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Role of polyoxometalate precursors and supports in the selective oxidation of methane into formaldehyde using supported metal oxide subnanocluster catalysts†

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The direct synthesis of useful chemicals from methane (CH_a) is desirable; however, the products are prone to nonselective overoxidation, leading to the formation of CO2. A previous study developed a supported iron oxide subnanocluster catalyst with high thermal stability using iron-containing polyoxometalates (POMs) as precursors to selectively produce formaldehyde (HCHO) and CO. Herein, we investigated various supported POM-based catalysts to further improve the selectivity to HCHO via CH₄ oxidation, specifically by suppressing the pyrolysis and overoxidation of HCHO. After examining various metal-containing POM precursors and supports, we found that catalysts prepared using mononuclear- and dinuclear-ironcontaining POM precursors supported on SiO₂ with a high specific surface area were effective and yielded relatively high quantities of HCHO. In situ diffuse reflectance infrared Fourier transform spectroscopy (DRIFTS) measurements under HCHO flow demonstrated that the pyrolysis and oxidation of HCHO were suppressed on SiO₂, while the pyrolysis of HCHO was promoted on Al₂O₃. Furthermore, in situ DRIFTS measurements conducted at different temperatures revealed that HCHO was not decomposed or oxidized at 500 °C in the absence of catalysts.

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Introduction

Methane (CH₄), the main component of natural gas, is abundantly available worldwide, including in unconventional resources, such as biogas (including bio-methanation gas)¹⁻³ and CH₄ hydrates. Furthermore, the production of carbonneutral CH4, such as in power-to-gas technology, is expected to increase in the future. This production involves synthesizing CH₄ via methanation processes that utilize green hydrogen produced from renewable energy sources, such as water electrolysis and CO2 emitted from power plants and factories.5-7 Thus, CH4 is gaining attention as a major energy resource and a promising chemical raw material.8-11 However, CH₄, with its tetrahedral shape and four identical C-H bonds, has the highest C-H bond energy (439 kJ mol⁻¹) among sp³hybridized hydrocarbons and is known as the least reactive alkane, making it very challenging to chemically convert into useful compounds. 12

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There are two methods for CH₄ conversion: indirect conversion via synthesis gas (syngas) and direct conversion to useful compounds, such as formaldehyde (HCHO), in one step. 13 Indirect conversion involves a syngas process, which is a high-temperature and high-pressure process that requires complicated and expensive equipment and significant energy consumption. Therefore, an economically viable direct conversion process of CH4 into high-value-added chemicals is desired. One approach to economically achieve direct CH4 conversion is to selectively oxidize CH₄ to oxygenates using O₂. However, the target products are susceptible to thermal decomposition or sequential oxidation under the harsh oxidation conditions required for CH₄ activation. Therefore, developing a catalytic process that selectively synthesizes the desired products while suppressing sequential reactions is an urgent challenge for ensuring the stable supply of CH4 oxidation-derived chemicals.

Catalysts for synthesizing HCHO from CH4 have been developed since the 1980s, and various types of catalyst systems have been proposed, such as oxide catalysts, 14,15 supported catalysts, 16-19 and zeolite catalysts. 20,21 Commonly investigated oxide catalysts include molybdenum,16,17,19,22 vanadium, 16,23 iron, 24 copper, 22,25 cobalt, 21 and tungsten. 14,26 However, bulk oxide and large nanoparticle catalysts have

demonstrated relatively low selectivity to HCHO. For example, large nanoparticles of VO_x on the VO_x/Al₂O₃ catalyst reduced both CH4 conversion and HCHO selectivity in CH4 oxidation.23 Additionally, Zhang et al. reported that in molybdenum oxide catalysts, CH4 conversion and HCHO selectivity were related to the density of Mo=O bonds; namely, Mo=O bonds on Zr(MoO₄)₂ were responsible for HCHO production, whereas excess lattice oxygen and bulk MoO3 caused overoxidation of HCHO.²⁷ In the case of iron oxide catalysts, isolated active lattice oxygen atoms on the surface of iron oxides helped to suppress overoxidation during CH4 oxidation.²⁸ Thus, supported metal oxide nanocluster catalysts have attracted attention, with reported examples including supported CuO_x , ²⁹ FeO_x , ^{30,31} MoO_x , ³² and VO_x nanocluster catalysts.33

The choice of catalyst supports is also significant in selective CH₄ oxidation, and various oxide supports have been investigated, with SiO2 being recognized as the most suitable support. Kobayashi et al. discovered that highly dispersed Fe³⁺ species on SiO₂ with isolated tetrahedrally coordinated Fe-O species significantly promoted selective HCHO production.³⁴ Although detailed mechanistic studies are still in progress, it is evident that highly dispersed active sites in FeO_x nanoclusters on SiO₂ are essential for achieving selective HCHO production. Similarly, there is ongoing debate regarding the reaction mechanism of pyrolysis overoxidation on SiO₂ and other supports.

