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# Synthesis and performance evaluation of slow-release fertilizers produced from inverse vulcanized copolymers obtained from industrial waste†

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To improve crop nutrient uptake efficacy (NUE) and better manage fertilization, slow-release fertilizers (SRFs) are developed by either coating the urea granules or making a composite. Several materials have already been developed, nevertheless, scalability of those materials is still a challenge due to their inherent drawbacks (such as hydrophilicity, crystallinity, non-biodegradability, etc.). Herein, we utilized a biodegradable, green and sustainable copolymer produced from industrial waste (sulfur-petroleum industry waste and myrcene-citrus industry waste) to coat the urea using a facile coating method to develop novel SRFs and achieve better agronomic and environmental advantages. The copolymer was first synthesized using a facile, solvent-free one-pot method called inverse vulcanization followed by Fourier transform infrared spectroscopy (FTIR) analysis to confirm the successful reaction between myrcene and sulfur subsequently coating the copolymer on urea granule. The morphology and coating thickness of coated fertilizers were analysed using scanning electron microscopy (SEM), followed by a nitrogen release test in distilled water and a soil burial test to confirm the biodegradability. The nitrogen release test revealed that the SRF with the maximum coating thickness of 1733  $\mu\text{m}$  releases only 16% of its total nitrogen after 4 days of incubation compared to the pristine urea which releases all its nutrient within 1 day. The soil burial test confirms the biodegradability of the copolymer, as after 50 days of incubation in soil the copolymer loses almost 18.25% of its total weight indicating that the copolymer is degrading.

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## 1 Introduction

The rapid increase in World population accompanied by a recent hike in inflation rate, swift urbanization, and land degradation due to massive floods are major roots for food shortage and a rise in food prices. These factors demand sustainable and efficient agricultural practices. To meet the food demand, agriculture sectors are consuming a massive amount of fertilizers. Urea is a well-known nitrogen (macro-nutrient required by plants) enriched fertilizer commonly used to boost the crop yield. However, owing to its hygroscopic

nature it releases the nutrients at a pace that is too fast for the plants to absorb, prompting the loss of nutrients (due to nitrification–denitrification, volatilization, surface run-off and nitrate leaching) and pollution of the environment through leaching to underground water sources leading to low nutrient use efficiency (NUE) (30–70%).<sup>1</sup> One remedy to control this issue and better manage fertilization, is to develop slow-release fertilizers (SRFs).<sup>2</sup>

Slow- or controlled-release fertilizers (SRFs/CRFs) are designed such that the active fertilizing agent is released in a delayed/controlled way, specifically taking into consideration plants' nutrient requirement therefore providing superior nutrient use efficiency (NUE) and halting the environmental pollution caused by leaching of surplus nutrient into soil and water.<sup>3,4</sup> NUE refers to the capacity of crops to absorb and consume nutrients for maximum yield.<sup>5</sup> An ideal SRF/CRF would be biodegradable, environmentally friendly, cost-effective, and would slow-down the nutrient release speed such that a single application would be enough for the entire life span of crop.

The approach used for the manufacture of SRFs is either coating the fertilizer with hydrophobic material or making

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composites by infusing urea in macromolecule that would slow down the nutrient release process.<sup>6</sup> Over the years, numerous classes of coating materials, ranging from inorganic materials to synthetic and natural polymers, have been explored. The drawbacks associated with these coatings include brittleness of inorganic materials,<sup>4</sup> hazardous and nonbiodegradable nature of synthetic polymers<sup>7,8</sup> and hydrophilicity of natural polymers.<sup>9</sup> Therefore, an investigation needs to be done in order to find a biodegradable, stable, and cheap alternative coating to mitigate these problems.

According to an estimate elemental sulfur is produced as a waste by-product of the crude oil and natural gas processing facilities at a rate of 70 million tons per year and this amount is expected to rise as more contaminated petroleum feedstocks are being utilized to meet global energy demand.<sup>10</sup> Sulfur is being used as a coating material for the production of SRFs for several decades.<sup>11</sup> It can reduce the caking tendency of several fertilizers during storage and transportation and delay the nutrient release with the additional advantage of being secondary plant nutrient required for plant growth.<sup>12</sup> However, crystalline nature of sulfur and exposure to higher temperature in the soil results in the formation of microscopic pores which leads to brittleness and higher friability which causes abrupt release of the nutrients.<sup>13</sup>

