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Preparation of highly dispersed SnO/TiO₂ catalysts and their performances in catalyzing polyol ester

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A series of stannous oxide supported on rutile titanium dioxide (SnO/TiO₂) were prepared by a conventional incipient wetness impregnation method, and their performance as catalysts for fatty acid esterification reactions was investigated. The effects of Sn precursors (SnCl₂·2H₂O or SnC₂O₄), loading amounts (5–15%), and treating ambiances (air and N₂) were explored. The optimized 10% SnO/TiO₂-Cl with SnCl₂·2H₂O as the Sn precursor and thermal treatment in N₂ showed the best esterification performance. Specifically, 10% SnO/TiO₂-Cl catalyzed the esterification process of trimethylolpropane and *n*-octanoic acid with a conversion of 99.6% over 5 h at 160 °C, and 10% SnO/TiO₂-Cl was efficient for six catalytic cycles. Based on the results of X-ray diffraction (XRD), Raman spectra, high-resolution transmission electron microscopy (HRTEM), infrared spectra of pyridine adsorption (Py-IR), and ammonia temperature programmed desorption (NH₃-TPD), the improved catalytic performance is supposed to be attributable to the high dispersion of the Sn species on 10% SnO/TiO₂-Cl as the moderate Lewis acid sites.

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1. Introduction

Polyol esters (POEs) are a class of green and renewable synthetic lubricants with high thermal stability, biodegradability, and viscosity index, which are mainly used for engine cooling and anti-wear, fire-resistant transformer cooling, industrial furnace chain lubrication, *etc.*^{1–3} Furthermore, compared with mineral-based lubricants obtained by adding additives to form a thick oil film between frictional substrates, POEs with polar functional groups easily adsorb to form a stable oil film, and thus have significantly better lubricating properties than mineral base oils.⁴ The esterification between polyols and carboxylic acids to POEs is generally catalyzed by conventional homogeneous acid catalysts, such as concentrated sulfuric acid, phosphoric acid, *p*-toluenesulfonic acid, *etc.*⁵ Although the conventional homogeneous acid catalysts have high catalytic efficiency and lower cost, they also suffer from serious equipment corrosion, more side reactions, and complicated subsequent separation processes.^{6–8} In contrast, heterogeneous acid catalysts are more economical and environmentally friendly than traditional homogeneous acid catalysts, and are becoming one of the research hotspots for catalytic reactions of POEs.

Recently, heterogeneous catalysts for fatty acid esterification reactions, such as immobilized mineral acids (sulfated ZrO₂, WO₃, SiO₂, SnO),^{9–11} heteropolyacids (H₃PW₁₂O₄₀/ZrO₂, H₃PW₁₂O₄₀),^{12–14} and oxides (SnO, WO₃/AlPO₄, WO₃/ZrO₂, SnO₂/SiO₂, zeolites),^{15–20} have been reported. Notably, stannous (Sn²⁺)-based catalysts such as stannous chloride (SnCl₂),²¹ stannous octanoate,²² and SnO¹⁶ have been increasingly considered as potential catalysts for esterification and ester exchange as a result of their high safety and low cost.^{23–25} However, as heterogeneous acidic catalysts, the main drawbacks of SnO are the low surface area and the inevitable presence of Sn residues in the product, which greatly restricts their practical application.²⁶ Despite substantial progress in supported Sn catalysts,^{27,28} more research is still necessary to improve the activity and stability of Sn-based catalysts for the synthesis of POEs.

Herein, a series of SnO supported on rutile titanium dioxide (SnO/TiO₂) catalysts were prepared by the conventional incipient wetness impregnation method to investigate the structure–activity relationship between Sn species and the esterification reaction. By adjusting the Sn precursors (SnCl₂·2H₂O or SnC₂O₄), loading amounts (5–15%), and treating ambiances (air and N₂), we found that 10% SnO/TiO₂-Cl with SnCl₂·2H₂O as precursors upon roasting in N₂ showed the excellent esterification properties. 10% SnO/TiO₂-Cl catalyzed the esterification of trimethylolpropane and *n*-octanoic acid with 99.6% conversion for 5 h at 160 °C, and 10% SnO/TiO₂-Cl was efficient after six catalytic cycles. The results indicated that the increased catalytic activity is due to the highly dispersed Sn species as the moderate Lewis acid sites on 10% SnO/TiO₂-Cl.

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Table 1 Summary of synthesis parameters and abbreviations of catalysts

Catalyst	Tin precursor	Heat treatment atmosphere
5% SnO/TiO ₂ -Cl	SnCl ₂ ·2H ₂ O (44 mg)	N ₂
7.5% SnO/TiO ₂ -Cl	SnCl ₂ ·2H ₂ O (68 mg)	N ₂
10% SnO/TiO ₂ -Cl	SnCl ₂ ·2H ₂ O (93 mg)	N ₂
15% SnO/TiO ₂ -Cl	SnCl ₂ ·2H ₂ O (148 mg)	N ₂
2.5% SnO/TiO ₂ -ox	SnC ₂ O ₄ (20 mg)	N ₂
10% SnO/TiO ₂ -ox	SnC ₂ O ₄ (85 mg)	N ₂
10% SnO/TiO ₂ -Cl-A	SnCl ₂ ·2H ₂ O (93 mg)	Air

