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Identifying the optimal oxidation state of Cu for electrocatalytic reduction of CO₂ to C₂₊ products†

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The electrocatalytic CO₂ reduction reaction (CO₂RR) to C₂₊ products is of great importance. It is known that the co-operation of Cu¹⁺ and Cu⁰ in the catalysts can yield a high faradaic efficiency (FE). However, it is very difficult to figure out the optimal ratio of Cu¹⁺ and Cu⁰ because Cu¹⁺ can be reduced to Cu⁰ during CO₂RR. To solve this problem and identify the optimal oxidation state of Cu, herein we propose a strategy to prepare Cu catalysts with different oxidation states, which could be stabilized by the pulsed electrolysis method during CO₂RR. On the basis of this method, we have studied the effect of the oxidation state of Cu on CO₂RR to form C₂₊ products. It has been found that the Cu catalyst with an oxidation state of +0.41 is the most efficient in our reaction system, and the FE of C₂₊ products is 70.3% in an H-type cell. This work provides a precise method to identify the optimal oxidation state of the catalysts that are not stable in the reaction.

The electrocatalytic CO₂ reduction reaction (CO₂RR) to form value-added carbon-based chemicals and fuels by utilizing renewable electricity is a promising technology to mitigate CO₂ emissions, fulfil the anthropogenic carbon cycle, and store excess renewable electricity as chemical energy.^{1–7} Among the products that can be generated from CO₂RR, C₂₊ products are the most desirable due to their high energy densities and industrial value as chemical feedstocks.^{8–12} However, their selectivity and activity are severely limited by multistep hydro-

genation and the sluggish kinetics of C–C coupling steps. To date, copper is known to be the most efficient electrocatalyst for selectively converting CO₂ to C₂₊ products.^{13–16} The synergism of Cu¹⁺ and Cu⁰ sites in copper catalysts has been verified to achieve high faradaic efficiency (FE) for CO₂-to-C₂₊ products.^{17–20} However, previous research has shown that Cu¹⁺ species are reduced to Cu⁰ in the reaction.^{21–23} Therefore, it is challenging to confirm the optimal Cu oxidation state for efficient electrocatalytic CO₂RR.

Various strategies have been used to tune the Cu electron structure, such as space confinement,^{24,25} the synthesis of alloys,²⁶ doping heteroatoms,^{27,28} and organic ligand modification.^{29,30} Among these, organic ligand modification (e.g. carboxylate and imidazole) has been reported to stabilize metal centers with appropriate oxidation states, and thus affects the intermediate adsorption during CO₂RR.^{30–36} Meanwhile, the Cu-based catalyst prepared by modifying the carboxylate ligand showed unique electrochemical CO₂ reduction selectivity toward C₂₊ products.^{37–39} This has made the method to be viewed as a good candidate for constructing Cu¹⁺ catalytic sites to promote the formation of C₂₊ products. However, the content of Cu¹⁺ drops dramatically during the potentiostatic electrolysis, resulting in the change of CO₂RR catalytic activity.

Pulsed potential electrolysis has emerged as a simple and effective method to increase the reaction durability and improve the product selectivity in CO₂RR *via* tuning the surface architecture, oxidation state, surface adsorbate coverage and local pH.^{40–44} Meanwhile, the pulsed electrochemical method is also a simple and quick method to prepare various materials, such as metals, alloys, metal chalcogenides and porous materials.^{45–49} Recently, based on the pulsed electrolysis method, our group proposed the “*in situ* periodic regeneration of catalyst (PR-C)” strategy to give long-term high efficiency of CO₂ electroreduction to generate C₂₊ products over the Cu catalyst by applying a positive potential pulse for a short time periodically in the halide-containing electrolyte.⁵⁰ At the same time, we also found that the Cu catalyst could be

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in situ regenerated to maintain the stability of the oxidation state of Cu *via* the pulsed potential electrocatalytic CO₂RR.

Identifying the optimal oxidation state of Cu in CO₂RR to form C₂₊ products is of importance from both fundamental and practical points of views. Herein, we designed and prepared several Cu_xC_yO_z catalysts with different Cu oxidation states using the pulsed electrochemical method. The oxidation state of Cu was stabilized by the pulsed potential in CO₂RR, and the optimal oxidation state of Cu for producing C₂₊ products was figured out. It was found that the catalyst with an average Cu valence state of 0.41 was most efficient, and the FE of C₂₊ products could reach 70.3% with a current density of 24.1 mA cm⁻² at -1.0 V *versus* the reversible hydrogen electrode (RHE).

