-PCNFs, carbon black and sodium alginate in deionized water in the weight ratios of 0.8 : 0 : 1 : 0.1 was coated on Cu foil and dried at 60 °C overnight and the electrode weight was approximately 1.5 mg cm<sup>-2</sup>. The electrodes were assembled in CR2032 coin type cells with Na metal and the electrolyte of 1 M NaClO<sub>4</sub> in EC : PC (1 : 1 in volume). Galvanostatic charge discharge measurements were performed *via* Neware and Hefa battery test systems at room temperature.

## 3. Results and discussion

### 3.1 Structural characterization

Schematic diagram of the preparation of nitrogen doped carbon nanoparticle on highly porous carbon nanofibers (N-PCNF) was presented in Fig. 1. Centrifugally spun PAN, PAN/PMMA, PAN/PS were prepared and then pyrrole was polymerized on PAN/PMMA and PAN/PS nanofibers by using APS as oxidant. Polypyrrole coated PAN/PMMA and PAN/PS nanofibers were used as precursors to synthesize nitrogen doped carbon nanoparticles on highly porous carbon nanofibers while PS and PMMA were used as pore generator. 1D porous channels observed from TEM images were fabricated *via* removal of PMMA or PS during heat treatment while PPy was used for N-doping and carbon nanoparticle precursors.

The morphologies of CNFs, PCNFs and N-PCNFs were investigated through SEM and TEM images. SEM images of CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5 are presented in Fig. 2. Fibrous structure was observed from all studied carbon nanofibers. In order to evaluate the effect of blend ratio and blend polymer type on the morphology, SEM images with higher magnification are presented in Fig. 3. CNFs had rough surface while PCNF1, PCNF2 and PCNF3 fabricated from PAN/PMMA and PAN/PS nanofibers had pores on fiber surfaces. SEM images of N-PCNF4 and N-PCNF5 show the nanoparticle-on-nanofiber morphology. Moreover, as shown in Fig. 4, TEM images of PCNF1, PCNF2 and PCNF3 depict highly porous structure with tubular structure. PCNFs synthesized from PAN/PMMA and PAN/PS showed similar tubular structure after carbonization without significant difference. The nanoparticle-on-nanofiber morphology with highly porous nanofibers was observed from TEM images of N-PCNF4 and N-PCNF5 as well. SEM and TEM images are consistent with each other and nanoparticle on nanofiber structure was observed from both SEM and TEM images of N-PCNF4 and N-PCNF5. It has been reported that tubular hollow structures can act as a buffer and



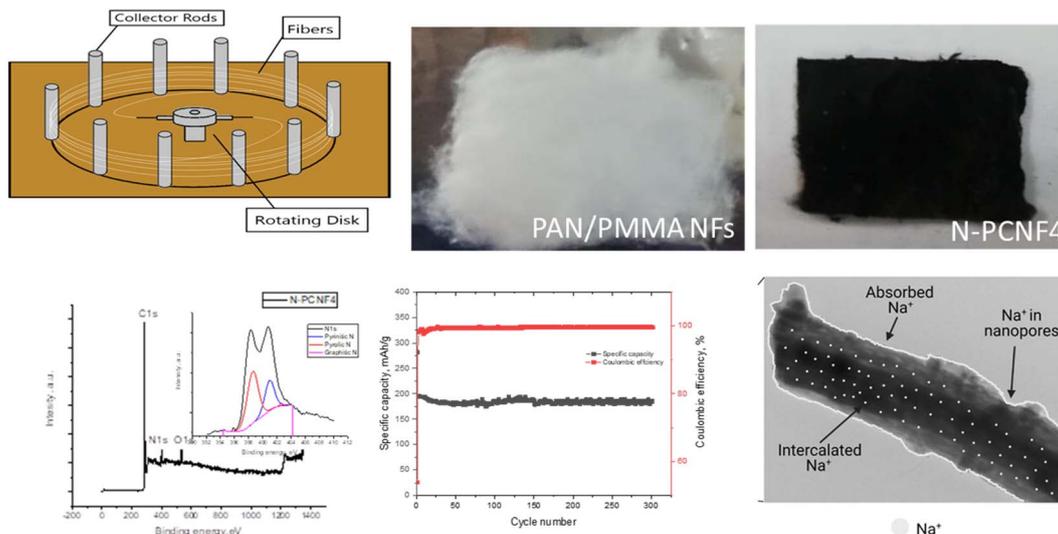


Fig. 1 Schematic diagram of the preparation of N-PCNF4.

