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Speeding up sustainable solution-phase peptide synthesis using T3P® as a green coupling reagent: methods and challenges†

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In peptide synthesis, the issues related to poor sustainability, long reaction times and high process mass intensity (PMI) are necessary to promote actions aimed at redefining procedural aspects projected towards more sustainable synthetic processes. Herein, we report a fast, widely applicable and green solution-phase peptide synthesis (GSolPPS) via a continuous protocol using propylphosphonic anhydride T3P® as the coupling reagent and *N*-benzyloxycarbonyl-protecting group (Z), which is easily removed by hydrogenation. Because *N,N*-dimethylformamide (DMF) replacement was a priority, the iterative process was performed in EtOAc, pushing further on overall sustainability. The efficiency of the synthetic protocol in terms of conversion, racemization and reaction times allowed extending the scope of the work to the synthesis of the standard peptide Leu-enkephalin as a proof of concept. Among the various explored procedures, the one-pot protocol (A_{cont plus}), avoiding work-ups, intermediate purification and any dispersion effect, allowed the achievement of PMI = 30 for each deprotection/coupling sequence necessary to introduce a single amino acid in the iterative process, without considering the possibility of solvent and base recovery. This value is the lowest reported for an oligopeptide synthesis protocol to date.

Peptide modality represents one of the emerging areas in drug discovery, as witnessed by the increasing number of peptide therapeutics approved by regulatory agencies¹ and entering different stages of clinical trial routes owing to their proven efficacy in interfering with upregulated or unwanted biological processes.²

The amide bond plays a key role in medicinal chemistry; therefore, various methods for its formation have been developed.³ Moving to peptide chemistry, the methodologies have

to be compatible with iterative synthesis, minimizing work-up and side reactions.

Therefore, the identification of a simple and sustainable iterative methodology that can induce the complete formation of the peptide bond in a short time at room temperature using easily removable *N*-protecting groups, affording oligopeptides with pharmaceutical grade purity, represents an important challenge for the sustainable production of this emerging class of therapeutics.

Short peptides (3–8 amino acids) used as cosmeceuticals owing to the scientific data, unlike cosmetics, support their role as active principles.⁴

After the identification and introduction of hormone peptides in the market in the early twentieth century, limited attention has been paid to the investigation of novel peptide drugs because of their difficult isolation from natural sources and limited bioavailability. Moreover, the short half-life due to enzyme degradation susceptibility strongly reduced the efficacy of peptides as therapeutics, requiring their administration via the intravenous route. The discovery of solid-phase synthetic technologies in the sixties⁵ and their optimization in the last six decades allowed us to overcome the previous bias and open access to the preparation of peptides on a large scale and to the introduction of non-coded residues, which can increase resistance to hydrolysis.⁶

Critical issues in peptide synthesis are poor greenness scores, such as the process mass intensity (PMI), and the common use of solvents and reagents, which are major concerns in large excess.

Dimethylformamide (DMF) is the best solvent for peptide synthesis, both in the solid- and solution-phase, because of its great performance in terms of the solubilization of reagents and by-products. It is reprotoxic and labelled as a hazardous solvent; accordingly, its replacement has been considered “advisable or requested” by the ACS Green Chemistry Institute® Pharmaceutical Round Table (GCIPR), which defined this goal as one of the hot topics in the list of the 12 green chemistry key research areas (KRAs).⁷ In addition,

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strong restrictions on its use in the industry will be applied from December 2023.⁸

Many efforts have been devoted in the last few years to greening solid-phase peptide synthesis (SPPS)⁹ by searching for alternatives to DMF.¹⁰

Moreover, piperidine,¹¹ the base of choice in Fmoc-protection removal in SPPS, is a substance under strict regulation used for the production of illegal narcotics,¹² and the need for a large excess of reagents and the amount of solvents required in washing steps to remove the coupling reagent's by-products represent weaknesses of this methodology.

With the introduction of SPPS, the traditional solution-phase approach was dismissed because a more complex impurity profile was generally obtained from the same synthetic sequence, making the isolation of intermediates by chromatography a key step to reaching acceptable purity for biomedical applications. Recently, liquid phase peptide synthesis (LPPS) has been reconsidered owing to the introduction of innovative technologies, which allow the precipitation of growing peptides or isolating them *via* membrane technology.¹³ Therefore, because of the great interest in reaching sustainable and efficient processes, LPPS recently received new attention in the synthesis of bioactive molecules, such as peptides¹⁴ and oligonucleotides.¹⁵ In particular, extensive improvements could arise from continuous approaches because few examples have been reported to date.^{9,16} An important requirement that would make solution-phase peptide synthesis (SolPPS) a much more "solid" technique to rely on, by combining efficiency and greenness score, would be the identification of greener effective coupling reagents,¹⁷ which can induce the complete formation of the peptide bond in a short time at room temperature.

