

PAPER

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[View Journal](#) | [View Issue](#)Cite this: *Nanoscale Adv.*, 2023, 5, 2639Molecular engineering on a MoS₂ interlayer for high-capacity and rapid-charging aqueous ion batteries†Xuefei Han,^{abc} Jing Yang,^{ID a} Yong-Wei Zhang^{ID *ab} and Zhi Gen Yu^{ID *ab}

Rechargeable aqueous ion batteries (AIBs) play essential roles in the increasing demand for high-performance energy storage systems, and yet they are hampered by the lack of suitable cathode materials because of the sluggish intercalation kinetics. In this work, we develop an effective and feasible strategy to enhance the performance of AIBs by broadening the interlayer spacing by using intercalated CO₂ molecules to promote the intercalation kinetics by using first principles simulations. Compared with pristine MoS₂, the intercalation of CO₂ molecules with a 3/4 ML coverage significantly increases the interlayer spacing to 9.383 Å from 6.369 Å and the diffusivity is boosted by 12 orders of magnitude for Zn ions, 13 orders for Mg ions and one order for Li ions. Moreover, the concentrations of intercalating Zn, Mg and Li ions are enhanced by 7, 1 and 5 orders of magnitude, respectively. The significantly increased diffusivity and intercalation concentration of metal ions signify that intercalating CO₂ bilayer MoS₂ is a promising cathode material to realize metal ion batteries with a rapid charging capability and high storage capacity. The strategy developed in this work can be generally applied to increase the metal ion storage capacity in transition metal dichalcogenide (TMD)- and other layered material-based cathodes and make them promising for next-generation rapidly rechargeable batteries.

Received 30th January 2023
Accepted 1st April 2023DOI: 10.1039/d3na00068k
rsc.li/nanoscale-advances

Introduction

Renewable and clean energy generation is a promising solution to counter the carbon dioxide emission mainly originating from burning fossil fuels. Several renewable energy harvesting technologies have been well developed.^{1–5} The rapid development and high demands of renewable and clean energy sources, portable electronic devices, and electric vehicles have triggered great ambition for low cost, large-scale, and high energy density battery systems for energy storage. Due to the high cost and safety hazards and the scarce source of metal Li, it is desirable to find alternative energy storage systems to replace Li-ion batteries (LIBs). In this regard, aqueous multivalent metal ion batteries (AMMIBs) are attracting tremendous attention and are considered promising substitutes for LIBs.^{6–13} Due to high safety, low cost, eco-friendliness, and high ionic conductivity (1000 times higher than organic electrolytes), rechargeable AMMIBs are promising batteries for grid-scale electrochemical

energy storage. Among rechargeable AMMIBs, aqueous zinc-ion batteries (ZIBs)^{14–17} and aqueous magnesium-ion batteries (MIBs)^{18–20} have attracted remarkable attention worldwide because they exhibit a high volumetric energy density of 5851 mA h mL^{−1} for ZIBs^{21,22} and 3833 mA h mL^{−1} for MIBs.^{23–25} However, the main challenge lies in developing suitable cathode materials for AMMIBs.

Several materials have been developed as promising cathode materials for AMMIBs. Vanadium oxides with a tunnel structure as cathode materials show high Zn-ion storage properties,^{26–31} but the dissolution of vanadium in water-based electrolytes remains a significant challenge.³² A composite of manganese dioxide and carbon molecular sieves (δ-MnO₂@CMS) with a core-shell structure and Chevrel phase Mo₆S₈ were reported as cathode materials for MIBs.^{33–35} The practicality of MIBs is hampered by the absence of suitable high-performance cathode materials with rapid Mg ion diffusion.³⁶ Transition metal dichalcogenides (TMDs) have attracted extensive attention due to their potential applications in cathode materials for rechargeable AMMIBs. The layer structure of TMDs and their weak interlayer interaction *via* van der Waals (vdW) force are appealing properties for multivalent metal ion diffusion and intercalation,^{37–41} especially for large metal ion-based AMMIBs. Among TMDs, MoS₂ is considered one of the most promising cathode materials for AMMIBs,^{42,43} which suggests the feasibility of employing MoS₂ as a functional AMMIB cathode material. This scarcity in the report suggests the heightened

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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3na00068k>

challenges in observing reversible metal ion storage in pristine MoS₂. The high intercalation energy barrier of metal ions leads to the low specific capacities recorded for the MoS₂ systems (for Zn 1–40 mA h g^{−1}).^{44–46} For example, it was found that Zn²⁺ diffusion across the MoS₂ framework was further hampered by its low electrochemical activity and low conductivity.⁴⁵ Hence, it is essential to explore effective modification methods to “activate” MoS₂ towards reversible metal ion storage for leveraging the advantages of MoS₂. Phase engineering was a highly appealing strategy for modulating the chemical and electrical properties. 1T-MoS₂, with a tetragonal symmetry, in which each Mo atom has an octahedral coordination with S atoms, has metallic conductivity, resulting in a lower metal ion diffusion barrier than in 2H-MoS₂. However, 1T MoS₂ is a metastable phase, and producing 1T MoS₂ on a large scale remains a big challenge. An alternative strategy is to tune the intercalation energy by changing the interlayer spacing. It was demonstrated that intercalation oxygen could increase their interlayer spacing (9.5 Å) and tune hydrophilicity, resulting in boosting the Zn ion diffusion kinetics by 3 orders of magnitude in MoS₂.⁴⁴ Sandwiched structures consisting of monolayer MoS₂ and carbon (MoS₂:C)⁴⁷ and MoS₂/graphene⁴⁸ were reported to realize high-performance sodium ion and Zn ion batteries thanks to the expanded interlayer spacing (11.6 Å). Although many efforts have been made to study layered MoS₂, the high capacity and long life-cycle times of MoS₂-based cathodes have not yet been coexisting for practical utilization. More work needs to be conducted to expand the interlayer spacing further and enhance the hydrophilicity of MoS₂ to realize high reversible capacity and superb durability, which remains a big challenge.