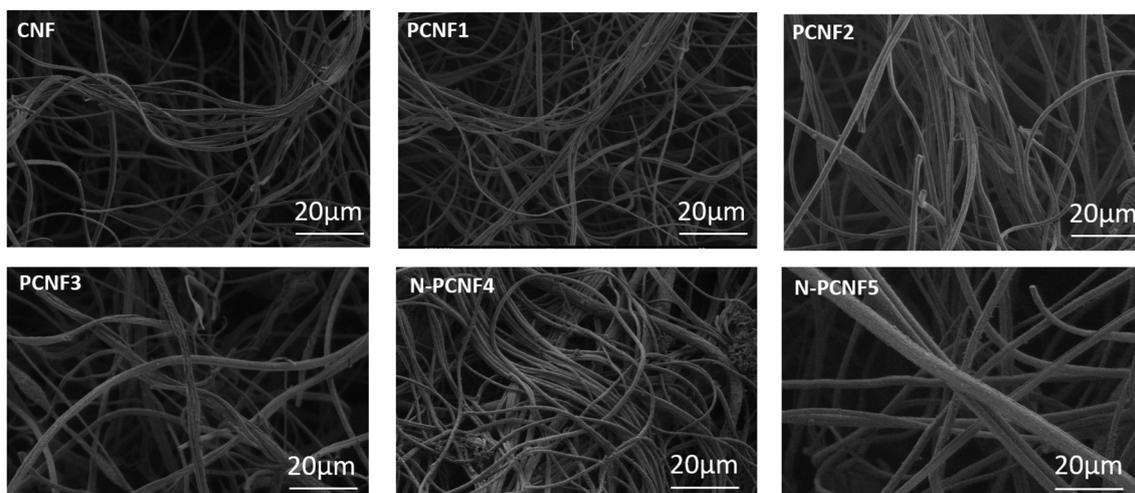


Fig. 2 SEM images of CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5.

relieve stresses caused by  $\text{Na}^+$  insertion and extraction during cycling and improve structural stability. Furthermore, larger surface areas and N doping could further improve the capacity and high rate cycling stability.<sup>13</sup>

The surface areas of the CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5 were 162, 176, 214, 183, 265, 219  $\text{m}^2 \text{g}^{-1}$ , respectively, as reported in Table 1.  $\text{N}_2$  adsorption/desorption curves were also presented in Fig. S1.† As compliance with the SEM and TEM images, increasing porosity and N doping led to higher surface areas which is beneficial for Na ion insertion and extraction since large surface area provides larger contact and improved transport of ions and electrons. Increasing surface area with N doping was also observed from N doped carbon framework by Huang *et al.*<sup>27</sup> and the result was ascribed to disturbed carbon framework and more exposed surface area caused by N doping. Pore volume was also

presented in Fig. S1b.† There are micro, meso and macropores in the electrodes. It has been reported that diverse nanoporosity increased the capacity by creating more nanochannels and thus let fast kinetics during  $\text{Na}^+$  intercalation and extraction.<sup>7</sup>

Raman spectra of CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5 are exhibited in Fig. 5a. Two peaks at 1340 and 1540  $\text{cm}^{-1}$  are ascribed to D band and G bands respectively. The intensity ratio between D and G band ( $I_D/I_G$ ) is sensitive to carbon microstructure, defects and degree of disorder. The  $I_D/I_G$  values for N-PCNF4 is 1.9 whereas other samples have the ratio of 1.7. The high intensity ratio proves increased disordered structure and defects and it could be ascribed to N doping.<sup>1,15</sup>

XRD patterns of CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5 are presented in Fig. 5b. Two broad peaks which are around  $25^\circ$  and  $44^\circ$  can be ascribed to (002) and (100) planes of graphite. The low intensity and broad peaks indicate highly