2. Experimental

2.1. Preparation of SnO/TiO₂ catalysts

A series of SnO/TiO₂ catalytic systems were prepared *via* an incipient wetness impregnation method. In a typical preparation of 10% SnO/TiO₂-Cl, a certain amount of SnCl₂·2H₂O (93 mg, Sinopharm) was dissolved in anhydrous ethanol (2.5 mL, Sinopharm). TiO₂ (0.5 g, rutile, Aladdin) was added into the above solution, and continuously stirred for 1 h. Subsequently, the above mixture was dried overnight in vacuum drying chamber at 30 °C. Finally, the dried mixture was heated to 450 °C at a rate of 10 °C min⁻¹ in N₂ (100 mL min⁻¹) and maintained for 2 h, to obtain 10% SnO/TiO₂-Cl. The preparation process of SnO/TiO₂-ox was similar as that of SnO/TiO₂-Cl except changing SnCl₂·2H₂O to stannous (II) oxalate (SnC₂O₄, Macklin). By turning calcination atmosphere (N₂) of SnO/TiO₂-Cl to air, SnO/TiO₂-Cl-A was obtained. The synthesis parameters and abbreviations of catalysts in this work are summarized in Table 1.

2.2. Catalyst characterization

Crystal structures of catalysts were obtained using X-ray diffraction (XRD, Bruker D8) with an automated transmission diffractometer, Cu Kα1 radiation (40 kV/40 mA), the scanning speed of 8° min⁻¹, and the scanning range of 5–80°. Transmission electron microscopy (TEM) and high-resolution TEM (HRTEM) were performed on a Tecnai G2 F30 S-Twin at an acceleration voltage of 200 kV. Raman spectra were used by HORIBA HR evolution Raman spectrometer with laser wavelength of 532 nm and the scanning range of 3500–55 cm⁻¹. X-ray photoelectron spectroscopy (XPS) were carried out on Thermo ESCALAB 250Xi with the excitation source of Al Kα ($h\nu$ = 1486.6 eV). To calibrate the sample's binding energy, the binding energy of the indeterminate carbon (C 1s) was set to 284.6 eV.

Infrared spectra of pyridine adsorption (Py-IR) were obtained by Bruker Tensor27 Fourier transform infrared spectrometer. The procedure is as follows: the sample was pretreated for 2 h at 450 °C and vacuum (10⁻² Pa), subsequently cooled to 40 °C, and background was collected. Then, pyridine was introduced into *in situ* IR cell, and the physically adsorbed pyridine was removed for 1 h with argon. Finally, a Py-IR spectrum was collected with resolution of 4 cm⁻¹, acquisition times of 64, and a scanning range of 1700–1400 cm⁻¹.

Ammonia temperature programmed desorption (NH₃-TPD) was performed on the TP-5080-B temperature-programmed chemisorption instrument. The procedure is as follows: the sample (100 mg) was pretreated for 1 h at 450 °C with helium (99.999%, 50 mL min⁻¹), subsequently cooled to 50 °C, and pure NH₃ (99.99%, 10 mL min⁻¹) was introduced for 0.5 h. Then, physically adsorbed NH₃ was removed for 2 h with helium. Finally, the sample was heated from 50 °C to 750 °C at a rate of 10 °C min⁻¹, and continuously record NH₃ in the effluent through a thermal conductivity detector.

2.3. Activity test

In a typical synthetic process of lubricating ester oils, catalyst (0.5 wt%, 35 mg), trimethylolpropane (TMP, 15 mmol, 2.0 g), and *n*-octanoic acid (45 mmol, 5.1 g) were added into a three-necked flask (100 mL). Then, 3 mL toluene was added as a water-removing agent and the reaction system was heated to 160 °C under stirring and maintained for the desired time. After the reaction, the catalyst was removed by filtration. The produce was obtained by rotational evaporation to remove water and unreacted substrates. The conversion rate of the reaction product was determined by testing the acid value.

3. Results and discussion

3.1. Catalytic performance

Esterification reactions between TMP and *n*-octanoic acid were carried out (Fig. 1a). Fig. 1b summarizes the effect of the preparation conditions (Sn precursor, loading amount, and treatment performed atmosphere) on the catalytic activity of these SnO/TiO₂. As a comparison, the conversions in the absence of catalyst were 72.0%, 82.0% and 88.7% at 2 h, 5 h and 10 h, respectively, while these conversions were 74.7%, 83.0% and 90.3% with the addition of TiO₂. This indicates that in the absence of catalyst or addition of TiO₂, the esterification takes a long time to reach complete conversion, which severely reduces the product efficiency. The addition of SnO/TiO₂ catalysts significantly improved the esterification process of TMP and *n*-octanoic acid. Among these SnO/TiO₂, 10% SnO/TiO₂-Cl showed the best esterification performance with conversions of 91.2% and 99.6% at 2 h and 5 h, which was equivalent to or better than the catalytic esterification performance reported in the literature (Table 2).^{29–36} By changing the Sn precursor from SnCl₂·2H₂O to SnC₂O₄, the conversion was reduced to 86.2% and 92.3% at 2 h and 5 h on 10% SnO/TiO₂-ox. Then, the effect of Sn loading on the catalytic activity was investigated. Catalytic activity showed a volcanic trend with increasing Sn loading. Moreover, the catalytic performance of 10% SnO/TiO₂-ox was slightly better than that of pure SnO (89.7% and 98.2% at 2 h and 5 h, 3.5 mg), using the reported preparation method.³² Finally, when SnO/TiO₂-Cl-A was used as the catalyst, the conversions decreased to 74.7% and 85.5% at 2 h and 5 h, respectively. Considering that SnO can be converted to SnO₂ by roasting in the air, the decrease in the activity of SnO/TiO₂-Cl-A should be due to the oxidation of SnO.



