



Cite this: *Green Chem.*, 2023, **25**, 8510

Received 4th July 2023,
Accepted 5th October 2023
DOI: 10.1039/d3gc02408c

rsc.li/greenchem

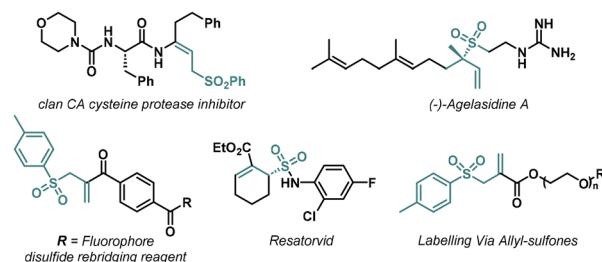
This work describes the electrochemical synthesis of allyl sulfones via sulfonation of allyl trifluoroborates. The process involves a radical addition to the alkene followed by the elimination of the trifluoroborate moiety, giving access to different substituted allyl sulfones. The efficacy of this reaction is attributed to the β - σ (C–B) stabilization of the electrochemically-generated intermediate carbocation, which plays a pivotal role in enforcing regioselectivity during the process. The reaction system is robust and general, allowing for compatibility with aromatic, heteroaromatic, and aliphatic sulfinic and allylic systems. The protocol is considered an inexpensive, safe, and environmentally friendly method for synthesizing this recognizable organic motifs in both synthetic and medicinal chemistry in absence of mediators or even additional supportive electrolyte, as only electrons applied through electrical current are required for the reaction to proceed. Moreover, the continuous flow conditions, which fits into green chemistry standards, have demonstrated the feasibility of electrochemically synthesizing allyl sulfones, resulting in highly favorable outcomes and enabling the potential scalability of the electrochemical process.

Introduction

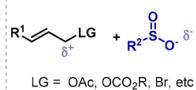
Sulfones represent functional groups with a long history in pharmaceutical development as they are prevalent moieties in biologically active molecules.¹ In particular, allyl sulfones could be identified as very relevant scaffolds due to their presence in pharmaceutically active derivatives such as^{2,3} protease inhibitors, fluorophores, the natural product ‘agelasidine’, or the antiinflammatory ‘resatorvid’ (Scheme 1A).¹ In addition, they are versatile and interesting building blocks in organic synthesis.^{4–6} Consequently, the synthesis of allyl sulfones have attracted the attention of organic and medicinal chemists over the years, turning the development of straightforward methodologies for the preparation of allylic sulfones a focal point in organic chemistry research.⁷

In this context, different strategies have been studied in the literature over the last decades. Although some impressive advancements in the synthesis of allyl sulfones have been

A. Allyl Sulfone derivatives as relevant scaffolds in pharmaceutically active compounds



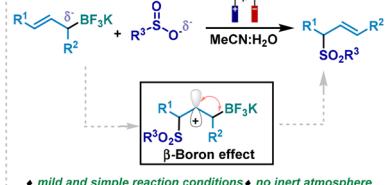
B. Traditional metal-catalyzed Tsuji-Trost allylation reaction



Reactivity

electrophile + nucleophile
polarity matched

C. This work. Electrochemical synthesis of allyl sulfones



nucleophile + nucleophile
polarity mismatched

REDOX INVERSION

radical + acceptor
polarity matched

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[†]Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3gc02408c>

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Scheme 1 Biologically active allyl sulfones (A), previous work (B) and this work (C).



