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# Room-temperature fabrication of defective $\text{CoO}_x\text{H}_y$ nanosheets with abundant oxygen vacancies and high porosity as efficient 5-hydroxymethylfurfural oxidation electrocatalysts†

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Recently, cobalt oxides/hydroxides have attracted increasing attention in the electrocatalytic oxidation reaction of 5-hydroxymethylfurfural (HMFOR) under ambient conditions for 2,5-furandicarboxylic acid (FDCA) production, but understanding of the interplay of defective sites (*i.e.*, oxygen vacancies and porosity) remains lacking. Herein, a series of defective cobalt oxide hydrate ( $\text{CoO}_x\text{H}_y$ ) nanosheets were fabricated *via* room-temperature reductive treatments with methylamine (MA) and/or  $\text{NaBH}_4$  (BH). These defective  $\text{CoO}_x\text{H}_y$  nanosheets possessed abundant oxygen vacancies in relation to high  $\text{Co}^{2+}/\text{Co}^{3+}$  ratios and high porosity, with a largely maintained ultrathin lamellar framework, and thus exhibited markedly improved catalytic activity and selectivity for HMFOR. DFT calculations also verified the beneficial role of oxygen vacancies towards HMF adsorption and activation, preferentially *via* the aldehyde group of HMF. In the optimal  $\text{CoO}_x\text{H}_y$ -MA, 98% FDCA yield and a faradaic efficiency of 83% were achieved within 200 min at a constant potential of 1.52 V vs. RHE. The mesoporosity mainly induced by MA improved the mass transportation of reactants and products, leading to a higher rate of HMFOR. Meanwhile, with the additional in-sheet micropores mainly induced by BH, the selectivity towards the oxidation intermediate 5-formyl-2-furancarboxylic acid (FFCA) significantly increased, probably due to the accelerated penetration of FFCA instead of further oxidation to FDCA. This work highlights the simultaneous regulation of the oxygen vacancies and porosity of metal oxide/hydroxide catalysts by facile reductive treatments for efficient electrochemical biomass conversion.

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## 1. Introduction

In recent decades, the production of high value-added chemicals or biofuels from widely available biomass has become an important strategy to alleviate energy and environmental crises.<sup>1,2</sup> Thermochemical methods have been widely employed in the biomass refining process, but they often entail high temperature, high pressure, and additional redox agents.<sup>2</sup> Meanwhile, electrochemical methods, such as electro-

chemical oxidation processes, provide an efficient, economical, and clean means to realize the conversion of biomass-derived platform molecules under low temperature and atmospheric pressure conditions.<sup>2-4</sup> The reactivity of the biomass-derived platform molecules can be facilely regulated by the electrode potential and transformed by the *in situ* generated oxidizing agents.<sup>5,6</sup> Therefore, the electrocatalytic method has received intensive interest as an emerging and burgeoning approach for bio-refinery.<sup>3,7,8</sup>

5-Hydroxymethylfurfural (HMF) is one of the representative platform furanic molecules from biomass valorization.<sup>1</sup> The electrocatalytic oxidation reaction of HMF (HMFOR) is a promising strategy for producing high-value fine chemicals, including 2,5-furandicarboxylic acid (FDCA), 5-hydroxymethyl-2-furancarboxylic acid (HMFCA), 5-formyl-2-furancarboxylic acid (FFCA), and diformylfuran (DFF).<sup>9,10</sup> Among them, FDCA is a near-market alternative monomer to petroleum-derived terephthalic acid for the polymeric industry.<sup>11,12</sup> The electrocatalytic oxidation reaction of HMFOR under alkaline conditions to produce FDCA as the main product with a high far-

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daic efficiency (FE) is a challenging process.<sup>13,14</sup> Furthermore, the thermodynamic oxidation potential of HMF is 0.30 V, which is much lower than 1.23 V of the oxygen evolution reaction (OER),<sup>3,14</sup> making HMFOR an effective alternative anodic reaction to the OER, when constructing a hybrid water electrolysis system for the co-production of hydrogen at a low overall energy consumption.<sup>15</sup>

In recent years, Earth-abundant cobalt-based electrocatalysts, such as their sulfides (Co-CoS<sub>x</sub><sup>16</sup>), phosphides (CoP<sup>17</sup>) and borides (CoB<sup>18</sup>), have been proposed as highly efficient HMFOR electrocatalysts to produce FDCA.<sup>4</sup> These non-oxides are inclined to surface reconstruction into their oxide/hydroxide counterparts, which play a real role in HMF oxidation.<sup>19</sup> Luo *et al.* used an electrochemistry-assisted hydrolysis approach to transform NiCo-MOF into a NiCo-layered double hydroxide (LDH) structure<sup>20</sup> and Kang *et al.* similarly reported a CoOOH catalyst from an *in situ* electrochemical oxidation strategy, demonstrating the real active species for HMFOR with cobalt oxides/hydroxides.<sup>21</sup> However, the cobalt oxides/hydroxides still show low catalytic efficiency, due to the limited number of catalytic sites<sup>13</sup> and the poor conductivity.<sup>22</sup> Aside from strategies such as interfacial electronic engineering (Co-CoS<sub>x</sub>,<sup>16</sup> CoP-CoOOH<sup>23</sup>) and introducing multimetallic components (NiCoFe-LDH,<sup>22</sup> NiCo<sub>2</sub>O<sub>4</sub><sup>24</sup>), constructing mesoporous Co<sub>3</sub>O<sub>4</sub><sup>12</sup> with facile mass diffusion and oxygen vacancy-doped Co<sub>3</sub>O<sub>4</sub><sup>11,25,26</sup> has been proposed to increase the electrochemically active surface area and electrical conductivity, thus optimizing the adsorption behavior of intermediates and the catalytic performance of HMFOR.<sup>8</sup> It can be seen that these two types of defective sites, *i.e.*, oxygen vacancies and porosity play a key role in HMF performance. Nevertheless, the interplay of oxygen vacancies and porosity and their co-influence on the HMFOR process have not been studied previously.

Vacancy incorporation has been demonstrated to be one of the most promising strategies for modulating electronic structures.<sup>7,8</sup> Our previous work reported a topotactic conversion of two-dimensional (2D) cobalt oxide hydrate (CoO<sub>x</sub>H<sub>y</sub>) to surfactant-free and polycrystalline Co<sub>3</sub>O<sub>4</sub> nanosheets with enriched oxygen vacancies at grain boundaries *via* rapid calcination at 300 °C for 5 min to improve the HMFOR performance.<sup>10</sup> The temperature of 300 °C was chosen to remove the surfactant to expose the catalyst surface and simultaneously maintain the ultrathin 2D structure. Nevertheless, heat treatment still caused the loss of active surfaces with increased crystallinity. Here, we report a facile room-temperature fabrication strategy towards the synthesis of defective CoO<sub>x</sub>H<sub>y</sub> nanosheets with abundant oxygen vacancies and high porosity through reductive treatments with methylamine (MA) and/or NaBH<sub>4</sub> (BH) (Fig. 1a), and the surfactant can be substituted during this ligand-exchange process for efficient electron transfer.<sup>27</sup> Compared to the pristine CoO<sub>x</sub>H<sub>y</sub> nanosheets, lower onset potential and faster kinetics of HMFOR can be observed with these reductively-treated nanosheets. Density functional theory (DFT) calculations demonstrate the optimized adsorption of HMF intermediates with the presence of oxygen vacancies. Mesoporosity provided accelerated mass transpor-

tation, increasing the accessibility to active sites, whereas the incorporation of in-sheet micropores increased the selectivity of 5-formyl-2-furancarboxylic acid (FFCA), probably by the tuned diffusive transport of the reaction intermediates.