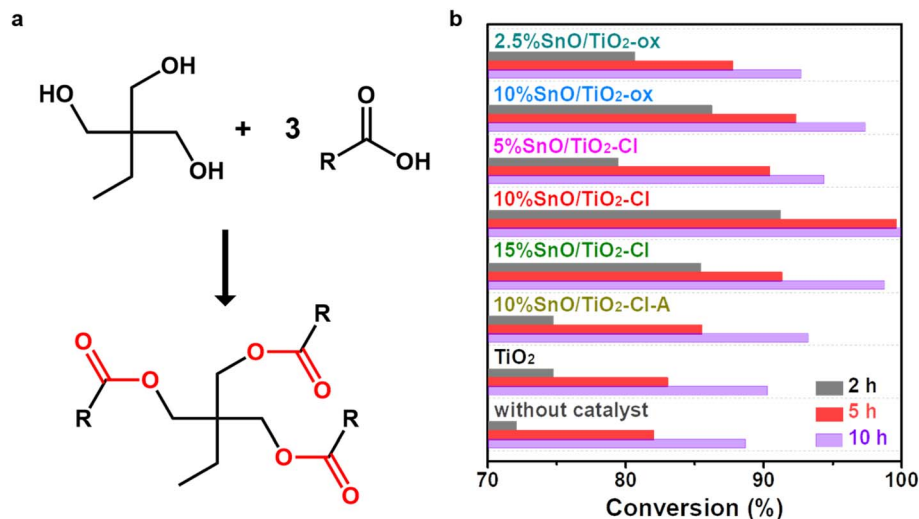


Fig. 1 (a) Schematic diagram of lubricating ester synthesis reaction. (b) Catalytic activities of different supported SnO/TiO₂ for lubricating ester synthesis reaction between TMP and *n*-octanoic acid.

Table 2 Comparison of catalytic performance for esterification reaction

Catalyst	Amount ^a (wt%)	Reactant	Tem. (°C)	Time (h)	Yield (%)	
10% SnO/TiO ₂ -Cl	0.5	TMP + <i>n</i> -octanoic acid	160	5	99.6	This work
<i>p</i> -toluene sulfonic acid	1.5	Fatty acid + TMP	130	5	89	29
SnCl ₂	0.1	TMP + undecylenic acid	150	8	96.2	30
SnO	0.9	Pentaerythritol + ricinoleic acid	170	4	99.7	31
SnO@GO	0.5	TMP + <i>n</i> -capric acid	160	5	98.3	32
COK-15b	2.2	Pentaerythritol + <i>n</i> -octanoic acid	160	6	97	33
Zr-ZSM-5-st	0.5	TMP + <i>n</i> -caprylic acid	160	5	94.41	34
ZrO ₂ -imidazolium sulfonic acid chloride	0.84	CH ₃ OH + oleic acid	60	1	98.9	35
Thermomyces lanuginosus lipase/Fe ₃ O ₄ -CA	5	Free fatty acids + neopentyl glycol	45	6	68	36

^a Amounts: mass percentage of the total amount of catalyst with substrate.

Following six cycles of TMP-*n*-caprylic esterification (Fig. 2), the obtained catalyst (10% SnO/TiO₂-Cl and 10% SnO/TiO₂-ox) exhibited reduced activity. The yield of TMP ester reflected

catalyst stability, and the value over 10% SnO/TiO₂-Cl was approximately 95%, higher than 10% SnO/TiO₂-ox (92.1%).

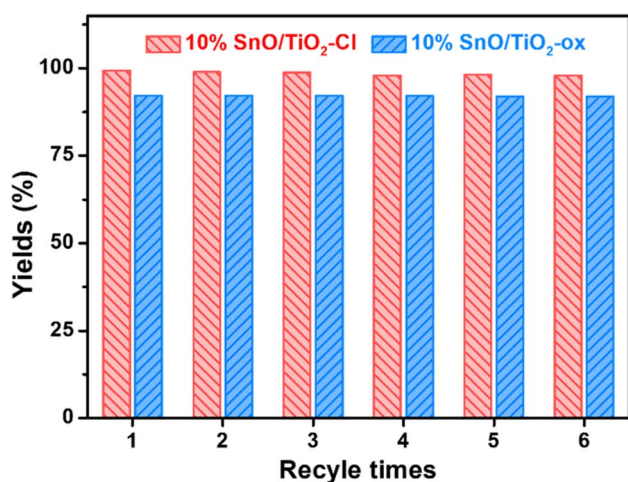


Fig. 2 The reusability of 10% SnO/TiO₂-Cl and 10% SnO/TiO₂-ox for esterification.

