

Cite this: *Nanoscale Adv.*, 2023, 5, 1124

# Photocatalytic formation of a gas permeable layer selectively deposited on supported metal nanoparticles for sintering-resistant thermal catalysis†

Ayato Takabayashi,<sup>a</sup> Fuminao Kishimoto,<sup>id</sup><sup>a</sup> Hiroto Tsuchiya,<sup>b</sup> Hitoshi Mikami<sup>b</sup> and Kazuhiro Takanabe<sup>id</sup><sup>\*a</sup>

Nanoparticle aggregation of supported metal catalysts at high temperatures is a serious problem that causes a drop in catalytic performance. This study investigates the protection of metal nanoparticles from sintering by selectively forming nanoscale SiO<sub>2</sub> shells on Pd supported on TiO<sub>2</sub> by ultraviolet (UV) light irradiation. The proton-coupled reduction reaction increases the local pH around Pd nanoparticles, resulting in hydrolysis of tetraethoxyorthosilicate (TEOS) in only the vicinity of the metal. An apparent quantum efficiency of only 0.6% is obtained for the Pd/TiO<sub>2</sub> catalyst in H<sub>2</sub> evolution from ethanol-containing water under 370 nm excitation light. Therefore, the pH of raw slurry solution should be precisely controlled to that slightly below the threshold value for the TEOS hydrolysis reaction before the photodeposition. Transmission electron microscopy (TEM) and energy dispersive X-ray spectroscopy (EDX) clearly show that the particle size of the Pd nanoparticles (~40 nm) with the SiO<sub>2</sub> shell (~20 nm) was almost unchanged by the high-temperature treatment at 900 °C in air, suggesting that the SiO<sub>2</sub> shell prevented thermal aggregation of Pd nanoparticles. The Pd/TiO<sub>2</sub> without SiO<sub>2</sub> shell decoration exhibited a drop in the number of active sites, which was likely due to aggregation of the Pd catalysts. However, the number of active sites on the Pd@SiO<sub>2</sub>/TiO<sub>2</sub> catalyst was maintained even after the catalyst was calcined at 900 °C. Consequently, the Pd@SiO<sub>2</sub>/TiO<sub>2</sub> catalyst maintained its catalytic performance for simulated exhaust gas purification even after treatment at 900 °C. This study presents a methodology to produce sintering-tolerant supported metal nanoparticles using the photocatalytic gas permeable layer fabrication method.

Received 12th October 2022  
Accepted 16th December 2022

DOI: 10.1039/d2na00703g

rsc.li/nanoscale-advances

## Introduction

An exhaust-gas purification system for automobiles uses an excessive amount of noble metals due to aggregation of the supported metal catalysts under harsh conditions, such as high-temperature hydrothermal conditions.<sup>1,2</sup> Aggregation is a phenomenon in which highly dispersed nanoparticles fuse on support materials at high temperature, resulting in particles with a large diameter and low dispersion.<sup>3,4</sup> The aggregation is typically induced by increased diffusion coefficients due to temperature beyond the Tammann temperature,<sup>5</sup> nanosizing of particles<sup>6</sup> (Ostwald ripening<sup>7</sup>), or direct particle migration and coalescence.<sup>8</sup> Suppressing the sintering could allow for high reactivity to be maintained for the exhaust-gas purification

reaction, which could minimize the amount of noble metals (Rh, Pd, and Pt) used in the system, thus reducing the system cost and mitigating security risks related to the rare metals.

Many studies have proposed ways to avoid the aggregation of metal particles on a support,<sup>9</sup> such as the development of a series of “super intelligent” catalysts,<sup>10</sup> enhanced metal-support interactions,<sup>11–13</sup> and coating metal nanoparticles with thin metal oxide layers.<sup>14–20</sup> However, the aggregation of the supported Pd nanoparticles under high-temperature treatment is complicated, and redispersion and sintering can occur during calcination in the presence of oxygen.<sup>21–23</sup> Two-dimensional surface Pd oxides during oxidation have been considered to be responsible for the redispersion of the Pd nanoparticles. Because the redispersion temperature depends on the Pd particle size, there is a very complex trade-off between the redispersion and aggregation of Pd nanoparticles. Such complex physicochemical phenomena lead us to fabricate an ensured protection layer on nanoparticles by encapsulation.<sup>24</sup>

Indeed, while there are many reports on the encapsulation of metal nanoparticles by metal oxide protection layers, most of them demonstrated self-assembly of metal alkoxides on

<sup>a</sup>Department of Chemical System Engineering, School of Engineering, The University of Tokyo, 7-3-1 Hongo, Bunkyo-ku, Tokyo, 113-8656, Japan. E-mail: takanabe@chemsys.t.u-tokyo.ac.jp

<sup>b</sup>Honda R&D Co., Ltd., 4630 Shimotakanezawa, Haga-machi, Hagagun, Tochigi 321-3393, Japan

† Electronic supplementary information (ESI) available: TEM and SEM images, and photocatalytic activity. See DOI: <https://doi.org/10.1039/d2na00703g>



colloidal metal nanoparticles.<sup>25–31</sup> The assembly of metal alkoxide precursors with a structure directing agent (SDA) can be induced by surfactant molecules or a microemulsion environment around the metal nanoparticles, resulting in the formation of gas-permeable metal oxide layers. Although these synthetic methods are well established in terms of materials chemistry, they are not suitable for industrial catalytic applications that require robust supports. Nevertheless, there are very limited examples of fabricating metal oxide protective layers on supported metal nanoparticle catalysts. Stair *et al.* developed a metal oxide decoration method on supported metal catalysts using atomic layer deposition (ALD).<sup>32,33</sup> Appropriate thermal treatment of the encapsulated catalyst can induce gas permeability of the metal oxide layers by the formation of cracks. However, large-scale manufacturing processes using ALD have not yet been established due to the requirement of vacuum conditions and extraordinary devices. Although strong metal–support interaction (SMSI) under reducing conditions that can also induce spontaneous encapsulation of supported metal nanoparticles is well studied,<sup>34,35</sup> the SMSI state is not stable during the oxidative exhaust-gas purification reaction.

In this study, we demonstrated a photocatalytic fabrication method of an ultrathin SiO<sub>2</sub> layer selectively decorated on already supported Pd metal nanoparticles to develop a sintering-resistant exhaust gas purification catalyst. The SiO<sub>2</sub> layer decoration was done using a slurry consisting of Pd/TiO<sub>2</sub> catalysts, tetraethylorthosilicate (TEOS) as a silica source, aqueous NaNO<sub>3</sub> as an electrolyte, and the tetramethylammonium cation (TMA<sup>+</sup>) as an SDA. Thermal decomposition of the SDA endows the SiO<sub>2</sub> layer with gas permeability, which enables the reactant exhaust gas components to get access to the Pd surface. This fabrication method is an application of our previous report on the decoration of a SiO<sub>2</sub> layer over the Pt/SrTiO<sub>3</sub> photocatalyst to develop efficient particulate photocatalysts for overall water splitting.<sup>36</sup> The role of the SiO<sub>2</sub> shell on Pt/SrTiO<sub>3</sub> is to function as a nano-membrane; H<sub>2</sub> generated inside the shell should escape through it, but O<sub>2</sub> generated outside cannot reach the Pt surface, preventing the reverse reaction of the overall water splitting. This reaction proceeded at close to room temperature in the liquid phase. The current paper investigates whether this nano-membrane selectively decorated on the supported metal particles can be applied to high-temperature exhaust gas purification. The physicochemical role of the SiO<sub>2</sub> layer in sintering-resistance of supported Pd nanoparticles was investigated. To apply the fabrication method of the SiO<sub>2</sub> layer for universal support materials, which are not specifically designed as a photocatalyst, precise control of the pH of the SiO<sub>2</sub> deposition solution is very crucial for effective SiO<sub>2</sub> decoration on Pd nanoparticles.

