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Gateway toward efficient and miniaturized $A_2B_2O_7$-type fluorite structure-based energy storage devices

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Defect fluorite structure with $A_2B_2O_7$ composition exhibits an intense potential for utilization in modern smart electrical devices. Efficient energy storage with low loss factors like leakage current makes them a prominent candidate for energy storage applications. Here we report a series of the form $Nd_{2-2x}La_{2x}Ce_2O_7$ with $x = 0.0, 0.2, 0.4, 0.6, 0.8$, and 1.0, synthesized via a sol−gel auto-combustion route. The fluorite structure of $Nd_2Ce_2O_7$ is slightly expanded with the incorporation of La without any phase transformation. A gradual replacement of Nd with La causes a decrease in grain size, which increases the surface energy and thus leads to grain agglomeration. The formation of exact composition without any impurity element is confirmed by energy-dispersive X-ray spectra. The polarization versus electric field loops, energy storage efficiency, leakage current, switching charge density, and normalized capacitance, which are considered key features of any ferroelectric material, are comprehensively examined. The highest energy storage efficiency, low leakage current, small switching charge density, and large value of normalized capacitance are observed for pure $Nd_2Ce_2O_7$. This reveals the enormous potential of the fluorite family for efficient energy storage devices. The temperature-dependent magnetic analysis exhibited very low transition temperatures throughout the series.

1. Introduction

The discovery of ferroelectric phenomena in 1920 by Rochelle Salt brought ferroelectric materials to the verge of technological innovation and opened a new door for their applications in many fields like electromechanical systems, memory devices, and radio frequency devices.[1,2] The ferroelectric materials exhibit polarization even in the absence of an applied electric field, known as remanent polarization.[3] In memory devices, positive remanent polarization is treated as 1 and negative as 0 state.[4] Many ferroelectric devices like ferroelectric random-access memory (FERAM) and energy storage devices are designed using materials belonging to pyrochlore, perovskite, and fluorite structures.[5] However, with the discovery of ferro-electricity in fluorite structure in 2011, modern researchers put their attention to using these materials for advanced applications because of their low dielectric constant, simple structure, low leakage current, and small unit cell volume compared to perovskite and pyrochlore materials.[6,7] Ongoing development in

smart electronic devices makes fluorite structure more imperative than others to be utilized for micromechanical and fast-switching applications.[8]

Generally, materials with fluorite structure are considered binary oxides with the general formula AO_2.[9] It is also found in the form of A_2O_4, A_4O_8, $AA'O_4$, $AA'BB'O_8$, $A_2A_2'O_8$, $A_2B_2O_8$, and $A_2BB'O_8$. All these families of fluorite structures belong to the $Fm3m$ space group.[10] Recently, it has been observed that ternary metal oxides such as $A_2B_2O_7$ (ABO) also exhibit a defect-fluorite structure with one oxygen vacancy.[11] The structure of the ABO-type compounds depends upon the cationic radius ratio of B and A-site cations. If the cationic radius ratio is greater than 1.3, it exhibits a pyrochlore structure, while if it is less than 1.3, then it transforms into a fluorite structure.[12] The structure of ABO fluorite compounds looks identical to the pyrochlore structure, with a different number of formula units in a unit cell. In fluorite structure, there are 4 formula units in a unit cell, while in pyrochlore structure, there are 8 formula units in a unit cell. This is the reason the lattice parameters of the ABO fluorite structure are almost half of the lattice parameters of the pyrochlore structure.[13] In addition, the crystallographic visualization also helps to distinguish these two structures based on occupancy of Wyckoff positions by cations and anions in three different ways,[14] (i) the A and B-site cations in pyrochlore structure occupy 16c and 16d sites, respectively while in the fluorite structure both A and B-site cations occupy a single 4a site, (ii) the oxygen anions in pyrochlore structure lie at 48f site

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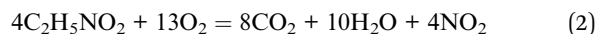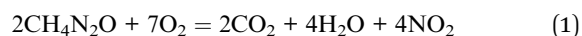
with coordinates $(x, 1/8, 1/8)$ while in fluorite structure, it lies at $8c$ site with coordinates $(1/4, 1/4, 1/4)$, and (iii) the oxygen anion which lies at $8b$ site in pyrochlore structure moves to $4a$ site in fluorite structure.[15] Due to these differences, the A and B-site cations are surrounded by 8 and 6 anions, respectively, in pyrochlore structure which transforms to an average of 7 coordinates in fluorite structure. This structural flexibility of ABO-type materials makes them prominent to be utilized in photoluminescence, anti-erosion applications, thermal barrier coating, magnetic devices, nuclear waste management, electrical device, gas sensors, and high-temperature catalysts.[16] Besides, the high dielectric constant, less energy loss, high power, and large energy density of ABO-based fluorites inspired us to investigate their new compositions for energy storage devices and fast-switching applications.[17]

In the recent past, ferroelectric and dielectric properties of ABO materials grabbed vast interest for their possible applications in smart and efficient electronic devices. For instance, the ferroelectric, dielectric, and piezoelectric properties of rare-earth doped fluorite oxides were studied by Zhou *et al.*, and they reported that the main cause of spontaneous polarization in fluorite structure was the rotation of octahedral oxygen and displacement of A-site ions.[18] Recently, Lee *et al.* reviewed the role of domains and domain dynamics on the ferroelectric properties of fluorite structure and proposed its utilization for ferroelectric field effect transistors due to switching of polarization by nucleation process in the fluorite structure.[19] As temperature plays a vital role in the ferroelectricity of any material, so, Park *et al.* investigated the kinetic and thermodynamic origin of ferroelectricity in fluorite structure and proposed a long-term device performance of the fluorite structure.[20] Meanwhile, as the size of the fluorite structure is almost half of the pyrochlore structure of ABO, therefore, Ali *et al.* explored the fluorite structure as a gateway of smart electronic devices due to its ferroelectric and antiferroelectric properties.[21] Moreover, antiferroelectric/ferroelectric type electrostatic nanocapacitors bearing fluorite structures have also been investigated by the same research group for energy storage applications.[1] The effect of static magnetic field and temperature has also been studied by Quader *et al.* on the dielectric properties of ABO fluorite structure.[22] Hence, a wide range of investigations of fluorite structures, especially in the field of ferroelectric energy storage applications, inspired us to investigate new ABO fluorite structures that can enhance the efficiency of energy storage devices. Therefore, in the present communication, our prime focus is to investigate the effect of La substitution in $Nd_2Ce_2O_7$ defected fluorite structure on its energy storage capabilities without compromising any phase transformation.