Propylphosphonic anhydride (T3P®), first reported by Wissmann & Kleiner in 1980,¹⁸ has been documented in the literature to be a suitable reagent in the formation of peptide bonds in solution¹⁹ and several other transformations.²⁰

This reagent has never been used for iterative continuous processes in solution because of its high sensitivity to water. All the methodologies reported to date perform only a single coupling step at a time and cannot be applied to a continuous deprotection-coupling iterative process.²¹ Recently, Albericio *et al.*²² reported the use of T3P® in SPPS by screening various solvents and temperatures to confirm the versatility of this coupling system. Under these conditions, where the use of water is not needed, this coupling reagent was insufficient to have the complete formation of the peptide bond and had to be combined with Oxyma Pure to afford excellent coupling reagent couples, taking the role usually played by carbodiimides.

T3P® has several practical advantages over other coupling reagents. It is not inflammable, not toxic (LD50 in rats >2000 mg kg⁻¹)²³ and the side product propylphosphonic acid can be easily removed by water washings. Moreover, this reagent is actually commercialized in solution in several solvents, such as ethyl acetate and alkyl carbonates, with a stability of 2 years.²⁴

In addition to the above reported characteristics of T3P®, its effect on preventing racemization in peptide bond for-

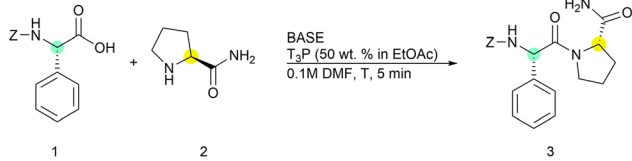
mation has been particularly addressed. All the reported applications displayed some drawbacks caused by long reaction times, low yields, and the formation of by-products. More recently, Dunetz *et al.*²⁵ confirmed the prevention of racemization in T3P®-mediated amide formation, selecting pyridine as the most suitable base.

We explored, herein, the possibility of applying this reagent to continuous solution-phase synthesis, eliminating the need for other coupling reagents and carbodiimides, which display several disadvantages in terms of toxicity and safety,²⁶ to develop a fast and green protocol for oligopeptide synthesis.

Consequently, we needed to avoid intermediate isolation, in particular with water involving work-ups, both in the coupling and protecting group removal steps. We identified the *N*-benzyloxycarbonyl (Z) group as the most suitable protection for the possibility of simple removal by hydrogenation without generating by-products that could affect the following steps. Recently, Lipshutz *et al.*²⁷ reported a single Z-deprotection/coupling 2-steps continuous reaction performed in water using oligopeptide fragments, adding a surfactant to create an efficient nano-micellar environment. From the above cited methodologies, any attempt to extend this procedure to a deprotection/coupling iterative process did not succeed.

First, we optimized the general conditions for solution-phase peptide synthesis using a standard reaction of the coupling between (L)-phenylglycine and (L)-prolinamide in DMF because this peptide bond is considered difficult and prone to racemization²⁸ (Table 1).

Table 1 Screening of conditions for the peptide coupling step



Entry ^a	T3P® (equiv.)	Base	Greenness score ^b	Conversion ^c (%)
		(Equiv.)		
1	1	DIPEA (1)	6.5	81
2	1	DIPEA (2)	6.5	88
3	1.5	DIPEA (3)	6.5	>99
4 ^d	1.5	DIPEA (3)	6.5	93
5 ^e	1.5	DIPEA (3)	6.5	80
6	1.5	2,6-Lutidine (3)	8.3	56
7	1.5	Pyridine (3)	7.5	42
8	1.5	TEA (3)	6.9	94
9	1.5	NMM (3)	6.9	90
10	1.5	^t BuNH ₂ (3)	6.5	93

^a Reactions were performed by dissolving Z-Phg in DMF (0.1 M conc) under nitrogen atmosphere and adding reagents to the following order: Pro-NH₂, DIPEA and finally T3P®. ^b According to the GSK base selection guide, see ref. 29. ^c Conversion was determined by LC-MS (see ESI for details†). ^d Reaction was performed at 0 °C. ^e Pre-activation of the acid was performed by adding T3P® and DIPEA before Pro-NH₂. The presence of Z-(D)-Phg-Pro-NH₂ was detected in a 5% amount.



The coupling reaction was performed under nitrogen atmosphere to reduce the presence of air moisture, which could eventually degrade T3P®. The selected results are reported in Table 1.

