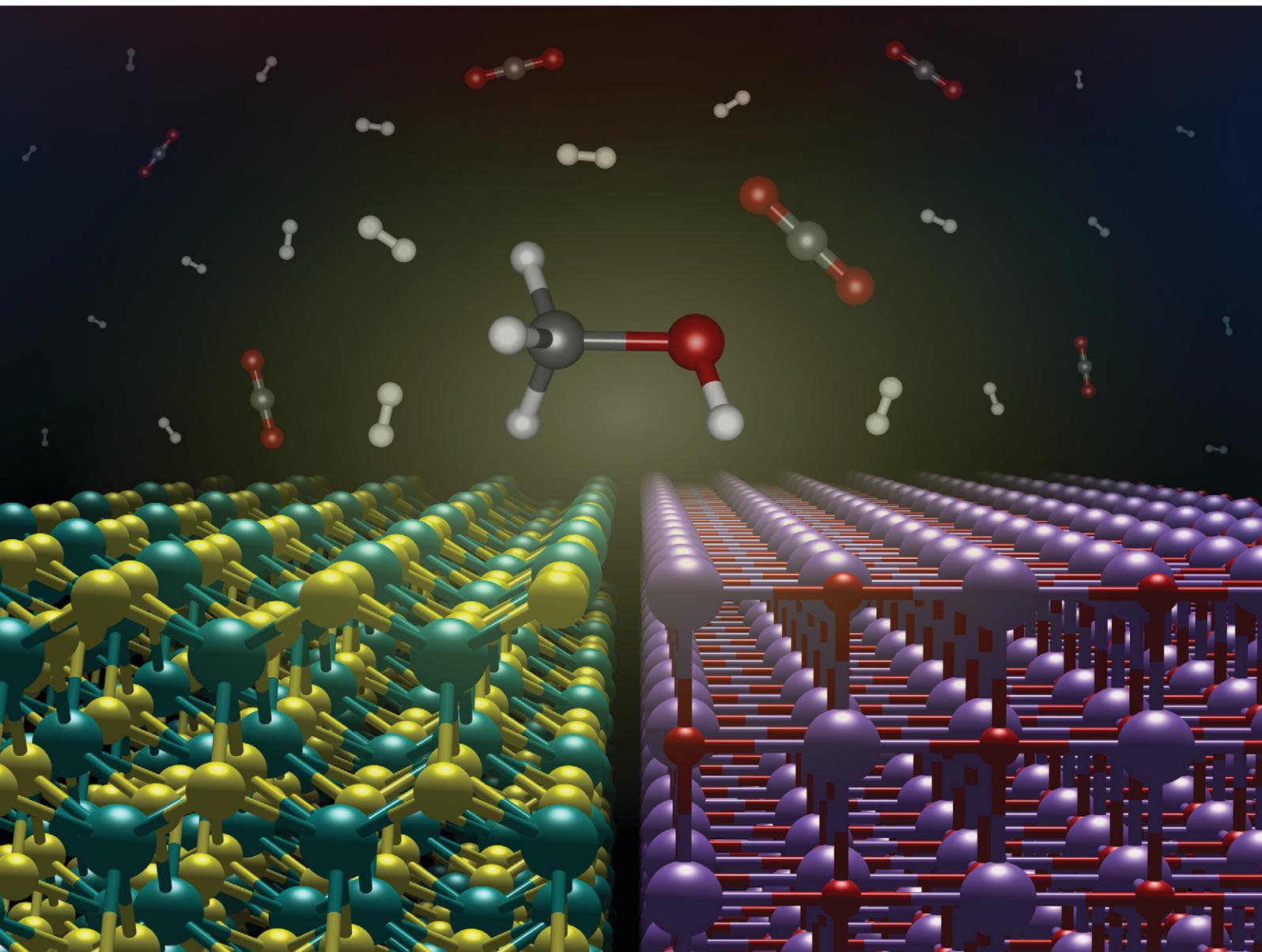


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Mn-promoted MoS₂ catalysts for CO₂ hydrogenation: enhanced methanol selectivity due to MoS₂/MnO_x interfaces†

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Considering the alarming scenario of climate change, CO₂ hydrogenation to methanol is considered a key process for phasing out fossil fuels by means of CO₂ utilization. In this context, MoS₂ catalysts have recently shown to be promising catalysts for this reaction, especially in the presence of abundant basal-plane sulfur vacancies and due to synergistic mechanisms with other phases. In this work, Mn-promoted MoS₂ prepared by a hydrothermal method presents considerable selectivity for CO₂ hydrogenation to methanol in comparison with pure MoS₂ and other promoters such as K and Co. Interestingly, if CO is used as a carbon source for the reaction, methanol production is remarkably lower, which suggests the absence of a CO intermediate during CO₂ hydrogenation to methanol. After optimization of synthesis parameters, a methanol selectivity of 64% is achieved at a CO₂ conversion of 2.8% under 180 °C. According to material characterization by X-ray Diffraction and X-ray Absorption, the Mn promoter is present mainly in the form of MnO and MnCO₃ phases, with the latter undergoing conversion to MnO upon H₂ pretreatment. However, following exposure to reaction conditions, X-ray photoelectron spectroscopy suggests that higher oxidation states of Mn may be present at the surface, suggesting that the improved catalytic activity for CO₂ hydrogenation to methanol arises from a synergy between MoS₂ and MnO_x at the catalyst surface.

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Introduction

With the alarming scenario of anthropogenic climate change due to persistently high emission levels of CO₂ from fossil fuels, Carbon Capture and Utilization (CCU) technologies may play a role in the transition of industry towards a net zero future.¹ Moving away from fossil fuels, CO₂ could be implemented as an alternative carbon feedstock in a variety of industrial processes.^{2,3} In this context, the production of methanol has been attracting particular interest, given its current role as a critical building block in the chemical industry and the development of sustainable aviation fuel in the near future.^{4,5}

Methanol synthesis from CO-rich syngas derived from natural gas using Cu/ZnO/Al₂O₃ catalysts has been a well-established industrial process over the past decades.