made, such as hydrosulfination, metal-catalyzed SO_2 insertions, *etc.*, tedious protocols, low atom economy and applicability may hamper the general application of this methodologies.^{8–20} Venerable Tsuji–Trost-type coupling can be clearly identified as the more prominent synthetic tool for the direct and rapid sulfonation of allyl derivatives.^{21–24} A plethora of metal-catalyzed systems have been described for the coupling of sulfonyl containing nucleophiles with various electrophilic allyl derivatives (acetates, carbonates, halides, alcohols, *etc.*) for the synthesis of allyl sulfones through allylic substitution reactions (Scheme 1B).^{25–30} These approaches involved a diverted allyl derivatives possessing electrophilic features, while the sulfonate anion would act as nucleophile, which engages a matched polarity (right, Scheme 1B). Aiming to discover novel approaches for synthesizing relevant allyl sulfones, we envisioned the use of sulfinate and allyl boron derivatives as allylating agent.^{31,32} We thought that a radical approach would circumvent the reactivity mismatched otherwise encountered when facing two nucleophilic species if a two electron process take place (Scheme 1B). Among the multiple strategies based on the single-electron activation of organic substrates, electrochemistry is quickly becoming one of the most popular avenues to access radical intermediates, as it utilizes “green” electrons as reagents through the application of an electric current.^{33–37} This technology enables mild conditions for redox reactions and enhances atom economy. Our hypothesis suggests an electrochemical protocol to alter the initial reactivity of sulfinate and allyl boronates by promoting a radical process through the anodic oxidation of sulfinate,^{38,39} allowing to include the borate counterpart to the portfolio of starting materials for allyl-sulfone synthesis. Similarly to well-known β -silicon effect (also known as $\sigma(\text{C}-\text{Si})-\pi$ -hyperconjugation),⁴⁰ β -boron effect can stabilize carbocations through β - $\sigma(\text{C}-\text{B})$ interactions, imposing the regioselectivity of the process.⁴¹ The suitability of allyl trifluoroborate salts as allylating agents was then recognized,^{31,32} and this system would provide an alternative for synthesizing allyl sulfones by encountering two nucleophilic species. After the generation of a sulfur-based radical derivative from sulfinate, and its addition to the double bond of the allyl-boron derivative, a carbocation would be formed upon a second anodic oxidation.³⁸ A trifluoroborate elimination process would then take place, generating a double bond in the vicinal carbon, involving an allylation process overall, resembling a formal Tsuji–Trost type allylation reaction (Scheme 1C). Contrary to what has been traditionally assumed in two electron processes, cross-coupling reactions where a sulfur-based nucleophile faced an electrophile bearing a leaving group, this system would confront two nucleophilic species as an alternative for the synthesis of allyl sulfones (Scheme 1C). This electrochemical approach would circumvent issues associated with previously described catalytic systems for the synthesis of allyl sulfones, as for example the use of metals as catalysts and the reaction conditions normally associated with their use, such as inert atmosphere or dry solvents, high temperatures or long reaction times.^{21–24} Therefore, an electrochemical system for

the synthesis of allyl sulfones would be highly desirable, pursuing and satisfying one of the main motives for the renaissance of electrochemistry in organic synthesis which has appeared as green alternative to conventional synthetic methods. This electrochemical approach that targets the synthesis of relevant organic molecules does offer both operational and practical advantages that make it a greener methodology (Scheme 1C). This new method is considered an inexpensive, safe, and environmentally friendly, that takes place under very mild and simple reaction conditions: (i) No need of stoichiometric oxidants, as only electrons, applied through electrical current, and the two substrates are needed. (ii) Avoids the use of metals as catalysts or promoters. (iii) Circumvents the necessity of using inert atmosphere or dry solvents, high temperatures or long reaction times traditionally required for the synthesis of allyl sulfones. (iv) The system eludes the use of supportive electrolyte, employs carbon-based electrodes and uses a reasonable benign mixture of solvents (H_2O and CH_3CN). In addition, this electrochemical synthesis of allyl sulfones has demonstrated its feasibility and potential for scalability under electrochemical continuous flow conditions, that fits into green chemistry standards as it allows the development of safer protocols and improves energy requirement of the system as shorter reaction times are needed.

Results

Optimization of the model reaction

We began the study by choosing tosyl sulfinate (**1a**) as sulfonyl radical precursor and allyl trifluoroborate salt (**2a**) as allyl source (Table 1, entry 1). By using graphite electrodes in both anode and cathode under 2.5 mA constant current for 3 h in

Table 1 Optimization of the electrochemical sulfonylation of allyl trifluoroborate salts