2. Experimental

Fluorite structure can be synthesized using various chemical routes, for instance, the hydrothermal method, solid-state method, and sol–gel auto-combustion method. However, here we used sol–gel auto-combustion approach to synthesize a series of $Nd_{2-2x}La_{2x}Ce_2O_7$ with ($x = 0.0, 0.2, 0.4, 0.6, 0.8$, and

1.0) because it is simple to execute, inexpensive, and ensures crystalline phase in a quick time. To proceed, $La(NO_3)_3 \cdot 6H_2O$ [purity > 99%], $Nd(NO_3)_3 \cdot 6H_2O$ [purity > 99%], and $Ce(NO_3)_3 \cdot 6H_2O$ [purity > 99%] were utilized as precursors along with fuel agents such as urea (CH_4N_2O) and glycine ($C_2H_5NO_2$). All the chemicals were purchased from Sigma Aldrich. The metal nitrates to fuel agent's ratio was maintained as 1 : 2. The stoichiometric amounts of all the precursors were weighed using a precise digital balance and dissolved separately into deionized water. The individual transparent solutions were combined into a beaker. The beaker was then placed on the hot plate with a magnetic stirrer inside it. The hotplate temperature was set at 95 °C, and stirred magnetically at 320 rpm. The solution thickened over time as a result of the continuous elimination of fumes. The solution was kept on the hot plate till stirring became difficult. Then, the stirring was stopped, magnetic stirrer was taken out, and in a few moments, frothing started that converted viscous liquid into gel. To eliminate the nitrogenous gases from the gel and to trigger the auto-combustion process, the hot plate's temperature was gradually increased up to 310 °C.[23] At this temperature urea and glycine catch fire and causes auto-combustion. This combustion yields the CO_2, NO_2, and water vapors according to the following balanced chemical equations:

$$2CH_4N_2O + 7O_2 = 2CO_2 + 4H_2O + 4NO_2 \tag{1}$$

$$4C_2H_5NO_2 + 13O_2 = 8CO_2 + 10H_2O + 4NO_2 \tag{2}$$

The hitting of this temperature gave birth to the flame inside the beaker, which instantly burnt all the gel and converted it into ash. As a result of this combustion, the temperature inside the beaker promptly increased which lead to following reaction for parent composition to occur:

$$4Ce(NO_3)_3 + 4Nd(NO_3)_3 = 2Nd_2Ce_2O_7 + 24NO_2 + 5O_2 \tag{3}$$

The molecular oxygen (O_2) and nitrogen dioxide (NO_2) were evolved as a byproduct in this reaction. For subsequent samples, the stoichiometric amount of $La(NO_3)_3 \cdot 6H_2O$ was substituted at $Nd(NO_3)_3 \cdot 6H_2O$ site. After that, the ash was put into an Agate mortar and pestle for grinding and converted into fine powder. The synthesized powder was then placed into ceramic cups and calcined at 800 °C for 3 h in a Nabertherm furnace to develop a pure phase.[24] This calcined powder was pressed using an Apex hydraulic press to make cylindrical pellets of 7 mm diameter and ~1 mm thickness by applying a force of 30 kN. All the pellets were then sintered at 350 °C for 1 h to make them hard.[25] The pictorial representation of this whole synthesis process is shown in Fig. 1.

An advanced Bruker D8 X-ray diffractometer (XRD) was used to analyze the crystalline phase of the synthesized series. A Nova NanoSEM-450 field emission scanning electron microscope (FESEM) was utilized to investigate the morphology and elemental composition. A Radiant's Technologies Inc., USA precision multiferroic tester was used to probe ferroelectric properties. Magnetic properties of synthesized series have been
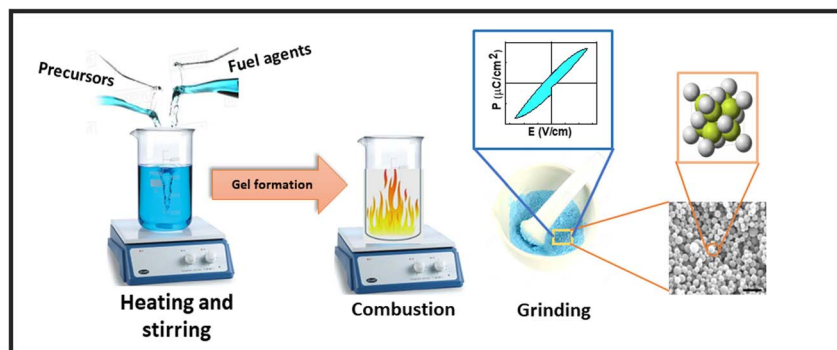
Fig. 1 Schematic illustration of sample synthesis.

carried out by Cryogenic vibrating samples magnetometer (VSM).

3. Results and discussion

The crystal structure is one of the key parameters which helps to control a material's characteristics. In the present communication, the crystal structure was analyzed using Cu Kα radiations with wavelength 1.54 Å obtained at 40 kV. The indexed XRD patterns of $Nd_{2-2x}La_{2x}Ce_2O_7$ ($x = 0.0, 0.2, 0.4, 0.6, 0.8,$ and 1.0) in the 2θ range of 20–80° are shown in Fig. 2. The patterns were indexed using the analytical approach as described by B. D. Cullity.[26] In the XRD pattern of pure $Nd_2Ce_2O_7$ (NCO), the diffracted intensity peaks were observed at 2θ values of 28.3°, 32.6°, 46.6°, 55.1°, 57.7°, 67.6° and 74.7° which were identified