In this study, we demonstrate a feasible and effective strategy to reduce metal ion diffusion barrier by using intercalation of CO₂ molecules to expand the MoS₂ interlayer spacing through density functional theory (DFT) simulations. Our comprehensive DFT results reveal that the intercalated MoS₂ by CO₂ is a promising cathode material for realizing rapidly chargeable metal ion batteries.

Results and discussion

AA' stacking bilayer MoS₂ was chosen to investigate the metal ion diffusion in this study since it has the lowest relative formation energy among five possible stacking configurations.⁴⁹ The optimized unit cell of AA' stacking bilayer MoS₂ is shown in Fig. S1a.† The optimized lattice constants are $a = b = 3.204$ Å and the layer spacing is 6.37 Å, which is quite close to the reported value of 6.21 Å.⁴⁹ The stacking formation E_s of bilayer MoS₂ was calculated based on the definition of the total energy difference per atom between the bilayer and two individual monolayers $E_s = \frac{E_{\text{bilayer}} - 2E_{\text{monolayer}}}{N}$, where N is the total number of atoms in bilayer MoS₂ unit cell (4 S and 2 Mo atoms). The calculated AA' stacking formation energy is −34.33 meV per atom, well agreeing with the reported value of −25.13 meV.⁴⁹ The tiny difference may originate from the van der Waals (vdW) correction methods.⁵⁰ The optimized unit cell was expanded to build $5 \times 5 \times 1$ supercells containing 50 Mo and 100 S atoms

shown in Fig. S1b,† and rectangle supercells ($a = 22.201$ Å and $b = 19.224$ Å) with 96 Mo and 192 S atoms shown in Fig. S1b.†

It was well explored that the interlayer spacing and coupling have a strong effect on the intercalation energy of metal ions.^{44,47,48} In this study, the intercalating CO₂ molecules were used to expand the interlayer spacing of bilayer MoS₂. First, the possible embedding configurations of the intercalating CO₂ molecules were investigated, and the optimized configurations and the calculated relative energies are shown in Fig. S2.† In which, the CO₂ molecule has two possible embedding sites, and the computed relative energies reveal that CO₂ prefers to stay at the bridge site rather than the hollow site. The intercalation energy of an intercalating CO₂ molecule E_{In} was calculated based on $E_{\text{In}} = E_{(\text{MoS}_2 + \text{CO}_2)} - (E_{\text{MoS}_2} + E_{\text{CO}_2})$, where, $E_{(\text{MoS}_2 + \text{CO}_2)}$, E_{MoS_2} and E_{CO_2} are the total energies of bilayer MoS₂ with one embedding CO₂ molecule, pristine bilayer MoS₂ and an isolated CO₂ molecule. For an isolated CO₂ molecule energy calculation, we put one CO₂ molecule in a cube model ($a = b = c = 15$ Å), and the calculated ground state energy is considered the energy of an isolated CO₂ molecule. The calculated intercalation energy of one embedding CO₂ molecule is 2.98 eV at the bridge site. With one intercalating CO₂ molecule, the interlayer spacing slightly increases to 7.61 Å from 6.37 Å (pristine bilayer MoS₂). The relatively high intercalation energy of the CO₂ molecule may originate from the strong interlayer coupling. Therefore, more CO₂ molecules may need to be intercalated to expand the interlayer spacing and reduce the interlayer coupling further. Meanwhile, we also investigate the intercalating CO₂ molecule diffusion in bilayer MoS₂. According to the symmetry of the 2H MoS₂ structure, the intercalating CO₂ has three possible diffusion pathways: armchair, zigzag and crossing, as shown in Fig. 1a. The calculated diffusion barriers are 0.209, 0.209 and 0.394 eV as shown in Fig. 1b. The computed results reveal that the CO₂ diffusion barriers along the zigzag and armchair directions are identical, revealing an isotropic diffusivity of the intercalating CO₂ in the bilayer MoS₂. Also, the direct crossing diffusion is more difficult due to its higher diffusion barrier than zigzag and armchair directions originating from the high relative energy at the hollow site. The calculated diffusion barriers of 0.209 and 0.394 eV also show that the intercalation CO₂ would be kinetically stable in the bilayer MoS₂ since the room temperature is about 0.0256 eV. Temperature and pressure contributions are not included in the results obtained from DFT simulations, which only are valid at $T = 0$ K and $P = 0$ atm. The results from DFT simulations can be used as an input to thermodynamics considerations to describe a situation of finite temperature and pressure. To investigate the thermodynamic stability of the CO₂ intercalated MoS₂, we performed DFT simulations considering appropriate thermodynamic functions, and the CO₂ intercalated stacking Gibbs free energy as a function of the temperature and pressure can be calculated

$$\Delta G(T, P) = E_s + \Delta E_{\text{ZPE}} - T\Delta S_{\text{vib}} + PV \quad (1)$$

where E_s is the normalized CO₂ intercalated stacking energy directly from DFT simulations, defined as