		MeCN/H ₂ O (1:1, 3 mL) <i>C</i> _{graf} (+) <i>C</i> _{graf} (-) 2.5 mA, 3 h, r.t.	
0.1 mmol	2.5 equiv.		
1	No deviation		80
2	TBAPF ₆ (0.5 equiv.)		83
3	TBAPF ₆ (1.0 equiv.)		56
4	1 mA vs. 2.5 mA		48
5	5 mA vs. 2.5 mA		52
6	MeCN/H ₂ O 5:1 vs. 1:1		52
7	H ₂ O vs. MeCN/H ₂ O 1:1		—
8	Ni foam (—) vs. <i>C</i> _{graf} (—)		50
9	2a : 1.5 equiv. vs. 2.5 equiv.		46
10	No current		0

^a Reaction conditions: **1a** (0.1 mmol) and **2a** (0.25 mmol) at constant current (2.5 mA) for 3 hours in MeCN : H₂O (1 : 1, 3 mL) at r.t., under air, using *C*_{graf} as anode and cathode. ^b Isolated yields after purification by column chromatography.

an undivided cell and absence of an additional electrolyte in a mixture 1 : 1 of MeCN and H₂O, the allyl sulfone **3a** was isolated in very good yield (80%, entry 1). For the development of a synthetically useful electrochemical synthesis of allyl sulfones, various parameters such as solvent, electrolyte, electrode material and electrochemical parameters have been studied and summarized in Table 1.

Firstly, the addition of an external electrolyte (entries 2 and 3) was studied. While the presence of 0.5 equiv. did not significantly improve the yield of the desired allyl sulfone **3a**, 1 equiv. of the electrolyte greatly decreased the efficiency of the process. When the reaction was carried out at lower or higher current intensities (entries 4 and 5), allyl sulfone **3a** was obtained in lower yield. Different solvent systems (entries 6 and 7) were then evaluated. Thus, when MeCN was employed in a major proportion (entry 6), the electrochemical system performed substantially worse. However, only H₂O as sole solvent did not improve either, showing that MeCN as co-solvent was necessary to competently perform the alkylation (entry 7). The electrode material selection had an important impact on the reaction. Ni cathode led to the allyl sulfone in moderate yield (entry 8). We then tried to reduce the excess of the allyl trifluoroborane salt **2a** (entry 9). However, in presence of 1.5 equiv. of it (**2a**), the reaction performance was notably

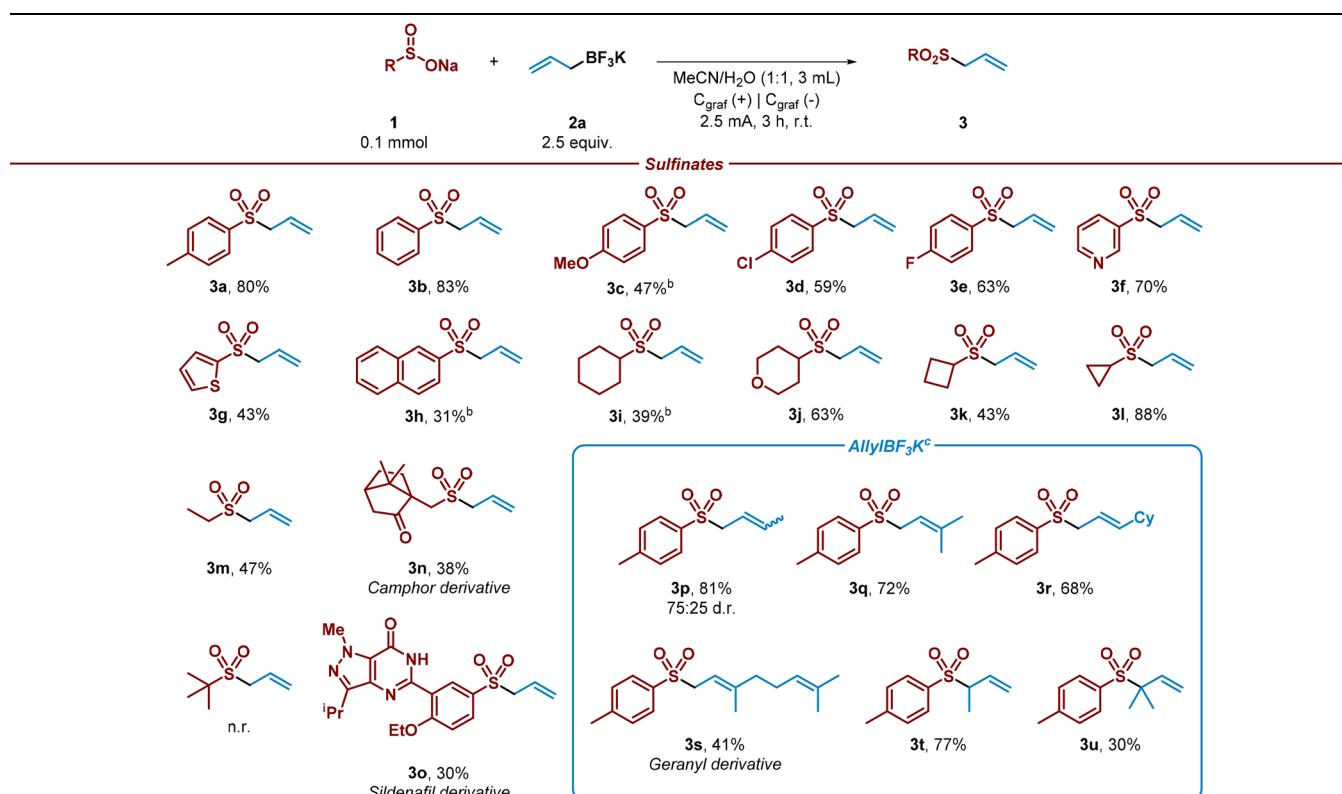
diminished as it plays a double role (reagent and electrolyte). Finally, the reaction did not take place in absence of electrical current (entry 10).