Inverse vulcanization, a solvent-free copolymerization process, is a relatively new technique which is being explored to produce sulfur-based polymers that are resistant to depolymerization by producing sustainable polymer chains. This technique can convert sulfur-the main monomer which acts as solvent and initiator in the molten form-into stable polymeric chains by addition of organic crosslinker or co-monomer.<sup>14</sup> The crosslinkers can be sourced from industrial feedstock (such as divinylbenzene (DVB),<sup>15</sup> 1,3-diisopropenylbenzene (DIB),<sup>16</sup> and dicyclopentadiene (DCPD)<sup>10</sup>) or from renewable bio-based sources (such as vegetable oil,<sup>17</sup> limonene,<sup>18</sup> myrcene,<sup>10</sup> and diallyl disulfide<sup>19</sup>). These copolymers have already been utilized in several application which includes Li-S batteries, water purification and CO<sub>2</sub> capturing. Recently Ghumman *et al.*,<sup>13,20,21</sup> utilized inverse vulcanized copolymer prepared from *Jatropha* and rubber seed oil as a coating material for urea to produce SRFs. The produced SRF demonstrated promising results in terms of biodegradability and nutrient release efficiency. However, the unreacted sulfur present in the structure of the copolymer causes the generation of pores which promotes the sharp release of nutrient compromising the nutrient release efficiency of the coated fertilizer. Therefore, it arises a need to produce a copolymer which does not possess any unreacted sulfur particle and can be considered as a sustainable copolymer.<sup>20</sup>

The current research work aims to produce urea as slow-release fertilizer (SRF) by dip-coating of inverse vulcanized copolymer composed of petroleum industry waste sulfur as monomer and citrus industry waste myrcene as bio-based organic cross-linker. Thus prepared, inverse vulcanized copolymer was analyzed by Fourier-Transform Infrared Spectroscopy (FTIR). The morphology of urea SRF was analyzed by means of Scanning Electron Microscopy (SEM) and the nutrient-release

efficiency of nitrogen was investigated in distilled water by using diacetyl monoxime (DAM) calorimetry method.

## 2 Methods

The overall flow of the research is illustrated in Fig. 1.

### 2.1 Materials

Chloroform, diacetyl monoxime, myrcene, elemental sulfur (reagent grade), urea (AR grade), thiosemicarbazide (TSC), phosphoric acid and sulfuric acid were procured from Merck, Malaysia. No material was further purified, they were used as received.

### 2.2 Methodology

**2.2.1 Synthesis of the copolymer (poly(S-M)).** Synthesis procedure for poly(S-M) has been adopted from previous literature on inverse vulcanized copolymers.<sup>22-24</sup> Sulfur (S) was melted to produce the thiyl radicals in a 30 ml glass vial using a thermoset oil at a temperature of 170 °C under continuous stirring.<sup>19</sup> The production of thiyl radical is indicated by the color change of the melted sulfur from yellow to orange. As the liquid turned to orange myrcene was added in the glass in a dropwise manner to avoid the sudden temperature drop. The plaque mixture was then allowed to react for 1 hour, after which the product was allowed to cool down to room temperature.<sup>22,23,25</sup> Afterwards the copolymer was extracted from the glass vial. Initial sulfur ratio was varied to produce two different copolymers, one with 50 wt% S and other with 60 wt% S.

**2.2.2 Characterization of copolymer.** Confirmation of the successful reaction of myrcene and sulfur, and chemical composition of the produced copolymer was analysed using Fourier transform infrared spectroscopy (FTIR) employing PerkinElmer frontier spectrometer. Using attenuated total reflectance (ATR) methods total 8 scans were performed at scanning range of 4000–500 cm<sup>-1</sup> frequency with 4 cm<sup>-1</sup> resolution.

**2.2.3 Coating of urea.** Coated urea was produced using a dip-coating method, for which 4 g of produced copolymer was overnight dissolved in 4 ml of chloroform solvent to obtain homogenous coating solution.<sup>20,21,26</sup> Afterwards, 8 g of urea granules (2–2.5 mm) was added in coating solution and carefully mixed using glass rod to ensure the uniform coating of the urea, followed by overnight drying. Same process was repeated twice or thrice to obtain different coating thickness.

**2.2.4 Morphology of the coated fertilizer.** Zeiss SUPRA 55VP microscope was used to estimate the coating thickness of the coated urea. The coated urea was dipped in liquid nitrogen and cut using a sharp knife to get the cross-section. Followed by the gold coating to the samples and placing in the device chamber sample holder and the gap was set to 4 mm.