## 2. Experimental

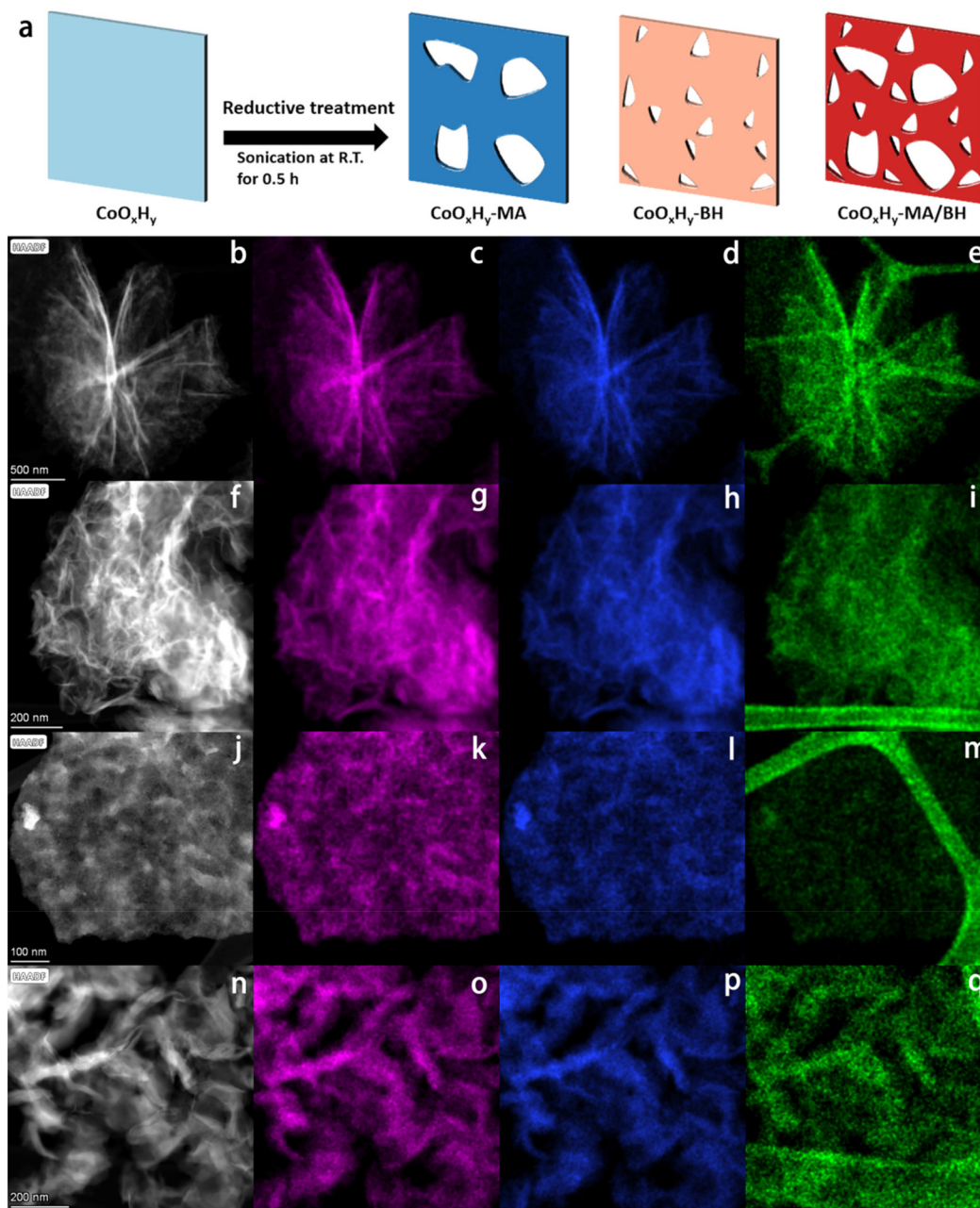
### 2.1 Preparation of catalysts

All chemicals with analytical grade were directly used without further purification. The CoO<sub>x</sub>H<sub>y</sub> nanosheets were synthesized by a simple template-assisted solvothermal method.<sup>10</sup> Concretely, 0.5 g P123 (EO<sub>20</sub>PO<sub>70</sub>EO<sub>20</sub>, *M<sub>w</sub>* = 5800 g mol<sup>-1</sup>) was dissolved in a mixed solvent composed of 7.5 g ethanol, 2.5 g H<sub>2</sub>O, and 30 mL ethylene glycol by sonication to form the template solution. 1.25 mmol of cobalt acetate tetrahydrate (Co(OOCH<sub>3</sub>)<sub>2</sub>·4H<sub>2</sub>O) and 1.25 mmol of hexamethylenetetramine (HMTA) were then added in the P123 solution under vigorous stirring for 30 min. The precursor solution was transferred into a 100 mL Teflon-lined stainless steel autoclave and heated at 170 °C for 2 h. After cooling to room temperature, the solid was collected by centrifugation, washed with ethanol and H<sub>2</sub>O, and finally dried at 60 °C overnight.

For room-temperature treatments, 50 mg of the as-synthesized CoO<sub>x</sub>H<sub>y</sub> nanosheets were dispersed in a 10 mL ethanol solution containing 1 mL of methylamine (MA) aqueous solution (40%). The molar ratio of MA to Co was approximately 20. After sonication for 0.5 h, the catalyst was obtained by washing with ethanol and drying at 60 °C overnight. Similarly, a freshly-prepared 10 mL ethanol solution containing 7.6 mg NaBH<sub>4</sub>, or methylamine/NaBH<sub>4</sub> (1 mL of methylamine and 7.6 mg NaBH<sub>4</sub>) was applied to the 50 mg as-prepared CoO<sub>x</sub>H<sub>y</sub> nanosheets under sonication for 0.5 h at room temperature, during which H<sub>2</sub> was released from NaBH<sub>4</sub> hydrolysis.<sup>28</sup> The catalysts were named CoO<sub>x</sub>H<sub>y</sub>-MA, CoO<sub>x</sub>H<sub>y</sub>-BH, and CoO<sub>x</sub>H<sub>y</sub>-MA/BH, respectively, according to the added reductant. Besides, Co<sub>3</sub>O<sub>4</sub>-300-MA, Co<sub>3</sub>O<sub>4</sub>-300-BH, and Co<sub>3</sub>O<sub>4</sub>-300-MA/BH were prepared in the same way except that Co<sub>3</sub>O<sub>4</sub>-300 nanosheets from the rapid calcination of CoO<sub>x</sub>H<sub>y</sub> nanosheets<sup>10</sup> were applied as the catalyst precursor.

### 2.2 Characterizations

The morphology and structure of the resulting materials were characterized by transmission electron microscopy (TEM, Titan Themis) at 200 kV, and the corresponding energy-dispersive X-ray spectroscopy (EDS) mapping was carried out using 4 in-column Super-X detectors. The specimen was prepared by ultrasonically suspending the sample in ethanol. A drop of the suspension was deposited on a carbon-coated copper grid and dried in air. The organic residue content was measured on a thermogravimetric analyzer at a heating rate of 10 °C min<sup>-1</sup> in air. The porous structure of the samples was characterized by nitrogen sorption at -196 °C on a Micromeritics ASAP 2020M analyzer. Prior to adsorption, the samples were outgassed at 80 °C overnight. The specific surface area (*S*<sub>BET</sub>) was calculated using the Brunauer-Emmett-Teller (BET) method. The total



**Fig. 1** Schematic illustration and TEM characterization of the pristine and reductively-treated  $\text{CoO}_x\text{H}_y$  nanosheets. (a) Schematics of the reductive treatments at room temperature for the formation of  $\text{CoO}_x\text{H}_y\text{-MA}$ ,  $\text{CoO}_x\text{H}_y\text{-BH}$ , and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$ . (b, f, j, n) The HAADF-STEM images, and the corresponding EDS mappings of (c–e)  $\text{CoO}_x\text{H}_y$ , (g–i)  $\text{CoO}_x\text{H}_y\text{-MA}$ , (k–m)  $\text{CoO}_x\text{H}_y\text{-BH}$ , and (o–q)  $\text{CoO}_x\text{H}_y\text{-MA/BH}$ . Co, O, and C elements are represented in magenta, blue, and green, respectively.

pore volume ( $V_{\text{tot}}$ ) was obtained according to the single point method at relative pressure  $P/P_0 = 0.99$ . The micropore surface area ( $S_{\text{micro}}$ ) and micropore volume ( $V_{\text{micro}}$ ) were determined by the  $t$ -plot method. The pore size distributions were estimated according to the non-local density functional theory (NLDFT) method. Zeta potential ( $\zeta$ ) was measured on a Zetasizer (Nano ZS) in an isopropanol–water solution. X-ray diffraction (XRD) analysis was performed by using an X'Pert-MPD diffractometer (PANalytical) using  $\text{Cu K}\alpha$  monochromatic

radiation ( $\lambda = 0.15418 \text{ nm}$ ) as an X-ray source. Fourier transform infrared spectroscopy (FTIR) was performed on a Thermo Scientific Nicolet iS50 spectrometer equipped with a MCT detector in the diffuse reflectance mode in the wavelength range of  $4000\text{--}400 \text{ cm}^{-1}$ . Raman spectra were collected using a Horiba spectrometer (LabRAM HR Evolution) and probed with a 532 nm laser. X-ray photoelectron spectra (XPS) were recorded on an ESCALab250Xi X-ray spectrometer with a monochromatic excitation source of  $\text{Al K}\alpha$  (1486.8 eV). All the

peaks were adjusted using the hydrocarbon C 1s peak at 284.8 eV from adventitious carbon as the binding energy reference. Electron paramagnetic resonance (EPR) measurements were performed on a Bruker A300 spectrometer at  $-196\text{ }^{\circ}\text{C}$ .

