

## RESEARCH ARTICLE

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# A stable nanotubular metal–organic framework as heterogeneous catalyst for efficient chemical fixation of CO<sub>2</sub>†

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The incorporation of high-density and easily accessible Lewis acid sites is of great significance to obtain high catalytic activity in the CO<sub>2</sub> cycloaddition reaction of epoxides. Nanotubular metal–organic frameworks (NTMOFs) have attracted widespread attention as heterogeneous catalysts because of their outstanding porosity, intriguing structural diversity and accessible active sites. However, NTMOFs used as heterogeneous catalysts are limited, mainly due to the synthetic difficulties and lack of chemical and physical stability. Herein, we constructed novel NTMOFs with an interior channel diameter of 1.8 nm and an exterior wall diameter of 3.0 nm. They exhibited excellent chemical resistance to both acid and alkaline solutions. The NTMOFs feature durable catalytic performance for CO<sub>2</sub> cycloaddition with epoxides at atmospheric pressure with good recyclability. The turnover frequency (TOF) (306 h<sup>−1</sup>) value is greater than any previously reported value for metal–organic framework (MOF)-based catalysts for the cycloaddition of CO<sub>2</sub> to epoxides under similar conditions. Experimental results and theoretical calculations reveal that the abundant coordinatively unsaturated open metal sites within the mesoporous nanotubular channels facilitates the sufficient contact of the catalytic active sites with the epoxide substrates, thus enhancing catalytic activity. More importantly, the new materials could be extended to CO<sub>2</sub> fixation by the use of raw power plant flue gas.

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## Introduction

The chemical conversion of carbon dioxide (CO<sub>2</sub>) into valuable organic compounds has attracted great attention for the control of the enormous environmental impact of increasing carbon dioxide concentrations in the atmosphere.<sup>1</sup> This strategy enables the reuse of emitted carbon dioxide to produce a variety of high-value-added chemicals such as organic carbonates,<sup>2</sup> carboxylic acids,<sup>3</sup> carbon oxide,<sup>4</sup> alkyl amines,<sup>5</sup> alcohols,<sup>6</sup> *etc.* It is quite promising for cyclic carbonate production through carbon dioxide cycloaddition with an epoxide in the context of atomic economy and green chemistry because of the widespread application of such products in industry and the

absence of side product generation during this reaction process.<sup>7</sup> Therefore, catalysts with high activity and heterogeneity are needed for the sustainable application.

Currently, various heterogeneous catalysts have been synthesized and used to promote the cycloaddition reaction of CO<sub>2</sub> with epoxides, including porous carbon,<sup>8</sup> porous silica,<sup>9</sup> zeolites,<sup>10</sup> porous organic polymers (POPs),<sup>11</sup> *etc.* However, they often are operated under harsh reaction conditions (initial CO<sub>2</sub> pressure >1.0 MPa and/or reaction temperatures >80 °C). There are many reports of this reaction under mild conditions,<sup>12</sup> such as the polar-group-functionalized porous ionic polymer (PIP) POP-PBH,<sup>13</sup> which could promote the cycloaddition reaction of CO<sub>2</sub> with epichlorohydrin as an efficient bifunctional heterogeneous catalyst with 1.0 mol% catalyst loading at a relatively low temperature (60 °C) and ambient pressure under metal-free conditions during a reaction time of 48 h. Therefore, increasing efforts are being devoted to developing efficient catalysts for the solventless synthesis of cyclic carbonates to take place under milder reaction conditions, including lower initial CO<sub>2</sub> pressure, lower catalyst loading and shorter reaction time.

Crystalline nanotubular metal–organic frameworks (NTMOFs)<sup>14–19</sup> with the advantages of outstanding porosity, large specific surface area and accessible active sites have

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attracted extensive attention, and are widely applied in gas storage,<sup>20–22</sup> sensing,<sup>23–25</sup> ice channels,<sup>26</sup> magnetic applications,<sup>27</sup> small molecular adsorption,<sup>28–31</sup> encapsulation of guest molecules,<sup>32</sup> proton conductivity<sup>33</sup> and catalysis.<sup>34–36</sup> However, their use has been limited by the synthetic difficulties and lack of chemical and physical stabilities of the NTMOFs. At present, one example of a single-walled metal-organic nanotube based heterogeneous catalyst has been applied in CO<sub>2</sub> cycloaddition with a pressure of 1 MPa at 100 °C for 12 h.<sup>35</sup> In order to explore milder-reaction-condition catalytic materials based on NTMOFs, herein, we report the successful design of a nanotube-based metal-organic framework through the solvothermal reaction of propeller-like 4,4',4''-tricarboxyltriphenylamine (H<sub>3</sub>tca) and cadmium nitrate, which exhibit high chemical and physical stabilities. The NTMOFs have good carbon dioxide cycloaddition activity within a relatively short reaction time under atmospheric pressure CO<sub>2</sub>, which is attributed to the abundant coordination unsaturated cadmium sites located in the pores of the NTMOFs. The reusability of the synthesized NTMOFs and direct CO<sub>2</sub> fixation from raw power plant flue gas were also studied.