As illustrated in Fig. 1A, the catalysts were prepared *via* a pulsed electrochemical method in 0.1 M KHCO₃ aqueous electrolyte containing 0.1 M potassium benzenedicarboxylate (K₂BDC). A typical H-type cell with three-electrode configuration was used in this work, which included a Cu foil working electrode, a Pt gauze counter electrode, and an Ag/AgCl reference electrode. Cu²⁺ ions were generated at the anode potential ($E_a = 1.25$ V *vs.* RHE), and they interacted with negatively charged carboxylate ligands to form the Cu complex. The Cu complex was then reduced to Cu_xC_yO_z at the cathode potential

($E_c = -1.0$ V *vs.* RHE). Anodic pulses (t_a) of 3 s followed by cathodic pulses (t_c) of 5 s were applied in this work. Such pulses were repeated for 80 cycles.

The scanning electron microscopy (SEM) image showed that the untreated Cu foil had a smooth surface (Fig. 1B). The Cu complex has a leaf-like structure generated by the stacking of lamellae due to the application of E_a (Fig. 1C). The X-ray diffraction (XRD) patterns in Fig. 1D confirmed the presence of a crystalline quasi-metal-organic framework (MOF).³⁹ After that, the electrochemical reconstruction^{51–53} was performed by reducing the Cu complex in an electrolyte to form the Cu_xC_yO_z catalyst. The leaf-like structure was converted into uniform nanoparticles (Fig. 1E). The high-resolution transmission electron microscopy (HR-TEM) image (Fig. 1F) confirmed that both metallic Cu and Cu₂O crystal lattices existed in the Cu_xC_yO_z catalyst, where 0.21 nm and 0.24 nm belong to Cu (111) and Cu₂O(111), respectively.^{54,55} This can also be confirmed from the XRD patterns (Fig. 1D). The corresponding energy-dispersive X-ray spectroscopy (EDS) results showed that the atomic ratio of Cu:C:O was approximately 4:2:1 (Table S1†). The EDS for elemental mapping (Fig. 1G) showed that the Cu, C and O elements were distributed homogeneously throughout the entire architectures.

The surface analysis of different samples was performed by quasi *in situ* X-ray photoelectron spectroscopy (XPS) (Fig. S1† and Fig. 1H). The Cu 2p XPS spectra and Cu Auger L₃M₄₅M₄₅ transition indicated that Cu²⁺ was the major species in the Cu complex, while Cu⁰ and Cu¹⁺ species existed in Cu_xC_yO_z. Furthermore, the detailed structural information of Cu was investigated by *in situ* X-ray absorption spectroscopy (XAS) and X-ray emission spectroscopy (XES). As shown in Fig. 1I, Cu K-edge X-ray absorption near-edge spectroscopy (XANES) of the Cu complex exhibited an edge profile similar to that of CuO or Cu(OH)₂ in the range from 8960 to 9020 eV, while the spectrum of Cu_xC_yO_z showed a close absorption edge with the Cu foil. These observations indicated that Cu in Cu_xC_yO_z has a lower oxidation state compared with that in the Cu complex. Both the Cu complex and Cu_xC_yO_z presented a characteristic Cu–Cu peak at 2.3 Å and a Cu–O peak at around 1.5 Å in the extended X-ray absorption fine-structure (EXAFS) spectra (Fig. S2†). However, the intensity of the Cu–O peak decreased and the intensity of the Cu–Cu peak increased in Cu_xC_yO_z compared to those of the Cu complex, suggesting that the oxidized copper in the Cu complex was partially reduced during the formation of Cu_xC_yO_z. Fig. 1J shows the Cu Kβ_{1,3} XES spectra of the Cu complex and Cu_xC_yO_z. When comparing the spectra with those of the reference samples, the *in situ* spectra for the Cu complex lay in between those for the Cu²⁺ and Cu⁺ references, whereas in the case of Cu_xC_yO_z, the spectra were in between those for Cu¹⁺ and Cu⁰. Both the XAS and XES results indicated the co-existence of Cu¹⁺ and Cu⁰ species in Cu_xC_yO_z, which is in line with the XPS data.