The attention was initially focused on the T3P®/base ratio and on the order of the addition of reagents to the vessel. To our surprise, the coupling reaction turned out to be very fast and afforded excellent conversions after only five minutes, supporting the idea of exploiting this protocol for iterative solution-phase peptide synthesis.

Moreover, when incomplete conversion was observed, little increase in dipeptide **1** formation occurred even after longer reaction times, suggesting that the acid activation by T3P® is almost instantaneous as well as the nucleophilic attack of the free amino ester.

Aiming at setting up the greenest possible protocol, we started adding only one equivalent for each reagent (entry 1), choosing DIPEA as the base. However, incomplete formation of the dipeptide was observed. With the only exception of entry 5, the base and T3P® were sequentially added to the mixture of the two amino acids. An increase in the conversion was observed by adding two equivalents of the base to neutralize both triphosphate salt functions (entries 2 and 3). The best results were obtained using a 1.5/3 ratio of the two reagents by achieving complete conversion in 5 minutes at room temperature (entries 3). The presence of 0.5% and 0.4% of Z-(D)-Phg-Pro-NH₂ was observed when the reaction was performed at rt and 0 °C, respectively (entries 3 and 4, see ESI for details†). Finally, when the preactivation of the acid was performed by adding T3P® and DIPEA before Pro-NH₂, only 80% conversion could be obtained, probably for the lability of the activated mixed anhydride (entry 5). Under these conditions, a 5% amount of racemized product was also detected (see ESI for details†). Having established optimized conditions, we screened a small library of bases and selected among those reported to be compatible with T3P® and those having a good greenness score according to the GSK selection guide. Although 2,6-lutidine and pyridine did not afford satisfactory results (entries 6 and 7) when the reaction was performed with triethylamine (TEA), *N*-methylmorpholine (NMM) or *t*-BuNH₂, peptide **3** was isolated with better yields (entries 8–10) even if slightly lower compared to DIPEA. This was confirmed to be the best base for obtaining a trade-off between greenness and peptide formation.

Concerning the solvent of choice, we then focused on replacing DMF with greener solvents²⁹ under the above identified optimized conditions. The selected results are reported in Table 2.

The screening confirmed the compatibility of T3P® with almost all the tested solvents even if the reactions that were performed in *N*-alkyl-pyrrolidones, such as *N*-octylpyrrolidone (NOP) and *N*-butylpyrrolidone (NBP), afforded limited conversions (entries 1 and 2) because of more difficult dissolution of activated mixed anhydride in low polar lipophilic solvents. Moreover, for all the solvents, the degree of racemization was very low if compared, for instance, with the commonly used

Table 2 Solvent screening and degree of racemization evaluation

Entry ^a	Solvent	Coupling reagents	Temperature	3 (%)	(D)/(L + D) (%)
1	NBP	T3P®	r.t.	56	0.6
2	NOP	T3P®	r.t.	56	n.d. ^b
3	DMC	T3P®	r.t.	93	0.2
4	GVL	T3P®	r.t.	89	0.7
5	ACN	T3P®	r.t.	96	0.2
6	THF	T3P®	r.t.	93	0.1
7	2-MeTHF	T3P®	r.t.	80	0.3
8	DCM	T3P®	r.t.	98	0.2
9	iPrOAc	T3P®	r.t.	69	1.1
10	EtOAc	T3P®	r.t.	94	0.5
11	EtOAc	T3P®	0 °C to r.t.	96	0.3
12	DMF	T3P®	r.t.	>99	0.5
13	DMF	T3P®	0 °C to r.t.	93	0.4
14	DMF	Oxyma pure/DIC	r.t.	73 ^c	1.0
15	DMF	Oxyma pure/DIC	0 °C to r.t.	70 ^c	1.0
16	EtOAc	Oxyma pure/DIC	r.t.	96	1.0
17	EtOAc	Oxyma pure/DIC	0 °C to r.t.	86	0.9

^a The reactions were performed under the conditions used in entry 3, Table 1, and conversion evaluated after 5 minutes. ^b n.d. = not detected. ^c Conversion was 90–93% after 1 h.

coupling system Oxyma Pure/DIC (entries 10 and 11 vs. entries 16 and 17 or entries 12 and 13 vs. entries 14 and 15).

Among the solvents with the best green score, EtOAc, THF and ACN exhibited the highest conversions. However, EtOAc has a lower affinity for water. This aspect needs to be particularly underlined considering the sensitivity to water displayed by T3P®. The reaction in EtOAc was also performed on a gram scale, affording excellent results (see ESI†).

EtOAc was selected as the solvent of choice to proceed with the investigation because, in addition to its technical characteristics, it is readily available at a competitive price with respect to DMF.