**2.2.5 Nitrogen release test in distilled water.** To estimate the nitrogen release from the coated urea, firstly the total available nitrogen content of the coated urea was evaluated using Kjeldahl. The detail of this method can be found elsewhere.<sup>27</sup> Diacetyl monoxime (DAM) calorimetry method was



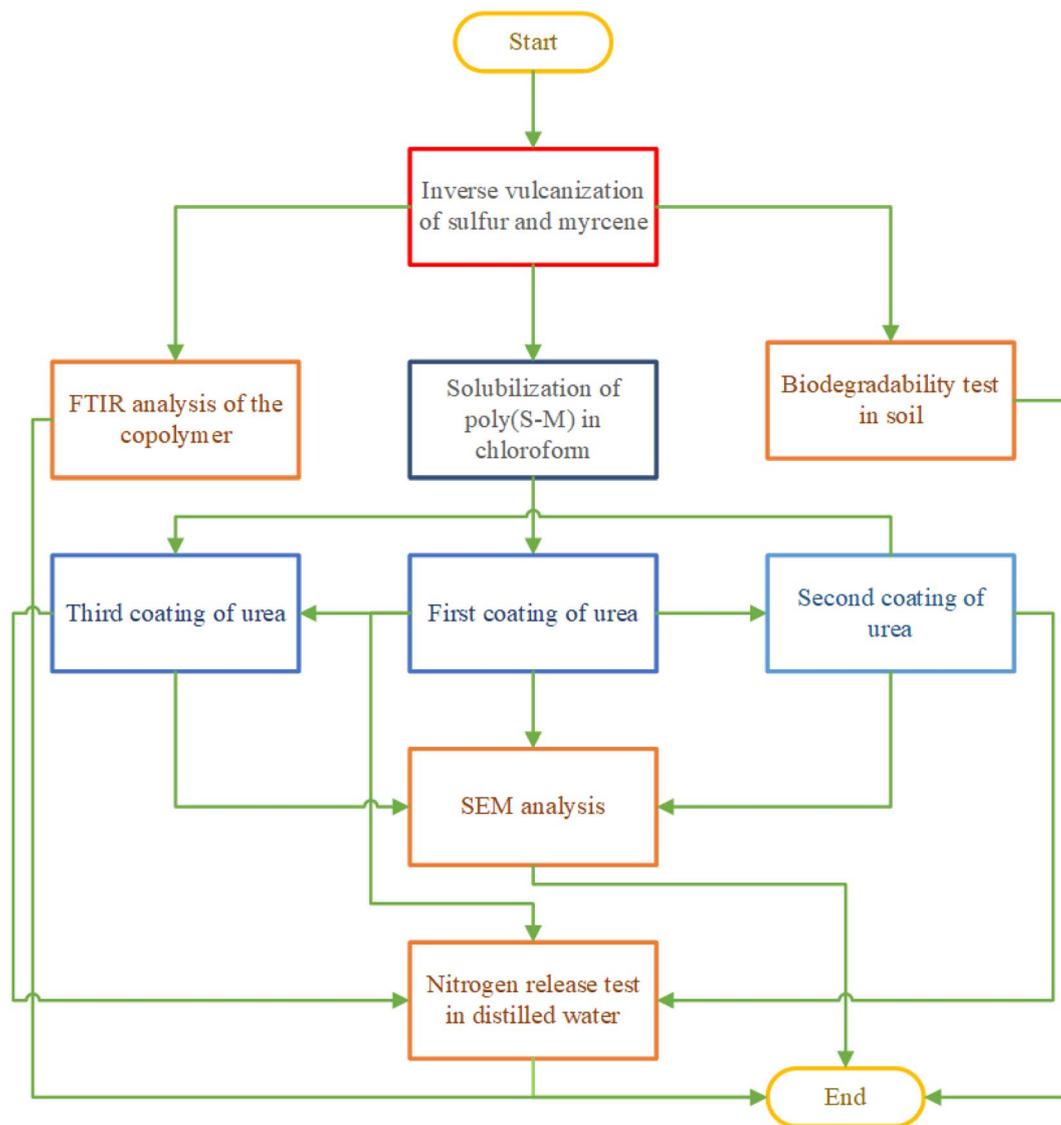


Fig. 1 Overall research flowchart.

used to estimate the amount of nitrogen leached out from the coated urea. This method involves the usage of color reagent, which was prepared by mixing the 25 ml of DAM solution (prepared by dissolving 2.5 g of diacetyl monoxime in 100 ml of distilled water), 15 ml of TSC solution (prepared by dissolving 0.25 g of thiosemicarbazide in 100 ml of distilled water) and 460 ml of acid reagent (prepared by mixing 10 ml of sulfuric acid, 240 ml of distilled water and 250 ml of phosphoric acid).<sup>20,21,28</sup> 2 g of coated urea was placed in a conical flask containing 200 ml of distilled water. To estimate the amount of the nitrogen released in distilled water from coated urea, after every 24 h 2.5 ml of the aliquot was taken and mixed with 7.5 ml of the color reagent and heated to a temperature of 85 °C for 30 min in a water bath followed by cooling at room temperature. The obtained red color solutions were then analysed using Shimadzu UV-1800 UV-Vis spectrophotometer at 526 nm wavelength.<sup>13,20,21</sup>