### 2.3 Electrochemical measurements

All the electrochemical measurements were performed in a standard three-electrode system on a CHI-760E electrochemical workstation. The working electrode was glassy carbon (3 mm in diameter) with the catalysts while the Pt plate ( $1 \times 1\text{ cm}^2$ ) and Hg/HgO electrode worked as the counter electrode and the reference electrode, respectively. Firstly, 3 mg of the catalyst was dispersed in 1.5 mL of isopropanol/water (1 : 3) solution and 10  $\mu\text{L}$  of Nafion (5%). After the mixture was well dispersed by sonication for 0.5 h, 6  $\mu\text{L}$  of the catalyst ink was dropped onto the clean glassy carbon electrode and dried in air at ambient temperature for a catalyst loading of  $0.167\text{ mg cm}^{-2}$ . Linear sweep voltammetry (LSV) measurements for HMF electrooxidation (HMFOR) and the OER were performed at a scanning rate of  $5\text{ mV s}^{-1}$  in 1 M KOH electrolyte solution with and without 5 mM HMF, respectively. The potential values were converted to potential *versus* the reversible hydrogen electrode (RHE) using eqn (1):

$$E_{\text{RHE}} = E_{\text{Hg/HgO}} + 0.059\text{ pH} + 0.098\text{ V.} \quad (1)$$

The Tafel plots were applied to assess the kinetics during the electrooxidation reactions, which were plotted as potential ( $E_{\text{RHE}}$ ) *vs.*  $\text{Log}(J/(\text{mA cm}^{-2}))$ . The electrochemically active surface area (ECSA) was estimated by the double-layer capacitance ( $C_{\text{dl}}$ ) *via* cyclic voltammetry (CV) in a non-faradaic potential region at scanning rates ranging from 10 to  $50\text{ mV s}^{-1}$  in 1 M KOH. The ECSA was calculated from the corresponding  $C_{\text{dl}}$ , based on eqn (2):

$$\text{ECSA} = C_{\text{dl}}/C_s. \quad (2)$$

The  $C_{\text{dl}}$  was calculated as the double-layer charging current divided by the scanning rate ( $C_{\text{dl}} = \Delta j/(2\nu)$ , where  $\Delta j = (j_a - j_c)$  and  $\nu$  is the scanning rate). The  $C_s$  is the specific capacitance, typically reported to be  $0.040\text{ mF cm}^{-2}$  in 1 M KOH.<sup>12</sup> The durability of the catalyst was evaluated by the chronoamperometric (CA) method in 1 M KOH at a constant potential of 1.52 V *vs.* RHE. The electrochemical impedance spectroscopy (EIS) tests were performed in the frequency range from  $10^5$  to 1 Hz with an AC voltage amplitude of 5 mV. The obtained Nyquist plots were fitted with a simplified Randles circuit. For reference, the electrocatalytic performance of commercial  $\text{IrO}_2$  was also checked under the same reaction conditions. It should be noted that no *iR* compensation was applied in all the above measurements.

To determine the HMF conversion, FDCA yield, and FE for FDCA, an H-type electrolyzer separated by a Nafion 117 membrane was applied for long-term HMF electrolysis. 200  $\mu\text{L}$  of the catalyst ink was transferred to a piece of carbon fiber cloth ( $1 \times 1\text{ cm}^2$ , 0.35 mm thick, HCP331N), dried in air, and employed as the working electrode. The catalyst mass loading

was  $0.40\text{ mg cm}^{-2}$ . A Pt plate ( $1 \times 1\text{ cm}^2$ ) and a Hg/HgO electrode were employed as the counter and reference electrodes, respectively. After passing a certain amount of charges, the sample was taken from the anode compartment for HPLC analysis. The HPLC was equipped with a Biorad 87H column ( $300\text{ mm} \times 7.8\text{ mm}$ ) and a refractive index detector (RID), using sulfuric acid (5 mM) as the mobile phase at a flow rate of  $0.60\text{ mL min}^{-1}$  and  $65\text{ }^{\circ}\text{C}$ . The concentrations of HMF, FDCA, DFF, HMFCa, and FFCA were calculated with respect to external standards.

Definitions of HMF conversion and product yield are as follows:

$$\text{HMF conversion} = \frac{\text{moles of HMF consumed}}{\text{moles of HMF loaded}} \times 100\% \quad (3)$$

$$\text{Product yield} = \frac{\text{moles of product formed}}{\text{moles of HMF loaded}} \times 100\%. \quad (4)$$

The definition of faradaic efficiency (FE) of product formation is as follows:

$$\text{FE} = \frac{\text{moles of product formed}}{\text{total charge passed}/(F \times n)} \times 100\% \quad (5)$$

where  $F$  is the Faraday constant ( $96485\text{ C mol}^{-1}$ ) and  $n$  is the electron transfer per mole of product.

### 2.4 DFT calculations

Perdew–Burke–Ernzerhof (PBE)<sup>29</sup> DFT calculations were performed for evaluating the adsorption behavior with the projector augmented wave (PAW)<sup>30</sup> on the Vienna *Ab initio* Simulation Package (VASP)<sup>31</sup> platform. PBE/PAW calculations were used to optimize the geometries and compute the energies of the adsorbate molecule, the surface adsorbent and the adsorption complex at a convergence energy threshold of  $10^{-5}$  eV. The vacuum space was set to be 15  $\text{\AA}$  along the  $z$  direction to avoid the interaction between the two neighboring cubic periodic boxes. The Hubbard  $U$  (DFT+ $U$ ) corrections were considered with a  $U$  value of 3 for the localized 3d orbitals of transition metal Co.<sup>32</sup>

The adsorption energy ( $E_{\text{ads}}$ ) of the adsorbate molecule is defined as eqn (6):

$$E_{\text{ads}} = E_{\text{complex}} - E_{\text{surface}} - E_{\text{molecule}} \quad (6)$$

where  $E_{\text{complex}}$ ,  $E_{\text{surface}}$ , and  $E_{\text{molecule}}$  are the energy of the adsorption complex, the surface adsorbent, and the adsorbate molecule, respectively.

## 3. Results and discussion

Ultrathin cobalt oxide hydrate ( $\text{CoO}_x\text{H}_y$ ) nanosheets were synthesized by a template-assisted solvothermal method, which have a flower-like layered assembly structure (Fig. 1b–e).<sup>10</sup> Due to the structure-directing by polymeric P123 inverse lamellar micelles,<sup>33</sup> the as-prepared  $\text{CoO}_x\text{H}_y$  nanosheets were encapsulated in organic residues of 31.2 wt% according to the TGA

results (Fig. S3, ESI†). Subsequently, room-temperature reductive treatments were used to fabricate defective  $\text{CoO}_x\text{H}_y$ -MA,  $\text{CoO}_x\text{H}_y$ -BH and  $\text{CoO}_x\text{H}_y$ -MA/BH by soaking the as-prepared  $\text{CoO}_x\text{H}_y$  in an ethanol solution of methylamine (MA) and/or  $\text{NaBH}_4$  (BH) and sonicating for 0.5 h. The polymeric organic residues significantly reduced after the reductive treatments by MA and/or BH, proving the simultaneous ligand exchange to expose the catalyst surface (Table 1, Fig. S3, ESI†).<sup>34</sup>

Scanning transmission electron microscopy coupled with energy-dispersive X-ray spectroscopy (STEM-EDS) mappings were used to verify the morphology and elemental distribution. As shown in Fig. 1f–q,  $\text{CoO}_x\text{H}_y$  in the reductively-treated samples formed different assembly structures, which is in line with the different surface charges, as verified by the different zeta potentials in Table 1, but remained in the ultrathin nanosheet framework structure. In addition, the elemental mappings of Co showed an isolated point-like distribution of deep contrast depth, which indicates the etching of  $\text{CoO}_x\text{H}_y$  by MA and/or BH. The dissolution reactions could follow the Hard Soft Acids Bases coordination theory<sup>27</sup> for the  $\text{Co}^{2+}$  and  $\text{Co}^{3+}$  sites in the as-prepared  $\text{CoO}_x\text{H}_y$ ;  $\text{Co}^{2+}$  was more inclined to dissolve than  $\text{Co}^{3+}$ . Due to the vulnerability of  $\text{CoO}_x\text{H}_y$  upon electron beam exposure,  $\text{Co}_3\text{O}_4$ -300 from the rapid calcination of the as-prepared  $\text{CoO}_x\text{H}_y$  at 300 °C for 5 min, and the catalysts after the reductive treatments of  $\text{Co}_3\text{O}_4$ -300 nanosheets were studied in parallel. Based on the STEM (Fig. S1†) and HRTEM images (Fig. S2†) of  $\text{Co}_3\text{O}_4$ -300,  $\text{Co}_3\text{O}_4$ -300-MA,  $\text{Co}_3\text{O}_4$ -300-BH, and  $\text{Co}_3\text{O}_4$ -300-MA/BH, MA and BH created nanodots by etching, where the lattice fringes were blurred, indicating the distortion of the lattice related to the oxygen vacancies.<sup>35</sup> MA was inclined to dissolve nanograin edges,<sup>36</sup> bringing in more mesopores, while BH led to more in-sheet micropores below 2 nm. These micropores should be formed during the addition of BH following hydrogen generation.<sup>28,37</sup>