The exceptional speed of the T3P® that promoted the coupling reaction prompted us to apply the optimized protocol to the difficult reactions of the sterically hindered Aib amino acid (Table 3).

Thus, Z-Aib was activated under previously optimized conditions and added to prolinamide (entry 1) or phenylalanine ester (entry 2).

Table 3 Coupling reactions involving hindered Aib amino acid

Entry	H ₂ N-AA-OPg	Z-AA-OH	T3P®/DIPEA (equiv.)	C ^a (%)
1	H ₂ N-Pro-NH ₂	Z-Aib	2/4	92
2	H ₂ N-Phe-O ^t Bu	Z-Aib	2/4	96
3	H ₂ N-FL-O ^t Bu	Z-Aib	2.5/5	98
4	H ₂ N-Aib-OMe	Z-Phe	2/4	78
5	H ₂ N-Aib-OMe	Z-Aib	2/4	4
6	H ₂ N-Aib-OMe (+H ₂ N-L-O ^t Bu)	Z-Aib	2/4	4 (92) ^b

^a Conversion calculated by HPLC. ^b By adding H₂N-L-O^tBu after 5 minutes, the conversion of the remaining activated Z-Aib afforded 92% of Z-Aib-L O^tBu dipeptide was detected.



In both cases, the coupling gave satisfactory results using a slight excess of the coupling mixture (2/4 equivalents), affording the dipeptides in 92–96% conversion.

These results are of remarkable interest considering that in SPPS, the introduction of the Aib amino acid generally requires a double coupling step.³⁰

An even better conversion was obtained in the reaction of Z-Aib with the dipeptide H-Phe-Leu-O^tBu **5** (98%, entry 3). On the contrary, when Aib was used as the nucleophilic amino acid and it reacted with activated Z-Phe, lower conversion was observed (78%, entry 4), suggesting that the hindrance of the triphosphate chain does not allow the approach of the bulky Aib nucleophile.

The coupling system turned out to be completely ineffective in the Aib–Aib dipeptide formation because only 4% conversion was obtained (entry 5).

However, when H₂N-Leu-O^tBu was added to the reaction mixture after 5 minutes, a mixture of 4% of Aib–Aib and 96% of Leu–Aib was obtained, showing that the low reactivity was not caused by difficult Aib activation but by only the clash between the hindering moieties when Aib is also the nucleophile (entry 6).

With the target of performing the deprotection/coupling iterative process, the Z-removal step was explored *via* hydrogenation in EtOAc as well (Table 4). To this purpose, we set up two different protocols, (i) performing the reaction in a vessel using Pd/C_{10%} catalyst powder at atmospheric pressure (method A) and (ii) with the H-Cube® flow system,³¹ that exploit a reusable cartridge containing the Pd/C_{10%} heterogeneous catalyst (method B). Method B has the advantage of *in situ* generating endogenous on-demand hydrogen from water, further increasing the greenness score of the synthetic

process. The optimal conditions were identified for both systems, as reported in Table 3, by performing Z-group removal from dipeptide Z-F-L-O^tBu **4**, which is the initial fragment of the pentapeptide Leu-enkephalin (H-YGGFL-OH), commonly used to test the efficiency of a new protocol for peptide synthesis.

To mimic continuous processes, we removed the amine protective group in the presence of possible contaminants derived from a previous coupling step through their direct addition.

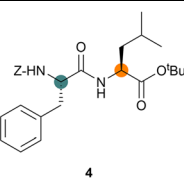
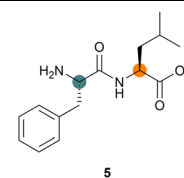
The mild conditions applied in method A and the use of Pd/C_{10%} catalyst, without any pre-treatment, allowed us to obtain the complete conversion of **4** into **5** in five hours (entry 1). Using a dehydrated catalyst, the reaction time decreased to 3 hours (entry 2). The presence of DIPEA in large excess afforded a much more reactive system (entry 4) but even the addition of 0.1 equivalent of base allowed to completely remove the Z protecting group in one hour (entry 3). On the contrary, when the triphosphate salt derived from T3P® treatment with water was added to the mixture, simulating the reacted coupling reagent, the transformation was strongly decreased (entry 5).

