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Ligand functionalization on Zr-MOFs enables efficient visible-light-driven H_2O_2 evolution in pure water[†]

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Different ligand functionalized UiO-66 (UiO-66-X, X = OH, $(\text{OH})_2$ and NH_2) were prepared and then modified by ZnIn_2S_4 lamellas to form $\text{ZnIn}_2\text{S}_4/\text{UiO-66-X}$ heterostructures for visible-light-driven H_2O_2 evolution in pure water. The H_2O_2 yields using $\text{ZnIn}_2\text{S}_4/\text{UiO-66-NH}_2$, $\text{ZnIn}_2\text{S}_4/\text{UiO-66-(OH)}_2$ and $\text{ZnIn}_2\text{S}_4/\text{UiO-66-OH}$ are 799, 733 and 165 $\mu\text{mol L}^{-1}$, respectively, which are 9.5, 8.7 and 2.0 times that of $\text{ZnIn}_2\text{S}_4/\text{UiO-66}$. The high performance of $\text{ZnIn}_2\text{S}_4/\text{UiO-66-NH}_2$ and $\text{ZnIn}_2\text{S}_4/\text{UiO-66-(OH)}_2$ is ascribed to the benign visible-light response and Z-scheme heterostructures, and the H_2O_2 evolution abides by indirect O_2 reduction with $\cdot\text{O}_2^-$ as an intermediate species. Additionally, H_2O_2 yields using $\text{ZnIn}_2\text{S}_4/\text{UiO-66-NH}_2$ on tap water and Xuanwu Lake (Nanjing, China) water can be comparable to that on the above-deionized water. This study sheds light on the great promise of functionalized MOFs and their applications on green (photo)catalytic energy conversion.

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1. Introduction

As an eco-friendly and versatile oxidant, hydrogen peroxide (H_2O_2) is of great significance in various areas, such as disinfection, pulp bleaching, organic synthesis and effluent treatment.^{1,2} Additionally, on account of its high energy capacity, facile storage and transportation, H_2O_2 has been deemed as a promising and ideal liquid fuel.³ Therefore, the market demand for H_2O_2 all over the world is pretty huge and will continuously increase. Currently, the main strategy to produce H_2O_2 in industry is the anthraquinone method, which undergoes consecutive hydrogenation and oxidation reactions, inducing large amounts of toxic byproduct generation and high energy input.⁴ Within this context, photocatalytic evolution of H_2O_2 from O_2 reduction has drawn plenty of interest for the last few years.^{5,6} In this respect, much effort on traditional semiconductors like CdS ,⁷ $\text{g-C}_3\text{N}_4$,⁸ Bi_2MoO_6 (ref. 9) and BiVO_4 (ref. 3) as photocatalysts was devoted to H_2O_2 generation and indeed, their corresponding performances were admirable. Nonetheless, most of these reactions were performed in the presence of sacrificial agents (alcohols) and pure O_2 , which makes the consequent separation of H_2O_2 difficult and requires extra energy input,

consequently going against green-chemistry principles. Additionally, these photocatalysts normally have a low O_2 capture/adsorption capacity due to their limited surface areas.

Thanks to an inherent large surface area, high porosity, tailorable structure and semiconducting behavior,¹⁰⁻¹² metal-organic frameworks (MOFs) are vastly acclaimed in gas-associated photocatalysis, like photocatalytic CO_2 reduction¹³ and N_2 fixation.¹⁴ Whereas, MOFs utilized in H_2O_2 generation from O_2 photocatalytic reduction received only sporadic attention.¹⁵ Yamashita and co-workers constructed hydrophobic MIL-125-NH₂ *via* ligand¹⁶ or metal cluster¹⁷ alkylation and achieved H_2O_2 production in a benzyl alcohol/water two-phase system, and later perylenetetracarboxylic diimide was grafted on MIL-125-NH₂ for H_2O_2 production in water.⁵ Additionally, MOF/semiconductor hybrids like MIL-125-NH₂(TiO_2)/ Ti_3C_2 ,¹⁸ ZIF-8/g-C₃N₄ (ref. 19) and MIL-125-NH₂@ZnS (ref. 20) were also fabricated for H_2O_2 photocatalytic evolution.

Ligand decoration on MOFs, especially MOFs with terephthalic acid as ligands, could facilely and effectively regulate their physicochemical properties, conferring charming versatility.¹⁰ The common route for ligand decoration is a substitution by functional groups, such as electron-donating amino and hydroxyl groups. The reasons are as follows: (1) uncoordinated hydroxyl and amino groups could serve as additional active sites to participate in the adsorption and activation of reactants,^{21,22} and may also induce interactions with guest molecules when using MOFs as hosts to prepare adsorbents or catalysts;²³ (2) amino or

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hydroxyl substituents could serve as auxochromic and bathochromic groups in aromatic rings, rendering a redshift of light absorption;^{24,25} (3) amino and hydroxyl groups are hydrophilic, which could promote the reaction in water. Given such a circumstance, it is believed that amino or hydroxyl-functionalized MOFs could possess great potential for photocatalytic H_2O_2 generation under visible light.