**2.2.6 Biodegradability test in soil.** For biodegradability test same procedure was followed as explained in our previous reports.<sup>13,20,21</sup>

## 3 Results and discussion

### 3.1 FTIR of the copolymer

Fig. 2 depicts FTIR-ATR spectra of produced copolymers and myrcene (monomer). The spectrum of myrcene shows the band representing double bond at 3090  $\text{cm}^{-1}$  (due to C-H stretching), 1646  $\text{cm}^{-1}$  (due to C-H bending) and at 947  $\text{cm}^{-1}$  and 885  $\text{cm}^{-1}$  (due to the out of plan vibrations). Spectrum of myrcene also showed signals between 2800–3000  $\text{cm}^{-1}$  representing the methyl carbon vibrations for C-H stretching and at 1399  $\text{cm}^{-1}$  for C-H stretching which belongs to  $-\text{CH}_2-$  group. However, the spectra of both copolymers evident the disappearance of the signals



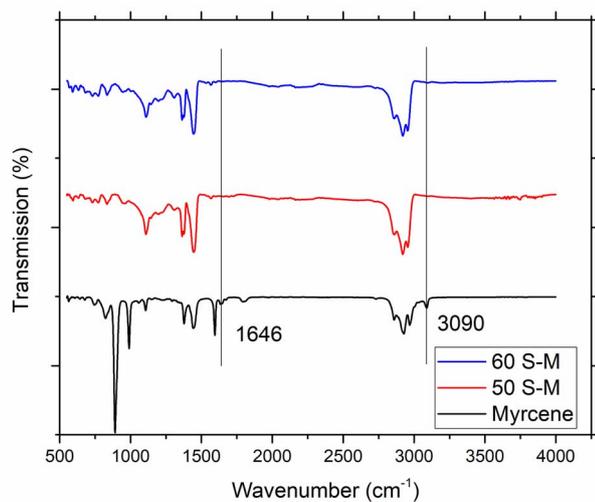


Fig. 2 FTIR-ATR spectra of myrcene, copolymer with 50 wt% S (50 S-M) and 60 wt% S (60 S-M).

representing double bonds (as shown in figure) and no new signal appeared. Findings are similar to the previous reported literature on inverse vulcanized copolymers.<sup>19,29,30</sup> Therefore, it can be concluded that, double bonds have been successfully utilized for C-S bond formation and the formation of copolymer is successful.

### 3.2 Morphology of coated fertilizers

Morphological properties and coating thickness of coated fertilizers were analyzed using scanning electron microscopy (SEM). Obtained SEM micrographs are shown in Fig. 3–6. Two copolymers with different sulfur ratio (*i.e.*, 50 wt% S and 60 wt% S) were utilized to coat the urea. Using 50 wt% S copolymer, three different coating thicknesses were achieved by repeating the same coating process and were named as 50 S-M-I (one coat), 50 S-M-II (two coat) and 50 S-M-III (three coat). In case of 60 wt% S copolymer coating process was not repeated as the adhesion of the copolymer was less compared to 50 wt% S copolymer which did not let the second coat to adhere with the previous coating, the fertilizer coated with this copolymer was named as 60 S-M-I. Fig. 3 depicts the cross-section of the coated fertilizers, in which red circle represents the urea and copolymer coating is marked by red circle.

For the 50 S-M-I coated fertilizer thickness was estimated to be 388.35  $\mu\text{m}$  and it was observed that the coating was consistent, and uniform compared to other coated fertilizers. With repeating the coating process significant increase in coating thickness can be evident from Fig. 5, which increased from 355.35 to 1593.72  $\mu\text{m}$ . However, with this increase the non-uniformity of coating also increased which may lead to abrupt release of the nutrient from the thinner side. The observed non-uniformity of coating can be because of adhering of the coated urea with other. Fig. 6 presents the 50 S-M-III coated fertilizer, from which it can be seen that the coating thickness is not same from all side of the coating. The coating thickness for this sample was estimated to be 1733.5  $\mu\text{m}$  from one side, the other

side observed to have 1585.6  $\mu\text{m}$  thick coating and the third side possesses 849.8  $\mu\text{m}$  thick layer. This observed inconsistency of the coating can cause damage to the integrity of the copolymer coating leading to sudden release of the nitrogen.