as emerging from (111), (002), (022), (113), (222), (400), and (133) planes, respectively and were well matched with inorganic crystal structure database (ICSD) reference card #01-075-0156 exhibiting a fluorite structure.[27] No impurity peaks were seen that confirmed the formation of a pure fluorite phase of parent NCO.[28,29] When La was introduced at Nd-site, it did not give birth to any new peak; however, the peaks already present in the diffraction patterns slightly moved toward lower 2θ values, which predicted the stability and expansion of the fluorite unit cell with the substitution of La. This unit cell expansion continued as we proceeded further in the series. This shifting of peaks is quite significant when we look at the diffracted pattern in the narrow 2θ range, as presented in Fig. 2(b).[30] The lattice constants $(a = \lambda\sqrt{h^2 + k^2 + l^2}/2\sin\theta)$ calculated for all the
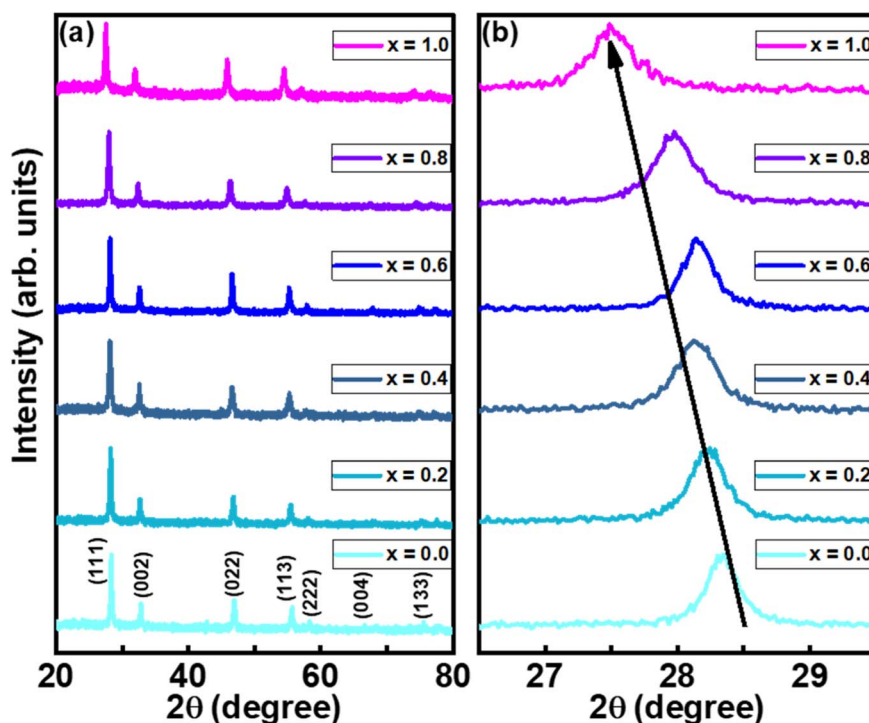

Fig. 2 (a) XRD patterns of $Nd_{2-2x}La_{2x}Ce_2O_7$ ($x = 0, 0.2, 0.4, 0.6, 0.8,$ and 1.0), and (b) shifting of (111) peak towards smaller 2θ values with substitution contents.
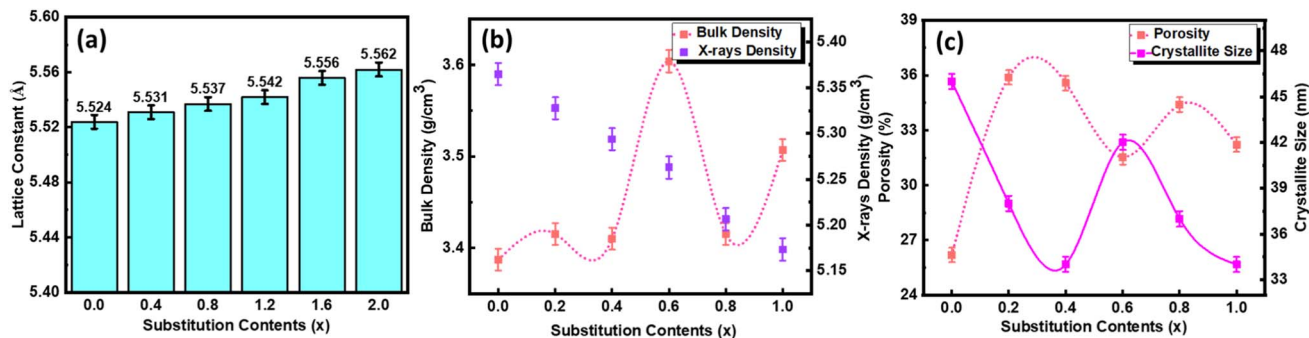
Fig. 3 (a) Variation in lattice constant, (b) bulk and X-rays densities and, (c) porosity and crystallite size as a function of substitution contents.

samples are plotted in Fig. 3(a), which reveal that lattice parameter (a) is slightly increased up to 5.56 Å from 5.54 Å by substituting La at the Nd-site. This increase in lattice parameter is attributed to the minutely larger ionic radius of La^{3+} (1.160 Å) as compared to Nd^{3+} (1.109 Å).[31] The calculation of lattice parameters leads to calculating the unit cell volume, X-ray density, and porosity of the samples. However, the calculation of porosity requires pre-calculation of bulk density which was determined by measuring the mass and volume of samples in the pellet form. In addition to all these parameters, the extent of crystallization is mapped by calculating the crystallite size using Scherrer's formula. The computed values of all these parameters are presented in Table 1, where the variation in various physical parameters like bulk and X-ray densities, crystallite size, and porosity with substitution contents of La is presented in Fig. 3.

To evaluate the reliability of structural data, the Rietveld refinement of all XRD patterns was carried out using Profex software. The working this software involves the matching of diffracted data of the sample with the standard data in terms of peak intensity, peak position, full width at half maxima, and generating the difference plot. In addition, it generates the value of lattice parameters, unit cell volume, atomic positions, Wyckoff positions, different profile factors, and goodness of fit.[32] The Rietveld refinement patterns are presented in Fig. 4, where it has been observed that all peaks were well matched with the reference pattern. The formation of the pure phase for all the compositions was confirmed because no considerable distortion appeared in any sample. The green line in the figure depicts the difference plot between the calculated and observed data where a minute change in intensity of peaks appeared,

which may be due to some sort of defective state or vacancies in samples. This refinement was quantified in terms of profile factors like R-profile, R-expected, R-Bragg, weighted R-profile, and goodness of fit (χ), as shown in Table 2. The numerical value of χ is an ultimate parameter that determines the reliability of Rietveld refinement. The value of χ for all samples was less than 2. In the present communication, the smaller value of χ confirmed the formation of the pure desired crystalline phase for all the samples. The unit cell of pure LCO and NCO are presented in Fig. 5.