Considering the unprecedented thermal, chemical and mechanical stability, UiO-66, a zirconium-based MOF (each $Zr_6O_4(OH)_4$ cluster coordinates with 12 organic ligands),²⁶ was selected as the matrix. A series of UiO-66-X (X = OH, $(OH)_2$ and NH_2) was synthesized through a facile solvothermal method and then decorated by $ZnIn_2S_4$ layers for visible-light-driven H_2O_2 evolution in pure water and ambient air. $ZnIn_2S_4$ is a typical layered semiconductor with favorable chemical and photostability, enviable visible-light absorption and delicate band configuration with a strong reduction ability that can reduce O_2 into H_2O_2 .^{27,28} Thereby, $ZnIn_2S_4$ was adopted to sensitize UiO-66-X and constructed heterostructures to optimize the photocatalytic performances. On account of the favorable visible-light capture and unique Z-scheme heterostructures, $ZnIn_2S_4$ modified UiO-66- NH_2 and UiO-66- $(OH)_2$ performed remarkable activities for photocatalytic H_2O_2 evolution. This study would motivate the development of MOF functionalization in green photo- or electrocatalysis.

2. Experimental section

2.1 Materials

Zirconium chloride ($ZrCl_4$, 98%), 2-hydroxyterephthalic acid ($\geq 98\%$) and 2,5-dihydroxyterephthalic acid ($\geq 98\%$) were bought from Aladdin Industrial Company. *N,N*-Dimethylformamide (DMF), acetic acid, terephthalic acid ($\geq 99\%$), ethanol, zinc acetate dihydrate ($\geq 99\%$), $Zn(CH_3COO)_2 \cdot 2H_2O$ and thioacetamide (TAA, $\geq 99\%$) were obtained from Sinopharm Chemical Reagent. 2-Aminoterephthalic acid (99%) was purchased from Sigma-Aldrich. Indium chloride tetrahydrate ($InCl_3 \cdot 4H_2O$, $\geq 99\%$) was bought from Shanghai Macklin.

2.2 Preparation of functionalized Zr-MOFs

UiO-66-OH, UiO-66- $(OH)_2$ and UiO-66- NH_2 were prepared based on our previous report.²⁶ Typically for UiO-66- NH_2 , 190.3 mg of $ZrCl_4$ (0.8 mmol), 147 mg of 2-aminoterephthalic acid (0.8 mmol) and 9.6 g of acetic acid (160 mmol) were added into 81.7 mL of DMF. After ultrasound treatment for 20 min and stirring for 2 h, the solution was transferred into a 150-mL autoclave and kept at 120 °C for 24 h. The products were collected by centrifugation, rinsed with DMF and methanol and finally dried at 80 °C overnight. UiO-66-OH, UiO-66- $(OH)_2$ and UiO-66 were prepared under the same conditions except that 2-aminoterephthalic acid was replaced by equimolar 2-hydroxyterephthalic acid, 2,5-dihydroxyterephthalic acid and terephthalic acid, respectively.

2.3 Zr-MOFs modified by $ZnIn_2S_4$ lamellas

In a typical synthesis, 169.3 mg of functionalized Zr-MOFs were dispersed in 15 mL of ethanol with ultrasound for 0.5 h. Meanwhile, 87.8 mg of $Zn(Ac)_2 \cdot 2H_2O$, 234.6 mg of $InCl_3 \cdot 4H_2O$ and 120.2 mg of TAA were dissolved into 15 mL of deionized water with stirring for 0.5 h. Subsequently, the above aqueous solution was dropped slowly into the ethanol suspension. After stirring for 2 h, the suspension was transferred into a 50 mL autoclave and kept at 180 °C for 24 h. The products were gathered by centrifugation, rinsed with water and ethanol and finally dried at 80 °C overnight. The obtained $ZnIn_2S_4$ /Zr-MOFs samples (50 wt% of $ZnIn_2S_4$ /Zr-MOFs) were named Z-UOH, Z-U(OH) $_2$ and Z-UN, respectively. As comparisons, $ZnIn_2S_4$ /UiO-66 (Z-U) and bare $ZnIn_2S_4$ nanosheets were also prepared under the same conditions with UiO-66 and without any Zr-MOFs, respectively.

2.4 Characterizations

X-ray diffraction (XRD) patterns were obtained on a Rigaku Ultima IV instrument. N_2 adsorption-desorption isotherms were measured using a Micromeritics TriStar II equipment, and corresponding specific surface areas were determined by the Brunauer-Emmett-Teller (BET) method. Scanning electron microscopy (SEM) and energy dispersive X-ray spectrometry (EDS) mapping were carried out with a Hitachi Regulus 8100 instrument. Transmission electron microscopy (TEM) was performed using JEOL JEM-1400 and JEM-2100 instruments. UV-vis diffuse reflectance spectra (DRS) were recorded on a Shimadzu UV-2600 spectrophotometer. Fourier transform infrared spectroscopy (FTIR) was conducted on a Thermo Electron Nicolet-360 instrument. Water contact angles were measured using a contact meter (KRUSS CAT, Germany). X-ray photoelectron spectroscopy (XPS) was executed on an AXIS UltraDLD instrument. Photocurrent, electrochemical impedance spectroscopy (EIS) and Mott-Schottky plots were tested using a CHI-760E electrochemical workstation. Electron paramagnetic resonance (EPR) was performed using a Bruker EMXPLUS spectrometer.

