


 Cite this: *Phys. Chem. Chem. Phys.*, 2023, 25, 30469

# A microfluidic labyrinth self-assembled by a chemical garden†

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Chemical gardens, self-assembling precipitates that spontaneously form when a metal salt is added to a solution of another precipitating anion, are of interest for various applications including producing reactive materials in controlled structures. Here, we report on two chemical garden reaction systems ( $\text{CuCl}_2$  and  $\text{Cu}(\text{NO}_3)_2$  seed crystals submerged in sodium silicate) that produced self-assembled microfluidic labyrinths in a vertical 2D Hele-Shaw reactor. The formation of labyrinths as well as the specific growth modes of the precipitate were dependent on the silicate concentration:  $\text{CuCl}_2$  labyrinths formed only at 3 and 4 M silicate and  $\text{Cu}(\text{NO}_3)_2$  labyrinths formed only at 4 and 5 M silicate. The labyrinth structures contained silicate on the exterior and crystalline material interpreted as hydrated minerals from the metal salt in their interiors. The bubble-guided tubes that form labyrinths can be controlled by changing the angle of the 2D reaction cell; this suggests that future experiments of this type could form self-organizing structures with controlled composition and orientation for use in microfluidics and various materials science applications.

 Received 23rd June 2023,  
 Accepted 26th October 2023

DOI: 10.1039/d3cp02929h

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## Introduction

Mazes and labyrinths – sometimes defined as a unicursal maze having only a single path<sup>1</sup> – in a fluid medium delineated by solid walls have been of great interest in microfluidics. A number of chemical-, physical- and biological-based maze solving techniques have been reported.<sup>2,3</sup> Chemical waves can find the optimal path,<sup>4</sup> as can chemotactic droplets.<sup>5,6</sup> Marangoni flow in a maze can solve the maze.<sup>7,8</sup> A liquid metal droplet in an electric field<sup>9</sup> is another maze-solving system. The nematode *Caenorhabditis elegans* has been investigated learning a microfluidic maze.<sup>10</sup> On the other hand, self-assembled mazes have been reported in slow drainage of a granular-fluid system in two-dimensional confinement.<sup>11</sup> Labyrinth-type domain structures were also observed in perovskite oxide films.<sup>12</sup> Moreover, maze structures have been noted to self-organize by wrinkling or through the fingering instabilities of macroscopic domains of magnetic so-called ferrofluids.<sup>13,14</sup>

Inorganic self-assembly has also been well studied in chemobronic systems,<sup>20</sup> including chemical gardens, which are precipitates that spontaneously form when a metal salt is added to a solution of another precipitating anion (often, sodium silicate). Chemical garden precipitates form as the metal salt ‘seed’ crystals dissolve and the metal cation precipitates with the  $\text{OH}^-$  and silicate from the external solution, forming a porous and semi-permeable mineral membrane whose physical characteristics vary depending on the reaction conditions. Osmotic pressure inside the structure eventually causes membrane rupture and re-precipitation, and so on until the metal salt is completely dissolved and a precipitate structure is formed. Chemical gardens have been studied in a variety of reaction systems, and researchers have attempted to control the morphology of the structures formed *via* methods such as bubble-guided tube growth<sup>21</sup> or growing confined precipitates in 2D reactors.<sup>22</sup>

Already in the 1920s there were attempts to control chemical-garden growth directions with light.<sup>23</sup> More recently, this idea has reappeared with microtubes of polyoxometalates, which are analogous to classical chemical gardens, and have been light-controlled with a laser.<sup>23–25</sup> Here, we report on a 2D chemical garden system that self-assembles a microfluidic labyrinth (Fig. 1).

