



Cite this: *RSC Adv.*, 2023, 13, 4803

Design strategies of Pt-based electrocatalysts and tolerance strategies in fuel cells: a review

Wenlei Luo,^a Yitian Jiang,^b Mengwei Wang,^b Dan Lu,^b Xiaohui Sun^b
and Huahui Zhang *^b

As highly efficient conversion devices, proton-exchange-membrane fuel cells (PEMFCs) can directly convert chemical energy to electrical energy with high efficiencies and lower or even zero emissions compared to combustion engines. However, the practical applications of PEMFCs have been seriously hindered by the intermediates (especially CO) poisoning of anodic Pt catalysts. Hence, how to improve the CO tolerance of the needed Pt catalysts and reveal their anti-CO poisoning mechanism are the key points to developing novel anti-toxic Pt-based electrocatalysts. To date, two main strategies have received increasing attention in improving the CO tolerance of Pt-based electrocatalysts, including alloying Pt with a second element and fabricating composites with geometry and interface engineering. Herein, we will first discuss the latest developments of Pt-based alloys and their anti-CO poisoning mechanism. Subsequently, a detailed description of Pt-based composites with enhanced CO tolerance by utilizing the synergistic effect between Pt and carriers is introduced. Finally, a brief perspective and new insights on the design of Pt-based electrocatalysts to inhibit CO poisoning in PEMFCs are also presented.

Received 1st December 2022

Accepted 29th January 2023

DOI: 10.1039/d2ra07644f

rsc.li/rsc-advances

1. Introduction

The increasing environmental issues and energy crisis severely threaten sustainable human society due to traditional energy consumption and declining fossil fuels. How to develop efficient storage technologies or conversion devices that can fully utilize sustainable and clean energy has drawn extensive attention.^{1,2}

Based on the guidance, many efforts have focused on developing highly efficient conversion devices, *e.g.*, fuel cells, which can effectively reduce traditional fossil energy consumption and its emissions. Proton exchange membrane fuel cells (PEMFCs) as highly efficient conversion devices have received more attention due to their high-power generation efficiency and environmentally sustainable regeneration.^{3–5} In PEMFCs, fuel molecules and oxidants can be converted into ions by the utilization of catalysts and migrated continuously in the electrolyte, thus continuously converting chemical energy into electrical energy. Despite the advantages, the anodizing process is very complicated, which involves the sluggish kinetics and formation of intermediates (especially CO) that can occupy the active sites of electrocatalysts,

^aNational Innovation Institute of Defense Technology, Academy of Military Science, Beijing 100071, China

^bState Key Laboratory of Space Power-sources Technology, Shanghai Institute of Space Power-Sources, 2965 Dongchuan Road, Shanghai 200245, China. E-mail: hhzhang2022@126.com



Luo Wenlei, graduated from the National University of Defense Technology with a master's degree, majored in aerospace science and technology. He received his doctor's degree in chemistry and process engineering from University of Leeds, the United Kingdom in 2015. He is a researcher in the energy and power. He is currently a member of research group "organic liquid hydrogen storage fuel cell", led the research on fuel cell technology.



Yitian Jiang is a senior engineer at Shanghai Institute of Space Power-sources. He received his master degree from Harbin Institute of Technology in 2012. His current research is focused on the construction and application of electrocatalysts for fuel cells.



thus leading to decreasing its catalytic performance. Therefore, it is urgently needed to design high-efficiency electrocatalysts for the complicated anodizing process.^{6–8}

As a state-of-the-art catalyst, platinum (Pt) can absorb many charged species and then accelerate the formation of active species on its surface, resulting in the rapid catalytic reaction. As a common Pt-based electrocatalyst, commercial Pt/C refers to commercial carbon supported Pt nanoparticles (NPs). To date, commercial Pt/C catalyst has been regarded as one of the most effective catalysts in PEMFCs, while its limited activity and poor stability are very difficult to meet the extensive applications.⁹ The main reason is that Pt is low abundance in nature and high price, and it is also very difficult to tune its morphology and structure, which cannot meet high-efficiency electrocatalysts in PEMFCs. Moreover, both active Pt NPs and carbon carriers are easily prone to be oxidized and dissolved under the electrochemical tests, leading to an irreversible structure transition compared to the initial structure. In contrast, the current densities of Pt-based alloys are tens or even hundreds of times than that of commercial Pt/C based on the previous reports.^{10–13} As an effective electrocatalyst, Pt-based alloys include Pt-based binary alloys, Pt-based multi-element alloys, and Pt-based high-entropy alloys. With the development of the technology, the following two

strategies have been used to prepare a series of Pt-based electrocatalysts: (1) alloying Pt with a second element, including Ru, Pd, Co, Ni, Fe and Cu, could reduce the usage of Pt and enhance its catalytic performance.^{14–18} That is because that the formation of binary or multi-element alloys can prevent the gradual deterioration of Pt and improve its CO tolerance. The enhanced CO tolerance usually involves in bifunctional mechanism and d-band theory. The former one emphasizes that the adsorbed CO species (CO_{ad}) could react with oxygen-containing species at oxophilic neighbour sites, and the latter one is an electronic effect, which can downshift the d-band center of Pt, thus weakening the interaction between Pt and CO_{ad}. (2) The improvement can also be made by fabricating Pt-based composites with carbon-based materials or metal oxides.^{19–21} In the composites, the synergy between Pt and carbon-based materials or metal oxides is believed to weaken the binding strength of CO. The synergy first involves absorbed hydroxide (OH_{ad}) species, which are obtained by promoting the dissociation of H₂O are essential for the complete oxidation of CO intermediate. Next, another effective method is to employ the charge-transfer mechanism in the composites. The appropriate supports are good for anchoring Pt NPs, and making them uniformly dispersed. Such Pt-based composites can transfer electron from Pt atoms to supports, thus receiving strong binding



Wang Mengwei received a MS degree in Inorganic Non-metallic Materials Engineering from University of Science and Technology Beijing (USTB) in 2009. She is now a researcher in fuel cell research group in Shanghai Institute of Space Power-Sources (SISP). Her research interests include the influence of surface chemistry of metal nanoparticles on catalytic activities and stability for fuel cells.



Sun Xiaohui, a senior engineer, received her bachelor's degree in applied chemistry from Tianjin University. She was engaged in the design, research and development, engineering application and quality management of chemical power supply in Shanghai Space Power Supply Research Institute. She participated in the exploration of improving the atomic utilization rate and reactivity of fuel cell

catalysts, and also carried out targeted quality management research in the research.



Lu Dan obtained a master's degree from Shanghai University in 2017. She mainly participated in the research on the preparation of fuel cell catalysts and electrodes. She has a strong interest in fuel cell-related technology research and engineering applications. She actively participated in the construction of flexible product production lines of fuel cell reactor collection systems, and carried out

extensive and valuable research.



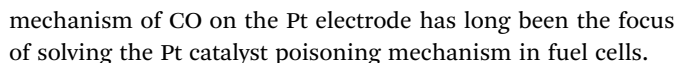
Zhang Huahui received her doctor's degree in material processing engineering from Beijing University of Aeronautics and Astronautics (BUAA) in September 2009. She is a researcher in the fields of development of Materials for chemical power. She is a member of group "organic liquid hydrogen storage fuel cell", and she participated in the research of fuel cell catalysts.

Her ambition is to conduct research to improve the performance of chemical power sources.



CO poisoning has been considered a notorious catalyst-deactivating process to limit the commercialization of PEMFCs.^{22–24} In order to design a novel anti-toxic Pt-based electrocatalyst, the electrochemistry behavior (*e.g.*, adsorption and oxidation) of CO on the Pt surface must be understood. In an acid electrolyte, the adsorption and oxidation of CO usually occur on the (100) and (110) facets of Pt. For CO adsorption, CO competes with fuels or other intermediates for the active sites, and CO adsorption heat is much higher than that of fuels or other intermediates, indicating CO is more easily adsorbed on the Pt active sites. In addition, there are different modes of CO adsorption on the Pt electrode, including linear or top bond, bridge or double bond, as well as hollow or triple bond.^{25–28} In contrast, the linear bond type dominates the mode of CO adsorption on the Pt electrode, and the adsorption isotherm is a Temkin isotherm, as follows:

where θ_{CO} represents CO coverage, ΔG_{o}^0 represents the standard free energy of adsorption, H represents Henry's low constant for CO solubility, and r represents the interaction parameter. CO oxidation is another significant electrochemistry behavior on the Pt electrode. In general, the rate of CO oxidation is relatively rapid at low coverage of CO, while the rate is decreased by high coverage of CO. The coverage rate of CO could reach a threshold on the surface of Pt by the reasonable design of Pt-based electrocatalysts. Therefore, this dependence on CO coverage could be explained when the CO oxidation involves adjacent active sites.^{29,30} As a result, many efforts have proposed the 'reactant pair' mechanism for CO oxidation. Based on the above, fully understanding the adsorption process and electro-oxidation



In this review, we will concisely summarize the latest developments of Pt-based electrocatalysts and its anti-CO poisoning mechanism in the following strategies, including (1) modifying the electronic structure of Pt by alloying Pt with transition metals to weaken the strong adsorption of CO and alleviate its CO poisoning, resulting in the improved activity and stability. (2) Synthesizing Pt-based composites to improve the conductivity and resistance against corrosion and oxidation. Importantly, utilizing the synergistic effect between Pt and carriers provides extra catalytic performance. In addition, we will discuss the advantages and challenges of Pt-based electrocatalysts for future development in PEMFCs. We hope that this review will call for more efforts to explore the design of Pt-based electrocatalysts to inhibit CO poisoning and propose new strategies for usage in PEMFCs (Fig. 1).

Based on the previous reports, the generated CO intermediate during the anodizing process can easily occupy the Pt active sites, thereby leading to decreased catalytic performance.^{31,32} Generally, the electrochemistry behaviors of CO include the adsorption and oxidation on the Pt surface. The linear bond type dominates the mode of CO adsorption, and the adsorption mode and coverage of CO determine the rate of CO oxidation. Therefore, increasing attention has focused on the adsorption and electrooxidation of CO for unravelling the anti-CO poisoning mechanism itself and accelerating the fabrication of anti-poisoning Pt-based electrocatalysts. For example, *in situ* electrochemical shell isolated nanoparticle-enhanced Raman spectroscopy is employed to explore the adsorption and electrooxidation of CO on the Pt(*hkl*) single crystal electrodes, and the decreasing sequence of CO electrooxidation activity is Pt(111) > Pt(100) > Pt(110) in acidic medium (Fig. 2a and b).³³ At the same time, both isotope substitution experiments and DFT calculations clarify that the formation of OH_{ad} and COOH_{ad} species could boost the CO electrooxidation process (Fig. 2c). As a result, fully understanding the adsorption process and electrooxidation mechanism of CO on the Pt electrode has long been the focus of solving the poisoning mechanisms of Pt-based electrocatalysts.

© 2023 The Author(s). Published by the Royal Society of Chemistry

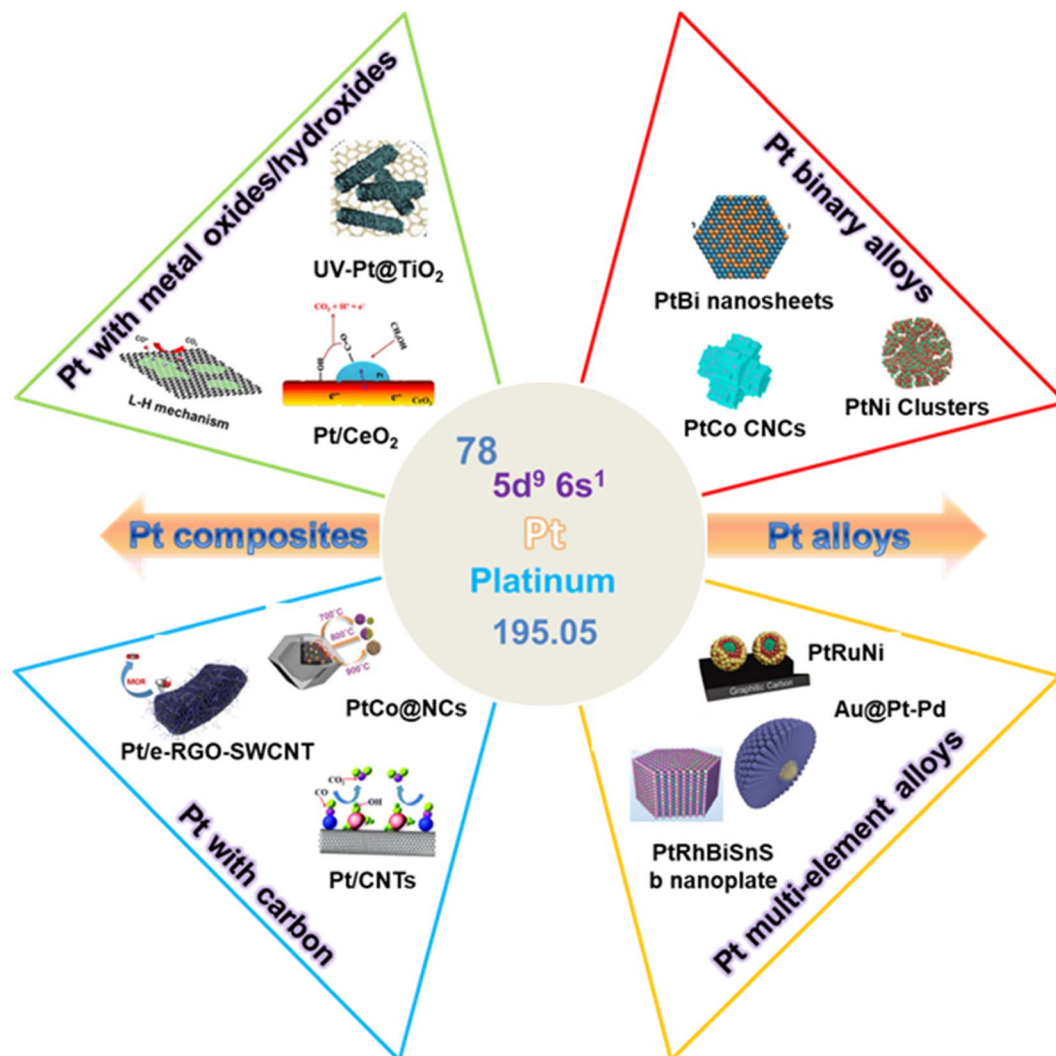


Fig. 1 Design strategies of Pt-based electrocatalysts and its outstanding strategies for CO tolerance.

their weaker adsorption for CO. Therefore, modifying Pt with a second metal could effectively adjust its electronic structure based on electron or ligand effects, including electron donation or reverse donation, thus reducing the strength of Pt–CO bond. More importantly, the adsorption energy of CO is related to the degree of CO coverage, and the oxidation rate of CO is determined by the electronic modification. At lower potentials, the d-band theory mechanism plays a dominant role because of the weak CO adsorption. The bifunctional mechanism occurs in the formation of oxygen-containing species in the oxygenophilic site of the second site at higher overpotentials, whose potential is less than that of Pt, which favors the oxidative removal of CO.^{37–40} Aside from the compositions, geometry and size effects of the catalyst also play an important role in improving the catalytic performance and weakening the Pt–CO bond. By adjusting the size and structure of Pt-based alloys, it can not only promote the exposure of Pt atoms, but also induce the formation of unsaturated coordination sites, thus greatly causing surface strain and space electric domain sites.

2.1 Pt-based binary alloys

In recent years, considerable interest has focused on Pt-based binary alloys owing to their great progress in both basic research and practical applications. The advantages of alloying Pt with a second metal can not only significantly reduce the usage of Pt but also effectively modify its electronic structure and change the coordination environment, thus contributing to the superior catalytic efficiency, selectivity and stability. Based on the guideline, different metal elements (*e.g.*, noble metal and transition metal elements) and geometric structures (*e.g.*, nanoplates, nanowires and porous structure) have been investigated to improve the CO tolerance and electrocatalytic performance of Pt. Both metal elements and geometric structures could predominate the catalytic effect under different conditions. For example, the surface component of Pt-based binary alloys with the same nanostructure is the key factor to determining the catalytic performance. In addition, the nanostructure of Pt-based binary alloys with the same surface component will predominate the catalytic effect.