## Experimental

### Materials

All chemicals were used as received. Rutile TiO<sub>2</sub> (<5 μm, ≥99.9%), tetramethylammonium bromide (TMAB, ≥98.0%), NaNO<sub>3</sub> (≥99.0%), sodium hydroxide (NaOH, 99.9%), hydrochloric acid (HCl, 37%), methanol (99.5%), and ethanol (99.5%) were purchased from Sigma–Aldrich. HNO<sub>3</sub> solution containing

Pd(NO<sub>3</sub>)<sub>2</sub> (0.216 M) was purchased from Kojima Chemicals Co., Ltd. Tetraethyl orthosilicate (TEOS, >97.0%) was purchased from Tokyo Chemical Industry.

### Preparation of Pd-impregnated TiO<sub>2</sub>

The Pd/TiO<sub>2</sub> catalyst was prepared by a simple impregnation method. TiO<sub>2</sub> rutile (0.950 g) was dispersed in a mixture of ultrapure water (20 mL) and Pd(NO<sub>3</sub>)<sub>2</sub>·HNO<sub>3</sub> solution (0.202 or 1.01 mL corresponding to 1 or 5 wt% Pd). The resulting slurry was ultrasonicated for 15 min and stirred while heating in a water bath at 100 °C until the water completely evaporated. Next, the sample was placed in a muffle furnace and calcined under static air for 2 h at 500 °C, 700 °C, or 900 °C with a heating rate of 5 °C min<sup>-1</sup>. Finally, the products were reduced under pure H<sub>2</sub> for 2 h at 150 °C with a heating rate of 5 °C min<sup>-1</sup>.<sup>36–38</sup> The obtained catalysts were denoted as Pd/TiO<sub>2</sub>-500, Pd/TiO<sub>2</sub>-700, and Pd/TiO<sub>2</sub>-900, respectively.

### Photocatalytic performance test of Pd/TiO<sub>2</sub>

A photocatalytic performance test was conducted on the 5 wt% Pd/TiO<sub>2</sub> catalyst to see whether the excited electrons from TiO<sub>2</sub> could be utilized in this reaction. The Pd/TiO<sub>2</sub> catalyst (0.010 g) was dispersed in 10 mL of an aqueous solution of 10 vol% MeOH. The slurry was then irradiated with UV light under ambient pressure with stirring. The amount of evolved H<sub>2</sub> gas was quantified using an online GC (GC-8A, Shimadzu) equipped with a thermal conductivity detector (TCD). The apparent quantum yield (AQY) of the catalyst was calculated using the following equation:

$$\text{AQY} = 2 \times \frac{N(\text{H}_2)}{N(\text{photons})}$$

where  $N(\text{H}_2)$  and  $N(\text{photons})$  denote the number of H<sub>2</sub> molecules produced and the number of photons reaching the surface of the reaction solution, respectively.

### SiO<sub>2</sub> precipitation by increasing pH

A solution was prepared for the SiO<sub>2</sub> precipitation test by dissolving TMAB (5.25 g) and NaNO<sub>3</sub> (0.852 g) in a mixture of ultrapure water (100 mL) and ethanol (100 mL). TEOS (15.2 mL) was then added into the solution after the pH was adjusted to 3 with HCl (1.0 M). After 2 h of stirring, NaOH aqueous solution (10 mM) was added drop by drop to increase the pH. With every increase in pH of 0.5, the UV-Vis absorption spectrum and dynamic light scattering (DLS) of the solution were measured using a JASCO UV-Vis V-770 (wavelength range of 200 nm to 2700 nm with a scan speed of 400 nm min<sup>-1</sup>) and HORIBA nanoPartica SZ-100, respectively. For the DLS measurement, the diameter of the precipitates was determined from the average of 10 measurements.

### Preparation of the Pd@SiO<sub>2</sub>/TiO<sub>2</sub> structure

SiO<sub>2</sub> layer deposition on Pd/TiO<sub>2</sub> was performed in a solution consisting of TMAB, NaNO<sub>3</sub>, and TEOS. TMAB (7.86 g) and NaNO<sub>3</sub> (1.28 g) were dissolved in a mixture of ultrapure water



(150 mL) and ethanol (150 mL), and then TEOS (22.8 mL) was added to the solution after the pH was adjusted to 3 with HCl (1.0 M). After 2 h of stirring, Pd/TiO<sub>2</sub>-500 (0.150 g) was dispersed in the solution, and the pH was adjusted to 5 by the addition of an aqueous solution of NaOH (0.10 M). The resulting slurry was irradiated with UV light (Asahi Spectra Co., Ltd., CL-H1-365-9-1) for 0.5 h with stirring and Ar bubbling. The UV light had a peak wavelength of 370 nm and a power of 870 mW.

The precipitates were separated by centrifugation and dried for 16 h at 110 °C with a heating rate of 5 °C min<sup>-1</sup>. They were then calcined under static air for 2 h at 500 °C, 700 °C, or 900 °C with a heating rate of 5 °C min<sup>-1</sup> in a muffle furnace. Finally, the catalysts were reduced under 30 mL min<sup>-1</sup> of pure H<sub>2</sub> for 2 h at 150 °C with a heating rate of 5 °C min<sup>-1</sup>. They were denoted as Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500, Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-700, or Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-900, respectively.

### Carbon monoxide chemisorption

The obtained catalysts (0.040 g) were packed in a U-shape quartz tube and pretreated under 4% H<sub>2</sub> balanced with Ar (10 mL min<sup>-1</sup>) at 150 °C for 30 min with a heating rate of 40 °C min<sup>-1</sup>. After purging with pure Ar (10 mL min<sup>-1</sup>) at 150 °C for 30 min, the catalysts were cooled down to 10 °C with ice packs. Subsequently, 500 μL of diluted CO gas (Ar balance, 0.4%) was injected every 8 min until CO was no longer adsorbed. The amount of CO that slipped through without being adsorbed was quantified using a gas chromatograph with a flame ionization detector (SHIMADZU GC-FID, GC-8A) and a methanizer (SHIMADZU MTN-1) and calculated from the peak area.

### Characterization

Transmission electron microscope (TEM) images were obtained using a JEOL 2000EX-II (acceleration voltage: 200 kV) equipped with a CCD camera. The specimens for TEM were prepared by placing drops of the sample dispersed in EtOH onto a 200-mesh copper grid coated with an amorphous carbon film with holes. Scanning transmission electron microscope images and energy dispersive spectroscopy mappings (STEM/EDS) were obtained using an FEI Talos F200X equipped with a Super-X (FEI) and SDD detector (Bruker). Nitrogen adsorption-desorption isotherms were recorded using a MicrotracBEL BELMETAL III. Prior to the measurement, the samples were dried under vacuum at 100 °C for 2 h.