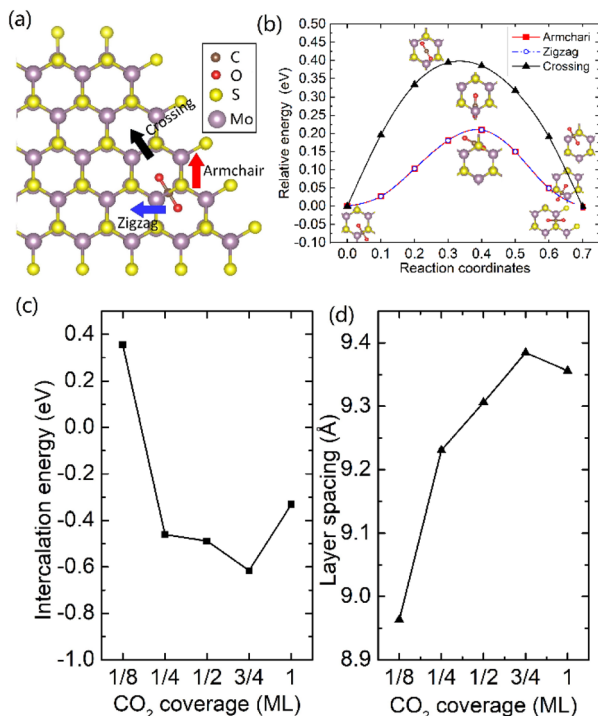


Fig. 1 (a) Three possible diffusion pathways of the intercalation of CO₂ molecule. (b) The diffusion barrier of CO₂ in bilayer MoS₂. The CO₂ molecule coverage-dependent intercalation energies (c) and layer spacing (d).

$$E_S = (E_{(\text{MoS}_2+n\text{CO}_2)} - (E_{(\text{MoS}_2)} + nE_{\text{CO}_2}))/n \quad (2)$$

where $E_{(\text{MoS}_2+n\text{CO}_2)}$, $E_{(\text{MoS}_2)}$ and E_{CO_2} are the energies of monolayer or bilayer MoS₂ with $n\text{CO}_2$ molecules, a clean monolayer, or bilayer MoS₂ and the CO₂ molecule, respectively. E_{ZPE} and S_{vib} are the zero point energy and the entropy contributed by vibration frequency ω_i and $T = 298$ K. E_{ZPE} can be calculated using $E_{\text{ZPE}} = \sum_i \frac{\hbar\omega_i}{2}$, and S_{vib} can be calculated using

$$S_{\text{vib}} = \sum_i \left[K_b T \ln \left(1 - \exp \left(-\frac{\hbar\omega_i}{K_b T} \right) \right) \right], \text{ where } \hbar \text{ is the reduced}$$

Planck's constant and ω_i is the i -th vibrational frequency of the diffusing ion. As a benchmark, we computed the entropy CO₂ molecule and the calculated entropy is 214.419 J (mol⁻¹ K⁻¹), which is comparable with the standard molar entropy of 213.79 J (mol⁻¹ K⁻¹). In this study, we proposed two pathways to intercalate CO₂ into MoS₂, directly intercalating CO₂ molecules into one bilayer MoS₂ (P1) or two monolayers (P2) one by one as shown in Fig. S6a.† Under the condition of $P = 0$ atm, we calculated the CO₂ intercalated stacking Gibbs free energies, and the results are shown in Fig. S6b.† In comparison, using two monolayers (P2) may easily realize the CO₂ intercalated MoS₂. As shown in Fig. S6b,† the stacking Gibbs free energy of one CO₂ intercalated MoS₂ is 3.48 eV using the P1 pathway and dramatically decreases to -3.50 eV. When the CO₂ coverage increases to 3/4 and 1 ML, the stacking Gibbs free energies of the CO₂ intercalated MoS₂ are 0.36 and 0.26 eV using the P1 pathway, much higher than those of -0.16 eV and -0.13 eV

using the P2 pathway, respectively. Based on the definition of the stacking Gibbs free energy, the negative value denotes that the final product of the CO₂ intercalated MoS₂ is thermodynamically stable, and the reaction is exothermic. It should be noted that the calculated stacking Gibbs free energies of the CO₂ intercalated MoS₂ are based on $P = 0$ atm. Based on eqn (1), we may estimate the pressure needed to realize 3/4 ML CO₂ intercalation when the Gibbs free energy is zero using the P1 process. The estimated pressure is 107 atm (0.0108 GPa) to stabilize CO₂ molecule (3/4 ML) intercalating in bilayer MoS₂. Furthermore, intercalating one CO₂ molecule may require higher pressure ($P = 961$ atm). Based on the results shown in Fig. S6b,† it can be seen that the intercalated structure appears unstable at a low coverage and becomes more stable at a high coverage, which reveals a substantial kinetic barrier in realizing the CO₂ molecules intercalated structure. Not surprisingly, the calculated CO₂ diffusion barriers of 0.209 and 0.394 eV shown in Fig. 1b can be considered the kinetic barrier for realizing a stable structure at a high coverage. The diffusion barriers also reveal that the intercalating CO₂ would be kinetically stable in the bilayer MoS₂ since the diffusion barriers are much higher than 0.0256 eV (room temperature). Also, we performed *ab initio* molecular dynamics (AIMD) for one CO₂ molecule intercalated MoS₂ and bilayer MoS₂ with a CO₂ coverage of 1 ML at 300 K within 1000 fs. Fig. S7† shows intercalated CO₂ molecules stabilized in the bilayer MoS₂, maintaining the 2H structural symmetry. The optimized models obtained from AIMD simulations also confirm the thermodynamical stability. Therefore, the CO₂ intercalated MoS₂ is thermodynamically stable and feasible. The diffusion barrier also indicates that intercalating CO₂ molecules can easily form a uniform distribution in bilayer MoS₂. The relative energy results also reveal that intercalating CO₂ molecules prefer to separate rather than cluster in the bilayer MoS₂.