A polyoxometalate (POM) is an anionic metal oxide cluster consisting of metal-oxygen polyhedral units, such as {WO₆}.^{35,36} The structure formed by eliminating some of the polyhedral units is called a lacunary POM, which functions as an inorganic multidentate ligand and allows the incorporation of different metals.37-39 Recent studies have reported a new synthesis method for introducing various multinuclear metal oxide clusters into lacunary POMs in organic solvents. 40-44 Furthermore, a previous study used a diironcontaining POM (Fe2) as a precursor to in situ form FeOx subnanoclusters on SiO2, which served as the active species to selectively converted CH₄ into HCHO and CO. 45 This catalyst maintained the catalytic activity at 600 °C for 72 h, because the FeO_x subnanoclusters were dispersed in tungsten oxide species and inhibited from aggregation and deactivation. Another study successfully developed a method to achieve highly dispersed POM tetra-n-butylammonium (TBA) salts on various supports.46 Therefore, using POMs with different multinuclear metal oxide cores as catalyst precursors allows the investigation of various highly stable metal oxide subnanoclusters for the selective conversion of CH₄ to HCHO.

In this study, to further enhance the HCHO yield, we investigated CH₄ oxidation by using various POMs supported on SiO₂ as catalyst precursors (Fig. 1). Catalysts prepared using mononuclear (Fe1)- and dinuclear (Fe2)-iron-containing POM precursors supported on SiO₂ produced relatively high yields of HCHO. Furthermore, we investigated CH4 oxidation using various supports. The results demonstrated that although CH₄ conversion increased when using Fe2/Al₂O₃ and Fe2/CeO2, overoxidation to CO2 was promoted. We observed that catalysts prepared using SiO₂ (i.e., Fe2/SiO₂) exhibited significantly higher HCHO yields than catalysts prepared using other supports. In addition, to elucidate the mechanism underlying the suppression of overoxidation of HCHO on Fe2/SiO2, we investigated HCHO oxidation using Al2O3 and SiO₂ with different specific surface areas. We found that SiO₂ suppressed HCHO pyrolysis, and the suppression was enhanced as the specific surface area of SiO2 increased. Moreover, we conducted temperature variation tests on HCHO oxidation and observed that HCHO was minimally decomposed or oxidized at 500 °C in the absence of catalysts.

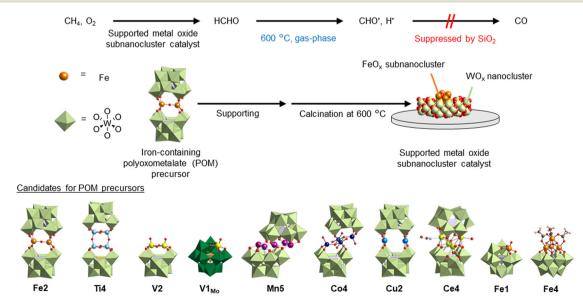


Fig. 1 Schematic diagram of the catalytic performance for CH₄ oxidation using supported metal oxide subnanocluster catalyst prepared from polyoxometalates.

Experimental

Paper

Catalyst characterization

Cold-spray ionization (CSI)-mass spectra were recorded on a JEOL JMS-T100CS spectrometer. Infrared (IR) spectra were measured on a JASCO FT/IR-4100 spectrophotometer using KBr disks. Inductively coupled plasma atomic emission spectroscopy (ICP-AES) analysis was performed on a Shimadzu ICPS-8100 instrument. Brunauer-Emmett-Teller (BET) surface areas were measured by N₂ adsorption at -196 °C using a Micromeritics TriStar II Plus instrument. Raman spectra were recorded on a JASCO NRS-5100 spectrometer. The measurement conditions included an irradiation laser wavelength of 532 nm and a laser power of 10.2 mW, and the data were collected twice with a measurement time of 3 min. X-ray absorption spectroscopy (XAS) of the Fe K-edge was performed using the transmission and fluorescence method at the BL14B2 beamline of SPring-8 (proposal numbers 2022B1656 and 2023A1512). The X-ray beam was monochromatized using a Si (111) monochromator, and the energy was calibrated using an Fe metal foil for the Fe K-edge XAS. X-ray absorption near-edge structure (XANES) and extended X-ray absorption fine structure (EXAFS) data were analyzed using Athena and Artemis software (Demeter, version 0.9.26; Bruce Ravel). The k^3 -weighted EXAFS spectra were Fouriertransformed into *R*-space in the range of 3–12 \mathring{A}^{-1} for Fe. The XAS measurement methods and EXAFS analyses are explained in detail in the ESI.†