Elemental analysis of the catalysts was performed using a scanning electron microscope and an energy dispersive X-ray spectrometer (SEM/EDX, Hitachi High-Tech S-4700-D). The length of the measurement was 120 s, and the acceleration voltage was 20.00 kV. Inductively coupled plasma optical emission spectrometry measurements were also performed (ICP-OES, Agilent 5110 VDV). X-ray photoelectron spectroscopy (XPS) was performed using a JEOL JPS-9010 MC with a Mg K $\alpha$  X-ray source. The powder sample was deposited on a metal nickel plate, and powder X-ray diffraction (XRD) patterns of the catalysts were collected on a D/texUltra UltimaIII operated at 40 kV and 40 mA. The samples were scanned from 20° to 60° at 10° min<sup>-1</sup>, and the average of 10 measurements was taken.

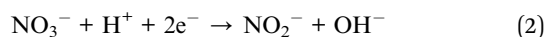
### Catalytic performance test

Catalytic reaction rates were measured using a fixed bed catalyst reactor system (MicrotracBEL, BELREA). First, 0.100 g of each catalyst was pretreated at 500 °C in both O<sub>2</sub> flow and H<sub>2</sub> flow for 15 min each. The catalytic performance was then tested in 400 mL min<sup>-1</sup> of simulated exhaust gas, which consisted of 500 ppm of NO, 5000 ppm of CO, 400 ppm of C<sub>3</sub>H<sub>6</sub>, 10% H<sub>2</sub>O, 14% CO<sub>2</sub>, 4900 ppm of O<sub>2</sub>, and 1700 ppm of H<sub>2</sub> in N<sub>2</sub> balance and ambient pressure. The reaction temperature was 200 °C to 500 °C. The outlet gas was analyzed by FT-IR (Best Instruments Co., Ltd., Bex-1000FT), an oxygen meter, and a hydrogen meter (Best Instruments Co., Ltd., Bex-520M).

## Results and discussion

### Photocatalytic shell formation scheme

Fig. 1 illustrates the shell formation scheme. Under UV light irradiation, the electrons generated by the photoexcitation of TiO<sub>2</sub> can be transferred to Pd nanoparticles. Next, the excited electrons in the Pd nanoparticles produce H<sub>2</sub> or reduce NO<sub>3</sub><sup>-</sup> as in the following equations, resulting in higher local pH in the vicinity of the Pd nanoparticles.



The locally increased pH induces the polymerization reaction of TEOS to form Si<sub>x</sub>O<sub>y</sub> oligomers, which are deposited on the Pd surface with TMA<sup>+</sup> by electrostatic force.<sup>39-41</sup> Finally, the removal of TMA<sup>+</sup> by the calcination process forms micropores in the SiO<sub>2</sub> shell, which allows gas diffusion to the core Pd.<sup>42,43</sup>

### Optimization of SiO<sub>2</sub> decoration conditions

The photocatalytic activity of 5 wt% Pd/TiO<sub>2</sub> is shown in Fig. S1,† for which the H<sub>2</sub> production rate was recorded in the presence of methanol as an electron donor. The 5 wt% Pd/TiO<sub>2</sub> catalyst (~0.6%) had an apparent quantum yield (AQY) at

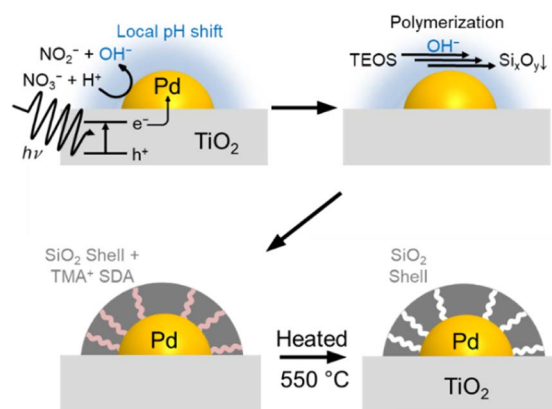


Fig. 1 Scheme of SiO<sub>2</sub> shell formation on supported Pd nanoparticles under UV light irradiation.



370 nm that was approximately one-hundredth lower than the reported AQY value for SrTiO<sub>3</sub>-based photocatalysts.<sup>39</sup> Hence, the local pH shift induced by the photocatalysis of 5 wt% Pd/TiO<sub>2</sub> was expected to be smaller than that of Pt/SrTiO<sub>3</sub>. Therefore, prior to the preparation of the catalyst, the threshold pH for SiO<sub>2</sub> precipitation was precisely measured. Based on the results, the initial pH for SiO<sub>2</sub> decoration on Pd/TiO<sub>2</sub> was determined to be slightly lower than the threshold pH.

Fig. 2a shows the UV-Vis transmission spectra of the solutions consisting of TEOS, NaNO<sub>3</sub>, and TMAB at various pH controlled by the addition of aqueous NaOH solution. The transmittance is hardly changed in the range of pH 3.5–5.5, but it begins to decrease slightly at pH 6.0 and significantly drops at pH 6.4. These results demonstrate that a small amount of Si<sub>x</sub>O<sub>y</sub> nanoparticles were formed at pH 6.0, and more were produced at pH 6.4. The formation of the nanoparticles was also observed by DLS measurements, as shown in Fig. 2b.

Below pH 6, almost no particles were present in the solution, which is consistent with the non-reproducible histograms derived from small fluctuations of the laser. Particles with a size distribution of 20–100 nm were observed at pH 6.0, which increased to 100–300 nm at pH 6.4. These results suggest that the formation of Si<sub>x</sub>O<sub>y</sub> precipitates by dehydration condensation of TEOS-derived Si(OH)<sub>4</sub> began at pH 5.5–6.0. Based on the results, the initial pH of the solution for SiO<sub>2</sub> deposition was set to 5.

### SiO<sub>2</sub> shell decoration on Pd/TiO<sub>2</sub>

Fig. 3 shows the TEM images of the 5 wt% Pd/TiO<sub>2</sub>-500, 5 wt% Pd/TiO<sub>2</sub>-900, 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500, and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-900 catalysts. Pd nanoparticles with a size of about 5–

30 nm are observed in 5 wt% Pd/TiO<sub>2</sub>-500 (Fig. 3a), but after 900 °C heat treatment in air, the Pd nanoparticles are sintered and have a particle size of more than 100 nm air (5 wt% Pd/TiO<sub>2</sub>-900, Fig. 3b). Histograms of Pd particle size on 5 wt% Pd/TiO<sub>2</sub>-500 and 5 wt% Pd/TiO<sub>2</sub>-900 obtained from TEM images are shown in Fig. 3k. The mean particle size increased from 16 nm to over 130 nm. The distribution of Pd particle size was statistically enlarged due to sintering induced by high-temperature heat treatment. The TEM image of 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 in Fig. 3c shows that Pd nanoparticles with a size of ~40 nm are covered by an amorphous shell with a thickness of ~15 nm. Other TEM images of 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 are shown in Fig. S2.† The particle size of Pd nanoparticles was roughly ~20 nm and the thickness of the SiO<sub>2</sub> layer was ~20 nm.