Substrate scope

Once the reaction conditions had been optimized, a variety of aryl and alkyl sulfinate precursors were tested under the electrochemical reaction conditions in an undivided cell with a readily available graphite anode/cathode, in absence of additional electrolyte and in a mixture of MeCN/H₂O as solvent system (Table 2). This study demonstrated that both aryl and alkyl sulfinate precursors efficiently performed the sulfonation, leading to the desired allyl sulfones in good yields (**3a**–**3o**). Phenyl sulfinate and aryl sulfinate decorated with electron-donating groups were well tolerated (**3a**–**3c**), although a diminished performance was observed with methoxy-bearing aryl sulfinate (**3c**). In addition, halides containing aryl sulfinate accomplished the desired reaction, leading to the sulfonated product in good yields (**3d** and **3e**). Notably, heterocyclic sulfinate showed decent performance under the optimized reaction conditions, conducting to the allyl sulfones (**3f** and **3g**).

2-Naphthyl substituted sulfinate was also subjected to the allylation reaction and led to the desired product with moderate yield (**3h**). Aliphatic sulfinate were then evaluated. To our

Table 2 Substrate scope of the electrochemical synthesis of allyl sulfones^a



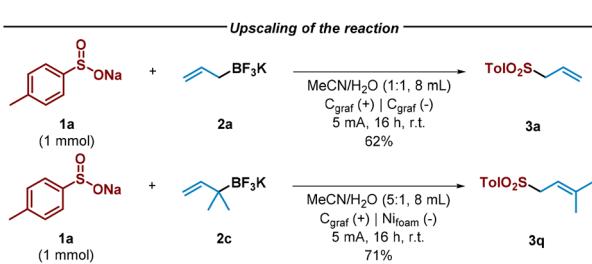
^a Reaction conditions: 1 (0.1 mmol) and 2 (0.25 mmol), CH₃CN : H₂O (1 : 1, 3 mL), r.t, under air, undivided cell (C_{graphite} anode, C_{graphite} cathode) at constant current (2.5 mA) for 3 h. Isolated yields. ^b Reaction conditions: 1 (0.1 mmol), 2 (0.25 mmol), TBAPF₆ (0.05 mmol), CH₃CN : H₂O (1 : 1, 3 mL), r.t, under air, undivided cell (C_{graphite} anode, C_{graphite} cathode) at constant current (2.5 mA) for 3 h. ^c Reaction conditions: 1 (0.1 mmol), 2 (0.25 mmol), CH₃CN : H₂O (5 : 1, 3 mL), r.t, under air, undivided cell (C_{graphite} anode, Ni foam cathode) at constant current (5 mA) for 2 h.