It is worth noting that application of different E_a values would affect the oxidation state of Cu in the Cu_xC_yO_z catalyst. As shown in Fig. 2A, we prepared a series of catalysts by changing the E_a . The SEM images are shown in Fig. S3,† and their

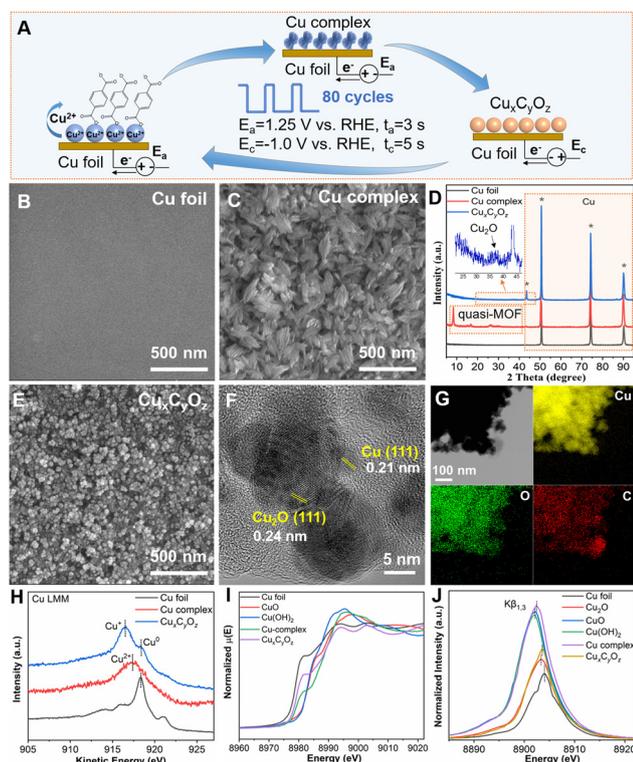


Fig. 1 (A) Schematic diagram of the preparation of the Cu_xC_yO_z catalyst. SEM images of (B) Cu foil, (C) Cu complex, and (E) Cu_xC_yO_z. (D) XRD patterns of different samples. (F) HR-TEM image and (G) elemental mappings images of Cu_xC_yO_z. (H) The quasi *in situ* XPS signals of Cu Auger LMM spectra for the Cu foil, Cu complex, and Cu_xC_yO_z. (I) XANES spectra and (J) XES spectra of different samples.

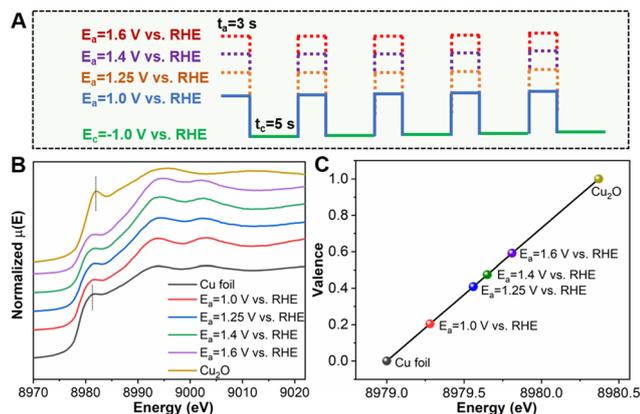


Fig. 2 (A) Schematic diagram of the preparation of the $\text{Cu}_x\text{C}_y\text{O}_z$ catalyst under different conditions ($E_a = 1.0$ V, 1.25 V, 1.4 V, and 1.6 V vs. RHE, $t_a = 3$ s, $E_c = -1.0$ V vs. RHE, $t_c = 5$ s). (B) XANES spectra of $\text{Cu}_x\text{C}_y\text{O}_z(0.41)$ reacted in route (1) for different times. (C) Average oxidation state of Cu in $\text{Cu}_x\text{C}_y\text{O}_z(0.41)$ under different conditions from Cu K-edge XANES.

morphologies were quite similar. The impact of the applied E_a on the Cu oxidation state was further investigated using XANES (Fig. 2B). The absorption edges of all the samples reside between those of Cu^0 and Cu^{1+} . We also acquired the oxidation state of Cu as a function of the Cu K-edge energy shift (Fig. 2C). The detailed calculation method for quantifying the oxidation average valence state of Cu is discussed in the ESI (Table S2†). The average valence of Cu increased gradually with the increase of the applied E_a , which was +0.20, +0.41, +0.47 and +0.59, when the applied E_a was 1.0 V, 1.25 V, 1.4 V, and 1.6 V vs. RHE, respectively. For making a clear distinction, these $\text{Cu}_x\text{C}_y\text{O}_z$ catalysts with different Cu oxidation states are denoted as $\text{Cu}_x\text{C}_y\text{O}_z(0.20)$, $\text{Cu}_x\text{C}_y\text{O}_z(0.41)$, $\text{Cu}_x\text{C}_y\text{O}_z(0.47)$, and $\text{Cu}_x\text{C}_y\text{O}_z(0.59)$.