2.5 Photocatalytic synthesis of H_2O_2

In a typical reaction setup, 20 mg of photocatalyst powder was dispersed in 40 mL of deionized water with stirring for 30 min in the dark. The photocatalytic reaction was initiated by irradiation of a 300 W xenon lamp (400–780 nm, 350 mW cm^{-2}) in ambient air. At designated time points, a certain suspension was taken and filtrated. H_2O_2 concentration was determined using the UV-2600 spectrophotometer at 350 nm according to iodometry (details in ESI†). The apparent quantum yield (AQY) at different irradiation wavelengths was determined using:

$$AQY = \frac{2 \times \text{the number of generated H}_2\text{O}_2}{\text{the number of incident photos}} \times 100\%$$

Meanwhile, the solar-to-H₂O₂ (STH) conversion efficiency was determined using:

$$STH = \frac{\Delta G(\text{H}_2\text{O}_2) \times [\text{H}_2\text{O}_2] \times V}{I \times A \times t} \times 100\%$$

where $\Delta G(\text{H}_2\text{O}_2)$ is the free energy for H₂O₂ formation (117 kJ mol⁻¹), I is the light intensity (100 mW cm⁻²) of simulated solar light, V is the volume of suspension, A is the irradiation area, and t is the reaction time.

3. Results and discussion

3.1 Photocatalytic performance on H₂O₂ evolution

The prepared ZnIn₂S₄/UiO-66-X photocatalysts were evaluated to catalyze H₂O₂ generation under visible light in pure water and ambient air. UiO-66-NH₂ and ZnIn₂S₄ have minimal H₂O₂ evolution after 2 h illumination (9 μmol L⁻¹ for UiO-66-NH₂ and 23 μmol L⁻¹ for ZnIn₂S₄, Fig. 1a), presumably due to the limited active sites and sluggish charge separation. After ZnIn₂S₄ *in situ* growth on UiO-66-NH₂, H₂O₂ evolution experienced a pronounced improvement for all ZnIn₂S₄/UiO-66-NH₂ hybrids without exception. Thereinto, the ZnIn₂S₄/UiO-66-NH₂ hybrid with 50 wt% of UiO-66-NH₂ performed the highest H₂O₂ evolution. Thus, this mass ratio was selected to prepare other ZnIn₂S₄/UiO-66-X photocatalysts, and their corresponding activities for H₂O₂ generation were displayed in Fig. 1b. The performance of Z-U is inferior. Whereas, the H₂O₂ yields catalyzed by Z-UN, Z-U(OH)₂ and

Z-UOH are 799, 733 and 165 μmol L⁻¹, which are 9.5, 8.7 and 2.0-fold enhancements compared with that of Z-U, respectively, affirming that the functionalization of UiO-66 by amino or hydroxyl groups could switch on the efficient photocatalytic generation of H₂O₂. Generally, the addition of sacrificial agents like alcohols and extra O₂ input are two prerequisites for efficient and continuous H₂O₂ generation by photocatalysis, yet is going against the green-chemistry principles. Comparatively, the H₂O₂ yield using Z-UN in pure water and ambient air can be even preferable to those of most MOF-based photocatalysts reacting with sacrificial agents and extra O₂ input (Table S1†), and Z-UN possesses favorable cycling performance (Fig. S1†). To study the effect of H₂O₂ decomposition, decomposition experiments with an initial H₂O₂ concentration of 1000 μmol L⁻¹ were conducted over ZnIn₂S₄ and Z-UN (Fig. S2†). The H₂O₂ decomposition rate is greater than the generation rate after 0.5 h illumination over ZnIn₂S₄. While for Z-UN, the generation rate is greater than the decomposition rate until 1.5 h illumination and close to the decomposition rate after 1.5 h illumination, demonstrating that the combination with MOFs promotes the H₂O₂ formation apparently.

