

CORRECTION

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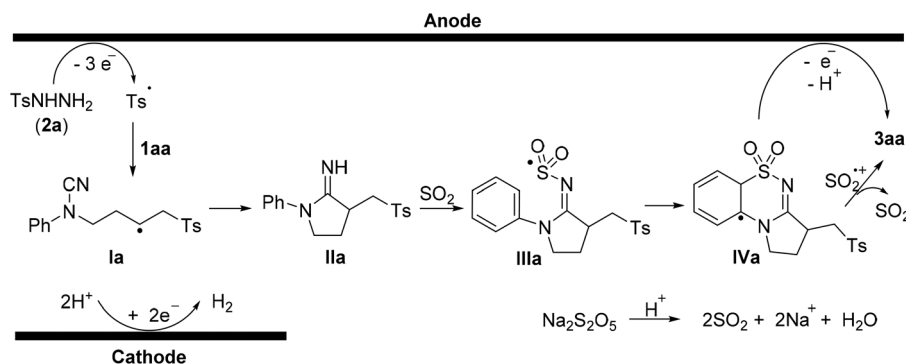
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Correction: Electrochemical tandem cyclization to access sulfonylated fused sultams *via* SO₂ insertion with sodium metabisulfiteYun-Hai Sun,^a Cheng-Jing Li,^a Ji-Ming Xi,^a Zhong-Lin Wei^a and Wei-Wei Liao^{*a,b}Correction for 'Electrochemical tandem cyclization to access sulfonylated fused sultams *via* SO₂ insertion with sodium metabisulfite' by Yun-Hai Sun *et al.*, *Org. Chem. Front.*, 2023, **10**, 705–711, <https://doi.org/10.1039/D2QO01821G>.

The authors regret that the proposed mechanism of anodic oxidation of Na₂S₂O₅ generating a sulfur dioxide radical cation was incorrect in Scheme 7 of the original article. In this work, the cyclic voltammetry (CV) experiments indicated that the solution of Na₂S₂O₅ in CH₃CN/H₂O displayed an obvious oxidative peak at 1.46 V (1.31 V in the presence of MsOH) (Fig. 1), while the mixture of Na₂S₂O₅ and **2a** displayed an obvious oxidative peak at 1.48 V under acidic conditions. Although the mechanism is vague at present, it is likely that the anodic oxidation of SO₂, which is generated from inorganic sulfite (Na₂S₂O₅), may generate a sulfur dioxide radical cation instead of directly removing one electron from the dianion S₂O₅²⁻, and the sulfur dioxide radical cation could subsequently oxidize **2a** to produce the corresponding arylsulfonyl radical as an alternative possible way. We are sorry that this led to some misunderstandings. The revised Scheme 7 is presented below in this correction notice, in which the anodic oxidation of Na₂S₂O₅ generating a sulfur dioxide radical cation was removed.



Scheme 7 Proposed mechanism.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

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