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Two-dimensional H– and F–BX (X = O, S, Se, and Te) photocatalysts with ultrawide bandgap and enhanced photocatalytic performance for water splitting†

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We theoretically propose a type of monolayer structure, H– or F–BX (X = As, Sb; Y = P, As), produced by surface hydrogenation or fluorination, with high stability, large band structures and high light absorption for photocatalytic water splitting. Based on first-principles calculations with the HSE06 functional, the electronic properties and optical properties were explored to reveal their potential performance in semiconductor devices. Additionally, owing to the Janus structure and high electronegativity of the monolayers, our calculations showed that surface fluorination can easily create an internal electric field compared with surface hydrogenation, which results in different trends of increasing bandgaps in monolayer H– and F–BX. We also found that the monolayers H– and F–BX have suitable band edges and high solar to hydrogen (STH) efficiency, enabling them to be photocatalysts for water splitting. Our work not only proposes eight monolayer semiconductors for expanding the number of two-dimensional semiconductors, but also provides a guide for how to regulate semiconductors for application in photocatalytic water splitting by using surface hydrogenation and fluorination.

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1. Introduction

Since the successful isolation of graphene,^{1–3} resulting in potential candidates for developing high performance electronic and optoelectronic nano-devices, the exploration of more novel 2D materials with various properties has attracted intense research interest. There are numerous 2D materials that have been reported, such as phosphorene,^{4–8} MXenes,^{9–11} transition metal dichalcogenides (TMDCs),^{12–19} and silicene.^{20–25} The use of these 2D materials in field-effect transistors, quantum devices, and energy storage, as well as conversion applications, has demonstrated excellent potential. In particular, because of their ultrathin atomic thickness and large specific surface area, these 2D materials are considered appealing candidates for use in flexible electronic devices^{26–28} and photocatalysts.^{29–35} However, the current 2D materials have two obvious drawbacks. First, although there are a few 2D materials with ultrawide bandgaps (>3.40 eV), the majority of 2D pristine materials have bandgaps between 0 and 3.40 eV,^{36–40} which limits their application in high-energy and high-frequency devices as well as insulators. Second, due to the rigorous constraints of band structures,

optical absorption, exciton recombination behaviors, and stability, there are only a few 2D materials, such as monolayer SiP₂ and B₆P₂, that perform well in photocatalytic water splitting.^{41–44} Due to the lack of 2D materials for high-energy and high-frequency devices, more effort should be made to investigate more 2D materials for use in ultrawide bandgap technologies and photocatalytic water splitting.

Recently, first-principles calculation has grown in importance. It can be separated into two parts. One part is the design of whole new structures, which is critical for enriching 2D materials,^{45,46} and the second part is the creation of new models and improvement of their performance by applying doping, heterojunction, or surface modifications.^{47–53} Due to the recent discovery of countless new 2D models, the first part is becoming increasingly challenging. Therefore, the second part is becoming more popular. Hydrogenation and fluorination are recognized as common but effective surface modification methods to tune the structures and properties of newly discovered materials. For example, the hydrogenation and fluorination of graphene can result in bandgaps of approximately 4.00 and 3.13 eV, respectively.^{53–56} In addition, applying hydrogenation to the top side of bilayer h–BN and fluorination to its bottom side simultaneously can form a new metal-free Janus photocatalyst that can drive infrared-light photocatalytic water splitting.⁵⁷ Furthermore, hydrogenation can also improve the stability of monoatomic-layer MX (M = Ga and In, X = S, Se, and Te) nanosheets.⁴⁶ In comparison to the bandgaps of typical

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monolayer MX, the bandgaps of hydrogenated MX nanosheets are direct and larger. As a result, we are driven to try to create new 2D structures with ultrawide bandgaps and appropriate efficiency for photocatalytic water splitting *via* surface hydrogenation and fluorination. If the outcome is favorable, it will encourage the development of 2D materials in ultrawide bandgap technology and photocatalytic water splitting.

Based on first-principles density functional theory, we have theoretically proposed the monolayers H- and F-BX (X = O, S, Se, and Te) by applying hydrogenation and fluorination to monoatomic-layer hexagonal BX. The binding energies and phonon spectra of these nanosheets were analyzed to identify their dynamics and energy stability. The calculations of their band structures and densities of states were performed to study their electronic structure properties. Furthermore, the edge levels, work functions, optical absorption spectra, and solar to hydrogen (STH) efficiencies of the H- and F-BX monolayers were studied to explore their performance in optical devices and photocatalysts for water splitting.

2. Computational details

Our entire analysis was based on first-principles density functional theory. The Vienna *Ab initio* Simulation Package (VASP)⁵⁸ was used to calculate structure optimizations, band structures, densities of states, and optical absorption spectra, utilizing projector-augmented-wave pseudopotentials.⁵⁹ The plane-wave cutoff energy was set to 500 eV. To relax our structure models, we used the generalized gradient approximation expressed by the Perdew–Burke–Ernzerhof (PBE) functional.⁶⁰ Once fully relaxed, the forces in the optimized monolayers H- and F-BX were less than 0.001 eV Å⁻¹ and the energy tolerances were less than 1 × 10⁻⁸ eV per atom. A 20 Å vacuum layer was adopted between adjacent layers. In the Monkhorst–Pack system,⁶¹ the Brillouin zone was represented by a 6 × 6 × 1 *k*-point sample.⁶¹ For studying the electrical structures and optical properties, the Heyd–Scuseria–Ernzerhof (HSE06) hybrid functional was used.⁶² The phonon spectra of the monolayers were computed using the Cambridge Series of Total Energy Package (CASTEP) and linear response approach.^{63,64} The norm-conserving pseudopotentials and a plane-wave cutoff energy of 940 eV were employed in the phonon spectra computations, while the other settings were the same as in VASP setting.

3. Results and discussion

3.1 Rationality and stability of monolayer H- and F-BX structures

From Fig. S2(a)–(d),† we can see that all the phonon spectra of monoatomic-layer BX (X = O, S, Se, and Te) clearly show a soft model in their Brillouin zones, thus we can confirm that monoatomic-layer hexagonal BX (X = O, S, Se, and Te) is unstable. In typical monolayer B₂X₂, two monoatomic-layer hexagonal BX structures stack directly, and the top sub-layer's B atoms connect with the bottom sub-layer's B atoms to saturate each other's unpaired electrons. As a result, we can use H and F atoms to saturate the unpaired electrons of the B atoms,

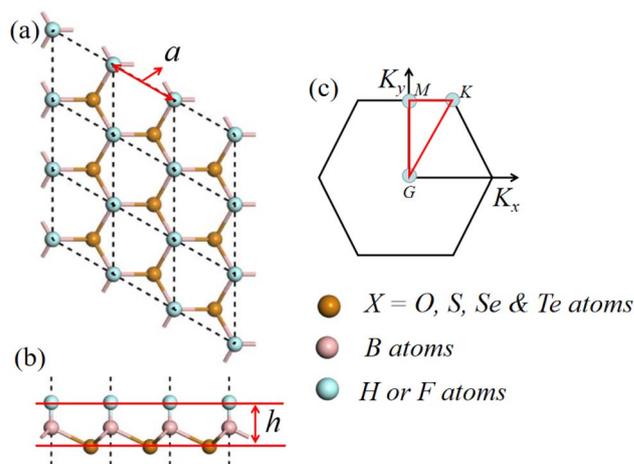


Fig. 1 (a) and (b) Structure model of the monolayers H- and F-BX (X = O, S, Se, and Te) and (c) Brillouin zone of monolayer H- and F-BX (X = O, S, Se, and Te) with high-symmetry points.