To expand the interlayer spacing and reduce the interlayer coupling further, more CO₂ molecules were intercalated into the bilayer MoS₂. The CO₂ coverage-dependent intercalation energy E_f was computed based on the definition of $E_f = E_{(\text{MoS}_2+n\text{CO}_2)} - E_{(\text{MoS}_2+(n-1)\text{CO}_2)} - E_{\text{CO}_2}$, where $E_{(\text{MoS}_2+n\text{CO}_2)}$, $E_{(\text{MoS}_2+(n-1)\text{CO}_2)}$ and E_{CO_2} are the energies of $n\text{CO}_2$ intercalating into MoS₂, $(n-1)\text{CO}_2$ intercalating into MoS₂ and an isolated CO₂ molecule. In this study, we only consider the CO₂ coverage of 1/8 ML, 1/4 ML, 1/2 ML, 3/4 ML and 1 ML (24 CO₂ molecules) using big rectangle supercells composed of 96 Mo and 192 S atoms, and the optimized models are shown in Fig. S5.† The calculated CO₂ molecule coverage-dependent intercalation energies and the corresponding layer spacings are shown in Fig. 1c and d. The simulation results show that the intercalation energy decreases with the increase in coverage, and the 3/4 ML coverage of the CO₂ molecules results in the lowest intercalation energy of -0.615 eV, and the intercalation energy is increased when the CO₂ coverage further increases. Not surprisingly, the interlayer spacing increases with the increase in coverage, resulting in the largest layer spacing of 9.384 Å among five considered coverages. A further increase in coverage (1 ML) makes the interlayer spacing decrease when the CO₂ coverage is higher than 3/4 ML. Based on the results shown in Fig. S5,† the



CO₂ intercalated MoS₂ keeps the structural stability with increased coverage even at the coverage of 1 ML. The calculated AA' stacking formation energy of bilayer MoS₂ with a CO₂ molecule coverage of 3/4 ML increases to −28.01 meV per atom, which is higher than that in pristine MoS₂ of −34.33 meV per atom. A comparison of these results reveals that the intercalation of CO₂ weakens the layer coupling. The calculated interlayer spacing of the bilayer MoS₂ with 3/4 ML coverage is 9.384 Å, slightly higher than that with a 1 ML coverage of 9.356 Å. It should be noted that the interlayer spacing is the average distance between two Mo atoms in the different layers. Based on the results shown in Fig. 1d, we may consider that the expanded interlayer spacing is converged when the CO₂ coverage is more than 3/4 ML. Therefore, we adopted the optimized bilayer MoS₂ models with a CO₂ coverage of 3/4 ML to investigate the diffusion of metal ions.

In this study, we investigated the diffusion barriers of three metal ions (Zn, Mg, and Li) in bilayer MoS₂. Based on the structure and 2H symmetry of the bilayer MoS₂, we proposed two possible embedding sites (tetrahedral *T_h* and octahedral *O_h*), as shown in Fig. S3†. Considering the Zn ion as a representative, we calculated the relative energies of two possible embedding sites, and the results are shown in Fig. S4† and proposed the possible diffusion pathway of metal ions as (b) → (c) → (d), as shown in Fig. S4.† The diffusion barriers of Zn, Mg and Li ions in the pristine bilayer MoS₂ were calculated, and the results are shown in Fig. 2a. The calculated diffusion barriers of Zn, Mg and Li ions in the pristine bilayer MoS₂ are 0.785, 0.942 and 0.356 eV, respectively. The temperature-dependent ion diffusivity of the three metal ions was predicted based on the computed diffusion barriers considering thermal corrections (ESI†), and the results are shown in Fig. 2b. The calculated ion diffusivities are $4.18 \times 10^{-16} \text{ cm}^2 \text{ s}^{-1}$, $1.01 \times 10^{-16} \text{ cm}^2 \text{ s}^{-1}$ and $6.90 \times 10^{-9} \text{ cm}^2 \text{ s}^{-1}$ at room temperature, corresponding to the Zn, Mg, and Li ions, respectively. These calculation results indicate that the pristine bilayer MoS₂ is a cathode material with limited performance for metal ion batteries. Clearly, new

strategies are highly demanded to reduce the ion diffusion barrier for high-performance batteries.

We further demonstrated a feasible and effective strategy to reduce the ion diffusion barrier by using intercalating CO₂ molecules to expand the MoS₂ interlayer spacing. With a high coverage of 3/4 ML CO₂, the interlayer spacing of MoS₂ is expanded to 9.383 Å. Compared with the interlayer spacing of 6.369 Å obtained from the pristine bilayer MoS₂, a bigger interlayer spacing results in a weak interaction between layers and lower diffusion barriers of metal Li, Zn and Mg ions. With the intercalating CO₂ molecules, the results shown in Fig. 2c and d reveal that the diffusion barriers of Zn and Mg were dramatically reduced to 0.026 and 0.064 eV, significantly boosting the Zn and Mg ion diffusion by 12 and 14 orders of magnitude at room temperature, while only by one order of magnitude for Li ions. Our comprehensive DFT results reveal that the CO₂ intercalated MoS₂ is a promising cathode material for realizing rapidly chargeable Zn and Mg metal ion batteries. Compared with previously reported results shown in Table 1, the calculated diffusivity of Zn in the bilayer MoS₂ is much smaller than the experimental one in the bulk MoS₂,⁵¹ and Mg has a higher diffusivity in the bilayer MoS₂ than in the bulk MoS₂. As for Li ions, they have nearly the same diffusivity in the bilayer MoS₂ as in the bulk MoS₂. It should be noted that the diffusivities of three ions are very low, showing that the pristine bilayer and bulk MoS₂ are not suitable for metal ion batteries. Comparing the experimental values shown in Table 1 and our theoretical values reveal that the CO₂ intercalated MoS₂ has significant potential for metal ion batteries. However, Zn and Mg ions have much higher diffusivity in the CO₂ intercalated MoS₂ than in O-modified⁵¹ and MoS₂/graphene heterojunctions.⁵² Not surprisingly, no significant improvement in Li ion diffusivity was found in our study compared with W and Mo alloyed MoS₂.⁵³