The different etching effects on the porous structure by MA and BH were also evidenced by the  $\text{N}_2$  adsorption–desorption results in Fig. 2a, b, and Table 1. As shown in Fig. 2a, these catalysts all exhibited IV-type isotherms with an H1-type hysteresis loop, which confirms the accumulation of thin  $\text{CoO}_x\text{H}_y$

nanosheets with slit-shaped pores.<sup>10,12</sup> After the MA treatment, the micropore volume decreased while the mesopore volume increased. It is possible that the  $\text{CoO}_x\text{H}_y$  nanosheets in  $\text{CoO}_x\text{H}_y$ -MA could reduce their size and re-assemble. The large mesopore volume of  $0.77 \text{ cm}^3 \text{ g}^{-1}$  of  $\text{CoO}_x\text{H}_y$ -MA can facilitate the mass transport of the reactants and products in the HMFOR reaction. In contrast, the BH treatment caused a significant increase in micropores while decreasing the mesopore volume. The MA/BH treatment superimposed the effects of MA and BH, leading to a hierarchically porous  $\text{CoO}_x\text{H}_y$  nanosheet assembly.

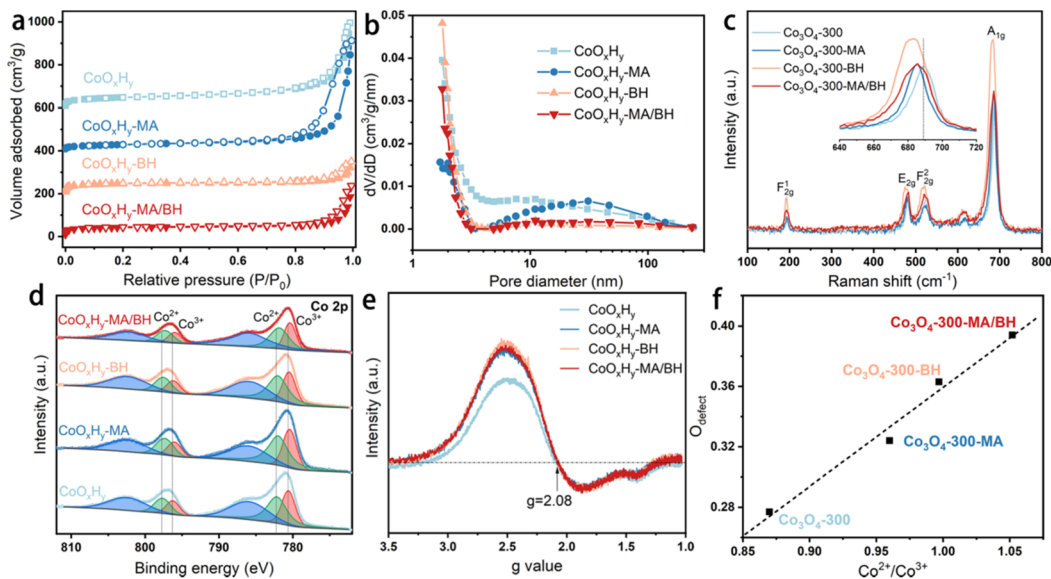
The crystal structure of the as-prepared  $\text{CoO}_x\text{H}_y$  nanosheets remained amorphous upon these reductive treatments, as confirmed by the XRD patterns (Fig. S4†). The functional groups were probed with FTIR spectroscopy (Fig. S5†).  $\text{CoO}_x\text{H}_y$ -MA showed the decrease of the bands at 3450 and 3345  $\text{cm}^{-1}$ , which were assigned to O–H, indicating the loss of surface OH after MA treatment.<sup>36</sup>  $\text{CoO}_x\text{H}_y$ -BH and  $\text{CoO}_x\text{H}_y$ -MA/BH showed relatively significant modifications in the shapes of the C–H bending peak at 1400  $\text{cm}^{-1}$  and the  $\text{Co}^{2+}$ –O stretching peak at 580  $\text{cm}^{-1}$ .<sup>38</sup> These results support the in-sheet etching of  $\text{CoO}_x\text{H}_y$  nanosheets by BH. The overall P123 residual-associated peaks were less intensive in these reductively-treated  $\text{CoO}_x\text{H}_y$  nanosheets, demonstrating the partial removal of organic residues *via* ligand exchange.

As for the investigation of the chemical constituents and oxygen defects, Raman spectroscopy, XPS, and EPR were performed.<sup>39</sup> Due to the vulnerability of  $\text{CoO}_x\text{H}_y$  nanosheets under the 532 nm laser,  $\text{Co}_3\text{O}_4$ -300 nanosheets before and after the reductive treatments were measured as a reference (Fig. 2c). The blue-shifts of the  $\text{A}_{1g}$  Raman peak induced by the reductive treatments indicate the presence of the defective structures of oxygen vacancies.<sup>39</sup> From the high-resolution Co 2p XPS spectra in Fig. 2d, the peaks at around 782.16 and 797.73 eV can be attributed to the  $\text{Co}^{2+}$  2p<sub>3/2</sub> and 2p<sub>1/2</sub> of  $\text{CoO}_x\text{H}_y$  nanosheets.<sup>9</sup> The other two peaks at about 780.59 and 796.28 eV belong to  $\text{Co}^{3+}$ . The binding energies of Co 2p were negatively shifted about 0.1–0.2 eV after the reductive treatments by MA and/or BH, revealing the modification in the electronic states of Co by MA and/or BH. In the O 1s XPS

**Table 1** The organic residual contents, textural properties, zeta potentials,  $\text{Co}^{2+}/\text{Co}^{3+}$  ratios, and the ECSAs of  $\text{CoO}_x\text{H}_y$ ,  $\text{CoO}_x\text{H}_y$ -MA,  $\text{CoO}_x\text{H}_y$ -BH, and  $\text{CoO}_x\text{H}_y$ -MA/BH

Catalyst	Organics <sup>a</sup> (wt%)	$S_{\text{BET}}^b$ ( $\text{m}^2 \text{ g}^{-1}$ )	$S_{\text{micro}}^c$ ( $\text{m}^2 \text{ g}^{-1}$ )	$V_{\text{tot}}^d$ ( $\text{cm}^3 \text{ g}^{-1}$ )	$V_{\text{meso}}^d$ ( $\text{cm}^3 \text{ g}^{-1}$ )	$V_{\text{micro}}^c$ ( $\text{cm}^3 \text{ g}^{-1}$ )	$D_{\text{pore}}^e$ (nm)	$\zeta^f$ (mV)	$\text{Co}^{2+}/\text{Co}^{3+g}$	ECSA <sup>h</sup>
$\text{CoO}_x\text{H}_y$	31.2	166	71	0.61	0.58	0.03	18.2	33	1.36	37.5
$\text{CoO}_x\text{H}_y$ -MA	21.0	101	40	0.79	0.77	0.02	32.8	30	1.53	65.5
$\text{CoO}_x\text{H}_y$ -BH	24.8	165	152	0.23	0.16	0.07	13.0	2	1.50	62.6
$\text{CoO}_x\text{H}_y$ -MA/BH	16.0	147	130	0.37	0.31	0.06	26.8	−7	1.48	73.5

<sup>a</sup> The organic residual content was determined by TGA as shown in Fig. S3.† <sup>b</sup> The surface area ( $S_{\text{BET}}$ ) was calculated by using the Brunauer–Emmett–Teller (BET) equation. <sup>c</sup> The area of micropores ( $S_{\text{micro}}$ ) and the micropore volume ( $V_{\text{micro}}$ ) were calculated by the *t*-plot method. <sup>d</sup> The total pore volume ( $V_{\text{tot}}$ ) was calculated from the saturation plateau at high relative pressures. The mesopore volume ( $V_{\text{meso}}$ ) was equal to  $V_{\text{tot}}$  minus  $V_{\text{micro}}$ . <sup>e</sup> The pore size ( $D_{\text{pore}}$ ) was calculated from the adsorption branch of the isotherms using the NLDFT model. <sup>f</sup> Zeta potential  $\zeta$  was measured in a dispersion of an isopropanol–water mixture. <sup>g</sup>  $\text{Co}^{2+}/\text{Co}^{3+}$  ratio was calculated from the peak deconvolution results of XPS Co 2p spectra in Fig. 2d. <sup>h</sup> ECSA value was obtained from the double-layer capacitance ( $C_{\text{dl}}$ ).