resulting in the formation of new monolayers H- and F-BX (X = O, S, Se, and Te), as shown in Fig. 1(a) and (b). The lattice parameters of the monolayers H- and F-BX after optimization of their entire structures are presented in Table 1. As can be seen in Fig. 1(a) and (b), the B₂X₂ monolayer clearly shows two different sublayers, an H or F sublayer and an X (X = O, S, Se, and Te) sublayer. This unique asymmetric geometry results in an out-of-plane built-in electric field of up to 0.96–1.31 Debye, as listed in Table 1. From looking at the dipoles of the H- and F-BX monolayers, we find that the dipoles of the F-BX monolayers are higher than those of the corresponding H-BX monolayers due to the strong electronegativity of the F atom. Meanwhile, the dipoles of the H- and F-BX monolayers increase with the decrease of the electronegativity of the X atom. Furthermore, the increase of the dipoles in the H-BX monolayers happens faster than that in the F-BX monolayers, which may further influence the change of the band structure. Although the monolayers H- and F-BX have suitable bonding character, further research into their energy and dynamics stability is required. To study the energy stability of the monolayers H- and F-BX, we calculated two types of binding energies (E_b). The first type E_{b1} is a typical one, which is calculated by the following formula:

$$E_{b1} = \frac{E_{\text{tot}} - E_{\text{atom}}(\text{B}) - E_{\text{atom}}(\text{HorF}) - E_{\text{atom}}(\text{X})}{3} \quad (1)$$

where E_{tot} is the total energy of the monolayers H- and F-BX and E_{atom} is the energy of the free atoms. The results are shown in Table 1. The E_{b1} values of the eight monolayers reveal that the monolayers H- and F-BX may have stable energies owing to their low binding energies. More evidence should be provided to support the claim that these eight monolayers have stable energies. Then, we calculated E_{b2} , which can provide the strength values of the H- or F-B bonds. The corresponding formula is as follows:

$$E_{b2} = E_{\text{tot}} - E(\text{BX}) - E_{\text{atom}}(\text{HorF}) \quad (2)$$



Table 1 The lattice constants and two types of binding energy in the monolayers H- and F-BX (X = O, S, Se, and Te)

	H-BO	H-BS	H-BSe	H-BTe	F-BO	F-BS	F-BSe	F-BTe
a (Å)	2.46	3.06	3.27	3.59	2.53	3.10	3.32	3.65
h (Å)	1.74	2.00	2.08	2.17	1.88	2.19	2.27	2.37
E_{b1} (eV per atom)	-6.32	-5.07	-4.78	-4.46	-6.97	-5.73	-5.45	-5.10
E_{b2} (eV)	-5.87	-5.43	-5.57	-5.32	-7.82	-7.42	-7.57	-7.23
p (Debye)	0.02	0.44	0.67	0.96	0.91	0.84	0.94	1.31

where $E(\text{BX})$ is the total energy of the monoatomic-layer hexagonal BX (X = O, S, Se, and Te) after full relaxation. The E_{b2} values of the monolayers H- and F-BX are shown in Table 1. We also calculated the E_{b2} values of the monolayer B_2X_2 (X = O, S, Se, and Te), which show the bond energies of the B-B bonds. The corresponding formula is as follows:

$$E_{b2} = E_{\text{tot}} - 2E(\text{BX}) \quad (3)$$

where E_{tot} is the total energy of the monolayer B_2X_2 . The E_{b2} values of the monolayers B_2O_2 , B_2S_2 , B_2Se_2 , and B_2Te_2 are -4.32, -3.65, -3.88, and -3.47 eV, respectively. It is clear that the E_{b2} values of the monolayer B_2X_2 are higher than those of the H- and F-BX monolayers. This means that H and F atoms bonding with the monoatomic-layer hexagonal BX are more stable than BX bonding with BX. Thus, we confirm that the monolayers H- and F-BX have stable energies.

Fig. 2 shows the phonon spectra of the monolayers H- and F-BX, which are expected to indicate the dynamics stability. The

phonon spectra without imaginary frequencies strongly indicate that the monolayers H- and F-BX have stable dynamics. Fig. 2 shows nine lattice wave branches in all the phonon spectra and a lack of imaginary frequencies in all the acoustic branches. Although we are unable to provide more evidence for monolayer thermodynamic stability and other types of stability owing to limited computing resources, the study of the binding energies and phonon spectra clearly indicates that the monolayers H- and F-BX are stable.

3.2 Electronic band structures of the monolayers H- and F-BX

Fig. 3 illustrates the band structures of the monolayers H- and F-BX as simulated by the HSE06 hybrid functional. Fig. 3(a) and (e) demonstrates that the bandgaps of the insulators H-BO and F-BO are 8.58 eV and 9.63 eV, respectively. The H-BS (4.36 eV), H-BSe (4.56 eV), F-BS (5.00 eV), and F-BSe (4.06 eV) are indirect, ultrawide bandgap semiconductors (>3.40), as shown in

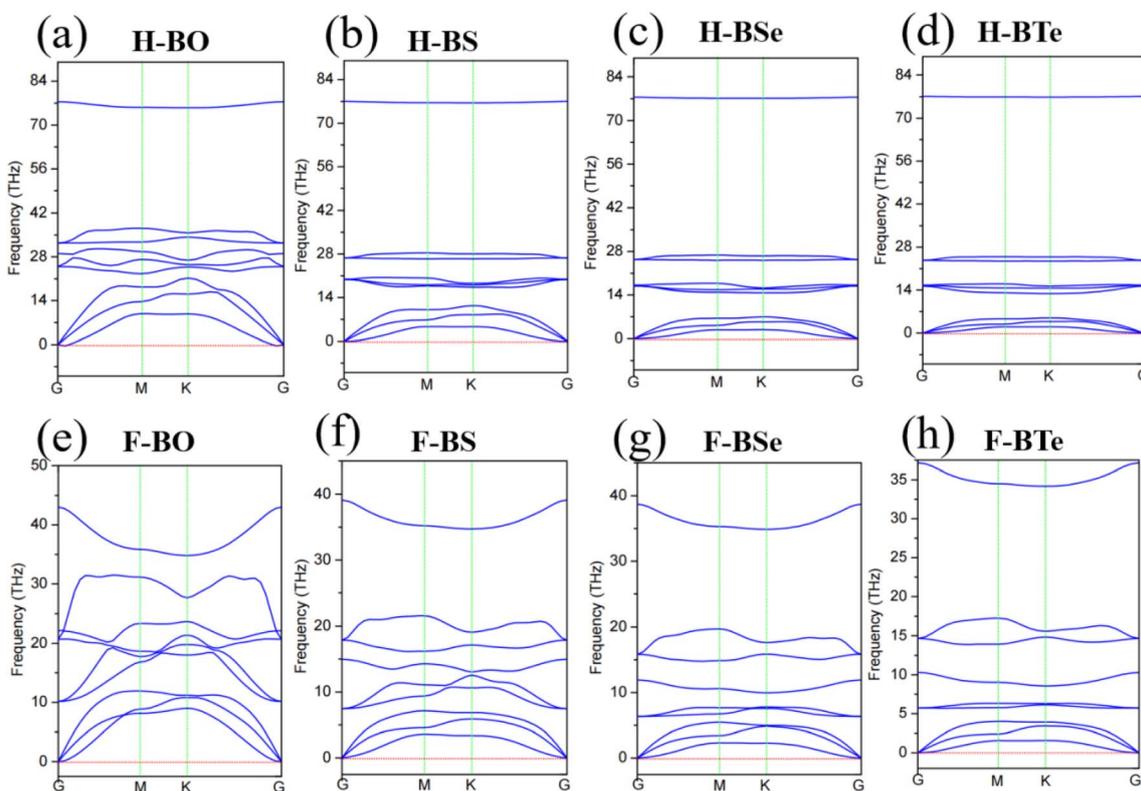


Fig. 2 Phonon spectra of monolayers H- and F-BX (X = O, S, Se, and Te).