Reagents

Dichloromethane, diethyl ether, and acetonitrile were purchased from Kanto Chemicals. SiO₂ (CARiACT Q-6, Q-10, Q-30, or Q-50, Fuji Silysia Chemical Ltd.), Al₂O₃ (KHS-46, Sumitomo Chemical Co. Ltd.), ZrO₂ (JRC-ZRO-6, Daiichi Kigenso Kagaku Kogyo Co., Ltd), CeO2 (JRC-CEO-5, Daiichi Kigenso Kagaku Kogyo Co., Ltd.), TiO2 (ST-01, ISHIHARA SANGYO KAISHA, Ltd), Nb₂O₅ (JRC-NBO-1, Companhia Brasileira de Metalurgia e Mineração), hydroxyapatite (HAP, Cat. No. 012-14882 FUJIFILM Wako Pure Chemical Corporation), and boron nitride (BN, Cat. No. 028-02281, FUJIFILM Wako Pure Chemical Corporation) were acquired from commercial sources. To represent SiO2 with different specific surface areas, CARiACT Q-6, CARiACT Q-30, and CARiACT Q-50 are denoted as SiO2-Q6, SiO2-Q30, and SiO2-Q50, respectively. SiO₂ without any specific indication refers to CARIACT Q-10. Fe₃O(CH₃CO₂)₇(H₂O)₃ (ref. 47) and TBA₄[SiW₉- $O_{28}(OCH_3)_6]$ (SiW9-OMe)^{48,49} were synthesized according to the reported procedures.

Synthesis of various POMs

 $TBA_8H_4Fe_2O(SiW_{10}O_{36})_2$ (Fe2),⁴⁵ $TBA_6Ti_4O_2(OH)_4(PW_{10}O_{36})_2$ (Ti4), 50 $TBA_4V_2O_2(OH)_2SiW_{10}O_{36}$ (V2),⁵¹ TBA₄VOPMo₁₁O₃₉ $TBA_7HMn_5(OH)_2(SiW_9O_{34})_2$ (Mn5), 42 $Co_4(OH)_4(SiW_{10}O_{36})_2$ (Co4),⁵³ $TBA_8H_4Cu_2(SiW_{10}O_{36})_2$ (Cu2),⁴⁰ $TBA_6Ce_4(H_2O)_2(CH_3CN)_2O(SiW_{10}O_{36})_2$ (Ce4), ⁴¹ and $TBA_4HFe(H_2-G_3)_2O(SiW_{10}O_{36})_2$ O)SiW₁₁O₃₉ (Fe1)^{54,55} were synthesized according to the reported procedures. In addition, TBA₄Fe₄(OH)₃(CH₃CO₂)₃SiW₉O₃₄ (Fe4) was synthesized according to the following procedure. SiW9-OMe (500 mg, 152 µmol) was added to a dichloromethane solution (40 mL) of Fe₃O(CH₃CO₂)₇(H₂O)₃ (134 mg, 205 μmol), and the resulting solution was stirred for 24 h at room temperature. Then, the remaining precipitate was filtered off, and the filtrate was dropped into diethyl ether (40 mL). Finally, the generated precipitate was collected by filtration through a membrane filter. After vacuum drying, a powder sample of Fe4 was obtained. The anionic structure of Fe4 was estimated from the IR spectra (Fig. S1†), ICP-AES elemental analysis, and CSI-mass spectrum (Fig. S2†), as described below. The IR spectra illustrated acetate peaks slightly shifted from those of Fe₃O(CH₃CO₂)₇(H₂O)₃. The elemental analysis confirmed that four Fe3+ ions were introduced into an [SiW₉O₃₄]¹⁰⁻ (SiW9) unit. The CSI-mass spectrum indicated the presence of a molecule containing four iron atoms, an SiW9 unit, three hydroxide ligands, and three acetate ligands (Fig. S2†). These results suggested that Fe4 had a cubane-type structure, as illustrated in Fig. 1, since a similar structure was reported for a manganese analogue according to the IR results.⁵⁶ IR (KBr pellet, cm⁻¹): 3441, 2962, 2874, 1631, 1581, 1544, 1485, 1460, 1382, 1275, 1152, 1106, 1000, 958, 909, 804, 672, 524, 376, 359, 324, 319, 303, 297, 290, 282, 278, 271, 257, 252. Elemental analysis: calcd (%) for TBA₄Fe₄(OH)₃(CH₃-CO₂)₃SiW₉O₃₄·2CH₂Cl₂, Si 0.74, Fe 5.85, W 43.33; found Si 0.74, Fe, 5.87, W 43.38. Positive-ion CSI-mass (acetonitrile): m/z =3891.36, $[TBA_5Fe_4(OH)_3(CH_3CO_2)_3SiW_9O_{34}]^+$ (theoretical m/z =3890.48).