3.2. Characterization

The crystal phases of these SnO/TiO₂ were characterized by XRD (Fig. 3). In addition to these XRD pattern peaks matching rutile TiO₂ (PDF No. 21-1276), no new diffraction peaks corresponding to Sn, SnO or SnO₂ appear in the XRD pattern of 10% SnO/TiO₂-Cl, indicating that Sn species are highly dispersed on 10% SnO/TiO₂-Cl.³⁷ In contrast, the 10% SnO/TiO₂-ox shows three new diffraction peaks at 29.8°, 33.2° and 57.3°, which correspond to (101), (110) and (211) diffraction planes of the romanite-type SnO (PDF No. 06-0395) with the tetragonal crystal structure.³⁸ In addition, SnO species on 2.5% SnO/TiO₂-ox and 10% SnO/TiO₂-ox have average particle sizes of 33.9 and 76.7 nm, calculated by Scherrer's formula ($D = K\lambda/B\cos\theta$, K is Scherrer's constant (0.89); B is the half-height width of the measured sample diffraction peak; θ is the diffraction angle; γ is X-ray wavelength, 0.154056 nm). Considering that the catalytic activity of 10% SnO/TiO₂-ox was lower than that of 10% SnO/TiO₂-Cl (Fig. 1b), the highly dispersed Sn species should be



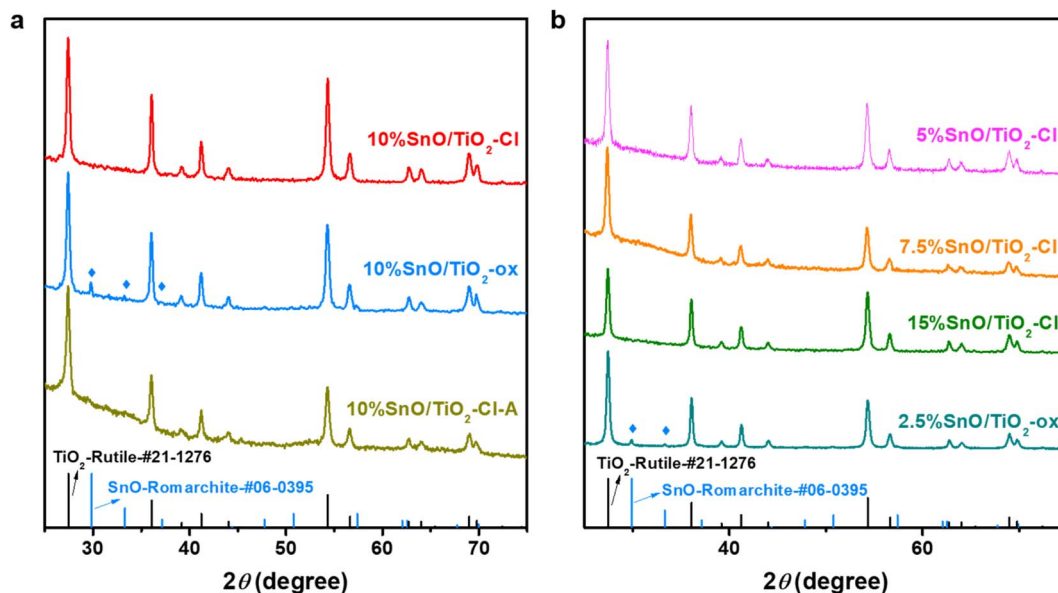


Fig. 3 XRD pattern of TiO_2 and different SnO/TiO_2 .

responsible for the improved catalytic performance. Furthermore, XRD results of $\text{SnO}/\text{TiO}_2\text{-Cl}$ with different SnO loadings indicated the highly dispersed SnO on rutile TiO_2 .

Fig. 4 illustrates Raman spectra of these SnO/TiO_2 catalysts. For Raman spectrum of rutile TiO_2 , three strong peaks at 250, 425, and 605 cm^{-1} are assigned to the second order effect (SOE) mode, E_g mode (asymmetric bending of the O-Ti-O bond in the {001} plane), and A_{1g} (symmetric stretching of O-Ti-O) mode of rutile TiO_2 , respectively.³⁹ In contrast, the peak of SnO (210 and 116 cm^{-1}) does not appear in the Raman spectrum of 10% $\text{SnO}/$

$\text{TiO}_2\text{-Cl}$, where the peak for E_g mode of vibration shifts to a higher Raman shift (431 cm^{-1}). This indicates that the structure of TiO_2 is slightly distorted, due to the insertion of Sn

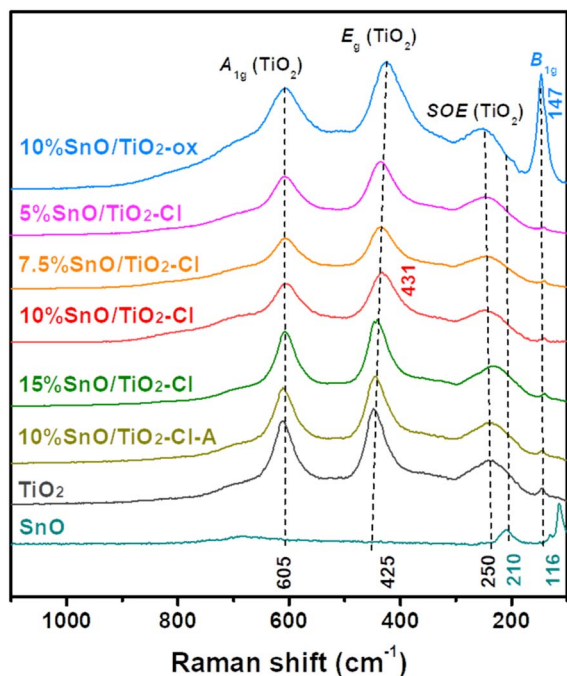


Fig. 4 Raman spectra of TiO_2 and different SnO/TiO_2 .