Next, quasi *in situ* XPS and Auger LMM transition measurements were performed to characterize the composition and structure changes of the catalysts during the CO_2RR in different routes. In route (1), pulsed electrolysis ($E_a = 1.25$ V vs. RHE, $t_a = 3$ s; $E_c = -1.0$ V vs. RHE, $t_c = 50$ s) was applied in 0.1 M $\text{KHCO}_3\text{-K}_2\text{BDC}$ electrolyte. The same electrolysis conditions as applied for the preparation of $\text{Cu}_x\text{C}_y\text{O}_z$ catalyst were followed for the CO_2RR , except that the time of t_c was extended. In route (2), potentiostatic electrolysis was applied in 0.1 M $\text{KHCO}_3\text{-K}_2\text{BDC}$ electrolyte with $E_c = -1.0$ V vs. RHE. As shown in Fig. 3A and B, the content of Cu^{1+} could be maintained unchanged in route (1), while it would be decreased gradually in route (2). To further investigate the changes in the Cu oxidation state, operando XAS at the Cu K-edge was carried out. The Cu K-edge XANES spectra of $\text{Cu}_x\text{C}_y\text{O}_z(0.41)$ used in route (1) for different times exhibited an edge profile similar to that of the catalyst before the reaction, while close to the absorption edge of the Cu foil when reacted in route (2) (Fig. 3C and D). On the other hand, the intensity of Cu–Cu coordination (2.3 Å) of $\text{Cu}_x\text{C}_y\text{O}_z(0.41)$ in route (1) was basically unchanged compared to that of the catalyst before the reaction, but it gradually increased in route (2) in the Fourier transform (FT)

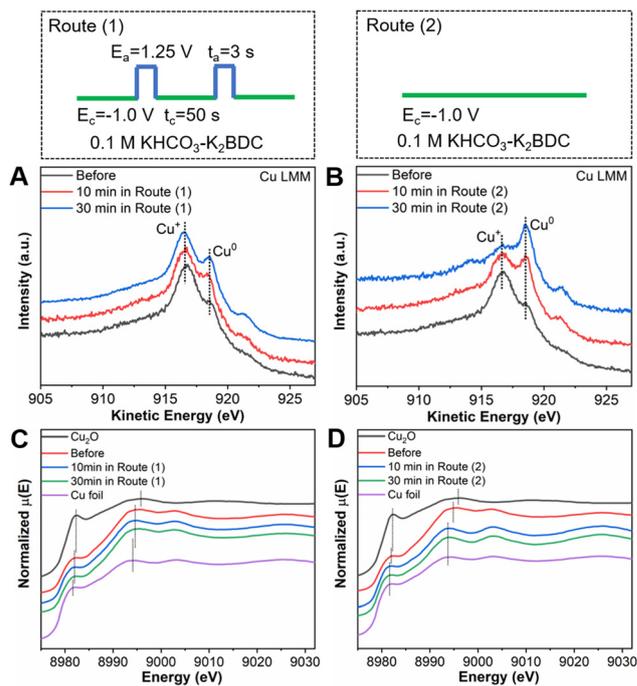


Fig. 3 The quasi *in situ* Auger LMM spectra of Cu for $\text{Cu}_x\text{C}_y\text{O}_z(0.41)$ reacted in (A) route (1) and (B) route (2) for different times. The *in situ* XANES spectra of $\text{Cu}_x\text{C}_y\text{O}_z(0.41)$ reacted in (C) route (1) and (D) route (2) for different times.

of the EXAFS spectra in R space (Fig. S4†). The results demonstrated that the Cu oxidation state can be maintained by the *in situ* regeneration of Cu^{1+} during the CO_2RR in route (1), while it declined gradually to tend to Cu^0 during the CO_2RR in route (2).

