


 Cite this: *RSC Adv.*, 2023, **13**, 2234

Arsenene as a promising sensor for the detection of H₂S: a first-principles study

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To explore the feasibility of arsenene in detecting H₂S gas, we employ the density-functional theory to investigate the geometry, electronic structure and magnetic properties of defected and doped arsenene. Point defects do not appreciably improve the sensing performance of arsenene due to small adsorption energies and charge transfer. The doping of transition metals (Ti, V, Cr, Mn, Co and Ni) introduces magnetic moments and narrows the band gap of arsenene. Transition metal (TM) dopants can enhance the interaction between H₂S and a modified arsenene substrate. Adsorption energies and charge transfers increase significantly, and the adsorption converts to chemisorption. After adsorption, the Ti and Cr-doped system's band gap change is twice that of the pristine and defective arsenene. The adsorption of H₂S changes the system properties of two TM-doped arsenenes: Ti-doped arsenene transforms from semiconductor to half-metal, and Ni-doped arsenene transforms from half-metal to conductor. Electrical signals can be observed in this process to detect H₂S molecules. Our calculations show that doping improves the detecting performance of arsenene to H₂S molecules more efficiently than defects. Our results indicate that arsenene has a promising future in developing H₂S gas sensors.

Received 19th October 2022

Accepted 3rd January 2023

DOI: 10.1039/d2ra06588f

rsc.li/rsc-advances

1. Introduction

Hydrogen sulfide (H₂S) gas is an inorganic compound, colourless and flammable, widely used in industrial production. It has high neurotoxic potency. Inhalation of small amounts of H₂S gas can be lethal. Researchers have recently investigated more and more sensing materials to detect toxic gases.^{1–3} Two-dimensional (2D) atom-thick materials (graphene,⁴ germanene,⁵ silicene,⁶ phosphorene,⁷ and antimonene⁸) are prospective nanometer gas sensors due to their unique characteristics and ultra-high specific surface area.^{9,10} However, many two-dimensional materials cannot exhibit high sensitivity and selectivity and recover quickly simultaneously. For example, low resistance and a lack of an inherent band gap limit the use of graphene in sensors.¹¹ Perfect graphene has limited application potential because the ordinary gas molecules are physically adsorbed on it.^{12,13} Materials with adjustable gas-sensitive properties offer the possibility to solve this problem.¹⁴

Researchers proposed arsenene in 2014,^{11,15} and its structure is usually buckled or puckered. In addition to high conductivity and low electrical noise, arsenene has a high surface-to-volume ratio. Arsenene can better adsorb some gases^{16–18} because of the sp³ hybridization and more active adsorption sites provided by the buckled surface structure. It can effectively and selectively

interact with gases.^{19,20} Previous studies have shown that arsenene has a tunable electronic band structure under strain and electric fields.^{21–23} It has promising potential in 2D nanospintronics devices²⁴ and gas sensing. Like black phosphorus, the carrier mobility of arsenene is as high as several thousand cm² V^{−1} s^{−1}.²⁵ This feature enables arsenene to respond quickly when detecting gases. In particular, the physical and chemical properties of arsenene can be regulated effectively by introducing adsorption and doping.^{26–28} Compared with graphene and silicene, arsenene has lower formation energy and is prone to point defects.^{29,30} The formation energy of Stone–Wales defects is 1.63 eV and smaller than other typical point defects. The relatively small value causes Stone–Wales defects easy to occur.^{31,32} Defects and doping modulate the properties of two-dimensional materials to a great extent. According to Wang *et al.*'s study, the defects in h-BN can exhibit strong defect-photon coupling and have exciting properties of photoluminescence (PL) upconversion (anti-Stokes process).³³ Tran *et al.* observed single photon emission in h-BN induced by B vacancy-related defects.³⁴ The magnetic moments of the system can be introduced by vacancy defects.³¹ The Pt vacancies and Ruderman–Kittel–Kasuya–Yoshida (RKKY) would induce surface magnetism. Therefore, layer-dependent magnetism can be observed.³⁵ Shawkat *et al.* reported band gap reduction and transitions from semiconductor to metal triggered by Se vacancies in PtSe₂.³⁶ Impurities atoms can modify the electronic and magnetic properties of 2D materials. Tullurene can be better employed in spintronic devices because of the

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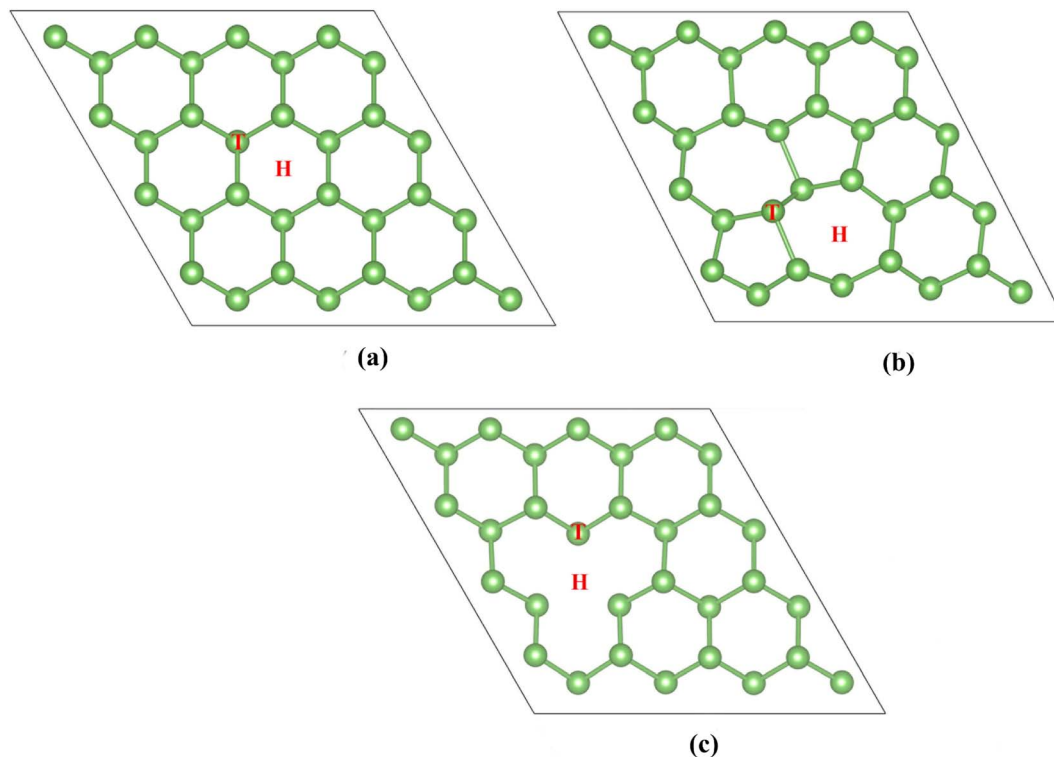


Fig. 1 Optimized structures of (a) P-As, (b) SW-As, and (c) SV-As.

contributions of TM dopants to magnetism.³⁷ The doping of O, S and Se in arsenene will behave as metallic characters.³⁸ For blue phosphorene (BP), dopants (Ce, Ti, Ti-Ti) will narrow the band gap.³⁹ V, Cr, and Mn doping in bismuthene can induce magnetic metal, semiconducting and half-metal properties, respectively.⁴⁰ The magnetic moment values and ground states of TM-doped arsenene differ according to density functional theory (DFT) calculations.^{27,41,42} Researchers investigated the properties of TM-doped arsenene in two ways: standard DFT and DFT+U. The constant U_{eff} was set as 2.5 eV for all TM atoms.^{27,41,42} Employing DFT+U with more accurate U_{eff} values to investigate the doping of transition metals is necessary. According to a previous report, the adsorption of pristine arsenene to small gas molecules (such as CO, NO, NO₂, and NH₃) is between physical and chemical adsorption.^{16,18,43–45}

Zhou *et al.* discovered that TM dopants could improve the interaction between arsenene substrate and HCN molecules more efficiently than defects. Therefore, more research is required to determine how defects and doping affect the sensitivity of arsenene to H₂S. Because in previous studies, the DFT method is relatively consistent with the experimental results. We can employ the DFT method to calculate 2D materials' properties.