To confirm that the amorphous shell consists of SiO<sub>2</sub>, STEM/EDS mapping images of Pd and Si distribution on the catalysts were acquired (Fig. 3e–g). The distribution of Si is localized around the Pd nanoparticles in both 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-900, demonstrating that the amorphous shell layer on Pd nanoparticles in Fig. 3c should be SiO<sub>2</sub>. Similar to the 5 wt% Pd/TiO<sub>2</sub>, SiO<sub>2</sub> decoration on 1 wt% Pd/TiO<sub>2</sub> was also demonstrated, as shown in Fig. S3.† The Pd nanoparticles with a size of about 5 nm were covered by amorphous SiO<sub>2</sub> shells with a thickness of a few nanometers. Fig. S4† shows the SEM image of 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500. The particle size of the TiO<sub>2</sub> was ~1 μm.

To investigate the importance of UV light irradiation for uniform decoration of the SiO<sub>2</sub> shell on Pd nanoparticles, TEOS treatment of Pd/TiO<sub>2</sub> catalysts under dark conditions was performed as a control experiment. The TEM image of the Pd/TiO<sub>2</sub> after TEOS treatment under dark conditions (Fig. S5†) shows sparse deposition of amorphous SiO<sub>2</sub> clusters on the entire Pd/TiO<sub>2</sub> surface. Unlike Pd@SiO<sub>2</sub>/TiO<sub>2</sub> catalysts prepared by UV-light irradiation, the Pd nanoparticles were not covered by the uniformly decorated SiO<sub>2</sub> nanoparticles. Therefore, the photocatalytic reaction induced by UV light irradiation is essential for the uniform SiO<sub>2</sub> shell decoration on Pd nanoparticles.

Weight fractions of Si and Pd in 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 were determined by ICP-OES measurement and the results are shown in Table 1. The weight percent of Pd was measured to be 5.0 wt%, which is equal to the Pd quantity in the impregnation solution. The weight fraction of SiO<sub>2</sub> was determined to be 5.9 wt%. No impurities of Na and Br were detected. According to the density of the Pd and SiO<sub>2</sub> (12.0 g cm<sup>-3</sup> and 2.65 g cm<sup>-3</sup>, respectively), the volume ratio between the Pd and SiO<sub>2</sub> was calculated as 1:5.3.

Assuming that all the SiO<sub>2</sub> forms a hemispherical shell on the hemispherical Pd nanoparticles, the thickness of the shell should be about 0.4 times the diameter of the Pd nanoparticles. Hence, if the Pd grain size is 40 nm, the thickness of the SiO<sub>2</sub> shell is calculated to be 16 nm, which is roughly consistent with the TEM result (Fig. 3c). Therefore, the weight content of the SiO<sub>2</sub> (~6 wt%) should be reasonable. SiO<sub>2</sub> particles were also observed on the TiO<sub>2</sub> surface because SiO<sub>2</sub> precursors are unavoidably adsorbed on the TiO<sub>2</sub> surface.

The Pd particle size is almost maintained even after heat treatment at 900 °C in air (5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-900, Fig. 3d).

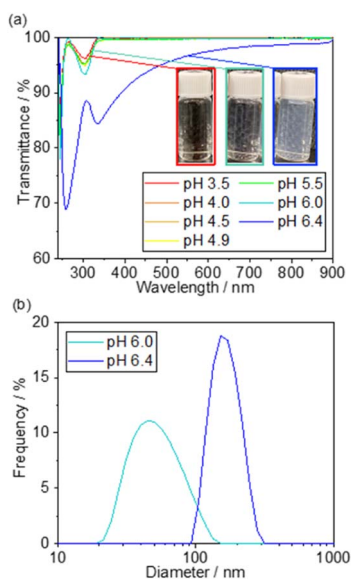


Fig. 2 SiO<sub>2</sub> precipitation test results. (a) UV-Vis transmission spectra of the solutions consisting of TEOS, NaNO<sub>3</sub>, and TMAB at various pH controlled by the addition of aqueous NaOH solution. (b) The particle size distribution of the precipitates in the solution at pH 6.0 and 6.4 obtained by DLS measurement.



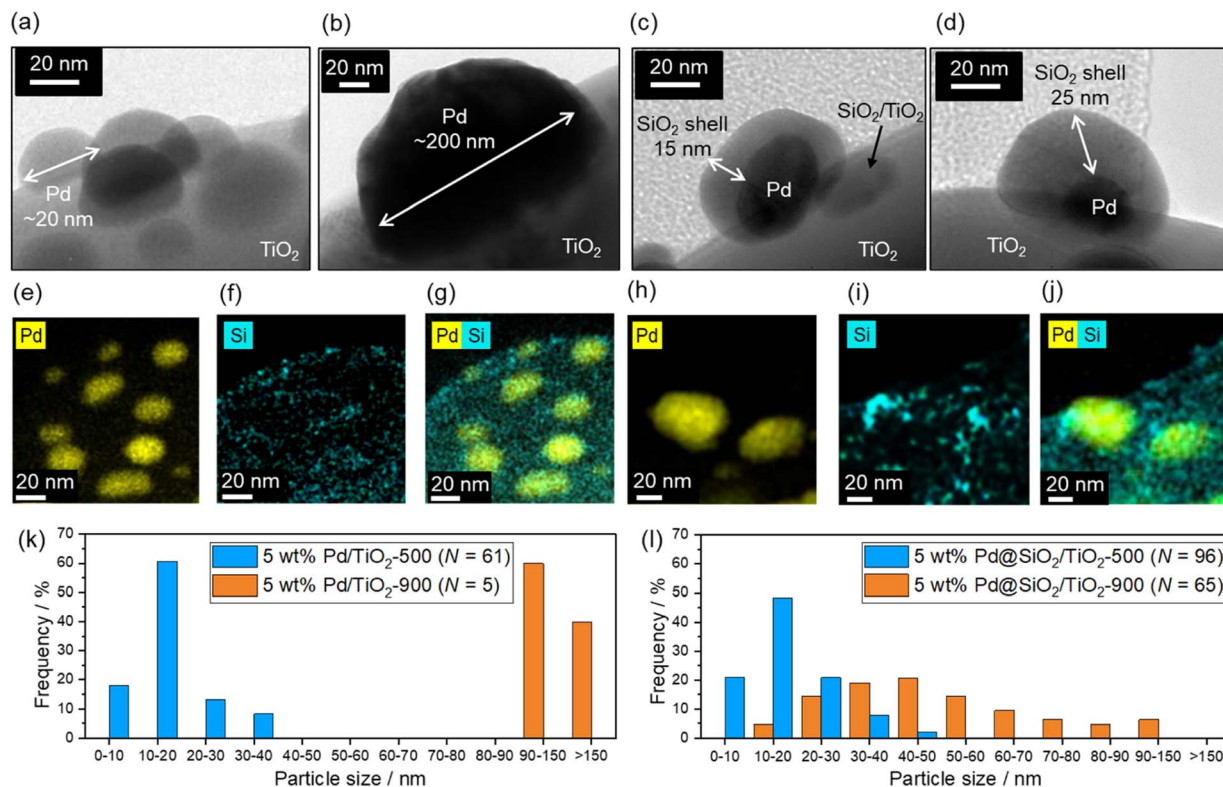


Fig. 3 TEM images of (a) 5 wt% Pd/TiO<sub>2</sub>-500, (b) 5 wt% Pd/TiO<sub>2</sub>-900, (c) 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500, and (d) 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-900, respectively. STEM/EDS mapping images of (e–g) 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 and (h–j) 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-900. Histogram of Pd particle size on (k) 5 wt% Pd/TiO<sub>2</sub>-500 and 5 wt% Pd/TiO<sub>2</sub>-900 obtained from TEM images, and (l) 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-900 obtained from STEM/EDS mappings.