Preparation of POM-supported catalysts

Various types of POMs were generally dispersed on supports with the loading amount of 10 wt% using the incipient wetness method.46 The preparation method of Fe2/SiO2 is described as a typical example. A TBA salt of Fe2 (100 mg) was dissolved in acetonitrile (2 mL). Then, the resulting solution was dropped onto a thin layer of SiO₂ (900 mg) spread on an evaporation tray. The resulting powder was dried at 100 °C for 5 h and then calcined at 600 °C for 5 h under air atmosphere, giving the Fe2/SiO2 catalyst. Various kinds of supported catalysts were prepared by the same method using POM precursors shown in Fig. 1. These supported catalysts were denoted using the abbreviation of the POM precursors and supports, such as Fe1/SiO2, Fe2/SiO2, and Fe4/SiO2. As for Fe2/SiO2, the catalyst with 35 wt% Fe2 (35Fe2/SiO2) was also prepared to align the loading amount of iron with Fe4/SiO2. Additionally, Cs-Fe2/SiO₂ was prepared following our previous report.⁴⁶ Briefly, TBA-Fe2 was supported on SiO₂ in acetonitrile using the incipient wetness method. Then, Cs-Fe2/SiO2 was obtained by cation exchange from TBA-Fe2/SiO₂ using cesium trifluoromethanesulfonate as the Cs source in ethanol.

Evaluation of CH₄ and HCHO oxidation performance

Catalytic performance tests for CH₄ oxidation were conducted using a fixed-bed flow-type reactor schematically illustrated in Fig. S3a.† After 100 mg of catalyst was loaded into a quartz tube reactor (6.0 mm i.d.), the reactant gas ($CH_4: O_2: Ar = 2:$ 1:7, total flow rate: 50 mL min⁻¹) was introduced into the reactor. Then, the input temperature of the tube furnace containing the catalyst was increased to 600 °C at a rate of 10 °C min⁻¹. The inlet and outlet gases were heated to 100 °C to reduce the temperature variation in the catalyst bed and to suppress the condensation of products. When the furnace temperature reached 600 °C, it was maintained for 1 h to allow the activity to stabilize. Then, HCHO was trapped in 20 mL of an aqueous solution containing 2 g of Na₂SO₃ and 50 µL of H₂SO₄, and the amount of trapped HCHO was determined by titration with an aqueous solution of NaOH (0.1 M). CO and CO2 were analyzed immediately before and after HCHO collection using a Nexis GC-2030 gas chromatograph equipped with a barrier discharge ionization detector (Shimadzu Corporation) and a Shincarbon-ST packed column. The respective calculation formulas for CH₄ conversion, product selectivity, and product yield in this study are as follows:

$$\label{eq:ch4} \begin{aligned} & \text{CH}_4 \text{ conv. (\%)} \\ &= \frac{\text{Total carbon moles of HCHO, CO, and CO}_2}{\text{A carbon mole of input CH}_4} \times 100 \end{aligned}$$

Product sel. (%) $= \frac{\text{A carbon mole of HCHO, CO, or CO}_2}{\text{Total carbon moles of HCHO, CO, and CO}_2} \times 100$

Product yield (%) = CH_4 conv. (%) × Product sel. (%) ÷ 100

In the pyrolysis and oxidation test of HCHO, gaseous HCHO was generated by flowing Ar or O2/Ar gas through 10 g of paraformaldehyde heated to 55 °C, as illustrated in Fig. S3b.†

In situ DRIFTS measurement

Diffuse reflectance infrared Fourier transform spectroscopy (DRIFTS) was performed using a JASCO FT/IR-6700 spectrometer equipped with an in situ sample cell in the range of 1500-4000 cm⁻¹ at a resolution of 4 cm⁻¹ (number of scans: 64). The samples were loaded onto a sample plate (6 mm diameter) and placed on the heater in the IR cell. The sample was pretreated by heating in a vacuum at 600 °C for 1 h, and the background spectrum was measured. After the pretreatment, the sample was exposed to approximately 1% HCHO in N₂ (20 mL min⁻¹) for 1 h. Then, the cell was evacuated into a vacuum at 600 °C, and the DRIFTS spectra were measured both in the reactant gas and in a vacuum. HCHO was generated in the same manner as described above.