Our work is mainly divided into the following parts: firstly, we investigate the effects of defects (SW-defects and vacancy) on the sensing performance of arsenene to H₂S and compare the best adsorption configuration. Then we further illustrate the effects of H₂S on TM-doped systems' magnetic and electronic properties. Our calculations find that TM doping is more effective in improving the sensing performance of arsenene to

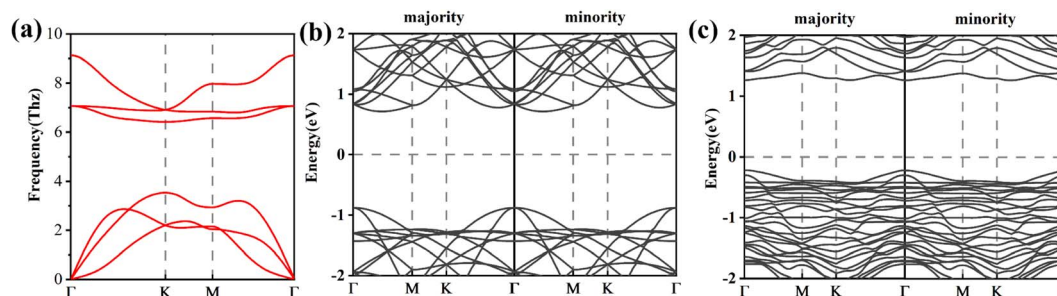


Fig. 2 (a) Phonon dispersion curves and (b) band structures of P-As. (c) Band structures of SW-As.



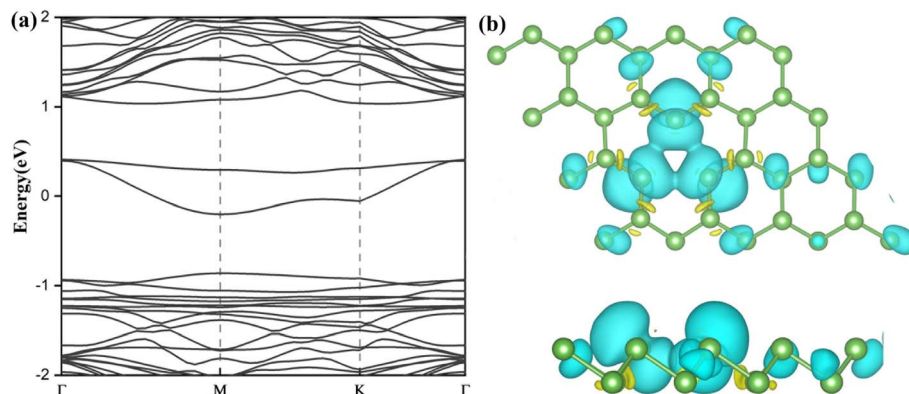


Fig. 3 (a) Band structures and (b) top and side views of spin densities of SV-As. The values of isosurface are set to $5 \times 10^{-7} \text{ e} \text{ \AA}^{-3}$. Yellow and blue colors represent spin-up and spin-down states, respectively.

H₂S than defects. The increased adsorption energy and charge transfer between the TM-doped arsenene and H₂S molecules indicate stronger interactions. We hope that this work could help develop arsenene-based H₂S gas sensors.

2. Computational methods

This paper performs first-principles calculations with the help of spin-polarised density functional theory (DFT), which is implemented in the Vienna Ab initio Simulation Package (VASP).^{46,47} The projected augmented wave (PAW) pseudopotential⁴⁹ is used to model the electron-ion interaction. To describe the exchange-correlation energy between electrons, we use the generalized gradient approximation (GGA) of the PBE functional.⁴⁸ Next, we optimize the structure until the system's total energy achieves an accuracy of 10^{-6} in structural relaxation. The residual forces of atoms are set as $0.02 \text{ eV} \text{ \AA}^{-1}$. A 15 \AA vacuum layer is set to the perpendicular direction to avoid interactions with neighboring slabs. Furthermore, we employ an $8 \times 8 \times 1$ Monkhorst-Pack k -point sampling and set the kinetic energy cutoff as 500 eV. The D3-Grimme correction (DFT-D3)⁵⁰ is used to consider the weak vdW interaction. A $4 \times 4 \times 1$ β -As supercell containing 32 As atoms is employed to model the arsenene monolayer.

Considering the electron correlation effects of the localized 3d orbitals, we add the Hubbard-U parameters to the energy functional of GGA.⁵¹ Referring to previous DFT+U research⁵²⁻⁵⁴ and our calculation details, the reasonable values of $U_{\text{eff}} = U - J$ for Ti, V, Cr, Mn, Co and Ni atoms are set to 2.5, 3.0, 3.55, 3.5, 2.0 and 2.45 eV, respectively.

3. Results and discussions

3.1 H₂S adsorbed on pristine and defective arsenene

The first part investigates the effect of point defects on the adsorption properties. Arsenene adopts a buckling structure and considers two common point defects: single vacancy and Stone-Wales. Fig. 1 shows the relaxed structures of perfect arsenene (P-As), single vacancy (SV-As) and Stone-Wales

defected arsenene (SW-As). The pristine arsenene's lattice constant is 3.61 \AA , and the bond length of the As-As is 2.52 \AA , identical to previous calculations.^{56,57}

Before structure relaxation, we rotate the bond of two neighboring As atoms by 90° to obtain SW-As. We remove an atom from the pristine arsenene and optimize the structure to get the SV defect. After structural optimization, the SW-As system forms the 55-77 structure. The SV defect is centrosymmetric. The three edge atoms of SV-As system are equivalent, and the distance between two edge atoms is 3.53 \AA . After introducing point defects, the changes in geometry structure of arsenene are analogous to that of other two-dimensional nanomaterials such as graphene,⁵⁸ phosphorene,⁵⁹ and antimonene.⁶⁰

We calculate the spin-polarized band structures of defected arsenene as depicted in Fig. 2 and 3. As an indirect semiconductor, the P-As system has a band gap of 1.70 eV. The conduct band minimum (CBM) is situated between Γ and M points. The valence band maximum (VBM) is located at the high symmetric Γ points. P-As system's parameters in our calculations are consistent with previous studies.^{61,62} At the same time, pristine arsenene is a nonmagnetic semiconductor because of

Table 1 The distance between the H₂S molecules and the arsenene substrate (d), adsorption energies (E_{ads}) and charge transfers (Q) of P-As, SV-As and SW-As

System	Site		d (\AA)	Q (e)	E_{ads} (eV)
P-As	Top	Vertical	3.88	0.016	-0.107
		Parallel	3.91	0.015	-0.112
	Hollow	Vertical	3.72	0.033	-0.161
		Parallel	3.47	0.040	-0.207
SV-As	Top	Vertical	3.48	0.052	-0.191
		Parallel	3.48	0.050	-0.250
	Hollow	Vertical	3.64	0.022	-0.158
		Parallel	3.75	0.012	-0.154
SW-As	Top	Vertical	3.53	0.039	-0.229
		Parallel	3.54	0.063	-0.233
	Hollow	Vertical	3.48	0.030	-0.207
		Parallel	3.47	0.033	-0.186