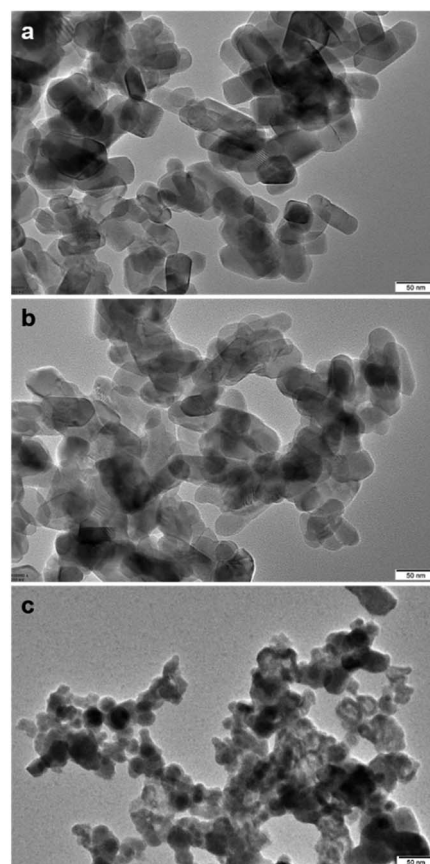


Fig. 5 TEM images of (a) 10% $\text{SnO}/\text{TiO}_2\text{-Cl}$, (b) 10% $\text{SnO}/\text{TiO}_2\text{-ox}$, and (c) SnO .



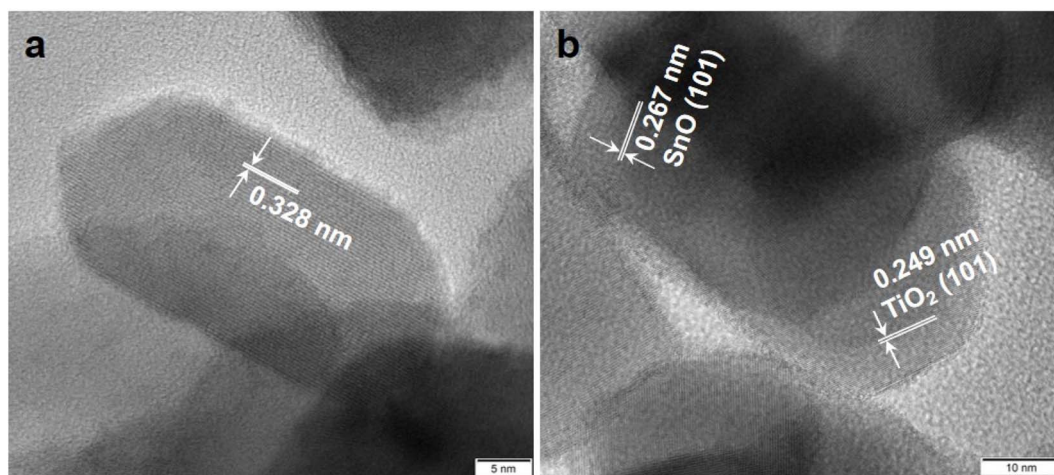


Fig. 6 HRTEM images of (a) 10% SnO/TiO₂-Cl and (b) 10% SnO/TiO₂-ox.

ions. To verify this observation, we examined the Raman spectra of SnO/TiO₂-Cl with different Sn contents, where the E_g peak shifts to 441 cm⁻¹ (15% SnO/TiO₂-Cl) from 436 cm⁻¹ (5% SnO/TiO₂-Cl). For Raman spectrum of 10% SnO/TiO₂-ox, one new peak at 147 cm⁻¹ (Fig. 4), assigned to a stoichiometric SnO_x phase ($1 < x < 2$).⁴⁰

As shown in Fig. 5, TEM images of 10% SnO/TiO₂-Cl and 10% SnO/TiO₂-ox show uniform polyhedral in the range of 50–150 nm. SnO shows uniform spheroidal in the range of 20–50 nm (Fig. 5c). After the magnification, HRTEM image of 10% SnO/TiO₂-Cl (Fig. 6a) shows a clear lattice distance of 0.328 nm, which is larger than the (110) plane (0.325 nm) of rutile TiO₂.