However, replacing fossil resources by alternative carbon sources such as biomass or captured CO₂ will require the development of new processes, among which the direct CO₂ hydrogenation may play a decisive role. In comparison with the traditional CO-based route, inherent limitations of the CO₂ hydrogenation reaction involve its less exothermic character, in addition to the simultaneous production of water.⁶ Nevertheless, reports from the first industrial plant devoted to CO₂ hydrogenation to methanol suggest significant advantages, being less energy-intensive and producing fewer reaction byproducts.⁷

Similarly as in the conventional methanol synthesis process, copper-based catalysts have been widely regarded as the most effective materials for CO₂ hydrogenation. Following extensive research, several structural-property relations have been identified for the Cu/ZnO system.^{8–10} Although some copper-based catalysts may produce methanol through the Reverse Water-Gas Shift + CO hydrogenation pathway,^{11,12} the Cu/ZnO interaction has been strongly associated with CO₂ hydrogenation *via* formate intermediate, according to theoretical studies and operando characterization.^{8,13,14}

In addition to such materials, other recent lines of research involve the search for novel catalysts, since the application of copper-based materials is still limited by the instability of the active phase,¹⁵ low selectivity to methanol¹⁶

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and surface poisoning in the presence of sulfur-containing gases.¹⁷ In virtue of these issues, non-metallic catalysts such as ZnZrOx (ref. 18) and In₂O₃ (ref. 19) have attracted growing interest as potential alternatives for this application due to their high stability and selectivity towards methanol.

More recently, similarly selective CO₂ hydrogenation to methanol has also been demonstrated using MoS₂ catalysts. In contrast with the aforementioned materials which optimally operate above 250 °C, MoS₂ can be active at remarkably low temperatures around 180 °C.²⁰ In addition to the more favorable thermodynamics for methanol formation, these mild operation conditions could offer an economical advantage with respect to other catalysts, in the context of large-scale applications. However, CO₂ hydrogenation to methanol can only be achieved in MoS₂ with specific properties, as under the typical reaction conditions most formulations of pure MoS₂ promote the formation of methane instead of methanol. In fact, only in the presence of basal plane S-vacancies CO₂ was shown to dissociate to CO and O at low temperatures, thus allowing for selective methanol production without further hydrogenation to methane.²⁰ This suggests that in favor of selective CO₂ hydrogenation to methanol, basal-plane sulfur vacancies should be obtained by preparation methods that induce the formation of sheet-like structures instead of edge-rich nanoparticles.

Moreover, in addition to the MoS₂ morphology, methanol selectivity can arise due to other effects. In a recent study, a MoS₂/ZnS catalyst produced by a solvothermal route involving a metal-organic framework precursor has shown high selectivity to methanol, as a possible result of ZnS blocking MoS₂ edge sites, thereby inhibiting CH₄ production.²¹ Therefore, these results suggest that the investigation of synergistic interactions between MoS₂ and other promoter compounds deserves closer attention.

In earlier works, MoS₂ has been combined with K,²² Co (ref. 23) and Ni (ref. 24) promoters leading to enhanced CO hydrogenation to higher alcohols at pressures around 100 bar. More recently, these promoters were incorporated into the hydrothermal synthesis of MoS₂, leading to considerable changes in its catalytic activity for CO and CO₂ hydrogenation, as the selectivity can be shifted from CH₄ to CO under lower reaction pressures of 20 bar.²⁵

In this work, K and Co promoted-MoS₂ are compared with Mn-promoted MoS₂ produced by a hydrothermal method. Following catalytic testing for CO₂ hydrogenation, Mn-promoted MoS₂ is further characterized by XRD, EXAFS, SEM, EDX, and XPS methods, seeking to investigate the main active phases that drive methanol production in the material.

Experimental

Catalyst synthesis

Mn-promoted MoS₂ was produced by a hydrothermal method, in which 2.2 g of ammonium heptamolybdate tetrahydrate ((NH₄)₆Mo₇O₂₄·4H₂O, Carl Roth, 99%) was

dissolved in approximately 20 mL water, together with the appropriate amounts of thiourea (CH₄N₂S, Merck, 99%) and manganese sulfate monohydrate (MnSO₄·H₂O, Merck, 99%). After stirring for 30 minutes, the mixture is transferred to an autoclave and kept at 200 °C during 16 h. The precipitate was filtered, washed with water and ethanol, dried under vacuum at room temperature and finally calcined under N₂ for 3 h.

For the comparison of pure MoS₂ with different promoters, Mn(0.5)-MoS₂, whose nomenclature refers to the nominal Mn/Mo molar ratio of 0.5, was produced using a CH₄N₂S:MnSO₄·H₂O:(NH₄)₆Mo₇O₂₄·4H₂O molar ratio of 32:3.5:1 and a calcination temperature of 500 °C. The analogous synthesis of K(0.5)-MoS₂, Co(0.5)-MoS₂ and MoS₂ is reported elsewhere.²⁵ The optimized Mn(0.3)-MoS₂ was produced using a CH₄N₂S:MnSO₄·H₂O:(NH₄)₆Mo₇O₂₄·4H₂O molar ratio of 24:2.1:1 and a calcination temperature of 400 °C.

Material characterization

X-ray powder diffraction (XRD) was performed in a PANalytical Empyrean. A Cu-LLF X-ray tube (45 kV, 40 mA, CuKα λ₁ = 1.5406 Å, λ₂ = 1.5444 Å) and a GaliPIX detector were used in Bragg-Brentano geometry. Scanning Electron Microscopy (SEM) was carried out using a FEI Quanta 250 FEG at a 5 kV voltage. Attached to the same device is an Octane Elite Super detector, which was employed for Energy Dispersive X-Ray (EDX) analysis, carried out at 20 kV.