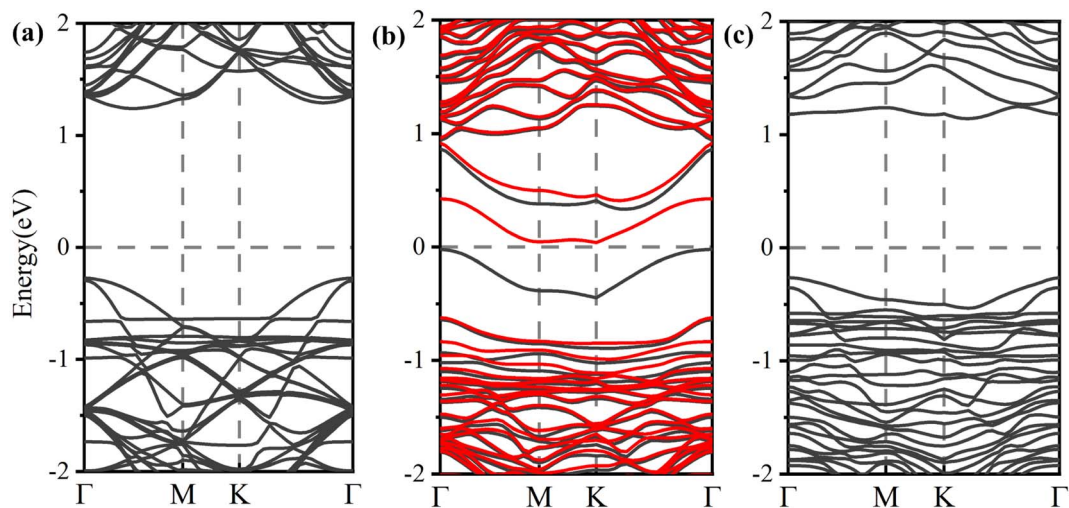


Fig. 4 Electronic band structures for (a) P-AS, (b) SV-AS, and (c) SW-AS with H₂S molecule adsorbed, respectively. The black lines represents spin-up states, and the red lines represents spin-down states.

its symmetrical band structure. The P-AS system is dynamically stable because its phonon spectrum (plotted in Fig. 2(a)) does not contain imaginary phonon modes.

The defective arsenene loses hexagonal lattice symmetry, and its electronic structure changes. Fig. 2(c) shows that the

indirect band gap of SW-As system decreases to 1.48 eV. The CBM moves to Γ point. The charge around the defect localizes strongly due to the flatter band structure. The state in gap (plotted in Fig. 3(a)) appears near the Fermi level because of dangling bonds induced by vacancy defect.

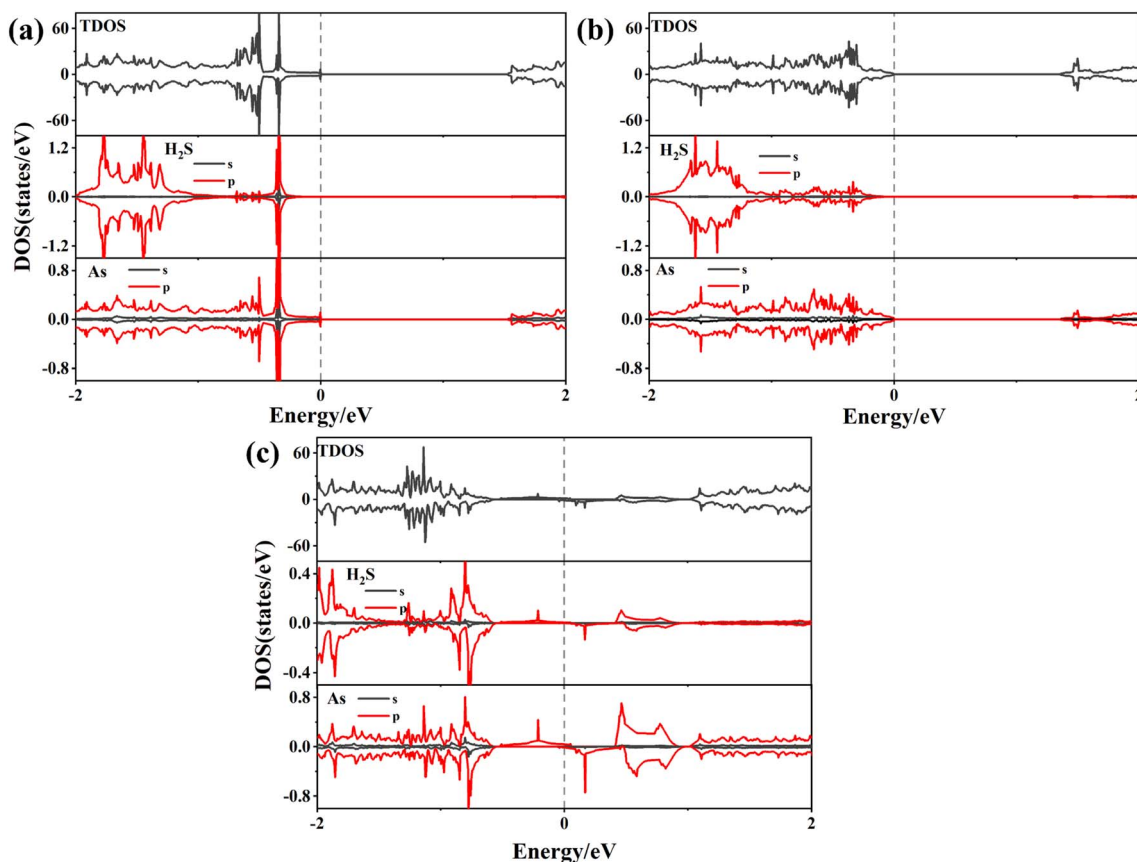


Fig. 5 Density of states of (a) H₂S/P-AS, (b) H₂S/SV-AS, and (c) H₂S/SW-AS.



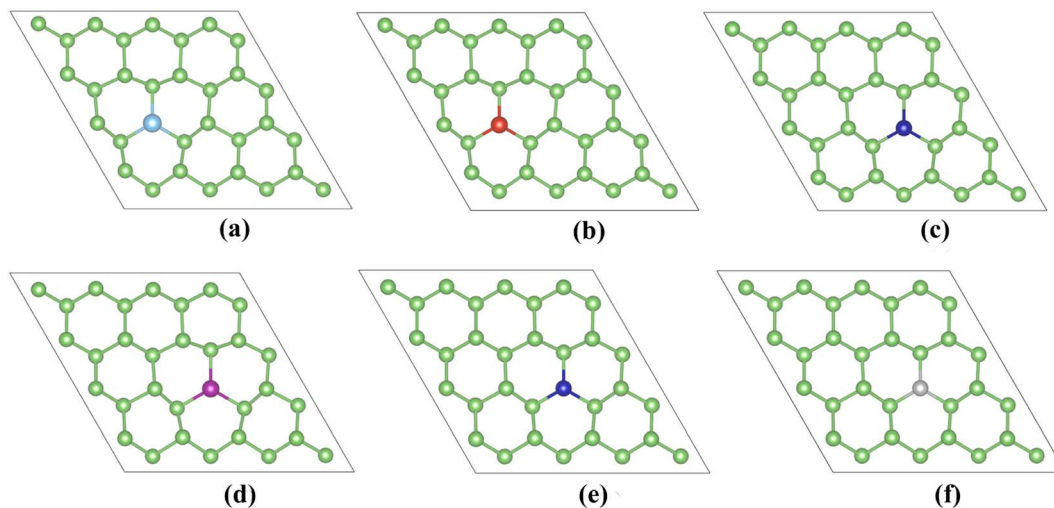


Fig. 6 Relaxed structures of (a) TiAs, (b) VAs, (c) CrAs, (d) MnAs, (e) CoAs and (f) NiAs.

We use the spin-polarized calculation to investigate the magnetic properties of SV-As system. Due to an unpaired electron and dangling bond around the vacancy defect, the SV-As system's total magnetic moment is $1\mu_B$. The structure of SV defect is centro-symmetrical, so the spin density is symmetrically localized (plotted in Fig. 3(b)). The electronic properties of perfect arsenene are changed by introducing point defects.

Next, we study the adsorption of H_2S on the pristine arsenene and defective arsenene. Obtaining the best adsorption configuration is beneficial to the following study. Considering the geometric structure of arsenene, we set two adsorption sites on the substrate (labelled in Fig. 1): hollow and top sites. Two adsorption models for H_2S are considered. One is that the H_2S molecule is perpendicular to the substrate, with the S atom close to the surface. The other is H_2S molecule parallel to the arsenene substrate. On the whole, we can observe 12 types of initial adsorption configurations. The adsorption energies of H_2S molecule on modified arsenene substrate are defined as:

$$E_{\text{ads}} = E_{H_2S/As} - E_{As} - E_{H_2S} \quad (1)$$

where $E_{H_2S/As}$ and E_{As} represent the total energy of the arsenene system with and without H_2S molecules adsorbed; E_{H_2S} is the energy of the isolated H_2S molecules. When calculating E_{H_2S} , we approximate it to the energy of an H_2S molecule in the cubic supercell (with a length of 15 \AA). The stability of H_2S adsorbed on arsenene substrate can be evaluated by this equation. This definition indicates that the negative value of E_{ads} represents the adsorption stability.