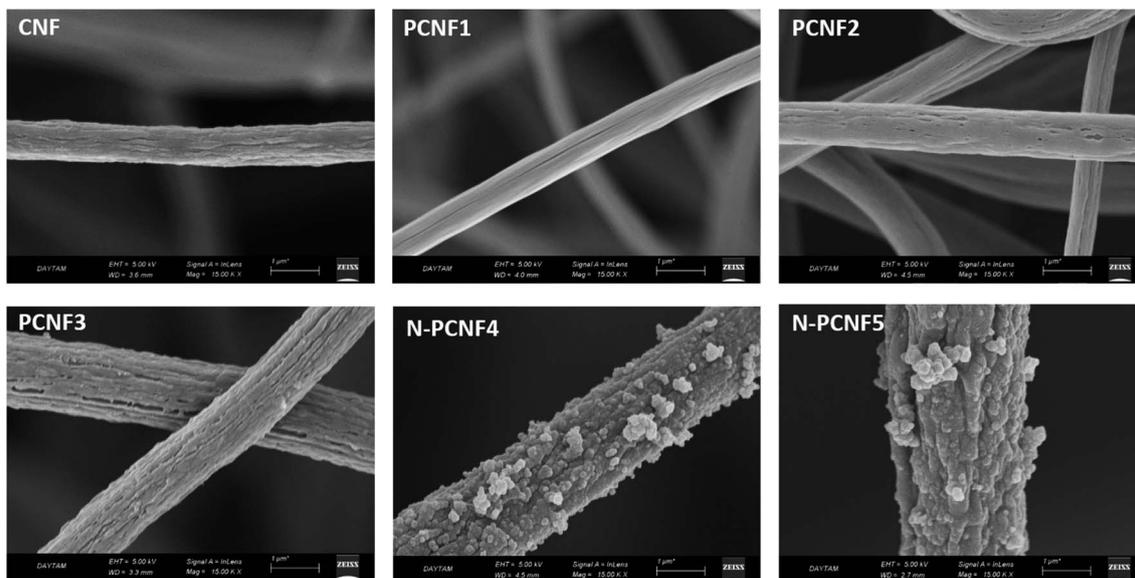


Fig. 3 SEM images of CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5 with higher magnification.

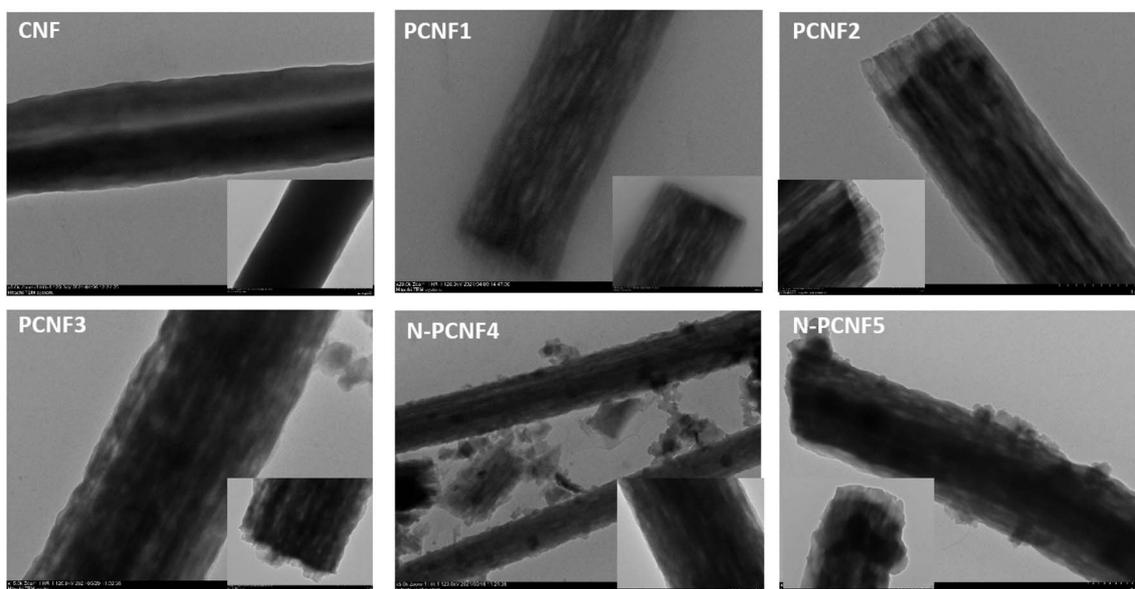


Fig. 4 TEM images of CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5.

Table 1 Specific surface area values of CNFs, PCNFs and N-PCNFs

	Surface area, $\text{m}^2 \text{g}^{-1}$
CNF	162
PCNF1	176
PCNF2	214
PCNF3	183
N-PCNF4	265
N-PCNF5	219

amorphous structure.<sup>1,26,28</sup> Moreover, interlayer spacings ( $d_{(002)}$ ) were found to be 0.36 for CNFs and PCNFs which is larger than that of graphite (0.34 nm) resulting better  $\text{Na}^+$  storage. For  $\text{Na}^+$  storage, highly disordered matrix is beneficial.<sup>16</sup> The large interlayer spacings of 0.37 are seen for N-PCNF4 and N-PCNF5 which is consistent with Raman spectra and hence proves N doping. Larger interlayer spacing caused by N doping was also reported for pitch derived carbons by Hao *et al.*<sup>29</sup> Both XRD and Raman spectra confirm larger  $d$  spacings and enhanced defects which is beneficial for  $\text{Na}^+$  storage properties.