By adding both DIPEA and T3P®, in the same amount and ratio used in the optimized coupling step, the reaction slightly accelerated compared with the one performed on pure **4** (entry 6 *vs.* entry 2). The reaction was also tested using the hydrochloride salt of DIPEA to verify the effect of the co-products derived from a coupling step involving the use of amino ester hydrochloride. In this case, a decrease in reaction speed was observed, and complete conversion was achieved in four hours. The free tertiary amine could speed up the reaction by favoring the displacement of the final product from the catalyst surface.³² To perform hydrogenation in the H-Cube® flow system (method B), we had to apply stronger conditions to increase speed and decrease time. The model reaction was performed on 4 mL of 0.1 M solution of **4** in EtOAc. The content of the Pd/C_{10%} catalyst charged in the cartridge was about twice the amount used in the reaction in the batch. However, the flow system allows for a high catalyst/substrate local concentration in the hydrogenation cell (roughly 2/1 ratio) and transforms high quantities of flowing substrate, thereby decreasing the impact of the catalyst. Excellent and rapid conversion was achieved under H₂ pressure (15–20 bar) at 60 °C with a flow of 1 mL min^{−1}, collecting compound **5** in 18 minutes (entry 8), but completeness was achieved after the addition of DIPEA to the flowing solution (entries 9 and 10). The effect of additives was similar to the results observed before (entry 11).

Based on the previous information, we decided to perform the solution-phase continuous synthesis (GSolPPS) of Leu-Enkephalin by testing both hydrogenation methods in the removal of the Z-protecting group (Fig. 1 and Table 5).

We first matched the coupling steps with DIPEA and T3P® with the “in flow” deprotection through the H-Cube® system (method B_{cont}). We performed the iterative process without isolating the intermediate peptides but checked conversion and crude composition by HPLC analysis for each step. The first

Table 4 Screening conditions for Z-F-L-O^tBu **4** deprotection

<div style="display: flex; align-items: center; justify-content: space-around;"> <div style="text-align: center;">  <p>4</p> </div> <div style="text-align: center;"> <p>Method A: Pd/C_{10%} (10% w/w) H₂ (1.013 bar) 0.1M EtOAc, 30 °C</p> <p>Method B: H-Cube cartridge Pd/C_{10%} H₂ (15 bar) 0.1M EtOAc, 60 °C, Cflow 1mL/min</p> </div> <div style="text-align: center;">  <p>5</p> </div> </div>				
Entry	Contaminant (equiv.)	Method	Time (min)	5 (%)
1	—	A ^a	300	>99
2	—	A	180	>99
3	DIPEA (0.1 eq.)	A	60	>99
4	DIPEA (3 eq.)	A	30	>99
5	T3P® : H ₂ O (1.5 eq.)	A	240	>99
6	DIPEA : T3P® (1.5/3)	A	120	>99
7	DIPEA : HCl (3 eq.)	A	240	>99
8	—	B ^b	18	>99
9	DIPEA (0.1 eq.)	B	16	>99
10	DIPEA (1 eq.)	B	18	>99
11	DIPEA : T3P® (1.5/3)	B	18	>99

^a Reaction performed with Pd/C_{10%} (10% w/w) not previously dried under vacuum. ^b Injection of 4 mL of 0.1 M sample. The H-Cube® setup generated 6 mL of death volume solvent and a complete recovery of **5** was obtained by collecting the solution after 18 minutes.