FESEM images of all the samples were taken at 200 000× magnification which helped to determine the shape and size of particles, distribution of particles, surface texture, and voids present on the surface of samples. The homogenously distributed round-shaped grains with relatively small porosity can be observed for the pure NCO sample. The grain size of samples was estimated using a Java-based ImageJ software, and the average grain size for pure NCO was recorded as approximately 27 nm. This grain size was increased when 20% of La was substituted at the Nd-site, and a further increase in the concentration of La caused a decrease in the grain size. This reduction of grain size increases the surface energy of the particles, which enhances their binding and thus leads to the agglomeration of particles; this is why we observe agglomeration in the form of clusters containing smaller grains as compared with the parent composition.[33] In Fig. 6, FESEM images of all the samples with histogram of particle size distribution as inset, are presented.

EDX spectra of all the samples are also presented in Fig. 6, where it can be noted that the EDX spectrum of pure NCO contains peaks related to Nd, O, Ce, Au, and C, which confirms

Table 1 Lattice constant, crystallite size, bulk density, X-rays density, and porosity of the samples

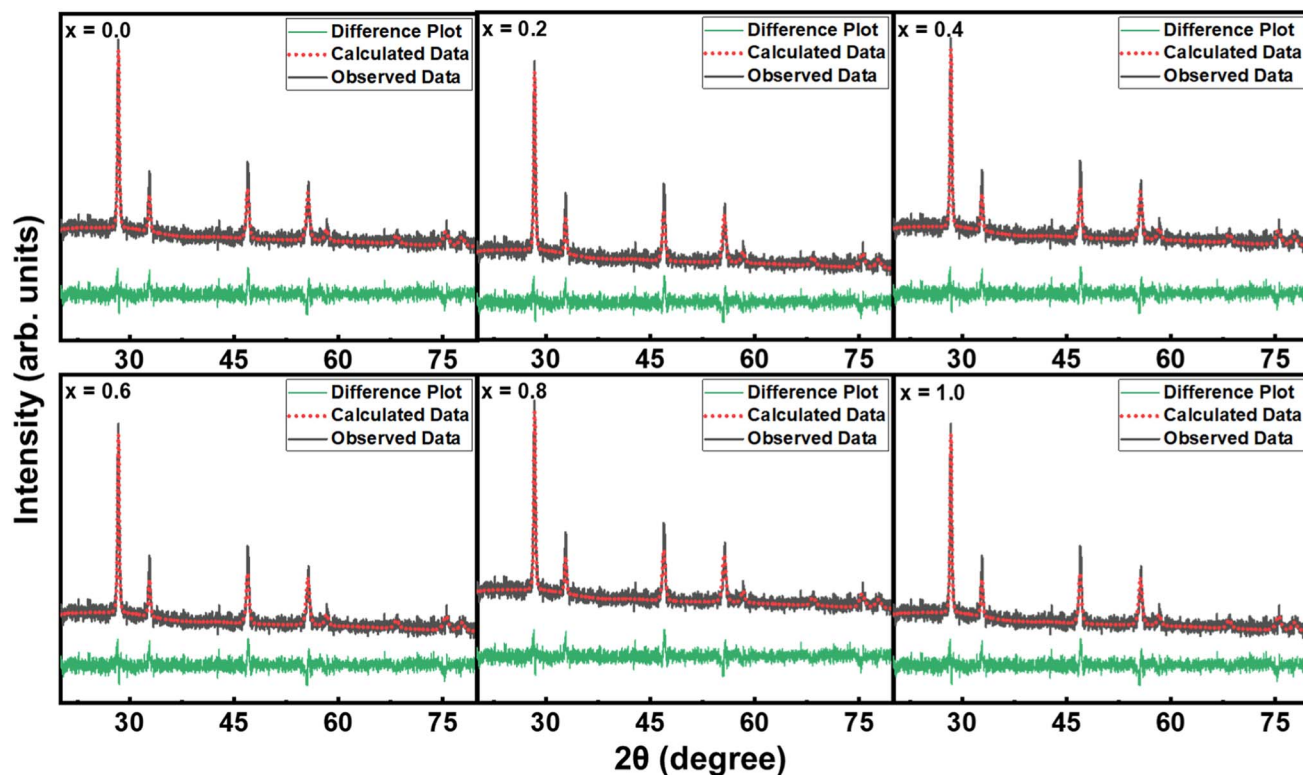| $Nd_{2-2x}La_{2x}Ce_2O_7$, $x=$	Lattice constant (Å)	Crystallite size (nm)	Bulk density (g cm^{-3})	X-rays density (g cm^{-3})	Porosity (%)
0.0	5.524	46	3.39	5.36	26.18
0.2	5.531	38	3.41	5.32	35.89
0.4	5.537	34	3.41	5.29	35.58
0.6	5.542	42	3.60	5.26	31.52
0.8	5.556	37	3.41	5.21	34.40
1.0	5.562	34	3.51	5.17	32.20

Fig. 4 Rietveld's refined XRD patterns of $Nd_{2-2x}La_{2x}Ce_2O_7$ analyzed using Profex software.

that there is no impurity element present in the sample. The presence of peaks belonging to C and Au did not wonder because the C peak came from the sticking tape used to hold the sample with a stub, whereas the Au peak is because of the coating of samples with a thin gold film to get better contrast.[34] The peaks belonging to La emerged when La was substituted at the Nd site, and the intensity of these peaks increased gradually, corresponding to the concentration of La. For pure LCO, when the whole Nd has been replaced with La, no peak corresponding to La appeared, which assured the complete replacement of Nd with La.

In the modern era, ferroelectric materials are considered as most promising materials which can bring a revolution in the energy storage world.[35] The intrinsic property of ferroelectricity of any material can be realized from its polarization *versus*
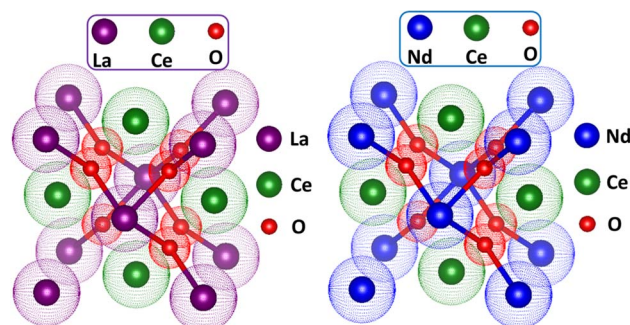

Fig. 5 Crystal structures of pure LCO and NCO.