To provide further structural analysis, XPS characterization was performed and XPS survey for N-PCNF4 and N-



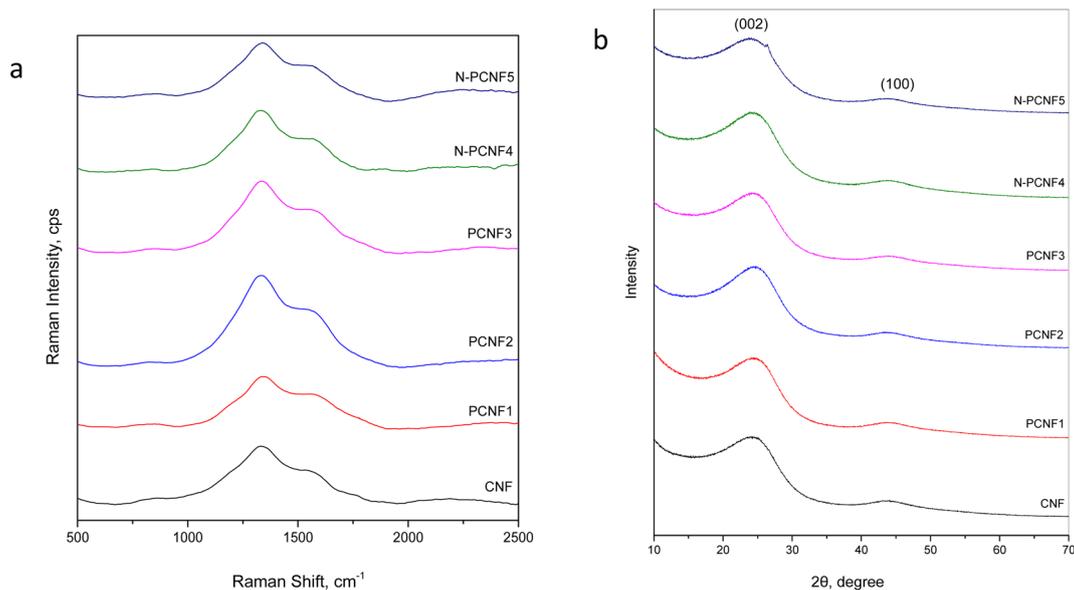


Fig. 5 Raman (a) spectra and XRD pattern (b) of CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5.