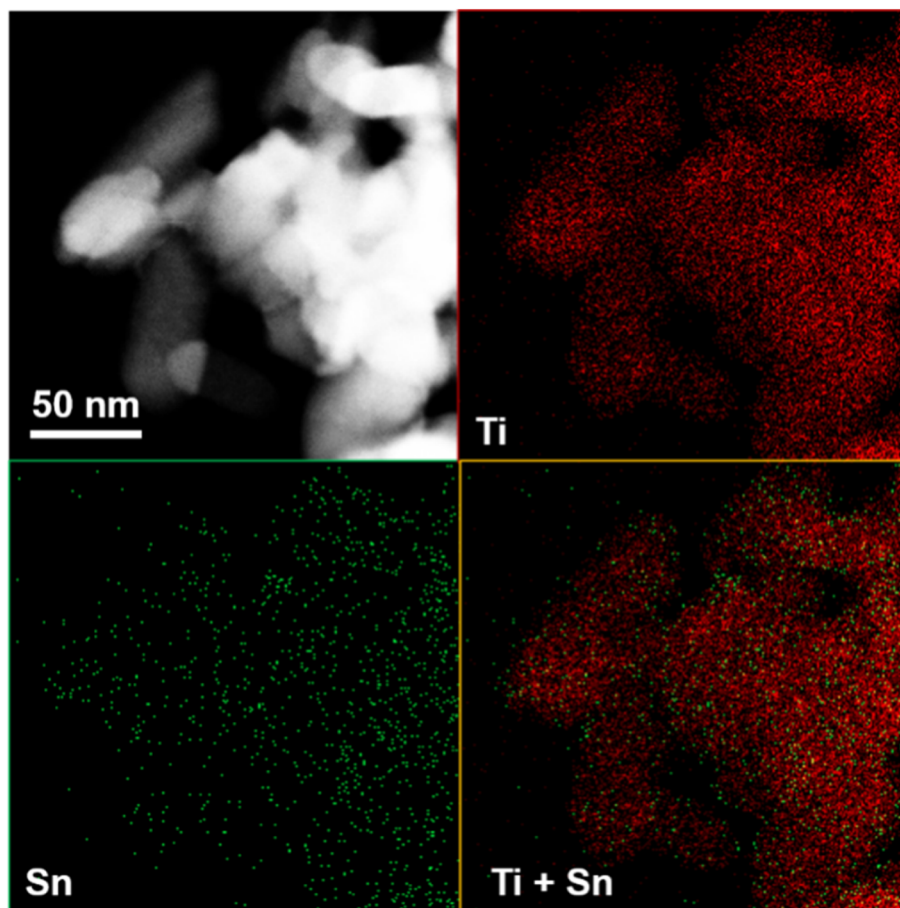


Fig. 7 HAADF-STEM-EDX elemental mapping images of 10% SnO/TiO₂-Cl.



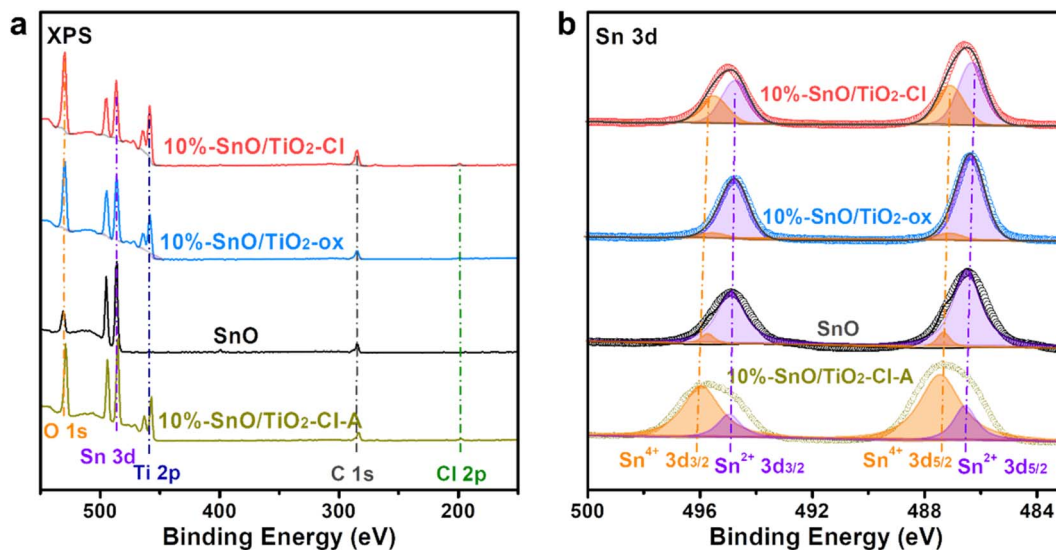


Fig. 8 (a) XPS full-spectra and (b) Sn 3d XPS spectra of 10% SnO/TiO₂-Cl, 10% SnO/TiO₂-ox, 10% SnO/TiO₂-Cl-A, and SnO.

Table 3 XPS spectra results of SnO and different SnO/TiO₂

Sample	Atom proportions				Sn 3d
	Sn (%)	O (%)	Ti (%)	Cl (%)	Sn ⁴⁺ /Sn ²⁺ (%)
10% SnO/TiO ₂ -Cl	36.9	35.5	26.7	0.75	41.6/58.4
10% SnO/TiO ₂ -ox	35.6	31.8	22.6	—	9.8/90.2
10% SnO/TiO ₂ -Cl-A	34.3	39.2	26.5	—	77.5/22.5
SnO	72.3	27.7	—	—	8.8/91.2

This demonstrates the entry of Sn ions into the lattice of rutile TiO₂, which is consistent with the XRD and Raman results (Fig. 3 and 4). In contrast, HRTEM image of 10% SnO/TiO₂-ox (Fig. 6b) displays clear lattice distances of 0.249 nm and 0.267 nm, which perfectly match with the (101) planes of rutile TiO₂ and romanite-type SnO. In HAADF-STEM-EDX element mapping images (Fig. 7), Sn elements (green) of 10% SnO/TiO₂-Cl are evenly distributed on TiO₂. In conclusion, the Sn species are highly dispersed on 10% SnO/TiO₂-Cl, whereas Sn species exist in the form of SnO NPs on 10% SnO/TiO₂-ox.