Results and discussion

CH₄ oxidation using various catalysts

First, the catalytic performance for CH₄ oxidation was investigated for various POM-based catalysts at 600 °C under atmospheric pressure, and the results are presented in Table 1. Ti4/SiO₂, V2/SiO₂, V1_{Mo}/SiO₂, Cu2/SiO₂, and Ce4/SiO₂ exhibited relatively low CH₄ conversion (Table 1, entries 2-4, 7, and 8). Although there have been several reports 16,17,19,22,23,25 on the oxidation of CH4 to HCHO using vanadium, molybdenum, and copper oxides, the CH₄ oxidation performance using these POM precursors supported on SiO2 was much lower than that using Fe2/SiO2 (Table 1, entry 1). Particularly, the use of Ce4/SiO2 resulted in minimal HCHO formation, with significant CO2 formation from complete CH4 oxidation (Table 1, entry 8). In contrast, although Fe2/SiO2, Mn5/SiO2, and Co4/SiO2 exhibited relatively high CH4 conversion, the overoxidation to CO_x (CO and CO₂) was more pronounced for Mn5/SiO₂ and Co4/SiO₂ than for Fe2/SiO₂ (Table 1, entries 1, 5, and 6). In the case of iron, reducing the cluster size promoted the partial oxidation of CH4 to HCHO, while the catalytic performance of manganese and cobalt nanoclusters remained similar to that of bulk oxides⁵⁷ or nanoparticles.⁵⁸ Thus, the catalytic performance for CH₄ oxidation on SiO₂supported metal oxide subnanocluster catalysts prepared from various POM precursors indicated that Fe2/SiO2 catalysts exhibited the highest HCHO yield among the metal oxide subnanocluster catalysts with 3d metal or Ce metal multinuclear structures.

Next, the catalytic performance for CH₄ oxidation was investigated for various Fe-POM-based catalysts prepared using iron-containing POM precursors with various multinuclear oxide cores at 600 °C under atmospheric pressure. The results are summarized in Table 1. The iron content of the supported iron catalysts is also summarized in Table S1.† Among the catalysts prepared from POMs with different numbers of iron nuclei, Fe1/SiO2 and Fe2/SiO2 exhibited relatively high HCHO yields of 0.73% and 0.57%, respectively (Table 1, entries 1, 9, and 10). Fe4/SiO2 exhibited a relatively high CH4 conversion of 3% but a low HCHO selectivity of 11%, resulting in a low HCHO yield of 0.33%. 35 wt% Fe2/SiO2 (35Fe2/ SiO₂) was prepared using Fe2 with the same iron loading amount as that of Fe4/SiO₂, and the reaction was performed using 35Fe2/SiO₂. As a result, 35Fe2/SiO₂ exhibited lower CO₂ selectivity than Fe4/SiO2 (Table 1, entries 10 and 11). In addition, Cs-Fe2/SiO₂ was prepared to improve the thermal stability of the anionic structure of Fe2,46 and the CH4 oxidation performance of Cs-Fe2/SiO₂ was investigated under the same reaction conditions. However, despite its high selectivity toward HCHO (73%), Cs-Fe2/SiO₂ exhibited relatively low CH₄ conversion (0.67%) (Table 1, entry 12).

To investigate the structural differences between Fe2/SiO₂, 35Fe2/SiO2 and Fe4/SiO2 after the reaction, Fe K-edge XAFS measurements were performed, with the XANES and EXAFS spectra displayed in Fig. S4 and S5,† respectively. The Fe K-edge XANES results indicated that the valence of Fe in each catalyst prepared from Fe2 and Fe4 was not different. In a previous study, Fe2/SiO2 was directly used for the reaction without any pretreatment.45 In contrast, Fe2/SiO2 in this study was used after calcination. The results of Fe K-edge XANES and EXAFS suggested that Fe2/SiO2 prepared by

Table 1 Catalytic performance in CH₄ oxidation for various catalysts prepared from POM precursors^a

Entry	Catalyst	Conv. [%]	Sel. [%]			Yield [%]
			НСНО	CO	CO_2	НСНО
1	Fe2/SiO ₂	1.4	41	45	14	0.57
2	Ti4/SiO ₂	0.40	82	5	13	0.33
3	$V2/SiO_2$	0.44	55	27	18	0.24
4	$V1_{Mo}/\overline{SiO}_2$	0.52	63	22	15	0.33
5	$Mn5/SiO_2$	0.97	25	45	30	0.24
6	$Co4/SiO_2$	1.4	22	52	26	0.30
7	Cu2/SiO ₂	0.69	47	27	26	0.33
8	Ce4/SiO ₂	0.56	_	66	34	_
9	Fe1/SiO ₂	1.4	54	32	13	0.73
10	Fe4/SiO ₂	3.0	11	62	28	0.33
11	35 Fe2 /SiO ₂	1.9	29	58	13	0.55
12	Cs-Fe2/SiO ₂	0.67	73	20	7	0.49

^a Reaction conditions: 10 wt% supported catalyst (100 mg), CH₄: O₂: Ar = 2:1:7, 50 mL min⁻¹, 1 atm, 600 °C.