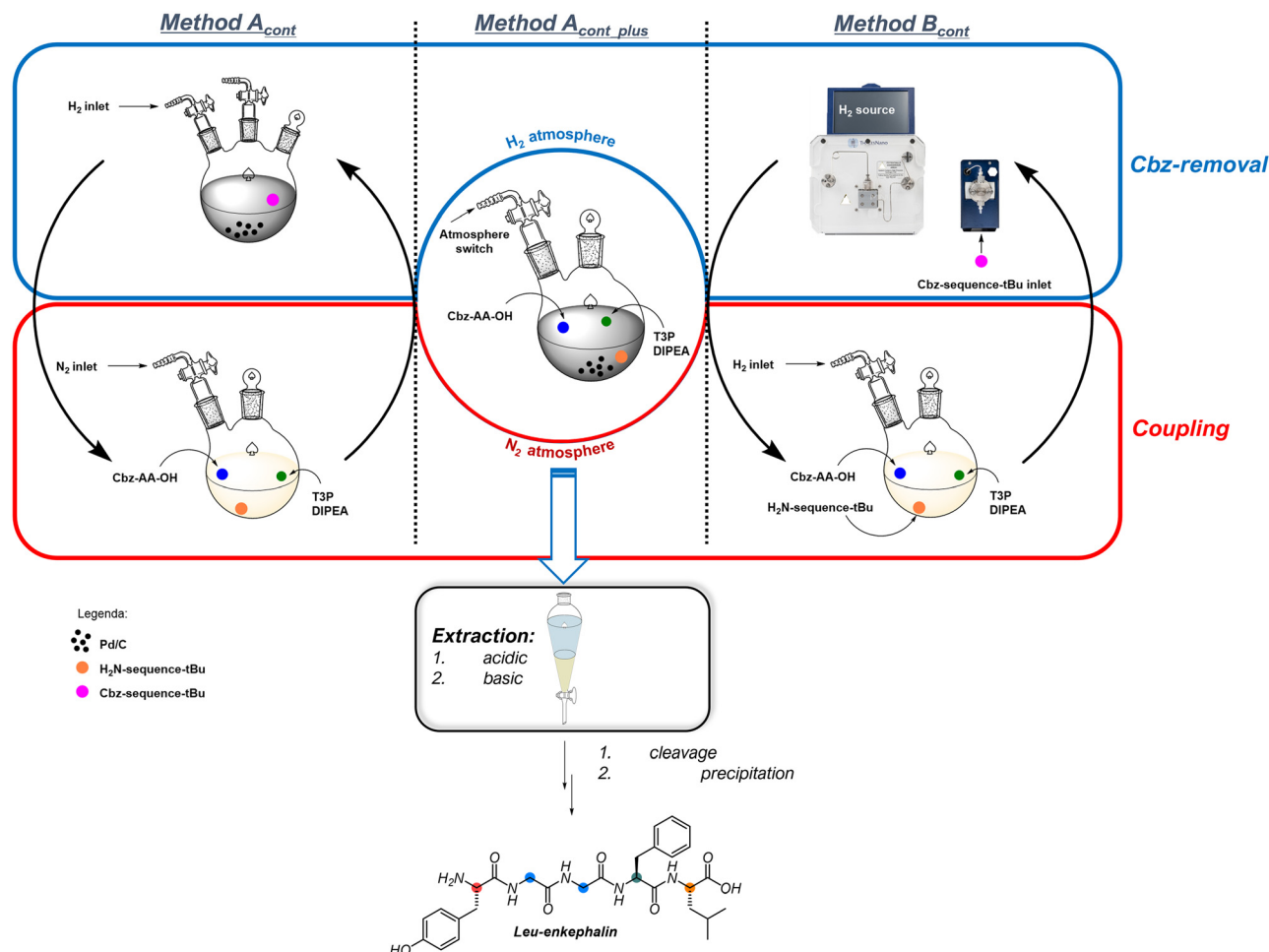


Fig. 1 Method A_{cont}, B_{cont} and A_{cont_plus} general procedures.

Table 5 Comparison of various methods in the synthesis of Leu-Enk by calculating relative Process Mass Intensity (PMI), recovery of solvent and base, and Process Mass Intensity after recovery (PMIr)

Entry	Method	Boc-Leu-Enk-O ^t Bu ^a (purity %)	Yield ^b (%)	Leu-Enk ^c (purity %)	PMI	Recovery ^d (yield %)	PMIr
1	A _{cont}	92	76	91	299	EtOAc (90) DIPEA (80)	180
2	B _{cont}	86	81	85	552	EtOAc (90) DIPEA (80)	200
3	A _{cont_plus}	86	78	86	197	EtOAc (90) DIPEA (80)	129

^a Calculated by HPLC (area % uncorrected with RRF). ^b Yield related only to Boc-Leu-Enk-O^tBu in the crude after washings. ^c Calculated by HPLC (area % uncorrected with RRF). ^d EtOAc and DIPEA were recovered by distillation of the organic phases after acid/basic water washings.

coupling step to produce **4** achieved >99% conversion in 5 minutes, and the crude solution was directly introduced into the hydrogenating flow system.

As reported above, treatment of **4** under these conditions easily afforded **5** with complete conversion.

Owing to the dispersion effect related to HCube® instrument setup, the volume of the collected solution was quadrupled compared to the starting solution.

Recently, practical issues regarding product diffusion in bolus injection mode for this system have been investigated, and stacked injections have been proposed as an alternative, but still dilution of the final product was observed.³³ Several factors can affect the diffusion such as the reactor geometry, the affinity of the reactant and product for the catalyst, the particle size of the catalyst and the ability of the system to reach a steady state. However, oligopeptides have multiple binding



sites for the Pd catalyst, and the design of the iterative process does not allow them to reach a steady state, thus limiting diffusion.

Therefore, before performing the following coupling step, EtOAc was distilled under an inert atmosphere to achieve a 0.1 M concentration, which is suitable for a further coupling step. Moreover, the recovery of EtOAc after each deprotection step reduced the impact of solvent volume on the PMI of the process.