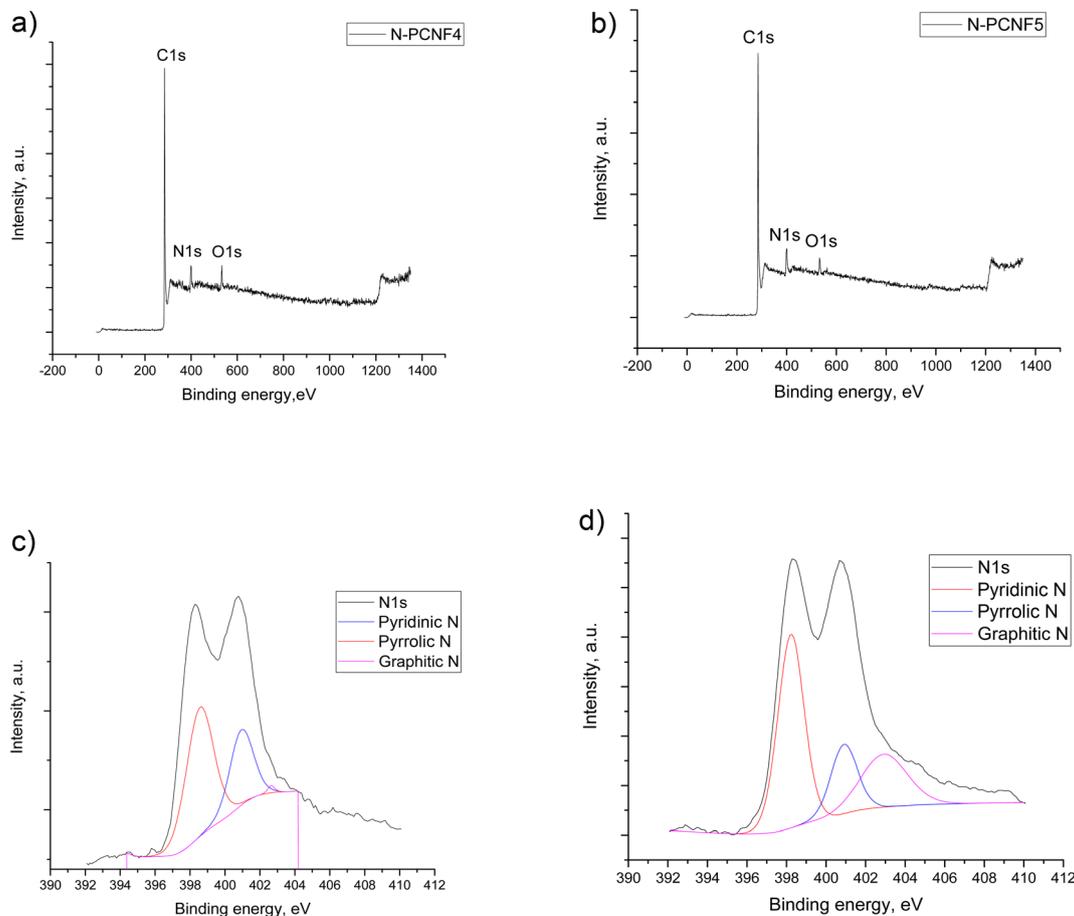


Fig. 6 XPS survey (a and b) and N 1s deconvoluted plots (c and d) of N-PCNF4 and N-PCNF5.



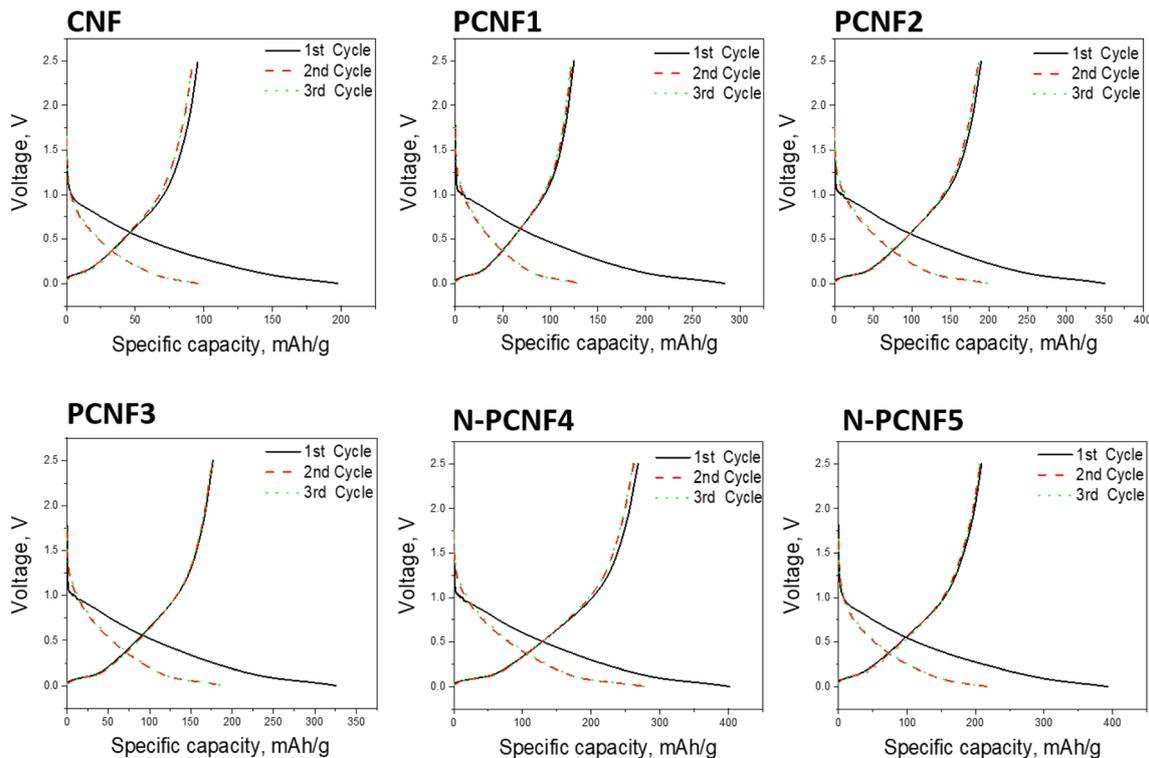


Fig. 7 First cycle charge discharge curves for CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5.

PCNF5 are presented in Fig. 6. Both N-PCNF4 and N-PCNF5 have clear N1s peaks, indicating a high amount of nitrogen doping. Nitrogen contents are 9 and 6%, respectively for N-PCNF4 and N-PCNF5. Moreover, the deconvolution of N1s peaks showed that the compositions of pyridinic and pyrrolic are high which result in more conductive and reversible sites for Na<sup>+</sup> ion adsorption.<sup>7,30</sup> The amounts of graphitic-N (402.6 eV), pyrrolic-N (400.8 eV) and pyridinic-N (398.5 eV) are 1, 33, and 66%, respectively for N-PCNF4 whereas those are 26, 20 and 54% respectively for N-PCNF5 which confirms Raman results that suggest higher disordered structure of N-PCNF4.