For comparison, 10 vol% ethanol was utilized to replace deionized water as a solution for photocatalytic H₂O₂ evolution (Fig. S3†), and the yield was boosted to 918 μmol L⁻¹ using Z-UN on account of the introduction of sacrificial agents. Of note, the H₂O₂ yields using tap water and Xuanwu Lake (Nanjing, China) water can be comparable to that of using deionized water (Fig. S3†), which further improves the economy and sustainability. Meanwhile, reactions in O₂ and N₂ were also conducted using Z-UN (Fig. 1c). By comparison, O₂ promoted the H₂O₂ generation while N₂ dramatically inhibited it, shedding light on that H₂O₂ photocatalytic generation stems mainly from O₂ reduction. H₂O₂ evolution in air is close to that in O₂, implying that enough O₂ in air was adsorbed and attached on Z-UN. The AQYs of photocatalytic H₂O₂ evolution using Z-UN at wavelengths of 360, 400, 420 and 550 nm in pure water and ambient air are 1.79, 1.67, 1.3 and 0.15, respectively, which is in accord with the UV-vis DRS trend of Z-UN (Fig. 1d), indicating good light utilization. Additionally, H₂O₂ generation can still reach 717 μmol L⁻¹ under simulated solar light (light filter: AM1.5, light intensity: 100 mW cm⁻², Fig. S4†), and the corresponding STH is calculated as 0.024%.

3.2 Structure characterizations of the prepared ZnIn₂S₄/UiO-66-X

The preparation processes of the ZnIn₂S₄/UiO-66-X hybrids were displayed in Fig. 2a, and the only difference in these processes is the organic ligands of UiO-66-X. The XRD pattern of ZnIn₂S₄ has four characteristic peaks at 21.7, 27.5, 30.3 and 47.1°, ascribing to the (006), (102), (104) and (110) crystal planes of hexagonal ZnIn₂S₄ (Fig. 2b).²⁹ All functionalized UiO-66 has similar characteristic peaks compared with those of UiO-66, indicating that they are topologically equivalent to

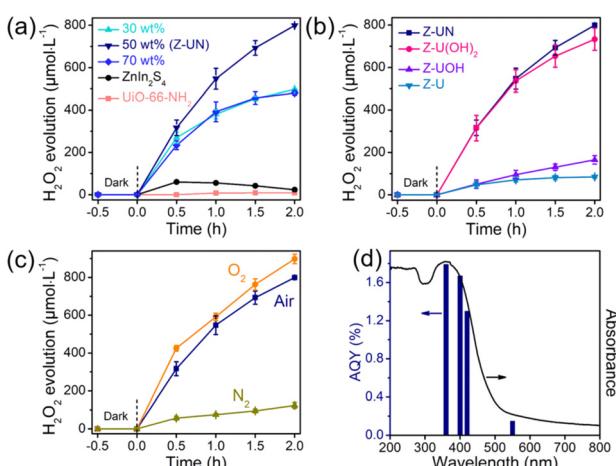


Fig. 1 Photocatalytic H₂O₂ evolution over ZnIn₂S₄, UiO-66-NH₂ and ZnIn₂S₄/UiO-66-NH₂ hybrids with 30, 50, 70 wt% of UiO-66-NH₂ (a), Z-U, Z-U(OH)₂, Z-UOH, Z-UN (b), photocatalytic H₂O₂ evolution over Z-UN in O₂, air and N₂ (c), AQYs of photocatalytic H₂O₂ evolution using Z-UN at different wavelengths and UV-vis DRS of Z-UN (d). Reaction conditions: 20 mg photocatalysts, 40 mL deionized water, 400–780 nm illumination for (a)–(c), ambient air for (a), (b) and (d), ambient temperature.

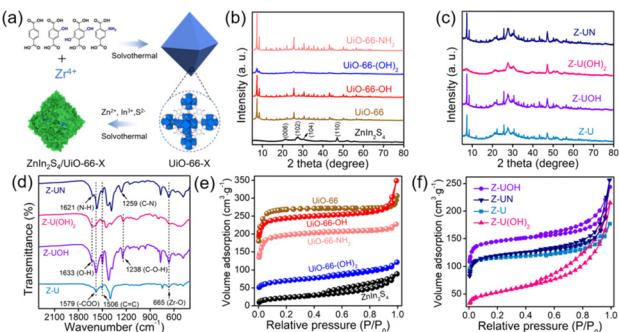


Fig. 2 Preparation diagram of ZnIn₂S₄/UiO-66-X hybrids (a), XRD patterns (b and c) and N₂ adsorption-desorption isotherms (e and f) of UiO-66-NH₂, UiO-66-(OH)₂, UiO-66-OH, UiO-66, ZnIn₂S₄, Z-UN, Z-U(OH)₂, Z-UOH and Z-U, FTIR spectra of Z-UN, Z-U(OH)₂, Z-UOH and Z-U (d).