Near-Ambient-Pressure X-ray Photoelectron Spectroscopy (NAP-XPS) was performed with a near ambient-pressure XPS system from SPECS (AlKα source, Phoibos 150 NAP detector) using a sample stage optimized for catalytic measurements.²⁶ The sample was mounted on a quartz sample carrier with steel backplate and heated with a laser. Pretreatment was carried out at 0.75 mbar H₂ and 400 °C and reaction conditions were simulated under 1 mbar H₂:CO₂ = 3:1 at 200 °C. *Ex situ* XPS was carried out using a SPECS u-Focus system (AlKα source, Phoibos 150 WAL detector). XPS data evaluation was carried out with CasaXPS software.²⁷ The peaks were fitted with Gauss-Lorentz (GL) sum functions and a Shirley background was used. All spectra were calibrated for S 2p_{3/2} = 162.0 eV.^{28,29}

XAS spectra of the Mn and Mo K-edge were collected at the SuperXAS beamline at PSI. It operated in top-up mode at 2.4 GeV and a ring current of 400 mA. A silicon-coated mirror (which also reduced higher-order harmonics) was used to collimate the polychromatic X-rays from a 2.9 T superbend magnet, which were subsequently monochromatized by a Si(111) channel-cut crystal. The monochromator was rocked at a frequency of 1 Hz resulting in two spectra per second. For Mn K-edge a Si-coated and for Mo K-edge a Rh-coated toroidal mirror focused the beam. The Mn, Mo K-edge absorption spectra were collected in transmission mode using 15 cm long ionization chambers filled with 1 bar of N₂ for Mn and 1 bar of N₂ plus 1 bar of Ar for Mo. A metal foil was measured simultaneously for absolute energy calibration.



Standard background subtraction, interpolation, and averaging were done with the Python-based software ProQEXAFS.³⁰ Normalization and analysis were performed with Demeter.³¹ The samples were measured in pellet form diluted by cellulose for optimized signal. In addition to the catalysts, commercial samples of MnO (ABCR, 99%) and MnCO₃ (ABCR, 95%) were analyzed for comparison.

Catalytic measurements

The catalytic activity was measured on a “micro effi” system of PID Eng&Tech in a fixed bed plug flow steel reactor. For the measurements, 1 g of the pure catalyst was pretreated at 21 bar H₂ at 400 °C. The reaction was conducted for 8 h at each temperature under 20% CO₂, 60% H₂ and 20% He at 21 bar, with a total flow of 5 mL_n min⁻¹. The product gas was analyzed by an Inficon Micro GC 3000 using a Plot Q column.

Results and discussion

Seeking to identify potentially promising promoters for MoS₂ catalysts in CO₂ hydrogenation, samples of M-promoted MoS₂ (M = K, Co or Mn) have been produced by analogous hydrothermal approaches. Subsequently, their catalytic activity for CO₂ hydrogenation was compared with pure MoS₂ obtained by the same method, as shown in Fig. 1a–c.

As expected, pure MoS₂ shows high selectivity for CH₄ between 180 and 320 °C, although the activity for CO production through the reverse water-gas-shift reaction becomes prominent above 220 °C. In contrast, the incorporation of K, Co or Mn precursors clearly lowers CH₄ and increases CO yields with respect to the pure material over the entire temperature range. Furthermore, Co–MoS₂

exhibits negligible methanol yield, while in K–MoS₂ some methanol production is observed around 280 °C.

Mn-promoted MoS₂, however, shows a distinct performance, with the highest methanol yield occurring at much lower temperatures around 200 °C. While pure MoS₂ shows a negligible methanol yield at 180 °C, Mn(0.5)–MoS₂ produces methanol with a selectivity of 45% at a CO₂ conversion of 0.8% under the same conditions. Furthermore, Fig. 1d shows that methanol production is mostly suppressed when CO is alternatively used as the feed gas, demonstrating that the catalyst cannot produce methanol from CO hydrogenation under these conditions. This observation suggests that a reaction mechanism involving CO as an intermediate can be ruled out.

Therefore, it is likely that methanol is produced *via* direct CO₂ hydrogenation *via* formate pathway in the presence of the Mn-promoted MoS₂ catalyst. Although this finding contrasts with the previously demonstrated CO-based pathway observed for pure MoS₂ nanosheets,²⁰ the coexistence of formate and CO intermediates has been already suggested for the MoS₂/ZnS catalyst.²¹ The divergence in these results support the idea that CO₂ hydrogenation to methanol in MoS₂-based catalysts may undergo distinct pathways due to structural changes and the addition of promoters to MoS₂, similarly as observed for copper-based catalysts, for example.^{11–14}

To obtain preliminary insights on the nature of the Mn promoter, characterization of Mn(0.5)–MoS₂ was conducted by XRD and *in situ* NAP-XPS. In Fig. S1,† the XRD pattern of Mn(0.5)–MoS₂ shows that the catalyst crystalline structure consists mostly of MoS₂,³² MnO,³³ and MnO₂.³⁴ Differently from previously reported K- and Co-promoted MoS₂,²⁵ no other sulfides and sulfates are formed in abundance, as MnS (ref. 35) appears with very low crystallinity in the XRD pattern.

In order to obtain more detailed insights into the surface composition of Mn(0.5)–MoS₂ under reaction-relevant conditions, NAP-XPS analysis was carried out firstly under vacuum at 200 °C, followed by H₂ flow at 400 °C and CO₂ + 3H₂ flow at 200 °C. Fig. 2a shows the region comprising Mo 3d and S 2s spectra. A single contribution is observed for sulfur at 226.4 eV, as typically reported for S²⁻ in MoS₂.^{28,29} The Mo 3d region, however, shows a more complex profile with three distinct species, which were fitted in accordance with the characteristic doublet separation of 3.14 eV. Coherently with the observation from XRD, the main Mo surface species consist of Mo⁴⁺ associated with MoS₂, evidenced by the 3d_{5/2} peak at 229.1 eV.^{28,29} Within a similar energy range, Mo⁴⁺ from MoO₂ is described by two neighboring doublets with different widths, in order to account for the typical asymmetry of MoO₂ peaks arising from the distinctive narrow band metallic character of this compound.³⁶ Accordingly, this species is fitted with neighboring 3d_{5/2} peaks at 229.3 eV and 231.0 eV, and respective 3d_{3/2} counterparts following area, width and position constraints reported in a previous study.