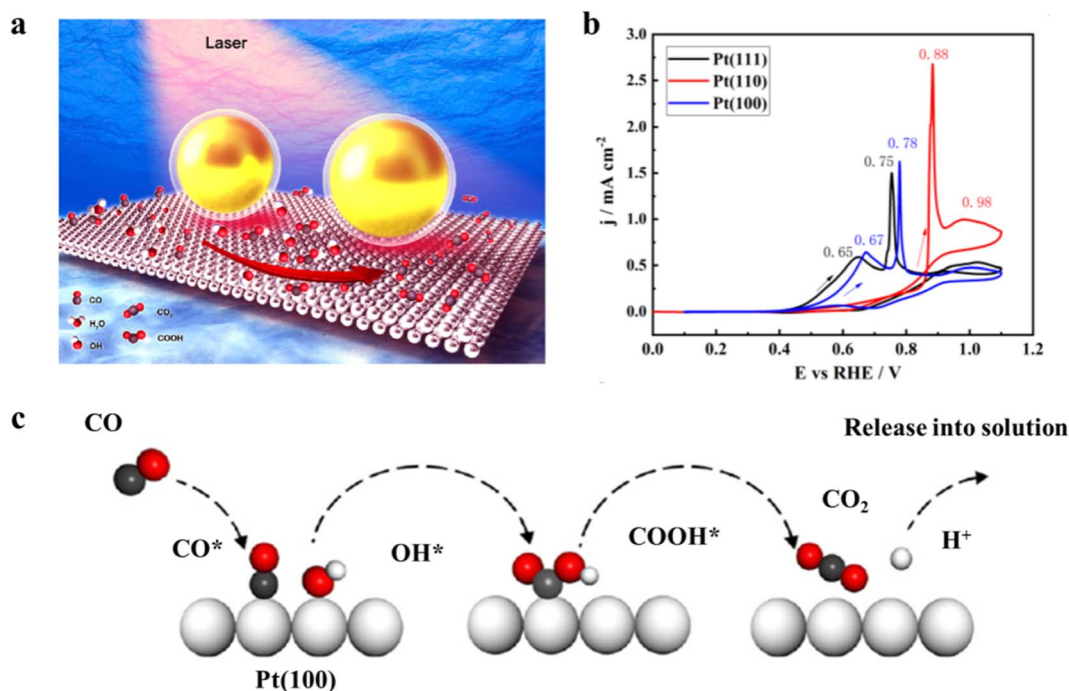


Fig. 2 (a) Schematic showing the *in situ* SHINERS study of the CO electrooxidation process on a Pt(100) surface. Large gold spheres indicate the nanoparticles; (b) CVs of CO electrooxidation on Pt(111) (black curve), Pt(100) (blue curve) and Pt(110) (red curve) single-crystal electrode surfaces in 0.1 M HClO₄ solution saturated with CO vs. RHE. Sweep rate was 50 mV s⁻¹; (c) the proposed mechanism of CO electrooxidation on Pt(100) surface. Reproduced with permission from ref. 34. Copyright 2020, Wiley-VCH.

Alloying Pt with noble metal elements, such as Pd, Au, Ag, Ir, and Ru, can effectively adjust the electronic structure and geometric structure of Pt. Among the noble metal elements, many efforts have focused on the synthesis of PtRu bimetallic alloys due to the famous bifunctional mechanism and electronic effect, which can effectively weaken the CO adsorption and enhance the CO oxidation rate.^{41–43} For instance, Wang *et al.* have prepared porous flower-like Pt₇₂Ru₂₈ nanoalloys assembled with sub-4.0 nm NPs by a facile aqueous method.⁴⁴ Thanks to the optimized molar ratio of Pt:Ru and unique nanostructures, the specific activity and mass activity of Pt₇₂Ru₂₈ nanoalloys can reach 10.98 mA cm⁻² and 1.70 A mg_{Pt}⁻¹ for methanol oxidation reaction (MOR), respectively. The enhanced oxidizing ability of Pt₇₂Ru₂₈ nanoalloys toward CO, HCOH, and HCOOH intermediates significantly makes them exhibit long-term durability for MOR. Another good example involves the fabrication of PtRu bimetallic NPs with adjustable alloying degree *via* a co-reduction method.⁴⁵ Based on *in situ* attenuated total reflection infrared (ATR-IR) spectroscopy, PtRu/PC-H can easily cleave C–H bond and accelerate the oxidative removal of CO_{ad} at a lower potential, indicating the enhanced CO tolerance. Therefore, porous PtRu bimetallic nanoalloys have long-term solvent durability after immersion in water for 16 months. In addition, Cai *et al.* have prepared PtRu/C by using a dimethylamine borane (DMAB)-based aqueous phase process.⁴⁶ The as-prepared PtRu/C exhibits superior MOR performance and output power density due to the enhanced CO resistance and CO₂ generation. Importantly, *in situ* ATR-IR spectroscopy was used to corroborate the improved MOR

performance and output power density at a molecular level. Based on the above, PtRu bimetallic alloys have been considered as one of the most promising catalysts in PEMFCs. In the bifunctional mechanism, Pt can provide the active sites, and the neighbouring Ru can activate water and produce reactive oxides for the oxidative removal of CO_{ad} adjacent to the Pt active sites. Interestingly, the well-defined facets enclosed PtRu nanocrystals exhibited the different CO-poisoning tolerance. Dong *et al.* have synthesized ultrathin PtRu nanocrystals with tuneable facets by a one-step wet chemical approach.⁴⁷ They found that the (111)-terminated PtRu nanocrystals showed much higher mass activity toward MOR, which was 2.28 times higher than (100)-terminated PtRu nanocrystals, indicating the superior CO tolerance of PtRu (111) facets relative to (100) facets.

In addition to the corporation of noble metal elements, many transition metal elements, such as Cu, Co, Ni, Fe, and Zn, have been employed to alloy Pt for improving the CO tolerance of Pt-based binary alloys.^{17,48,49} For example, ultrastable self-supported PtCu nanowires with abundant Cu-vacancies were prepared by combining a co-reduction method and an electrochemical dealloying strategy (Fig. 3a).⁵⁰ A series of CO poisoning measurements and DFT results proved that the optimized PtCu (111) with Cu-vacancy defects could effectively weaken the adsorption of CO and simultaneously enhance the adsorption of OH_{ad}, resulting in outstanding MOR performance (Fig. 3b–g). In addition, Yang *et al.* have successfully prepared Pt–Co nanoframes by etching Co from solid rhombic dodecahedra.⁵¹ The MOR mass activity of Pt–Co nanoframes can reach 4.28 A mg_{Pt}⁻¹ in alkaline media, which is 4 times higher than that of



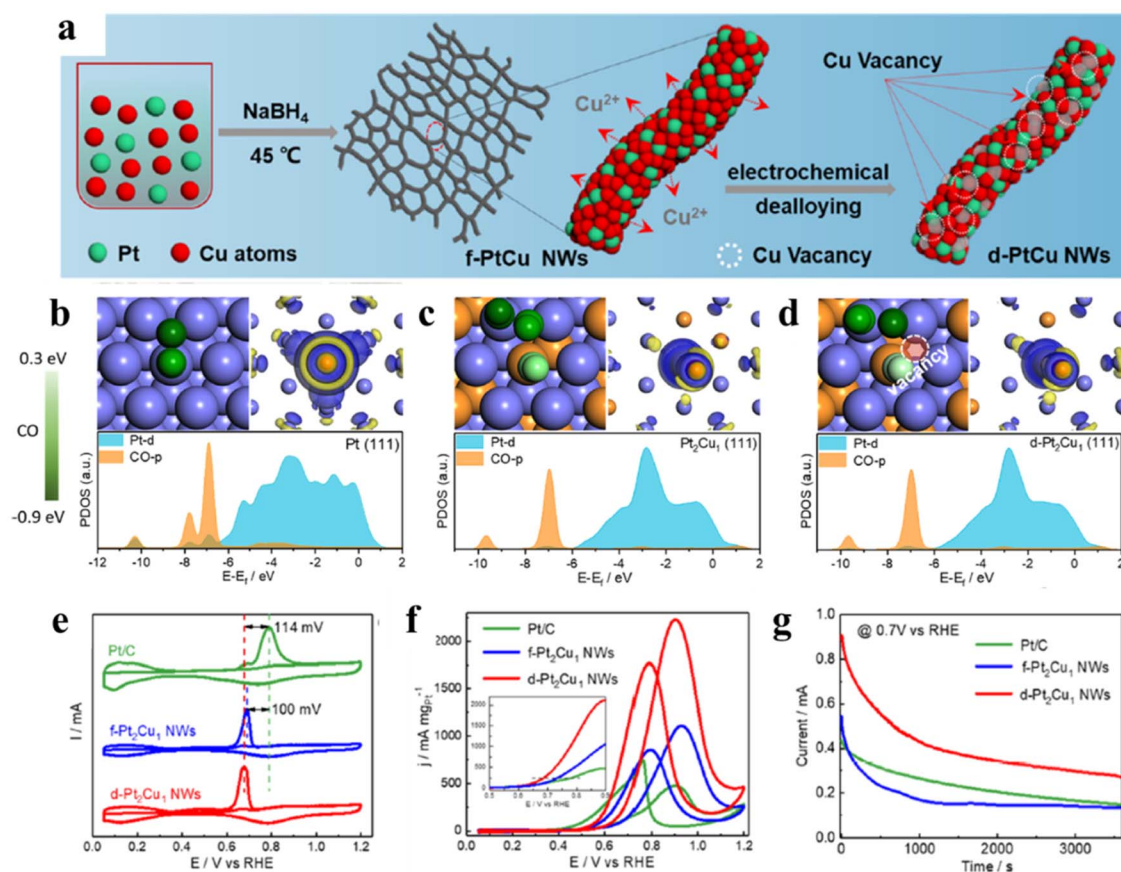


Fig. 3 (a) Schematic illustration of the surface structure evolution of d-Pt₂Cu₁ NWs with the electrochemical dealloying process; (b–d) electronic structure analyses of Pt(111), Pt₂Cu₁(111), and Pt₂Cu₁(111) with Cu-vacancy surfaces. Models for the adsorption of CO (left sides of (b–d)). Charge density differences of CO (right sides of (b–d)). The yellow and blue shadows represent charge accumulation and depletion in the space, respectively. Calculated projected density of states (PDOSs) of coupling CO (below sides of (b–d)) on different surfaces. (e) CO stripping performed at a scan rate of 20 mV s⁻¹ in a N₂-saturated 0.1 M HClO₄ + 0.5 M CH₃OH solution at 50 mV s⁻¹; the inset CV curves show the MOR histograms of the onset potential. (f) MOR CVs normalized to the Pt mass recorded in a N₂-saturated 0.1 M HClO₄ + 0.5 M CH₃OH solution at 50 mV s⁻¹; the inset CV curves show the MOR histograms of the onset potential. (g) Chronoamperometric (CA) measurements at 0.7 V vs. RHE for 3600 s. Reproduced with permission from ref. 50. Copyright 2021, American Chemical Society.

commercial Pt/C. The enhanced MOR performance can be attributed to the robust surface and optimized binding strength of CO. Similarly, PtNi colloidal nanocrystal clusters (CNCs) assembled by many PtNi nanocrystals were synthesized by a simple one-pot solvothermal method.⁵² The corporation of Ni could adjust the catalytic structure and downshift the d-band center, which contributes to weakening the CO adsorption and improving MOR performance.

Bi as the representative main group element has been employed to alloy Pt for improving the CO tolerance of Pt. The incorporation of Bi can regulate the electronic structure of Pt and easily generate metal oxides on the surface of Pt-based alloys, which can promote the bifunctional mechanism. Based on the previous reports, considerable studies mainly focused on the addition of Bi to alloy Pt, which could enable PtBi alloys with high tolerance to CO poisoning. Based on this guideline, PtBi–Pt core-shell nanoplates with controlled composition and structure have been synthesized by using a simple solution-phase reaction.⁵³ The incorporation of Bi species not only

accelerates the formation of ordered intermetallics but also downshifts the d-band center, thereby resulting the enhanced MOR activity and stability. Moreover, the core-shell structure contributes to the strain effect due to lattice mismatch and contraction between Pt and Bi, further promoting the MOR performance. Similarly, Guo *et al.* have developed a new class of 2D nanoplates by a facile one-pot approach, which is composed of an intermetallic PtBi core and ultrathin Pt shell.⁵⁴ Thanks to the strain effect, the as-prepared catalysts also exhibit enhanced MOR performance in both acid and alkaline electrolytes. In addition, significant studies have proved that Bi can irreversibly adsorb on the Pt electrode. For example, Shin *et al.* have obtained Bi-modified Pt supported on carbon black by mixing Bi₂O₃, 40 wt% commercial Pt/C and 0.5 M H₂SO₄.⁵⁵ Bi is irreversibly adsorbed on Pt/C, and the amount of Bi can be tuned by the concentration of Bi³⁺ solution. Therefore, the current density of Bi-modified Pt/C for formic acid oxidation reaction (FAOR) is 13 times higher than that of Pt/C. The enhanced FAOR activity can be attributed to the fact that the adsorbed Bi works



as a third body, making Pt ensembles and thus significantly decreasing CO intermediate' generation. Another interesting finding is that porous Pt nanoframes decorated with $\text{Bi}(\text{OH})_3$ have been fabricated by a two-step method. On the one hand, incorporating Bi species could effectively tune the geometric structure and downshift the Pt d-band center. On the other hand, the decoration of $\text{Bi}(\text{OH})_3$ can favor the oxidative removal of CO by using the formation of OH_{ad} species.⁵⁶ In general, the decoration of Bi species could greatly enhance the electrocatalytic performance of Pt by geometrically separating adjacent Pt surface sites and adsorbing Bi adatoms on the Pt surface.

In order to further enhance the utilization of Pt atoms, accelerate the electron and mass transfer, and expose more active sites, many efforts have been paid attention to the design and fabrication of various Pt-based nanostructures, such as one-dimensional nanowires, two-dimensional nanoplates, and three-dimensional polyhedra and nanoframes.^{23,57–59} For example, Yang and co-authors have developed bimetallic $\text{Pt}_{95}\text{Co}_5$ nanowires by a one-pot surfactant-free method.⁶⁰ The formation of one-dimensional nanowires could not only reduce the cost but also contribute to the superior electrochemical properties due to the modification of Co, which can improve the tolerance to carbonaceous species, such as CO_{ad} , CHO_{ad} , and COOH_{ad} intermediates. Moreover, Tang *et al.* employed a facile one-pot hydrothermal method to prepare porous Pt_3Cu nanobowls that are assembled by ultrafine NPs (≈ 2.9 nm).⁶¹

Interestingly, *N,N'*-methylenebisacrylamide works as a crucial structure-directing agent, which plays an important role in the deposition and growth of well-defined bowl-shaped nanostructures. Therefore, the self-supported Pt_3Cu nanobowls show significant MOR activity and anti-poisoning capability. A series of characterizations and measurements further confirm that highly open geometric configuration, the number of the energetic low-coordinated sites and compressive lattice strain contribute to the high utilization efficiency of Pt and induce the compressive lattice strain, thus tuning the Pt d-band center. In addition, TEM image of Pt_3Cu nanobowls after the stability test shows the perfect maintenance of their initial bowl-like nanostructure, illustrating the bowl-like nanostructure could not only reduce the Pt usage but also contribute to the superior CO tolerance.