calcination under air atmosphere formed a structure similar to that of Fe2/SiO2 treated by the reaction gas (Fig. S4a, S5a and b†).45 In our previous report on Fe2/SiO2 (used for CH4 oxidation without precalcination), 45 it was revealed that Fe2 was decomposed into FeOx subnanoclusters (<1 nm) and WO_x nanoclusters (approximately 3 nm) on SiO₂ under CH₄ oxidation conditions at 600 °C. Various control experiments characterizations also revealed that the subnanoclusters were the active species. Furthermore, the FeO_x subnanoclusters were dispersed on the WO_x nanoclusters, and as a result, excessive aggregation and deactivation of the effective FeO_x subnanoclusters were suppressed. Thus, a similar structure was expected to form over Fe4/SiO₂. The Fe K-edge XANES spectra of 35Fe2/SiO2 and Fe4/SiO2 suggested that they also had similar structures to the previously reported Fe2/SiO₂ (Fig. S4b†).⁴⁵ On the other hand, we found that the peak intensity corresponding to the Fe-O bond in the first coordination sphere for Fe2/SiO2 and 35Fe2/SiO2 was higher than that for Fe4/SiO2 in the Fe K-edge EXAFS spectra (Fig. S5c and d†). This may be due to the larger number of WO_x nanoclusters near the FeO_x subnanoclusters in $Fe2/SiO_2$ and 35Fe2/SiO2 than that in Fe4/SiO2 since the amount of tungsten per iron in Fe2 was larger than that in Fe4. Furthermore, in the Raman spectrum of Fe2/SiO2, three main vibrational peaks were observed, corresponding to v(W=0) at 980–960 cm⁻¹, ν (W–O_a–W) at 840–780 cm⁻¹, and ν (W–O_b–W) at 720-680 cm⁻¹ (Fig. S6†).^{59,60} The terminal oxo species (W=O) corresponds to an isolated tungsten species. In contrast, in the spectrum of Fe4/SiO2, two peaks corresponding to $v(W-O_a-W)$ and $v(W-O_b-W)$ were observed and no $\nu(W=O)$ peak was observed. Considering that Fe2/SiO₂ and Fe4/SiO₂ contained almost the same amount of tungsten, 5.3 wt% and 4.5 wt%, respectively, there was a sufficient amount of WO_x species near the FeO_x subnanoclusters. The remaining WO_x species on Fe2/SiO₂ formed the isolated tungsten species, possibly because the amount of tungsten per iron of Fe2 was larger than that of Fe4. Thus, these results were consistent with the EXAFS spectra (Fig. S5†). Therefore, the difference in the iron oxide subnanocluster structures formed

on SiO2 was reflected in the catalytic performance. In addition, 35Fe2/SiO₂ exhibited higher HCHO selectivity than Fe4/ SiO2, possibly due to the higher tungsten content and the presence of a larger number of tungsten oxide clusters, which inhibited the aggregation of FeO_x subnanoclusters. In summary, the iron oxide subnanocluster catalysts prepared from Fe1 and Fe2 exhibited the highest HCHO yield in CH4 oxidation.

Support effect

The catalytic performance for CH₄ oxidation was investigated for Fe2-based catalysts prepared using various supports at 600 °C under atmospheric pressure. The results and specific surface areas of supported are summarized in Table 2. The results indicated that higher specific surface areas of SiO2 supports resulted in higher CH₄ conversion (Table 2, entries 1-4). This suggested that a higher dispersion state of the Fe2 precursor on SiO₂ led to the generation of more dispersed iron oxide nanoclusters, resulting in increased CH4 conversion and HCHO yield. The supported Fe2 catalyst prepared using SiO2-Q6, which had a very high specific surface area of 386 m² g⁻¹, exhibited the highest CH₄ conversion and HCHO yield of 2.0% and 0.73%, respectively. In contrast, despite having lower specific surface areas than SiO2, the supported Fe2 catalysts prepared using Al₂O₃, ZrO₂, and CeO₂ exhibited higher CH₄ conversion, resulting in the production of only CO_x (Table 2, entries 5-7). In addition, the reaction was performed at 500 °C with Fe2/CeO2, but the selectivity to HCHO did not increase, although the CH4 conversion decreased (Table 2, entry 8). Similar catalytic performance has been reported for FePO₄ catalysts supported on Al₂O₃, SiO₂, TiO₂, and ZrO2, suggesting that the choice of support affects the overoxidation of HCHO.18

The surface acidity of these supports may promote the easy activation of CH₄, resulting in overoxidation of the products. 61 In addition, HAP, BN, Nb2O5, and TiO2 had smaller specific surface areas than SiO2, resulting in lower CH4 conversion when used as supports for Fe2 catalysts (Table 2,