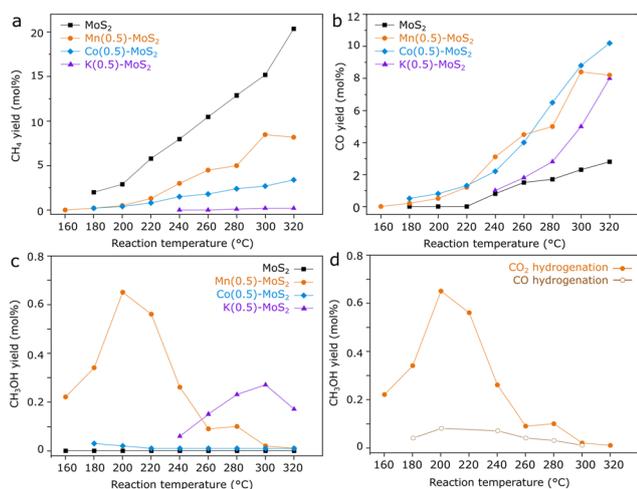


Fig. 1 a) CH₄, b) CO, c) CH₃OH yield obtained from catalytic reaction with 1 g MoS₂, Mn(0.5)-, Co(0.5)-, and K(0.5)-promoted MoS₂ under 1 mL min⁻¹ CO₂ + 3 mL min⁻¹ H₂ + 1 mL min⁻¹ He at 21 bar and d) CH₃OH yield under CO₂ hydrogenation conditions, compared with an analogous experiment using CO as carbon source.



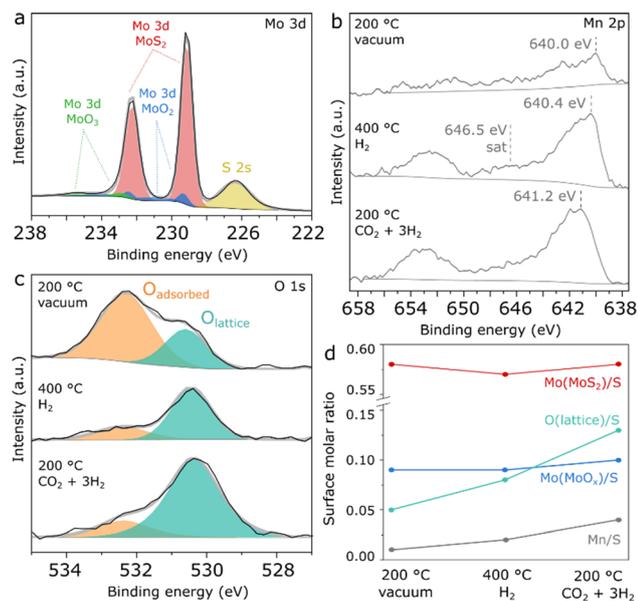


Fig. 2 NAP-XPS spectra of Mn(0.5)–MoS₂ showing the regions a) Mo 3d under vacuum at 200 °C, b) Mn 2p and c) O 1s under vacuum at 200 °C, H₂ at 400 °C and CO₂ + 3H₂ at 200 °C and d) summarized quantification of surface molar ratios with respect to S.

Furthermore, a minor contribution from another doublet is verified, with a 3d_{5/2} peak at 232.9 eV as a possible result of surface Mo⁶⁺ oxides.^{36,37} Correspondingly, MoS₂ produced by an analogous hydrothermal approach without the Mn promoter also presents similar evidences for surface Mo oxides, as seen in Fig. S2.† Since this pure MoS₂ sample has shown negligible methanol selectivity, this finding suggests that MoO_x should not be responsible for the promoting effect.

As shown in Fig. S3,† the chemical environment of surface Mo and S experiences insignificant changes upon H₂ pretreatment and reaction conditions. In Fig. 2b, the Mn 2p region initially shows a doublet with very low intensity, which suggests a low surface concentration of the Mn promoter. However, under H₂ treatment at 400 °C, the Mn 2p doublet becomes more prominent, exhibiting a 2p_{3/2} maximum at 640.4 eV and a satellite feature around 646.5 eV, as possible indicators of Mn²⁺ in MnO.³⁸ Subsequently, under CO₂ + 3H₂ flow at 200 °C the Mn 2p doublet is shifted towards higher binding energy by almost 1 eV, suggesting the formation of higher oxidation states such as Mn³⁺ and Mn⁴⁺ at the surface under reaction conditions. However, precise determination and quantification of these phases is challenging in a spectrum with such low intensity due to the complex and asymmetric character of the components associated with manganese oxides.³⁸

In Fig. 2c, the O 1s spectrum shows surface oxygen described by two species: the component at approximately 530.6 eV can be associated with lattice oxygen from Mo or Mn oxides, while the one at 532.3 eV typically refers to surface hydroxyl or organic species.³⁹ Although initially the lattice oxygen component is smaller, it becomes much more

prominent than the adsorbed species under pretreatment and reaction conditions.

As a summary of the NAP-XPS analysis, Fig. 2d shows a quantification based on combined survey spectra and high-resolution Mo 3d and O 1s regions. While the components related to MoS₂ and MoO_x show negligible changes, the Mn/S surface molar ratio increases concurrently with the O_{lattice}/S ratio during H₂ pretreatment and reaction conditions, suggesting the formation of surface Mn oxides at the surface under reaction-relevant conditions. Interestingly, Mn²⁺ appears to coexist with higher Mn oxidation states despite the reducing environment created by H₂ and CO₂, as a possible consequence of oxygen transferred from bulk to surface or due to the dissociation of CO₂ on MnO. Although precise identification of this surface oxidation mechanism is challenging, NAP-XPS analysis suggests that Mn oxides are the key Mn-containing phase under reaction conditions.

Given the promising catalytic activity of Mn-promoted MoS₂, the hydrothermal synthesis was further optimized in terms of calcination temperature, content of thiourea and content of Mn precursor. As shown in Fig. S4,† the methanol yield is improved by lowering calcination temperature from 500 °C to 400 °C, the MnSO₄·H₂O:(NH₄)₆Mo₇O₂₄·4H₂O molar ratio from 3.5:1 to 2.1:1 and CH₄N₂S:(NH₄)₆Mo₇O₂₄·4H₂O from 32:1 to 24:1.