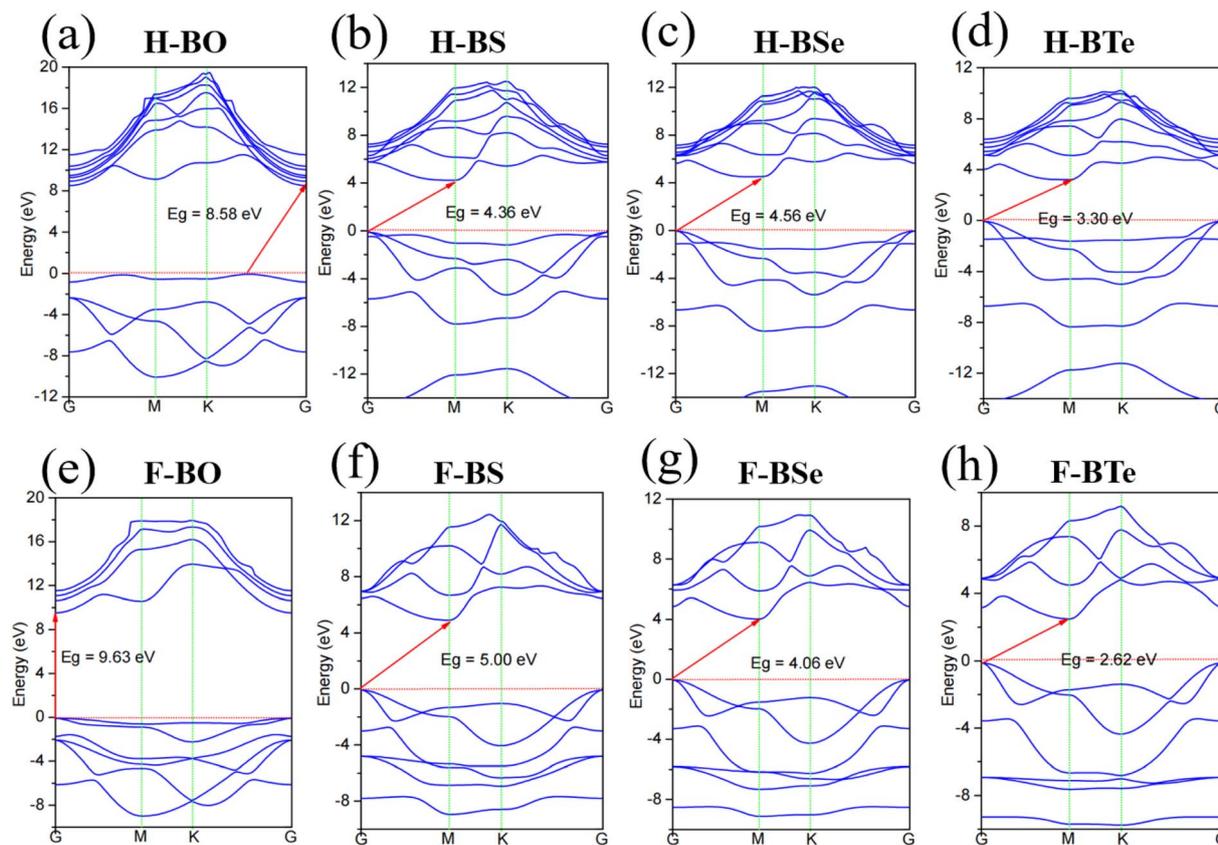


Fig. 3 Band structures of the monolayers H- and F-BX ($X = \text{O}, \text{S}, \text{Se}, \text{and Te}$).

Fig. 3(b), (c), (f) and (g). Additionally, the H-BTe and F-BTe (see Fig. 3(d) and (h)) are indirect, large bandgap semiconductors with values of 3.30 and 2.62 eV, respectively. From the bandgaps of F-BX and H-BX, we notice that bandgaps of the F-BX monolayers change faster than those of H-BX, resulting in the bandgap of H-BO being smaller than that of F-BO, but the bandgap of H-BTe being larger than that of F-BTe. These different trends in the bandgaps could be caused by the different trends of the increase of the built-in electric field. Meanwhile, we also found that the valence band maximum (VBM) of F-BO is located at the G point, whereas the VBM of H-BO occurs along the K - G direction. The conduction band minimum (CBM) values of both H-BO and F-BO lie at the G point. All the VBM values of the other monolayers H- and F-BX lie at the G point, while all their CBM values lie at the M point. For the above band structure analysis, the corresponding Brillouin zone of H-BO and F-BO with high-symmetry points can be seen in Fig. 1(c). We also calculated the band structure of the 1T-phase B_2X_2 as shown in Fig. S3† and the bandgaps of the B_2X_2 monolayers match well with those found in previous work.^{67,68} Compared with the corresponding monolayer B_2X_2 , the monolayers H- and F-BX have much larger bandgaps, implying that hydrogenation and fluorination of monoatomic-layer hexagonal BX can enlarge their bandgaps and broaden their applications in insulator and semiconductor technologies.

As shown in Fig. 4 and 5, we calculated the projected density of states (PDOS) of the monolayers H- and F-BX in order to

determine the contribution of each atom to the electronic states. For H-BO and H-BS (shown in Fig. 4(a) and (b), respectively), the occupied states near the VBM are filled by B, O (S), and H atoms, while their occupied states near the CBM are mainly filled by B and O (S) atoms. Fig. 4(c) and (d) demonstrates that B and Se (Te) atoms are significant contributors to the occupied states adjacent to the VBMs and CBMs, while H atoms are practically insignificant. Therefore, as shown in Fig. 4, the carriers in H-BX mainly transport on the BX planes, which may cause a hard separation of electron and hole pairs. For F-BX, the occupied states near the VBM and CBM, as shown in Fig. 5, are also mainly contributed to by B and X atoms. However, in contrast to H-BX, although it is not so strongly, F atoms in F-BX contribute to the occupied states near the VBM and CBM, which means that F atoms can provide carriers. Due to the noncoplanarity of the B, X, and F atoms, these phenomena allow F-BX to be more effective as a photocatalyst for separating electrons and holes.