The CO_2RR performances of $\text{Cu}_x\text{C}_y\text{O}_z$ catalysts were then investigated in an H-type cell under different electrolysis conditions for 2 h. We considered the effect of K^+ concentration on the CO_2RR performance, and found that the increase of the K^+ concentration in the electrolyte did not promote the formation of C_{2+} products (Fig. S5†). We also carried out the electrolysis experiments under a N_2 atmosphere (without CO_2). No carbon-based reduction product could be detected, indicating that CO_2 was the carbon source in this work. For $\text{Cu}_x\text{C}_y\text{O}_z(0.41)$, it had a higher current density at -0.9 to -1.25 V vs. RHE in route (1) than that in route (2) (Fig. 4A and B). Route (1) also showed a significant difference in the distributions of CO_2RR products compared with route (2). C_1 products (CO, formate and methane), C_{2+} products (ethylene, ethanol and *n*-propanol) and H_2 can be detected by ^1H nuclear magnetic resonance (NMR) spectroscopy and gas chromatography (GC). The FE for C_{2+} products in route (1) was much higher than that in route (2). In route (1), the $\text{Cu}_x\text{C}_y\text{O}_z(0.41)$ catalyst exhibits a volcano-shaped dependence of total FE for C_{2+} products at different E_c values, and the maximum FE (C_{2+} products) could reach up to 70.3% at -1.0 V vs. RHE with a current density of 24.1 mA cm^{-2} , while the highest FE for C_{2+} products was only 46.6% in route (2). The catalyst also showed

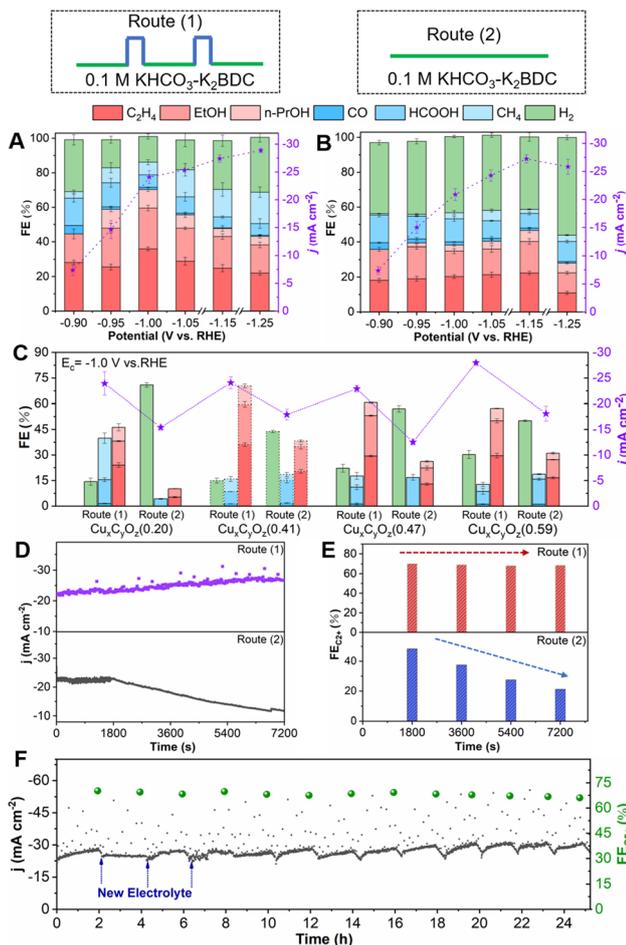


Fig. 4 CO₂RR product distribution and current density (j) of Cu_{*x*}C_{*y*}O_{*z*}(0.41) in (A) route (1) and (B) route (2). (C) CO₂RR product distribution and current density of Cu_{*x*}C_{*y*}O_{*z*} with different oxidation states of Cu. (D and E) Dependence of the current density (j) and FE of Cu_{*x*}C_{*y*}O_{*z*}(0.41) on time in route (1) and route (2). (F) Long-term stability of CO₂ pulsed electroreduction over Cu_{*x*}C_{*y*}O_{*z*}(0.41).

lower CO₂RR performance to form C₂₊ products in routes (3) and (4) in 0.1 M KHCO₃ (Fig. S6†).