Table 2 Data obtained through Rietveld refinement of the samples

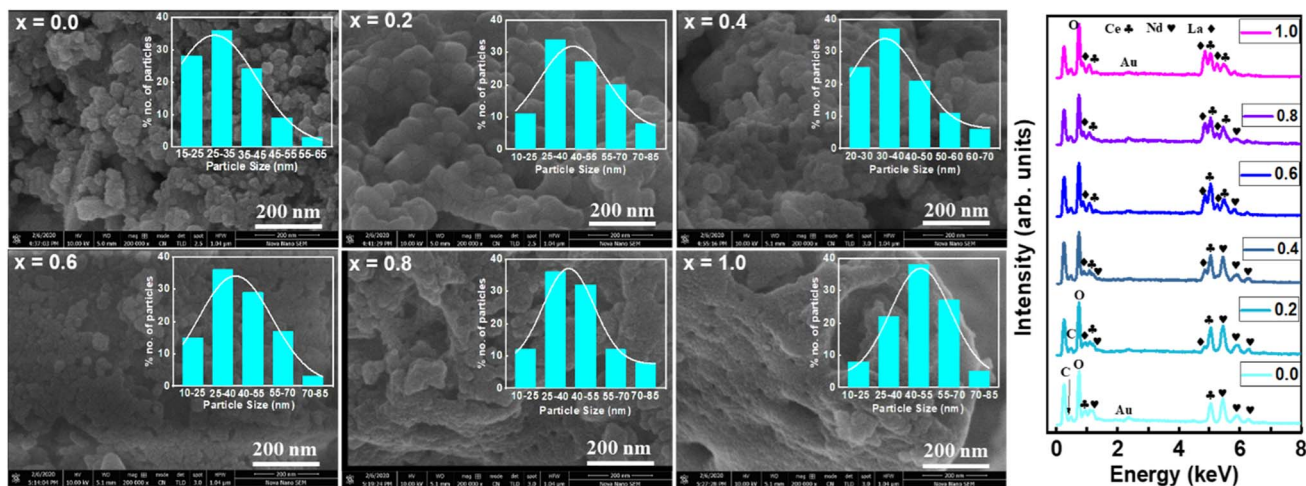| Sample $Nd_{2-x}La_{2-2x}Ce_2O_7$	$x = 0.0$	$x = 0.2$	$x = 0.4$	$x = 0.6$	$x = 0.8$	$x = 1.0$
Cell parameters						
$a = b = c$ (Å)	5.524	5.531	5.537	5.542	5.556	5.562
Volume (Å)^3	168.56	169.20	169.75	170.21	171.51	172.06
$\alpha = \beta = \gamma$	90	90	90	90	90	90
R-Factors (%)						
R_{exp}	13.63	13.04	12.14	13.21	12.55	12.68
R_{wp}	15.06	16.68	15.85	15.65	19.30	16.41
χ^2	1.22	1.6362	1.705	1.403	2.365	1.674
GoF (χ)	1.105	1.279	1.306	1.185	1.537	1.294

Fig. 6 FESEM images and EDX spectra of $Nd_{2-2x}La_{2x}Ce_2O_7$ (x = 0, 0.2, 0.4, 0.6, 0.8, and 1.0) samples.

electric field (PE) plots. For that purpose, PE loops of all the samples were taken up to 800 V cm^{-1} electric field strength. The unsaturated PE loops for all the samples are presented in Fig. 7(a). It has been observed that polarization increased by increasing the applied electric field and reached 0.0017 μC cm^{-2} for pure NCO. When the strength of the applied electric field is reduced, then polarization also decreases, but some extent of polarization remains within the material even after the complete removal of the external electric field; that polarization is known as remanence polarization (P_r).[36] For pure NCO, the value of P_r was 0.00043 μC cm^{-2}. The value of maximum polarization (P_m) and P_r decreased gradually by increasing the

substitution of La up to 60% and then started increasing. Among all these compositions, the pure NCO exhibited the highest value of P_m. The variation of P_m and P_r as a function of substitution contents can be seen in Fig. 7(b). The decrease in P_m with increasing substitution may be due to the smaller atomic polarization of La than Nd.[37] Similar decreasing behavior with increasing substitution is also witnessed in P_r and H_c up to 60%; after that, it increased with the increase of substitution. The decrease in polarization in intermediate compositions may be due to the presence of both La and Nd atoms at the A-site in the sample. The presence of a single element at A-site ensures the better coupling among the electric
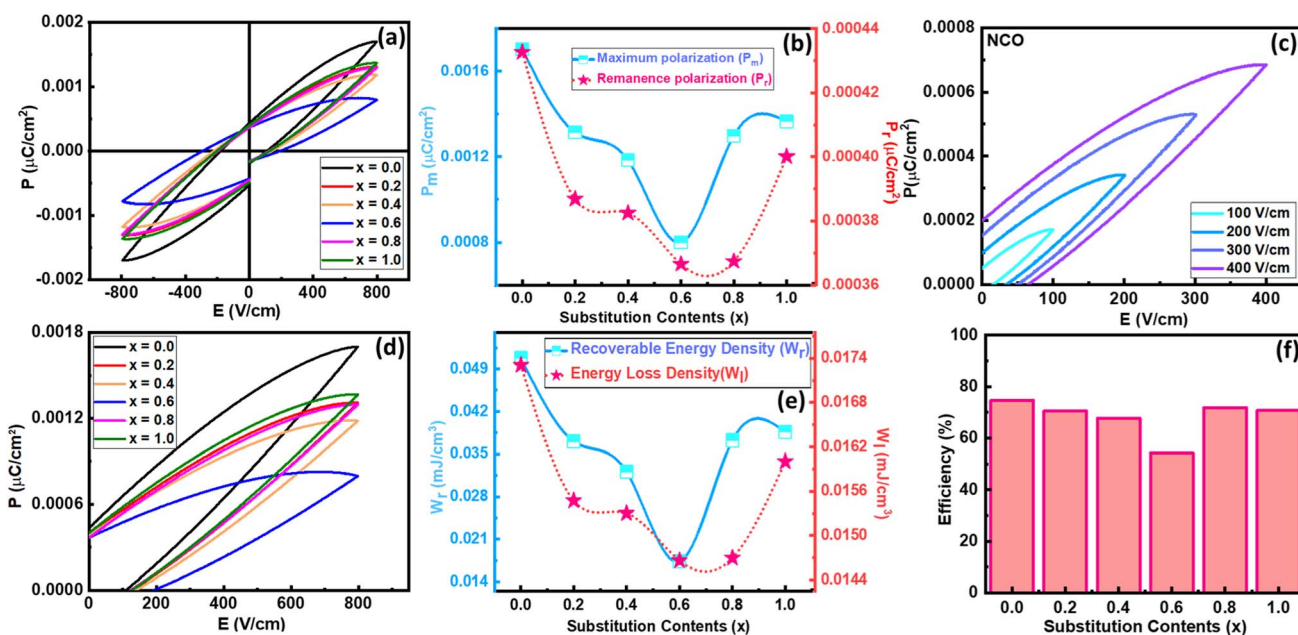