the face-centered cubic lattice of the UiO-66 structure and this result is in line with the previous reports.^{30,31} Evidently, the peaks attributed to ZnIn₂S₄ and UiO-66-X appeared on each pattern of the ZnIn₂S₄/UiO-66-X composites, demonstrating their successful integrations (Fig. 2c and S5†). The weak XRD peaks of UiO-66-(OH)₂ are triggered by its low crystallinity. In addition, the good stability of Z-UN before and after photocatalytic H₂O₂ evolution was also verified by XRD (Fig. S6†). The functional groups were identified by FTIR, and the vibration bands between UiO-66-X and the corresponding ZnIn₂S₄/UiO-66-X are similar (Fig. S7, S8† and 2d). All hybrids show the common bands at 1579, 1506 and 665 cm⁻¹, which are initiated by the vibrations of carboxylate groups from the terephthalic ligand, C=C bonds from the benzene ring and Zr-O bonds from Zr₆ clusters, respectively.²⁴ Two bands at 1621 and 1259 cm⁻¹ in the spectrum of Z-UN separately correspond to the N-H bending vibration and C-N stretching vibration, demonstrating the presence of amino groups.³² In the case of Z-U(OH)₂ and Z-UOH, the ν (O-H) bands and C-O stretching of hydroxyl on the benzene rings could be found at 1633 and 1238 cm⁻¹.³³ The XRD and FTIR results collectively validate the successful preparation of functionalized UiO-66 decorated by ZnIn₂S₄. It is known that the hydroxyl and amino functionalization would enhance the hydrophilicity. To determine this, water contact angles of Z-U, Z-UOH, Z-U(OH)₂ and Z-UN were measured (Fig. S9†). All samples performed good hydrophilicity because of plenty of hydroxyl groups in the metal cluster (Zr₆O₄(OH)₄) of Zr-MOFs. Even so, the hydroxyl and amino functionalization could still improve the hydrophilicity to a certain degree, especially for Z-U(OH)₂. The reinforced hydrophilicity is conducive to the H₂O₂ generation in water.

MOFs are known for their large surface area and high porosity,³⁴ which are advantageous for mass transfer, O₂ adsorption and attachment. The BET surface areas of UiO-66, UiO-66-OH, UiO-66-NH₂ and UiO-66-(OH)₂ were determined as 928, 745, 587 and 221 m² g⁻¹, respectively (Fig. 2e). In contrast with UiO-66-X, ZnIn₂S₄ possesses a small surface

area of 81 m² g⁻¹. When ZnIn₂S₄ integrates with UiO-66-X, the surface areas could achieve noticeable magnifications to 352, 448, 361 and 178 m² g⁻¹ for Z-U, Z-UOH, Z-UN and Z-U(OH)₂, respectively (Fig. 2f).

The effect of UiO-66 functionalization on the microstructure was observed by SEM. First of all, ZnIn₂S₄ is a distinct layer structure while it has a serious aggregation (Fig. 3a).³⁵ Under the same synthetic conditions except for the organic ligand categories, UiO-66-NH₂ (Fig. 3b), UiO-66-OH (Fig. 3c) and UiO-66 (Fig. 3d) present regular and uniform octahedrons, while UiO-66-(OH)₂ has a tiny nanoparticle morphology caused by the low crystallinity (Fig. 3e).²⁴ In addition, the particle sizes follow a sequence of UiO-66-(OH)₂ (~40 nm) < UiO-66-NH₂ (~450 nm) < UiO-66-OH (~650 nm) < UiO-66 (~950 nm), which probably induced by the steric-hindrance effect of functional groups during the crystal nucleus growth. As to Z-UN (Fig. 3f), Z-UOH (Fig. 3g) and Z-U (Fig. 3h), the corresponding UiO-66-NH₂, UiO-66-OH and UiO-66 octahedrons were evenly wrapped by ZnIn₂S₄ nanosheets. Conversely, for Z-U(OH)₂, UiO-66-(OH)₂ nanoparticles were loaded on ZnIn₂S₄ layers (Fig. 3i). Besides, the growth of ZnIn₂S₄ nanosheets on UiO-66-NH₂ could be tuned (Fig. S10†) and the morphology of Z-UN relatively remained unchanged after the photocatalytic H₂O₂ generation (Fig. S11†).

To further characterize the microstructure of Z-UN, TEM images of UiO-66-NH₂, ZnIn₂S₄ and Z-UN were conducted. The regular octahedron of UiO-66-NH₂ (Fig. 4a) and ultra-thin layers of ZnIn₂S₄ (Fig. 4b) could be indubitably reflected. What's more, their spatial distribution and interface contact in Z-UN can be clearly observed (Fig. 4c and d). The lattice fringe of 0.295 nm corresponding to the ZnIn₂S₄ (104) crystal plane in high-resolution TEM of Z-UN further validates the outer layer of ZnIn₂S₄ (Fig. 4d). Likewise, the (110), (104) and (102) crystal planes of ZnIn₂S₄ could be also indicated by the diffraction

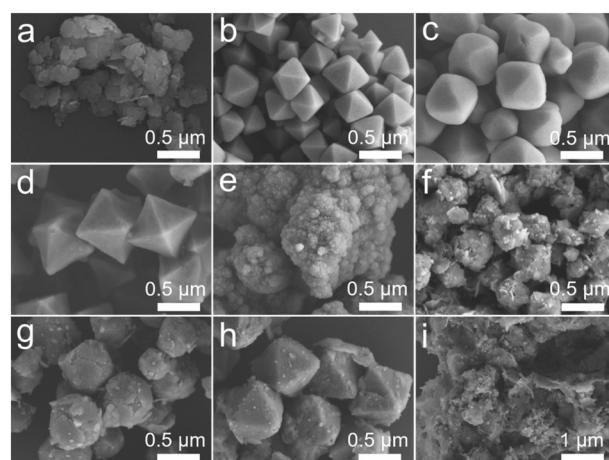


Fig. 3 SEM images of ZnIn₂S₄ (a) UiO-66-NH₂ (b), UiO-66-OH (c), UiO-66 (d), UiO-66-(OH)₂ (e), Z-UN (f), Z-UOH (g), Z-U (h) and Z-U(OH)₂ (i).