Table 2 Catalytic performance over supported Fe2 catalysts on various supports

Entry	Support	BET surface	Conv. [%]	Sel. [%]			Yield [%]
		area [m ² g ⁻¹]	$\overline{\mathrm{CH_4}}$	НСНО	CO	CO_2	НСНО
1	SiO ₂ -Q50	67	0.71	57	35	8	0.41
2	SiO ₂ -Q30	100	0.96	43	44	13	0.42
3	SiO_2	278	1.4	41	45	14	0.57
4	SiO ₂ -Q6	386	2.0	36	46	18	0.73
5	Al_2O_3	158	6.1	_	82	18	_
6	ZrO_2	40	3.3	_	78	22	_
7	CeO_2	43	6.8	_	65	35	_
8	CeO_2^{-b}	43	0.79	_	69	31	_
9	Nb_2O_5	145	0.84	39	51	10	0.33
10	TiO_2	72	1.2	7	64	29	0.08
11	BN	6	0.28	57	24	19	0.16
12	HAP	9	0.12	67	n.d.	33	0.08

^a Reaction conditions: supported Fe2 catalyst (100 mg), CH₄:O₂:Ar = 2:1:7, 50 mL min⁻¹, 1 atm, 600 °C. ^b 500 °C.

entries 9-12). Therefore, the effect of the support on CH₄ oxidation was found to depend on the type of support and its specific surface area. Moreover, it was found that the Fe2 catalyst supported on SiO2, which was inert and had a high specific surface area, exhibited the highest HCHO yield.

Pyrolysis and oxidation tests of HCHO

As the support significantly affected the overoxidation of HCHO, pyrolysis and oxidation tests of HCHO on various supports and Fe2/SiO2 were investigated (Fig. 2). The results indicated that HCHO was decomposed by 80% or more even in the absence of catalysts at 600 °C (Fig. 2, entry 1). We found that HCHO pyrolysis was significantly suppressed in the presence of SiO₂ (Fig. 2, entry 2), and this suppression increased with the specific surface area of SiO₂ (Fig. 2, entry 3). Additionally, HCHO pyrolysis was suppressed by using Fe2/ SiO₂ (Fig. 2, entry 4). In contrast, HCHO pyrolysis on Al₂O₃ was promoted (Fig. 2, entry 5). HCHO overoxidation on SiO₂

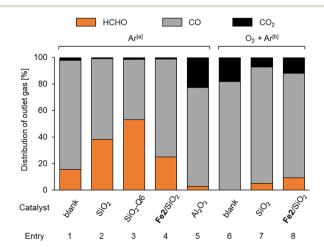


Fig. 2 HCHO pyrolysis and oxidation tests on various supports and Fe2/SiO₂. Reaction conditions: catalyst (100 mg), HCHO (1%), 1 atm, 600 °C. [a] Ar (50 mL min⁻¹), [b] O₂ (5 mL min⁻¹), Ar (45 mL min⁻¹).

was more suppressed even in the presence of O2 than in the absence of catalysts (Fig. 2, entries 6 and 7). Furthermore, HCHO conversion using Fe2/SiO2 was lower than that in the absence of catalysts (Fig. 2, entry 8). These findings suggested that the reason for the high HCHO yield when using Fe2/SiO2 was the suppression of HCHO pyrolysis by SiO2 and the use of Fe2 as a catalyst precursor. Meanwhile, even with an inert support such as SiO2, HCHO was sequentially oxidized in the presence of O2 at 600 °C. Since a significant amount of HCHO was thermally decomposed even at 600 °C, pyrolysis and oxidation tests without catalysts were conducted at various temperatures, as illustrated in Fig. 3. The results indicated that both pyrolysis and oxidation were suppressed at lower temperatures and that almost no pyrolysis or oxidation occurred at 500 °C. In our previous study, however, the CH₄ conversion over Fe2/SiO2 at 550 °C was lower than 1%, and the CH₄ conversion at 500 °C predicted based on the Arrhenius plots was below 0.4%. Therefore, FeO_x subnanocluster active site should be modified and improved to archive higher HCHO yield at 500 °C.

Proposed mechanism of the suppression of HCHO pyrolysis

To elucidate the mechanism by which pyrolysis of HCHO was promoted on Al₂O₃ and suppressed on SiO₂, in situ DRIFTS measurements were conducted at 600 °C. The DRIFTS spectra measured 60 min after evacuation following HCHO flow were presented in Fig. 4. The observed peaks of DRIFTS spectra and their assignments are summarized in Table S2.†62-64 Furthermore, time variation of in situ DRIFTS spectra of Al₂O₃ and SiO2 at 600 °C is illustrated in Fig. S7.† In the spectrum of Al₂O₃, three peaks observed at 1572, 2906, and 2997 cm⁻¹ were attributed to formate species (Fig. 4a). In addition, the peaks at 1630 and 3751 cm⁻¹ were attributed to H₂O and an aluminol group, respectively, indicating that these species possibly involved in the mechanism.