Charge analysis can intuitively obtain the electron flow direction after the interaction of each fragment. Moreover, to explore the nature of chemical bonds, we utilize Bader charge analysis⁵⁵ and calculate the charge density difference (CDD). The definition is defined as follows:

$$\Delta\rho = \rho_{H_2S/As} - \rho_{As} - \rho_{H_2S} \quad (2)$$

$\rho_{H_2S/As}$, ρ_{As} are the charge densities of the arsenene system with and without H_2S molecules adsorbed; ρ_{H_2S} represents the isolated H_2S molecules' charge densities.

After complete structure optimization, we summarize the corresponding parameters in Table 1. The variation of adsorption energy (E_{ads}) is from -0.107 eV to -0.250 eV . It indicates that H_2S molecules weakly interact with arsenene substrates. Bader charge analysis obtains the charge transfer (Q) ($0.012e$ to $0.063e$) between the arsenene substrate and H_2S molecule. The positive values of Q indicated that H_2S molecules obtained charge as the acceptor.

The low E_{ads} and Q values suggest the adsorption by physical force. When H_2S molecules are parallel to the top sites of SW-As and SV-As, the adsorption energies are maximum. The parameters in Table 1 show that this is the most stable adsorption structure for SW-As and SV-As. The values of E_{ads} are equal to -0.233 eV (SW-As) and -0.250 eV (SV-As), which are larger than the E_{ads} of P-As (-0.112 eV). After structural optimization, the H_2S molecule is placed above the arsenene substrate in an inclined structure. The interaction distance between the H_2S molecule and arsenene substrate is relatively long ($3.47\text{--}3.91 \text{ \AA}$) in all adsorption systems. In addition, the charge transfer Q of SW-AS ($0.063e$) and SV-AS ($0.050e$) are larger than that of P-AS

Table 2 Calculated averaged bond length d_{TM-As} ($TM-As_1 = TM-As_2 = TM-As_3$), bond angles (θ) ($As_1-TM-As_2 = As_2-TM-As_3 = As_3-TM-As_1$), binding energy (E_b), total magnetic moment (M_{tot}), local magnetic moment of TM dopants (M_{TM}) and local magnetic moment of the nearest-surrounding As atoms (M_{As}) for TM-doped arsenene

System	d_{TM-As} (\AA)	θ ($^\circ$)	E_b (eV)	M_{tot} (μ_B)	M_{TM} (μ_B)	M_{As} (μ_B)
TiAs	2.58	103.07	-4.84	1.0	1.164	-0.050
VAs	2.55	101.00	-7.60	2.0	2.561	-0.152
CrAs	2.56	100.96	-3.44	3.0	4.060	-0.271
MnAs	2.56	112.23	-3.67	4.0	4.648	-0.170
CoAs	2.37	108.24	-4.30	2.0	2.033	-0.039
NiAs	2.31	103.53	-4.10	1.0	0.641	0.051



(0.040e). In conclusion, the introduction of SW- and SV-defects improves the adsorption stability to a certain extent. However, the adsorption is still physisorption.

The modulation of H₂S adsorption on the electronic band structure of the arsenene system needs further study. The band structures (depicted in Fig. 4) and the density of states (DOS) (plotted in Fig. 5) are calculated by spin-polarized calculations. The band gaps after H₂S adsorption decrease by 0.078 eV (P-As), 0.076 eV (SV-As) and 0.044 eV (SW-As), respectively. As shown in Fig. 5, the H₂S 2p orbital and As 4p orbital hybridizes near the Fermi level. The band gaps decrease due to the shift of Fermi level after adsorption. Fig. 4(b) illustrates that the degeneracy of two channels at the Γ point vanishes after adsorption. Impurity states arise above the Fermi level due to adsorption. After adsorption, P-As and SW-As are still non-magnetic systems without spin splitting. The magnetic moment of SV-As remained 1 μ_B . The two primary determinants of electrical conductivity, carrier concentration and mobility, are impacted by band gaps. The following equation can define the relationship between band gap and electrical conductivity:

$$\sigma \propto \exp\left(\frac{-E_g}{2kT}\right) \quad (3)$$

where σ represents the electrical conductivity of the systems, T and k are the thermodynamic temperature and Boltzmann's constant, respectively. After adsorption, the conductivity of the arsenene substrate changes. However, the change amplitude is too small to detect. Therefore, we need methods to make arsenene more effective in detecting H₂S.

3.2 H₂S adsorbed on TM-doped arsenene

Due to the limited improvement of arsenene sensing performance by point defects, we introduce transition metal dopants (Ti, V, Cr, Mn, Co, and Ni). TM-doped arsenene is obtained by substituting an As atom from a 4 × 4 × 1 pristine arsenene supercell with a TM atom. Fig. 6 depicts the doped structure after structure relaxation. The geometric deformation is not apparent after structure relaxation.

After structure optimization, the bond lengths of TM-As (shown in Table 2) for TM-doped arsenene are 2.58 Å (Ti-doped), 2.55 Å (V-doped), 2.56 Å (Cr-doped), 2.56 Å (Mn-doped), 2.37 Å

(Co-doped) and 2.31 Å (Ni-doped). The difference in ionic radius from Ti to Ni may result in the variation of the bond lengths of TM-As. Our calculation results are quite different from the GGA results of previous studies.⁴¹ The difference may be because strongly correlated systems containing 3d electrons cannot be accurately described by the GGA functional. The stability of TM-doped arsenene can be verified by binding energy (E_b). We can utilize the following equation to calculate the binding energy (E_b):

$$E_b = E_{\text{Doped-As}} - E_{\text{V-As}} - E_{\text{TM}} \quad (4)$$

in which $E_{\text{Doped-As}}$ is the total energy of TM-doped arsenene, $E_{\text{V-As}}$ is the energy of the arsenene substrate with one As atom removed, and E_{TM} represents the energy of an isolated TM atom at the same position as in the TM-doped arsenene, respectively. The binding energies of TiAs, VAs, CrAs, MnAs, CoAs and NiAs are calculated as -4.84, -7.60, -3.44, -3.67, -4.30 and -4.10 eV by eqn (4). The negative and large values of the E_b indicate that the TM-doped systems are stable. In addition, we explore the kinetic stability of the TM-doped systems. The phonon dispersion curves of the TiAs and CrAs are calculated. Fig. 7 shows that the phonon spectrum has no imaginary frequencies and exhibits linear dispersion near the gamma point. The phonon spectrum further illustrates the structural stability of the TM-doped systems. The ground states of TM-doped systems are magnetic. The total magnetic moments of TiAs, VAs, CrAs, MnAs, CoAs and NiAs are 1.0, 2.0, 3.0, 4.0, 2.0 and 1.0 μ_B per supercell, respectively.

Differently, the ground state of the Ni-doped system was reported to be non-magnetic by previous studies.^{41,42,63} The difference in magnetic moments is due to previous studies taken by the GGA functional. This functional will ignore the Coulomb interactions between the localized 3d orbitals. Table 2 shows the antiferromagnetic interactions between TM and the adjacent As atoms in the TiAs, VAs, CrAs, MnAs and CoAs. At the same time, we can observe long-range ferromagnetic interactions between Ni and adjacent As atoms.