In Fig. 3a, as a result of the optimized hydrothermal method, Mn(0.3)–MoS₂ shows a methanol selectivity of 64% with an improved CO₂ conversion of 2.8% at 180 °C. At higher temperatures, the formation of CO and CH₄ becomes dominant, although further cooling to 180 °C does not lead to an expressive decrease in methanol yield and selectivity. Furthermore, the material also exhibits stable catalytic activity at 180 °C during a 50-hour experiment, as shown in Fig. 3b.

To identify the main Mn-containing phases in the optimized catalyst, XRD and EXAFS have been performed in Mn(0.3)–MoS₂ catalyst, following exposure to relevant reaction conditions. In Fig. 4, as-synthesized Mn(0.3)–MoS₂ presents a variety of phases. As expected, the main contribution consists of MoS₂ (ref. 32) with low crystallinity,

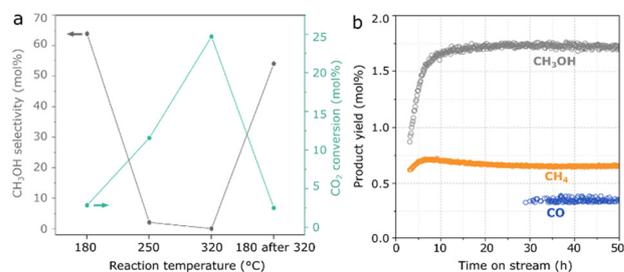


Fig. 3 Methanol selectivity and total CO₂ conversion as a result of a) 8 h catalytic reaction at 180 °C, 250 °C, 320 °C and 180 °C after 320 °C with 1 g Mn(0.3)–MoS₂ under 1 mL min⁻¹ CO₂ + 3 mL min⁻¹ H₂ + 1 mL min⁻¹ He at 21 bar and b) CH₃OH, CH₄ and CO yields under the same conditions during a 50-hour reaction.



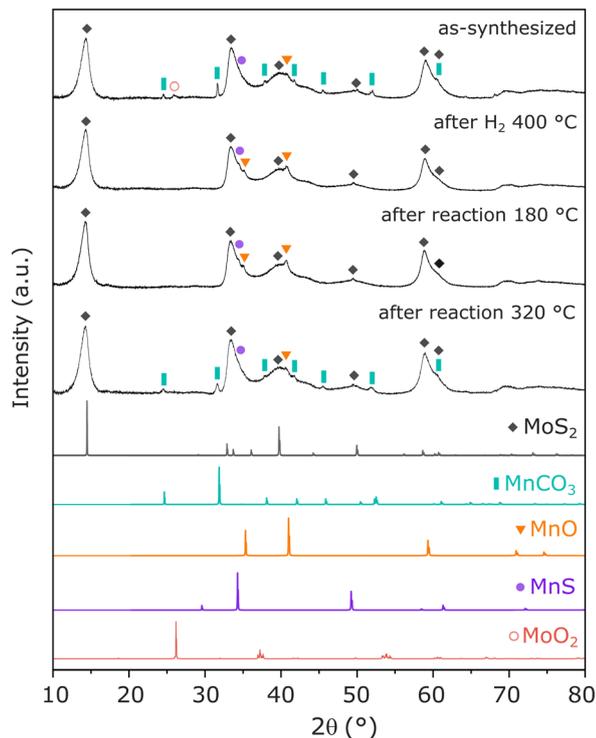


Fig. 4 XRD pattern of Mn(0.3)-MoS₂ following synthesis, H₂ pretreatment at 400 °C and catalytic reaction at 180 °C and 320 °C, in comparison with reference data for MoS₂ (COD-ID 1010993), MnCO₃ (COD-ID 1011228), MnO (COD-ID 1514099), MnS (COD-ID 1011351) and MoO₂ (COD-ID 9009090).

which is a typical outcome from the hydrothermal method. The promoter Mn is mostly present in the form of MnCO₃ (ref. 40) as a result of the reaction between the Mn precursor and thiourea during the hydrothermal synthesis. Additionally, the diffractogram suggests that minor contributions from MoO₂ (ref. 32) and nearly amorphous MnO (ref. 33) and MnS (ref. 35) are also present in the material.

The pretreatment under H₂ at 400 °C significantly alters the catalyst structure: although the patterns related to MoS₂ and MnS do not experience significant changes, MnCO₃ is absent in the pretreated material while a noticeable profile related to MnO (ref. 33) arises. This observation indicates that the pretreatment converts MnCO₃ into MnO, in an analogous manner as previously reported under H₂ at similar temperatures.^{41,42} Furthermore, the feature related to MoO₂ disappears, possibly indicating amorphization under reducing conditions. All these features are maintained after the material undergoes catalytic reaction at 180 °C, demonstrated to be the optimal condition for CO₂ hydrogenation to methanol.

Moreover, after reaction at 320 °C, the material exhibits once again the pattern from MnCO₃, indicating that the higher temperature favors the carbonation reaction of MnO in the presence of the CO₂ + 3H₂ mixture at 21 bar. This phase transformation may be closely associated with the

slight catalyst deactivation after reaction at 320 °C, already shown in Fig. 3a.

Given the observation of MnO following catalytic reaction at the ideal conditions for methanol production, this phase can be pointed out as a likely key feature behind the promoting effect of Mn. Nevertheless, due to the possible effect of amorphous Mn-containing species that could remain undetected by XRD, such as other oxides, sulfides or Mn intercalated within MoS₂ layers, EXAFS analysis was carried out in order to elucidate the coordination environment of Mn atoms in the catalyst.

In Fig. 5a, XANES spectra of the Mo edge show negligible differences between the samples. Fig. 5b and S5† show that the Fourier transform of the Mo K-edge exhibits mainly two features related to Mo-S and Mo-Mo, coherently with MoS₂.⁴³ As suggested in the EXAFS fitting in Fig. S5 and Table S1,† the higher prominence of the Mo-S coordination with respect to Mo-Mo may be associated with the low crystallinity of MoS₂, in line with the XRD patterns and previous EXAFS reports.^{44,45} Despite the differences observed for Mn, the Mo K-edges remain unchanged regardless of exposure to pretreatment and reaction conditions. Therefore, the result confirms the high stability of MoS₂ following reaction conditions and rules out major contributions from Mo oxides to the catalyst composition.