### 3.2 Electrochemical characterization

The sodium storage properties of the electrodes were investigated in 2032 type half cells using sodium as the counter electrode. Fig. 7 shows the discharge and charge curves for the first cycles for all studied electrodes. During the first cycle, the discharge capacities are 197, 284, 350, 325, 402 and 393 mA h g<sup>-1</sup> respectively for CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5 while the charge capacities are 95, 136, 189, 176, 266 and 212 mA h g<sup>-1</sup> respectively for CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5. The large irreversible capacity in the first cycle can be attributed to the electrolyte decomposition and formation of SEI layer.<sup>15</sup> In the following charge discharge cycles, capacities became stable, indicating reversible intercalation/deintercalation of Na<sup>+</sup> (Table S1 and Fig. S2†). Moreover, the small plateau and large sloping capacity

regions indicate that a large part of the capacity was contributed by adsorption of Na<sup>+</sup> ions into the nanopores and the reversible redox reaction between Na<sup>+</sup> ions and functional groups, whereas a minor capacity was shared by intercalation in graphitic nanolayers as reported before for hard carbon included carbon nanofiber electrodes.<sup>7,31</sup>

The cycling performance of the electrodes at the current density of 50 mA g<sup>-1</sup> was displayed in Fig. 8. The reversible capacities of around 95, 120, 190, 170, 270 and 220 mA h g<sup>-1</sup> were seen from CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5 respectively. The higher capacities were observed from PCNFs compared to CNF and N doping further enhanced the reversible capacities up to 270 mA h g<sup>-1</sup>. The highest reversible capacity of around 270 mA h g<sup>-1</sup> was achieved by N-PCNF 4 electrode after 100 cycles among all studied electrodes. Enhanced capacities could be attributed to increased defects, and increased N content which well consistent with Raman and XRD results. Large amount of disordered matrix beneficial for Na<sup>+</sup> storage and large content of disordered part implies larger Na<sup>+</sup> storage active sites thus results in higher conductivity.<sup>28</sup> Qu *et al.*<sup>13</sup> also synthesized N doped carbon fiber electrodes *via* electrospinning and dopamine coating and enhanced reversible capacity of around 200 mA h g<sup>-1</sup> at 50 mA h g<sup>-1</sup> obtained from N doped carbon electrodes compared to that of (around 100 mA h g<sup>-1</sup>) carbon electrodes. Enhanced performance was attributed to porous structure and N doping. N containing functional groups promote redox reactions and defects enhance sodium ion adsorption. Moreover, porous structure contacts better with electrolyte and shorten ion pathway as well.