The composition and valence of various elements on the surface of SnO and SnO/TiO₂ were determined by XPS. Fig. 8a shows the XPS full spectra of these samples. For SnO/TiO₂, four groups of XPS peaks at ~529.7, 486.2, 458.3, and 284.8 eV, correspond to O 1s, Sn 3d, Ti 2p, and C 1s. In contrast, a weak peak in 10% SnO/TiO₂-Cl (Fig. 8a) appears at 198.9 eV, assigned to Cl 2p, due to the introduction of SnCl₂·2H₂O as the precursor. The atomic ratio of Sn/Cl (36.9%/0.75%, Table 3) is much higher than that of SnCl₂·2H₂O (1/2), suggesting that SnCl₂·2H₂O should transform into Sn oxides during synthesis. Moreover, atomic ratios of Sn/O on the surface of 10% SnO/TiO₂-Cl and 10% SnO/TiO₂-ox are 36.9%/35.5% and 35.6%/31.8%, indicating that Sn species should exist in the form of SnO. As shown in Fig. 8b, after calibrating and fitting of Sn 3d

XPS, two groups of peaks at 486.3, 494.8 eV, and 487.1, 495.5 eV correspond to Sn²⁺ 3d_{5/2}, 3d_{3/2} and Sn⁴⁺ 3d_{5/2}, 3d_{3/2}, respectively.⁴¹ The ratios of Sn⁴⁺/Sn²⁺ on 10% SnO/TiO₂-Cl, 10% SnO/TiO₂-ox, 10% SnO/TiO₂-Cl-A, and SnO are 41.6/58.4, 9.8/90.2, 77.5/22.5, 8.8/91.2, indicating the coexistence of Sn⁴⁺ and Sn²⁺. Combining the result of XRD, Raman, HRTEM, and XPS with the fact that the ion radii of Sn⁴⁺ (69 pm) is slightly larger than that of Ti⁴⁺ (60.5 pm), we deduce that Sn⁴⁺ should enter rutile TiO₂ lattice to lead to slight crystal expansion, meanwhile, highly dispersed SnO species are present on the surface of 10% SnO/TiO₂-ox.

Fig. 9 shows Py-IR adsorption spectra on TiO₂ and SnO/TiO₂. Firstly, no obvious ≡N⁺-H and C=N infrared adsorption peaks of pyridine hydrogen ion (PyH⁺) species were observed at 1540

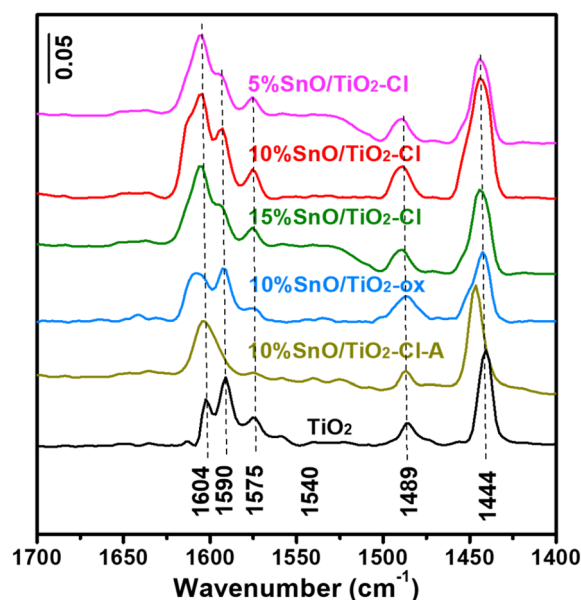
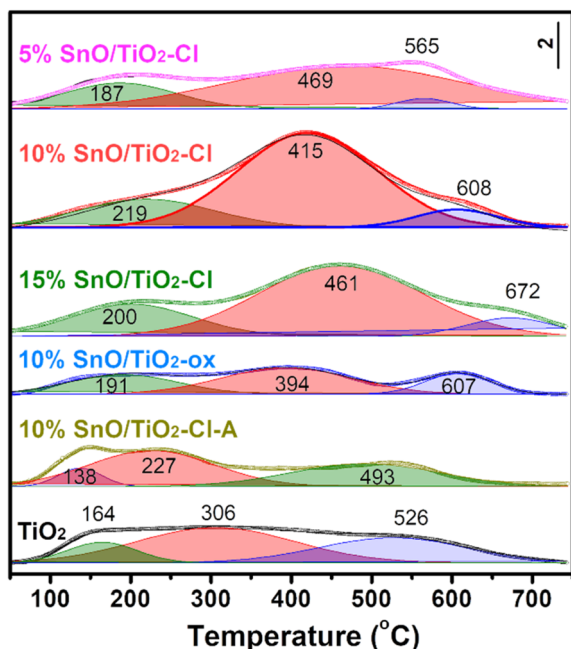


Fig. 9 Py-IR adsorption spectra of TiO₂ and different SnO/TiO₂.