### 3.3 Nitrogen release test

Fig. 7 depicts the cumulative nitrogen release from developed coated fertilizers and pristine urea in distilled water at room temperature (The primary data for nitrogen release test is given in Table S1†). Initial nutrient release rate reflects the robustness of the coating film on urea. The consistent, thicker, and meticulous the film is, lesser will be the initial nutrient release rate. The pristine urea released almost 99.6% of its total nutrients in less than 24 h whereas coated fertilizers 50 S-M-I, 50 S-M-II, 50 S-M-III and 60 S-M-I released almost 45% of its total nitrogen in 9 days, 15 days, 17 days, and 8 days, respectively.

As Fig. 7 illustrate that in case of coated fertilizer 50 S-M-III sharp increase in nutrient release at some points which could be because of the inconsistent coating film achieved as shown in Fig. 6, the thickness of this coated ranges from 849–1733  $\mu\text{m}$ . In case of 60 S-M-I initial release rate was high compared to other coated, it may be due to the higher percentage of the sulfur in the copolymer which increases the crystallinity of copolymer resulting in pores generation of the film surface promoting the nitrogen loss. Similarly results were achieved in our previous work.<sup>13,20,21</sup>

### 3.4 Biodegradability test

Fig. 8 depicts the weight loss % of the 50 S-M copolymer in soil (The primary data for biodegradability test is given in Table S2†). Figure illustrates that with the increase of incubation time weight loss increases. The weight loss of the 50 S-M copolymer reaches to 18.25% on the 50<sup>th</sup> day of the incubation, which reveals that the 50 S-M copolymer is decaying steadily in soil but will take more time to get fully decomposed. This slow degradation may be happening because of the sulfur oxidation and/or because of depolymerization of the loosely bonded S-S chain as observed during experimentation. Bacteria called *Aspergillus niger* is present in soil which helps the sulfur to oxidize and convert into sulfate which is a secondary nutrient required by the plants. Thus, adding another benefit in using copolymer as an effective coating as it will not only reduce the nutrient loss but also will provide secondary for plant without producing any soil or water pollution. Further this test establish that this copolymer is biodegradable.

## 4 Comparison of newly developed and other coated urea

To elaborate the properties of prepared coated urea fertilizer the comparison has been made with the other coated urea reported in literature as summarized in Table 1.

Coating materials can be categorized into 3 types namely, (i) polymer-based (natural and synthetic), (ii) hybrid and (iii) sulfur-based coating. Synthetic polymer-based materials have been employed for control release of nutrient and only few of