Beyond the metal ion diffusion barrier and diffusivity, another challenge is to improve the storage capacity of AMMIBs, mainly determined by the embedding concentration of ions. Meanwhile the ion concentration or solubility in cathode materials is governed by the intercalation energy of metal ions.⁴⁹ Consequently, we calculated the intercalation energy of embedded metal ions in the pristine and the intercalating CO₂ bilayer MoS₂ (ESI†). The results are shown in Fig. 3a reveal that the intercalation energy of Zn ion significantly decreases to 0.411 eV from 1.279 eV in the pristine bilayer MoS₂, and the intercalation energy of Li dramatically decreases to 0.204 eV from 0.898 eV. However, the intercalation energy of Mg slightly decreases to 0.463 eV from 0.613 eV. Based on the calculated intercalation energies of metal ions in the pristine and CO₂

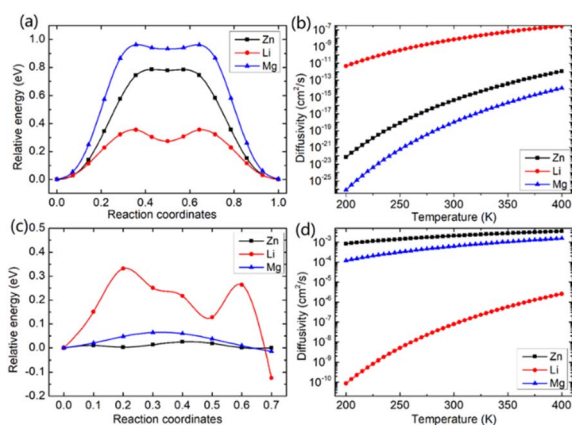


Fig. 2 The computed diffusion barriers (a) and (c) and diffusivities (b) and (d) of metal Zn, Li, and Mg ions in pristine bilayer MoS₂ and in bilayer MoS₂ with an intercalating CO₂ coverage of 3/4 ML, respectively.

Table 1 The comparison between the computed ion diffusivity ($\text{cm}^2 \text{ s}^{-1}$) with reported experimental values shown in brackets

Ions	Bilayer MoS ₂	Bulk MoS ₂	CO ₂ -MoS ₂	Modified MoS ₂
Zn	4.18×10^{-16}	(8×10^{-12})	2.08×10^{-3}	$(9 \times 10^{-8} \text{ to } 10^{-9})^{51}$
Mg	1.01×10^{-18}	(1×10^{-20})	6.26×10^{-4}	$(3.24 \times 10^{-9})^{52}$
Li	6.90×10^{-9}	(3.78×10^{-9})	7.84×10^{-8}	$(1.2 \times 10^{-8})^{53}$



embedded bilayer MoS₂, we computed the temperature-dependent intercalation ion concentration (ESI[†]), and the results are shown in Fig. 3b. The Zn ion embedding concentration is enhanced by 7 orders of magnitude to $3.63 \times 10^{11} \text{ cm}^{-2}$, and the Li ion embedding concentration is enhanced by 5 orders of magnitude to $1.98 \times 10^{13} \text{ cm}^{-2}$ at 300 K. However, the Mg ion embedding concentration is slightly boosted by just one order of magnitude to $1.25 \times 10^{11} \text{ cm}^{-2}$. In contrast, the intercalating CO₂ molecules have significant contributions to the enhancement of the embedding concentration of Zn and Li ions but have a feeble effect on Mg ions. Hence, we consider that the intercalation of CO₂ molecules significantly promotes bilayer MoS₂ to be a promising cathode material for high-capacity Zn and Li ion batteries, but only slightly for Mg ion batteries.

To gain an in-depth understanding of the diffusion barrier decrease originating from the intercalating CO₂ and the possible chemical reaction between CO₂ and embedding metal ions, we analyzed the charge difference between embedding metal ions and bilayer MoS₂. As a benchmark, the charge differences of metal ions in the pristine bilayer MoS₂ were evaluated as well. We calculate the charge density difference between the embedding ions and CO₂ molecule intercalated bilayer MoS₂ systems by using the formula of $\Delta\rho = \rho_{(\text{host}+\text{M})} - (\rho_{\text{host}} + \rho_{\text{M}})$, where $\rho_{(\text{host}+\text{M})}$, ρ_{host} and ρ_{M} represent the charge density of the ion embedded MoS₂, the pristine MoS₂ and the isolated metal ion, respectively. The simulation results are shown in Fig. 4, in which the purple color corresponds to the charge accumulation forming bonding states, and the green color represents the charge depletion forming anti-bonding states. Fewer bonding states and more anti-bonding states make the interaction stronger between embedded ions and their host systems. For the Zn ion embedding system as shown in Fig. 4a and d, the intercalating CO₂ molecules result in more charge depletion than in the pristine MoS₂, resulting in a decrease of the Zn ion intercalation energy. A similar tendency is also found in the Li ion embedding system shown in Fig. 4c and f. However, no significant change can be seen in the Mg ion embedding system shown in Fig. 4b and e. The results shown in Fig. 4 indicate that the intercalation of CO₂ molecules has a great contribution to the intercalation energy of Zn and Li ions only, consistent with the calculated intercalation energies of metal ions as shown in Fig. 3a. Also, we found no electron

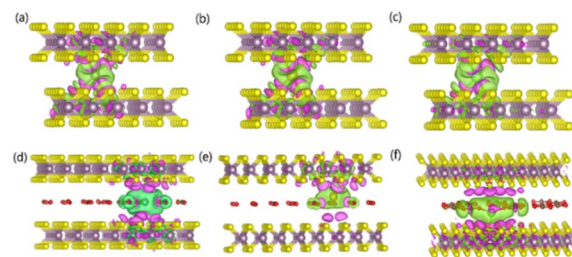


Fig. 4 Charge density difference isosurfaces for (a) Zn, (b) Mg and (c) Li interacting with pristine MoS₂ and the charge density difference isosurfaces for (d) Zn, (e) Mg and (f) Li interacting with CO₂ embedded MoS₂. The purple (green) color corresponds to charge accumulation (depletion). The isosurface is taken as $6 \times 10^{-5} e/\text{\AA}$.