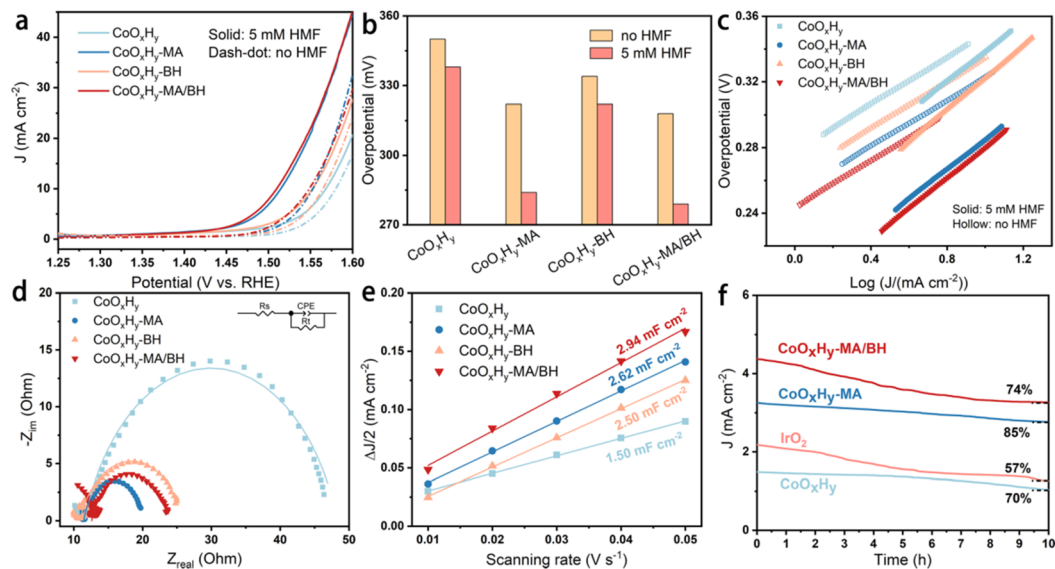


**Fig. 2** Structural characterization results of the pristine and reductively-treated  $\text{CoO}_x\text{H}_y$  and  $\text{Co}_3\text{O}_4$ -300 nanosheets. (a)  $\text{N}_2$  sorption isotherms and (b) the corresponding pore size distributions of  $\text{CoO}_x\text{H}_y$ ,  $\text{CoO}_x\text{H}_y$ -MA,  $\text{CoO}_x\text{H}_y$ -BH, and  $\text{CoO}_x\text{H}_y$ -MA/BH; (c) Raman spectra of  $\text{Co}_3\text{O}_4$ -300,  $\text{Co}_3\text{O}_4$ -300-MA,  $\text{Co}_3\text{O}_4$ -300-BH, and  $\text{Co}_3\text{O}_4$ -300-MA/BH; (d) XPS Co 2p spectra and (e) EPR signals against the  $g$  values of  $\text{CoO}_x\text{H}_y$ ,  $\text{CoO}_x\text{H}_y$ -MA,  $\text{CoO}_x\text{H}_y$ -BH, and  $\text{CoO}_x\text{H}_y$ -MA/BH; (f) The linear relationship between  $O_{\text{defect}}$  and  $\text{Co}^{2+}/\text{Co}^{3+}$  ratios, calculated from the XPS Co 2p and O 1s spectra of  $\text{Co}_3\text{O}_4$ -300,  $\text{Co}_3\text{O}_4$ -300-MA,  $\text{Co}_3\text{O}_4$ -300-BH, and  $\text{Co}_3\text{O}_4$ -300-MA/BH, as shown in Fig. S6.†

spectra (Fig. S6†), the O 1s peak can be deconvoluted into three characteristic peaks, which are located at 533.60, 531.21, and 529.72 eV, corresponding to adsorbed  $\text{H}_2\text{O}$ , O adsorbed in the defective sites ( $O_{\text{defect}}$ ), and lattice O,<sup>40</sup> respectively. As shown in Fig. 2f, the percentage of  $O_{\text{defect}}$  can be positively related to the  $\text{Co}^{2+}/\text{Co}^{3+}$  ratio, which agrees well with the previous observations.<sup>10,42,43</sup> Part of  $\text{Co}^{2+}$  was induced by an oxygen vacancy defective site to balance the charge, and the oxygen vacancy facilitated the reconstruction of a higher valence state of Co species such as  $\text{CoOOH}$  to improve the HMFOR performance.<sup>42,43</sup> The  $\text{Co}^{2+}/\text{Co}^{3+}$  ratios were 1.53, 1.50, and 1.48 for  $\text{CoO}_x\text{H}_y$ -MA,  $\text{CoO}_x\text{H}_y$ -BH, and  $\text{CoO}_x\text{H}_y$ -MA/BH, respectively, higher than 1.36 for the pristine  $\text{CoO}_x\text{H}_y$ . The EPR signals were contributed by the unpaired electrons from both  $\text{Co}^{2+}$  species and oxygen vacancies,<sup>10,39</sup> and the EPR lineshape of the transition metal is typically asymmetric arising from a mixture of absorption and dispersion effects.<sup>44,45</sup> A  $g$ -value of 2.08 was obtained (Fig. 2e) from the extreme points of the magnetic field strength diagram in Fig. S7.† The same  $g$ -values of these  $\text{CoO}_x\text{H}_y$  nanosheets evidenced similar signal sources from both  $\text{Co}^{2+}$  species and oxygen vacancies. However, the reductively-treated  $\text{CoO}_x\text{H}_y$  nanosheets exhibited stronger EPR signals compared to the pristine  $\text{CoO}_x\text{H}_y$ , confirming the higher  $\text{Co}^{2+}/\text{Co}^{3+}$  ratios and the correspondingly higher surface oxygen vacancies in these samples.<sup>10,43</sup> The above characterization results demonstrated that the treatment by MA and/or BH can create high porosity and abundant oxygen vacancies in  $\text{CoO}_x\text{H}_y$  nanosheets, which provides a promising opportunity for promoting the electrocatalytic activity. Additionally, the residual surface N contents were quite low, about 2 at% according to the N 1s XPS spectra in Fig. S8,† implying that

the electron-donating effect from  $\text{NH}_2$  coordination should not be a dominating factor for modifying the electronic properties of MA and MA/BH treated catalysts.

The electrocatalytic performance of the as-prepared and reductively-treated  $\text{CoO}_x\text{H}_y$  nanosheets for HMFOR was evaluated in 1 M KOH containing 5 mM HMF by LSV. As shown by the different HMFOR LSV plots in Fig. S10,† the catalyst performance was influenced by the catalyst loading. Therefore the catalyst loading on the glassy carbon electrode was controlled to be  $0.167 \text{ mg cm}^{-2}$  for the convenience of comparison. In the results presented in Fig. 3a–c, the pristine  $\text{CoO}_x\text{H}_y$  showed overpotentials of 338 and 350 mV to reach the current density of  $10 \text{ mA cm}^{-2}$  for HMFOR and OER activity, respectively. The MA and/or BH reductively-treated  $\text{CoO}_x\text{H}_y$  nanosheets exhibited better electrocatalytic oxidation performance, with more significant improvement in HMFOR than the OER. A similar phenomenon was also observed when comparing the electro-oxidation activities of  $\text{Co}_3\text{O}_4$ -300 and  $\text{CoO}_x\text{H}_y$ ; the more significantly improved HMFOR performance was ascribed to the increased accessible active sites and oxygen vacancies.<sup>10</sup> At  $10 \text{ mA cm}^{-2}$ , the overpotentials of  $\text{CoO}_x\text{H}_y$ -MA and  $\text{CoO}_x\text{H}_y$ -MA/BH were only 284 and 279 mV for HMFOR, respectively, much lower than those of the reported  $\text{CoOOH}$  thin film<sup>13</sup> and the pure Co LDH catalyst.<sup>22</sup> Moreover, the Tafel slopes of all  $\text{CoO}_x\text{H}_y$  nanosheets in HMFOR were about  $90 \text{ mV dec}^{-1}$ , larger than the values in the OER, indicating the more complex electron transfer kinetics for HMFOR.<sup>10</sup> The Tafel slopes agree well with the reported Co-based electrocatalysts with the partial formation of  $\text{Co}^{4+}$  as the catalytically active species.<sup>9</sup> The exchange current densities of  $\text{CoO}_x\text{H}_y$ -MA ( $6.7 \times 10^{-3} \text{ mA cm}^{-2}$ ),  $\text{CoO}_x\text{H}_y$ -BH ( $4.6 \times 10^{-3} \text{ mA cm}^{-2}$ ), and  $\text{CoO}_x\text{H}_y$ -MA/BH