3.3 Optical and photocatalytic properties of the monolayers H- and F-BX

Due to their large bandgaps, the monolayers H- and F-BX are likely to be suitable photocatalysts for water splitting. Thus, calculating the work function of these eight monolayers allowed us to determine the vacuum level alignment of the band edge levels, as shown in Fig. 6(a). When we calculated the work function, we applied the dipole correction owing to the Janus



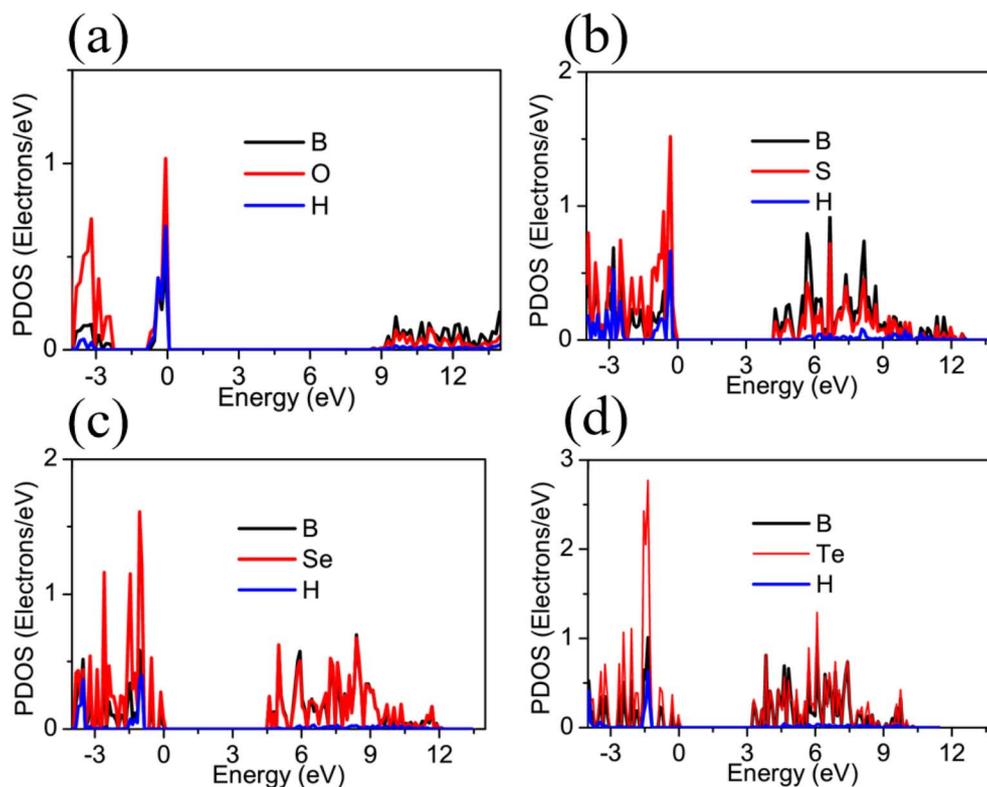


Fig. 4 PDOS of the monolayer H-BX (X = O, S, Se, and Te).

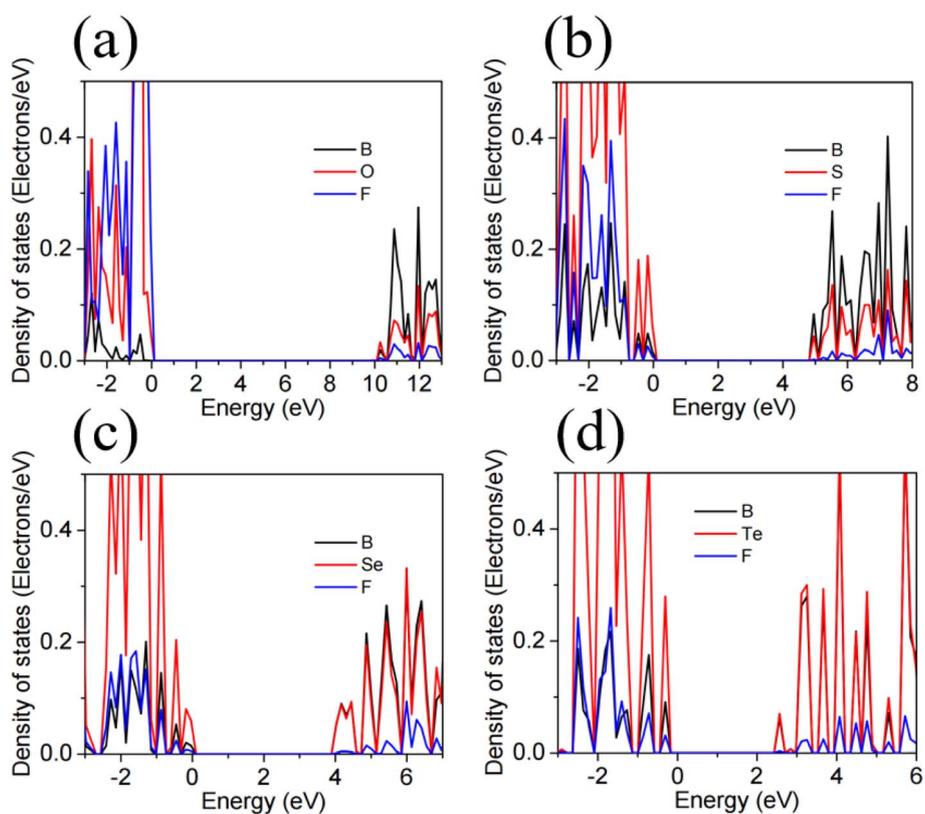


Fig. 5 PDOS of the monolayer F-BX (X = O, S, Se, and Te).



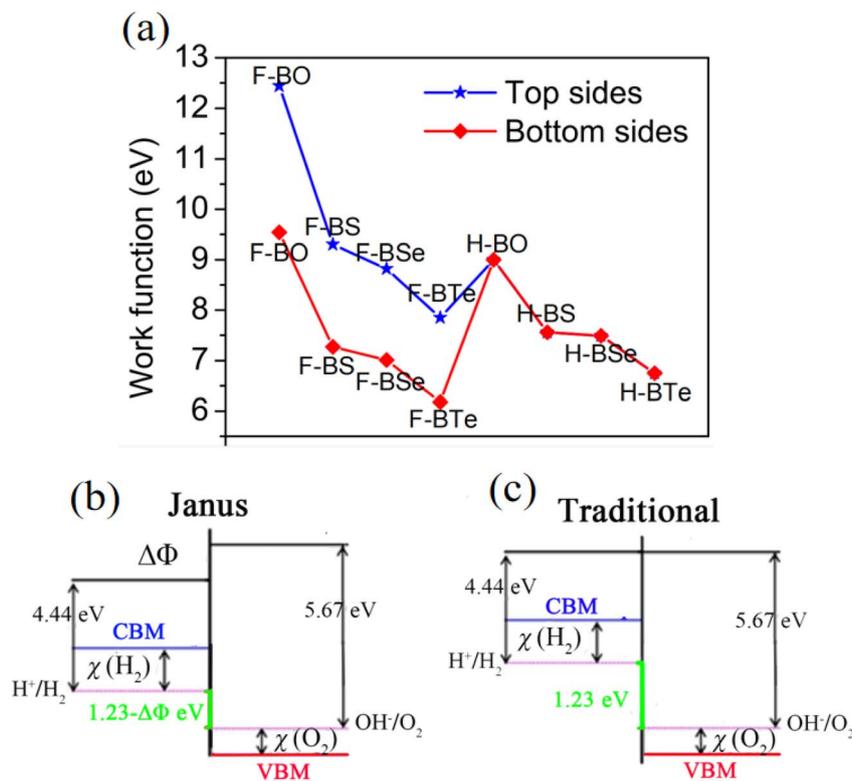


Fig. 6 (a) Work function of the monolayers H- and F-BX (X = O, S, Se, and Te) on the top sides and bottom sides; (b) energy level diagram for photocatalytic water splitting, which is suitable for the monolayer F-BX (X = O, S, Se, and Te); and (c) schematic diagram suitable for the monolayer H-BX (X = O, S, Se, and Te).