transfer and bond formation between intercalating CO₂ and embedded Zn and Mg ions, except for the Li-ion system. We also calculated the projected density of states (PDOS) of the pristine and intercalated CO₂ (3/4 ML) MoS₂ as shown in Fig. 5a and b, respectively. It can be seen that O-p orbitals have strong interaction with Mo-d orbitals. The results reveal that CO₂ can be stabilized in the bilayer MoS₂. We also computed d-band centers of Mo from the intercalated CO₂ (3/4 ML) and pristine MoS₂, and the results are shown in Fig. 5c. Intercalated CO₂ molecules push the d-band of Mo to a higher energy state and result in a lower d-band center than in the pristine MoS₂, and further enhance the metal ion solubility in intercalating CO₂ bilayer MoS₂. Bader charge⁵⁴ provides the definition of the chemical bond for charge analysis, which is based on the

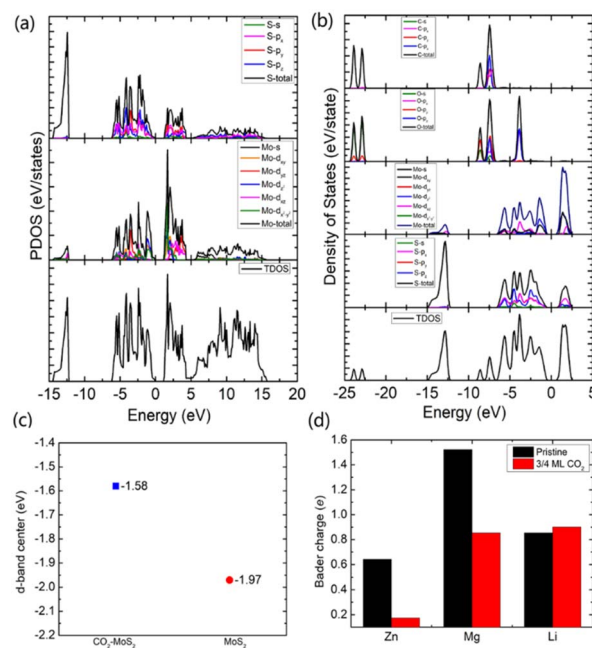


Fig. 5 The calculated projected density of states (PDOS) of the pristine MoS₂ (a) and CO₂ intercalated MoS₂ (b). (c) The calculated Mo d-band center with respect to the Fermi level (set to zero) of CO₂ intercalated and pristine MoS₂. (d) The calculated Bader charge of metal ions in the pristine and CO₂ intercalated (3/4 ML) MoS₂.

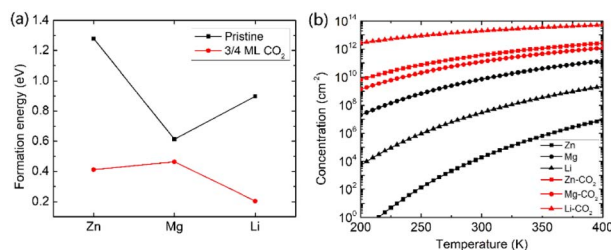


Fig. 3 The intercalation energies (a) and temperature-dependent concentration (b) of metal ions of Zn, Mg and Li in the pristine and modulated bilayer MoS₂ with an intercalating CO₂ coverage of 3/4 ML.

electronic charge density. In Fig. 5d, the Bader charges for Zn and Mg are significantly decreased but there is a little increase for Li ions with CO₂ embedding. The result reveal that intercalating CO₂ molecules weaken the ion interaction of Zn and Mg with host elements, while the slightly increased Bader charge of Li ions reveals that Li ions have strong chemical interaction with CO₂ as shown in Fig. 4f, which well explains the multi-saddle points in Li NEB results shown in Fig. 2c. We consider that those embedding CO₂ molecules would not react with Zn and Mg and may form chemical bonds with Li ions. Therefore, we believe that it is feasible to use CO₂ intercalation for improving the performance of MoS₂-based Zn and Mg ion batteries.

To gain an in-depth understanding of the location of CO₂ insertion at active sites, we computed the work function of pristine and CO₂ with/without metal ion intercalation bilayer MoS₂. The work function Φ is defined as the vacuum energy respective to the Fermi level ($\Phi = E_{\text{vac}} - E_{\text{f}}$). The work function of bilayer MoS₂ highly depends on the crystal orientation, the presence of impurities, defects, and doping. The calculated local potentials along the z-direction are shown in Fig. S8a.† It should be noted that the results shown in Fig. S8a† are simply plotted in one figure without considering the common alignment reference. Based on the calculated local potential along the z-direction shown in Fig. S8a,† we computed the work functions of selected five cases and the results are shown in Fig. S8b.† It can be seen that the calculated work function of the pristine bilayer MoS₂ is 6.02 eV, and the same as that of 3/4 ML CO₂ embedded MoS₂. The results reveal that intercalated CO₂ molecules do not affect the work function. Therefore, we considered that CO₂ is a promising material to expand the interlayer spacing of layered MoS₂. With metallic ion (Zn, Mg and Li) insertion, the work function decreases to 5.65 eV in Zn intercalated 3/4 ML CO₂ embedded MoS₂, further decreases to 4.78 eV in Mg intercalated 3/4 ML CO₂ embedded MoS₂, and further decreases to 4.53 eV in Li intercalated 3/4 ML CO₂ embedded MoS₂.