Fig. 7 (a) PE loops of all the samples *versus* applied electric field, (b) variation in maximum and remanence polarization, (c) charging and discharging curves of NCO, (d) unipolar PE loops of all compositions, (e) variation in recoverable and energy loss densities as a function of substitution contents, and (f) efficiency of all samples.
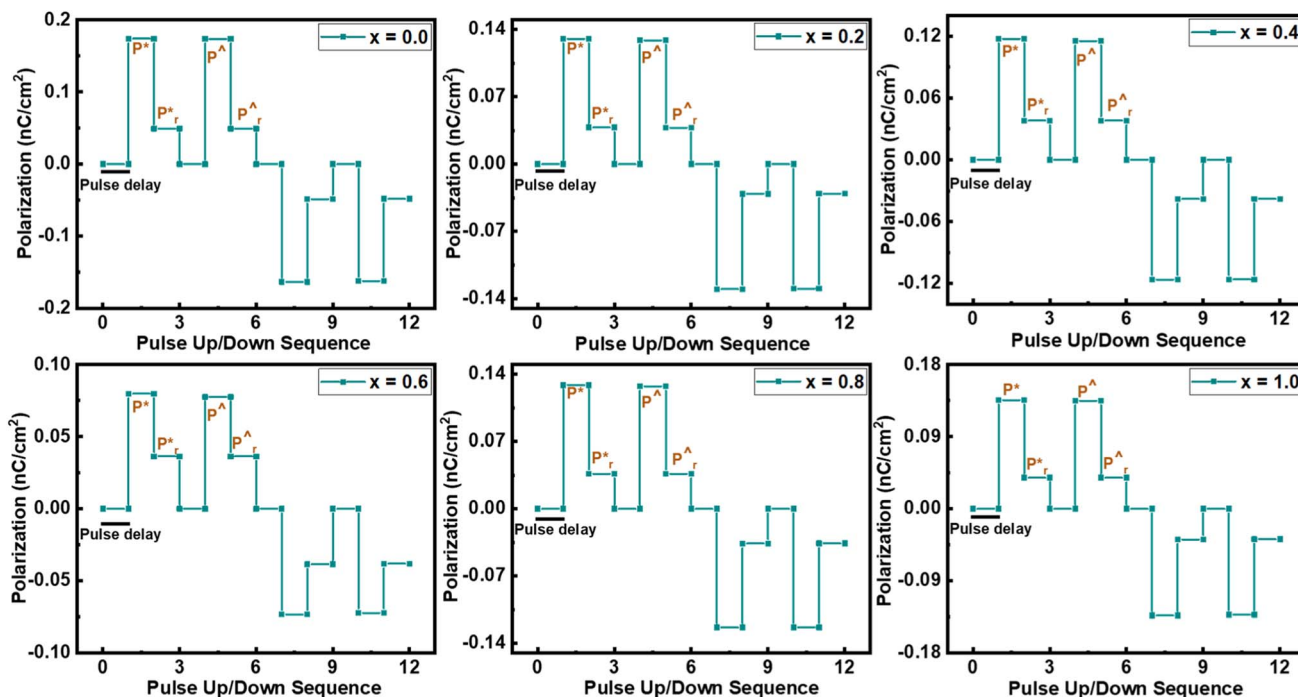
Fig. 8 Positive up and negative down pulse sequence of $Nd_{2-2x}La_{2x}Ce_2O_7$ samples.

dipole moments, due to which pure LCO and NCO exhibit better ferroelectric properties. Since pure NCO revealed the maximum polarization in the given field strength, therefore, its polarization response was also recorded at different voltages to analyze the effect of voltage on electric polarization. The unipolar loops of NCO at different field strengths and loops of all compositions are presented in Fig. 7(c) and (d). The PE loops also give information about energy densities like total energy density (W_T), energy loss density (W_L), and recoverable energy density (W_R) which can be calculated using the following equations:

$$W_T = \frac{1}{2}P_{max}E \quad (4)$$

$$W_R = \int_{P_r}^{P_m} E\,dp \quad (5)$$

$$W_L = W_T - W_R \quad (6)$$

here E is the applied electric field and dp is the difference of P_m and P_r. The area of the loop gives information about W_L, while the area bounded by the discharge curve and polarization axis give information about W_R.[38] The sum of W_R and W_L yields the value of W_T. The variation in the value of W_R and W_L as a function of substitution contents can be seen in Fig. 7(e). Where the maximum value of W_R and W_L were recorded for pure NCO samples. From these values, energy storage efficiency (η) was calculated, and the maximum efficiency of 74% was exhibited by pure NCO, while pure LCO exhibited 71% efficiency, which can be seen in Fig. 7(f).