Table 1 ICP results of 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-900

Sample	Component ratio/wt%			
	Pd	SiO <sub>2</sub>	TiO <sub>2</sub>	Si/Pd ratio
5 wt% Pd@SiO <sub>2</sub> /TiO <sub>2</sub> -500	5.0	5.9	89	2.1
5 wt% Pd@SiO <sub>2</sub> /TiO <sub>2</sub> -900	5.1	6.3	89	2.2

The diameter of the Pd nanoparticles seems to be slightly increased after heat treatment at 900 °C (Fig. 3h) in comparison to the STEM/EDS mapping images of Pd in 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 (Fig. 3e). Histograms of Pd particle size on 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-900 were obtained from the STEM/EDS mapping and shown in Fig. 3l. Although the mean particle size increased from 23 nm to 53 nm after 900 °C heat treatment, the Pd particle size of 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-900 was clearly statistically smaller compared to 5 wt% Pd/TiO<sub>2</sub>-900. The localization of the Si distribution around Pd particles was also measured (Fig. 3i and j), and the weight percent of Pd and SiO<sub>2</sub> was not changed after heat treatment at 900 °C in air (Table 1). This suggests that the SiO<sub>2</sub> shell prevents aggregation of the Pd nanoparticles.

Table 2 summarizes the amount of CO adsorbed on 5 wt% Pd/TiO<sub>2</sub> and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub> after heat treatment in air at various temperatures. The amount of CO adsorbed on 5 wt%

Pd/TiO<sub>2</sub> after heat treatment at 500 °C drastically dropped from 15 μmol g<sub>cat</sub><sup>-1</sup> to 2.8 μmol g<sub>cat</sub><sup>-1</sup> after heat treatment at 900 °C. The rate of the decrease of the surface Pd site after heat treatment at 900 °C in air was calculated to be 22%. The calculated particle diameter with an assumption of hemispherical Pd particles was 25 nm after heat treatment at 500 °C and 139 nm after heat treatment at 900 °C, which is consistent with the TEM results.

The CO adsorption capacity on 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub> was 94% retained even after heat treatment at 900 °C, suggesting that the SiO<sub>2</sub> decoration suppressed the aggregation of Pd nanoparticles and achieved excellent high-temperature stability. Although the CO adsorption amount decreases for

Table 2 Results of CO chemisorption capacity (unit: μmol g<sub>cat</sub><sup>-1</sup>)

Sample	Heat treatment temperature			Rate of decrease by 900 °C heating
	500 °C	700 °C	900 °C	
5 wt% Pd/TiO <sub>2</sub>	15	13	2.8	19%
5 wt% Pd@SiO <sub>2</sub> /TiO <sub>2</sub>	5.0	5.2	4.7	94%
1 wt% Pd/TiO <sub>2</sub>	5.4	n.t. <sup>a</sup>	1.2	22%
1 wt% Pd@SiO <sub>2</sub> /TiO <sub>2</sub>	1.7	n.t. <sup>a</sup>	1.7	100%

<sup>a</sup> Not tested.



5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 (5.0 μmol g<sub>cat</sub><sup>-1</sup>) compared to 5 wt% Pd/TiO<sub>2</sub>-500 (15 μmol g<sub>cat</sub><sup>-1</sup>), this result may not simply be due to the SiO<sub>2</sub> layer inhibiting the adsorption of CO onto the Pd surface. As will be discussed later, the catalytic activity of exhaust gas purification reactions is almost unchanged with or without the SiO<sub>2</sub> layer. Hence, it can be concluded that SiO<sub>2</sub> does not block the surface sites, and the decrease in CO adsorption by the SiO<sub>2</sub> layer modification can be attributed to the difference in the interaction between CO and the Pd surface.

A similar trend was observed when the loaded Pd amount was decreased to 1 wt%. The amount of CO adsorbed on 1 wt% Pd/TiO<sub>2</sub> significantly drops to 22% after heat treatment at 900 °C in air, but the CO adsorption amount on 1 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub> is well maintained (~100%) after the heat treatment. Therefore, aggregation of Pd nanoparticles at elevated temperature can be effectively prevented by the shell.

X-ray photoemission spectroscopy was also performed on 5 wt% Pd/TiO<sub>2</sub>-500 and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500. In the survey scan (Fig. 4a), Si peaks that cannot be seen for 5 wt% Pd/SiO<sub>2</sub>/TiO<sub>2</sub>-500 clearly appear for 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500, indicating that SiO<sub>2</sub> was deposited on the catalyst. The detailed chemical state of Si species on the Pd/TiO<sub>2</sub> catalyst was characterized using the Si 2p region X-ray photoemission spectrum (Fig. 4b). Based on the binding energy of the Si 2p peaks, the deposited Si species on both Pd@SiO<sub>2</sub>/TiO<sub>2</sub> and Pd/TiO<sub>2</sub> after TEOS treatment under dark conditions can be identified as SiO<sub>2</sub>. The spectra show that the Si 2p peak position of Pd@SiO<sub>2</sub>/TiO<sub>2</sub> was positively shifted compared to the Si 2p peak of the Pd/TiO<sub>2</sub> catalyst after TEOS treatment under dark conditions. Unlike the SiO<sub>2</sub> clusters sparsely deposited on the Pd/TiO<sub>2</sub> catalyst, the SiO<sub>2</sub> shell localized on the Pd nanoparticles became positively charged probably due to electron withdrawal by Pd nanoparticles.

Fig. 4c shows the detailed spectra of the region of Pd 3d peaks. The peak area decreased after the SiO<sub>2</sub> decoration, suggesting that the Pd nanoparticles were successfully covered by SiO<sub>2</sub>. Pd 3d peaks were also slightly decreased by the TEOS treatment, suggesting that the Pd nanoparticles should be

partially coated with the SiO<sub>2</sub> clusters. However, the peak intensity is much higher than that of the Pd@SiO<sub>2</sub>/TiO<sub>2</sub> catalysts prepared by UV light irradiation. Therefore, the photocatalytic reaction triggers uniform SiO<sub>2</sub> shell formation on Pd nanoparticles, which corresponds to the TEM observation.

Deconvolution revealed the components of metallic Pd (340.5 eV and 335.3 eV) and PdO (342.3 eV and 337.0 eV) in the spectra of both 5 wt% Pd/TiO<sub>2</sub>-500 and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500. The ratio of the PdO peak to the Pd metal peak is higher for 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500, which is probably due to the selective SiO<sub>2</sub> decoration on metallic Pd or oxidation of the surface Pd species resulting from the formation of an interface with the SiO<sub>2</sub> shell.