Furthermore, the CO₂RR performance of other Cu_{*x*}C_{*y*}O_{*z*} catalysts was also evaluated. The Cu_{*x*}C_{*y*}O_{*z*}(0.41) catalyst exhibited an optimal CO₂-to-C₂₊ product performance with the average Cu valence state of 0.41 under an E_a of 1.25 V (Fig. 4C). In the control experiments, we optimized the duration of the pulse (t_a and t_c) and the concentration of the electrolyte. The detailed discussion is shown in the ESI (Fig. S7–S12†). The final optimized reaction conditions are $E_a = 1.25$ V vs. RHE, $t_a = 3$ s, $E_c = -1.0$ V vs. RHE, $t_c = 50$ s, and the FE of C₂₊ products was the highest in the electrolyte containing 0.1 M K₂BDC.

Based on the above analysis, some important observations are summarized below.

(1) The pulsed electrolysis can effectively inhibit the hydrogen evolution reaction (HER) and promote the formation of C₂₊ products.

(2) The addition of K₂BDC in the electrolyte did not contribute to the formation of C₂₊ products by comparing to that in routes (2) and (4) under potentiostatic electrolysis conditions (Fig. 4B and Fig. S6B†), while the addition of K₂BDC in the electrolyte combined with pulsed electrolysis can achieve the *in situ* regeneration of the Cu_{*x*}C_{*y*}O_{*z*} catalyst to stabilize the oxidation state of Cu during the pulsed CO₂RR, leading to the improved CO₂-to-C₂₊ performance.

(3) The oxidation state of Cu in Cu_{*x*}C_{*y*}O_{*z*} catalysts could regulate the CO₂-to-C₂₊ product performance. The optimal oxidation state of Cu in Cu_{*x*}C_{*y*}O_{*z*} catalysts was +0.41 corresponding to the best CO₂-to-C₂₊ performance.

(4) The SEM images show that there was no obvious change after different reaction times in route (1) and route (2) (Fig. S13†), suggesting that the K₂BDC in the electrolyte did not change the morphology of the catalyst, which also confirmed that the change in CO₂RR performance did not originate from the variation of the catalyst morphology.

The stability was crucial for the application of CO₂RR. First, the current density of CO₂RR and the FE of C₂₊ products, which depend on the reaction time within 2 h, were investigated. Obviously, the current density and FE of C₂₊ products over the Cu_{*x*}C_{*y*}O_{*z*}(0.41) catalyst did not change significantly with time in route (1), while both of them decreased continuously in route (2) (Fig. 4D and E). The Cu_{*x*}C_{*y*}O_{*z*} catalysts with the other oxidation states of Cu also show a similar phenomenon (Fig. S14†). Furthermore, consecutive cycles were carried out to determine the long-term CO₂RR stability of the Cu_{*x*}C_{*y*}O_{*z*}(0.41) catalyst in route (1). The current density and FE of C₂₊ products did not change notably over 25 hours in route (1) (Fig. 4F). The results demonstrated that Cu_{*x*}C_{*y*}O_{*z*}(0.41) exhibited outstanding catalytic activity and stability toward CO₂RR in route (1) due to the fact that Cu in the catalyst has an optimal oxidation state and can be maintained by the *in situ* regeneration of Cu during the reaction.

Conclusions

In summary, the Cu_{*x*}C_{*y*}O_{*z*} catalysts with different Cu oxidation states have been synthesized *via* the pulsed electrochemical method. The oxidation state of Cu can be stabilized by the pulsed anode potential in CO₂RR, which allows us to study the effect of the oxidation state of Cu on the performance of the catalysts more precisely. It is found that the FEs of C₂₊ products depend strongly on the oxidation state of Cu. The catalyst with a Cu oxidation state of +0.41 yields the highest C₂₊ FE of 70.3% with a current density of 24.1 mA cm⁻² in an H-type cell. This work provides a precise method to identify the optimal oxidation state of the catalysts. This method is specifically favorable for studying the catalysts that are not stable during the electrochemical reaction due to the reduction of the active species. Obviously, it is also useful for designing efficient catalysts with a suitable oxidation state for CO₂RR.

Author contributions

L. X., X. F. S. and B. X. H. proposed the project, designed the experiments, and wrote the manuscript. L. X. performed the whole experiments. J. Q. F., L. M. W., X. N. S. and X. X. T. performed the analysis of experimental data. L. B. Z., X. D. M. and S. H. J. conducted a part of the characterization study. J. D. and A. B. C. participated in discussions. X. F. S. and B. X. H. supervised the whole project.

Conflicts of interest

There are no conflicts to declare.

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