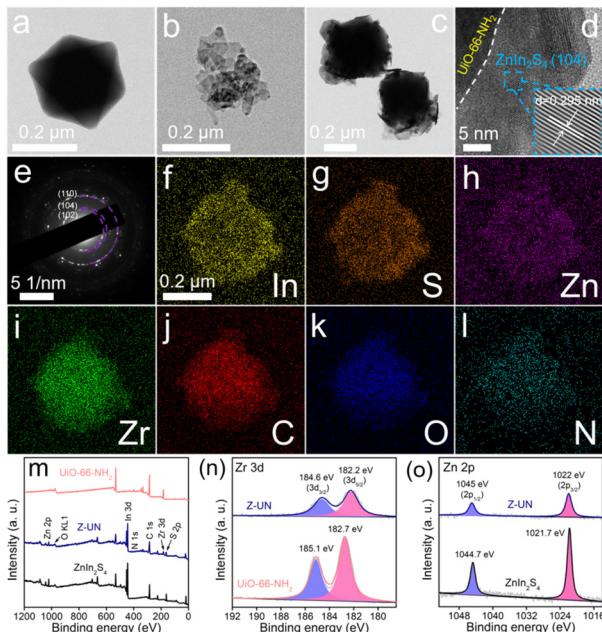


Fig. 4 TEM images of UiO-66-NH_2 (a), ZnIn_2S_4 (b) and Z-UN (c and d). SAED pattern (e) and EDS mapping images (f–l) of Z-UN, XPS spectra of UiO-66-NH_2 , Z-UN and ZnIn_2S_4 : survey (m), Zr 3d (n) and Zn 2p (o).

fringes of selected area electron diffraction (SAED) patterns (Fig. 4e). EDS mapping images of Z-UN uncover the uniform distribution of In, S, Zn, Zr, C, O and N elements (Fig. 4f–l), and such results indicate the formation of homogeneous interfacial junctions between ZnIn_2S_4 nanosheets and UiO-66-NH_2 octahedrons.

The elemental composition of In, S, Zn, Zr, C, O and N for Z-UN was further reflected by the XPS survey spectrum (Fig. 4m). The Zr 3d curve of UiO-66-NH_2 can be split into two peaks at 182.7 and 185.1 eV (Fig. 4n), which are assigned to Zr $3d_{5/2}$ and $3d_{3/2}$, respectively. Nevertheless, the two related peaks of Z-UN underwent negative shifts to 182.2 and 184.6 eV, implying a net gain of electrons for Zr^{4+} . In contrast to the peak shifts of Zr^{4+} , peaks of Zn $2p_{3/2}$ and $2p_{1/2}$ experienced high-frequency shifts from 1021.7 and 1044.7 eV of ZnIn_2S_4 to 1022 and 1045 eV of Z-UN, respectively (Fig. 4o), suggesting an electron loss of Zn^{2+} . Based on the gleaned above results, it is inferred that the electrons moved from ZnIn_2S_4 to UiO-66-NH_2 with their integration, which demonstrates the successful construction of an internal electric field between ZnIn_2S_4 and UiO-66-NH_2 .

3.3 Mechanism discussion

The functionalization of MOFs has a huge effect on their light absorption because the functional groups could serve as auxochromic and bathochromic groups in aromatic rings.^{24,25} To determine this, UV-vis DRS of UiO-66-X were performed first (Fig. 5a). UiO-66 has the smallest light absorption mainly in 200–320 nm that is aroused by the ligand-to-metal charge transfer, suggesting the bonding between carboxylate oxygen

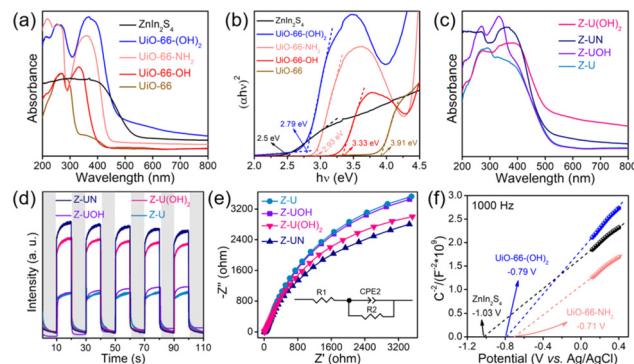


Fig. 5 UV-vis DRS (a) and $(\alpha h v)^2$ versus $h v$ curves (b) of ZnIn_2S_4 , UiO-66-(OH)_2 , UiO-66-NH_2 , UiO-66-OH and UiO-66 , UV-vis DRS (c), photocurrent spectra (d) and EIS Nyquist plots (e) of Z-UN, Z- U(OH)_2 , Z-UOH and Z-U, Mott-Schottky plots of ZnIn_2S_4 , UiO-66-(OH)_2 and UiO-66-NH_2 (f).