The surface area and porosity of the SiO<sub>2</sub> shell were examined using nitrogen adsorption–desorption isotherms and carbon monoxide pulsed adsorption measurement, as shown in Fig. 5. The isotherm of 5 wt% Pd/TiO<sub>2</sub>-500 exhibits type II features and a Brunauer–Emmett–Teller (BET) surface area of 4.5 m<sup>2</sup> g<sub>cat</sub><sup>-1</sup> (Fig. 5). However, the isotherm of 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub> shows a steep rise in the lower relative pressure region, suggesting the presence of micropores. The BET surface area of Pd@SiO<sub>2</sub>/TiO<sub>2</sub> was 45 m<sup>2</sup> g<sub>cat</sub><sup>-1</sup>. The 10-fold increase in the surface area of 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub> compared to Pd/TiO<sub>2</sub> is attributed to the porous nature of the SiO<sub>2</sub> shell on Pd.

Fig. 6 shows the XRD patterns of 5 wt% Pd/TiO<sub>2</sub> and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub> after heating at various temperatures. All patterns exhibited a small peak around 40.1° corresponding to Pd<sup>0</sup>(111). The pattern of 5 wt% Pd/TiO<sub>2</sub> demonstrated that the TiO<sub>2</sub> support consisted of a major rutile phase and a minor anatase phase. The minor anatase peaks of 5 wt% Pd/TiO<sub>2</sub> disappeared after heat treatment at 900 °C. Typically, the transition temperature of TiO<sub>2</sub> without any modifications is around 800 °C. However, 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub> showed the anatase peaks even after heat treatment at 900 °C, suggesting that SiO<sub>2</sub> deposition on the Pd/TiO<sub>2</sub> shifted the anatase-to-rutile transition temperature of the TiO<sub>2</sub> support to a higher temperature.

Okada *et al.* reported an upper shift of the transition temperature of TiO<sub>2</sub> by a few atomic percent with Si doping. This was due to the suppression of surface nucleation sites for rutile by the formation of an amorphous SiO<sub>2</sub> surface layer.<sup>44</sup> Similar to this report, the SiO<sub>2</sub> partially covering the Pd/TiO<sub>2</sub>

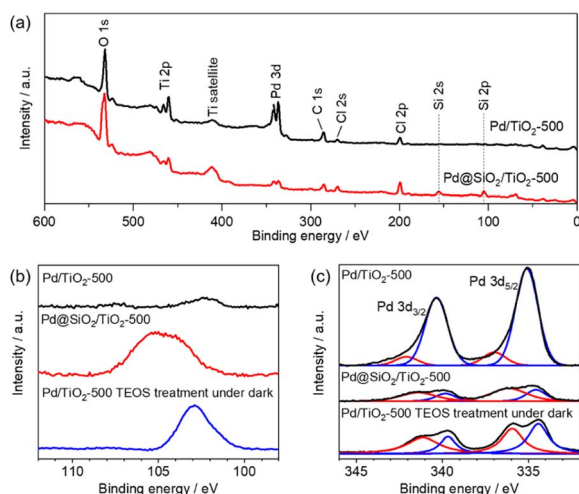


Fig. 4 XPS results of 5 wt% Pd/TiO<sub>2</sub>-500 and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500. (a) Survey scan, (b) Si 2p region and (c) Pd 3d region.

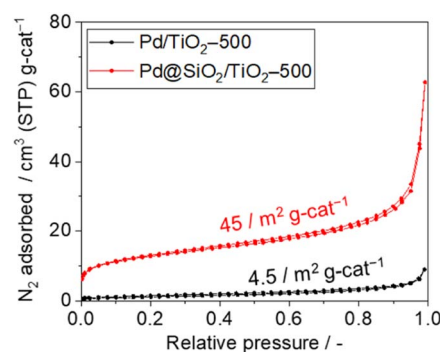


Fig. 5 Nitrogen adsorption–desorption isotherms of 5 wt% Pd/TiO<sub>2</sub>-500 and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500.



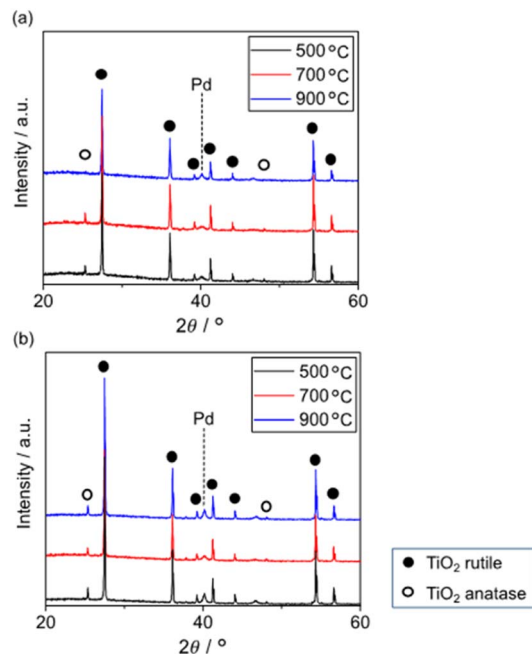


Fig. 6 XRD patterns of (a) 5 wt% Pd/TiO<sub>2</sub>-500, 700, 900 and (b) 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500, 700, 900.

applied *via* photodeposition seemed to have the ability to suppress the nucleation sites of rutile TiO<sub>2</sub>.

### Catalytic performance of Pd@SiO<sub>2</sub>/TiO<sub>2</sub>

The purification reaction of simulated exhaust gas was examined to investigate the influence of the high-temperature

treatment of 5 wt% Pd/TiO<sub>2</sub> and 5 wt% Pd@SiO<sub>2</sub>/TiO<sub>2</sub> on the catalytic performance (Fig. 7). The test was carried out under a gas flow of NO (500 ppm), CO (5000 ppm), C<sub>3</sub>H<sub>6</sub> (400 ppm), H<sub>2</sub>O (10%), CO<sub>2</sub> (14%), O<sub>2</sub> (4900 ppm), and H<sub>2</sub> (1700 ppm) with N<sub>2</sub> balance at ambient pressure. The conversion of NO reduction, CO oxidation, and C<sub>3</sub>H<sub>6</sub> oxidation was estimated from the concentrations of NO, CO, and C<sub>3</sub>H<sub>6</sub> in outlet gas, respectively, while N<sub>2</sub> reduction was estimated using the NH<sub>3</sub>, NO<sub>2</sub>, and N<sub>2</sub>O concentrations in the outlet gas.