In addition to the above-mentioned strategies, there is an important approach to improve the Pt utilization and its anti-poisoning properties, which involves the design of point defects (atomic vacancies) and surface defects (grain boundaries). That is because the generated defects could regulate the surface electronic structure of Pt and contribute greatly to improving their adsorption capacity for CO intermediate.^{62,63} For example, porous PtAg nanoflowers have been successfully prepared by combining liquid reduction and chemical etching method. Wang *et al.* first synthesized different element proportions of PtAg alloys by adjusting the feeding ratios of Pt

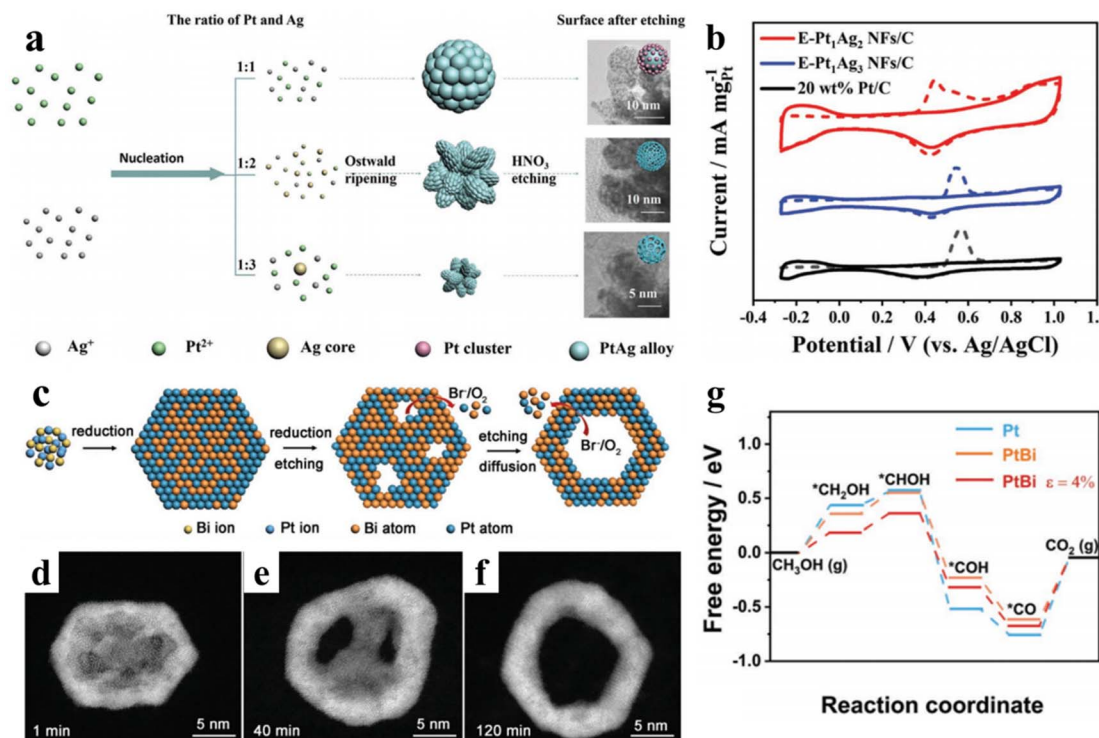


Fig. 4 (a) The schematic diagram of the synthesis of E-PtAg NFs; (b) CO-stripping voltammograms in 0.1 M HClO_4 solution at a scan rate of 50 mV s^{-1} for E-Pt₁Ag₂ NFs/C, E-Pt₁Ag₃ NFs/C, and commercial Pt/C, respectively; reproduced with permission from ref. 64. Copyright 2022, Wiley-VCH. (c) Schematic illustration of the formation mechanism of PtBi NRs; (d–f) HAADF-STEM images of time-dependent products obtained at various reaction time; (g) free energy diagram for MOR on Pt, PtBi, and PtBi $\epsilon = 4\%$ surfaces. Reproduced with permission from ref. 65. Copyright 2022, Wiley-VCH.

and Ag precursors, resulting in uneven element distribution.⁶⁴ Next, a chemical etching method was used to corrode silver by the addition of nitric acid, finally leading to the formation of porous PtAg nanoflowers. More importantly, the molar ratio of Pt : Ag can greatly affect the size and morphology of porous PtAg nanoflowers. Further chemical etching method determines the forming process of porous crystal (Fig. 4a). Therefore, porous PtAg nanoflowers manifest a mass activity of 1136 mA mg_{Pt}⁻¹, which is 2.6 times higher than that of commercial Pt/C. Both CO-stripping voltammograms and forward/reverse peak current density ratio exhibit high anti-CO poisoning capability due to the incorporation of Ag, which could alter the Pt electronic structure and promote the oxidative removal of CO species (Fig. 4b). Therefore, it is a very effective strategy to adjust the adsorption/desorption properties of CO intermediate by fabricating porous nanostructures with abundant surface defects. Another impressive sample focused on the fabrication of bimetallic PtBi nanorings with inhomogeneous tensile strain and abundant low coordination atoms.⁶⁵ PtBi nanoplates were first prepared by a facile solvothermal method, and the porous nanoplates were generated by a self-etching effect with Br⁻ and O₂ pairs as etching agents. The mass activity and specific activity of PtBi nanorings for MOR can reach 6.42 A mg⁻¹ and 11.93 mA cm⁻², respectively. Both CO stripping measurements and DFT results have demonstrated that the improved MOR performance of PtBi NRs could be ascribed to the tensile strain and ligand effect, which arise from the well-defined crystal structure and modification of Pt electronic structure owing to the doping of Bi (Fig. 4c–g). In addition, the morphology of the PtBi NRs were well maintained after the CA test, indicating their excellent structural stability.

Except for bimetallic Pt-based alloys, the single-atom alloy (SAA) strategy could significantly reduce the amount of Pt loading and promote the catalytic performance by effectively decreasing the binding strength of CO.^{22,66} Therefore, one of the biggest advantages of SAA catalysts is that each Pt atom is highly exposed on the surface of electrocatalysts, which can improve the Pt utilization rate and make them highly active catalysts for electrochemical reactions. To promote the fast conversion of CO to CO₂ at low temperature in fuel cells, Flytzani-Stephanopoulos and co-workers have systematically investigated the interaction between CO and different Pt ensembles on Pt–Cu SAA by using a variety of surface science and catalytic techniques.²⁴ They found that CO is weakly adsorbed on single Pt atoms, while both the larger Pt ensembles and monometallic Pt exhibit stronger adsorption capacity for CO. As a result, more CO-free Pt sites are available on the SAA electrode compared to larger Pt ensembles and monometallic Pt, thereby contributing to obtaining higher hydrogenation and hydrogen electro-oxidation activity. In addition, single Pt atoms dispersed on Ru NPs have the advantage of providing OH_{ad} species based on the bifunctional mechanism, which can easily remove the CO intermediate from the active sites. For example, Tilley *et al.* successfully synthesized single Pt atoms on the surface of Ru by the slow reduction of a Pt(II) precursor with oleylamine. More importantly, they employ a high temperature treatment to drive the rearrangement of Pt islands on the branched Ru, resulting

in the formation of discrete Pt atoms on the Ru electrode.⁶⁷ Therefore, the forming OH_{ad} species adjacent to the active sites of Ru can easily remove the CO intermediate when CO molecules adsorb on the Pt atoms, thus leading to improved MOR activity and stability. DFT calculations further confirm that the structural properties of single Pt atoms could strongly promote the adsorption of CH₃OH and reduce the adsorption strength of CO intermediate, thereby limiting their CO poisoning and improving the MOR performance.

In summary, alloying Pt with a second element (Pd, Zn, Co, Cu, Au *etc.*) has considered as one of the most effective ways to tune its geometric structure, electronic structure and interactions with reaction intermediates.⁶⁸ Importantly, the second element and special nanostructures can induce synergistic effects to abundant active sites, high mass utilization, and fast mass transfer and weaken the affinity of Pt to the poisonous species. Despite the success, the second element is easily oxidized and dissolved, thus leading to unstable nanostructures, which could reduce the CO tolerance of Pt-based binary alloys. As a result, optimizing the amount and species of the second element is especially important for the anti-CO poisoning.

2.2 Pt-based multi-element alloys

Based on the previous reports, the increased activity and stability of Pt-based electrocatalysts by alloying Pt with transition metals could be attributed to downshifting the Pt d-band center, thus optimizing the adsorption of CO intermediate. Therefore, the catalytic activity and stability of Pt-based electrocatalysts can be further enhanced by the synergistic interaction between different active components, which have many advantages for the design of Pt-based multi-element alloys.^{69–72} First, the incorporation of less-expensive metals could greatly reduce the usage of Pt. Second, the multi-element properties can effectively provide multiple sites to adsorb different intermediates, respectively. Third, the electronic structures of Pt active sites can be finely adjusted by the surrounding atoms. For example, Xie and Lin *et al.* have successfully prepared trimetallic PtNiCo branched nanocages by etching Ni-rich cores of PtNiCo nanomultipods (Fig. 5a).⁷³ The mass activity of trimetallic PtNiCo nanocages is 2.8 A mg_{Pt}⁻¹ in acidic medium, which is much higher than that of commercial Pt/C. According to *in situ* FTIR spectra, trimetallic PtNiCo nanocages show the weaker interaction of Pt–CO_{ad} binding strength. As a result, the alloying effects of Co and Ni could optimize the Pt d-band center and improve the CO tolerance of Pt, leading to the enhanced MOR performance on the trimetallic PtNiCo nanocages. (Fig. 5b and c). MOR durability tests were conducted by CA tests for 3600 s. PtNiCo nanocages possessed the highest mass activity compared to PtNi/C or Pt/C even after the durability testing, demonstrating the enhanced CO tolerance is beneficial to improve the catalytic stability of PtNiCo nanocages.

The rational design of ternary Pt-based electrocatalysts is of vital significance for the practical applications of fuel cells. Tian and his co-authors have developed a series of two-dimensional ternary PtBiM nanoplates (M = Co, Ni, Cu, Zn, Sn) *via* a facile



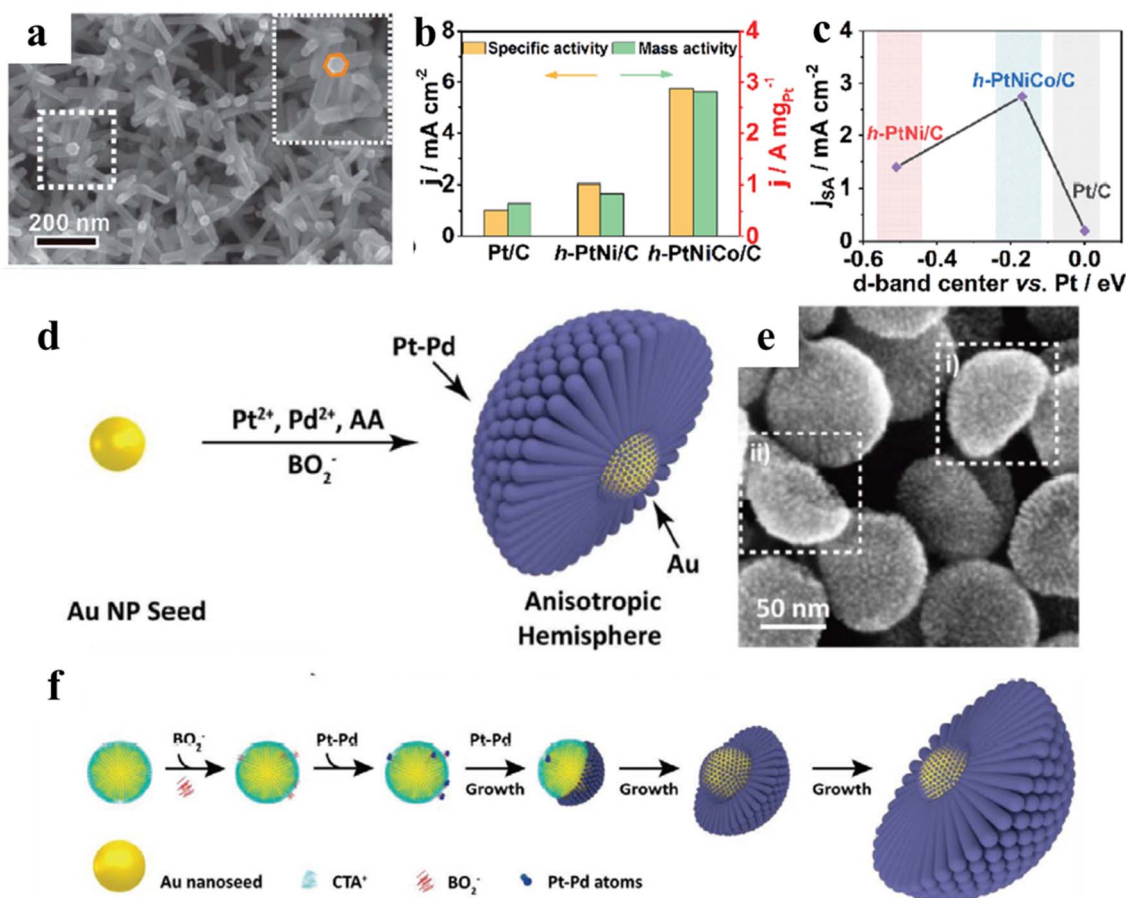


Fig. 5 (a) SEM image of trimetallic PtNiCo nanomultipods with a core-shell structure; (b) comparison of specific activity and mass activity at 0.9 V (vs. RHE); (c) the oxygen reduction reaction (ORR) specific activity for catalysts associated with the shift in d-band center vs. Pt/C; reproduced with permission from ref. 73. Copyright 2021, Royal Society of Chemistry. (d) The synthetic schematic diagram of Au@Pt-Pd H-Ss; (e) SEM image of the Au@Pt-Pd H-Ss; (f) schematic diagram of the growth mechanism of Au@Pt-Pd. Reproduced with permission from ref. 81. Copyright 2021, Wiley-VCH.

solvothermal method.⁷⁴ Both MOR activity and durability exhibit significant improvement with the incorporation of metal compositions. Notably, $\text{Pt}_3\text{Bi}_3\text{Zn}$ nanoplates exhibited a much slower decay than that of the other catalysts after one thousand cycles, which further indicated a significant improvement in the stability. The prominent performance can be attributed to the optimized electronic structure of Pt and generated OH_{ad} species, which will result in weak adsorption between Pt and CO and the oxidative removal of CO. Up to now, single-atom tailoring metal alloys have been extensively used to maximize the Pt active center because of its unsaturated coordinative environment and unique electronic structure, which benefit the oxidative removal of CO. For instance, Sun and his co-authors have developed PtNi NPs with the modification of Ru single atoms by combining chemical etching and a selective atomic layer deposition (ALD) technique.⁷⁵ Impressively, the mass activity of Ru-ca-PtNi hybrid catalysts can reach 2.01 A mg^{-1} for MOR, and only 16% activity loss of its initial activity can be observed during the tests, indicating a significant durability. DFT calculation and FTIR analysis prove that the

confined Ru atoms in cavities could efficiently promote the removal of CO by tuning the Pt d-band center. On the basis of ternary alloys, the addition of special functional metal has been believed to make the ternary alloys have the corresponding functions. Based on the guideline, Xie *et al.* have fabricated 1.5 nm-thin quarter metallic PtCoNiRh nanowires with high atomic-exposure by using an oleylamine system.⁷⁶ PtCoNi nanowires were used as a control, and Rh as a highly anticorrosive metal was used to construct Pt-Rh sites on the surface of quaternary metallic PtCoNiRh nanowires. Compared to PtCoNi nanowires, the mass activity and specific activity of PtCoNiRh nanowires show the significant improvements in MOR. Moreover, the onset potential of MOR is distinctly negative, and the prominent CO tolerance also can be observed on the PtCoNiRh nanowires. Both *in situ* FTIR study and DFT calculation have proved that the incorporation of Rh could change the CO binding from linear mode to bridge mode, and Pt-Rh sites belong to the thermodynamical sites for adsorption and dehydrogenation of CH_3OH . It is generally accepted that the enhancement of catalytic activity stems from the lattice strain

caused by the lattice mismatch between Pt and other metals, which can weaken the interaction between Pt atoms and the adsorbed species, and thus release more active sites on the surface. A variety of elements doped in Pt lattice can cause lattice strain due to their different lattice and directly affect the local coordination environment mainly by regulating the non-uniform elements, thus greatly changing the adsorption properties of nanostructures.