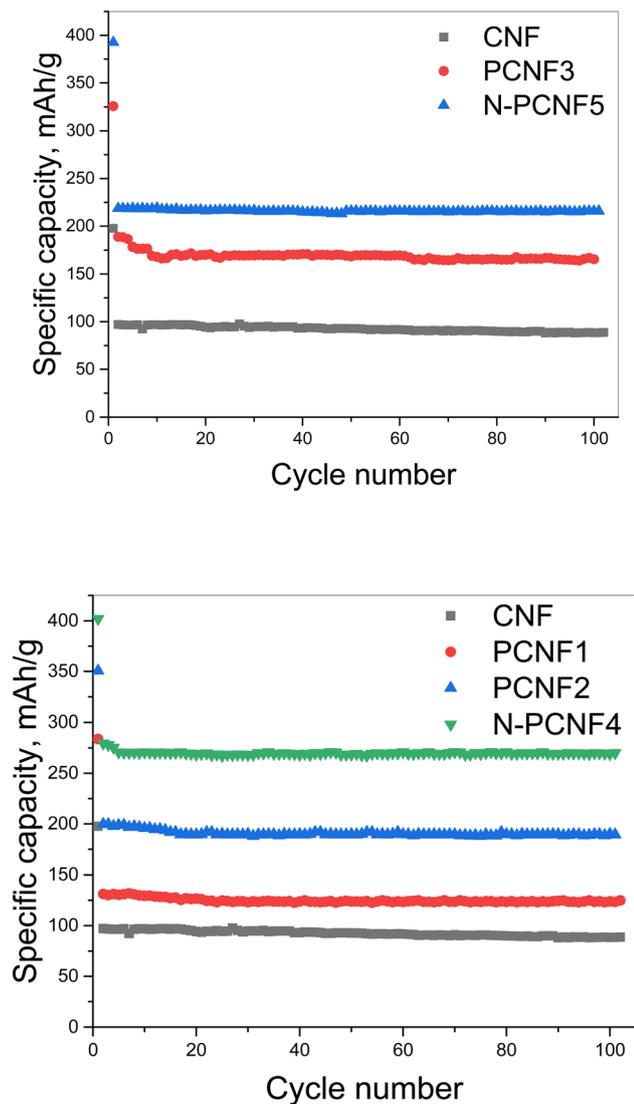


Fig. 8 Cycling performance of CNF, PCNF1, PCNF2, PCNF3, N-PCNF4 and N-PCNF5 electrodes at a current density of  $50 \text{ mA g}^{-1}$ .

N doped nanoparticle on nanofiber electrode with tubular like porous structure and expanded interlayer spacings showed the highest reversible capacity among all studied electrodes. Moreover, good cycling performance of N-PCNF4 electrode at high current density of  $1000 \text{ mA g}^{-1}$  was also studied and the result was shown in Fig. 9a. At  $1000 \text{ mA g}^{-1}$ , in 300 cycles reversible capacity was around  $172 \text{ mA h g}^{-1}$  with a high capacity retention over 88%. A high coulombic efficiency over 98% was also achieved. It has been reported that the doped N can enlarge the spacing distance and improve reversibility during charge discharge cycles.<sup>29</sup>

C-rate performance of the electrodes were investigated and the results are shown in Fig. 9b. The superior capacities of around 270, 256, 238, 220, 185  $\text{mA h g}^{-1}$  were observed from N-PCNF 4 electrode at the current densities of 0.05, 0.1, 0.2, 0.4 and 1 A  $\text{h g}^{-1}$  whereas those of CNF electrodes were 98, 88, 69,

49 and 35  $\text{mA h g}^{-1}$ . The capacity retention exceeded 60% even after 20-fold increase in current density. Furthermore, when the C rate was back to  $50 \text{ mA h g}^{-1}$ , the capacity of  $270 \text{ mA h g}^{-1}$  was observed from N-PCNF4.

The electrochemical impedance spectroscopy (EIS) spectra at an alternating current (AC) voltage of 5 mV amplitude in the 100 kHz to 0.01 Hz were also performed and presented in Fig. 9c. All of the studied electrodes showed semicircle arc shaped curves representing the charge-transfer resistance on electrode/electrolyte interface and the inclined line in the low frequency region, the Warburg impedance, is associated with the sodium-ion diffusion process in the electrodes. The diameter of semi circle is approximately 500 ohm for CNFs however that of is approximately 300 ohm for PCNFs. Moreover, the diameter decreased to 150 and 100 ohm for N-PCNF5 and N-PCNF4, respectively. Introducing porosity decreased the charge transfer resistance and N doping further decreased the resistance owing to larger accessible area and improved electronic conductivity with N inclusion which was proved by XPS survey. Lower resistance with N doping was also reported for carbon anodes and the result was attributed to enhanced conductivity.<sup>32</sup>

The best performance was seen from N-PCNF4 and this result could be ascribed to superior morphology that combine highly porous carbon nanofibers and PPy based N doped carbon nanoparticles. PCNFs provide 3d interconnected highly conductive electrode structure while N doped carbon nanoparticles provide high capacity along with fast diffusion kinetics. It has been reported that N doping change the electronic state of the carbon surface and thus improve electrical conductivity of the materials. Moreover, it increases defects on the surface of the carbon materials, thereby enhancing the capacity.<sup>21</sup> Ghani *et al.*<sup>7</sup> was also studied hard carbon based electrodes and reported superior capacity with high C rate performance. In N-PCNF4, high amount of nitrogen functional groups and porous fiber structure provided more active sites, enhanced electrical conductivity, and surface polarity, all led to excellent cycling performance. The excellent C rate performance of nitrogen doped carbon nanoparticle on highly porous nanofiber electrodes could be ascribed to novel morphology with more active sites owing to large surface area between electrodes and electrolytes, short ion diffusion distance and increased interlayer spacing which overcomes kinetic limitations. In disordered carbon electrodes,  $\text{Na}^+$  insert between carbon layers and  $\text{Na}^+$  are accommodated into mesopores. Large interlayer spacing between layers accelerates  $\text{Na}^+$  transport and storage. Moreover, N doping generate defects and form disordered carbon structure and thus further enhance  $\text{Na}^+$  absorption properties. Additional N increase electronic conductivity as well. In summary, more active sites with low charge transfer resistance, expanded interlayer spacing *via* N doping, large electrode electrolyte interface and short path for ions and electrons due to nanoparticle-on-nanofiber morphology, porous structure and N doping led to high cycling and C-rate performance.