delight, secondary and primary sulfinites showed good performance, leading to the desired coupled product in good yields (**3i–3m**). Sterically encumbered *tert*-butyl sulfinate did not lead to the desired allylated product. In addition, as we wished to extend the applicability of the method, recognizable scaffolds decorated with a sulfinate moiety were also studied under the optimized electrochemical conditions. To our delight, camphor and sildenafil sulfinate derivatives provided the corresponding allylated products (**3n** and **3o**). The potential of the electrochemical method was further evaluated by testing various allylic trifluoroborate salts. Secondary trifluoroborate derivatives α -substituted with methyl and sterically hindered cyclohexyl group were very well tolerated (**3p** and **3r**). Remarkably, tertiary salts led efficiently to the corresponding allylic sulfones (**3s** and **3q**). To our delight, β -substituted (**3t**) and β -disubstituted (**3u**) trifluoroborate salts conducted to the corresponding allylated products successfully. It should be noted, that iso-prenyl sulfone **3u** has been utilized in the literature as ideal prenyl transfer reagent.⁴² This may help to overcome limitations in prenylation reactions, which is considered a challenging and essential reaction on which nature relies to modify properties of molecules and build terpenoids. In addition, and considering the electrons consumed and the isolated yields under the optimised reaction conditions, this process presents respectable faradaic efficiency (23–69%).⁴³

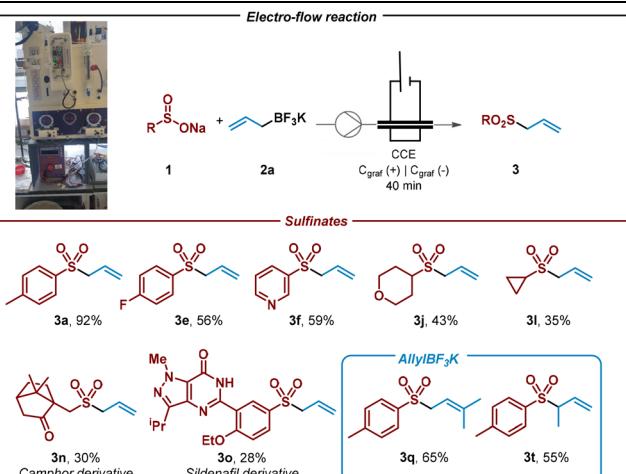
Notably, the reaction between sulfinate **1a** and allyl trifluoroborate salts **2a** and **2c** exhibited remarkable efficiency, even when scaled up by a factor of 10, resulting in the formation of desired allyl sulfones in good yield after 16 hours with slightly modified reaction conditions (Scheme 2).

Given the encouraging results achieved and the synthetic potential of the developed method, we decided to perform the electrochemical reaction under continuous flow conditions (Table 3, see ESI†). Flow chemistry fits into the current regimen that requires the development of new high safety, low waste generation, and energy efficiency processes. Gratifyingly, the reaction showed a remarkable enhancement in the efficiency of the process for allyl sulfone **3a** that could be isolated in 92% yield after 40 minutes. In addition, and to test the robustness of the method, the flow reaction was tested with aromatic, heteroaromatic and aliphatic sulfonates, including more sophisticated camphor and sildenafil sulfonate derivatives, providing the corresponding allylated products



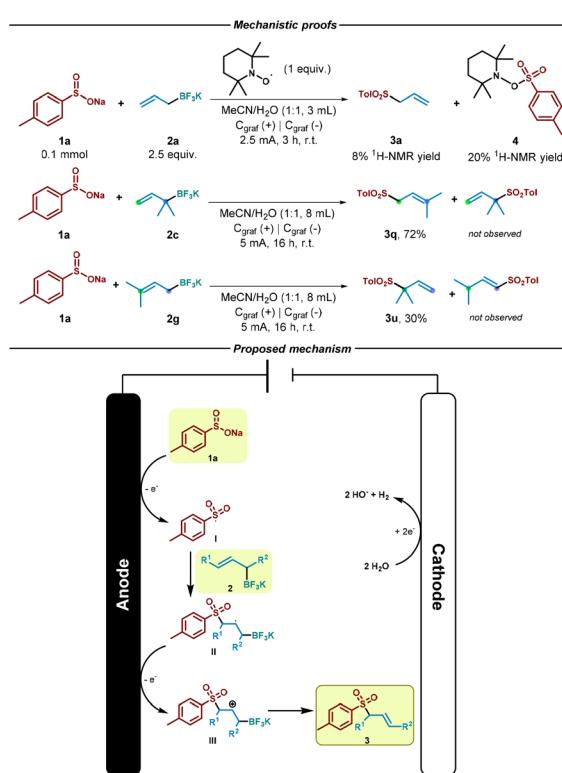
Scheme 2 Upscaling of the reaction and synthesis of allyl sulfone **3a** and **3g**.