In all reactions, the conversion increased at an elevated reaction temperature, but there was a relatively high conversion of NO and N<sub>2</sub> reduction at lower temperature due to the selective reduction by H<sub>2</sub> contained in the reaction gas. In other words, the reducing agent for NO and N<sub>2</sub> reduction was changed from H<sub>2</sub> at lower temperature to other gases (*e.g.*, CO or C<sub>3</sub>H<sub>8</sub>) at higher temperature. This occurred because the H<sub>2</sub> was consumed by the reaction with O<sub>2</sub> at higher temperature.<sup>45</sup>

The conversion of NO reduction, N<sub>2</sub> reduction, and C<sub>3</sub>H<sub>8</sub> oxidation over 5 wt% Pd/TiO<sub>2</sub> significantly dropped after heat treatment at 900 °C in air due to the aggregation of the Pd nanoparticles. In contrast, Pd@SiO<sub>2</sub>/TiO<sub>2</sub> maintained the conversion of the NO and N<sub>2</sub> reduction or improved the conversion of C<sub>3</sub>H<sub>8</sub> oxidation even after the treatment. This demonstrates the protective role of the SiO<sub>2</sub> shell on Pd nanoparticles in preventing the thermal drop of the catalytic performance. The catalytic performance of Pd@SiO<sub>2</sub>/TiO<sub>2</sub>-500 for the NO reduction, N<sub>2</sub> reduction and C<sub>3</sub>H<sub>8</sub> oxidation was almost corresponding to the performance of Pd/TiO<sub>2</sub>-500. Therefore, the gas permeability of the SiO<sub>2</sub> shell was sufficiently high not to prevent the catalytic reaction at Pd nanoparticles. Moreover, SiO<sub>2</sub> decoration does not inhibit the Pd surface site

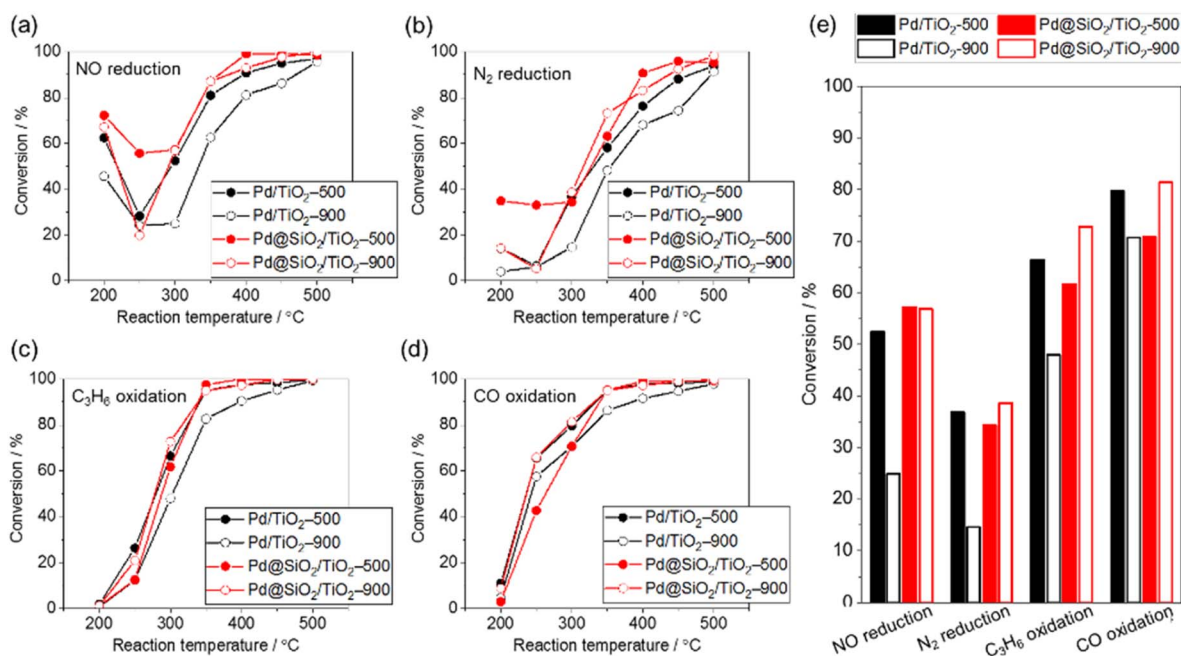


Fig. 7 Catalytic performance in simulated exhaust gas (500 ppm NO, 5000 ppm CO, 400 ppm C<sub>3</sub>H<sub>6</sub>, 10% H<sub>2</sub>O, 14% CO<sub>2</sub>, 4900 ppm O<sub>2</sub>, and 1700 ppm H<sub>2</sub> in N<sub>2</sub> balance, 400 mL min<sup>-1</sup>, ambient pressure). (a) NO reduction, (b) N<sub>2</sub> reduction, (c) hydrocarbon oxidation, (d) CO oxidation, and (e) summarized bar graph at 300 °C.



for activation of exhaust gas. On the other hand, the performance of CO oxidation is not significantly changed by the SiO<sub>2</sub> decoration and probably due to CO diffusion limitation of the CO oxidation kinetics above 277 °C.<sup>46</sup>

There have been no previous demonstrations of the use of these photocatalytically fabricated gas-permeable layers for high-temperature thermal catalytic reactions, and this paper is the first report of such a demonstration. The current photocatalytic fabrication method can be applied for a large variety of gas permeable layers, such as chromium-,<sup>47</sup> molybdenum-,<sup>48</sup> titanium-,<sup>49,50</sup> tantalum-based oxides,<sup>49,50</sup> and so on. Furthermore, this photo-assisted fabrication method does not require special vacuum conditions and can be applied for mass production using a continuous-flow synthesis system.<sup>51</sup> This study bridges strategies of photocatalysis-assisted nanostructure synthesis and thermal catalysts with sintering tolerance.

## Conclusions

In this work, ultrathin SiO<sub>2</sub> shells were successfully synthesized on Pd nanoparticles supported by TiO<sub>2</sub> by a local pH shift near the supported Pd catalyst using a UV light-induced photocatalytic reaction on the TiO<sub>2</sub> support. TEM and STEM/EDS mapping images showed that the particle size of the Pd nanoparticles (~40 nm) with the SiO<sub>2</sub> shell (~20 nm) was almost unchanged by high-temperature treatment at 900 °C in air. This suggested that the SiO<sub>2</sub> shell prevented thermal aggregation of the nanoparticles. The Pd/TiO<sub>2</sub> without the SiO<sub>2</sub> shell exhibited a drop in the number of active sites, which was likely due to aggregation of the -Pd catalysts. However, the number of active sites on the Pd@SiO<sub>2</sub>/TiO<sub>2</sub> catalyst was maintained even after the catalyst was calcined at 900 °C. Consequently, the SiO<sub>2</sub> shell on Pd nanoparticles prevented the drop in the catalytic performance of the Pd/TiO<sub>2</sub> catalyst in simulated exhaust gas purification without disturbing the catalytic performance.

## Author contributions

K. T. and H. M. led the research and contributed to the experiment planning, results analysis, and supervision of the research project. A. T. and H. T. carried out the experiments. F. K. and H. T. were involved in the design of experiments, results analysis, and paper writing. All the authors worked on the preparation of the manuscript and provided useful comments.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

F. K. would appreciate the financial support of the Leading Initiative for Excellent Young Researchers (LEADER) program by the Ministry of Education, Culture, Sports, Science and Technology (MEXT), Japan. The study was financially supported by Honda R&D Co., Ltd.