## Materials and methods

The experiments were carried out using a Hele-Shaw cell consisting of two squared acrylic plates and a strip of rubber

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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3cp02929h>



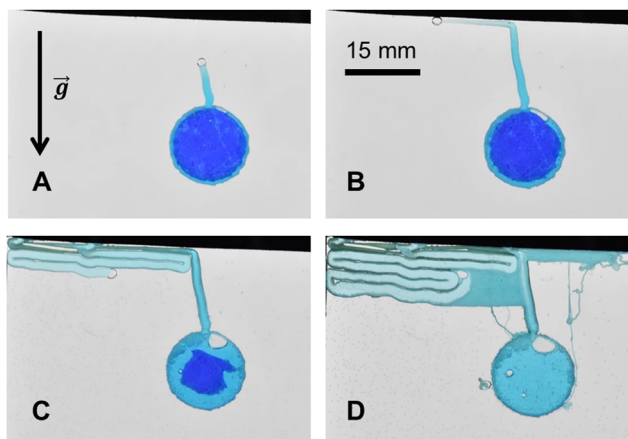


Fig. 1 Example of the four stages of the labyrinth growth in the case of a  $\text{CuCl}_2$  seed in a 4 M sodium silicate solution: (A) bubble-guided growth, (B) roof growth, (C) Labyrinth growth, and (d) new precipitates growth.

to isolate the system setting the plates 1 mm apart. A metal salt pellet (either  $\text{CuCl}_2$  or  $\text{Cu}(\text{NO}_3)_2$ ), made in a pill press with 300 mg of salt and diameter of 13 mm, was placed at the centre of the cell. Then, the cell was filled with  $\sim 5$  mL of sodium silicate solution of a concentration varying from 1 M to 5 M. The evolution of the chemical garden was recorded using a Nikon D3400 DSLR camera taking pictures every 60 s. The time the system takes to complete the chemical reaction varied between 20 and 50 hours. The memory card of the camera limited the number of pictures per experiment, causing some experiments not to be completely filmed (as seen in the supplementary information). In most experiments the cell was held vertically upright throughout the experiment; however, for a select successful experiment (5 M sodium silicate +  $\text{Cu}(\text{NO}_3)_2$  pill seed) we started the experiment with the Hele-Shaw cell rotated  $45^\circ$  to the left, and then  $\sim 20$  minutes after the start of the reaction the cell was rotated  $45^\circ$  to the right, in order to test whether the direction of tube growth can be controlled.

To measure the physical characteristics of the precipitates, the experimental photos were scaled to real size with ImageJ code. The width of the channels was determined by taking five measurements of each horizontal channel at different points and calculating the mean. Velocities were calculated by dividing the difference in channel distance at two points by the 60s time interval between photos. To measure the growth of precipitate, we binarized the images so that the pixels were either black (labyrinth/precipitate) or white (background). Then, we processed the images with a Python program to calculate the percentage of black pixels in each image, showing the increase in area of the labyrinth over time.

We also performed chemical microanalysis on the materials obtained for each Cu salt experiment with the silicate concentration that produced the most stable labyrinths: 3 M silicate for  $\text{CuCl}_2$  and 5 M silicate for  $\text{Cu}(\text{NO}_3)_2$ . The precipitates were extracted from the cells by taking out the silicate, opening each cell and drying the material inside. The samples were not rinsed with water because they were so delicate that the water

destroyed the tubes. X-ray diffraction (XRD) analysis was performed at  $25^\circ\text{C}$ , using a BRUKER D8 Discover diffractometer with the PILATUS3R 100K-A detector and the  $\text{Cu K}\alpha$  radiation ( $\lambda = 1.5406 \text{ \AA}$ ) and steps of  $0.02^\circ$  ( $2\theta$  units) and  $40 \text{ s step}^{-1}$  exposition time. Raman spectroscopy analysis was carried out with a JASCO NRS-5100 microspectrometer using a 2.6 mW 785 nm laser. Additionally, a laser attenuator was added to allow for longer collection times. Scanning electron microscopy (SEM) analysis was performed with a FEI microscope, QuantaScan650F, coupled with an energy dispersive X-ray spectroscopy (EDX) detector for elemental analysis. EDX was used to create complete chemical maps as well as point-by-point chemical analysis.