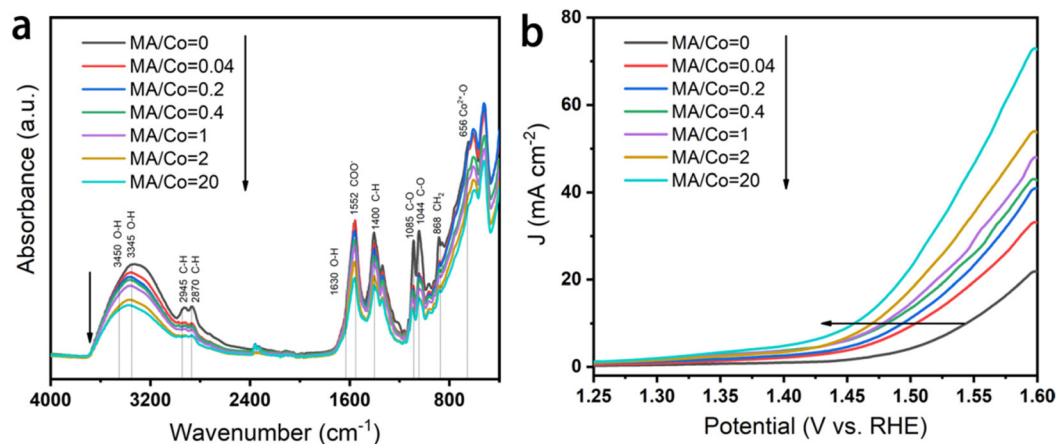


**Fig. 3** Electrochemical performance of  $\text{CoO}_x\text{H}_y$ ,  $\text{CoO}_x\text{H}_y\text{-MA}$ ,  $\text{CoO}_x\text{H}_y\text{-BH}$ , and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$ . (a) Polarization curves at a scanning rate of  $5 \text{ mA cm}^{-2}$  in  $1 \text{ M KOH}$  with and without  $5 \text{ mM HMF}$ , (b) overpotentials at a current density of  $10 \text{ mA cm}^{-2}$ , and (c) the corresponding Tafel plots; (d) EIS Nyquist plots at  $1.62 \text{ V vs. RHE}$ , (e) linear fitting of half of the difference between the oxidation and reduction current densities against the scanning rate, derived from CVs in Fig. S9,† and (f)  $i-t$  tests in  $1 \text{ M KOH}$  at  $1.52 \text{ V vs. RHE}$ .

( $9.6 \times 10^{-3} \text{ mA cm}^{-2}$ ) were higher than that of the pristine  $\text{CoO}_x\text{H}_y$  ( $1.8 \times 10^{-3} \text{ mA cm}^{-2}$ ), implying the higher intrinsic electron transfer rates at equilibrium<sup>23,40</sup> with these reductively-treated  $\text{CoO}_x\text{H}_y$  nanosheets. MA and/or BH created abundant defective active sites by ligand exchange and etching, which governed the superior HMFOR reactivity. As shown in Fig. 4, by increasing the MA/Co molar ratio from 0 to 20, the increased extent of ligand exchange and etching resulted in steadily improved HMFOR performance.

The origin of the remarkable HMFOR performance was further investigated by electrochemical impedance spectra (EIS) and the electrochemical surface area (ECSA). As shown in Fig. 3d and Table S1,† the charge transfer resistances ( $R_{ct}$ ) of

$\text{CoO}_x\text{H}_y\text{-MA}$  ( $8.1 \Omega$ ),  $\text{CoO}_x\text{H}_y\text{-BH}$  ( $15.1 \Omega$ ), and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$  ( $11.1 \Omega$ ) were smaller than that of  $\text{CoO}_x\text{H}_y$  ( $36.7 \Omega$ ), implying more rapid electronic transport in these reductively-treated  $\text{CoO}_x\text{H}_y$  nanosheets. The ECSA obtained by double-layer capacitances ( $C_{dl}$ ) showed the  $C_{dl}$  values of  $\text{CoO}_x\text{H}_y$  ( $1.50 \text{ mF cm}^{-2}$ ),  $\text{CoO}_x\text{H}_y\text{-MA}$  ( $2.62 \text{ mF cm}^{-2}$ ),  $\text{CoO}_x\text{H}_y\text{-BH}$  ( $2.50 \text{ mF cm}^{-2}$ ), and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$  ( $2.94 \text{ mF cm}^{-2}$ ), indicating higher densities of exposed catalytically active sites like oxygen vacancy-induced active sites<sup>26</sup> after these reductive treatments. At an overpotential of  $300 \text{ mV}$ , the  $\text{CoO}_x\text{H}_y$ ,  $\text{CoO}_x\text{H}_y\text{-MA}$ ,  $\text{CoO}_x\text{H}_y\text{-BH}$ , and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$  displayed current densities of  $0.10$ ,  $0.22$ ,  $0.09$ , and  $0.21 \text{ mA cm}^{-2}_{\text{ECSA}}$  for HMFOR (Fig. S11†), respectively. These results indicate that oxygen vacancies increase the



**Fig. 4** Influence of the MA/Co molar ratio on the electrocatalytic performance of  $\text{CoO}_x\text{H}_y$ . (a) FTIR spectra and (b) LSVs in  $1 \text{ M KOH}$  solution containing  $10 \text{ mM HMF}$  at a scanning rate of  $10 \text{ mV s}^{-1}$ .

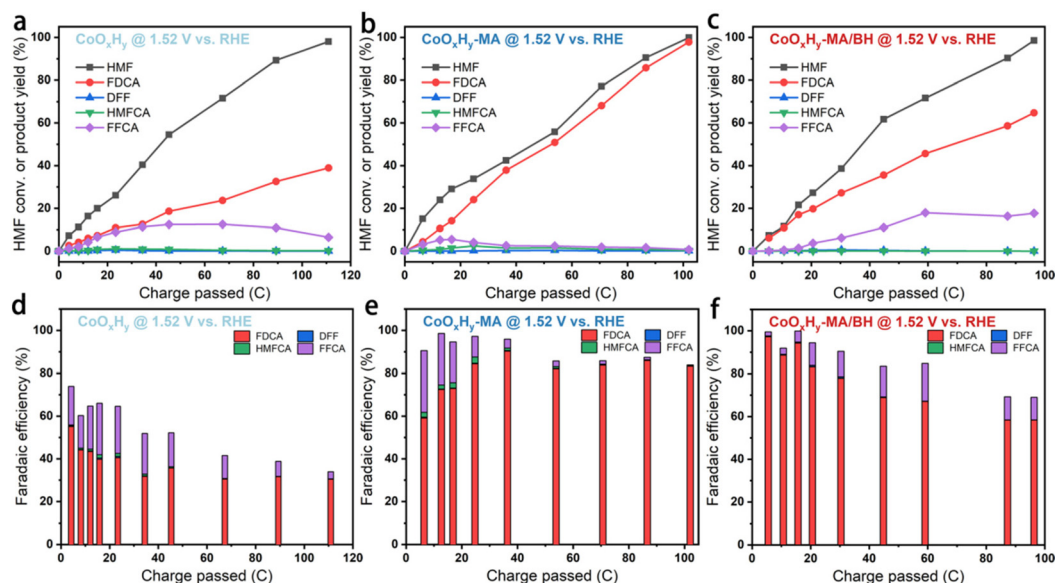


intrinsic HMFOR activities, and pore structures influence the catalytic performance by changing the mass transport. With the abundant oxygen vacancies and large mesopore volume in  $\text{CoO}_x\text{H}_y\text{-MA}$  and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$ , the mass transport of the reactants and products to the oxygen vacancy-induced active sites can be increased and the normalized HMFOR activity was enhanced. However, in  $\text{CoO}_x\text{H}_y\text{-BH}$ , although with abundant oxygen vacancies, the normalized HMFOR activity was even slightly lower than the pristine  $\text{CoO}_x\text{H}_y$ , indicating that micropores mainly induced by BH can significantly influence the mass transport during the HMFOR process. The electrocatalytic stability was evaluated by chronoamperometry at a constant potential of 1.52 V vs. RHE. The deactivation commonly occurs at a high anodic potential due to structural changes or detachment of redox-active species.<sup>41</sup> Nevertheless, after a 10 h chronoamperometric test, the current density decreased by 15% and 24% for  $\text{CoO}_x\text{H}_y\text{-MA}$  and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$ , respectively, indicating the limited change and decent stability of these defective  $\text{CoO}_x\text{H}_y$  nanosheets (Fig. 3f).