## References

- 1 A. Beniya and S. Higashi, *Nat. Catal.*, 2019, **2**, 590.
- 2 R. J. Farrauto, M. Deeba and S. Alerasool, *Nat. Catal.*, 2019, **2**, 603.
- 3 S. E. Wanke and P. C. Flynn, *Catal. Rev.*, 1975, **12**, 93.
- 4 E. Ruckenstein and B. Pulvermacher, *J. Catal.*, 1973, **29**, 224.
- 5 T. J. Gray, *J. Am. Ceram. Soc.*, 1954, **37**, 534.
- 6 K. Dick, T. Dhanasekaran, Z. Zhang and D. Meisel, *J. Am. Chem. Soc.*, 2002, **124**, 2312.
- 7 S. B. Simonsen, I. Chorkendorff, S. Dahl, M. Skoglundh, J. Sehested and S. Helveg, *J. Catal.*, 2011, **281**, 147.
- 8 T. W. Hansen, A. T. Delariva, S. R. Challa and A. K. Datye, *Acc. Chem. Res.*, 2013, **46**, 1806.
- 9 A. K. Datye and M. Votsmeier, *Nat. Mater.*, 2021, **20**, 1049.
- 10 Y. Nishihata, J. Mizuki, T. Akao, H. Tanaka, M. Uenishi, M. Kimura, T. Okamoto and N. Hamada, *Nature*, 2002, **148**, 164.
- 11 H. Jeong, O. Kwon, B. Kim, J. Bae, S. Shin, H. Kim, J. Kim and H. Lee, *Nat. Catal.*, 2020, **3**, 368.
- 12 E. J. Peterson, A. T. DeLaRiva, S. Lin, R. S. Johnson, H. Guo, J. T. Miller, J. H. Kwak, C. H. F. Peden, B. Kiefer, L. F. Allard, F. H. Ribeiro and A. K. Datye, *Nat. Commun.*, 2014, **5**, 4885.
- 13 V. Muravev, G. Spezzati, Y. Q. Su, A. Parastaev, F. K. Chiang, A. Longo, C. Escudero, N. Kosinow and E. J. Hensen, *Nat. Catal.*, 2021, **4**, 469.
- 14 Y. Lu, Y. Yin, Z. Li and Y. Xia, *Nano Lett.*, 2002, **2**, 785.
- 15 J. N. Park, A. J. Forman, W. Tang, J. Cheng, Y. S. Hu, H. Lin and E. W. McFarland, *Small*, 2008, **4**, 1694.
- 16 N. Almana, S. P. Phivilay, P. Laveille, M. N. Hedhili, P. Fornasiero, K. Takanabe and J. Basset, *J. Catal.*, 2016, **340**, 368.
- 17 R. G. Chaudhuri and S. Paria, *Chem. Rev.*, 2012, **112**, 2373.
- 18 M. B. Gawande, A. Goswami, T. Asefa, H. Guo, A. V. Biradar, D. Peng, R. Zboril and R. S. Varma, *Chem. Soc. Rev.*, 2015, **44**, 7540.
- 19 J. Li, Y. Zhang, S. Ding, R. Panneerselvam and Z. Tian, *Chem. Rev.*, 2017, **117**, 5002.
- 20 H. O. Otor, J. B. Steiner, C. Garcia-Sancho and A. C. Albarubio, *ACS Catal.*, 2020, **10**, 7630.
- 21 E. Ruckenstein and J. J. Chen, *J. Phys. Chem.*, 1981, **85**, 1606.
- 22 E. Ruckenstein and J. J. Chen, *J. Colloid Interface Sci.*, 1982, **86**, 1.
- 23 H. Lieske and J. Völter, *J. Phys. Chem.*, 1985, **89**, 1841.
- 24 A. Aitbekova, *et al.*, *Nat. Mater.*, 2022, **21**, 1290.
- 25 Q. Zhang, I. Lee, J. B. Joo, F. Zaera and Y. Yin, *Acc. Chem. Res.*, 2013, **46**, 1816.
- 26 C. Y. Seo, X. Chen, K. Sun, L. F. Allard, G. B. Fisher and J. W. Schwank, *Catal. Commun.*, 2018, **108**, 73.
- 27 A. J. Hill, C. Y. Seo, X. Chen, A. Bhat, G. B. Fisher, A. Lenert and J. W. Schwank, *ACS Catal.*, 2019, **10**, 1731.
- 28 S. H. Joo, J. Y. Park, C. K. Tsung, Y. Yamada, P. Yang and G. A. Somorjai, *Nat. Mater.*, 2009, **8**, 126.
- 29 M. Cargnello, N. L. Wieder, T. Montini, R. J. Gorte and P. Fornasiero, *J. Am. Chem. Soc.*, 2010, **132**, 1402.





- 30 K. A. Dahlberg and J. W. Schwank, *Chem. Mater.*, 2012, **24**, 2635.
- 31 Z. Niu and Y. Li, *Chem. Mater.*, 2014, **26**, 72.
- 32 J. Lu, B. Fu, M. C. Kung, G. Xiao, J. W. Elam, H. H. Kung and P. C. Stair, *Science*, 2012, **335**, 1205.
- 33 J. Lu, J. W. Elam and P. C. Stair, *Acc. Chem. Res.*, 2013, **46**, 1806.
- 34 S. J. Tauster, S. C. Fung and R. L. Garten, *J. Am. Chem. Soc.*, 1978, **10**, 170.
- 35 H. Xin, R. Li, D. Li, T. Song, R. Mu, Q. Fu and X. Bao, *J. Am. Chem. Soc.*, 2022, **144**, 4874.
- 36 H. Zhu, Z. Qin, W. Shan, W. Shen and J. Wang, *J. Catal.*, 2005, **233**, 41.
- 37 H. Zhu, Z. Qin, W. Shan, W. Shen and J. Wang, *Catal. Today*, 2007, **126**, 382.
- 38 F. Liang, H. Zhu, Z. Qin, G. Wand and J. Wand, *Catal. Commun.*, 2009, **10**, 737.
- 39 J. A. Bau and K. Takanabe, *ACS Catal.*, 2017, **7**, 7931.
- 40 A. Walcarius, E. Sibottier, M. Etienne and J. Ghanbaja, *Nat. Mater.*, 2007, **6**, 602.
- 41 A. Goux, M. Etienne, E. Aubert, C. Lecomte, J. Ghanbaja and A. Walcarius, *Chem. Mater.*, 2009, **21**, 731.
- 42 G. Herzog, E. Sibottier, M. Etienne and A. Walcarius, *Faraday Discuss.*, 2013, **164**, 259.
- 43 S.-H. Wu, C.-Y. Mou and H.-P. Lin, *Chem. Soc. Rev.*, 2013, **42**, 3862.
- 44 K. Okada, N. Yamamoto, Y. Kameshima, A. Yasumori and K. J. D. MacKenzie, *J. Am. Ceram. Soc.*, 2001, **84**, 1591.
- 45 H. Hamada and M. Haneda, *Appl. Catal., A*, 2012, **421**, 1.
- 46 B. Michel, *Kinetics of Heterogeneous Catalytic Reactions*, Princeton University Press, 1984.
- 47 K. Maeda, K. Teramura, D. Lu, N. Saito, Y. Inoue and K. Domen, *Angew. Chem., Int. Ed.*, 2006, **45**, 7806.
- 48 A.-T. Garcia-Esparza, *et al.*, *Angew. Chem., Int. Ed.*, 2017, **56**, 5780–5784.
- 49 T. Takata, C. Pan, M. Nakabayashi, N. Shibata and K. Domen, *J. Am. Chem. Soc.*, 2015, **137**, 9627.
- 50 T. Suguro, *et al.*, *Nat. Commun.*, 2022, **13**, 5698.
- 51 D. Cambie, C. Bottecchia, N. J. Straathof, V. Hessel and T. Noel, *Chem. Rev.*, 2016, **116**, 10276.