structures of the monolayers H- and F-BX. Fig. 6(a) indicates that all the monolayers H- and F-BX have a large work function (>6.0 eV). In combination with their large enough bandgaps as mentioned above, we propose that these monolayers may be used in photocatalytic water splitting. Furthermore, it is interesting to note that the work functions of the monolayer F-BX for the top (F atoms) and bottom (B and X atoms) sides are different. In contrast, the work functions of the monolayer H-BX on the top (H atoms) and bottom (B and X atoms) sides are the same. This phenomenon will cause the monolayers F-BX and H-BX to exhibit different photocatalytic mechanisms. The monolayer H-BX utilizes the conventional photocatalytic mechanism and is anticipated to be used in photocatalytic water splitting, as shown in Fig. 6(c). Fig. 6(c) demonstrates that in semiconductor photocatalysts, the CBM relative to the vacuum level must be lower than the oxidation potential [$V(\text{O}_2/\text{H}_2\text{O}) = -5.67 \text{ eV}$] and the VBM relative to the vacuum level must be greater than the reduction potential [$V(\text{H}_2/\text{H}^+) = -4.44 \text{ eV}$]. The Janus photocatalytic process used by the monolayer F-BX is depicted in Fig. 6(b), and it shows that the internal electric field and vacuum level difference will help separate electrons and holes on the top and bottom sides and increase the driving energy of the redox reaction in water splitting. Next, we will study the photocatalytic properties of the monolayers H- and F-BX for water splitting.

As can be seen in Fig. 7(a), comparing the CBM and VBM of the monolayers H- and F-BX with the redox potentials in water splitting, the band edge positions of the monolayers H- and F-

BX match with the band structure requirements for photocatalytic water splitting. Additionally, Fig. 7(a) demonstrates that every monolayer of H- and F-BX has a sufficiently large driving energy value (>0.4 eV) to support both the reduction of H₂O to O₂ and H⁺ to H₂. According to this discovery, the monolayers H- and F-BX might be thought of as potential 2D semiconductors for photocatalytic water splitting. However, even though these monolayers meet the band structure requirements for photocatalytic water splitting, it is still vital to consider the solar to hydrogen efficiency and optical absorption, which are crucial standards for the photocatalytic efficiency. The STH efficiency (η_{sth}) of the monolayer H-BX can be calculated by the following formula:⁶⁵

$$\eta_{\text{sth}} = \frac{\Delta G \int_E^{\infty} \frac{P(\hbar\omega)}{\hbar\omega} d(\hbar\omega)}{\int_0^{\infty} P(\hbar\omega) d(\hbar\omega)} \quad (4)$$

where $P(\hbar\omega)$ denotes the AM1.5G solar energy flux, which varies with photon energy $\hbar\omega$, and ΔG is set to 1.23 eV. The minimum energy E of the photons is given as follows:

$$E = \begin{cases} E_g & (\chi(\text{H}_2) \geq 0.2, \chi(\text{O}_2) \geq 0.6) \\ E_g + 0.2 - \chi(\text{H}_2) & (\chi(\text{H}_2) < 0.2, \chi(\text{O}_2) \geq 0.6) \\ E_g + 0.6 - \chi(\text{O}_2) & (\chi(\text{H}_2) \geq 0.2, \chi(\text{H}_2) < 0.6) \\ E_g + 0.8 - \chi(\text{H}_2) - \chi(\text{O}_2) & (\chi(\text{H}_2) < 0.2, \chi(\text{O}_2) < 0.6) \end{cases} \quad (5)$$



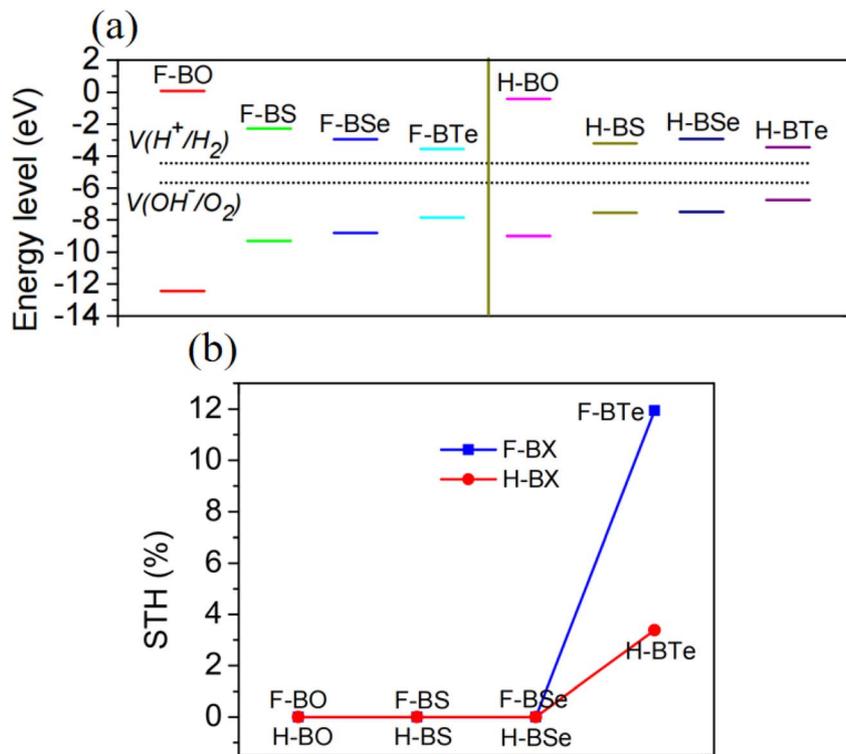


Fig. 7 (a) VBM and CBM of the monolayers H- and F-BX (X = O, S, Se, and Te) compared with the standard redox potentials for water splitting and (b) STH efficiency of the monolayers H- and F-BX (X = O, S, Se, and Te).

where E_g is the bandgap of the monolayer H-BX, $\chi(O_2)$ denotes the driving energy of reactions involving the evolution of oxygen, and $\chi(H_2)$ denotes the driving energy of reactions involving the evolution of hydrogen, which can be seen in Fig. 6(b) and (c). The STH efficiencies of H-BO, H-BS, and H-BSe tend to zero, whereas the H-BTe has a STH efficiency value of 3.38%, as shown in Fig. 7(b). The monolayer H-BX (X = O, S, Se, and Te) exhibits unsatisfactory performance because of large bandgaps. The STH efficiency of the monolayer F-BX needs to be adjusted because of its internal electric field, which influences electron-hole separation during photocatalytic water splitting in a positive way.⁶⁶

$$\eta'_{\text{sth}} = \eta_{\text{sth}} \times \frac{\int_0^{\infty} P(\hbar\omega) d(\hbar\omega)}{\int_0^{\infty} P(\hbar\omega) d(\hbar\omega) + \Delta\phi \int_{E_g}^{\infty} \frac{P(\hbar\omega)}{\hbar\omega} d(\hbar\omega)} \quad (6)$$

where $\Delta\phi$ (shown in Fig. 6(b)) is the electrostatic potential difference defined as $|V_{\text{bottom}} - V_{\text{top}}|$. Fig. 7(b) indicates that the STH efficiencies of F-BO, F-BS, and F-BSe also tend to zero owing to their large bandgaps. However, the STH efficiency of F-BTe reaches 11.94% (>10%), which means that it has certain economic value when applied in photocatalytic hydrogen production.