Remarkably, many studies have focused on the design of the structure and morphology of Pt-based multi-element alloys based on the incorporation of three or even more elements, especially core-shell structures, ordered alloys, graded interfacial nanolayers, and open-structure nanoframes.^{77–79} Among the unique catalytic structure, fabricating shape-controlled core-shell structures with functional Pt surfaces could provide an effective routine to improve the catalytic properties of Pt-based alloys, which will involve the synergistic effect of multi-component alloys. However, how to develop a facile approach is the critical factor for the fabrication of multi-component core-shell structures, which usually include the shape control of metallic core and thickness control of Pt-based shell. Based on the above, Fang *et al.* have prepared octahedral CuNi@Pt–Cu core-shell nanocrystals.⁸⁰ In the core-shell structure, the formation of Pt-based shell should precisely control the experimental parameters, including Pt precursor/capping ligand, temperature and heating rate. The resultant CuNi@Pt–Cu catalysts exhibit superior MOR activity compared to commercial Pt/C. That is because that the onset potential of CO electro-oxidation peaks on the CuNi@Pt–Cu electrode is more negative than that on the commercial Pt/C, indicating a weaker CO adsorption affinity. Similarly, Jiang *et al.* have successfully fabricated anisotropic Au@Pt–Pd hemispherical nanostructures, in which Au NPs are the core, and Pt–Pd NPs are the shell (Fig. 5d and e).⁸¹ In the synthetic process, BO^{2-} with ultra-low concentration will result in their non-uniform adsorption on Au seed surface. Therefore, both the non-uniform adsorption of BO^{2-} and rapid growth of Pt–Pd atoms promote the formation of anisotropic hemispherical nanostructures, as shown in (Fig. 5f). Compared to commercial Pt/C, Au@Pt–Pd hemispherical nanostructures manifest superior electrocatalytic activity and durability for alcohols oxidation and FAOR. The great improvement could be attributed to the existence of surface defects and unsaturated coordination environment, which can weaken the adsorption of CO, thus improving their catalytic performance. It is well known that small organic molecules oxidation reaction at the anode belongs to a surface structure-sensitive reaction, and thus the surface crystal structure of Pt-based electrocatalysts can determine their activity and selectivity. For instance, Pt (110) facets can effectively promote the cleavage of C–C bond in acidic solution.⁸² In addition to the surface crystal structure and compositions, strain engineering can tune the Pt electronic structure and adsorption energy of CO intermediate. Based on the theoretical guidance, Cui *et al.* have constructed a class of core-shell hexagonal nanoplates (PtPd@PtIr1 HNPs) enclosed by Pt (110) facets with a 7.2% tensile strain.

PtPd@PtIr1 HNPs show a top-ranked activity and a Faraday efficiency of 57.93% for C1 pathway for ethanol oxidation reaction (EOR).⁸³ DFT results confirm that the strong synergy between Ir1 sites and tensile strain contributes to breaking the C–C bond of CH_2CO^* and facilitating oxidative removal of CO. As a result, it can be seen that shape-controlled multi-element nanocrystals with low-coordination active sites could effectively tune their catalytic properties.

2.3 Pt-based high-entropy alloys

It is well known that alloying Pt strategy with relatively small amounts of secondary elements has long been used to confer desirable properties to Pt-based alloys. A novel alloying strategy has been developed to combine multiple elements in high concentrations to generate high-entropy alloys (HEAs). In general, HEAs contain five or more metal elements with similar atomic ratios, and each component has a different atomic size, thereby easily leading to lattice distortion. Compared to conventional alloys, major effects of HEAs include high entropy, lattice distortion, slow diffusion, synergic effect, and high organizational stability.^{84–86} Therefore, HEAs with unique structural properties and a significant high-entropy effect will break through the bottleneck of electrochemical catalytic materials in fuel cells.

In the past few years, HEAs as catalytic materials have arisen extensive interest owing to their unique characteristics. Moreover, the catalytic properties of HEAs can be tuned by adjusting their compositions and phase structure. Specially, for the uniform and ultrasmall (~ 3.4 nm) HEAs, their dissimilar element interaction could exhibit novel catalytic properties. For example, Wang and co-authors have prepared the uniform and ultrasmall $\text{Pt}_{18}\text{Ni}_{26}\text{Fe}_{15}\text{Co}_{14}\text{Cu}_{27}$ HEAs by a simple low-temperature oil phase strategy.⁸⁷ The transition metal elements (Ni, Fe, Co, and Cu) are selected because they are in high abundance and can easily form solid solutions with Pt. Therefore, $\text{Pt}_{18}\text{Ni}_{26}\text{Fe}_{15}\text{Co}_{14}\text{Cu}_{27}$ HEAs show the mass activity of $15.04 \text{ A mg}_{\text{Pt}}^{-1}$ for MOR, and the I_f/I_b (I_f represents the forward current density, and I_b represents the reverse current density) ratio of $\text{Pt}_{18}\text{Ni}_{26}\text{Fe}_{15}\text{Co}_{14}\text{Cu}_{27}$ HEAs (3.26) is much higher than that of commercial Pt/C (2.31), indicating the superior CO tolerance of $\text{Pt}_{18}\text{Ni}_{26}\text{Fe}_{15}\text{Co}_{14}\text{Cu}_{27}$ HEAs. DFT results have proved that the active sites on the HEA surfaces exhibit very varied adsorption preferences for CO intermediate. Therefore, a larger energy barrier of 0.81 eV and activation energy (0.94 eV) are obtained for the formation of CO, resulting in suppressing the CO poisoning. Compared to conventional alloys with typically disordered nanocrystals, the intermetallics possess highly ordered structures, which can effectively clarify the structure–property relationship in practical applications. Therefore, how to combine the advantages of intermetallics and HEAs have become promising candidate as multifunctional electrocatalysts. Based on the design, Quan group has successfully synthesized PtRhBiSnSb HEAs with hexagonal close-packed structure by a facile wet-chemistry method, in which the formation of PtBi-type (PtRh)(BiSnSb) arises from the substitution of Sn/Sb atoms at the Bi position



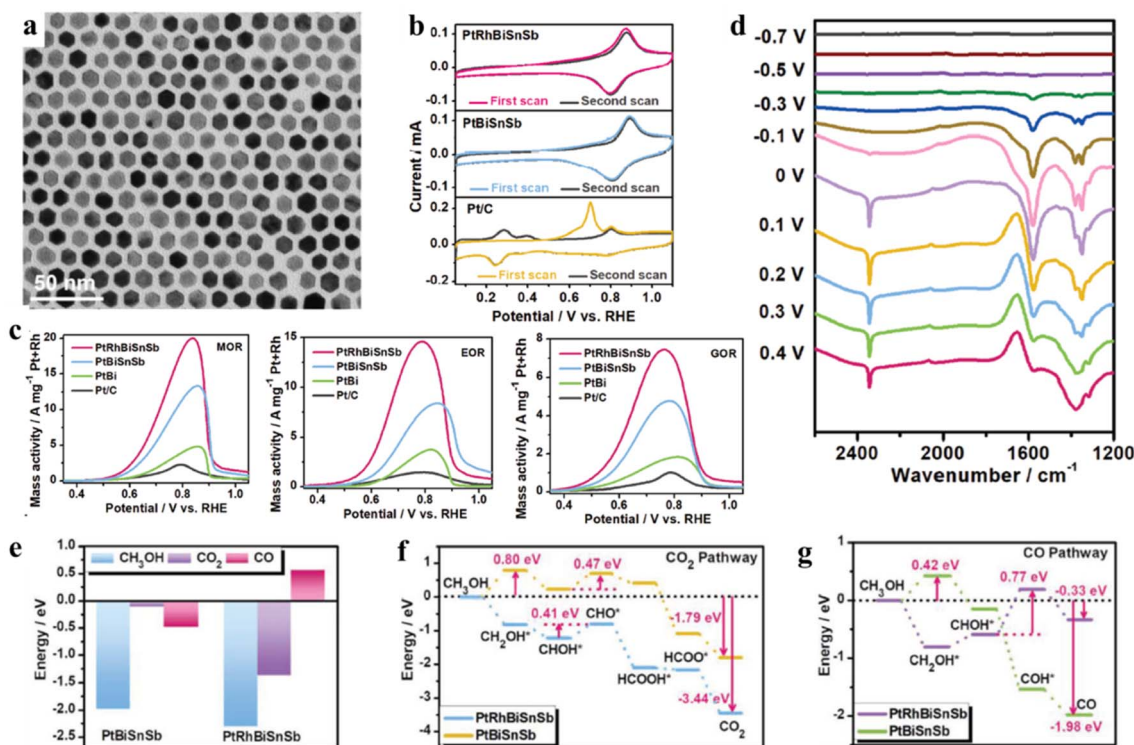


Fig. 6 (a) TEM image; (b) CO-stripping curves of PtRhBiSnSb HEI nanoplates, PtBiSnSb nanoplates, and Pt/C catalysts recorded in Ar-saturated 1.0 M KOH at a scan rate of 50 mV s^{-1} ; (c) MOR, EOR and GOR positive-going polarization curves of different catalysts recorded at a scan rate of 50 mV s^{-1} ; (d) the *in situ* FTIR spectra of PtRhBiSnSb HEI nanoplates recorded in Ar-saturated 1.0 M KOH + 1.0 M CH_3OH solution using a Hg/HgO as the reference electrode; (e) the adsorption energy comparison of CH_3OH , CO_2 , and CO on PtRhBiSnSb HEI nanoplates and PtBiSnSb nanoplates. (f and g) The reaction energy comparison of CO_2 pathway (f) and CO pathway (g) for MOR process on PtRhBiSnSb HEI nanoplates and PtBiSnSb nanoplates. Reproduced with permission from ref. 88. Copyright 2022, Wiley-VCH.

and Rh atoms at the Pt position (Fig. 6a).⁸⁸ The as-obtained PtBi-type HEAs exhibit the advantages of HEAs and intermetallics, thereby demonstrating the isolated and combined multimetallic elements, stronger interatomic interaction, enlarged Pt–Pt distance, and better etching resistance. As a result, PtRhBiSnSb HEAs show not only ultra-high mass activities in MOR, EOR and GOR (Fig. 6c) but also long-term stability. The oxidation peaks of CO_{ad} at 0.7 V vs. RHE are significantly suppressed on the PtRhBiSnSb and PtBiSnSb electrodes compared to the Pt/C, which can be attributed to the isolated Pt sites on their surface without contiguous Pt active sites and enlarged Pt–Pt distance. DFT calculation has further proved that CO_{ad} species have been significantly inhibited at 0.7 V owing to the optimal electronic structure of Pt and enlarged Pt–Pt distance, resulting in stronger selectivity and durability for the CO_2 reaction pathways (Fig. 6b). In addition, exposure and modification of oxophilic Bi/Sn/Sb species also indirectly weaken the CO poisoning on PtBi-type HEAs (Fig. 6d–g). Impressively, the hexagonal shape and multimetallic compositions of PtRhBiSnSb HEAs can be basically maintained, showing their excellent structural and compositional stability. As we known, Pt-group metals (PGMs) contain six neighbouring elements in the periodic table, and each PGM could efficiently accelerate catalytic reactions. Therefore, alloying Pt with PGMs will promote the formation

of Pt-based HEAs with an excellent structure for complex or multistep reactions, which involve many reactants or intermediates. For example, Kitagawa and his co-authors have developed Pt-based HEAs with all six PGMs (denoted as PGM-HEAs) by using facile chemical synthesis.⁸⁹ They found that PGM-HEAs show the record-high activity and stability for EOR. The enhanced EOR performance can be attributed to alloying six elements together, which might provide an optimal surface for the adsorption/desorption of reactants or CO intermediate.

The above results indicate that Pt-based HEAs formed by incorporating different metal elements can effectively improve their catalytic performance and anti-CO poisoning pathway. However, the design of Pt-based HEAs don't focus on the precise control of elemental component and species at the atomic level, including the modification of different ligands, multiatomic sites and nanoclusters with atomic precision and single atom deposition. Some innovative synthesis strategies are still needed to further develop various advanced Pt-based HEAs. In addition, it is greatly need to investigate the synergy in Pt and other added elements due to the lack information at the atomic level. More importantly, it is crucial process to optimize the design of current reactors and systems for practical applications. Except for the specific shortcomings of the current synthetic Pt-based HEAs method, it is still urgent to investigate their catalytic mechanism in the reaction process.

3. Pt-based composites

According to the above discussion, it is feasible to alloy Pt with metal elements to tune the adsorption/desorption of CO on the Pt-based electrocatalysts. In addition to the alloying strategy, fabricating Pt with carbon-based materials or metal oxides have been considered as one of the most effective way to improve catalytic performance and promote the wide applications of Pt-based electrocatalysts in the field of catalysis.^{28,52,90} In the composites, the synergy between Pt and carriers is believed to mainly contribute to the enhanced performance due to the optimal electronic structure of Pt and the formation of OH_{ad} species adjacent the Pt active sites, which can be determined by both selecting the optimized carriers and constructing suitable interface structure.⁹¹ The former one could weaken the binding strength of CO owing to the formation of electron-rich structure on the Pt electrode, and latter one is essential for the complete oxidation of CO intermediate.^{92,93} In addition, appropriate carriers can favor the better dispersion of Pt on their surface, which can promote electron transfer between Pt atoms and carriers, thus receiving strong binding of single Pt and positively charged Pt atoms. Therefore, the fabrication of Pt-based composites provides an important opportunity to understand the catalytic process and mechanism and explores the relationship between structure–activity–performance and CO intermediate.^{94–96}

3.1 Carbon-supported Pt-based electrocatalysts

To improve the utilization and catalytic performance of Pt, carbon-based materials with high surface area have been considered as the promising support. That is because carbon-based materials with a large specific surface area could benefit the dispersion of Pt NPs and lower their cost.^{97,98} Many carbon-based materials, such as graphene, carbon nanotubes (CNTs), carbon black (CB), and carbon clothes *etc.*, have received increasing attention because of their intrinsic properties, such as a large surface area, porous structure and good electric conductivity. As a very valuable catalyst support, graphene materials with special electronic structure can result in uneven distribution of surface charge, which can promote the formation of a large number of active sites. In addition, their conjugate structure and cavity structure make them possess excellent transport performance and semiconductor properties. Therefore, the accurate design and manufacture of novel graphene-supported Pt catalysts with precise atomic compositions and efficient morphology control have extensive practical prospects.