The electronic and magnetic properties of arsenene can be influenced by introducing TM dopants. We analyze the electronic band structures (shown in Fig. 8) and density of states (DOS) (depicted in Fig. 10) of doping systems before and after

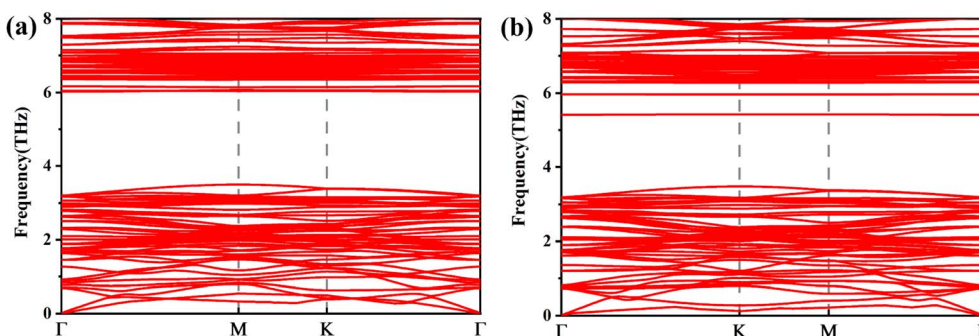


Fig. 7 Phonon dispersion curves of (a) TiAs, (b) CrAs.



adsorption. The structure of DOS for TM-doped systems is asymmetric. TM impurities can introduce magnetism into arsenene. The corresponding band structure is spin splitting.

Ti-doped arsenene is an indirect semiconductor (Fig. 8(a)). The majority channels display a band gap of 1.18 eV. The indirect band gap in the minority states is 1.42 eV. The spin-up channels of V-doped arsenene have a direct band gap of 1.12 eV and behave metallic character (Fig. 8(b)). On the contrary, the spin-down channels display an indirect band gap of 1.45 eV and indicate the half-metallic property of the V-doped system. The spin-up channels of MnAs, CoAs and NiAs (see Fig. 8(d)–(f)) show a two-fold degenerate state crossing the Fermi level (E_F) and display metallic properties. However, the spin-down channels are semiconducting and indicate the half-metallic behaviors of MnAs, CoAs and NiAs. The CrAs has two majority degenerate channels above the Fermi level. Two minority channels degenerate below the Fermi level. This phenomenon is due to impurity states generated by doping (Fig. 8(c)). The minority states display a direct band gap of 1.41 eV and indicate a half-metallic character. The impurity

states near the Fermi level (E_F) suggest that VAs, CrAs, MnAs, CoAs and NiAs are spin half-metallic materials. To sum up, TM dopants can effectively regulate the electronic and magnetic properties of arsenene.

The H_2S molecules are set parallel to doped arsenene because Table 1 shows that this configuration is the most stable. Besides, the S atom is located above the TM atom, so the influence of TM dopants on adsorption can be considered simultaneously.

We summarize the S–TM bond lengths in Table 3. The bond lengths range from 2.45 Å–2.61 Å. The interaction distance of TM-doped arsenene is shorter than that of pristine and defective arsenene. This suggests a stronger interaction between the H_2S molecules and the TM dopants. In addition, the E_{ads} of TM-doped arsenene varies from -0.89 eV to -1.21 eV, effectively enlarged by doping.

Meanwhile, the charge transfer (Q) of the H_2S molecule significantly increases. For Ti, V, Cr and Mn atoms that easily lose electrons, the H_2S molecule obtains $0.136e$, $0.118e$, $0.168e$ and $0.093e$ charges from TM-doped arsenene, respectively. For

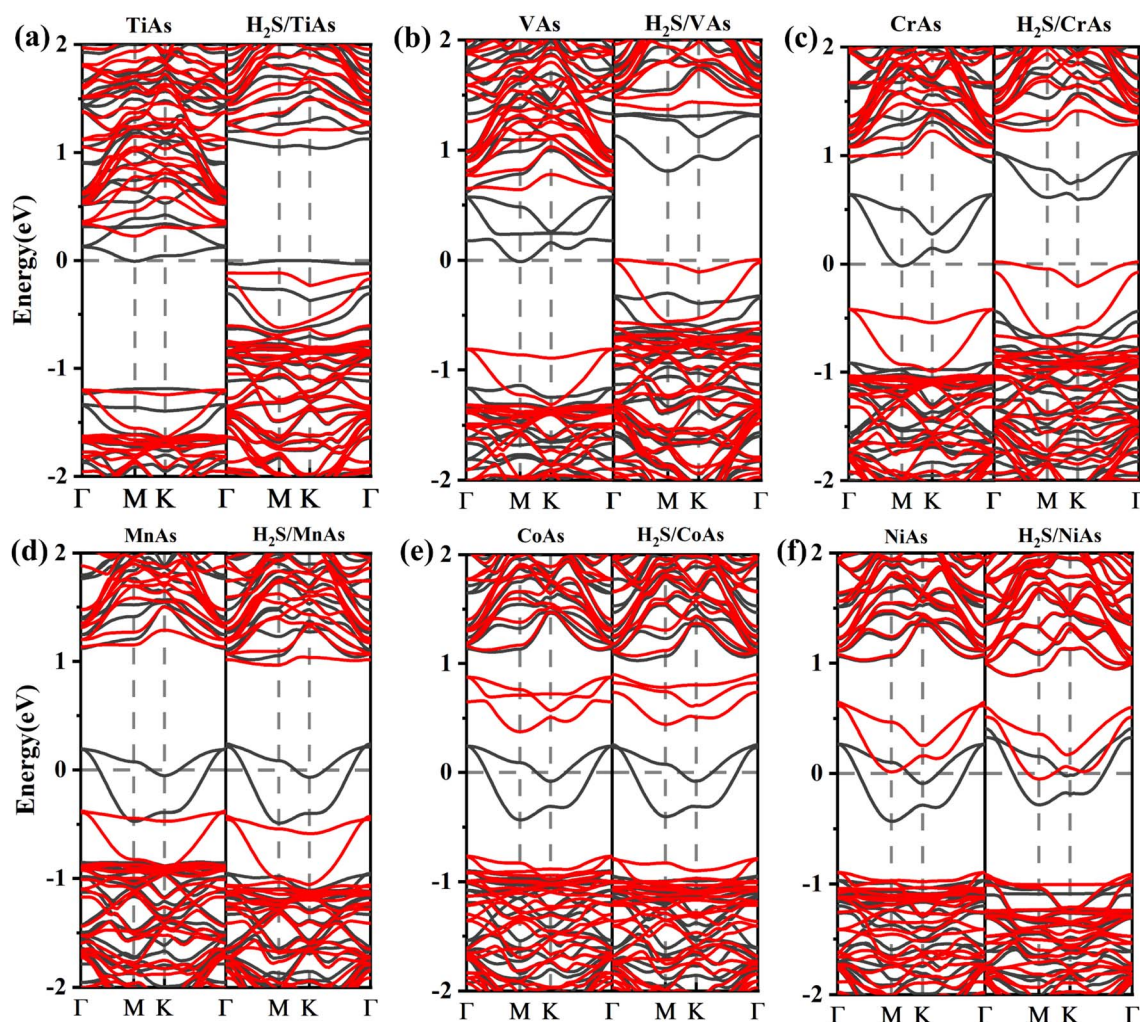


Fig. 8 (a)–(f) Electronic band structures for TiAs, VAs, CrAs, MnAs, CoAs and NiAs with and without H_2S molecule adsorbed, respectively. The black lines show spin-up band structures, and the red lines represent spin down band structures.



Table 3 The adsorption energy (E_{ads}), charge transfer (Q), and bond lengths of S–TM ($d_{\text{S-TM}}$) of TM-doped arsenene with H_2S adsorbed. The total magnetic moment (M_{tot}), the local magnetic moments of TM atoms (M_{TM}), H_2S molecule ($M_{\text{H}_2\text{S}}$) and As atoms (M_{As}) are also listed

System	E_{ads} (eV)	Q (e)	$d_{\text{S-TM}}$ (Å)	M_{tot} (μ_{B})	M_{TM} (μ_{B})	$M_{\text{H}_2\text{S}}$ (μ_{B})	M_{As} (μ_{B})		
							As ₁	As ₂	As ₃
TiAs	−1.16	0.136	2.61	1.0	1.014	0.040	−0.040	−0.030	−0.024
VAs	−1.21	0.118	2.57	2.0	2.406	0.029	−0.136	−0.120	−0.117
CrAs	−1.01	0.168	2.57	3.0	3.859	0.028	−0.257	−0.255	−0.202
MnAs	−0.98	0.093	2.61	4.0	4.507	0.057	−0.165	−0.159	−0.155
CoAs	−1.07	−0.107	2.45	2.0	1.956	0.063	−0.040	−0.028	−0.028
NiAs	−0.89	−0.101	2.53	0.7	0.468	0.011	0.044	0.040	0.028

Co and Ni atoms that easily gain electrons, the H_2S molecules lose 0.107e and 0.101e charge, respectively. These results show that TM dopants can significantly improve arsenene's sensing performance for detecting H_2S molecules. Meanwhile, the adsorption shifts to chemisorption (Fig. 9).