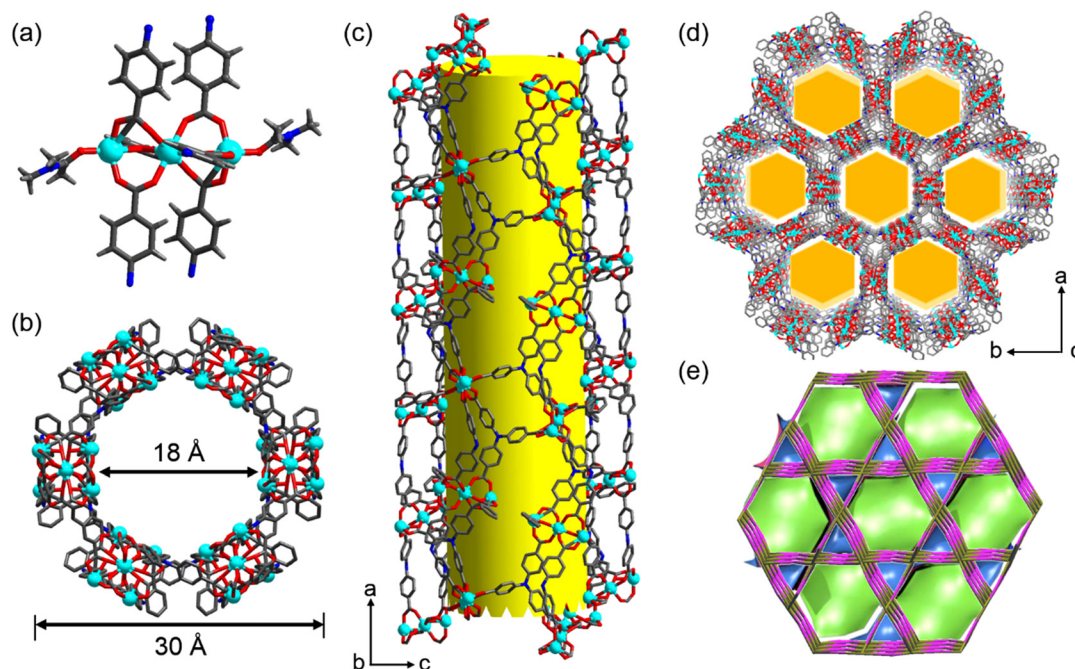
## Results and discussion

### Synthesis and characterization of NTMOF materials

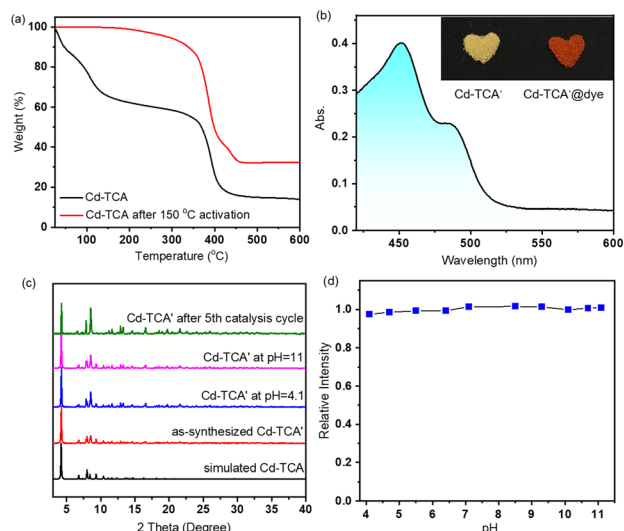
A solvothermal reaction of 4,4',4''-tricarboxyltriphenylamine (H<sub>3</sub>tca) with cadmium nitrate in water and *N,N*-dimethyl-