For example, Li *et al.* synthesized single-nickel-atom-alloyed platinum hexagonal nanocrystals (NiPtSAA/GDY) by a simple electrochemical deposition strategy, which induces the porous GDY hexagonal ring by the domain-limiting effect (Fig. 7a).⁹⁹ Both the blue and red lines represent the lattice spacing of Pt hexagonal nanocrystals (0.230 nm, HPtNs/GDY, Fig. 7b). The green circle in Fig. 7c represents the Ni atoms, and different lines represent the lattice spacing of NiPtSAA/GDY (0.212 nm). Notably, the GDY can easily tune the size and morphology of Pt

NPs, and alloying a single Ni atom with Pt NPs will promote their lattice distortion, thus leading to reducing the reaction kinetic barrier. Therefore, the low valence of Pt in NiPtSAA/GDY shows a high resistance to CO poisoning and enhanced MOR performance. Similarly, Sun *et al.* have proposed a facile treatment strategy for controlling PtRu/PC-H nanocrystals with high alloying temperature.¹⁰⁰ First, a simple hydrogen co-reduction strategy was performed at 300 °C, and then homogeneous PtRu NPs on porous graphitic carbon surfaces were obtained by high temperature-treatment at 700 °C. *In situ* FTIR spectroscopy shows the appearance of a CO₂ signal (2345 cm^{−1}) at −150 mV on PtRu-PC-H, which is much lower than that of PtRu/PC-L and Pt/PC, indicating the decreased CO oxidation potential. As an effective carrier, another important carbon-based material is carbon nanotubes, which possess a one-dimensional hollow tubular structure, and their properties mainly depend on the curled way of the graphene sheet, which will determine the metal properties or semiconductor properties. At the same time, as an excellent catalytic carrier, carbon nanotubes also have a great specific surface, stable chemical properties, excellent mechanical properties, and unique pore structure. Based on the above, Hou *et al.* have developed a facile strategy to synthesize a free-standing 3D interconnected embossed graphene (e-RGO)-single wall carbon nanotubes (SWCNTs) paper supported Pt NPs (Pt/e-RGO-SWCNT). SWCNTs were assembled to insert through the film to promote the penetration of electrolyte into the inner structure of three-dimensional porous electrode.¹⁰¹ Benefiting from the structural properties, Pt/e-RGO-SWCNT exhibits a current density of 191.7 mA mg^{−1}, demonstrating excellent MOR performance. The generation of OH_{ad} species promotes the oxidative removal of CO on the Pt/e-RGO-SWCNT, thus contributing to the improved MOR performance (Fig. 7d–f). Moreover, a bottom-up approach has been developed to prepare ultrafine Pt NPs dispersed on three-dimensional low-defect carbon nanotube/nitrogen-doped graphene hybrid aerogel (Pt/LDCNT-NG).¹⁰² The main synthetic process contains two steps: (1) the self-assembly of three-dimensional LDCNT-NG aerogels and dispersion of LDCNT on GO surface *via* π – π supramolecular interaction; (2) controllable growth of Pt NPs on the LDCNT-NG backbones. Therefore, Pt/LDCNT-NG manifests high mass activity and reliable long-term stability. They found that N dopants can accelerate the formation of a large number of OH_{ad} species, thereby removing CO-like poisoning intermediates. In general, increasing attention has been paid to the interaction between Pt and carbon nanotubes, which can determine the forming process, structure, and dispersion of Pt NPs. It is well-known that carbon nanotubes as a carrier can alter the catalytic system's Galvani potential and lower the Fermi level, thus benefiting the electron transfer at the electrode–electrolyte interface. As a result, to obtain highly catalytic catalysts, the effects and interaction mechanism need to be clearly understood.

In recent years, many efforts have focused on the fabrication of carbon-based materials by converting metal–organic frameworks (MOFs) into N-doped structures owing to their superior structural properties, including a large specific surface area, high porosity as well as adjustable pore structures.¹⁰³ Based on



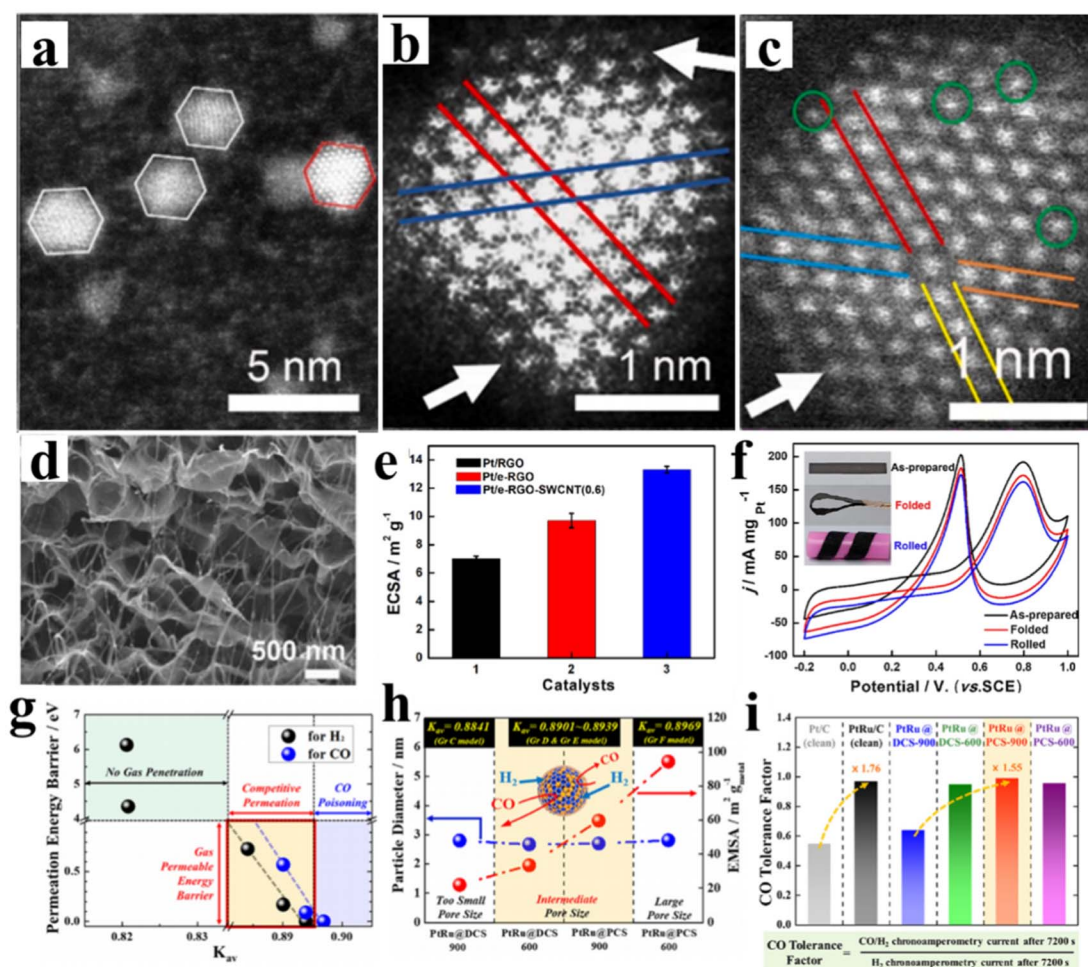


Fig. 7 (a and b) HAADF-STEM images of HPtNP/GDY; (c) HAADF-STEM images of NiPtSAA/GDY; reproduced with permission from ref. 99. Copyright 2020, Elsevier. (d) SEM images of the cross-section of Pt/e-RGO-SWCNT. (e) The histograms of ECSAs of as-prepared samples. (f) CVs of Pt/e-RGO-SWCNT catalyst toward MOR in 0.5 M H₂SO₄ + 1.0 M CH₃OH at 50 mV s⁻¹ with different deformed states: as-prepared, folded, and rolled. Note that the rolled film was immobilized on insulating plastic substrates. Reproduced with permission from ref. 94. Copyright 2019, Elsevier. (g) DFT-computed permeation energy barriers for H₂ and CO gases vs. K_{av} of the different types of defective graphene. (h) Correlation between particle diameter and EMSA of PtRu@C/C catalysts including the available volume fraction (K_{av} = available volume/total volume) values estimated from DFT calculations. (i) CO tolerance factor of the corresponding catalysts calculated from the CA currents in H₂/CO gas mixture (100 ppm CO) and pure H₂ gas atmospheres, respectively. Reproduced with permission from ref. 106. Copyright 2022, Elsevier.

the guideline, Pt-based electrocatalysts on N-doped CNT coated by CoP NPs with high specific surface areas (Pt-CoP-NCZ/CNT) have been developed by *in situ* growth of CoZn-ZIF on CNT materials. Pt-CoP-NCZ/CNT catalysts possess superior MOR activity and long-term stability in both acid medium and alkaline medium.¹⁰⁴ The excellent MOR performance can be mainly attributed to the N-doped structure and the introduction of Zn²⁺. The former benefits the assembly of Pt NPs, and the latter can suppress the Co agglomeration, thereby reducing CO poisoning. Similarly, Hu *et al.* have developed an *in situ* reduction-fusion method to synthesize different Pt-Co nanostructures embedded on porous N-doped carbon carriers (PtCo@NCs).¹⁰⁵ A pyrolysis process of ZnCo-MOF precursor was precisely controlled, in which the metallic Co source serves as the conductive carbon carrier. The PtCo@NCs manifests a mass activity of 2.30 A mg_{Pt}⁻¹ for MOR, demonstrating an excellent MOR performance. The CO tolerance of PtCo@NCs has been

significantly enhanced owing to the synergistic effect in the Pt-Co sites, which can accelerate the rate-determining step of methanol dehydrogenation and oxidative removal of CO. Encapsulating Pt NPs with carbon-based materials have been considered as one of the most promising strategies for inhibiting CO poisoning. Meanwhile, the CA test results showed that PtCo@NCs with precisely controlled nanostructures possessed the optimized durability compared to other PtCo catalysts and commercial Pt/C. Moreover, Jung *et al.* have fabricated PtRu NPs by using a molecular sieve effect, in which PtRu NPs were encapsulated by carbon shells with controlled pore thickness and size.¹⁰⁶ The carbon shell plays an important role in the hybrid system. On the one hand, carbon shell could efficiently promote the hydrogen oxidation reaction (HOR) by selectively permeating H₂ gas, thus decreasing CO poisoning effect (Fig. 7g). On the other hand, carbon shell also can effectively prevent the Ru dissolution, and the ligand effect by alloying Pt



NPs with Ru was still operative. Therefore, carbon shell-encapsulated PtRu NPs show the improved activity and stability (Fig. 7h and i).

Although the carbon carriers or carbon layers have achieved great progress in anti-CO poisoning, carbon-based materials are easily oxidized into CO or CO₂ in the electrochemical reactions. Then, the oxidized carbon carriers could cause the aggregation of Pt electrocatalysts. More seriously, CO produced from carbon-based materials will further accelerate the poisoning of Pt. To solve the above issues, increasing attention has been paid to investigating the mixture of carbon and semiconductor carriers. For example, Yang and his co-authors successfully synthesized the atomic carbon layers on the surface of SiC NPs by a facile dry-etching approach, and Pt NPs were then anchored on their surface.²² A porous carbon layer with an enlarged surface area can be generated by selectively removing silica atoms from the surface of SiC NPs. In addition, the etching time can determine the thickness of atomic carbon layers. Significantly, Pt/SiC catalysts show high MOR activity, long-term stability and enhanced anti-CO poisoning ability.¹⁰⁷ DFT calculation and XAFS measurements were used to clarify the catalytic mechanism. They found that the reduction of CO adsorption energy and enhanced amount of OH_{ad} species on the atomic carbon layers promote the high MOR performance. As a result, this novel catalyst can achieve the integration of maximum MOR performance with minimized CO poisoning.

Besides the success of fabricating carbon-based carriers, there are still challenges and urgent needs for developing advanced carbon-based materials: (1) Pt-based nanocrystals were generally physically mixed with carbon-based materials or deposited on the carriers, which leads to the random distribution between Pt NPs and carriers. (2) Carbon-based materials are prone to the structural changes during reaction process, such as corrosion and agglomeration, which is not beneficial for the catalytic stability. Future works need to focus on the structural stability of carbon-based materials. (3) The big electrolyte-electrode area can increase the side reactions.

3.2 Metal oxides or hydroxides supported Pt-based electrocatalysts

Based on the previous reports, combining Pt NPs with metal oxides has been widely used in practical applications because it can promote high activity and selectivity for significant electrochemical reactions. Therefore, metal oxides or hydroxides as the semiconductor carriers have received more and more attention due to their low cost and strong bonding force with Pt. More importantly, the synergy between Pt and metal oxides can weaken the binding strength of CO due to the formation of OH_{ad} and the modified electronic structure of Pt. However, as a semiconductor, the conductivity of metal oxides is very low, and they also cannot be operated in acidic solution. To solve the issues, a facile strategy has been developed to fabricate two-dimensional Pt/SnO₂/rGO nanocomposites by a facile photo-assisted reduction method.¹⁰⁸ In the work, the photo-excited electrons from SnO₂ NPs can promote the reduction of ultra-fine Pt NPs with an average size of 1–2 nm on the surface of

SnO₂. Therefore, Pt/SnO₂/rGO nanocomposites manifest higher MOR activity and better long-term stability for MOR compared to Pt/rGO catalysts. With the incorporation of SnO₂, the OH_{ad} species can be easily generated adjacent to the Pt active sites, thus benefiting the oxidative removal of CO-like species and then release the Pt active sites for further MOR. Similarly, Pt₃Sn NPs enriched with Pt₃Sn/SnO₂ interfaces have been synthesized by using a surface restructuring engineering under H₂ and air atmospheres.¹⁰⁹ The polyol method was first used to deposit Pt₂Sn NPs on nitrogen-doped graphene (NG). Next, SnO₂/Pt₃Sn interfaces were fabricated through thermal treatment method in a H₂ atmosphere, followed by annealing under air atmosphere. The incorporation of Sn could not only tune the Pt electronic structure but also increase the atomic distance of Pt–Pt, thereby lowering the CO affinity on the Pt active sites. Importantly, the ultra-small SnO₂ can promote the formation of OH_{ad} species at low potentials. Thanks to the unique structure, Pt₃Sn@u-SnO₂/NG catalysts exhibit superior EOR and MOR activity and excellent stability. Besides, other transition metal oxides, such as TiO₂, CeO₂, and Co₃O₄, were used as the optimal alternatives due to their strong metal-support interactions (SMSI).^{110–113} Among these metal oxides, TiO₂ carriers could fix Pt NPs on its surface, and tune the d-band center of Pt due to its wide band gap. More importantly, hydrophilic TiO₂ carriers are conducive to the formation of OH_{ad} in both acid and alkaline media. Based on the above, Yan *et al.* have fabricated novel porous UV-Pt@TiO₂/graphene hybrids by a simple hydrothermal synthesis.¹¹⁴ They found that TiO₂ nanorods with the (001) and (110) facets could strengthen their trapping ability for Pt NPs, thus improving their adsorption for methanol molecules and weakening the CO poisoning compared to TiO₂ NPs. At the same time, the strong metal-support interaction facilitates the adjustment of surface properties and electronic structure of Pt, thereby demonstrating superior catalytic activity and long-term durability for MOR.

Recently, considerable works have proved that sub-nanometre clusters possess better catalytic performance than nanometre-sized crystals. Therefore, combining metal oxides and Pt sub-nanometre clusters has been considered as the promising strategy for the improvement of CO tolerance on the Pt electrode. For example, Zhang *et al.* have fabricated single Pt atoms that uniformly disperses on a FeO_x carriers with a large surface area.¹¹⁰ DFT results confirm that the electron transfer from Pt atoms to FeO_x surface promotes the formation of vacant d orbitals. The generated vacant d orbitals not only benefit the strong binding and stability of single Pt atom, but also provide positively charged Pt atoms, leading to reducing the CO adsorption energy and the activation barrier of CO oxidation. As a result, Pt₁/FeO_x catalyst shows excellent catalytic activity. To investigate the catalytic activity in monatomic catalysts that really comes from the separated sub-atoms, Parkinson *et al.* have employed atomically resolved scanning probe microscopy, surface-sensitive spectroscopy, and DFT calculation to investigate CO oxidation on a Pt/Fe₃O₄ electrode.¹¹⁵ They have proved that individual CO molecules can adsorb on a sub-nanocluster, and the oxidation of CO can occur from the metastable (PtCO)₂. Generally, ultrathin structure (<2 nm) can benefit the exposure



of Pt active sites and high utilization of Pt. Moreover, Mo as an oxophilic metal could effectively alleviate the CO poisoning due to the formation of OH_{ad} at lower potential. To combine these two advantages, Guo *et al.* have reported a series of sub-monolayer YO_x/MoO_x -surface co-decorated Pt nanowires (YO_x/MoO_x -Pt NWs) by developing a new strategy (Fig. 8a).¹¹⁶ The forming mechanism of YO_x/MoO_x -Pt NWs has been investigated systematically by collecting the intermediates during the preparation process. The addition of CTAB plays an important role in the formation of Pt NWs, and carbonyl group decomposed from the $\text{Mo}(\text{CO})_6$ and its concentration determine the formation of ultrathin NWs and their uniformity, respectively (Fig. 8b–d). In addition, FTIR spectroscopy confirms that YO_x/MoO_x improves the anti-CO poisoning ability of YO_x/MoO_x -Pt NWs. Interestingly, the advanced systems can inhibit the CO poisoning of YO_x/MoO_x -Pt NWs by a new decoupling process, in which the free energy of CO_{ad} and COOH_{ad} can be decoupled on the YO_x/MoO_x -Pt NWs due to the high oxygen affinity of Mo and Y (Fig. 8e). The CO onset oxidation potential of 5% MoO_x -Pt and 22% YO_x/MoO_x -Pt shifts negatively compared to commercial Pt/C, indicating the enhanced CO tolerance due to the high affinity of Y and Mo to oxygen. Importantly, YO_x/MoO_x -Pt NWs can be well preserved after the stability tests, while commercial Pt/C exhibited severe aggregation under the same condition. As a result, it can be seen that both the enhanced CO tolerance and stable nanostructures benefit the excellent catalytic performance.