## Results and discussion

### Precipitation formation stages

We observed formation of chemical garden precipitates from the dissolution of the metal salt pellet into the sodium silicate solution, growing due to a combination of buoyancy and internal osmotic fluid pressure as observed in previous studies.<sup>20</sup> However in this case the confinement of the precipitation reaction in the 2D Hele-Shaw cell led to several distinct stages of precipitate formation, though not all stages were observed in all experiments. In Stage 1 (bubble-guided growth), an air bubble on the pellet's surface (from air trapped inside the pressed pellet) guided the growth of a precipitate tube upwards due to buoyancy (Fig. 1A).<sup>20</sup> In Stage 2 (roof growth), once the tube reached the top of the cell, the bubble vanished and the fluid pressure from the dissolving metal salt pill became the dominant force for forming new precipitate, creating a "roof" of precipitate mass that covered the top surface (Fig. 1B). In Stage 3 (labyrinth growth), new air bubbles became trapped on the edges of the cell as the roof formed, which produced new bubble-guided tubes (Fig. 1C). The bubble cannot move upwards and reach the liquid-air interface, so it has to move horizontally or downwards pushed by the fluid flow. The bubble conserves its state of movement until it hits an obstacle, *i.e.* a tube or the edge of the cell. This causes the tube to form an alternating maze-like pathway, with the bubble ensuring that each level of the tube was closely layered with the level above. In Stage 4 (more tubes and dendritic pattern precipitates), for some sodium silicate concentrations, new precipitate structures grew after the labyrinth formation was finished (Fig. 1D). The labyrinths may remind some of the chemical garden spirals observed by Hussein *et al.*<sup>26</sup> and the layering tubes of Bentley *et al.*<sup>27</sup>

### Effect of metal salt on precipitate formation

Both of the salts tested formed chemical gardens at all silicate concentrations, but  $\text{CuCl}_2$  and  $\text{Cu}(\text{NO}_3)_2$  only formed labyrinth structures at specific silicate concentrations (Fig. 1 and 2). Thus, the formation of labyrinth structures is a property affected by both silicate concentration and the anion of the Cu salt. As these are complex systems, we theorize that various



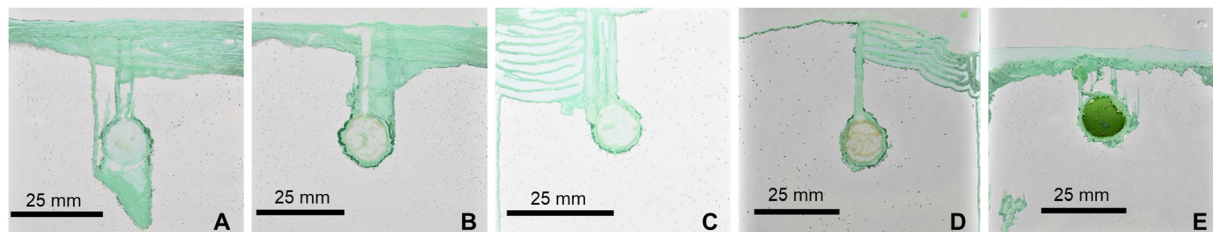


Fig. 2 Labyrinth experiment  $\sim 20$  hours after the start of the reaction, using a copper chloride ( $\text{CuCl}_2$ ) pill inside a Hele-Shaw cell and submerged in (A) 1 M, (B) 2 M, (C) 3 M, (D) 4 M, and (E) 5 M sodium silicate solution.

characteristics should be taken into account for the formation of labyrinth structures, including for example, gravity, density, buoyancy, or solubility of the metal salt; it is unclear what properties determine whether the specific stages of structures will form from a particular reaction chemistry in this system.

### Copper chloride [ $\text{CuCl}_2$ ] labyrinths

In experiments with  $\text{CuCl}_2$  we observed that different structures formed for each sodium silicate concentration tested (Fig. 2a–e). 1 M (Fig. 2a) and 2 M (Fig. 2b) silicate experiments exhibited vertical precipitate tubes growing upwards and precipitation of a “roof” (Stages 1 and 2). The width of the labyrinth tube in Stage 1 was determined by the bubble size; however, during Stage 3 (reached at 3 M silicate) the growth followed a laminar pattern with layer separation of  $\sim 0.2$  mm (Fig. 2c). Well-defined labyrinths were only observed in 3 M (Fig. 2c) and 4 M (Fig. 2d) silicate experiments. The thickness of the tubes forming the labyrinth varied from 0.8 mm to 1.5 mm for 3 M and from 1 mm to 2.5 mm for 4 M silicate. At 5 M silicate (Fig. 2e), no labyrinths formed, but thinner bubble-guided tubes were observed growing downwards or upwards.