As shown in Fig. 5c, XANES spectra from the Mn K-edge region of Mn(0.3)MoS₂ give evidence for considerable changes between as-synthesized and spent catalysts. In Fig. 5d, the corresponding Fourier transforms show 2 main features related to Mn-O and Mn-Mn coordination shells.^{46,48} Consistently with the presence of MnO observed by XRD after treatment with H₂ at 400 °C and reaction at 180

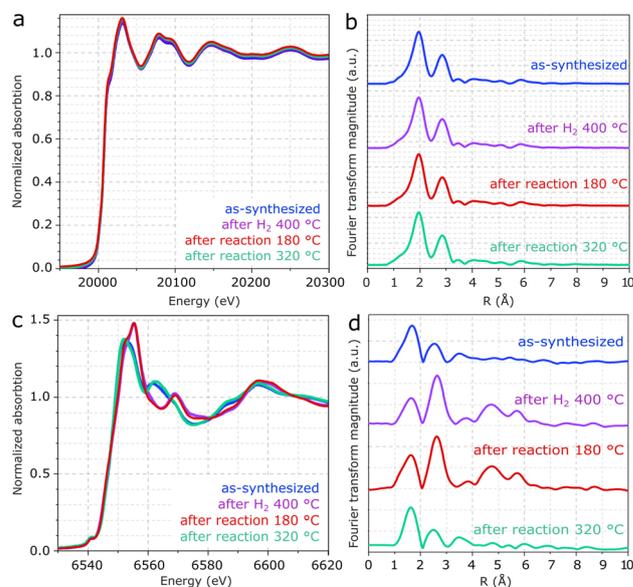


Fig. 5 a and b) XANES spectra and the associated EXAFS analysis of the Mo and c and d) Mn K-edges from Mn(0.3)-MoS₂ after synthesis, H₂ pretreatment at 400 °C and catalytic reaction at 180 °C and 320 °C.



°C, the EXAFS spectrum clearly shows the characteristic profile associated with MnO (ref. 47 and 48) in these two samples, as confirmed by the fitting presented in Fig. S6 and Table S2.[†]^{43,45}

On the other hand, as-synthesized Mn(0.3)MoS₂ exhibits a similar XANES spectrum from the one observed after 320 °C reaction, as shown in Fig. 5c. Despite some similarity with the typical XANES profile observed in MnCO₃,⁴⁹ the respective EXAFS data shows considerable differences when compared to a MnCO₃ reference in Fig. S7,[†] which raises the possibility of a more complex composition. In fact, Fig. 6 shows that these XANES spectra may be described as a linear combination of MnCO₃ and MnO. A clear indication of this similarity is the distinctive feature of MnO around 6569 eV, which makes the XANES spectra easily distinguishable from those of MnS and other Mn oxides.⁵⁰ In view of this finding, as-synthesized Mn(0.3)MoS₂ is considered to present the Mn promoter in the form of polycrystalline MnCO₃ in combination with low-crystallinity MnO, similarly as observed for the material exposed to reaction at 320 °C, since both samples show a weak contribution of the oxide in the XRD patterns in Fig. 4.

Material characterization by XRD and XAS strongly suggests that MoS₂ and MnO are the key components related to CO₂ hydrogenation to methanol in this catalyst. In order to unveil how these phases are present at the catalyst surface, further characterization was performed by SEM and XPS.

Fig. 7a and S8[†] show that the Mn(0.3)MoS₂ surface is mostly composed of thin MoS₂ sheets arranged in a nanoflower morphology, highly similar to MoS₂ produced by an analogous hydrothermal synthesis. This aspect is consistent with the broad XRD pattern observed for MoS₂, which suggests small crystallites and sparse stacking of MoS₂

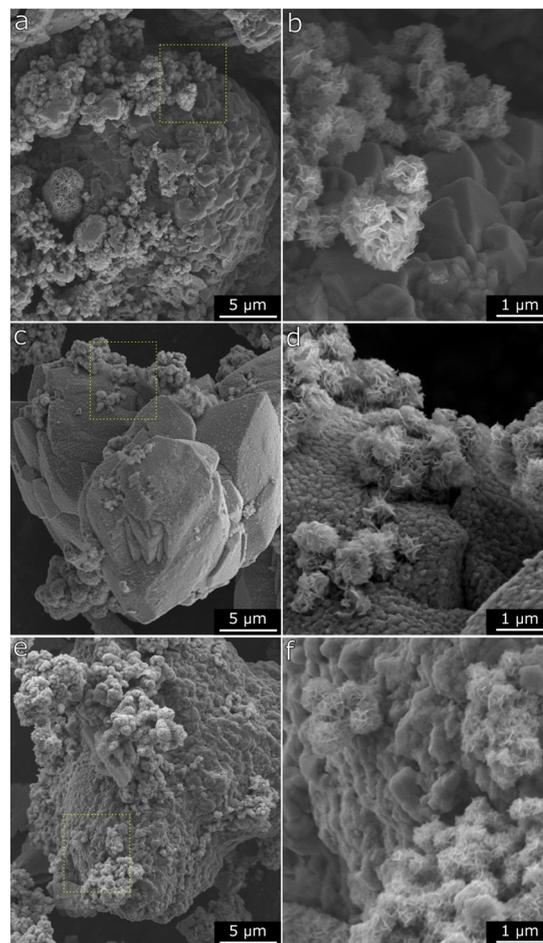


Fig. 7 a and b) SEM micrographs of Mn(0.3)-MoS₂ after synthesis, c and d) after H₂ pretreatment at 400 °C and reaction at 180 °C, e and f) after H₂ pretreatment at 400 °C and reaction at 320 °C.