To better understand the detailed changes in the bandgap properties of the TM-doped arsenene after adsorption, the spin-polarized electronic band structures (shown in Fig. 8) and total density of states (TDOS) (depicted in Fig. 10) are calculated. Furthermore, we calculate the projected density of the states (PDOS) of H_2S molecules, TM dopant atoms and adjacent As atoms (illustrated in Fig. 11). The Fermi level of Ti-doped system moves down after adsorption (Fig. 10(a)). It makes the system transforms from an n-type doped semiconductor to half metal. The size of band gap of the system decreases. At the Fermi level, the H_2S s orbitals hybridize with Ti 3d and As 4p orbitals, which

mainly contributes to the band near the Fermi level. At the bottom of the conduction band, the s and p orbitals of H_2S hybridize with Ti 3d and As 4p orbitals. After adsorption, the band gap increases, and the Fermi level of V-doped arsenene moves down from CBM to VBM. The half-metallic behaviors of the system remain unchanged. Before adsorption, the Fermi level passes through the majority states. The spin-down band structure is semiconducting. After adsorption, the Fermi level passes through the minority states. The spin-up band structure is semiconducting. The spin-down channels near Fermi level mainly originate from the hybridization of V 3d and As 4p orbitals. The Fermi level (E_{F}) of CrAs moves down, and the size of band gap increases after adsorption. The degeneracy of spin-down channel at Γ before adsorption disappears. The Cr 3d, H_2S 3p, and As 4p orbitals hybridize at around 0.5 eV. The hybridization make Fermi level pass through the spin-down

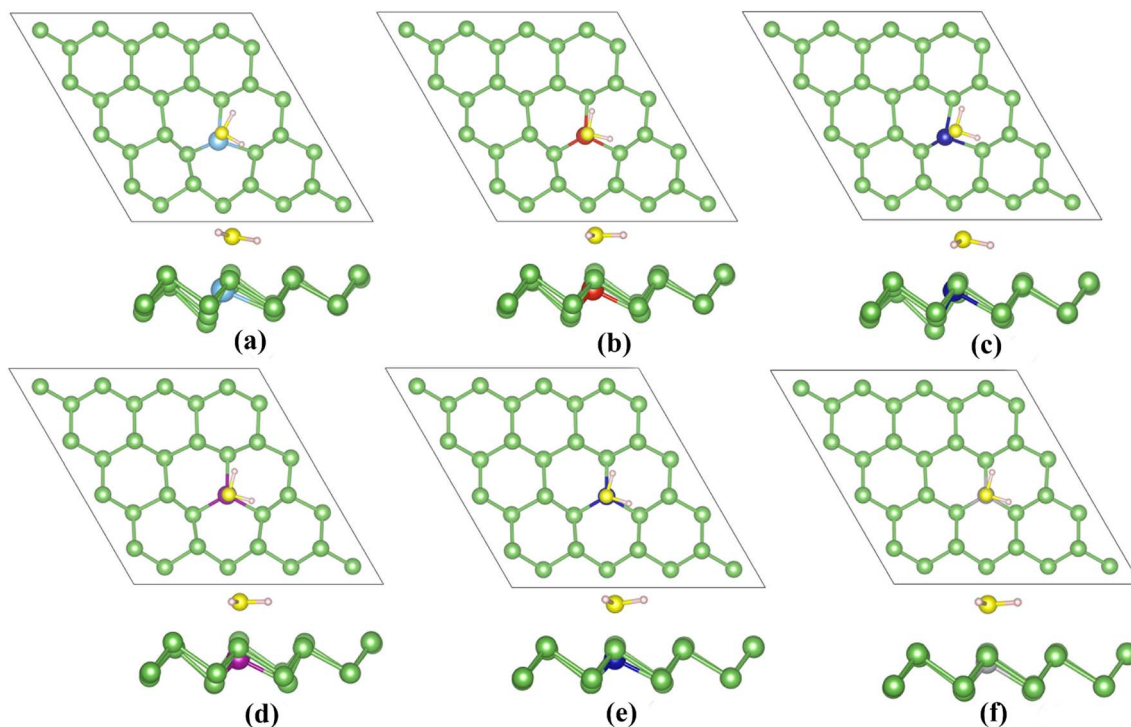


Fig. 9 (a)–(f) Relaxed adsorption structures of (a) $\text{H}_2\text{S}/\text{TiAs}$, (b) $\text{H}_2\text{S}/\text{VAs}$, (c) $\text{H}_2\text{S}/\text{CrAs}$, (d) $\text{H}_2\text{S}/\text{MnAs}$, (e) $\text{H}_2\text{S}/\text{CoAs}$ and (f) $\text{H}_2\text{S}/\text{NiAs}$.



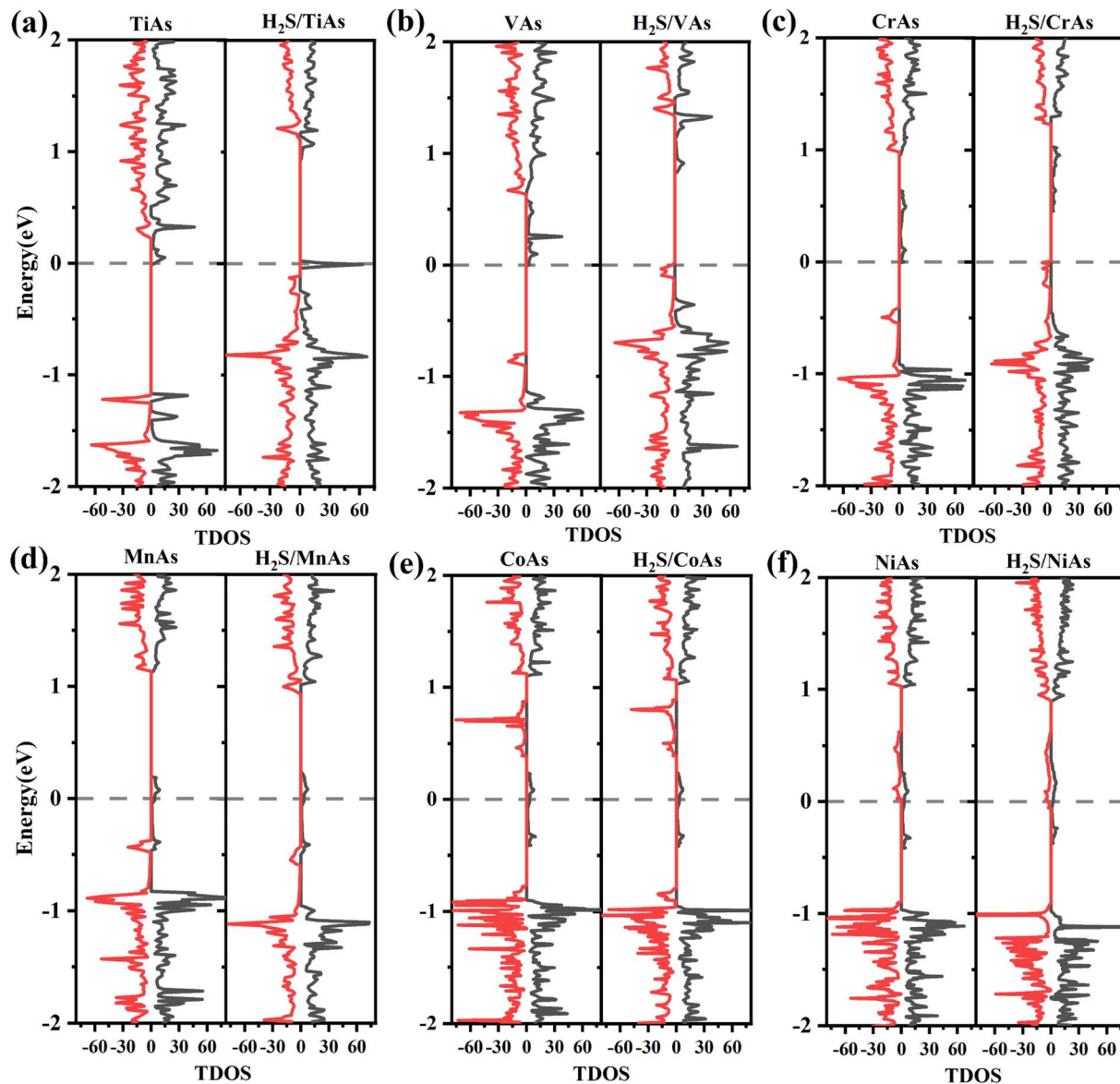


Fig. 10 (a)–(f) Spin-polarized density of states (DOS) for TiAs, VAs, CrAs, MnAs, CoAs and NiAs with and without H₂S molecule adsorbed, respectively.