acetamide (DMA) led to a light-yellow block crystalline product, Cd-TCA. Single-crystal X-ray analysis reveals that Cd-TCA crystallizes in the centrosymmetric rhombohedral *R*3̄*c* space group, and it represents a new three-dimensional binodal 3,6-*c* topological network formed by three-fold {4<sup>2</sup>·6} and six-fold {4<sup>4</sup>·6<sup>2</sup>·8<sup>8</sup>·10} nodes (in Schläfli notation) (Fig. 1e).<sup>37</sup> The asymmetric unit consists of one tca<sup>3−</sup> ligand, one-and-a-half crystallographically independent Cd(II) cations and one coordinated DMA molecule (Fig. S1, ESI†). The inorganic secondary building unit is a trinuclear Cd cluster (Fig. 1a). The Cd1 atom adopts a perfect hexacoordinated octahedral geometry (Cd–O: 2.242–2.261 Å), is located at the centre of symmetry, and coordinates with four O atoms from four different tca<sup>3−</sup> ligands through four monodentate μ<sub>2</sub>–η<sup>1</sup>:η<sup>1</sup> modes and two bridging μ<sub>2</sub>–η<sup>2</sup>:η<sup>1</sup> modes (Fig. S2a, ESI†).

The two equivalent terminal Cd2 atoms in the SBU are five-coordinated with a triangular pyramidal environment; four of them come from the carboxylate oxygen atoms of tca<sup>3−</sup> ligands, and the fifth is the oxygen atom of a coordinating DMA solvent molecule. The Cd...Cd separation is 3.64(3) Å. The interphenyl dihedral angles for the tca<sup>3−</sup> ligand are 60.7(1)°, 68.3(7)° and 74.9(6)° with a μ<sub>6</sub>–η<sup>1</sup>:η<sup>1</sup>:η<sup>1</sup>:η<sup>1</sup>:η<sup>1</sup>:η<sup>2</sup> coordination mode (Fig. S2b, ESI†). The extended coordination of six such trinuclear cadmium SBUs by the carboxyl groups of the tca<sup>3−</sup> ligands results in the formation of “crown ether”-like hexagonal metallo-macrocycles, which are further connected infinitely with adjacent channels along the *ab* plane sharing the corrugated walls, forming NTMOFs having an outer diameter of 30 Å and an effective pore size of approximately 18 Å



**Fig. 1** (a) Cd coordination environment in Cd-TCA. (b) Each nanochannel is composed of trinuclear cadmium SBUs that are linked through tca<sup>3−</sup> ligands (H atoms are omitted for clarity). (c) Nanotubular structure of Cd-TCA showing the tube interior (yellow column) along the *b* direction. (d) Packing diagram of Cd-TCA with 1D opening channels along the *c* direction. (e) The 3,6-*c* topology for the NTMOF structure in Cd-TCA (the light-green and violet nodes in the network represent Cd(II) ions and the center of the tca<sup>3−</sup> ligands, respectively).



**Fig. 2** (a) The TG curve of Cd-TCA from room temperature to 600 °C. (b) UV-vis monitoring of 2',7'-dichlorofluorescein dye released from Cd-TCA'. Insert: optical photographs showing the solids of Cd-TCA' (left) and Cd-TCA' after 2',7'-dichlorofluorescein soaking treatment (right). (c) Powder XRD patterns: simulation from Cd-TCA single-crystal data, as-synthesized Cd-TCA', Cd-TCA' immersed in solutions with different pH values and Cd-TCA' after catalysis. (d) Fluorescence intensity of Cd-TCA' at different pH values in water.

(Fig. 1b–d and Fig. S3, ESI†). The topological analysis of Cd-TCA using ToposPro showed that the  $\text{Cd}_3\text{O}$  SBUs can be regarded as six connected nodes and the  $\text{tca}^{3-}$  ligands can be regarded as three connected nodes. The point symbol for the net of Cd-TCA is  $\{4^2 \cdot 6\}^2 \{4^4 \cdot 6^2 \cdot 8^8 \cdot 10\}$ ; it possesses a novel topology 3,6-c net with stoichiometry (3-c)2(6-c) (2-nodal net; Fig. 1e). The solvent cavity calculated using the program Platon is 57.4% of the total crystal volume. TGA of Cd-TCA shows that a weight loss of 40.5% was observed between 30 and 200 °C, corresponding to the loss of DMA molecules (expected: 39.6%) (Fig. 2a). These results are perfectly consistent with the elemental analysis data.