Conclusions

In conclusion, we developed a simple and effective strategy by tuning the interlayer spacing of MoS₂ and the intercalation energy of Zn, Mg and Li ions for cathode materials of rapidly chargeable batteries. Employing DFT simulations, we demonstrate that the intercalation of CO₂ molecules with a 3/4 ML coverage effectively expands the interlayer spacing, reduces layer coupling of bilayer MoS₂, and significantly reduces the diffusion barrier and intercalation energies of metal ions, achieving high performance metal ion batteries. As a result, the diffusivities of Zn and Li increase by 12 and 13 orders of magnitude, and the intercalating ion concentration or the storage capacities of Zn and Li ion batteries are boosted by 7 and 5 orders of magnitude. Our simulation result demonstrated that intercalated CO₂ molecules significantly contribute to both diffusivity and embedding concentration of Zn ions but they contribute to only the diffusivity of Mg and the embedding concentration of Li ions. Therefore, we demonstrate that CO₂

molecule intercalated bilayer MoS₂ is a promising cathode material for high-capacity and rapid-charging ZIBs. The strategy of expanding interlayer spacing and reducing intercalation energy developed in this study can be generally applied to increase the ion storage capability in layered structure-based electrode materials and sheds light on the development of advanced materials for next-generation high-performance energy storage.

Author contributions

Xuefei Han: methodology, investigation, writing – original draft, and writing – review & editing. Jing Yang: writing – review & editing. Yong-Wei Zhang: conceptualization and writing – review & editing. Zhi Gen Yu: conceptualization, methodology, writing – review & editing, and supervision.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This research was sponsored by the Agency for Science, Technology and Research (A*STAR), and the Department of Materials Science and Engineering (MSE), the National University of Singapore (NUS). The computational resource was provided by the A*STAR Computational Resource Centre, Singapore (A*CRC) and the National Supercomputing Centre, Singapore (<https://www.nsc.sg>). Y.-W. Z acknowledges the support from A*STAR-SERC-CRF Award.

Notes and references

- 1 P. K. Santra and P. V. Kamat, *J. Am. Chem. Soc.*, 2012, **134**, 2508–2511.
- 2 B. C. H. Steele and A. Heinzl, *Nature*, 2001, **414**, 345–352.
- 3 Y. Yang, K. C. Pradel, Q. Jing, J. M. Wu, F. Zhang, Y. Zhou, Y. Zhang and Z. L. Wang, *ACS Nano*, 2012, **6**, 6984–6989.
- 4 S. Xu, Y. Wei, J. Liu, R. Yang and Z. L. Wang, *Nano Lett.*, 2008, **8**, 4027–4032.
- 5 Z. L. Wang, *ACS Nano*, 2013, **7**, 9533–9557.
- 6 J. B. Goodenough and Y. Kim, *Chem. Mater.*, 2010, **22**, 587–603.
- 7 V. Etachri, R. Marom, R. Elazari, G. Salitra and D. Aurbach, *Energy Environ. Sci.*, 2011, **4**, 3243–3262.
- 8 Y. Zhang, P. Wang, T. Zheng, D. Li, G. Li and Y. Yue, *Nano Energy*, 2018, **49**, 596–602.
- 9 H. M. A. Hamid and Z. Çelik-Butler, *Nano Energy*, 2018, **50**, 159–168.
- 10 Y. Liang, H. D. Yoo, Y. Li, J. Shuai, H. A. Calderon, F. C. R. Hernandez, L. C. Grabow and Y. Yao, *Nano Lett.*, 2015, **15**, 2194–2202.
- 11 W. Ye, F. F. Wu, N. X. Shi, H. Zhou, Q. Q. Chi, W. H. Chen, S. Y. Du, P. Gao, H. B. Li and S. L. Xiong, *Small*, 2020, **16**, 1906607.
- 12 Z. Hu, L. Wang, K. Zhang, J. Wang, F. Cheng, Z. Tao and J. Chen, *Angew. Chem., Int. Ed.*, 2014, **53**, 12794–12798.