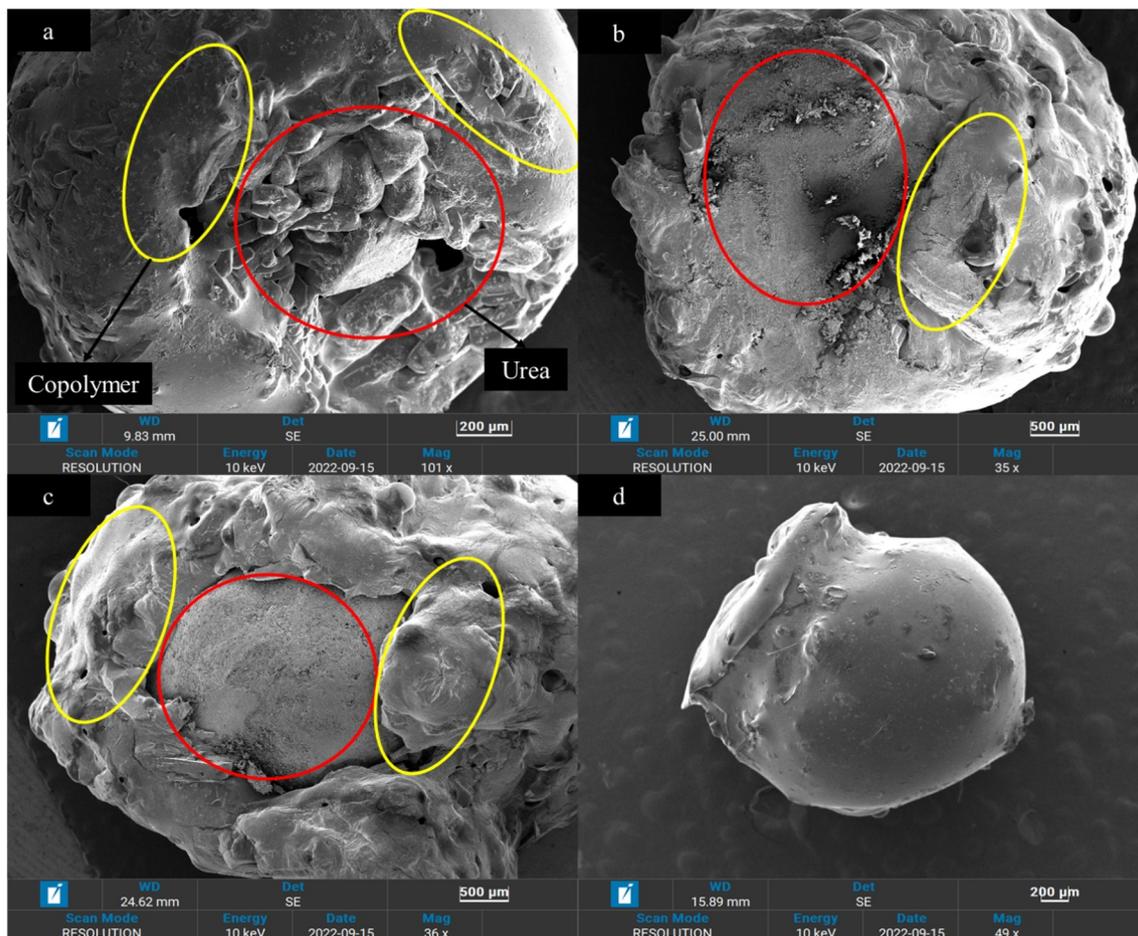


Fig. 3 Cross-section of the coated fertilizers (a) 50 S-M-I, (b) 50 S-M-II, (c) 50 S-M-III and (d) 60 S-M-I.

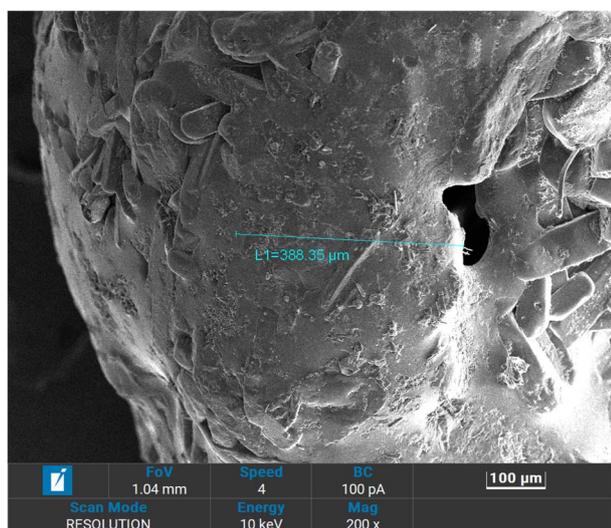


Fig. 4 Thickness of coating for 50 S-M-I.

them showed satisfactory control release rate polyester/urea inclusion complex released 82% N in 1 hour which doesn't meet the European Standard for slow-release fertilizer (>15% in

24 h).<sup>31</sup> In addition, the rate of release of nutrients from synthetic polymer coated urea fertilizer can be difficult to control and may not match the requirements of different crops, leading to inefficient use of fertilizer. Moreover, most reported

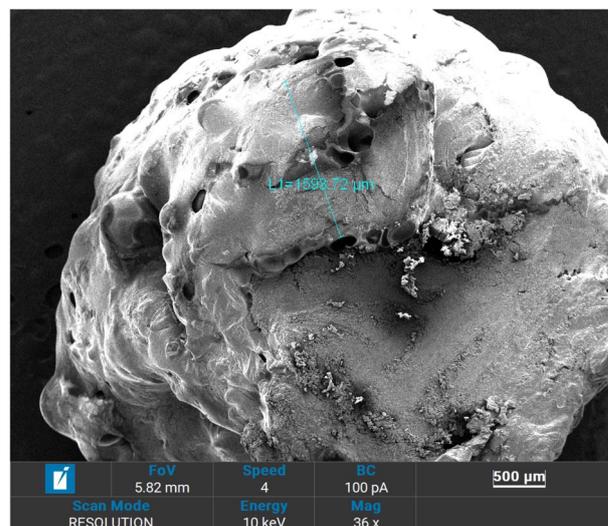


Fig. 5 Thickness of coating for 50 S-M-II.



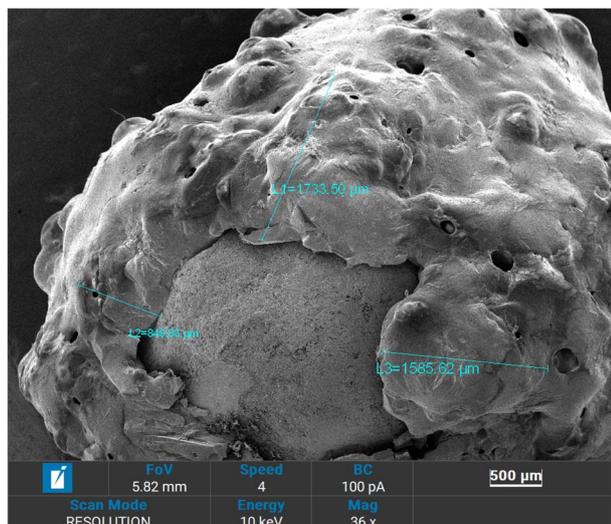


Fig. 6 Thickness of coating for 50 S-M-III.