The activated compound Cd-TCA' was generated through the desolvation process by heating as-synthesized Cd-TCA at 150 °C under vacuum for 12 h, and the resulting PXRD pattern indicated that the host framework remained stable after removing the guest molecules (Fig. 2c). TGA and FT-IR also clearly indicated that the coordinated DMA molecules were completely excluded (Fig. 2a and S4 in ESI†). These coordinatively unsaturated open metal sites are conducive to interaction with guest molecules that enter the framework pores. The porosity of Cd-TCA' was tested using activated samples *via* dye-sorption measurements. Cd-TCA' samples were soaked in a methanol solution of 2',7'-dichlorofluorescein, and the filtered solids demonstrated obvious adsorption capacity for these dyes *via* UV-vis absorption spectrum monitoring. This experiment gave a quantum uptake equivalent to 6.6% of the Cd-TCA' weight (Fig. 2b), and the Cd-TCA' samples showed an obvious color change from yellow to orange after 2',7'-dichlorofluorescein soaking treatment (Fig. 2b, insert), all of which

indicated that the channels of the Cd-TCA' are accessible to large reagents typically used for catalysis, as the dye molecules have very large estimated cross-sections. The nanotubular channel structure in Cd-TCA' is conducive to the full exposure of the abundant coordinatively unsaturated open metal sites in the channels to facilitate sufficient contact of the catalytic active sites with the epoxide substrates to enhance catalytic activity, suggesting that Cd-TCA' should function as an active heterogeneous catalyst for the cycloaddition of carbon dioxide.

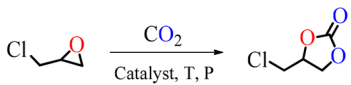
In order to explore the chemical resistance of Cd-TCA', Cd-TCA' samples were placed in a pH = 4.1 or pH = 11 aqueous solution for one day. The structures of the samples were evaluated using powder XRD (Fig. 2c). It is noted that Cd-TCA' was stable under these two tested conditions, and thus its structural and chemical integrity at different pH values can be confirmed. Additionally, investigation of the pH-dependent fluorescence response of Cd-TCA' showed that there was almost no change in its emission intensity in the pH range of 4.1–11.1 (Fig. 2d). Encouraged by considerably high thermal stability and chemical stability, the catalytic activities of the NTMOFs were further studied.

### Catalytic performance in the cycloaddition of $\text{CO}_2$ to epoxides

Our catalytic experiments focused on  $\text{CO}_2$  cycloaddition to epoxides, and the optimized catalyst dosage was 1.8  $\mu\text{mol}$  and reaction temperature was 60 °C under 1 atm  $\text{CO}_2$  (Fig. S5, ESI†). The initial catalytic activity was assessed with epichlorohydrin (3.3 mmol) as the substrate and  $\text{CO}_2$  (balloon pressure, 1 atm), using 1.8  $\mu\text{mol}$  (5.5 mmol%) Cd-TCA', in the presence of tetrabutylammonium bromide (TBABr, 0.3 mmol) as the co-catalyst at a temperature of 60 °C in a solvent-free environment. As can be seen from Table 2, an almost complete conversion can be achieved within 6 h over Cd-TCA' (entry 1 in Table 2). The yields of the cyclic carbonates from a small aliquot of the supernatant reaction mixture were calculated using  $^1\text{H}$  NMR. The turnover number (TON) per mole of Cd-TCA' for epichlorohydrin was about 1833, while the turnover frequency (TOF) per mole of Cd-TCA' per hour for epichlorohydrin was approximately 306. To the best of our knowledge, this TOF value is greater than any previously reported value for MOF-based catalysts for the cycloaddition of  $\text{CO}_2$  to epoxides under similar conditions (see Table S3 in ESI for comparison†).

Control experiments were further carried out to investigate the factors influencing this reaction in detail. No 4-(chloromethyl)-1,3-dioxolan-2-one product was generated in the absence of catalysts (Table 1, entry 1). Cd-TCA' or cadmium nitrate resulted in only 2% and 6% catalytic yield without a co-catalyst under the desired conditions (Table 1, entries 2 and 3). When TBABr alone was used as the catalyst, the activity was also negligible (Table 1, entry 4). Furthermore, the subcomponents of Cd-TCA' were then investigated independently. In the presence of 4,4',4''-tricarboxyltriphenylamine ( $\text{H}_3\text{tca}$ ) or cadmium nitrate instead of Cd-TCA', the conversion of epichlorohydrin was sharply reduced to 7% and 45%, respectively, under the same conditions (Table 1, entries 5 and 6). Thus, these results demonstrated that our Cd-TCA' was essential for  $\text{CO}_2$  cycloaddition reaction with epoxides.