To sense the exact value of P_m and P_r, the up and down pulses are applied to the material. For the first time, Scott reported this
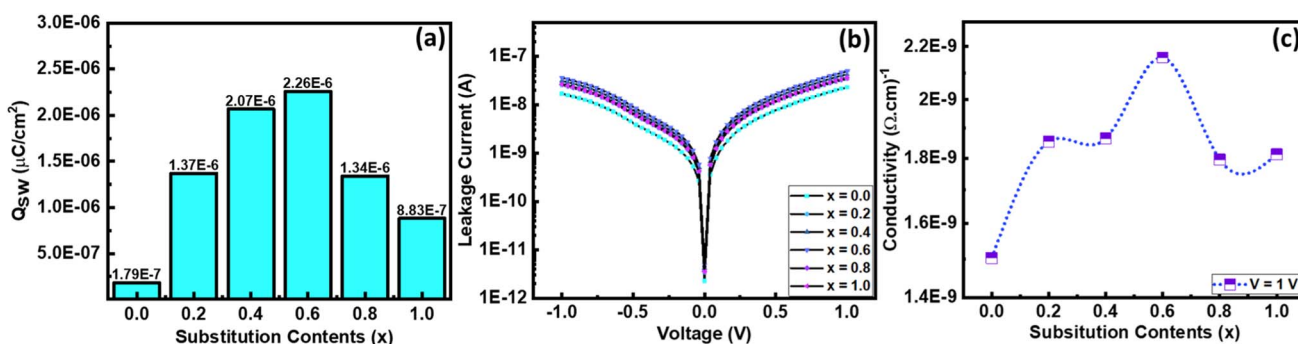

Fig. 9 (a) Switching charge density as a function of substitution contents, (b) $I–V$ curves of all the samples at 1 V, and (c) conductivity of all the samples as a function of substitution contents.
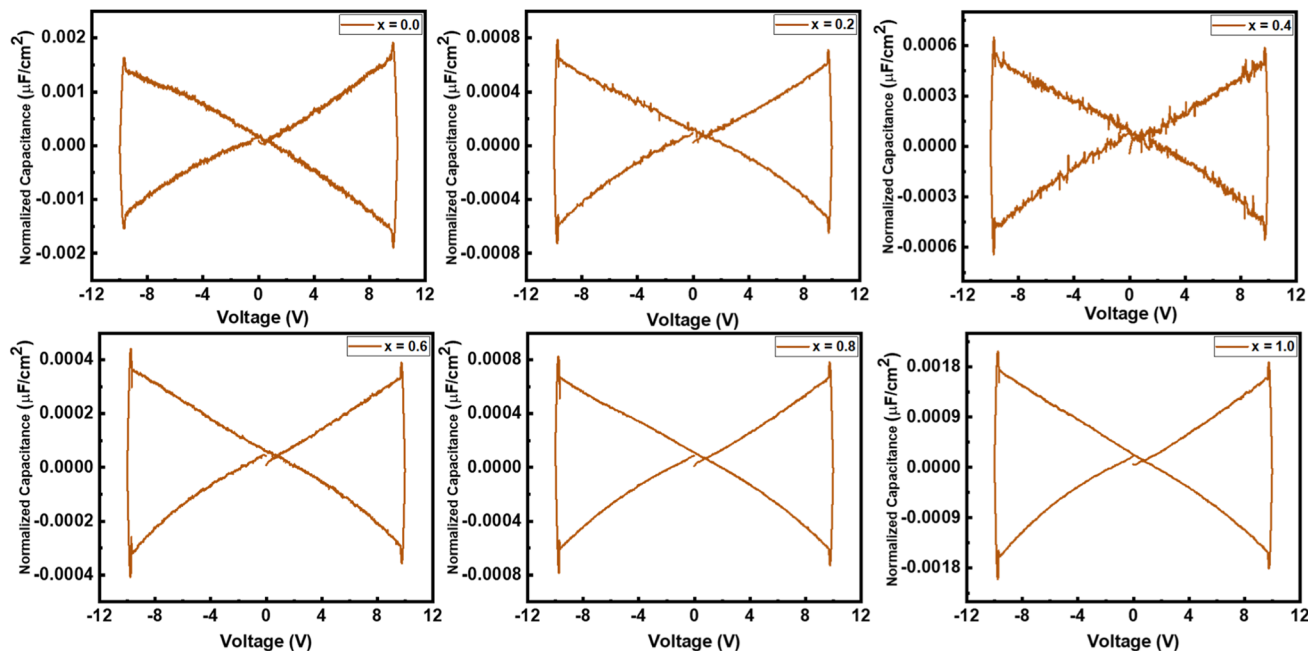
Fig. 10 Normalized capacitance of $Nd_{2-2x}La_{2x}Ce_2O_7$ ($x = 0, 0.2, 0.4, 0.6, 0.8,$ and 1.0) samples.

method and this measurement is known as the positive up and negative down (PUND) measurement.[39] In the present communication, all the samples are examined at 100 V cm^{-1} electric field strength along 1 ms pulse width, keeping 1 s pulse delay, and pictorial representation is provided in Fig. 8. Here, total polarization is the combination of leakage polarization as well as switching polarization. In the first pulse, both P_m and P_r are recorded, including switching charge density (Q_{sw}).[40] When a second pulse is applied, then both polarizations are recorded, excluding switching charge density.[41] The same behavior of polarization was also recorded by applying a reverse electric field. From this measurement, the value of switching charge density is calculated by taking the difference in maximum polarization of the second pulse from the first pulse.[42] The variation in the value of Q_{sw} concerning substitution contents can be seen in Fig. 9(a), which revealed that pure NCO exhibited a minimum value of Q_{sw}.