The oxophilic metals, such as Bi, Pb, Ru, and Rh, could effectively alleviate the CO poisoning due to the formation of OH_{ad} at lower potential.^{117–120} In particular, some of them (*e.g.* Bi) can interact with Pt to form intermetallic compounds with long-

range ordered nature and stable compositions. For example, hexagonal PtBi intermetallic nanoplates with the modification of trace Pb atoms ($\text{PtBi}@6.7\%\text{Pb}$) *via* a facile wet-chemistry method (Fig. 9).¹²⁰ DFT results confirm that Pb oxyhydroxide on the surface can not only promote the electron transfer but also inhibit the CO poisoning. Therefore, the mass activity of $\text{PtBi}@6.7\%\text{Pb}$ can reach $51.07 \text{ A mg}_{\text{Pt}}^{-1}$ with excellent durability and high CO tolerance in the alkaline electrolyte. At the same time, the p–d coupling of Bi and Pb sites can improve the MOR process with low reaction barriers. In addition, it has been reported that the incorporation of $\text{Bi}(\text{OH})_3$ species can adjust the geometric structure and downshift the d-band center of Pt. Especially, the surface decoration of $\text{Bi}(\text{OH})_3$ species can accelerate the decomposition of H_2O to generate OH_{ad} , resulting the removal of CO and other intermediates.^{63,121–123} On this basis, porous Pt nanoframes decorated with $\text{Bi}(\text{OH})_3$ species have been prepared by using a two-step method.¹²⁴ The porous structure can improve the utilization efficiency of Pt, and $\text{Bi}(\text{OH})_3$ species can enhance the anti-CO poisoning. CO stripping experiments and *in situ* ATR-IR confirm the key role of $\text{Bi}(\text{OH})_3$ species. Therefore, The mass activity of Pt– $\text{Bi}(\text{OH})_3$ catalyst is 13.5 times higher than that of Pt/C, and the long-term durability of EOR on the Pt– $\text{Bi}(\text{OH})_3$ is much stable in both alkaline and acidic condition. In recent years, many efforts have focused on the decoration of Pt surface with oxides or other compounds, thus forming inverse heterogeneous nanostructures with an inverse interface structure. Therefore, a bismuth oxyhydroxide ($\text{BiO}_x(\text{OH})_y$)-Pt inverse interface has been developed *via* an electrochemical reconstruction strategy.¹²⁵ The as-prepared ($\text{BiO}_x(\text{OH})_y$)-Pt inverse interface could induce the electron deficiency adjacent Pt active sites, resulting in decreasing the CO adsorption and strengthening

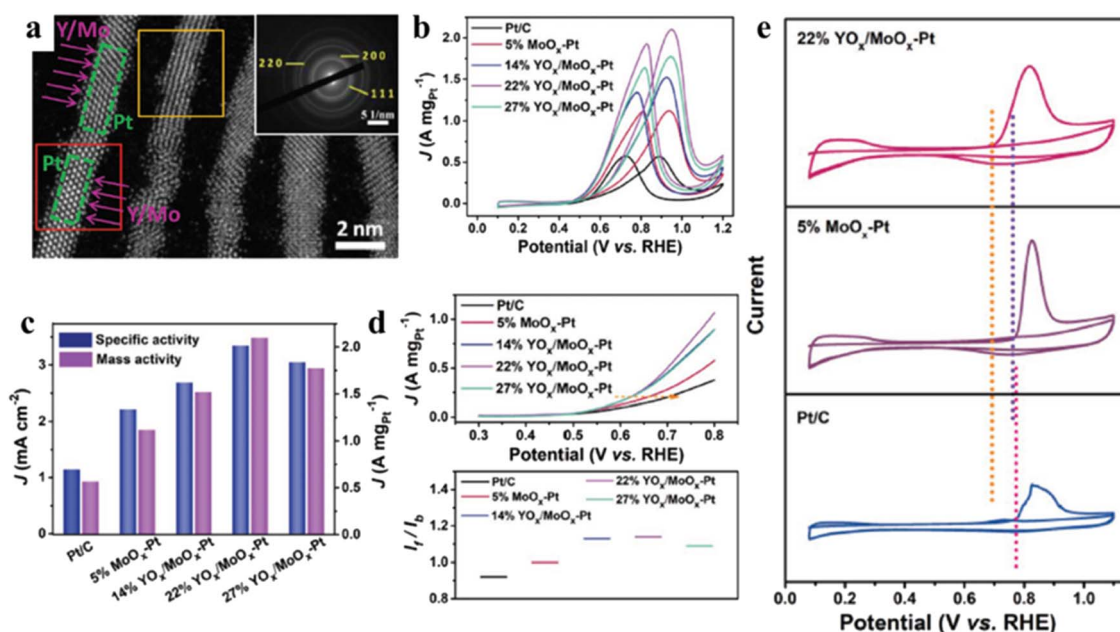


Fig. 8 (a) Atomic-resolution HAADF-STEM image (inset is the corresponding SAED pattern) of 22% YO_x/MoO_x -Pt NWs; (b) Pt mass-normalized MOR; (c) linear sweep voltammogram (LSV) curves, (d) the I_t/I_b ratios derived from the MOR and mass activities after different scan cycles. (e) CO stripping curves of 22% YO_x/MoO_x -Pt, 5% MoO_x -Pt, and commercial Pt/C. Reproduced with permission from ref. 116. Copyright 2021, Wiley-VCH.

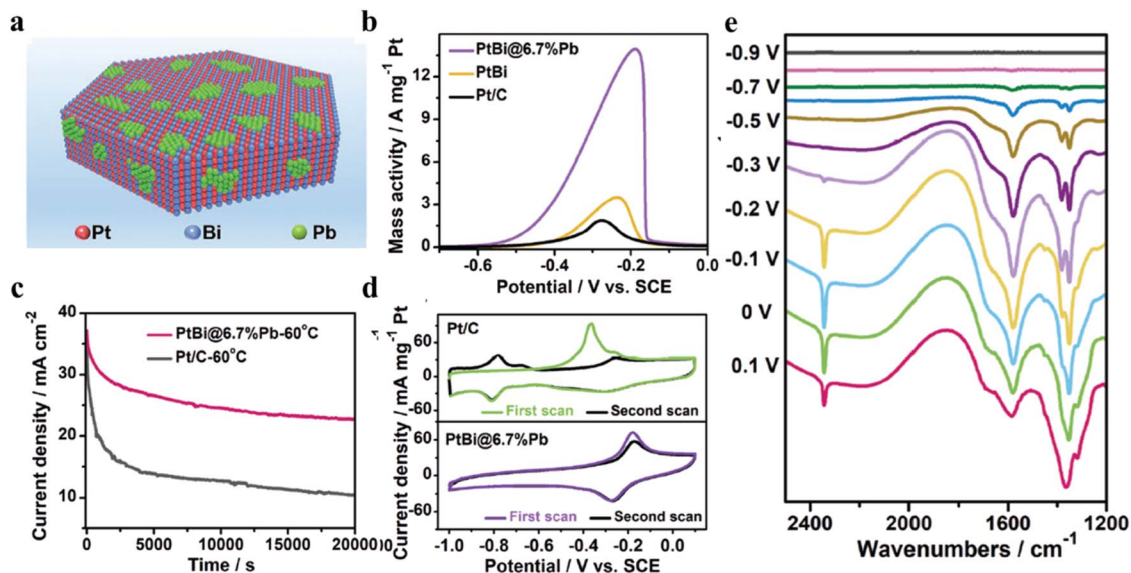


Fig. 9 (a) The schematic illustration showing the atomic arrangement of Pt/Bi/Pb atoms in the PtBi@6.7%Pb nanoplates; (b) the positive-going MOR polarization curves of different catalysts recorded at a scan rate of 50 mV s^{-1} ; (c) long-term chronoamperometry of PtBi@6.7%Pb nanoplates and Pt/C recorded at -0.3 V vs. SCE at 60°C ; (d) CO-stripping curves of PtBi@6.7% Pb nanoplates and Pt/C catalysts; (e) the *in situ* FTIR spectra of PtBi@6.7% Pb nanoplates recorded in Ar-saturated $1 \text{ M KOH} + 1 \text{ M CH}_3\text{OH}$ solution. Reproduced with permission from ref. 120. Copyright 2022, Wiley-VCH.

OH_{ad} adsorption, leading to high MOR activity and good stability. In general, this work emphasizes the importance of inverse interface structure and its electronic effect on Pt.

Although metal hydroxides show great advantages for CO tolerance on the Pt electrode, poor electronic conductivity is their inevitable defect, which can affect the electron transfer and limit the catalytic activity. In addition, the interaction between metal hydroxides and OH_{ad} species is neither too strong nor too weak, which can benefit the removal of CO intermediate. Based on the above, amorphous $\text{Ni}(\text{OH})_2$ and graphene have been mixed to support Pt NPs.¹²⁶ As a result, the Pt-based ternary hybrids show superior activity and durability for MOR. The incorporation of amorphous $\text{Ni}(\text{OH})_2$ species significantly accelerates the dissociative adsorption of H_2O and then promotes the oxidative removal of CO *via* the Langmuir–Hinshelwood reaction pathway. Similarly, three-dimensional porous Pt/ $\text{Ni}(\text{OH})_2$ /NG catalysts have been prepared by a facile route.¹²⁷ In the composites, the N-doped graphene structure can promote the electron transfer and electrolyte diffusion, and Pt NPs can expose more active sites. Importantly, $\text{Ni}(\text{OH})_2$ species provide a large number of reactive oxygen species. Therefore, the mass activity of Pt/ $\text{Ni}(\text{OH})_2$ /NG can reach $2.99 \text{ A mg}_{\text{Pt}}^{-1}$, and their catalytic durability and CO tolerance were also significantly improved due to the abundant OH_{ad} species adjacent to the Pt active sites.

Based on the above, the synergy between Pt and metal oxides or hydroxides can weaken the binding strength of CO due to the formation of OH_{ad} and the modified electronic structure of Pt. However, the conductivity of metal oxides or hydroxides as the semiconductor carriers is much lower compared to carbon-based carriers. The usage of carbon-based materials could effectively improve the conductivity of metal oxides or

hydroxides. Moreover, the structural stability of composites is its inherent defect. For example, Pt-based composites were generally physically mixed with carriers or deposited on the carriers, which leads to the random distribution between Pt NPs and carriers. Modifying metal oxides or hydroxides on the Pt surface could promote the formation of strong coupling interface structure, and no significant reduction of conductivity can be observed due to the trace metal oxides or hydroxides.

4. Summary and outlooks

In summary, fabricating Pt-based electrocatalysts with high CO tolerance is essential for the widespread commercialization of PEMFCs. Two effective strategies have been developed to improve the CO tolerance of Pt-based electrocatalysts, including alloying Pt with a second element and fabricating Pt-based composites with geometry and interface engineering. Alloying strategy not only could reduce the usage of Pt but also tune the Pt electronic structure based on the electron or ligand effects, including electron donation or reverse donation, thus reducing the strength of Pt–CO bond. Importantly, alloying strategy can effectively tune the geometric structure of Pt, which make Pt-based electrocatalysts exhibit a much stable nanostructure. Despite the success, the second element is easily oxidized and dissolved under long-term electrochemical reactions, thus leading to unstable nanostructures, which could reduce the CO tolerance of Pt-based alloys. In the Pt-based composites, the synergistic effect between Pt and carriers is believed to mainly optimize the Pt electronic structure and promote the formation of OH_{ad} species adjacent the Pt active sites. The former one could weaken the binding strength of CO, and the latter one is essential for the complete oxidation of CO intermediate. As



Table 1 Summary of the performance of Pt-based electrocatalysts in MOR

| Strategies | Catalysts | Electrolytes | Mass activity (A mg ⁻¹) | Ref. |
|--|--|---|-------------------------------------|------|
| Pt-based binary alloys | Pt ₇₂ Ru ₂₈ | 0.1 M HClO ₄ + 0.5 M CH ₃ OH | 1.70 | 44 |
| | PtRu/C(DMAB) | 0.5 M H ₂ SO ₄ + 1 M CH ₃ OH | 0.53 | 46 |
| | d-Pt ₂ Cu ₁ | 0.1 M HClO ₄ + 0.5 M CH ₃ OH | 2.23 | 50 |
| | Pt-Co nanoframes | 1 M KOH + 1 M CH ₃ OH | 4.28 | 51 |
| | PtNi CNCs | 0.5 M H ₂ SO ₄ + 0.5 M CH ₃ OH | 0.69 | 52 |
| | PtBi nanoplates/C | 0.1 M HClO ₄ + 0.1 M CH ₃ OH | 1.10 | 53 |
| | Pt ₉₅ Co ₅ | 0.5 M H ₂ SO ₄ + 1 M CH ₃ OH | 0.49 | 60 |
| | Pt ₃ Cu NBS | 0.5 M H ₂ SO ₄ + 1 M CH ₃ OH | 0.53 | 61 |
| | PtAg nanoflowers | 0.1 M HClO ₄ + 0.5 M CH ₃ OH | 1.13 | 64 |
| | PtBi nanorings | 1 M KOH + 1 M CH ₃ OH | 6.42 | 65 |
| Pt-based multi-element alloys | h-PtNiCo/C | 0.1 M HClO ₄ + 0.5 M CH ₃ OH | 2.82 | 73 |
| | Pt ₃ Bi ₃ Zn NPLs | 0.1 M HClO ₄ + 0.5 M CH ₃ OH | 3.29 | 74 |
| | Ru-ca-PtNi | 0.1 M HClO ₄ + 0.5 M CH ₃ OH | 2.01 | 75 |
| | PtCoNiRh NWs | 0.1 M HClO ₄ + 0.5 M CH ₃ OH | 1.36 | 76 |
| | Cu-Ni@Pt-Cu | 0.1 M HClO ₄ + 1 M CH ₃ OH | 0.99 | 80 |
| | Au@Pt-Pd | 1 M KOH + 1 M CH ₃ OH | 4.38 | 81 |
| | Pt ₁₈ Ni ₂₆ Fe ₁₅ Co ₁₄ Cu ₂₇ | 1 M KOH + 1 M CH ₃ OH | 15.04 | 87 |
| Pt-based high-entropy alloys | PtRhBiSnSb | 1 M KOH + 1 M CH ₃ OH | 19.53 | 88 |
| | NiPtSAA/GDY | 1 M KOH + 1 M CH ₃ OH | 4.40 | 99 |
| | PtRu/PC-H | 0.1 M HClO ₄ + 0.5 M CH ₃ OH | 1.67 | 100 |
| | CoP-NCZ/CNT | 1 M KOH + 1 M CH ₃ OH | 3.62 | 104 |
| | PtCo@NCs | 0.1 M HClO ₄ + 1 M CH ₃ OH | 2.30 | 105 |
| Carbon-supported Pt-based electrocatalysts | Pt/SiC | 0.5 M H ₂ SO ₄ + 0.5 M CH ₃ OH | 0.65 | 22 |
| | Pt/SnO ₂ /rGO | 0.5 M H ₂ SO ₄ + 0.5 M CH ₃ OH | 0.63 | 108 |
| Metal oxides or hydroxides supported Pt-based electrocatalysts | UV-Pt@TiO ₂ /graphene | 0.5 M H ₂ SO ₄ + 1 M CH ₃ OH | 1.94 | 114 |
| | YO _x /MoO _x -Pt NWs | 0.1 M HClO ₄ + 0.5 M CH ₃ OH | 2.1 | 116 |
| | PtBi@6.7%Pb | 1 M KOH + 1 M CH ₃ OH | 13.93 | 120 |
| | Pt/Ni(OH) ₂ /NG | 1 M KOH + 1 M CH ₃ OH | 2.99 | 127 |

a result, Pt-based electrocatalysts synthesized by two strategies have excellent electrochemical performance, as shown in Table 1.