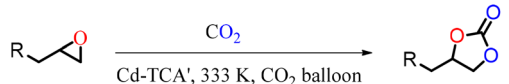










**Table 1** Control experiments of coupling of epichlorohydrin with CO<sub>2</sub><sup>a</sup>

			
Entry	Catalyst	Co-catalyst	Yield <sup>b</sup> (%)
1	None	None	n.d.
2	Cd-TCA'	None	2
3	Cd(NO <sub>3</sub> ) <sub>2</sub>	None	6
4	None	TBABr	n.d.
5	H <sub>3</sub> tca	TBABr	7
6	Cd(NO <sub>3</sub> ) <sub>2</sub>	TBABr	45

<sup>a</sup> Reaction conditions: epichlorohydrin (3.3 mmol), catalyst (1.8 μmol based on Cd<sub>1.5</sub>, 1.8 μmol based on H<sub>3</sub>tca, and 2.7 μmol based on Cd(NO<sub>3</sub>)<sub>2</sub>), and TBABr (0.3 mmol) with 1 atm CO<sub>2</sub> at 333 K for 6 h. <sup>b</sup> The yield of products was determined using <sup>1</sup>H NMR.

To assess the applicability of Cd-TCA' catalyst, CO<sub>2</sub> cycloaddition to different epoxides (Table S4, ESI†) was carried out under similar conditions. As shown in Table 2, the reaction yields catalyzed by Cd-TCA' from related aliphatic epoxides were 98% for epibromohydrin, 100% for glycidol, 93% for allyl glycidyl ether, 92% for butyl glycidyl ether and 97% for glycol diglycidyl ether with corresponding TOF values of 299, 306, 284, 281 and 296 per h per Cd<sub>1.5</sub> cluster, respectively (Table 2, entries 2–6). When the benzene-containing molecules 2-(phe-

**Table 2** Cd-TCA' catalyzed coupling of epoxides with CO<sub>2</sub><sup>a</sup>

				
Entry	Epoxides	Yield <sup>b</sup> (%)	TON <sup>c</sup>	TOF <sup>d</sup>
1		100	1833	306
2		98	1796	299
3		100	1833	306
4		93	1705	284
5		92	1687	281
6		97	1778	296
7		97	1778	296
8		78	1430	238
9		86	1577	263
10		72	1320	220

<sup>a</sup> Reaction conditions: Cd-TCA' (1.8 μmol, based on Cd<sub>1.5</sub>), epoxide (3.3 mmol), and TBABr (0.3 mmol) with 1 atm CO<sub>2</sub> at 333 K and 6.0 h. <sup>b</sup> The yield of products was determined by <sup>1</sup>H NMR. <sup>c</sup> Moles of product per mole of Cd-TCA'. <sup>d</sup> Moles of product per mole of Cd-TCA' per hour.

noxyethyl)oxirane, 2-((4-methoxyphenoxy)methyl)-oxirane, 2-[[4-(nitrophenoxy)methyl]oxirane and resorcinol diglycidyl ether were used as substrates, the epoxy carbonates were produced with yields reaching 97%, 78%, 86% and 72% and corresponding TOF values of 296, 238, 263 and 220 per h per Cd<sub>1.5</sub> cluster, respectively (Table 2, entries 7–10). Thus, all the epoxides exhibit high yields (>70%). These results demonstrated that Cd-TCA' displays excellent performance in this catalytic transformation of CO<sub>2</sub> to cyclic carbonates due to its large mesoporous diameter, which is beneficial for the substrates to sufficiently come into contact with the catalytic active sites. The recycling behavior of Cd-TCA' in the CO<sub>2</sub> fixation reaction of epichlorohydrin was studied in five consecutive runs. After each catalysis, Cd-TCA' was collected by centrifugation and washed with diethyl ether, dried under vacuum, and reused for the next catalytic run. There was almost no observable change in the catalytic activity of the Cd-TCA' catalyst; the yield of the product 4-(chloromethyl)-1,3-dioxolan-2-one remained at about 98% for five consecutive runs (Fig. 5a). Furthermore, the PXRD result exhibited barely any difference before and after the catalytic reaction (Fig. 2c), explicitly confirming the structural stability of the Cd-TCA' during catalysis, which was also further supported by the SEM images (Fig. S6, ESI†). In addition, the leaching of Cd in this reaction system as determined using inductively coupled plasma (ICP) analysis was also very low (<1%), which further verified the heterogeneity of Cd-TCA'. All of the above results suggest the favorable stability of Cd-TCA', which is very important for practical applications. Therefore, since Cd-TCA' has high stability and good recyclability, we carried out a gram-scale experiment for the fixation reaction of epichlorohydrin and CO<sub>2</sub> catalyzed by Cd-TCA'. The reaction was performed by adding 10 mmol of epichlorohydrin and extending the reaction time to 18 h; the other reaction conditions were not changed. It is worth noting that a yield of up to 97.1% can be achieved, giving a total of 1.33 g of product.