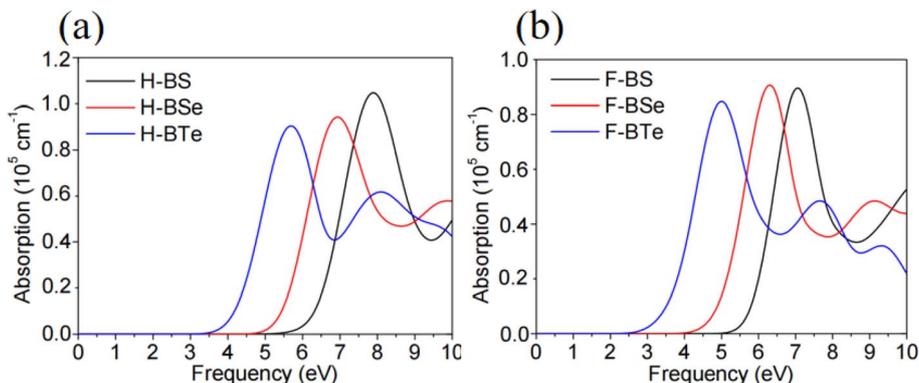


Fig. 8 (a) and (b) Optical absorption spectra of the monolayers H-BX and F-BX (X = S, Se, and Te), respectively.



Additionally, as can be seen in Fig. 8, we estimated the optical absorption spectra to investigate the optical absorption coefficients of the monolayers H- and F-BX. The derived frequency-dependent dielectric function can be written as $\varepsilon(\omega) = \varepsilon_1(\omega) + i\varepsilon_2(\omega)$, which can be used to determine the absorption coefficient using the following expression:

$$\alpha(E) = \frac{2e}{\hbar c} \left[\frac{(\varepsilon_1^2 + \varepsilon_2^2)^{1/2} - \varepsilon_1}{2} \right]^{1/2} \quad (7)$$

Studying the optical absorption of H- and F-BO makes little sense due to their insulating characteristics. As a result, we only examine the optical absorption spectra of the monolayers H- and F-BX (X = S, Se, and Te) in this research. The optical absorption spectra of the monolayers H-BX and F-BX (X = S, Se, and Te) are shown in Fig. 8(a) and (b), respectively. Fig. 8 shows that the optical absorption zones of the monolayers H-BX (X = S, Se, and Te) and F-BX (X = S and Se) exhibit considerable optical absorption in the high-frequency ultraviolet range. It follows that the monolayers H-BX (X = S, Se, and Te) and F-BX (X = S and Se) might play a significant role in the development of UV light-emitting diodes (LEDs), high-frequency, high-temperature, and high-power electronic devices, and UV photocatalysts. By combining its band edge state with the fact that F-BTe exhibits apparent optical absorption in the visible-light region, the monolayer can be employed as a photocatalyst in visible-light water splitting, as shown in Fig. 8(b).

4. Conclusion

In summary, using first-principles calculations, we have studied the monolayers H- and F-BX (X = O, S, Se, and Te). The conclusions are as follows: first, the calculation of the binding energies and phonon spectra of the monolayers H- and F-BX indicate that the monolayers H- and F-BX have both highly stable energies and dynamics.

Second, upon studying their electronic properties, we find that H-BO and F-BO are insulators, while H-BS, H-BSe, F-BS, and F-BSe are indirect, ultrawide bandgap semiconductors.

Additionally, H-BTe and F-BTe are indirect semiconductors with broad bandgaps that fulfill the demands of photocatalysts for water splitting. Our research also reveals that F-BTe is a potential candidate for visible-light photocatalytic water splitting because of its excellent photocatalytic qualities, which include high light harvesting capacity, internal electric field, and desirable STH efficiency.

Our work verifies the beneficial effects of surface hydrogenation and fluorination in adjusting the characteristics of 2D materials and also introduces eight new monolayers for use in ultrawide bandgap technology and photocatalytic water splitting.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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References

- 1 K. S. Novoselov, A. K. Geim, S. V. Morozov, D. Jiang, Y. Zhang, S. V. Dubonos, I. V. Grigorieva and A. A. Firsov, Electric Field Effect in Atomically Thin Carbon Films, *Science*, 2004, **306**, 666.
- 2 A. K. Geim and K. S. Novoselov, The Rise of Graphene, *Nat. Mater.*, 2007, **6**, 183.
- 3 A. H. Castro Neto, F. Guinea, N. M. R. Peres, K. S. Novoselov and A. K. Geim, The Electronic Properties of Graphene, *Rev. Mod. Phys.*, 2009, **81**, 109.
- 4 L. Li, Y. Yu, G. J. Ye, Q. Ge, X. Ou, H. Wu, D. Feng, X. H. Chen and Y. Zhang, Black Phosphorus Field-Effect Transistors, *Nat. Nanotechnol.*, 2014, **9**, 372.
- 5 M. Buscema, D. J. Groenendijk, S. I. Blanter, G. A. Steele, H. S. J. van der Zant and A. Castellanos-Gomez, Fast and Broadband Photoresponse of Few-Layer Black Phosphorus Field-Effect Transistors, *Nano Lett.*, 2014, **14**, 3347.
- 6 J. Qiao, X. Kong, Z.-X. Hu, F. Yang and W. Ji, High-Mobility Transport Anisotropy and Linear Dichroism in Few-Layer Black Phosphorus, *Nat. Commun.*, 2014, **5**, 4475.
- 7 V. Tran, R. Soklaski, Y. Liang and L. Yang, Layer-Controlled Bandgap and Anisotropic Excitons in Few-Layer Black Phosphorus, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2014, **89**, 235319.
- 8 H. Liu, A. T. Neal, Z. Zhu, Z. Luo, X. Xu, D. Tománek and P. D. Ye, Phosphorene: An Unexplored 2D Semiconductor with a High Hole Mobility, *ACS Nano*, 2014, **8**, 4033.
- 9 B. Anasori, M. R. Lukatskaya and Y. Gogotsi, 2D Metal Carbides and Nitrides (MXenes) for Energy Storage, *Nat. Rev. Mater.*, 2017, **2**, 16098.
- 10 M. Naguib, V. N. Mochalin, M. W. Barsoum and Y. Gogotsi, 25th Anniversary Article: MXenes: A New Family of Two-Dimensional Materials, *Adv. Mater.*, 2014, **26**, 992.
- 11 F. Shahzad, M. Alhabeab, C. B. Hatter, B. Anasori, S. Man Hong, C. M. Koo and Y. Gogotsi, Electromagnetic Interference Shielding with 2D Transition Metal Carbides (MXenes), *Science*, 2016, **353**, 1137.
- 12 S. Manzeli, D. Ovchinnikov, D. Pasquier, O. V. Yazyev and A. Kis, 2D Transition Metal Dichalcogenides, *Nat. Rev. Mater.*, 2017, **2**, 17033.
- 13 Q. H. Wang, K. Kalantar-Zadeh, A. Kis, J. N. Coleman and M. S. Strano, Electronics and Optoelectronics of Two-Dimensional Transition Metal Dichalcogenides, *Nat. Nanotechnol.*, 2012, **7**, 699.
- 14 M.-Y. Li, *et al.*, Epitaxial Growth of a Monolayer WSe₂-MoS₂ Lateral p-n Junction with an Atomically Sharp Interface, *Science*, 2015, **349**, 524.