On the other hand, in the spectrum of SiO₂, two sharp peaks at 2959 and 2858 cm⁻¹ were observed (Fig. 4b). From the previously reported data summarized in Table. S2,† these

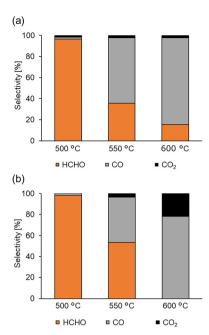


Fig. 3 Pyrolysis and oxidation tests without catalysts at various temperatures. Reaction conditions: no catalyst, 1 atm, HCHO (ca. 0.6%), (a) Ar (50 mL min⁻¹), (b) O_2 (5 mL min⁻¹), Ar (45 mL min⁻¹).

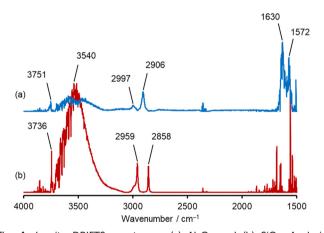


Fig. 4 In situ DRIFTS spectra on (a) Al₂O₃ and (b) SiO₂. Analysis conditions: HCHO/N2 mixture gas (20 mL min-1) flowed at 600 °C for 60 min. Then, the cell was evacuated, and DRIFTS spectra were measured 60 min later. The absorbance intensity on the vertical axis was Kubelka-Munk-transformed.

peaks can be attributed to dioxymethylene species. As well as Al₂O₃, the peaks derived from adsorbed H₂O (ref. 65) at 3540 cm⁻¹ and a silanol group⁶⁶ at 3736 cm⁻¹ were observed. These findings suggested that the adsorbed HCHO was decomposed into formate on Al₂O₃, while not on SiO₂.

We proposed a surface reaction mechanism of HCHO on Al2O3 and SiO2 based on the results of in situ DRIFTS, as displayed in Fig. 5. In the gas phase, CO is generated by HCHO pyrolysis via the formyl radical.⁶⁷ In the presence of Al₂O₃, HCHO reacts with aluminol groups on the surface to form dioxymethylene species. Then, the dioxymethylene spe-

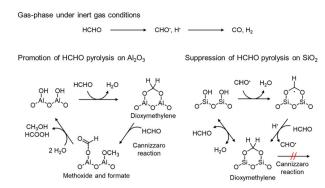


Fig. 5 Proposed mechanism of the suppression of HCHO pyrolysis on SiO₂ and the promotion of HCHO pyrolysis on Al₂O₃.

cies reacts with another HCHO to afford formate and methoxide species via the Cannizzaro reaction. 62 Then, these species further react with H2O to form formic acid and methanol; however, due to the high temperature of 600 °C, they are possibly sequentially oxidized to CO_x. Therefore, Al₂O₃ undesirably promotes the overoxidation of HCHO to CO_r.

In contrast, the dioxymethylene species adsorbed on SiO₂ do not cause the Cannizzaro reaction because SiO2 possesses no strong acidic or basic sites. Moreover, the formyl radical reacts with silanol species, and adsorbed on SiO₂. Then, the adsorbed radical species reacts with hydrogen atom or accepts hydrogen atom from HCHO to form dioxymethylene. Since the adsorbed dioxymethylene species reacts with H₂O to generate HCHO, we consider that HCHO pyrolysis can be suppressed on SiO₂. SiO₂-Q6 suppressed HCHO pyrolysis more than SiO₂ possibly because SiO₂-Q6 possessed more silanol groups to react with formyl radicals due to higher specific surface area.

Conclusions

The catalytic performance for CH₄ oxidation using SiO₂-supported metal oxide subnanocluster catalysts prepared from various POM precursors indicated that Fe1/SiO2 and Fe2/SiO2 exhibited the highest HCHO yields among oxide subnanocluster catalysts with 3d metal or Ce metal multinuclear oxide cores. In addition, we found that Fe2-based catalysts prepared using SiO2 with a high specific surface area improved both the CH₄ conversion and HCHO yield. Pyrolysis and oxidation tests of HCHO on various supports revealed that HCHO was almost completely decomposed even in the absence of catalysts at 600 °C. Moreover, the pyrolysis and overoxidation of HCHO were suppressed when SiO₂ was used as the support. In situ DRIFTS measurement under HCHO flow demonstrated that the pyrolysis and oxidation of HCHO were suppressed on SiO2, while the pyrolysis of HCHO was promoted on Al₂O₃. Thus, to achieve even higher HCHO yields in the future, it is necessary to suppress the pyrolysis and oxidation of HCHO in the gas phase. As pyrolysis and oxidation did not occur below 500 °C in the absence of catalysts, it is necessary to develop more effective catalysts that can lower the temperature of CH₄ activation.

Conflicts of interest

There are no conflicts to declare.

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