### Mechanism study

Previous experimental and theoretical studies have demonstrated that the ring-opening of the epoxide was the rate-determining step for the cycloaddition reaction.<sup>38,39</sup> Thus, the activation of epoxide by Cd-TCA' was first studied. The FT-IR spectra of Cd-TCA' after immersion in a methanol solution of epichlorohydrin displayed the C–H stretching vibration at the range from 2963 to 2860 cm<sup>−1</sup>, with an apparent red-shift from the range of 3066 to 2921 cm<sup>−1</sup> (free epichlorohydrin), indicating that the epoxide can be adsorbed and activated by Cd-TCA' (Fig. S7, ESI†).

Additionally, XPS spectra further illustrated the coordination environment of Cd in Cd-TCA' before and after immersion in an epichlorohydrin solution. The Cd 2d spectra suggested a positive shift for Cd-TCA' after immersion in epichlorohydrin (two peaks at 405.79 eV and 412.66 eV), compared with that for Cd-TCA' (405.58 eV and 412.38 eV) (Fig. 3a, insert). This result should be attributed to the decreased electron density around the Cd atoms, which most probably



derived from the weak bonding interaction between the  $\text{Cd}^{2+}$  center and epichlorohydrin.<sup>40</sup> A new peak of Cl 2p could be found in the XPS full spectrum of Cd-TCA' (Fig. 3a), which further confirmed that epichlorohydrin was adsorbed into the Cd-TCA'. Extensive molecular-force-field-based calculations were implemented for Cd-TCA' in order to monitor its affinity toward epichlorohydrin. Bonding interactions were observed between Cd centers and the O atom of epichlorohydrin, with the shortest interatomic separation being 2.53(7) Å (Fig. 3b). Additionally, SEM images revealed retention of the morphology without the formation of surface-deposited solids throughout the epichlorohydrin immersion process (Fig. 3c and d). Moreover, the EDX elemental mapping revealed that Cd and Cl are uniformly distributed in epichlorohydrin-treated Cd-TCA' (Fig. 3e). All the results supported the conclusion that Cd-TCA' could adsorb and activate epichlorohydrin.

Based on the experimental results and previous reports,<sup>7,35,41,42</sup> a Cd-TCA'-catalyzed  $\text{CO}_2$  cycloaddition process is proposed (Fig. 4). First, the epoxide binds to the Lewis acid sites in Cd-TCA' through the O atom of the epoxide, which leads to the activation of the epoxy ring. The carbon atom with less steric hindrance is attacked by  $\text{Br}^-$  that is generated from TBABr to open the epoxy ring. The oxygen anion on the opened epoxy ring then combines with carbon dioxide to form an alkyl carbonate anion. Finally, cyclization occurs *via* intramolecular nucleophilic substitution to generate the cyclic carbonate and regenerate the catalytic species to complete the catalytic cycle.

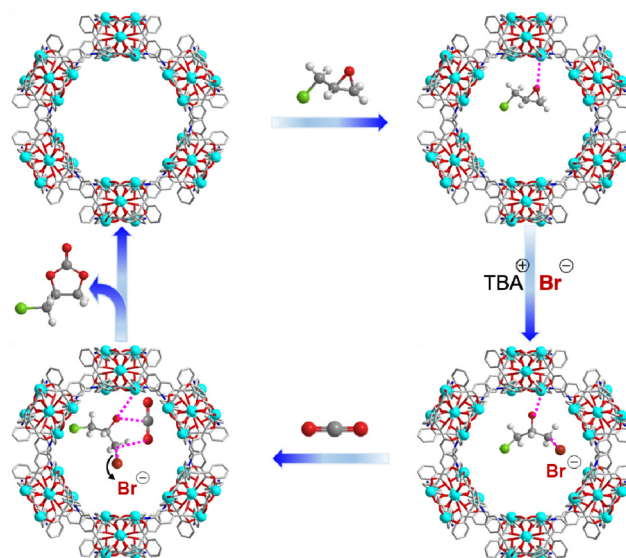


Fig. 4 Proposed mechanism of Cd-TCA'-catalyzed  $\text{CO}_2$  cycloaddition with epoxides.