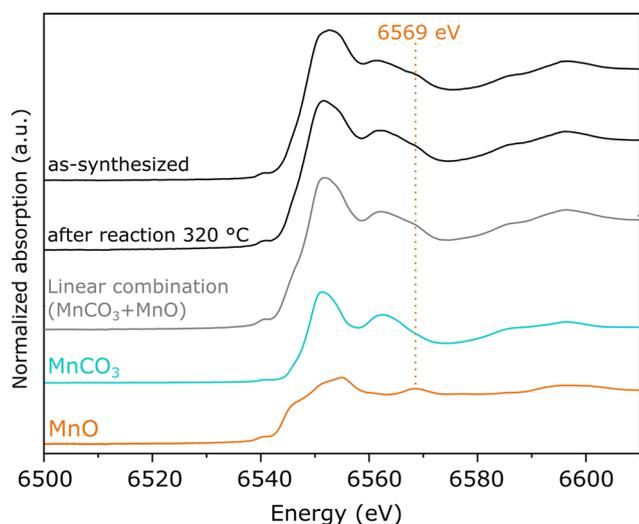


Fig. 6 XANES spectra of the Mn K-edge from Mn(0.3)-MoS₂ after synthesis and catalytic reaction at 320 °C, compared to MnO and MnCO₃ references and their linear combination considering equal proportions of each phase.

layers. Moreover, these features coexist with another evident morphology consisting of the micrometer-sized particles shown in more detail in Fig. 6b. After H₂ pretreatment and subsequent catalytic reaction at 180 °C, the nanoflower structure associated with MoS₂ remains unchanged, as shown in Fig. 6c and d. On the other hand, the larger particles are significantly altered, now featuring a rougher surface with abundant pores in the nanometer range. This finding is consistent with the conversion of MnCO₃ into MnO, which releases CO and CO₂, thus forming the characteristic porous surface. Such morphology change has also been already reported following exposure of MnCO₃ to reducing conditions under similar temperatures.^{41,42} As demonstrated by XRD and EXAFS analysis of Mn(0.3)MoS₂, further increasing reaction temperature to 320 °C induces the partial carbonation of MnO into MnCO₃. This effect can also be correlated with the SEM data in Fig. 6e and f, where the characteristic porous surface of MnO is again less prominent.

Further evidence for the MoS₂/MnO system after H₂ pretreatment is shown in Fig. 8, in which the EDX mapping confirms the elemental composition of the nanosheets as



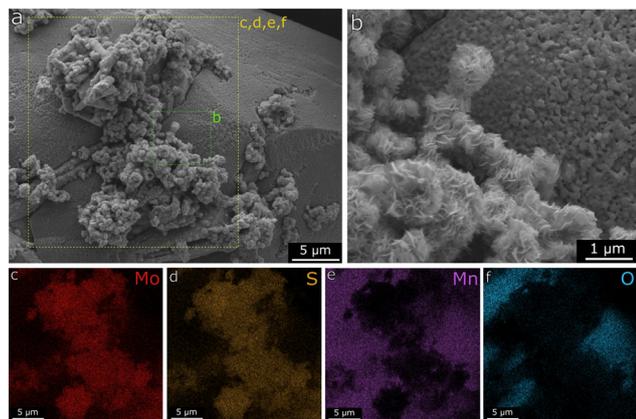


Fig. 8 a and b) SEM micrograph with the respective EDX element mapping for c) Mo, d) S, e) Mn and f) O atoms for Mn(0.3)-MoS₂, after H₂ pretreatment at 400 °C.

mostly Mo and S, while the larger particles are rich in Mn and O. A summary of the EDX spectra contained in the evaluated region is displayed in Fig. S9.†

Finally, further insights into the surface composition of Mn(0.3)-MoS₂ are provided by XPS analysis of fresh and used catalysts. In Fig. 9a, the S 2p spectrum shows a unique doublet with the characteristic splitting of 1.2 eV and S 2p_{3/2} located at 162.0 eV, as typically reported for sulfides such as MoS₂. Similarly as observed in the NAP-XPS analysis of Mn(0.5)-MoS₂, Fig. 9b indicates that MoS₂ (Mo 3d_{5/2} at 229.1 eV)^{28,29} coexists with surface MoO₂ (Mo 3d_{5/2} at 229.3 eV and 231.0 eV) as well as MoO₃ (Mo 3d_{5/2} at 232.9 eV),^{36,37} although MoS₂ brings again the most prominent contribution.

Accordingly, Fig. 9c shows that the O 1s spectrum indicates two distinct contributions: the minor peak at 532.3 eV can be ascribed to adsorbed oxygen from water or surface hydroxyl species, while the most prominent is located at 530.6 eV, associated with Mo and Mn oxides.³⁹ The low

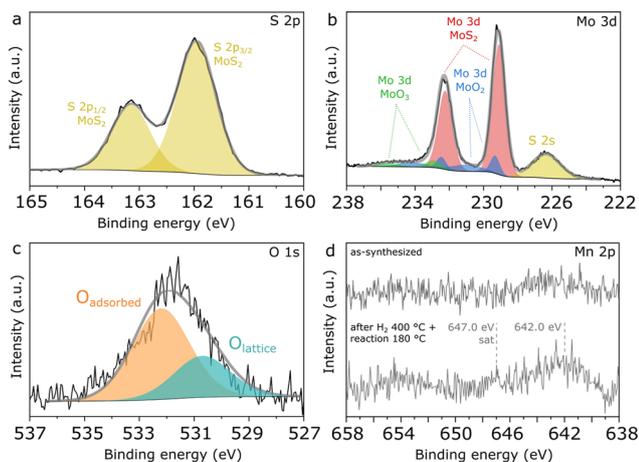


Fig. 9 High-resolution XPS spectra of Mn(0.3)-MoS₂ showing a) S 2p, b) Mo 3d, c) O 1s and d) Mn 2p regions.

intensity of the peak related to lattice oxygen can be correlated with the minor contributions from oxides in the Mo 3d and Mn 2p spectra, thus confirming that metal oxides consist only of a minor contribution to the total surface composition, with MoS₂ being the dominant species.