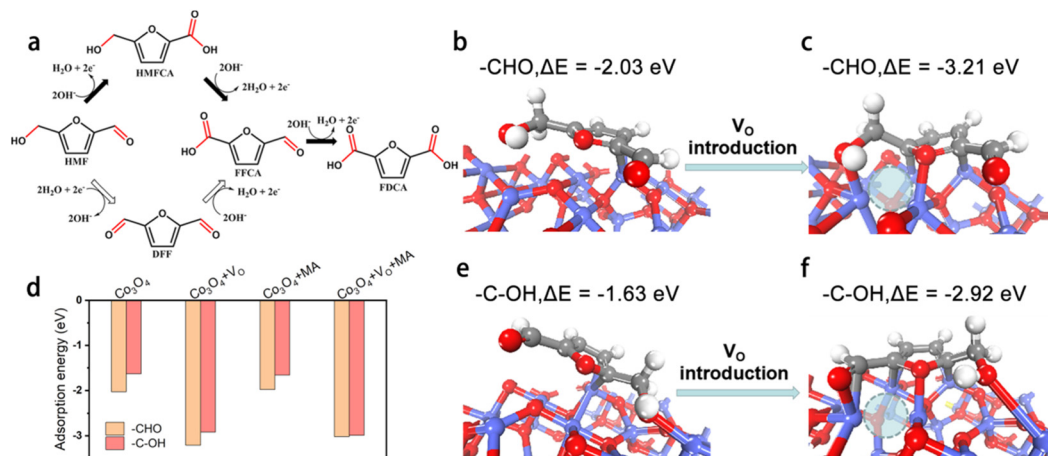
HMF electrolysis was also conducted in 1 M KOH containing 5 mM HMF at a constant potential of 1.52 V vs. RHE. The degradation of HMF to humin-type products in 1 M KOH is shown in Fig. S12.† Excessive degradation of HMF can be inhibited to a certain extent due to the high electrochemical reactivity under alkaline conditions.<sup>13</sup> Here, the thermodegradation of HMF was about 26% after 8 h. The catalytic activities and faradaic efficiencies (FEs) of the as-prepared  $\text{CoO}_x\text{H}_y$ ,  $\text{CoO}_x\text{H}_y\text{-MA}$ , and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$  towards electrochemical HMF oxidation were investigated and compared (Fig. 5). At 1.52 V vs. RHE, full conversion of HMF was realized in about 300 min for the pristine  $\text{CoO}_x\text{H}_y$ , while  $\text{CoO}_x\text{H}_y\text{-MA}$  and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$  required less than 200 min. The significantly

enhanced activities for HMFOR over  $\text{CoO}_x\text{H}_y\text{-MA}$  and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$  again demonstrate the vital role of the oxygen vacancies and the mesopores that provide high accessibility to reactants and products. The FE and yield for FDCA formation were up to 39% and 30% for the pristine  $\text{CoO}_x\text{H}_y$ , respectively, still much higher than those of the  $\text{CoOOH}$  thin film<sup>13</sup> and the pure Co LDH catalyst,<sup>22</sup> owing to its ultrathin nanosheet structure. However, compared to the reductively-treated  $\text{CoO}_x\text{H}_y$  nanosheets in  $\text{CoO}_x\text{H}_y\text{-MA}$  and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$ , the selectivity for FDCA was low due to the longer duration with more side degradation reactions of HMF and the completing oxygen evolution reaction.  $\text{CoO}_x\text{H}_y\text{-MA}$  can optimally achieve 98% FDCA yield and a FE of 83%, indicating the selective oxidation of HMF to FDCA with only minor losses due to decomposition and the OER. The room-temperature reductive treatment with MA can induce abundant oxygen vacancies and high mesoporosity in  $\text{CoO}_x\text{H}_y$  nanosheets through ligand exchange and etching, which offer highly accessible active sites for HMFOR. Varying the applied potential would decrease the FDCA yield and FE due to the increased side reactions of HMF degradation at 1.47 V or the OER at 1.57 V (Fig. S13†). On the other hand, with the additional in-sheet micropores in  $\text{CoO}_x\text{H}_y\text{-MA/BH}$ , the product distributions were distinct from those of  $\text{CoO}_x\text{H}_y\text{-MA}$ , affording FDCA and FFCA yields of 65% and 18%, respectively, at 99% conversion of HMF. It is thus speculated that these in-sheet micropores in  $\text{CoO}_x\text{H}_y\text{-MA/BH}$  may change the diffusion behavior of the intermediate product FFCA. An apparent proportion of FFCA could penetrate these micropores and desorb from the catalyst surface, avoiding being further oxidized.

Furthermore, DFT calculations were performed to establish a basic understanding of the HMFOR process over  $\text{CoO}_x\text{H}_y$



**Fig. 5** Results of HMFOR using  $\text{CoO}_x\text{H}_y$ ,  $\text{CoO}_x\text{H}_y\text{-MA}$ , and  $\text{CoO}_x\text{H}_y\text{-MA/BH}$  at a constant potential of 1.52 V vs. RHE in 1 M KOH containing 5 mM HMF, after passing various amounts of charges. (a, b and c) The conversion of HMF and the yields of oxidation products, and (d, e and f) the corresponding faradaic efficiencies of (a and d)  $\text{CoO}_x\text{H}_y$ , (b and e)  $\text{CoO}_x\text{H}_y\text{-MA}$ , and (c and f)  $\text{CoO}_x\text{H}_y\text{-MA/BH}$ , respectively.



**Fig. 6** Effect of surface oxygen vacancy ( $\text{V}_\text{O}$ ) sites and chemisorbed methylamine (MA) on HMFOR. (a) Two reaction pathways of HMFOR; HMF adsorption modes and energies on the  $\text{Co}_3\text{O}_4$  (3 1 1) surface (b, e) with no  $\text{V}_\text{O}$  and (c, f) in the presence of  $\text{V}_\text{O}$ .  $\text{V}_\text{O}$  is denoted with dashed circles; (d) adsorption energies of HMF on various surface sites, based on the adsorption configurations in Fig. S14.†

nanosheets and clarify the nature of the HMFOR performance enhancement by the MA and/or BH treatments. Since the adsorption of HMF on the active sites has been regarded as a rate-determining step of HMFOR,<sup>8,14,46</sup> the adsorption free energy changes ( $\Delta E$ ) of HMF adsorbed on the catalyst surface with different end groups (aldehyde  $-\text{CHO}$  group and the alcohol hydroxyl  $-\text{C-OH}$  group) were calculated (Fig. 6 and S12†). As shown in Fig. 6b and e, HMF was adsorbed on the  $\text{CoO}_x\text{H}_y$  nanosheets in two adsorption modes *via* the  $-\text{CHO}$  and  $-\text{C-OH}$  groups with a  $\Delta E$  of  $-2.03$  and  $-1.63$  eV, respectively. The introduction of an oxygen vacancy ( $\text{V}_\text{O}$ ) can evidently reduce the energetics to  $-3.21$  and  $-2.92$  eV, respectively, for the  $-\text{CHO}$  and  $-\text{C-OH}$  adsorption configurations (Fig. 6c and f). With the co-existence of MA in the neighbouring position, the energetics changed from  $-1.98$  to  $-3.02$  eV and from  $-1.66$  to  $-2.99$  eV after the introduction of  $\text{V}_\text{O}$ , for the adsorption modes of  $-\text{CHO}$  and  $-\text{C-OH}$ , respectively. These results reveal that the oxygen vacancies directly promote the adsorption and activation of HMF, especially for the  $-\text{CHO}$  adsorption mode, leading to the facilitated HMFOR process. The previously proposed functions of  $\text{V}_\text{O}$  like promoting  $\text{OH}^-$  adsorption,<sup>11</sup> increasing electronic conductivity,<sup>25</sup> and forming a higher valence state of Co species,<sup>42,43</sup> may still be present in the reductively-treated  $\text{CoO}_x\text{H}_y$  nanosheets and contribute to the improved intrinsic HMFOR activity. Meanwhile, the MA molecule appears to have little impact on the adsorption energies of HMF. Given the small amount of MA indicated by the XPS data, the coordination effect of residual MA in  $\text{CoO}_x\text{H}_y$ -MA and  $\text{CoO}_x\text{H}_y$ -MA/BH can therefore be neglected. The above computation results also show that the adsorption energies *via*  $-\text{CHO}$  were larger than *via*  $-\text{C-OH}$ . These energy differences indicate that the HMFOR preferentially follows the reaction pathway of forming HMFCFA as an intermediate instead of DFF, as shown in Fig. 6a, in accordance with the HPLC results that much more HMFCFA was formed than DFF during the electrolysis process. The main HMFCFA pathway on the  $\text{CoO}_x\text{H}_y$ ,

nanosheets agrees with most of the reported cobalt-based HMFOR electrocatalysts under strong alkaline conditions.<sup>11–14,25</sup>