- 13 Y. Xu, F. Bahmani, M. Zhou, Y. Li, C. Zhang, F. Liang, S. Habib Kazemi, U. Kaiser, G. Meng and Y. Le, *Nanoscale Horiz.*, 2019, **4**, 202–207.
- 14 T. Xiong, Z. G. Yu, H. Wu, Y. Du, Q. Xie, J. Chen, Y.-W. Zhang, S. J. Pennycook, W. S. V. Lee and J. Xue, *Adv. Energy Mater.*, 2019, **9**, 1803815.
- 15 M. Liu, Q. Zhao, H. Liu, J. Yang, X. Chen, L. Yang, Y. Cui, W. Huang, W. Zhao, A. Song, Y. Wang, S. Ding, Y. Song, G. Qian, H. Chen and F. Pan, *Nano Energy*, 2019, **64**, 103942.
- 16 P. He, M. Yan, G. Zhang, R. Sun, L. Chen, Q. An and L. Mai, *Adv. Energy Mater.*, 2017, **7**, 1601920.
- 17 M. Song, H. Tan, D. Chao and H. J. Fan, *Adv. Funct. Mater.*, 2018, **28**, 1802564.
- 18 J. Yang, J. Wang, X. Dong, L. Zhu, D. Hou, W. Zeng and J. Wang, *Appl. Surf. Sci.*, 2021, **544**, 148775.
- 19 L. Zhou, F. Xiong, S. Tan, Q. An, Z. Wang, W. Yang, Z. Tao, Y. Yao, J. Chen and L. Mai, *Nano Energy*, 2018, **54**, 360–366.
- 20 N. Wu, Y. Lyu, R. Xiao, X. Yu, Y.-X. Yin, X.-Q. Yang, H. Li, L. Gu and Y.-G. Gu, *NPG Asia Mater.*, 2014, **6**, e120.
- 21 X. Jia, C. Liu, Z. G. Neale, J. Yang and G. Cao, *Chem. Rev.*, 2020, **120**, 7795–7866.
- 22 F. Wan and Z. Niu, *Angew. Chem., Int. Ed.*, 2019, **58**, 16358–16367.
- 23 D. Aurbach, D. Aurbach, Z. Lu, A. Schechter, Y. Gofer, H. Gizbar, R. Turgeman, Y. Cohen, M. Moshkovich and E. Levi, *Nature*, 2000, **407**, 724–727.
- 24 J. Muldoon, C. B. Bucur and T. Gregory, *Chem. Rev.*, 2014, **114**, 11683–11720.
- 25 H. D. Yoo, I. Shterenberg, Y. Gofer, G. Gershinsky, N. Poura and D. Aurbach, *Energy Environ. Sci.*, 2013, **6**, 2265–2279.
- 26 Q. Wang, T. Sun, S. Zheng, L. Li, T. Ma and J. Liang, *Inorg. Chem. Front.*, 2021, **8**, 4497–4506.
- 27 F. Ming, H. Liang, Y. Lei, S. Kandambeth, M. Eddaoudi and H. N. Alshareef, *ACS Energy Lett.*, 2018, **3**, 2602–2609.
- 28 D. Kundu, B. D. Adams, V. Duffort, S. H. Vajargah and L. F. Nazar, *Nat. Energy*, 2016, **1**, 16119.
- 29 C. Xia, J. Guo, Y. Lei, H. Liang, C. Zhao and H. N. Alshareef, *Adv. Mater.*, 2018, **30**, 1705580.
- 30 D. Chao, C. Zhu, M. Song, P. Liang, X. Zhang, N. H. Tiep, H. Zhao, J. Wang, R. Wang, H. Zhang and H. J. Fan, *Adv. Mater.*, 2018, **30**, 1803181.
- 31 W. Shi, B. Yin, Y. Yang, M. B. Sullivan, J. Wang, Y.-W. Zhang, Z. G. Yu, W. Si, V. Lee and J. Xue, *ACS Nano*, 2021, **15**, 1273–1281.
- 32 F. Wan, L. Zhang, X. Dai, X. Wang, Z. Niu and J. Chen, *Nat. Commun.*, 2018, **9**, 1656.
- 33 H. Zhang, D. Cao and X. Bai, *J. Power Sources*, 2019, **444**, 227299.
- 34 E. Levi, Y. Gofer and D. Aurbach, *Chem. Mater.*, 2010, **22**, 860–868.
- 35 R. Zhang, T. S. Arthur, C. Ling and F. Mizuno, *J. Power Sources*, 2015, **282**, 630–638.
- 36 I. Shterenberg, M. Salama, Y. Gofer, E. Levi and D. Aurbach, *MRS Bull.*, 2014, **39**, 453–460.
- 37 M. Chhowalla, H. S. Shin, G. Eda, L.-J. Li, K. P. Loh and H. Zhang, *Nat. Chem.*, 2013, **5**, 263–275.
- 38 X. Huang, Z. Zeng and H. Zhang, *Chem. Soc. Rev.*, 2013, **42**, 1934–1946.
- 39 J. Xu, J. Zhang, W. Zhang and C.-S. Lee, *Adv. Energy Mater.*, 2017, **7**, 1700571.
- 40 Y. Liu, S. Liu, X. Xie, Z. Li, P. Wang, B. Lu, S. Liang, Y. Tang and J. Zhou, *InfoMat*, 2023, **5**, e12374.
- 41 X. Chen, P. Ruan, X. Wu, S. Liang and J. Zhou, *Acta Phys.-Chim. Sin.*, 2022, **38**, 2111003.
- 42 V. P. Hoang Huy, Y. N. Ahn and J. Hur, *Nanomaterials*, 2021, **11**, 1517.
- 43 H. Li, L. Ma, C. Han, Z. Wang, Z. Liu, Z. Tang and C. Zhi, *Nano Energy*, 2019, **62**, 550–587.
- 44 H. Liang, Z. Cao, F. Ming, W. Zhang, D. H. Anjum, Y. Cui, L. Cavallo and H. N. Alshareef, *Nano Lett.*, 2019, **19**, 3199–3206.
- 45 H. Li, Q. Yang, F. Mo, G. Liang, Z. Liu, Z. Tang, L. Ma, J. Liu, Z. Shi and C. Zhi, *Energy Storage Mater.*, 2019, **19**, 94–101.
- 46 W. Xu, C. Sun, K. Zhao, X. Cheng, S. Rawal, Y. Xu and Y. Wang, *Energy Storage Mater.*, 2019, **16**, 527–534.
- 47 Z.-T. Shi, W. Kang, J. Xu, Y.-W. Sun, M. Jiang, T.-W. Ng, H.-T. Xue, D. Y. W. Yu, W. Zhang and C.-S. Lee, *Nano Energy*, 2016, **22**, 27–37.
- 48 S. Li, Y. Liu, X. Zhao, Q. Shen, W. Zhao, Q. Tan, N. Zhang, P. Li, L. Jiao and X. Qu, *Adv. Mater.*, 2021, **33**, 2007480.
- 49 P. Tao, H.-H. Guo, T. Yang and Z.-D. Zhang, *Chin. Phys. B*, 2014, **10**, 106801.
- 50 J. He, K. Hummer and C. Franchini, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2014, **89**, 075409.
- 51 H. Liang, Z. Cao, F. Ming, W. Zhang, D. H. Anjum, Y. Cui, L. Cavallo and H. N. Alshareef, *Nano Lett.*, 2019, **19**, 3199–3206.
- 52 C. Wu, G. Zhao, X. Yu, C. Liu, P. Lyu, G. Maurin, S. Le, K. Sun and N. Zhang, *Chem. Eng. J.*, 2021, **412**, 128736.
- 53 S. Bhoyate, J. Kim, E. Lee, B. Park, E. Lee, J. Park, S. H. Oh, J. Kim and W. Choi, *J. Mater. Chem. A*, 2020, **8**, 12436–12445.
- 54 G. Henkelman, A. Arnaldsson and H. Jónsson, *Comput. Mater. Sci.*, 2006, **36**, 254–360.