Despite the great progress has been achieved in Pt-based alloys and composites as highly effective catalysts against CO poisoning, there are still great challenges that need to be resolved: (1) it is difficult to choose the type and number of metallic elements and precisely control their proportions. (2) Strict environmental conditions, such as acid solution, alkali solution and high temperature, can easily cause the agglomeration and elemental dissolution of Pt-based alloys and composites. (3) The limitations of computational software and models make it difficult to theoretically calculate the top-level design of Pt-based electrocatalysts. (4) The complexity of surface compositions, phase state, and catalytic reactions makes it difficult to monitor the real-time evolution behavior on the Pt electrode.

Based on the above discussion, controllable synthesis of Pt-based electrocatalysts and its investigation of CO tolerance can be mainly focused on five aspects, as follows: (1) the optimal catalytic system can be constructed by means of theoretical calculation, including the selection of type, quantity and proportion of metallic elements. (2) Focusing on the design and synthesis of stable structures, including intermetallic compounds, high-entropy alloys, and strong coupling interface

structures *etc.* (3) Combining *in situ* characterizations, the real-time structural evolution process can be observed, thereby guaranteeing the detection of the intrinsic active sites and catalytic mechanism. (4) It is also important to design and develop simpler synthesis processes to reduce costs and improve the utilization of components in the catalyst. (5) Large-scale synthesis paths need to be developed for the next-generation fuel cells. Based on the study of Pt-based electrocatalysts, we believe that many effective approaches can be developed to design Pt-based electrocatalysts with high CO tolerance, which can be widely used in PEMFCs in the future.

Conflicts of interest

There are no conflicts to declare.

References

- 1 M. K. Debe, *Nature*, 2012, **486**, 43–51.
- 2 C. Z. Li, Q. Yuan, B. Ni, T. He, S. M. Zhang, Y. Long, L. Gu and X. Wang, *Nat. Commun.*, 2018, **9**, 3702.
- 3 S. Wang and S. P. Jiang, *Natl. Sci. Rev.*, 2017, **4**, 163–166.
- 4 H. Yang, A. C. Zhang, Y. F. Bai, M. Y. Chu, H. Li, Y. Liu, P. Zhu, X. L. Chen, C. W. Deng and X. L. Yuan, *Inorg. Chem.*, 2022, **61**, 14419–14427.



- 5 M. A. Abdelkareem, K. Elsaid, T. Wilberforce, M. Kamil, E. T. Sayed and A. Olabi, *Nat. Energy*, 2021, **6**, 451.
- 6 J. N. Zhou, C. H. Zeng, H. Ou, Q. Y. Yang, Q. Y. Xie, A. Zeb, X. M. Lin, Z. Ali and L. Hu, *J. Mater. Chem. C*, 2021, **9**, 11030–11058.
- 7 S. N. Upadhyay and S. Pakhira, *J. Mater. Chem. C*, 2021, **9**, 11331–11342.
- 8 F. Xiao, Y. C. Wang, Z.-P. Wu, G. Chen, F. Yang, S. Zhu, K. Siddharth, Z. Kong, A. Lu, J. C. Li, C. J. Zhong, Z. Y. Zhou and M. Shao, *Adv. Mater.*, 2021, **33**, 2006292.
- 9 J. Tollefson, *Nature*, 2008, **451**, 877.
- 10 J. Choi, J. Cho, C. W. Roh, B. S. Kim, M. S. Choi, H. Jeong, H. C. Ham and H. Lee, *Appl. Catal., B*, 2019, **247**, 142–149.
- 11 L. C. Lin, C. H. Kuo, Y. H. Hsu, L. C. Hsu, H. Y. Chen, J. L. Chen and Y. T. Pan, *Appl. Catal., B*, 2022, **317**, 121767.
- 12 M. Gong, D. Xiao, Z. Deng, R. Zhang, W. Xia, T. Zhao, X. Liu, T. Shen, Y. Hu, Y. Lu, X. Zhao, H. Xin and D. Wang, *Appl. Catal., B*, 2020, **282**, 119617.
- 13 J. Zhai, J. Gao, J. Zhang, D. Liu, S. Gao, Y. Yan, K. Zhang, K. Cai, F. Yu, M. Lin and J. Li, *Chem. Eng. J.*, 2022, **442**, 136172.
- 14 Y. Lu, S. Du and R. Steinberger-Wilckens, *Appl. Catal., B*, 2016, **187**, 108–114.
- 15 Y. X. Xiao, J. Ying, G. Tian, Y. Tao, H. Wei, S. Y. Fan, Z. H. Sun, W. J. Zou, J. Hu, G. G. Chang, W. Li, X. Y. Yang and C. Janiak, *Appl. Catal., B*, 2019, **258**, 118080.
- 16 Y. Y. Zhou, C. H. Liu, J. Liu, X. L. Cai, Y. Lu, H. Zhang, X. H. Sun and S. D. Wang, *Nano-Micro Lett.*, 2016, **8**, 371–380.
- 17 T. W. Song, M. X. Chen, P. Yin, L. Tong, M. Zuo, S. Q. Chu, P. Chen and H. W. Liang, *Small*, 2022, **18**, 2202916.
- 18 J. Lee, C. Yeon, J. Oh, G. Han, J. Do Yoo, H. J. Yun, C. W. Lee, K. T. Lee and J. Bae, *Appl. Catal., B*, 2022, **316**, 121645.
- 19 D. Sebastián, I. Suelves, E. Pastor, R. Moliner and M. J. Lázaro, *Appl. Catal., B*, 2012, **132–133**, 13–21.
- 20 X. Yang, J. Zheng, M. Zhen, X. Meng, F. Jiang, T. Wang, C. Shu, L. Jiang and C. Wang, *Appl. Catal., B*, 2012, **121–122**, 57–64.
- 21 Y. M. Xi, Y. Zhang, X. T. Cai, Z. X. Fan, K. F. Wang, W. R. Dong, Y. Shen, S. X. Zhong, L. Yang and S. Bai, *Appl. Catal., B*, 2022, **305**, 121069.
- 22 G. Bai, C. Liu, Z. Gao, B. Lu, X. Tong, X. Guo and N. Yang, *Small*, 2019, **15**, 202951.
- 23 W. Chen, J. Cao, W. Fu, J. Zhang, G. Qian, J. Yang, D. Chen, X. Zhou, W. Yuan and X. Duan, *Angew. Chem., Int. Ed.*, 2022, **61**, e202200190.
- 24 J. Liu, F. R. Lucci, M. Yang, S. Lee, M. D. Marcinkowski, A. J. Therrien, C. T. Williams, E. C. H. Sykes and M. Flytzani-Stephanopoulos, *J. Am. Chem. Soc.*, 2016, **138**, 6396–6399.
- 25 P. Ferrari, L. M. Molina, V. E. Kaydashev, J. A. Alonso, P. Lievens and E. Janssens, *Angew. Chem., Int. Ed.*, 2016, **55**, 11059–11063.
- 26 J. Zhang, S. Deo, M. J. Janik and J. W. Medlin, *J. Am. Chem. Soc.*, 2020, **142**, 5184–5193.
- 27 P. Trens, R. Durand, B. Coq, C. Coutanceau, S. Rousseau and C. Lamy, *Appl. Catal., B*, 2009, **92**, 280–284.
- 28 V. F. Valdés-López, T. Mason, P. R. Shearing and D. J. L. Brett, *Prog. Energy Combust. Sci.*, 2020, **79**, 100842.
- 29 M. J. Hazlett, M. Moses-Debusk, J. E. Parks, L. F. Allard and W. S. Epling, *Appl. Catal., B*, 2016, **202**, 404–417.
- 30 J. Yeston, *Science*, 2016, **9**, 2155–5435.
- 31 H. Chen, J. Liu, X. Wu, C. Ye, J. Zhang, J. L. Luo and X.-Z. Fu, *Small*, 2022, **18**, 202204100.
- 32 A. M. Baena-Moncada, A. Bazan-Aguilar, E. Pastor and G. Á. Planes, *J. Power Sources*, 2019, **437**, 226915.
- 33 M. Su, J. C. Dong, J. B. Le, Y. Zhao, W. M. Yang, Z. L. Yang, G. Attard, G. K. Liu, J. Cheng, Y. M. Wei, Z. Q. Tian and J. F. Li, *Angew. Chem., Int. Ed.*, 2020, **59**, 23554–23558.
- 34 L. Huang, S. Zaman, X. Tian, Z. Wang, W. Fang and B. Y. Xia, *Acc. Chem. Res.*, 2021, **54**, 311–322.
- 35 W. Liang, Y. Wang, L. Zhao, W. Guo, D. Li, W. Qin, H. Wu, Y. Sun and L. Jiang, *Adv. Mater.*, 2021, **33**, e2100713.
- 36 Z. Wang, H. Wang, Z. Zhang, G. Yang, T. He, Y. Yin and M. Jin, *ACS Nano*, 2017, **11**, 163–170.
- 37 S. Q. Liu, H. R. Wen, G. Ying, Y. W. Zhu, X. Z. Fu, R. Sun and C. P. Wong, *Nano Energy*, 2018, **44**, 7–14.
- 38 A. L. Wang, L. Zhu, Q. Yun, S. Han, L. Zeng, W. Cao, X. Meng, J. Xia and Q. Lu, *Small*, 2020, **16**, e2003782.
- 39 C. Fang, T. Bi, Q. Ding, Z. Cui, N. Yu, X. Xu and B. Geng, *ACS Appl. Mater. Interfaces*, 2019, **11**, 20117–20124.
- 40 A. L. Wang, H. Xu, J. X. Feng, L. X. Ding, Y. X. Tong and G. R. Li, *J. Am. Chem. Soc.*, 2013, **135**, 10703–10709.
- 41 A. Rodríguez-Gómez, E. Lepre, L. Sánchez-Silva, N. López-Salas and A. Raquelde la Osa, *J. Energy Chem.*, 2022, **66**, 168–180.
- 42 J. S. Guo, G. Q. Sun, S. G. Sun, S. Y. Yan, W. Q. Yang, J. Qi, Y. S. Yan and Q. Xin, *J. Power Sources*, 2007, **168**, 299–306.
- 43 S. L. Lu, K. Eid, D. H. Ge, J. Guo, L. Wang, H. J. Wang and H. W. Gu, *Nanoscale*, 2017, **9**, 1033–1039.
- 44 W. Y. Zhao, B. Ni, Q. Yuan, P. L. He, Y. Gong, L. Gu and X. Wang, *Adv. Energy Mater.*, 2017, **7**, 1601593.
- 45 J. M. Zhang, X. M. Qu, Y. Han, L. F. Shen, S. H. Yin, G. Li, Y. X. Jiang and S. G. Sun, *Appl. Catal., B*, 2020, **263**, 118345.
- 46 Q. J. Wang, Y. W. Zhou, Z. Jin, C. G. Chen, H. Li and W. B. Cai, *Catalysts*, 2021, **11**, 925.
- 47 L. Huang, X. P. Zhang, Q. Q. Wang, Y. J. Han, Y. X. Fang and S. J. Dong, *J. Am. Chem. Soc.*, 2018, **140**, 1142–1147.
- 48 T. Fu, J. Huang, S. Lai, S. Zhang, J. Fang and J. Zhao, *J. Power Sources*, 2017, **365**, 17–25.
- 49 Y. S. Kang, K. H. Choi, D. Ahn, M. J. Lee, J. Baik, D. Y. Chung, M. J. Kim, S. Y. Lee, M. Kim, H. Shin, K. S. Lee and Y.-E. Sung, *J. Power Sources*, 2016, **303**, 234–242.
- 50 K. L. Wang, D. Y. Huang, Y. C. Guan, F. Liu, J. He and Y. Ding, *ACS Catal.*, 2021, **11**, 14428–14438.
- 51 S. P. Chen, M. F. Li, M. Y. Gao, J. B. Jin, M. A. van Spronsen, M. B. Salmeron and P. D. Yang, *Nano Lett.*, 2020, **3**, 1974–1979.
- 52 P. P. Yang, X. L. Yuan, H. C. Hu, Y. L. Liu, H. W. Zheng, D. Yang, L. Chen, M. H. Cao, Y. Xu, Y. L. Min, Y. G. Li and Q. Zhang, *Adv. Funct. Mater.*, 2018, **28**, 1704774.