Table 3 Substrate scope of the electrochemical synthesis of allyl sulphones under electro-flow conditions^a



^a Reaction conditions: **1** (0.1 mmol) and **2** (0.25 mmol), CH₃CN : H₂O (1 : 1, 4 mL), r.t, under air, undivided cell (C_{graphite} anode, C_{graphite} cathode) at constant current (2 mA). Isolated yields

with similar efficiency when compared with batch conditions (**3a**, **3e**, **3f**, **3j**, **3l**, **3n** and **3o**). In addition, tertiary and β -substituted trifluoroborate salts led efficiently to the corresponding allyl sulfones (**3q** and **3t**) under continuous flow conditions. It should be noted that the space–time yield calculated for the preparation of **3a** under flow reactions conditions is



Scheme 3 Control experiments and reaction mechanism.

44.7 g per (Lh), which leads to a more efficient process when compared with batch conditions (1.7 g per (Lh)).^{44,45}

We then performed a reaction in presence of TEMPO (2,2,6,6-tetramethylpiperidinoxy) as control experiment in order to gain an insight into the reaction mechanism. As shown in Scheme 3, in the presence of this radical scavenger, the reaction was almost inhibited as only 8% of the allyl sulfone was observed by NMR. In addition, the corresponding TEMPO adduct (**4**) resulted from the trapping event of the sulfonyl radical species could be identified (see ESI†). This result strongly indicate that radical species are involved in the reaction pathway. Based on this experiment, the literature and the previously showed results, we propose a mechanism for the electrochemical reaction (Scheme 3). At first, sulfinate (**1a**) gets oxidized ($E_{\text{ox}} = -0.37$ V vs. SCE, see ESI†)⁴⁶ at the anode, generating the sulfonyl radical (**I**). Then, addition of the sulfur-based radical to the alkene (**2**) generates secondary alkyl radical **II**, which undergoes anodic oxidation to generate the cation intermediate **III**.^{38,47–49} Finally, elimination of BF_3 as in a Sakurai-type allylation reaction⁵⁰ would generate the desired allylsulfone **3**. We believe that the reaction should follow an addition/oxidation sequence.

Conclusions

In summary, we have presented a versatile and straightforward electrochemical system for the synthesis of allyl sulfones using readily accessible sulfinate and allyl trifluoroborate salts. The reaction system has exhibited robustness and generality, accommodating aromatic, heteroaromatic, and aliphatic sulfonates. Considering these advantages, this protocol offers an economical, safe, and environmentally friendly approach to produce allyl sulfones, which serve as prominent organic motifs in both synthetic and medicinal chemistry. In addition, the synthesis of allyl sulfones has been efficiently implemented to an electrochemical continuous flow system.

Author contributions

L.M., R.R.-R., J.A.F.-S. and J.A. designed the experiments and analyzed the data. L.M., and R.R.-R. performed the experiments. R.R.-R and J.A.F.-S. described the ESI.† J.A.F.-S. and J.A. wrote the manuscript. All authors have given approval to the final version of the manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

Financial support was provided by the Spanish Government (PID2021-122299NB-I00, TED2021-130470B-I00, TED2021-

129999B-C32), European Structural Funds (S2018/NMT-4367), Proyectos Sinérgicos I + D (Y2020/NMT6469) and 'Comunidad Autónoma de Madrid' (SI1/PJI/2019-00237). J. A. F.-S. thanks the Spanish Government for a Ramón y Cajal contract.

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