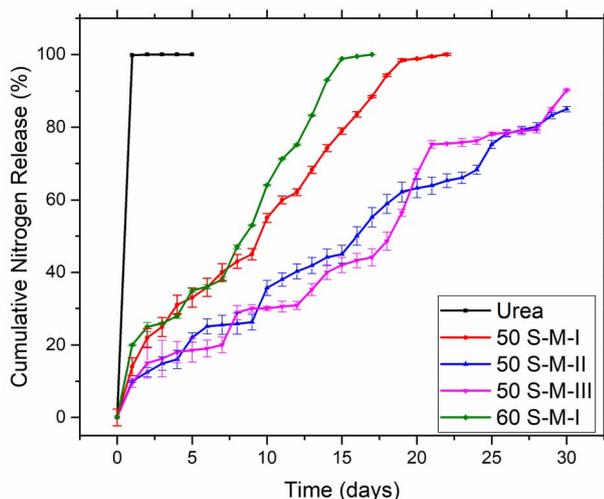


Fig. 7 Cumulative nitrogen release of the developed slow-release fertilizers in distilled water.

such coating materials were not tested for their biodegradable nature which limits their practical applicability. So, synthetic coatings are expensive, non-biodegradable and toxic (due to chemicals involved in synthesis) and not effective for control/slow-release fertilizers.

Biopolymers attracted lot of attention recently because of their non-toxicity and biodegradability but they have inherent shortcoming of being hydrophilic in nature which needs further modification for slow release of nutrients. Gungula *et al.*<sup>32</sup> prepared hydrogels by incorporating urea into *Borassus aethiopicum* starch and *Maesopsis eminii* starch by using borax as binding material which is not ecofriendly. The nitrogen release efficiency (NSE) of prepared hydrogels was also higher than the minimum criteria in first 24 h. Although many natural polymers based slow-release fertilizers have shown excellent nutrient release behavior, the coating method (spray coating in rotary

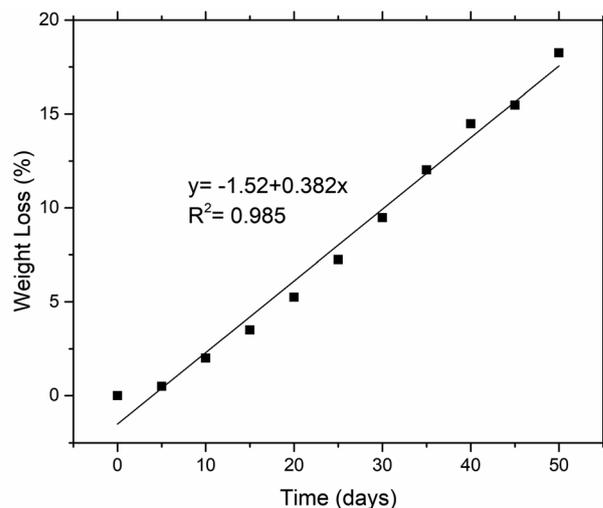


Fig. 8 Weight loss of the copolymer in soil over time.

drum) requires organic solvent which not only enhances the capital and recovery cost of solvent but also poses threat to the environment.<sup>33,34</sup> Another significant point is that there is a lack of accord regarding definition of biopolymer some researchers grafted bio-based polymers into synthetic latex and called it bio polymer which raises serious questions regarding environmental safety when it comes to the practical soil applications.<sup>35</sup>

Researchers have also focused on employing hybrid coating materials (polymer/wax composites, polymer/clay composites and polymer zeolite composites) for urea to improve release of nutrient, reduce leaching of nitrogen and extend fertilization period.<sup>36</sup> However, complexity of production, compatibility of different materials, difficulty in controlling release rate and environment concerns related to disposal are major drawback in this field.

Sulfur was initially employed as coating material, owing to that fact that it is also essential secondary macronutrient for plant growth, for urea granules to synthesize SRF/CRF and results demonstrated slow nutrient release of only 1% in first 24 h.<sup>37</sup> However, sulfur coating was found not to be uniform, even with sealant sub-coating, that can lead to formation of fissures which can cause early nutrient release. The vulcanized castor oil-based polyurethane prepared by one-pot method showed smooth coating surface, but nutrient release was 60% in 7 days which is faster than required. Urea coated with poly(eugenol sulfone) released only 65% N in 30 days, but sulfur oxidation was not improved.<sup>38</sup> Sulfur and soybean oil base coating demonstrated improved sulfur oxidation and uptake as fertilizer, but soybean is an edible oil and not a cost-effective option.<sup>39</sup> Previously, our group has prepared inverse vulcanized copolymers by using *Jatropha* oil and rubber seed oil (both are nonedible) by mechanical mixing and dip-coating method, respectively and demonstrated their potential as SRF.