- 15 S. Chuang, *et al.*, MoS₂ P-Type Transistors and Diodes Enabled by High Work Function MoO_x Contacts, *Nano Lett.*, 2014, **14**, 1337.
- 16 F. Xia, H. Wang and Y. Jia, Rediscovering Black Phosphorus as an Anisotropic Layered Material for Optoelectronics and Electronics, *Nat. Commun.*, 2014, **5**, 4458.
- 17 A. Kormányos, V. Zólyomi, N. D. Drummond and G. Burkard, Spin-Orbit Coupling, Quantum Dots, and Qubits in Monolayer Transition Metal Dichalcogenides, *Phys. Rev. X*, 2014, **4**, 011034.
- 18 L. Britnell, *et al.*, Strong Light-Matter Interactions in Heterostructures of Atomically Thin Films, *Science*, 2013, **340**, 1311.
- 19 Z. Hu, Z. Wu, C. Han, J. He, Z. Ni and W. Chen, Two-Dimensional Transition Metal Dichalcogenides: Interface and Defect Engineering, *Chem. Soc. Rev.*, 2018, **47**, 3100.
- 20 A. Fleurence, R. Friedlein, T. Ozaki, H. Kawai, Y. Wang and Y. Yamada-Takamura, Experimental Evidence for Epitaxial Silicene on Diboride Thin Films, *Phys. Rev. Lett.*, 2012, **108**, 245501.
- 21 L. Tao, E. Cinquanta, D. Chiappe, C. Grazianetti, M. Fanciulli, M. Dubey, A. Molle and D. Akinwande, Silicene Field-Effect Transistors Operating at Room Temperature, *Nat. Nanotechnol.*, 2015, **10**, 227.
- 22 P. Vogt, P. De Padova, C. Quaresima, J. Avila, E. Frantzeskakis, M. C. Asensio, A. Resta, B. Ealet and G. Le Lay, Silicene: Compelling Experimental Evidence for Graphenelike Two-Dimensional Silicon, *Phys. Rev. Lett.*, 2012, **108**, 155501.
- 23 Z. Ni, Q. Liu, K. Tang, J. Zheng, J. Zhou, R. Qin, Z. Gao, D. Yu and J. Lu, Tunable Bandgap in Silicene and Germanene, *Nano Lett.*, 2012, **12**, 113.
- 24 B. Radisavljevic and A. Kis, Mobility Engineering and a Metal-Insulator Transition in Monolayer MoS₂, *Nat. Mater.*, 2013, **12**, 815.
- 25 J. R. Schaibley, H. Yu, G. Clark, P. Rivera, J. S. Ross, K. L. Seyler, W. Yao and X. Xu, Valleytronics in 2D Materials, *Nat. Rev. Mater.*, 2016, **1**, 16055.
- 26 K. S. Novoselov, A. Mishchenko, A. Carvalho and A. H. Castro Neto, 2D Materials and van Der Waals Heterostructures, *Science*, 2016, **353**, aac9439.
- 27 G. Fiori, F. Bonaccorso, G. Iannaccone, T. Palacios, D. Neumaier, A. Seabaugh, S. K. Banerjee and L. Colombo, Electronics Based on Two-Dimensional Materials, *Nat. Nanotechnol.*, 2014, **9**, 768.
- 28 H. Tsai, *et al.*, High-Efficiency Two-Dimensional Ruddlesden-Popper Perovskite Solar Cells, *Nature*, 2016, **536**, 312.
- 29 M. Z. Rahman, C. W. Kwong, K. Davey and S. Z. Qiao, 2D Phosphorene as a Water Splitting Photocatalyst: Fundamentals to Applications, *Energy Environ. Sci.*, 2016, **9**, 709.
- 30 P. Xia, B. Zhu, B. Cheng, J. Yu and J. Xu, 2D/2D g-C₃N₄/MnO₂ Nanocomposite as a Direct Z-Scheme Photocatalyst for Enhanced Photocatalytic Activity, *ACS Sustainable Chem. Eng.*, 2018, **6**, 965.
- 31 X. Li, J. Yu, S. Wageh, A. A. Al-Ghamdi and J. Xie, Graphene in Photocatalysis: A Review, *Small*, 2016, **12**, 6640.
- 32 B. Luo, G. Liu and L. Wang, Recent Advances in 2D Materials for Photocatalysis, *Nanoscale*, 2016, **8**, 6904.
- 33 T. Su, Q. Shao, Z. Qin, Z. Guo and Z. Wu, Role of Interfaces in Two-Dimensional Photocatalyst for Water Splitting, *ACS Catal.*, 2018, **8**, 2253.
- 34 J. Fu, Q. Xu, J. Low, C. Jiang and J. Yu, Ultrathin 2D/2D WO₃/g-C₃N₄ Step-Scheme H₂-Production Photocatalyst, *Appl. Catal. B Environ.*, 2019, **243**, 556.
- 35 W.-J. Ong, L.-L. Tan, Y. H. Ng, S.-T. Yong and S.-P. Chai, Graphitic Carbon Nitride (g-C₃N₄)-Based Photocatalysts for Artificial Photosynthesis and Environmental Remediation: Are We a Step Closer To Achieving Sustainability?, *Chem. Rev.*, 2016, **116**, 7159.
- 36 L. Britnell, *et al.*, Electron Tunneling through Ultrathin Boron Nitride Crystalline Barriers, *Nano Lett.*, 2012, **12**, 1707.
- 37 K. Zhang, Y. Feng, F. Wang, Z. Yang and J. Wang, Two Dimensional Hexagonal Boron Nitride (2D-hBN): Synthesis, Properties and Applications, *J. Mater. Chem. C*, 2017, **5**, 11992.
- 38 X.-J. Ye, Z.-X. Zhu, L. Meng and C.-S. Liu, Two-Dimensional CaFCl: Ultra-Wide Bandgap, Strong Interlayer Quantum Confinement, and n-Type Doping, *Phys. Chem. Chem. Phys.*, 2020, **22**, 17213.
- 39 S. Zhang, Z. Yan, Y. Li, Z. Chen and H. Zeng, Atomically Thin Arsenene and Antimonene: Semimetal-Semiconductor and Indirect-Direct bandgap Transitions, *Angew. Chem., Int. Ed.*, 2015, **54**, 3112.
- 40 P. Johari and V. B. Shenoy, Tuning the Electronic Properties of Semiconducting Transition Metal Dichalcogenides by Applying Mechanical Strains, *ACS Nano*, 2012, **6**, 5449.
- 41 T. Yu, C. Wang, X. Yan, G. Yang and U. Schwingenschlög, Anisotropic Janus SiP₂ Monolayer as a Photocatalyst for Water Splitting, *J. Phys. Chem. Lett.*, 2021, **12**, 2464.
- 42 M. Sun and U. Schwingenschlög, B2P6: a two-dimensional anisotropic Janus material with potential in photocatalytic water splitting and metal-ion batteries, *Chem. Mater.*, 2020, **32**(11), 4795–4800.
- 43 C.-F. Fu, J. Sun, Q. Luo, X. Li, W. Hu and J. Yang, Intrinsic Electric Fields in Two-Dimensional Materials Boost the Solar-to-Hydrogen Efficiency for Photocatalytic Water Splitting, *Nano Lett.*, 2018, **18**, 6312.
- 44 J. Lin, T. Zhang, B. Zhang and X. Wang, Wide-bandgaps Two Dimensional C₃XN (X = N and P) for Metal-Free Photocatalytic Water Splitting, *Appl. Surf. Sci.*, 2021, **542**, 148597.
- 45 Y. Li, Y.-L. Li, B. Sa and R. Ahuja, Review of Two-Dimensional Materials for Photocatalytic Water Splitting from a Theoretical Perspective, *Catal. Sci. Technol.*, 2017, **7**, 545.
- 46 J. Lin, B. Zhang and T. Zheng, Half-Hydrogenated Monolayer Group III Monochalcogenides: Wide Direct Bandgaps and Potential Photocatalysts for Water Splitting, *J. Phys. Chem. C*, 2019, **123**, 27697.
- 47 L. Zhang, J. Yu, M. Yang, Q. Xie, H. Peng and Z. Liu, Janus Graphene from Asymmetric Two-Dimensional Chemistry, *Nat. Commun.*, 2013, **4**, 1443.