## 4. Conclusions

In summary, we have successfully fabricated defective  $\text{CoO}_x\text{H}_y$  nanosheets with abundant oxygen vacancies and high porosity as an efficient HMFOR electrocatalyst based on a room-temperature reductive treatment with methylamine (MA) and/or  $\text{NaBH}_4$  (BH). These defective  $\text{CoO}_x\text{H}_y$  nanosheets in  $\text{CoO}_x\text{H}_y$ -MA and  $\text{CoO}_x\text{H}_y$ -MA/BH show remarkable activity with small overpotentials of 284 and 279 mV at a current density of 10  $\text{mA cm}^{-2}$  in 1 M KOH containing 5 mM HMF, respectively, nearly 60 mV lower than the that of the precursor  $\text{CoO}_x\text{H}_y$ . The improved catalytic performance should result from the larger electrochemically active surface areas and the faster electron transfer rates, owing to the enriched oxygen vacancies created by etching. The DFT calculations reveal the facilitated adsorption and activation with the incorporation of oxygen vacancies, preferentially through the aldehyde group. For long-term HMF electrolysis, 98% FDCA yield and a faradaic efficiency of 83% could be obtained on  $\text{CoO}_x\text{H}_y$ -MA with a large mesoporosity facilitating mass transport, whereas the additional in-sheet micropores in  $\text{CoO}_x\text{H}_y$ -MA/BH regulated the selectivity more towards the intermediate product FFCA. This facile reductive treatment with MA and/or BH can be a promising post-treatment strategy for developing cost-effective catalytic systems possessing tunable oxygen vacancies and porosity for electrochemical HMFOR.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

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## References

- P. Sudarsanam, R. Zhong, S. Van den Bosch, S. M. Coman, V. I. Parvulescu and B. F. Sels, *Chem. Soc. Rev.*, 2018, **47**, 8349–8402.
- Z. Zhang and G. W. Huber, *Chem. Soc. Rev.*, 2018, **47**, 1351–1390.
- D. Yan, C. Mebrahtu, S. Wang and R. Palkovits, *Angew. Chem., Int. Ed.*, 2023, **62**, e202214333.
- O. Simoska, Z. Rhodes, S. Weliwatte, J. R. Cabrera-Pardo, E. M. Gaffney, K. Lim and S. D. Minter, *ChemSusChem*, 2021, **14**, 1674–1686.
- M. T. Bender, Y. C. Lam, S. Hammes-Schiffer and K. Choi, *J. Am. Chem. Soc.*, 2020, **142**, 21538–21547.
- C. Chen, L. Wang, B. Zhu, Z. Zhou, S. I. El-Hout, J. Yang and J. Zhang, *J. Energy Chem.*, 2021, **54**, 528–554.
- G. Yang, Y. Jiao, H. Yan, C. Tian and H. Fu, *Small Struct.*, 2021, **2**, 2100095.
- C. Tang, Y. Zheng, M. Jaroniec and S. Qiao, *Angew. Chem., Int. Ed.*, 2021, **60**, 19572–19590.
- X. Deng, G. Y. Xu, Y. J. Zhang, L. Wang, J. Zhang, J. F. Li, X. Z. Fu and J. L. Luo, *Angew. Chem., Int. Ed.*, 2021, **60**, 20535–20542.
- R. Zhong, Q. Wang, L. Du, Y. Pu, S. Ye, M. Gu, Z. C. Zhang and L. Huang, *Appl. Surf. Sci.*, 2022, **584**, 152553.
- Y. Lu, T. Liu, C. L. Dong, C. Yang, L. Zhou, Y. C. Huang, Y. Li, B. Zhou, Y. Zou and S. Wang, *Adv. Mater.*, 2021, 2107185.
- C. Wang, H. J. Bongard, M. Yu and F. Schuth, *ChemSusChem*, 2021, **14**, 5199–5206.
- B. J. Taitt, D. Nam and K. Choi, *ACS Catal.*, 2019, **9**, 660–670.
- Y. Yang and T. Mu, *Green Chem.*, 2021, **23**, 4228–4254.
- B. You and Y. Sun, *Acc. Chem. Res.*, 2018, **51**, 1571–1580.
- J. Chen, Y. Wang, M. Zhou and Y. Li, *Chem. Sci.*, 2022, **13**, 4647–4653.
- N. Jiang, B. You, R. Boonstra, I. M. Terrero Rodriguez and Y. Sun, *ACS Energy Lett.*, 2016, **1**, 386–390.
- J. Weidner, S. Barwe, K. Sliozberg, S. Piontek, J. Masa, U. P. Apfel and W. Schuhmann, *Beilstein J. Org. Chem.*, 2018, **14**, 1436–1445.
- R. Luo, Y. Li, L. Xing, N. Wang, R. Zhong, Z. Qian, C. Du, G. Yin, Y. Wang and L. Du, *Appl. Catal., B*, 2022, **311**, 121357.
- X. Deng, M. Li, Y. Fan, L. Wang, X. Fu and J. Luo, *Appl. Catal., B*, 2020, **278**, 119339.
- R. Zhang, S. Jiang, Y. Rao, S. Chen, Q. Yue and Y. Kang, *Green Chem.*, 2021, **23**, 2525–2530.
- M. Zhang, Y. Liu, B. Liu, Z. Chen, H. Xu and K. Yan, *ACS Catal.*, 2020, **10**, 5179–5189.
- H. Wang, Y. Zhou and S. Tao, *Appl. Catal., B*, 2022, **315**, 121588.
- L. Gao, Y. Bao, S. Gan, Z. Sun, Z. Song, D. Han, F. Li and L. Niu, *ChemSusChem*, 2018, **11**, 2547–2553.
- X. Huang, J. Song, M. Hua, Z. Xie, S. Liu, T. Wu, G. Yang and B. Han, *Green Chem.*, 2020, **22**, 843–849.
- K. Gu, D. Wang, C. Xie, T. Wang, G. Huang, Y. Liu, Y. Zou, L. Tao and S. Wang, *Angew. Chem., Int. Ed.*, 2021, **60**, 20253–20258.
- M. A. Boles, D. Ling, T. Hyeon and D. V. Talapin, *Nat. Mater.*, 2016, **15**, 141–153.
- L. Yu, P. Pellechia and M. A. Matthews, *Int. J. Hydrogen Energy*, 2014, **39**, 442–448.
- J. P. Perdew, J. A. Chevary, S. H. Vosko, K. A. Jackson, M. R. Pederson, D. J. Singh and C. Fiolhais, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1992, **46**, 6671–6687.
- P. E. Blöchl, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1994, **50**, 17953–17979.
- G. Kresse and J. Furthmüller, *Comput. Mater. Sci.*, 1996, **6**, 15–50.
- L. Gong, D. Zhang, C. Y. Lin, Y. Zhu, Y. Shen, J. Zhang, X. Han, L. Zhang and Z. Xia, *Adv. Energy Mater.*, 2019, **9**, 1902625.
- Y. Dou, T. Liao, Z. Ma, D. Tian, Q. Liu, F. Xiao, Z. Sun, J. H. Kim and S. X. Dou, *Nano Energy*, 2016, **30**, 267–275.
- N. Naresh, F. G. S. Wasim, B. P. Ladewig and M. Neergat, *J. Mater. Chem. A*, 2013, **1**, 8553–8559.
- R. Gao, Z. Li, X. Zhang, J. Zhang, Z. Hu and X. Liu, *ACS Catal.*, 2015, **6**, 400–406.
- Y. H. Lee, S. Park, K. G. Lee, M. Y. Lee, K. H. Cho, S. J. Kim and K. T. Nam, *ChemCatChem*, 2019, **11**, 1665–1672.
- M. Liao, G. Zeng, T. Luo, Z. Jin, Y. Wang, X. Kou and D. Xiao, *Electrochim. Acta*, 2016, **194**, 59–66.
- Y. Cai, J. Xu, Y. Guo and J. Liu, *ACS Catal.*, 2019, **9**, 2558–2567.
- A. Badreldin, A. E. Abusrafa and A. Abdel-Wahab, *ChemSusChem*, 2021, **14**, 10–32.
- S. Bae, J. Kim, H. Randriamahazaka, S. Moon, J. Park and I. Oh, *Adv. Energy Mater.*, 2017, **7**, 1601492.

- 41 A. Martín, S. Mitchell, C. Mondelli, S. Jaydev and J. Perez-Ramírez, *Nat. Catal.*, 2022, **5**, 854–866.
- 42 M. Sun, Y. Wang, C. Sun, Y. Qi, J. Cheng, Y. Song and L. Zhang, *Chin. Chem. Lett.*, 2022, **33**, 385–389.
- 43 Z. Xiao, Y. Huang, C. Dong, C. Xie, Z. Liu, S. Du, W. Chen, D. Yan, L. Tao, Z. Shu, G. Zhang, H. Duan, Y. Wang, Y. Zou, R. Chen and S. Wang, *J. Am. Chem. Soc.*, 2020, **142**, 12087–12095.
- 44 F. Dyson, *Phys. Rev.*, 1955, **98**, 349–359.
- 45 R. Webb, *Phys. Rev.*, 1967, **158**, 225–233.
- 46 Y. Lu, T. Liu, C. L. Dong, Y. C. Huang, Y. Li, J. Chen, Y. Zou and S. Wang, *Adv. Mater.*, 2021, **33**, 2007056.