and metal.³⁶ By comparison, the light absorption edges of all functionalized UiO-66 underwent redshifts, especially for UiO-66-NH_2 and UiO-66-(OH)_2 shifting to the visible-light range, which is initiated by the conjugated π electron transition from the amino or hydroxyl-auxochromic chromophores to the Zr centers.³⁷ Correspondingly, the bandgaps of UiO-66 , UiO-66-OH , UiO-66-(OH)_2 and UiO-66-NH_2 are 3.91, 3.33, 2.79 and 2.93 eV, respectively, according to the curves of $(\alpha h v)^2$ versus $h v$ (Fig. 5b). It should be stressed that ZnIn_2S_4 has an outstanding visible-light absorption with a narrow bandgap of 2.5 eV, and all $\text{ZnIn}_2\text{S}_4/\text{UiO-66-X}$ heterostructures exhibit striking visible-light absorption with edges about 500 nm similar to that of ZnIn_2S_4 (Fig. 5c and S12†).

With the aim to validate the charge separation and transfer over $\text{ZnIn}_2\text{S}_4/\text{UiO-66-X}$ heterostructures, the photocurrent and EIS were tested. At first, in comparison with ZnIn_2S_4 and UiO-66-NH_2 , Z-UN performed with a stronger photocurrent signal and smaller arc radius of the EIS Nyquist plot (Fig. S13 and S14†), which signifies the reinforced separation of photo-excited charge carriers and diminished resistance of charge transfer. In the case of $\text{ZnIn}_2\text{S}_4/\text{UiO-66-X}$ heterostructures, Z-UN and Z- U(OH)_2 possess more efficient charge separation and lower resistance of charge transportation than those of Z-UOH and Z-U counterparts (Fig. 5d and e), which accord well with the corresponding results of photocatalytic H_2O_2 generation. The charge transfer routes meet the thermodynamic requirements, thus the band configurations of ZnIn_2S_4 , UiO-66-(OH)_2 and UiO-66-NH_2 are indispensable to be uncovered. To this end, Mott-Schottky plots of ZnIn_2S_4 , UiO-66-(OH)_2 and UiO-66-NH_2 were monitored at 1000 and 2000 Hz (Fig. 5f and S15†). The flat-band potentials are -1.03 , -0.79 and -0.71 V vs. Ag/AgCl (-0.833 , -0.593 , -0.513 V vs. NHE, $E_{\text{NHE}} = E_{\text{Ag/AgCl}} + 0.197$ V) for ZnIn_2S_4 , UiO-66-(OH)_2 and UiO-66-NH_2 , respectively. It is regarded that the conduction band (CB) or lowest unoccupied molecular orbital (LUMO) potentials for n-type semiconductors are more negative by 0.1 V than their flat-band potentials.³⁸ As such, the CB potential of ZnIn_2S_4 ,

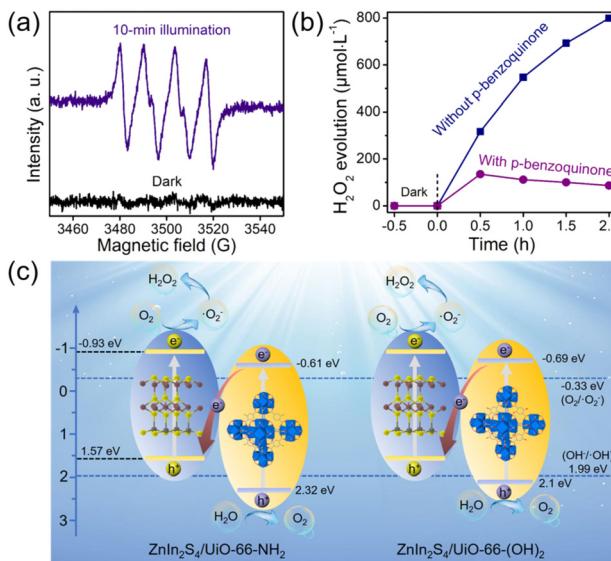


Fig. 6 EPR spectra of DMPO-·O₂⁻ adducts using Z-UN at dark and 10 min illumination (a), photocatalytic H₂O₂ evolution using Z-UN with and without *p*-benzoquinone (b), the proposed mechanism for photocatalytic H₂O₂ evolution using Z-UN and Z-U(OH)₂ (c).