### Copper nitrate [ $\text{Cu}(\text{NO}_3)_2$ ] labyrinths

Similar to the  $\text{CuCl}_2$  experiments, different concentrations of silicate led to formation of different structures with  $\text{Cu}(\text{NO}_3)_2$  (Fig. 3). At 1 M silicate (Fig. 3a), all bubbles separated from the pill before forming the initial thick tube (contrary to what we have seen in  $\text{CuCl}_2$  at 1M), and so the precipitate could not complete Stage 2 (roof growth). With 2 M (Fig. 3b) and 3 M silicate (Fig. 3c) experiments showed similar structures to those obtained with  $\text{CuCl}_2$ : some bubbles were created after filling the cell with the silicate and bubble-guided growth of the initial tube occurred. Afterwards, however, it did not create a periodic labyrinth structure; instead, most of the material formed a laminar structure with small layer separation ( $\sim 0.2$  mm). On

the other hand, with 4 M (Fig. 3d) and 5 M (Fig. 3e) silicate we observed formation of much more defined and periodic labyrinths.

Because the width of the tube is controlled by the size of the initial bubble, this width could be reproduced by controlling the bubble. Knowing these variables, we could also decide the approximate length of the tubes, which also depends on the mass of the pill and its distance of the cell roof and sides. On the other hand, the windings and of the labyrinth and the new precipitates are not completely reproducible at the moment, except if we control the growth.

### Quantifying the effect of silicate concentration on labyrinths growth with Cu salts

There were similarities between the evolution of the labyrinths depending on silicate concentration for both Cu salts. Both had a range of silicate concentrations where labyrinths were stable, with wider, periodic channels (3 M and 4 M silicate for  $\text{CuCl}_2$ , and 4 M and 5 M silicate for  $\text{Cu}(\text{NO}_3)_2$ ). There was also a silicate concentration range for each in which the precipitate created a layered laminar pattern instead of the usual wider canals (1 M and 2 M silicate for  $\text{CuCl}_2$ , and 2 M and 3 M silicate for  $\text{Cu}(\text{NO}_3)_2$ ). In the case of  $\text{CuCl}_2$  at 5 M silicate, the precipitate reverted back to this laminar pattern (Fig. 2e).

For  $\text{CuCl}_2$ , the labyrinth channel widths increased with time (due to the pressure of the internal fluid) (Fig. 4A); the lower silicate concentration (3 M) created wider channels. The labyrinths obtained with 3 M and 4 M silicate and  $\text{CuCl}_2$  grew slowly, with velocities on the order of  $10^{-3}$  mm  $\text{s}^{-1}$ , with velocity decreasing exponentially over time (Fig. 4B). This may be the result of the mass of the pill decreasing as it dissolved, thus decreasing the internal fluid pressure that pushes the bubble up. The growth rate of the rest of the  $\text{CuCl}_2$  experiments that show only laminar growth (1 M, 2 M, and 5 M silicate) (Fig. 4C) all show a similar tendency, where the rate of growth

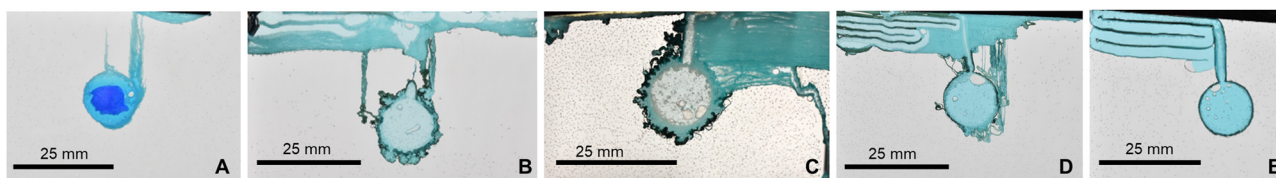


Fig. 3 Labyrinth experiment  $\sim 20$  hours after the start of the reaction, using a copper nitrate ( $\text{Cu}(\text{NO}_3)_2$ ) pill inside a Hele-Shaw cell and submerged in (A) 1 M, (B) 2 M, (C) 3 M, (D) 4 M, and (E) 5 M sodium silicate solution.