- 53 X. L. Yuan, X. J. Jiang, M. H. Cao, L. Chen, K. Q. Nie, Y. Zhang, Y. Xu, X. H. Sun, Y. G. Li and Q. Zhang, *Nano Res.*, 2019, **12**, 429–436.
- 54 Y. N. Qin, M. C. Luo, Y. J. Sun, C. J. Li, B. L. Huang, Y. Yang, Y. J. Li, L. Wang and S. J. Guo, *ACS Catal.*, 2018, **8**, 5581–5590.
- 55 M. Choi, C. Ahn, H. Lee, J. Kim, S. Oh, W. Hwang, S. Yang, J. Kim, O. Kim, I. Choi, Y. Sung, Y. Cho, C. Rhee and W. Shin, *Appl. Catal., B*, 2019, **253**, 187–195.
- 56 M. Y. Chu, J. L. Huang, J. Gong, Y. Qu, G. L. Chen, H. Yang, X. C. Wang, Q. X. Zhong, C. W. Deng, M. H. Cao, J. X. Chen, X. L. Yuan and Q. Zhang, *Nano Res.*, 2022, **15**, 3920–3926.
- 57 Y. C. Yan, H. Shan, G. Li, F. Xiao, Y. Y. Jiang, Y. Y. Yan, C. H. Jin, H. Zhang, J. B. Wu and D. R. Yang, *Nano Lett.*, 2016, **16**, 7999–8004.
- 58 A. Mahmood, D. Q. He, S. H. Talib, Y. He, Z. Q. Song, Z. B. Li, D. X. Han and L. Niu, *Adv. Funct. Mater.*, 2022, **32**, 2205223.
- 59 E. B. Zhu, X. C. Yan, S. Y. Wang, M. J. Xu, C. Wang, H. T. Liu, J. Huang, W. Xue, J. Cai, H. Heinz, Y. J. Li and Y. Huang, *Nano Lett.*, 2019, **19**, 3730–3736.
- 60 Q. Q. Lu, L. T. Sun, X. Zhao, J. S. Huang, C. Han and X. R. Yang, *Nano Res.*, 2018, **11**, 2562–2572.
- 61 Z. B. Zhang, J. Li, S. Y. Liu, X. Zhou, L. Xu, X. L. Tian, J. Yang and Y. W. Tang, *Small*, 2022, **18**, 2202782.
- 62 D. Kim, M. Chung, S. Kim, K. Yun, W. Cha, R. Harder and H. Kim, *Nano Lett.*, 2019, **19**, 5044–5052.
- 63 Y. J. Chen, J. J. Pei, Z. Chen, A. Li, S. F. Ji, H. P. Rong, Q. Xu, T. Wang, A. J. Zhang, H. L. Tang, J. F. Zhu, X. D. Han, Z. B. Zhuang, G. Zhou and D. S. Wang, *Nano Lett.*, 2022, **22**, 7563–7571.
- 64 M. Qiao, H. Wu, F. Y. Meng, Z. B. Zhuang and J. X. Wang, *Small*, 2022, **18**, 2106643.
- 65 S. M. Han, Y. Ma, Q. B. Yun, A. L. Wang, Q. S. Zhu, H. Zhang, C. H. He, J. Xia, X. M. Meng, L. Gao, W. B. Cao and Q. P. Lu, *Adv. Funct. Mater.*, 2022, 2208760.
- 66 L. Zhang, H. S. Liu, S. H. Liu, M. N. Banis, Z. X. Song, J. J. Li, L. J. Yang, M. Markiewicz, Y. Zhao, R. Y. Li, M. Zheng, S. Y. Ye, Z. J. Zhao, G. A. Botton and X. L. Sun, *ACS Catal.*, 2019, **9**, 9350–9358.
- 67 A. R. Poerwoprajitno, L. Gloag, J. Watt, S. Cheong, X. Tan, H. Lei, H. A. Tahini, A. Henson, B. Subhash, N. M. Bedford, B. K. Miller, P. B. O'Mara, T. M. Benedetti, D. L. Huber, W. Zhang, S. C. Smith, J. J. Gooding, W. Schuhmann and R. D. Tilley, *Nat. Catal.*, 2022, **5**, 231–237.
- 68 X. F. Ren, Q. Y. Lv, L. F. Liu, B. H. Liu, Y. Wang, A. M. Liu and G. Wu, *Sustainable Energy Fuels*, 2020, **4**, 15–30.
- 69 Z. W. Seh, J. Kibsgaard, C. F. Dickens, I. Chorkendorff, J. K. Nørskov and T. F. Jaramillo, *Science*, 2017, **355**, eaad4998.
- 70 X.-X. Li, P.-Y. Zhu, Q. Li, Y.-X. Xu, Y. Zhao and H. Pang, *Rare Met.*, 2020, **39**, 680–687.
- 71 J. Liang, N. Li, Z. Zhao, L. Ma, X. Wang, S. Li, X. Liu, T. Wang, Y. Du, G. Lu, J. Han, Y. Huang, D. Su and Q. Li, *Angew. Chem., Int. Ed.*, 2019, **58**, 2–9.
- 72 R. M. Aran-Ais, F. Dionigi, T. Merzdorf, M. Gocyla, M. Heggen, R. E. Dunin-Borkowski, M. Gliech, J. Solla-Gullon, E. Herrero, J. M. Feliu and P. Strasser, *Nano Lett.*, 2015, **15**, 7473–7480.
- 73 H. R. Ma, Z. P. Zheng, H. S. Zhao, C. Shen, H. M. Chen, H. Q. Li, Z. M. Cao, Q. Kuang, H. X. Lin and Z. X. Xie, *J. Mater. Chem. A*, 2021, **9**, 23444.
- 74 H. Tian, D. X. Wu, J. Li, J. M. Luo, C. M. Jia, Z. X. Liu, W. Huang, Q. Chen, C. M. Shim, P. L. Deng, Y. J. Shen and X. L. Tian, *J. Energy Chem.*, 2022, **70**, 230–235.
- 75 F. P. Kong, X. Z. Liu, Y. J. Song, Z. Y. Qian, J. J. Li, L. Zhang, G. P. Yin, J. J. Wang, D. Su and X. L. Sun, *Angew. Chem., Int. Ed.*, 2022, **61**, e202207524.
- 76 W. Wang, X. W. Chen, X. Zhang, J. Y. Ye, F. Xue, C. Zhen, X. Y. Liao, H. Q. Li, P. T. Li, M. C. Liu, Q. Kuang, Z. X. Xie and S. F. Xie, *Nano Energy*, 2020, **71**, 104623.
- 77 S. Xie, S.-I. Choi, N. Lu, L. T. Roling, J. A. Herron, L. Zhang, J. Park, J. Wang, M. J. Kim, Z. Xie, M. Mavrikakis and Y. Xia, *Nano Lett.*, 2014, **14**, 3570–3576.
- 78 J. Park, L. Zhang, S.-I. Choi, L. T. Roling, N. Lu, J. A. Herron, S. Xie, J. Wang, M. J. Kim, M. Mavrikakis and Y. Xia, *ACS Nano*, 2015, **9**, 2635–2647.
- 79 Y. Xia, K. D. Gilroy, H. C. Peng and X. Xia, *Angew. Chem., Int. Ed.*, 2017, **56**, 60–95.
- 80 C. Li, X. B. Chen, L. H. Zhang, S. H. Yan, A. Sharma, B. Zhao, A. Kumbhar, G. W. Zhou and J. Y. Fang, *Angew. Chem., Int. Ed.*, 2021, **60**, 7675–7680.
- 81 W. K. Liang, Y. W. Wang, L. Zhao, W. Guo, D. Li, W. Qin, H. H. Wu, Y. H. Sun and L. Jiang, *Adv. Mater.*, 2021, **33**, 2100713.
- 82 H. Peng, J. Ren, Y. Wang, Y. Xiong, Q. Wang, Q. Li, X. Zhao, L. Zhan, L. Zheng and Y. Tang, *Nano Energy*, 2021, **88**, 106307.
- 83 G. L. Zhang, D. J. Cao, S. Y. Guo, Y. Fang, Q. Wang, S. Cheng, W. S. Zuo, Z. Z. Yang and P. Cui, *Small*, 2022, **18**, 2202587.
- 84 P. Wang, Y. Q. Bu, J. B. Liu, Q. Q. Li, H. T. Wang and W. Yang, *Mater. Today*, 2020, **37**, 64–73.
- 85 Z. W. Chen, L. X. Chen, Z. Garipey, X. Yao and C. V. Singh, *Trends Chem.*, 2022, **4**, 577–579.
- 86 M. Bondesgaard, N. L. N. Broge, A. Mamakhel, M. Bremholm and B. B. Iversen, *Adv. Funct. Mater.*, 2019, **29**, 1905933.
- 87 H. D. Li, Y. Han, H. Zhao, W. J. Qi, D. Zhang, Y. D. Yu, W. W. Cai, S. X. Li, J. P. Lai, B. L. Huang and L. Wang, *Nat. Commun.*, 2020, **11**, 5437.
- 88 W. Chen, S. P. Luo, M. Z. Sun, X. Y. Wu, Y. S. Zhou, Y. J. Liao, M. Tang, X. K. Fan, B. L. Huang and Z. W. Quan, *Adv. Mater.*, 2022, **34**, 2206276.
- 89 D. S. Wu, K. H. Kusada, T. Yamamoto, T. Toriyama, S. Matsumura, S. Kawaguchi, Y. Kubota and H. Kitagawa, *J. Am. Chem. Soc.*, 2020, **142**, 13833–13838.
- 90 Y. J. Wang, W. Y. Long, L. L. Wang, R. S. Yuan, A. Ignaszak, B. Z. Fang and D. P. Wilkinson, *Energy Environ. Sci.*, 2018, **11**, 258–275.
- 91 B. W. Zhang, Y. X. Jiang, J. Ren, X. M. Qu, G. L. Xu and S. G. Sun, *Electrochim. Acta*, 2015, **162**, 254–262.



- 92 J. Zheng, G. Li, J. M. Zhang, N. Y. Cheng, L. F. Ji, J. Yang, J. L. Zhang, B. W. Zhang, Y. X. Jiang and S. G. Sun, *ACS Catal.*, 2021, **11**, 9317–9332.
- 93 J. Zheng, G. Li, J. M. Zhang, N. Y. Cheng, L. F. Ji, J. Yang, J. L. Zhang, B. W. Zhang, Y. X. Jiang and S. G. Sun, *Sci. China: Chem.*, 2023, **66**, 279–288.
- 94 K. Kodama, T. Nagai, A. Kuwaki, R. Jinnouchi and Y. Morimoto, *Nat. Nanotechnol.*, 2021, **16**, 140–147.
- 95 C.-Y. Ahn, J. E. Park, S. Kim, O.-H. Kim, W. Hwang, M. Her, S. Y. Kang, S. B. Park, O. J. Kwon, H. S. Park, Y. Cho and Y.-E. Sung, *Chem. Rev.*, 2021, **121**, 15075–15140.
- 96 X. Guo, M. G. Lia, L. Y. Qiu, F. Y. Tian, L. He, S. Geng, Y. Q. Qiu, Y. Song, W. W. Yang and Y. S. Yu, *Chem. Eng. J.*, 2023, **453**, 1397996.
- 97 Q. C. Tran, V.-D. Dao, H. Y. Kim, K.-D. Jun and H.-S. Choi, *Appl. Catal., B*, 2021, **204**, 365–373.
- 98 G. Han, M. G. Li, W. Y. Zhang, L. He, F. Y. Tian, Y. Q. Liu, Y. S. Yu, W. W. Yang and S. J. Guo, *Adv. Mater.*, 2022, **34**, 2202943.
- 99 L. Hui, Y. R. Xue, C. Y. Xing, Y. X. Liu, Y. C. Du, Y. Fang, H. D. Yu, C. Zhang, F. He and Y. L. Li, *Nano Energy*, 2022, **95**, 106984.
- 100 J. M. Zhang, X. M. Qu, Y. Han, L. F. Shen, S. H. Yin, G. Li, Y. X. Jiang and S. G. Sun, *Appl. Catal., B*, 2020, **263**, 118345.
- 101 Q. Zhang, F. Yue, L. J. Xu, C. X. Yao, R. D. Priestley and S. F. Hou, *Appl. Catal., B*, 2019, **257**, 117886.
- 102 M. M. Yan, Q. G. Jiang, T. Zhang, J. Y. Wang, L. Yang, Z. Y. Lu, H. Y. He, Y. S. Fu, X. Wang and H. J. Huang, *J. Mater. Chem. A*, 2018, **6**, 18165–18172.
- 103 J. Fonseca, T. H. Gong, L. Jiao and H. L. Jiang, *J. Mater. Chem. A*, 2021, **9**, 10562–10611.
- 104 W. Zhan, L. Ma, M. Y. Gan, J. J. Ding, S. C. Han, D. Y. Wei, J. Shen and C. L. Zhou, *Int. J. Hydrogen Energy*, 2020, **45**, 15630–15641.
- 105 G. F. Hu, L. Shang, T. Sheng, Y. G. Chen and L. Y. Wang, *Adv. Funct. Mater.*, 2020, **30**, 2002281.
- 106 J. Min, S. Kim, A. A. Jeffery, H. Shin, Y. S. Kang, Y. Kim, J. Jang, S. Lee, S. Park, G. Park, S. J. Yoo, S. Yim and N. Jung, *Mater. Today Energy*, 2022, **29**, 101124.
- 107 G. L. Bai, C. Liu, Z. Gao, B. Y. Lu, X. L. Tong, X. Y. Guo and N. J. Yang, *Small*, 2019, **15**, 1902951.
- 108 S. L. Wu, J. Liu, D. W. Liang, H. M. Sun, Y. X. Ye, Z. F. Tian and C. H. Liang, *Nano Energy*, 2016, **26**, 699–707.
- 109 Z. C. Wang, L. Wang, W. B. Zhu, T. Zeng, W. Wu, Z. Lei, Y. Y. Tan, H. F. Lv and N. C. Cheng, *Nanoscale Adv.*, 2021, **3**, 5062–5067.
- 110 B. T. Qiao, A. Q. Wang, X. F. Yang, L. F. Allard, Z. Jiang, Y. T. Cui, J. Y. Liu, J. Li and T. Zhang, *Nat. Chem.*, 2011, **3**, 634–641.
- 111 M. Y. Gao, F. Y. Tian, Z. Guo, X. Zhang, Z. J. Li, J. Zhou, X. Zhou, Y. S. Yu and W. W. Yang, *Chem. Eng. J.*, 2022, **446**, 137127.
- 112 B. D. Chandler, *Nat. Chem.*, 2017, **9**, 108–109.
- 113 Z. X. Luo, G. Q. Zhao, H. G. Pan and W. P. Sun, *Adv. Energy Mater.*, 2022, **12**, 2201395.
- 114 K. F. Zhang, J. Qiu, J. Wu, Y. Q. Deng, Y. H. Wu and L. F. Yan, *J. Mater. Chem. A*, 2022, **10**, 4254.
- 115 M. Meier, J. Hulva, Z. Jakub, F. Kraushofer, M. Bobić, R. Bliem, M. Setvin, M. Schmid, U. Diebold, C. Franchini and G. S. Parkinson, *Sci. Adv.*, 2022, **8**, eabn4580.
- 116 M. G. Li, Z. L. Zhao, W. Y. Zhang, M. C. Luo, L. Tao, Y. J. Sun, Z. H. Xia, Y. G. Chao, K. Yin, Q. H. Zhang, L. Gu, W. W. Yang, Y. S. Yu, G. Lu and S. J. Guo, *Adv. Mater.*, 2021, **33**, 2103762.
- 117 Y. X. Liang, Y. J. Sun, X. Y. Wang, E. G. Fu, J. Zhang, J. L. Du, X. D. Wen and S. J. Guo, *Nanoscale*, 2018, **10**, 11357–11364.
- 118 S. P. Chen, M. F. Li, X. Zhang, Z. J. Li, J. Zhou, X. Zhou, Y. S. Yu and W. W. Yang, *Nano Res.*, 2022, **15**, 1809–1816.
- 119 Y. J. Sun, Y. X. Liang, M. C. Luo, F. Lv, Y. N. Qin, L. Wang, C. Xu, E. G. Fu and S. J. Guo, *Small*, 2017, **14**, 1702259.
- 120 W. Chen, S. P. Luo, M. Z. Sun, M. Tang, X. K. Fan, Y. Cheng, X. Y. Wu, Y. J. Liao, B. L. Huang and Z. W. Quan, *Small*, 2022, **18**, 2107803.
- 121 B. Lan, Q. L. Wang, Z. X. Ma, Y. J. Wu, X. L. Jiang, W. S. Jia, C. X. Zhou and Y. Y. Yang, *Appl. Catal., B*, 2022, **300**, 120728.
- 122 M. G. Li, Z. G. Xia, M. C. Luo, L. He, L. Tao, W. W. Yang, Y. S. Yu and S. J. Guo, *Small Sci.*, 2021, **13**, 2100061.
- 123 Y. L. Yuan, Q. Y. Wang, Y. Qiao, X. L. Chen, Z. L. Yang, W. C. Lai, T. W. Chen, G. H. Zhang, H. G. Duan, M. Liu and H. W. Huang, *Adv. Energy Mater.*, 2022, **12**, 2200970.
- 124 X. L. Yuan, B. Jiang, M. H. Cao, C. Y. Zhang, X. Z. Liu, Q. H. Zhang, F. L. Lyu, L. Gu and Q. Zhang, *Nano Res.*, 2020, **13**, 265–272.
- 125 X. C. Wang, M. Xie, F. L. Lyu, Y. M. Yiu, Z. Q. Wang, J. T. Chen, L. Y. Chang, Y. J. Xia, Q. X. Zhong, M. Y. Chu, H. Yang, T. Cheng, T. K. Sham and Q. Zhang, *Nano Lett.*, 2020, **20**, 7751–7759.
- 126 G. Yuan, L. Wang, X. W. Zhang and Q. F. Wang, *J. Colloid Interface Sci.*, 2019, **536**, 189–195.
- 127 K. F. Zhang, H. F. Wang, J. Qiu, J. Wu, H. J. Wang, J. W. Shao, Y. Q. Deng and L. F. Yan, *Chem. Eng. J.*, 2021, **421**, 127786.