- 48 A.-Y. Lu, *et al.*, Janus Monolayers of Transition Metal Dichalcogenides, *Nat. Nanotechnol.*, 2017, **12**, 744.
- 49 J. Zhang, S. Jia, I. Kholmanov, L. Dong, D. Er, W. Chen, H. Guo, Z. Jin, V. B. Shenoy, L. Shi and J. Lou, Janus Monolayer Transition-Metal Dichalcogenides, *ACS Nano*, 2017, **11**(8), 8192–8198.
- 50 Y. Bai, Q. Zhang, N. Xu, K. Deng and E. Kan, The Janus Structures of Group-III Chalcogenide Monolayers as Promising Photocatalysts for Water Splitting, *Appl. Surf. Sci.*, 2019, **478**, 522.
- 51 C. Zhu, L. Zhang, B. Jiang, J. Zheng, P. Hu, S. Li, M. Wu and W. Wu, Fabrication of Z-scheme Ag₃PO₄/MoS₂ composites with enhanced photocatalytic activity and stability for organic pollutant degradation, *Appl. Surf. Sci.*, 2016, **377**, 99–108.
- 52 S. Wang, X. Yang, X. Zhang, X. Ding, Z. Yang, K. Dai and H. Chen, A plate-on-plate sandwiched Z-scheme heterojunction photocatalyst: biOBr–Bi₂MoO₆ with enhanced photocatalytic performance, *Appl. Surf. Sci.*, 2017, **391**, 194–201.
- 53 S. Kwon, J.-H. Ko, K.-J. Jeon, Y.-H. Kim and J. Y. Park, Enhanced Nanoscale Friction on Fluorinated Graphene, *Nano Lett.*, 2012, **12**, 6043.
- 54 F. Withers, M. Dubois and A. K. Savchenko, Electron Properties of Fluorinated Single-Layer Graphene Transistors, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2010, **82**, 073403.
- 55 M. Batzill, The Surface Science of Graphene: Metal Interfaces, CVD Synthesis, Nanoribbons, Chemical Modifications, and Defects, *Surf. Sci. Rep.*, 2012, **67**, 83.
- 56 S.-H. Cheng, K. Zou, F. Okino, H. R. Gutierrez, A. Gupta, N. Shen, P. C. Eklund, J. O. Sofo and J. Zhu, Reversible Fluorination of Graphene: Evidence of a Two-Dimensional Wide Bandgap Semiconductor, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2010, **81**, 205435.
- 57 Y. Xu, S. Qi and W. Mi, Electronic Structure and Magnetic Properties of Two-Dimensional h-BN/Janus 2H-VSeX (X = S, Te) van Der Waals Heterostructures, *Appl. Surf. Sci.*, 2021, **537**, 147898.
- 58 G. Kresse and J. Furthmüller, Efficient iterative schemes for ab initio total-energy calculations using a plane-wave basis set, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1996, **54**, 11169–11186.
- 59 G. Kresse and D. Joubert, From Ultrasoft Pseudopotentials to the Projector Augmented-Wave Method, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1999, **59**, 1758.
- 60 J. P. Perdew, K. Burke and M. Ernzerhof, Generalized Gradient Approximation Made Simple, *Phys. Rev. Lett.*, 1996, **77**, 3865–3868.
- 61 H. J. Monkhorst and J. D. Pack, Special Points for Brillouin-Zone Integrations, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1977, **16**(4), 1748–1749.
- 62 J. Heyd, G. E. Scuseria and M. Ernzerhof, Hybrid Functionals Based on a Screened Coulomb Potential, *J. Chem. Phys.*, 2003, **118**, 8207–8215.
- 63 M. D. Segall, P. J. D. Lindan, M. J. Probert, C. J. Pickard, P. J. Hasnip, S. J. Clark and M. C. Payne, First-principles simulation: ideas, illustrations and the CASTEP code, *J. Phys.: Condens. Matter*, 2002, **14**, 2717–2744.
- 64 S. Baroni, S. de Gironcoli, A. dal Corso and P. Giannozzi, Phonons and related crystal properties of extended systems from density-functional perturbation theory, *Rev. Mod. Phys.*, 2001, **73**, 515–562.
- 65 X. Li, Z. Li and J. Yang, Proposed Photosynthesis Method for Producing Hydrogen from Dissociated Water Molecules Using Incident Near-Infrared Light, *Phys. Rev. Lett.*, 2014, **112**, 018301.
- 66 C. Fu, J. Sun, Q. Luo, X. Li, W. Hu and J. Yang, Intrinsic electric fields in twodimensional materials boost the solar-to-hydrogen efficiency for photocatalytic water splitting, *Nano Lett.*, 2018, **18**(10), 6312–6317.
- 67 B. Mortazavi and T. Rabczuk, Boron Monochalcogenides; Stable and Strong Two-Dimensional Wide Band-Gap Semiconductors, *Energies*, 2018, **11**, 1573.
- 68 P. Mishra, D. Singh, Y. Sonvane and R. Ahuja, Two-dimensional boron monochalcogenide monolayer for thermoelectric material, *Sustain. Energy Fuels*, 2020, **4**, 2363–2369.