LUMO potentials of UiO-66-(OH)₂ and UiO-66-NH₂ are about -0.93, -0.69 and -0.61 eV, corresponding to the VB or highest occupied molecular orbital (HOMO) potentials of 1.57, 2.1 and 2.32 eV, respectively ($E_{VB\ or\ HOMO} = E_{Bandgap} + E_{CB\ or\ LUMO}$).

Although the related band configurations were known, the routes of photocatalytic H₂O₂ generation must be confirmed before elucidating the whole mechanism. Generally, it is deemed that the photocatalytic O₂ reduction into H₂O₂ experiences a direct (O₂ + 2H⁺ + 2e⁻ → H₂O₂) or indirect (O₂ + e⁻ → ·O₂⁻, ·O₂⁻ + e⁻ + 2H⁺ → H₂O₂) way.^{2,20} To unveil this, the generation of ·O₂⁻ under visible-light illumination using Z-UN was first evidenced by EPR (Fig. 6a). Subsequently, *p*-benzoquinone was employed as a scavenger of ·O₂⁻ in the H₂O₂ evolution process using Z-UN. As illustrated in Fig. 6b, it is noticeable that the H₂O₂ yield has a stark decrease in the presence of *p*-benzoquinone, affirming that ·O₂⁻ serves as a significant intermediate species for H₂O₂ generation. Likewise, the same trend also happened on Z-U(OH)₂ (Fig. S16†). These results collectively suggest that photocatalytic H₂O₂ generation using Z-UN and Z-U(OH)₂ underwent an indirect O₂ reduction reaction.

At last, the proposed mechanism for photocatalytic H₂O₂ evolution using ZnIn₂S₄/UiO-66-X was discussed based on the above analyses. By means of the narrow bandgaps, ZnIn₂S₄ (2.5 eV), UiO-66-NH₂ (2.93 eV) and UiO-66-(OH)₂ (2.79 eV) are capable to be excited to generate electrons and holes by visible-light illumination (Fig. 6c). At this time, according to the band configuration of each monomer, the transfer routes of charge carriers should be determined using Z-UN and Z-U(OH)₂, namely, conforming to type-II or Z-scheme heterostructure. To this end, ·OH generation was tested since the corresponding potential ($E(OH^-/OH) = 1.99$ eV (ref. 39))

is in between the VB potential of ZnIn₂S₄ (1.57 eV) and HOMO potential of UiO-66-NH₂ (2.32 eV) or UiO-66-(OH)₂ (2.1 eV). As a consequence, ·OH generation was demonstrated by EPR spectra using Z-UN (Fig. S17†) and Z-U(OH)₂ (Fig. S18†), suggesting that the holes on the HOMO of UiO-66-NH₂ and UiO-66-(OH)₂ were kept. Thereby, Z-scheme heterostructures were formed in Z-UN and Z-U(OH)₂. While for Z-U and Z-UOH, the electrons would transfer from the CB of ZnIn₂S₄ to LUMO of UiO-66 or UiO-66-OH (Fig. S19†) because of no generation of electrons and holes by UiO-66 and UiO-66-OH. Ultimately, electrons on the CB of ZnIn₂S₄ in Z-UN and Z-U(OH)₂ possess sufficient abilities to reduce O₂ into ·O₂⁻ ($E(O_2^-/O_2) = -0.33$ eV) and further evolve into H₂O₂. Appreciably, the conspicuous visible-light absorption and unique Z-scheme heterostructures decide the superiority of Z-UN and Z-U(OH)₂ on photocatalytic H₂O₂ evolution in comparison with those of Z-U and Z-UOH. That is to say, the functionalization of UiO-66 does enable the efficient photocatalytic generation of H₂O₂.

4. Conclusions

To sum up, various ZnIn₂S₄/UiO-66-X heterostructures were fabricated for photocatalytic H₂O₂ generation under visible light in pure water. The H₂O₂ yields using ZnIn₂S₄/UiO-66-NH₂, ZnIn₂S₄/UiO-66-(OH)₂ and ZnIn₂S₄/UiO-66-OH are 799, 733 and 165 μmol L⁻¹, respectively, while the corresponding yield using ZnIn₂S₄/UiO-66 is only 84 μmol L⁻¹. The outstanding performances by using ZnIn₂S₄/UiO-66-NH₂ and ZnIn₂S₄/UiO-66-(OH)₂ are attributed to the favorable visible-light absorption and Z-scheme heterostructures. Furthermore, it is concluded that the H₂O₂ generation using ZnIn₂S₄/UiO-66-NH₂ and ZnIn₂S₄/UiO-66-(OH)₂ followed indirect O₂ reduction with ·O₂⁻ as the intermediate species. This study could provide new insight into MOF functionalization and a new doorway to green energy conversion by (photo)catalysis.

Author contributions

JQ: conceptualization, investigation, funding acquisition, and writing – original draft; LZ: data curation, formal analysis, investigation, and writing – original draft; GX, DD, YT: investigation; JY: conceptualization, resources, funding acquisition, and writing – review & editing.

Conflicts of interest

There are no conflicts of interest.

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