channels after adsorption. The band gap of the Mn-doped arsenene is reduced owing to the interactions of H₂S 3p, Mn 3d and As 4p orbitals. New spin-down channels appear, but the system is still half-metallic. The Co-doped arsenene remains half-metallic before and after adsorption. The Fermi level passes through two spin-up channels that degenerate at the Γ point before adsorption. After adsorption, the degeneracy disappears and the corresponding density of states increases. Fig. 11 shows that this change is due to the interactions of H₂S 3p, Co 3d and As 4p orbitals. After adsorption, the Ni-doped system's band gap decreases, and the system changes from half-metal to conductor. The degeneracy at Γ points disappears after adsorption. Ni 3d and As 4p orbitals mainly contribute to the DOS near the Fermi level. The strong interaction between the dopants and H₂S molecules leads to the conversion of

physisorption to chemisorption. The band gaps of the system change significantly after adsorption. As shown in eqn (3) above, the electrical conductivity of the system varies greatly. After adsorption, the Ti-doped arsenene changes from semiconductor to half metal, and the Ni-doped arsenene changes from half metal to conductor. The change in electronic properties of the system will generate electrical signals in the substrate. We can detect this signal and apply arsenene to sensors.

The magnetic properties of TM-doped arsenene are affected by H₂S adsorption. The magnetization density distribution of TM-doped systems with H₂S adsorbed is plotted in Fig. 12. It mainly includes the TM atom, H₂S molecule and three surrounding As atoms. Moreover, we calculate the corresponding magnetic moments of TM dopants, H₂S molecules and



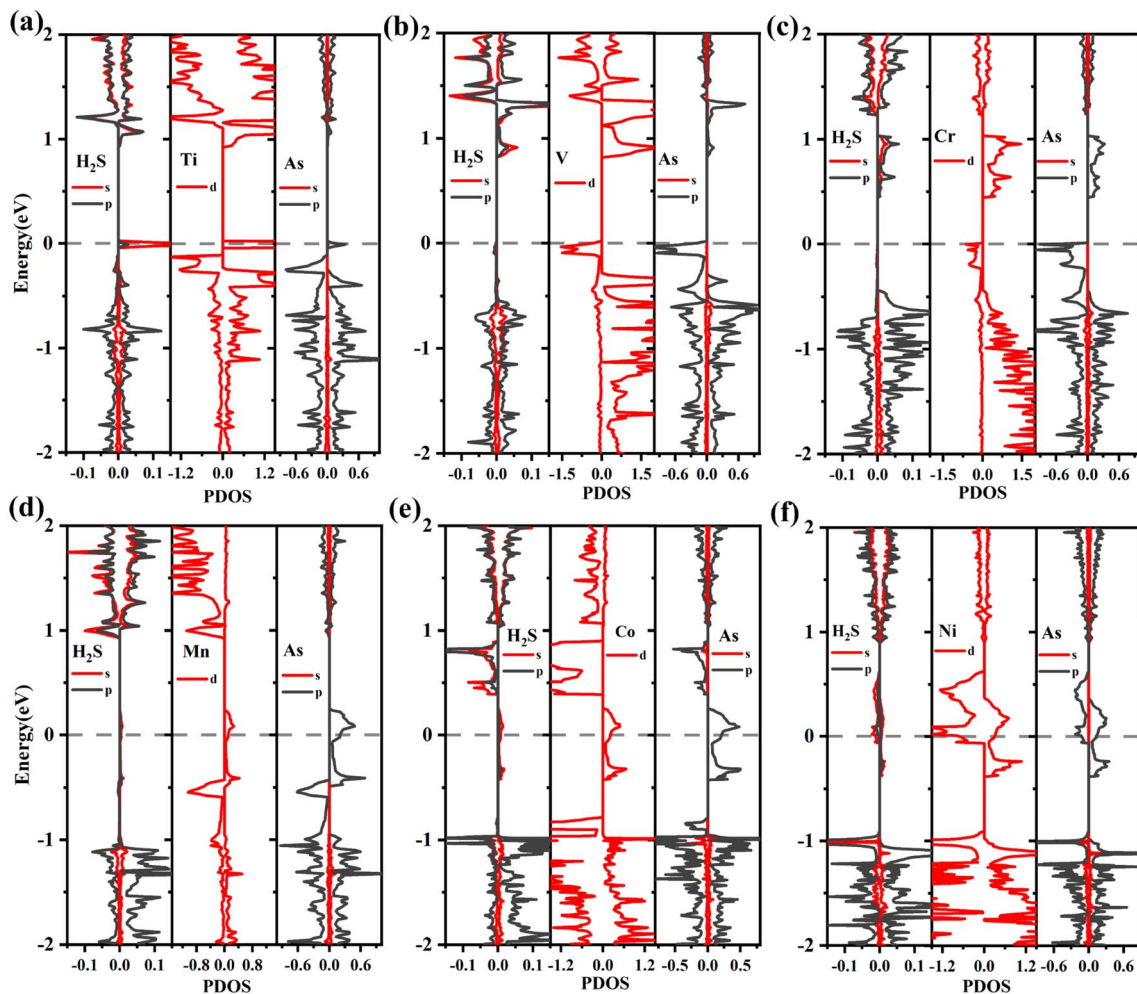


Fig. 11 (a)–(f) The PDOS of TM-doped arsenene with H₂S adsorbed: (a) H₂S/TiAs, (b) H₂S/VAs, (c) H₂S/CrAs, (d) H₂S/MnAs, (e) H₂S/CoAs and (f) H₂S/NiAs.

adjacent As atoms. We label the adjacent As atoms as As₁, As₂, and As₃ in Table 3, where As₁ has the shortest distance from the TM atom.

The ground states of TiAs, VAs, CrAs, MnAs, CoAs and NiAs have magnetism after adsorption. The total magnetic moments are 1.0 (TiAs), 2.0 (VAs), 3.0 (CrAs), 4.0 (MnAs), 2.0 (CoAs) and 0.5 μ_B (NiAs) per supercell. The magnetism of Ni-doped arsenene weakens. The adsorption reduces the Ni-doped system's magnetic moment. As seen in Tables 2 and 3, the contribution of Ni atom and adjacent As atoms to the magnetism of NiAs decreases after adsorption. Other TM-doped systems' total magnetic moments hardly change at all; however, the TM dopant and As atoms' corresponding local magnetic moments decrease. The reduction of local magnetic moment is due to the spin transfer torque. The H₂S molecules' local magnetic moments are in the same direction as the TM atom. The TM 3d and adjacent As 4p orbitals mainly contribute to the magnetism (plotted in Fig. 11). Moreover, the asymmetrical structure of H₂S molecules' PDOS originates from the hybridization with TM dopant and As atoms.

Eventually, we calculate the charge density differences (CDD) (shown in Fig. 13) to understand the charge distribution during adsorption. H₂S molecules gain electrons in TiAs, VAs, CrAs and MnAs. The blue region around the TM atoms indicates the loss of electrons. In the case of Co and Ni-doped arsenene, the yellow region around the TM atoms suggests the accumulation of charge. The corresponding H₂S molecules lost charge. In addition, the region where electrons converge is concentrated between the H₂S molecules and dopant atoms. This aggregation of charge indicates the formation of a covalent bond. In conclusion, strong interactions between the H₂S molecules and the TM-doped systems conform to the strong hybridization of the orbitals and the enormous adsorption energy.