In the following iterative coupling/deprotection steps, to ensure complete conversions and solvent dryness, a slightly higher excess of coupling reagents (2 equiv. of T3P® and 4 equiv. of DIPEA) was used. The Z-protection removal steps were always complete under the optimized conditions, but the dilution of the deprotected peptide always occurred, requiring partial concentration of the solution under vacuum before residue addition. At the end of this continuous synthesis, Boc-Leu-Enk-O'Bu was isolated in 81% yield with 86% purity. In the last step, Boc-Tyr(O'Bu)-OH was used to obtain the contemporary cleavage of all the protecting groups in a single reaction. For the entire reaction sequence, the solution remained clear and the DIPEA and T3P® deriving salts did not interfere with the multistep process completion or give backpressure failures related to increased viscosity of the solvent flow entering the H-Cube® catalyst cartridge.

To contain the PMI, we explored the deprotection/coupling iterative process in batches. We performed the solution-phase synthesis of the peptide by alternating the coupling reaction under nitrogen with the hydrogenation performed in a round bottom flask using an internal filtering septum that offers the possibility to remove Pd/C_{10%} under nitrogen atmosphere, thus avoiding air moisture contamination (method A_{cont}). Monitoring by HPLC each coupling and deprotection step revealed that under these conditions, each hydrogenation was completed in 1 hour as the couplings were performed in 5 minutes. The Pd catalyst was filtered and dried under a vacuum to be recycled for further use. The minimal amount of EtOAc added to wash the catalyst was removed under nitrogen flow to return to the standard volume. As for method B_{cont}, the last step was carried out using Boc-Tyr(O'Bu)-OH, giving Boc-Leu-Enk-O'Bu in 76% isolated yield with 92% purity. The synthesis of Boc-Leu-Enk-O'Bu was also achieved with a [2 + 2 + 1] sequence using method A_{cont} and by coupling the dipeptides Z-Gly-Gly and H-Phe-Leu-O'Bu before adding the last Boc-Tyr-OH residue. The satisfactory results obtained (98% conversion and 93% purity, see ESI for details†) confirmed that T3P® is an excellent reagent for coupling oligomers obtained by parallel synthesis, which is a common approach applied in medicinal chemistry for difficult sequence production.

At the end of the synthetic sequence, Pd/C_{10%} catalyst was filtered and recovered. By submitting the organic phase containing the peptide to acid washing (HCl 0.1 M), all the unreacted amines were removed, and the unreacted acids were eliminated by basic washing (NaHCO₃ 0.05 M). DIPEA was then recovered by turning the pH of the acid water phase to basic and extracting it with EtOAc. The yield of the protected

peptide was calculated after distillation and the recovery of the organic phase under vacuum.

A comparison of the two methodologies did not display great advantages of one over the other because the shorter reaction time in the deprotection for method B_{cont} was compensated for by the need for a larger amount of solvent and related distillation.

The data obtained from both methods represented the starting point for the development of a really green and sustainable solution-phase peptide synthesis (GSolPPS).

To decrease the amount of solvent and simplify the operations, allowing the decrease in investment for industrial scale application, we used a one-pot procedure.

Because the side products and unreacted reagents/catalysts of the two steps were proven not to interfere with the desired reactivity, we performed the synthesis of the pentapeptide using Method A_{cont} without any intermediate work-up, just adding the reagent sequentially in the proper order (method A_{cont_plus}).

During the multistep synthesis, no addition of solvent was required, and no addition of the Pd/C_{10%} catalyst was performed. This is because the activity remained unaffected in every hydrogenation step, and the atmosphere was simply switched from hydrogen to nitrogen and *vice versa*.

Therefore, after the first peptide coupling with amino acids, DIPEA and T3P® (1:1:4:1.5 ratio) in EtOAc (0.1 M) for 5 minutes, Pd/C_{10%} (10% w/w) with respect to the dipeptide, corresponding to 4.4 mol%, was directly added into the flask, and the atmosphere was switched from nitrogen to hydrogen.

When the deprotection step was complete after 1 hour, hydrogen was removed, and the new Z-amino acid (1 equiv.) was followed by DIPEA (4 equiv.) and T3P® (2 equiv.) were added and stirred for 5 minutes. This iterative sequence of additions was repeated until the pentapeptide was finally synthesized.

A 0.1 equiv. amount of DIPEA was added before each hydrogenation to reproduce the optimal conditions (see Table 3 entry 3). The recovery of Pd and base was performed as reported above for Method A_{cont}. From this continuous approach, peptide Boc-Leu-Enk-O'Bu was obtained in 86% purity and 78% yield (Table 5 entry 3).