The now prepared sulfur-myrcene copolymer exhibit excellent N release rate of only 18% in 5 days and enhanced sulfur oxidation potential. Myrcene is by-product of citrus industry, hence it's a biodegradable and sustainable option to





Table 1 Summary of developed slow-release fertilizers

Material	Base fertilizer	Nitrogen release	Biodegradability	Coating method	Reference
<b>Synthetic polymer based</b>					
Polyester/urea inclusion complex	Urea	82% in 1 h	No	Mechanical mixing	31
Polyurethane nanocomposite	Urea	1.9% in 24 h	Not studied	Rotating drum	40
Epoxidized vegetable oil cured by triethylenetetramine	Urea	50% in 120 min	Yes	Fluidized bed coating	6
Polyvinyl alcohol hydrogels	Urea	10% in 4 days	Not studied	<i>In situ</i> polymerization	41
<b>Natural polymer based</b>					
Bio polyurethane	Urea	8% in 24 h	Yes	Rotary drum	3
Lignin-clay modified bio polyurethane	Urea	80% in 30 days 3.6% in 24 h	Yes	Rotary drum	3
<i>Borassus aethiopicum</i> starch (BAS-SRF)	Urea	70% in 30 days 76% in 24 h	Yes	Chemical gelation	32
<i>Maesopsis eminii</i> (ME-SRF) hydrogels	Urea	68% in 96 h	Yes	Chemical gelation	32
Lignin/montmorillonite	Urea	34.6% in 4 days	Yes	Dip coating	34
Chitosan/alginate (Ch-Al) composites	Urea	10% N release in 3 h and 14% in 5 h	Yes	Copolymerization	42
Gelatin/alginate composites	Urea	6% in 3 h and 8% in 5 h	Yes	Copolymerization	42
Modified bentonite/polyurethane composite	Urea	3% in 24 h	Not studied	Rotary drum	33
<b>Hybrid coatings</b>					
Hydrogels of cassava starch-g-polyacrylic acid/natural rubber/polyvinyl alcohol	Urea	38.5% at 30 days	Yes	Rotary drum	36
Starch-g-poly(styrene-co-butylacrylate)	Urea	45% in 60 min	Not studied	Rotary drum	35
<b>Inorganic/S based</b>					
Vulcanized castor oil based polyurethane	Urea	60% in 7 days	Yes, sulfur oxidation not studied	Rotary drum	11
Polysulfone (SO <sub>2</sub> and eugenol based)	Urea	65% in 30 days	Yes	Copolymerization	38
Inverse vulcanized copolymer using <i>Jatropha</i> oil	Urea	13.7% in 24 h	Yes	<i>Ex situ</i> mixing	13
Sulfur and rubber seed oil {poly(S-RSO)}	Urea	17% in 20 days 65% in 43 days	Yes, and promotes sulfur oxidation	Dip coating method	21
Sulfur and myrcene (S-M)	Urea	18% in 5 days 78% in 30 days	Yes, promotes sulfur oxidation	Dip coating method	This study

use in inverse vulcanization process. However, morphology study shows non-uniformity which can lead to sudden release of N. So, there is a need to further investigate it. Researchers need to focus on applicability of these coated fertilizers in real soil conditions where lot of factors can affect the biodegradation and nutrient release rate such as pH, toxicity, and action of microorganisms.

## 5 Conclusions

Novel coated urea was developed using dip-coating method utilizing sustainable copolymer poly(S-M) as a coating material. The SEM images of the coated fertilizer revealed the non-uniform coating which could be because of the adhering coated urea with other particles. Which may pose challenge to its saleable process. The nitrogen release tests the SRF with maximum coating thickness of 1733  $\mu\text{m}$  releases only 16% of its total nitrogen after 4 days of incubation compared to the pristine urea which releases all its nutrient within 1 day. Which shows that the newly developed coated fertilizer perfectly follows the European Standard (EN 1326/2001) as it releases less than 15% nutrient in first 24 h. The soil burial test confirms the biodegradability of the copolymer, as after 50 days of incubation in soil the copolymer losses almost 18.25% of its total weights indicating that the copolymer is degrading. The inconsistency of the coating thickness affects the nutrient release rate which should be addressed in future study.

## Conflicts of interest

The authors declare that they have no conflict of interest.

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