In order to verify possible modifications at the surface of Mn(0.3)MoS₂, XPS analysis was also conducted after H₂ pretreatment at 400 °C and reaction at 180 °C. As shown in Fig. S10,† these conditions do not promote expressive changes in S 2p, Mo 3d and O 1s regions, suggesting that the surface remains rich in stable MoS₂, in line with the strong similarities in the nanosheet morphology observed by SEM both before and after reaction. On the other hand, Fig. 9d shows that a significant change is observed in the Mn 2p region, as the doublet is only visible in the spent catalyst. This effect may be associated with the phase transition from MnCO₃ to MnO experienced by the catalyst upon H₂ pretreatment, as the higher surface area of the porous MnO could enhance the photoelectron signal related to the Mn species. Even after pretreatment and reaction, the Mn/Mo surface atomic ratio is still approximately 0.1, much lower than the nominal value of 0.3, as a possible outcome of the extensive covering of Mn phases by MoS₂ sheets demonstrated by the SEM data.

Furthermore, in the spent catalyst the Mn 2p spectrum is consistent with concurrent MnO and MnO_x with higher oxidation states, given the combination of a faint satellite feature related to MnO near 647.0 eV and the 2p_{3/2} peak located around 242.0 eV.³⁸ Since XRD and EXAFS characterization give evidence of MnO as the only Mn oxide phase in the material, these oxidized MnO_x species are understood to be limited to the catalyst surface.

In previous research, S-vacancies in MoS₂ have been strongly associated with its catalytic activity, and their presence at the surface can be usually verified by XPS. According to quantification based on Mo 3d and S 2s regions shown in Fig. 9b, an approximate S/Mo ratio of 1.7 is calculated considering only the Mo component associated with MoS₂. This low value may be associated with abundant sulfur vacancies formed during calcination under N₂, as it does not change significantly after H₂ pretreatment and reaction.

Moreover, given that abundant basal plane sulfur vacancies have been strongly suggested as active sites for CO₂ hydrogenation to methanol,²⁰ an oxygen chemisorption experiment coupled with *in situ* DRIFTS has been performed in an attempt to distinguish between edge- and basal plane sulfur vacancies. However, as shown in Fig. S11,† the results indicate no visible changes in the vibrational spectrum before and after O₂ flow. Although this suggests a limited concentration of basal plane sulfur vacancies in the catalyst with respect to previous reports, the presence of other surface metal oxides such as MoO_x and MnO_x introduces overlapping vibrational bands that may hinder the detection of the typical features related to oxygen chemisorption on sulfur vacancies.



In summary, characterization of Mn(0.3)-MoS₂ upon hydrothermal synthesis indicates that the Mn promoter is initially present as a combination of MnCO₃ and low-crystallinity MnO. Subsequently, H₂ pretreatment fully converts the carbonate into MnO, although higher oxidation states for Mn may be present at the surface. This MoS₂/MnO_x character is maintained after reaction at 180 °C. Even though increasing reaction temperature to 320 °C leads to partial carbonation of MnO, only a mild catalyst deactivation is observed.

Despite the correlation of MoS₂/MnO_x with catalytic activity, it is challenging to evaluate if MnCO₃ plays any role in catalytic activity, as here this phase always coexists with MnO, as demonstrated by XANES analysis. In light of these findings, CO₂ hydrogenation to methanol may be associated with the synergy between MoS₂ and MnO_x, while pure MoS₂ obtained by an analogous synthesis approach shows negligible selectivity for methanol.

In view of recent findings, MoS₂ growth in the vicinity of MnO_x might explain the improved methanol selectivity, since edge-blocking effects could inhibit the production of CH₄ in MoS₂ edges.²¹ Therefore, further studies with simpler MoS₂/MnO_x systems and detailed characterization of S-vacancies will be important for understanding the interplay between MoS₂ and MnO_x during CO₂ hydrogenation to methanol.

In comparison with other MoS₂-based materials presented in Table S3† such as few-layer MoS₂,²⁰ MoS₂/ZnS (ref. 21) and Cu/MoS₂@SiO₂,⁵¹ the Mn(0.3)-MoS₂ catalyst has a moderate methanol selectivity at lower or comparable CO₂ conversion levels. Accordingly, as shown in the SEM results, MoS₂ is not well dispersed with the Mn promoter phase, which would explain the notable production of the CH₄ byproduct related to pure MoS₂. Therefore, this suggests MoS₂/MnO_x catalysts could be further improved by employing synthesis methods that enhance the dispersion of these phases and their surface area.

Conclusions

In summary, this work demonstrates that Mn-promoted MoS₂ presents promising properties as a catalyst for CO₂ hydrogenation to methanol. Even though its catalytic activity is lower than in some other MoS₂-based catalysts, a sharp increase in methanol selectivity is observed in comparison with pure MoS₂ obtained by an analogous hydrothermal synthesis method. This improvement suggests a promoting effect of Mn, which may be closely associated with the presence of Mn oxides, according to material characterization. More specifically, the optimized catalyst contains MnO as the main Mn phase, although surface characterization indicates that Mn²⁺ coexists with higher oxidation states under reaction-relevant conditions. This finding suggests that the limited selectivity to methanol

could be further improved in catalysts with abundant MoS₂/MnO_x interfaces.

Furthermore, given the importance of basal plane S-vacancies in MoS₂ catalysts for CO₂ hydrogenation to methanol,²⁰ it can be speculated that MoS₂/MnO_x interfaces facilitate this reaction in Mn-promoted MoS₂, possibly due the blockage of MoS₂ edge sites, as previously observed for a MoS₂/ZnS system.²¹ Therefore, further characterization focused on S-vacancies is necessary for a deeper understanding of such edge-blocking mechanism between MoS₂ and metal-oxides. Moreover, the evidence for direct CO₂ hydrogenation without a CO intermediate in Mn-promoted MoS₂, differently as in pure MoS₂ nanosheets,²⁰ emphasizes the importance of investigating reaction mechanisms with more detail in future work.

Author contributions

Conceptualization: G. P., K. F. and G. A. S. A. Methodology: G. P. and K. F. Formal analysis: G. A. S. A., G. P. and S. P. Investigation: G. A. S. A., G. P., S. P., T. W., C. R. and R. R. Writing – original draft preparation: G. A. S. A. and G. P. Writing – review & editing: G. A. S. A. and K. F. Supervision: K. F. Funding acquisition: K. F.

Conflicts of interest

There are no conflicts to declare.

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