This last protocol could be very useful for the synthesis of short peptides. For example, those commonly used in the cosmetic industry.³⁴ Thus, the synthesis of Pentapeptide-18 (Leuphasyl), composed of the sequence H-tyr-Ala-Gly-Phe-Leu-OH (H-YAGFL-OH), was performed following method A_{cont}. This peptide is an active peptide ingredient present in anti-aging formulations owing to its effectiveness in wrinkles reduction, which decreases the ACh secretion in the synaptic cleft similar to enkephalins. Under the one-pot synthetic conditions, H-YAGFL-OH was obtained at 89% purity. In the field of the green synthesis of oligopeptides, several strategies have been applied to improve the environmental footprint of these molecules.

To further confirm the sustainability of the synthetic process, we decided to calculate the process mass intensity



and the recovery of the reaction solvent and base (Table 5, PMI and PMIr).

As reported above, EtOAc was recovered by simple distillation under vacuum, while DIPEA was recovered after extraction from the basified water phase and distillation. The reported values include the deprotection step using typical TFA/TIS/water cocktail cleavage and the precipitation of the free peptide with diisopropyl ether (DIPE). Detailed calculations are reported in the ESI.† The downstream process for the isolation of pure peptides to achieve pharmaceutical-grade purity is not reported herein. However, the described coupling conditions were able to minimize the formation of diastereoisomers, which are the most critical impurities to be removed in the downstream process.³⁵

Among the three methodologies, the one-pot method A_{cont+plus} is certainly the greenest because a lower PMI was obtained (197), and owing to the recovery of solvent and base, the value of PMIr was further decreased to 129 (Table 5, entry 3). For methods A_{cont} and B_{cont}, the PMI values were higher (entries 1 and 2) but still very convenient, as the PMIr values were <200.

To correctly compare these methodologies with SPPS and LPPS,^{10a,36,37} we calculated the average PMI for the introduction of a single amino acid in the iterative process (Table 6), considering only solvents, reagents and water, being all the contribution of other factors (resin, tag and cleavage cocktail) depending on the length of the synthesized peptide. Among the LPPS technologies, Ajiphase® was selected, being the most efficient, even though it uses toxic solvents (if chloroform and DMF) and potentially explosive coupling reagents, such as HOBT.^{9a}

The one-pot method A_{cont+plus} (entry 3) was confirmed to be more sustainable (PMI/aa = 30). The GSPPS (entry 4), the LPPS, performed with Ajiphase® tag technology (entry 5), and the MW-assisted SPPS (entry 6) have higher PMI/aa values of 86, 200 and 180, respectively.

On these bases, to the best of our knowledge, the one-pot iterative deprotection/coupling method A_{cont+plus} represents the first example of a totally green solution-phase approach competitive to SPPS.

Table 6 Contribution to Process Mass Intensity (PMI) for single amino acid deprotection/coupling sequence

Entry	Method	PMI/aa
1	A _{cont}	46
2	B _{cont}	89
3	A _{cont+plus}	30
4	GSPPS ^a	86
5	LPPS ^b	200
6	MW-SPPS ^c	180

^aData calculated from the average PMI/aa in the synthesis of Octreotide (see ref. 10a). ^bData calculated from the synthesis of Bivaluridin with LPPS using Ajiphase® technology (see ref. 36). ^cData calculated from the synthesis of 65–74 ACP under microwave-assisted SPPS (see ref. 37).

Conclusions

Greening peptide synthesis represents an important research goal owing to the growing number of peptide drugs on the market. The reported processes were designed using solvent, reagents and reaction conditions with low environmental (PMI and ecotoxicity), health (human toxicity) and safety impact, requiring low energy usage. To complete the puzzle of sustainable reagents in this field, we explored the reactivity of T3P® as a coupling reagent in EtOAc solution, optimizing three methodologies that allowed the synthesis of peptides with a very good purity grade in a remarkably short time. Moreover, owing to the presence of a limited number of impurities, an easy downstream process can be foreseen. The use of the Z-protecting group was guaranteed to maintain dryness during all the steps, affording conditions suitable for using T3P® as a coupling reagent in iterative processes. Any other deprotection methodology using bases or acids, which require water washing, is not compatible with this green reagent.

The main characteristics of the described process are simplicity, short reaction time, and the use of green reagents and solvents at room temperature. Solvent and base can be easily recovered without great energy usage. Finally, the comparison of the PMI values for a single amino acid introduced with other iterative methodologies confirms the remarkable advancement in terms of sustainability for one-pot oligopeptide synthesis.

Author contributions

Alexia Mattellone and Dario Corbisiero equally contributed to the execution and analysis of the experiments. All authors have given approval to the final version of the manuscript.

Conflicts of interest

There are no conflicts to declare.

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