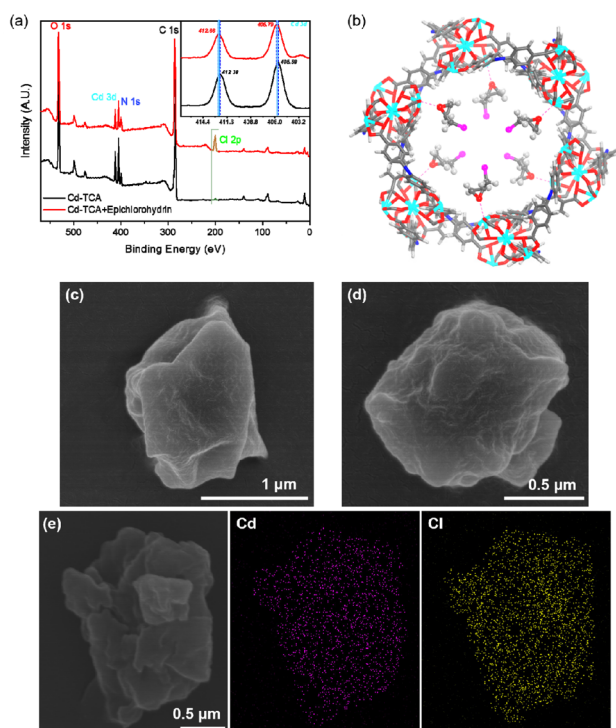
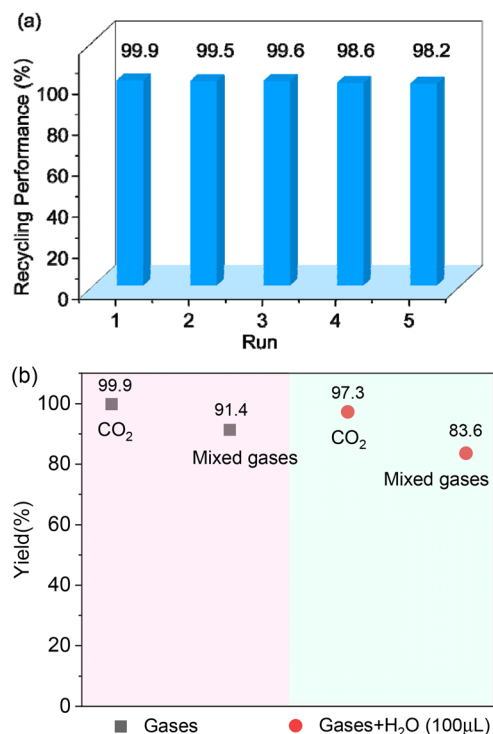


Fig. 3 (a) XPS spectra for Cd-TCA' and Cd-TCA' after treated with epichlorohydrin. Insert shows the Cd 3d region of Cd-TCA' and Cd-TCA' after immersion in an epichlorohydrin solution. (b) The calculated interatomic interaction between Cd-TCA' and epichlorohydrin. (c and d) SEM images of Cd-TCA' and epichlorohydrin-treated Cd-TCA'. (e) SEM/EDX images of epichlorohydrin-treated Cd-TCA'.