Table 4 The Py-IR results of TiO₂ and different SnO/TiO₂

Sample	Adsorbed Py on Lewis acid					Acidity ($\mu\text{mol}_{\text{Py}} \text{g}^{-1}$)
	8a (A ₁)		8b (B ₁)	19a (A ₁)	19a (B ₁)	
TiO ₂	1604	1590	1575	1489	1444	34.5
5% SnO/TiO ₂ -Cl	1604	1595	1576	1489	1444	36.3
10% SnO/TiO ₂ -Cl	1608	1593	1576	1487	1442	69.8
15% SnO/TiO ₂ -Cl	1606	1595	1576	1489	1444	38.1
10% SnO/TiO ₂ -ox	1603	1591	1576	1485	1441	32.8
10% SnO/TiO ₂ -Cl-A	1605	—	1576	1487	1446	35.5

Fig. 10 NH₃-TPD of TiO₂ and different SnO/TiO₂.Table 5 The NH₃-TPD results of TiO₂ and different SnO/TiO₂

Sample	NH ₃ -TPD temperature (°C)			Acidity ($\text{mmol}_{\text{NH}_3} \text{g}^{-1}$)		
	Weak	Middle	Strong	Weak	Middle	Strong
TiO ₂	164	306	526	0.05	0.20	0.14
5% SnO/TiO ₂ -Cl	187	469	565	0.08	0.29	0.02
10% SnO/TiO ₂ -Cl	219	415	608	0.14	0.61	0.04
15% SnO/TiO ₂ -Cl	200	461	672	0.12	0.35	0.05
10% SnO/TiO ₂ -ox	191	394	607	0.12	0.20	0.08
10% SnO/TiO ₂ -Cl-A	138	227	493	0.03	0.14	0.11

and 1635 cm^{-1} on these samples, indicating the absence of Brønsted acid sites.⁴² Secondly, these Py-IR adsorption peaks on TiO₂ at 1444, 1489, 1575, 1590 and 1604 cm^{-1} are all caused by the vibration of Py adsorbed on Lewis acid sites with the vibration modes of 19a (B₁), 19a (A₁), 8b (B₁), and 8a (A₁), respectively.⁴³ Compared with TiO₂, the 19a (B₁) and 8a (A₁) vibrational adsorption peaks of SnO/TiO₂ became significantly

broaden and enhanced. The amounts of Lewis acid on these samples were semi-quantitatively analyzed by Py-IR results (Table 4). It turned out that 10% SnO/TiO₂-Cl has the highest Lewis acid amount on (69.8 $\mu\text{mol}_{\text{Py}} \text{g}^{-1}$), which is 2.0 and 2.1 times than that on TiO₂ (34.5 $\mu\text{mol}_{\text{Py}} \text{g}^{-1}$) and 10% SnO/TiO₂-ox (32.8 $\mu\text{mol}_{\text{Py}} \text{g}^{-1}$), indicating that the highly dispersed Sn species provided more Lewis acid sites.

NH₃-TPD was used to quantify the acid amount (Fig. 10). The peaks at low temperature region (α , 100–250 °C), medium temperature zone (β , 250–500 °C), and high temperature area (γ , 500–700 °C) corresponded to weak acid sites, medium strong acid sites, and strong acid sites, respectively. The acid distribution and amount of 10% SnO/TiO₂-ox and 10% SnO/TiO₂-Cl-A are similar as that of TiO₂, while NH₃-TPD curve of SnO/TiO₂ has changed greatly with the obvious increase of medium strong acid sites. Through the peak fitting of NH₃-TPD curves (Table 5), the amount of weak and strong acid sites on SnO/TiO₂-Cl were as much as those of TiO₂ and 10% SnO/TiO₂-ox, while the amount of medium strong acid sites (0.29–0.61 $\text{mmol}_{\text{NH}_3} \text{g}^{-1}$) on SnO/TiO₂-Cl was higher than those of TiO₂ and 10% SnO/TiO₂-ox (0.20 $\text{mmol}_{\text{NH}_3} \text{g}^{-1}$). With the increase of Sn content, the amount of medium strong acid sites from 0.29 $\text{mmol}_{\text{NH}_3} \text{g}^{-1}$ (5% SnO/TiO₂-Cl) first increases to 0.61 $\text{mmol}_{\text{NH}_3} \text{g}^{-1}$ (10% SnO/TiO₂-Cl), and then decreases to 0.35 $\text{mmol}_{\text{NH}_3} \text{g}^{-1}$ (15% SnO/TiO₂-Cl).

4. Conclusion

The catalytic behaviors of esterifying fatty acid and TMP over SnO/TiO₂ were studied. The structure–activity relationships between Sn species and lubricating ester synthesis reaction were dissected by changing Sn precursors, loading amounts, and treating ambiances. The optimized 10% SnO/TiO₂-Cl with SnCl₂·2H₂O as precursor after roasting in N₂ reveals the excellent esterification performance. The detailed structure and surface acidity characterization results show that Sn species highly dispersed on 10% SnO/TiO₂-Cl, as the moderate Lewis acid sites to accelerate esterification reactions between TMP and *n*-octanoic acid.

Conflicts of interest

The authors declare no competing financial interest.



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