The introduction of TM dopants makes arsenene have a stronger affinity for H₂S molecules. The interactions between arsenene substrate and H₂S molecules can be reinforced by TM dopants. In all the doping cases considered, TM-doped arsenene can be employed to detect H₂S because of the charge transfer. The cationic of TM renders plural molecules attached to each adsorption site. Therefore, TM-doped arsenene has



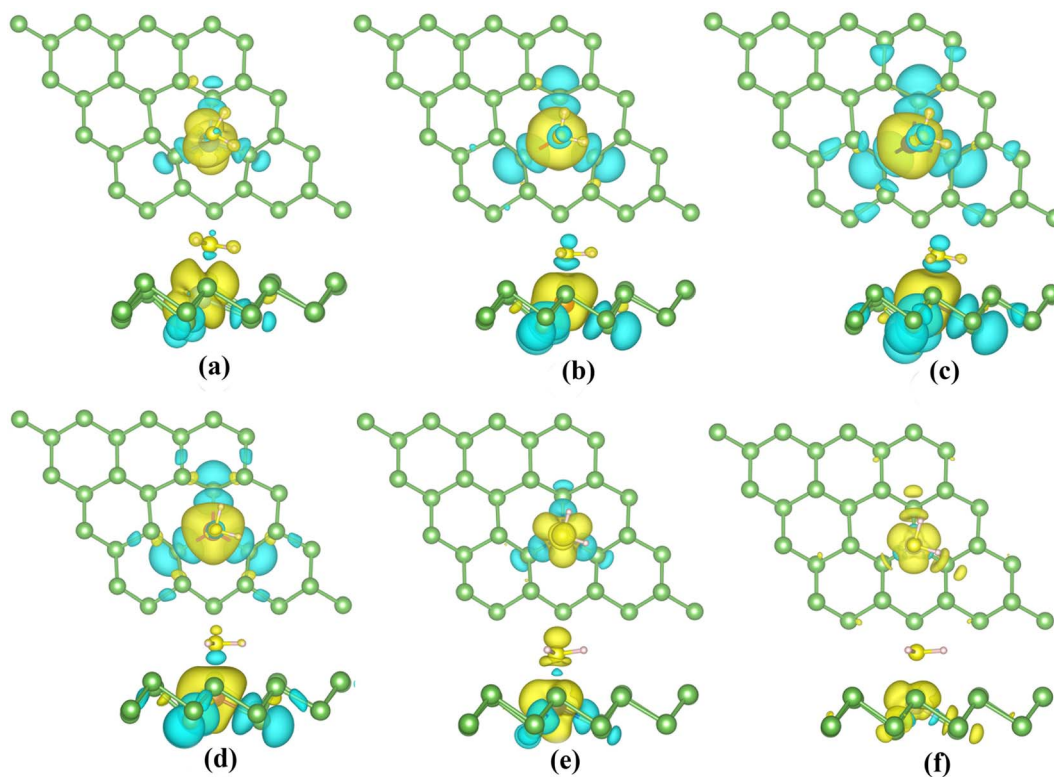


Fig. 12 (a)–(f) Top and side views of the spin densities plots of TiAs, VAs, CrAs, MnAs, CoAs and NiAs with H_2S adsorbed, where yellow colors represents spin-up states and blue colors represents spin-down states. The isosurface values are set to $0.001 \text{ e } \text{\AA}^{-3}$.

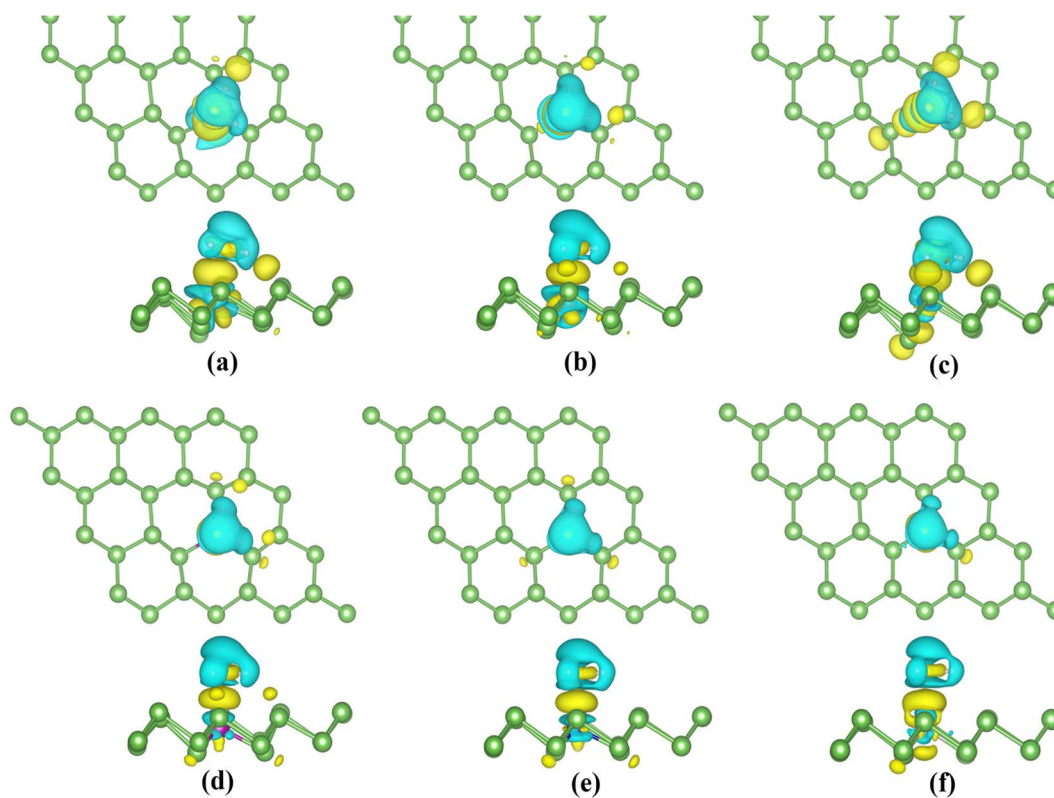


Fig. 13 (a)–(f) Top and side views of the charge density differences between H_2S molecules and TiAs, VAs, CrAs, MnAs, CoAs and NiAs. Yellow represents the accumulation of charge, and blue represents the diminishing of charge. The isosurface value is set as $0.001 \text{ e } \text{\AA}^{-3}$.



improved H₂S sensing performance significantly. By detecting electrical signals, TM-doped arsenene has a promising application in capturing H₂S.

4. Conclusion

In conclusion, we employ first-principles calculations to analyze the feasibility of defective arsenene (SW-As and SV-As) and TM-doped arsenene (TiAs, VAs, CrAs, MnAs, CoAs and NiAs) in detecting H₂S molecules. The interaction between arsenene substrate and H₂S can be enhanced by introducing SW- and SV-defects. However, the relatively small adsorption energies (−0.107 eV to −0.250 eV) and charge transfer suggest physisorption. After introducing dopants, the sensitivity is effectively improved because adsorption energies (−0.89 eV to −1.21 eV) and charge transfer have significantly increased. The H₂S molecules are chemically adsorbed on the TM-doped arsenene. The band gap change of Ti and Cr-doped system after adsorption is 0.1411 eV and 0.1685 eV. The change for the pristine and defective arsenene is less than 0.1 eV. After adsorption, Ti-doped arsenene changes from semiconductor to half-metal, and Ni-doped arsenene transforms from half-metal to conductor. We can observe remarkable electrical signals to detect H₂S gas in the adsorption process of TM-doped systems. Therefore, the TM dopants can increase the sensing performance of arsenene to H₂S gas more effectively than introducing point defects. Our results suggest that arsenene-based gas sensors have a good application prospect in the future.

Conflicts of interest

There are no conflicts to declare.

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