### $\text{CO}_2$ cycloaddition of epoxides using mixed gases

Coal-fired power stations are the main source of  $\text{CO}_2$  accumulation in the Earth's atmosphere. The exhaust gas emitted by coal-fired power stations is a mixture of about 80%  $\text{N}_2$ , 14%  $\text{CO}_2$  and 6%  $\text{O}_2$ .<sup>43</sup> We decided to further explore the application of Cd-TCA' in  $\text{CO}_2$  fixation in simulated samples by simulating the above mixed gas. Under the same experimental conditions, except that mixed gases (1.0 MPa, ensuring that the  $\text{CO}_2$  pressure is about 1 atm) were used instead of high-purity  $\text{CO}_2$  gas, the yield of the product 4-(chloromethyl)-1,3-dioxolan-2-one was 91.4% (Fig. 5b). This efficiency is almost consistent with that using pure  $\text{CO}_2$  gas, indicating that the activity of the catalyst in  $\text{CO}_2$  chemical fixation was not decreased due to the use of the mixed gases without further separation and purification. In addition, the typical components of coal-fired flue gas are 76.0–77.0%  $\text{N}_2$ , 12.5%–12.8%  $\text{CO}_2$ , 4.4%  $\text{O}_2$ , 6.2%  $\text{H}_2\text{O}$ , 420 ppm  $\text{NO}_x$ , 50 ppm CO and 420 ppm  $\text{SO}_2$ .<sup>44,45</sup> In terms of studying the catalytic conversion application of carbon dioxide in actual samples, the influence of  $\text{N}_2$  and trace gases ( $\text{SO}_2$ ,  $\text{NO}_x$ , CO) can be ignored, and  $\text{O}_2$  and  $\text{H}_2\text{O}$  may have the greatest impact on the reaction. In order to promote the practical application of this material in industry, the catalytic conversion of  $\text{CO}_2$  using high-purity  $\text{CO}_2$  and mixed gases in the presence of moisture (100  $\mu\text{L}$ ) was also explored. The yield of the product 4-(chloromethyl)-1,3-dioxolan-2-one was 97.3% and 83.6% using high-purity  $\text{CO}_2$  and mixed gas, respectively, demonstrating that the catalyst activity was not affected by moisture, and Cd-TCA' has clear feasibility for practical application in the chemical industry.



**Fig. 5** (a) Recycling experiment of Cd-TCA' for CO<sub>2</sub> cycloaddition with epichlorohydrin. (b) CO<sub>2</sub> cycloaddition of epoxides using mixed gases and mixed wet gases, respectively.

## Conclusions

In conclusion, we have successfully prepared a new nanotubular metal-organic framework material with the propeller-like triphenylamine moiety as the backbone, and it possesses a new 3,6-conn topology and can act as a heterogeneous catalyst for CO<sub>2</sub> cycloaddition with various epoxides. It features strong stability and superior catalytic performance with the highest activity of a turnover frequency reaching 1500 per mole of catalyst per hour after five repeated reactions using a CO<sub>2</sub> balloon. The high-density and easily accessible Lewis acid sites within the mesoporous nanotubular channels facilitate sufficient contact of the catalytic active sites with the substrates, thus enhancing catalytic activity. It is worth noting that this material can also be applied to simulate the fixation of actual CO<sub>2</sub> gas in raw power station flue gas. This work provides an efficient approach to use the nanotubes for the CO<sub>2</sub> cycloaddition reaction with cyclic carbonates under mild conditions.

## Experimental

### Synthesis of Cd-TCA

A mixture of H<sub>2</sub>O/*N,N*-dimethylacetamide (10 mL) containing 4,4',4''-tricarboxyltriphenylamine (H<sub>3</sub>tca) (23.0 mg) and Cd (NO<sub>3</sub>)<sub>2</sub>·4H<sub>2</sub>O (88.0 mg) was placed in an oven and heated at 110 °C for 2 days; when cooled, light-yellow bulk crystals were produced and collected. Yield: 76%. Anal calc. for [Cd<sub>1.5</sub>(tca)

(DMA)]·3DMA (C<sub>37</sub>H<sub>48</sub>N<sub>5</sub>O<sub>10</sub>Cd<sub>1.5</sub>): C 49.85, H 5.43, N 7.86%; found: C 49.72, H 5.38, N 7.77%.

### Typical procedure for the CO<sub>2</sub> cycloaddition reaction with epoxides

In a 10 mL Schlenk tube, the catalytic reaction was performed with a steady reaction tube pressure controlled by a rubber bladder containing CO<sub>2</sub> (0.1 MPa). The catalyst Cd-TCA' (1.8 μmol, 5.5 mmol%, based on Cd<sub>1.5</sub>), epoxide (3.3 mmol), and the cocatalyst tetra-*n*-tert-butylammonium bromide (TBAB, 0.3 mmol) were added, and the reaction mixture was stirred at 333 K for 6.0 h. When the reaction was completed, the supernatant solution obtained after the catalysts were centrifuged was used for <sup>1</sup>H NMR analysis, and the reaction yields were calculated.

## Conflicts of interest

There are no conflicts to declare.

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