The PE loops of all samples are not completely closed; there is some sort of breakage that occurred at zero external applied electric field along the negative polarization axis. This gap in PE loops occurs due to the leakage current.[43] To examine the role of leakage current, a graph between current and voltage was recorded using a precision multiferroic tester. Fig. 9(b) shows the trend of leakage current for all the compositions with both positive and negative applied fields. However, the trend in leakage current seems similar but not perfectly identical. An increase in leakage current was noted by increasing the substitution contents of La, and the maximum value of leakage current was recorded for composition with 60% La substitution. After that, it decreased gradually with further increasing substitution of La. From these IV graphs, the value of conductivity (σ) was also calculated[44] and presented as a function of substitution contents in Fig. 9(c), and its maximum value was reported for 60% La substitution.
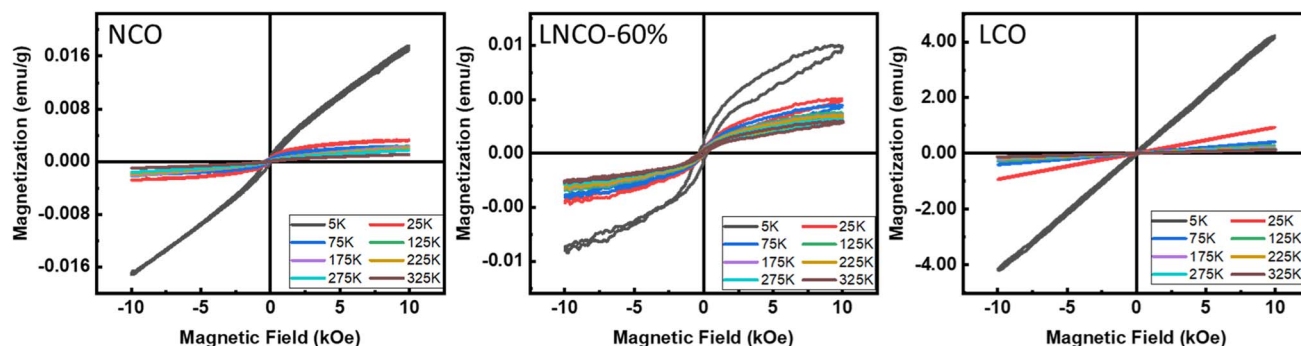

Fig. 11 Temperature dependent magnetic hysteresis loops of NCO, LNCO-60%, and LCO.

A multiferroic tester was used to find the normalized capacitance against applied voltage for all the samples. The normalized capacitance is the linear function of measured capacitance. The polarization of the samples was recorded at 10 V in the presence of the magnetic field. The value of normalized capacitance was calculated by taking the derivate of polarization with voltage. Fig. 10 shows the plot of normalized capacitance against the voltage for all the samples. A decreasing trend in normalized capacitance can be noted by increasing substitution contents up to 60%. As pure NCO showed the highest efficiency, in a similar way, it also showed the largest value of normalized capacitance among others.

To study the temperature-dependent magnetic properties of these materials, we selected three compositions, i.e., pure NCO, 60% La substituted NCO, and pure LCO, and Cryogenic VSM was used to record the magnetic hysteresis loops, which were presented in Fig. 11. The MH loops at various cryogenic temperatures, i.e., 5, 25, 75, 125, 175, 225, 275, and 325 K, were recorded by applying a magnetic field of strength 1 T. For pure NCO, it can be noted that at room temperature, it exhibits weak ferrimagnetic behavior that is due to the weak magnetic response of the fluorite structure.[45] When the temperature is reduced, the saturated weak ferrimagnetic properties of pure NCO are improved and converted into unsaturated MH loops. The maximum magnetization increased gradually by reducing the temperature. Maximum magnetization of 0.017 emu g^{-1} recorded at 5 K temperature for pure NCO. An abrupt shift of the saturated MH loop to the unsaturated MH loop can be seen after reducing the temperature from 25 K. By substituting La at Nd site, it can be noted that the behavior of temperature on magnetization almost remains the same. 60% La substituted NCO sample exhibited ferrimagnetic nature for all the applied temperatures, but net magnetization reduced as compared to pure NCO. At 5 K temperature, La substituted sample exhibits a maximum magnetization of 0.0054 emu g^{-1}. When complete La was replaced at Nd site, and an interesting magnetic response was recorded. It can be noted that for pure LCO sample, at room temperature it exhibits paramagnetic behavior with very less magnetization. However, when the temperature reduces gradually, its magnetic response starts improving, and maximum magnetization increased by keeping samples in less temperature. A gradual improvement in magnetization was recorded up to 25 K temperature. However, by further reducing the temperature, an unsaturated ferrimagnetic response was recorded with a very small remanence magnetization and coercive field. The maximum value of magnetization of pure LCO was recorded as 4.17 emu g^{-1} at 10 kOe.

4. Conclusion

To analyze the structural, morphological, electrical, and magnetic response of ABO-defected fluorite structure, a series of Nd$_{2-2x}$La$_{2x}$Ce$_2$O$_7$ along x = 0.0, 0.2, 0.4, 0.6, 0.8, and 1.0 was synthesized via cost effective sol–gel auto-combustion process. The face-centered cubic crystal structure with slight variation in lattice parameter from 5.54–5.56 Å was noted due to the substitution of La with a larger ionic radius at the Nd site in the parent composition. The crystallite size of the samples varied in the range of 34–46 nm, while X-ray density and bulk density varied in the range of 5.15–5.41 g cm^{-3} and 3.4–3.6 g cm^{-3}, respectively. The Rietveld refinement process further validated the structural reliability in terms of various profile factors and goodness of fit. The average grain size of the parent composition was recorded as 26 nm, and after the substitution, the agglomeration started due to a decrease in grain size. The PE loops revealed that polarization increased by increasing the applied electric field, and pure samples like pure NCO and LCO exhibited good ferroelectric response. The substituted samples in which both La and Nd atoms are present at the A-site exhibited a small decrease in ferroelectric behavior. The maximum value of recoverable density was noted for pure NCO and LCO. Among all these compositions, pure NCO exhibited the best energy storage efficiency of 74%. In PUND measurement, NCO exhibited the lowest value of switching charge density. The leakage current versus voltage graphs uncovered the increase in leakage current with an increase in substitution up to 60%, which then decreased by further increasing substitution. The conductivity was also calculated from I–V characteristics that showed an increase in conductivity by increasing substitution. Normalized capacitance showed maximum value for pure NCO. The temperature depended on the magnetic properties of pure NCO, and 60% La substituted NCO revealed that all the samples showed an improved magnetic response by decreasing the temperature. At room temperature pure NCO and 60% La substituted NCO exhibited ferrimagnetic nature while pure LCO showed paramagnetic nature. However, all the samples show good unsaturated ferrimagnetic nature at 5 K temperature. The maximum magnetization of 4.17 emu g^{-1} was recorded for pure LCO samples at 10 kOe keeping sample at 5 K temperature. Less leakage current, efficient energy storage capabilities, and high normalized capacitance with a lower switching charge density of pure NCO make it a potential candidate for energy storage and fast switching devices.

Conflicts of interest

There are no conflicts of interest to declare for this submission.

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