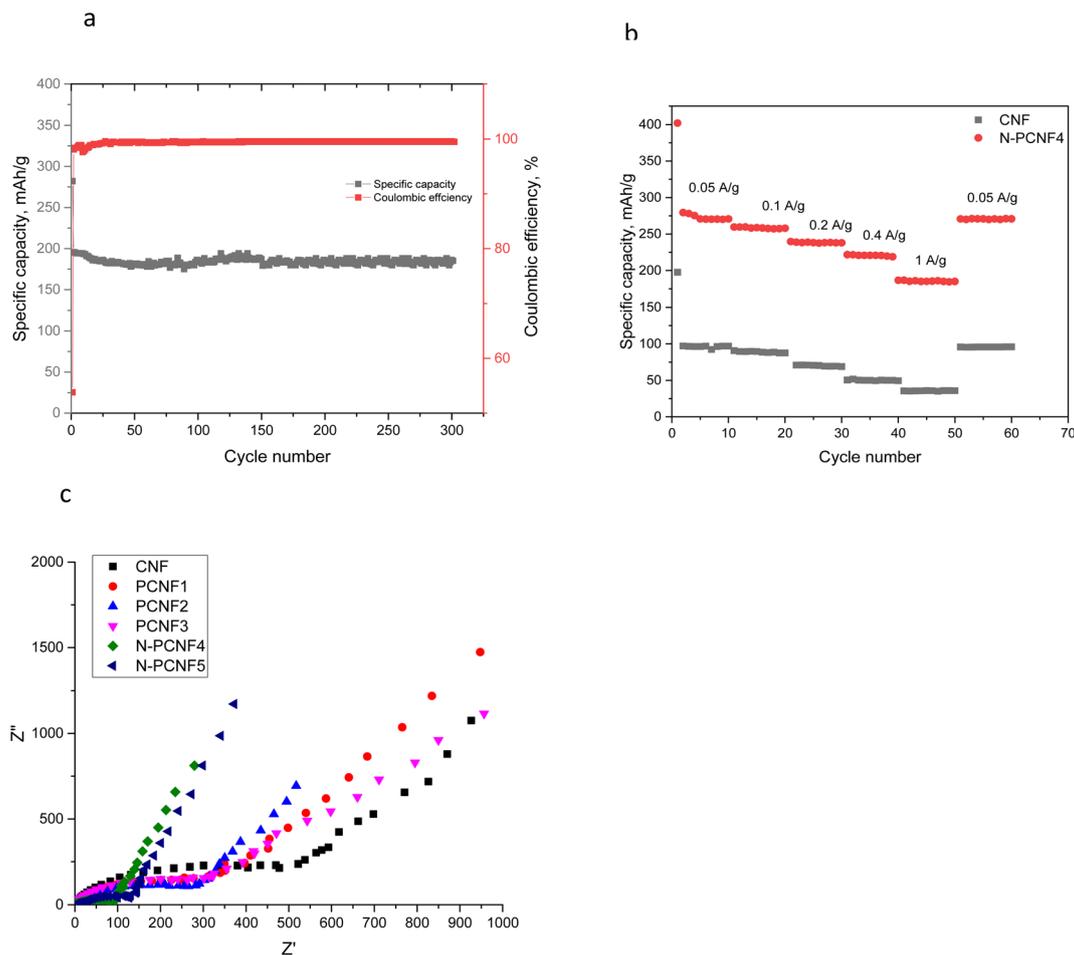


Fig. 9 Cycling performance of N-PCNF4 electrodes at a current density of  $1000 \text{ mA g}^{-1}$  (a) C-rate performance of CNF and N-PCNF4 electrodes (b) and EIS spectra of the electrodes (c).

## 4. Conclusion

Nitrogen doped carbon nanoparticle on highly porous carbon nanofiber electrodes were synthesized *via* low cost and facile approach for sodium ion batteries. Ppy coating, centrifugal spinning and heat treatment was employed to create high performance electrodes with novel nanoparticle-on-nanofiber morphology for the first time. High reversible capacity of  $270 \text{ mA h g}^{-1}$  with good cycling and C rate performance was ascribed to the highly porous, disordered carbon structure and N doping. This facile way could be employed for the development of next generation electrodes for energy storage systems.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

This research was supported by The Scientific and Technology Research Council of Turkey (TUBITAK) (219M348) and Korea Institute of Industrial Technology (PEM23020).

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