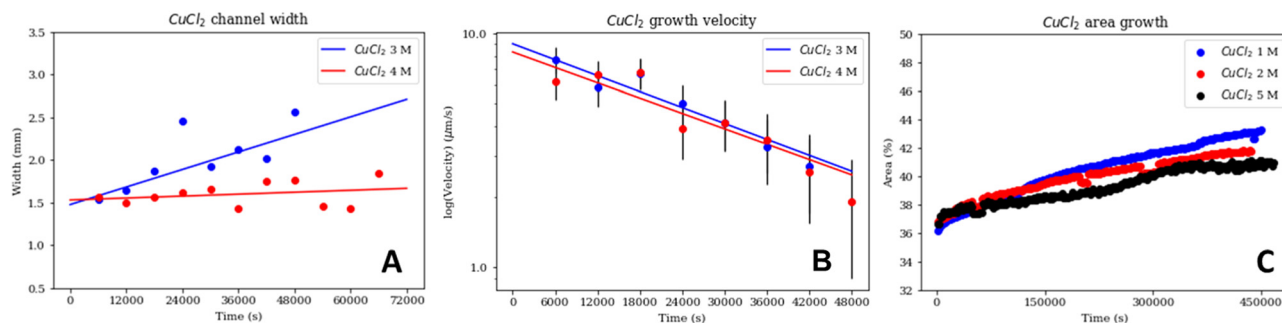


Fig. 4 Graphical representation of the evolution of the (A) 2D tube or channel width and (B) tube growth velocity of the  $\text{CuCl}_2$  + 3 M (blue) and 4 M (red) sodium silicate labyrinths during the first  $\sim 20$  hours with an error of  $1 \mu\text{m}$ ; as well as of the (C) rate of growth of the labyrinth through the percentage of the cell area covered by the  $\text{CuCl}_2$  + 1 M (blue), 2 M (red), and 5 M (black) sodium silicate chemical gardens. The coefficients of determination for the fit lines in (A) are  $R^2 = 0.567$  (blue) and  $R^2 = 0.693$  (red), and in (B)  $R^2 = 0.923$  (blue) and  $R^2 = 0.869$  (red). The error in the width in 4a is approximately 3%, not great enough to be represented. Note that the lines in 4c do not start at 0% because they represent the area in the cell with material, which could include the initial pill and the cell borders.

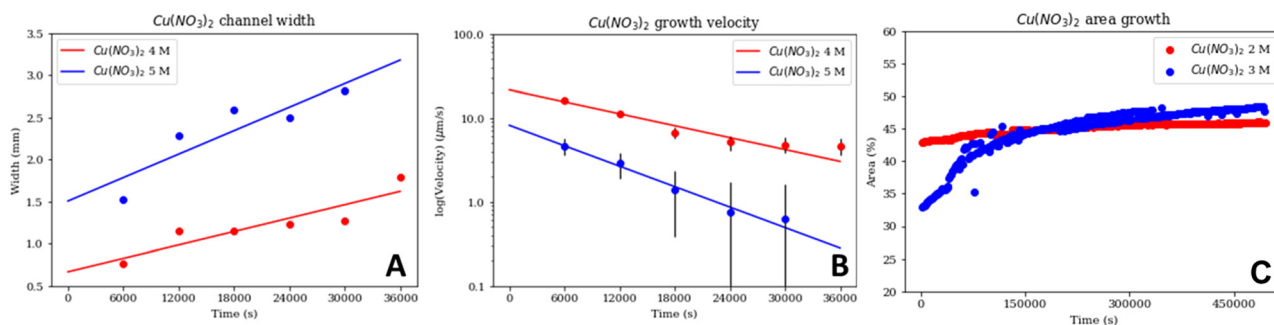


Fig. 5 Graphical representation of the evolution of the (A) tube width and (B) tube growth velocity of the  $\text{Cu}(\text{NO}_3)_2$  + 4 M (red) and 5 M (blue) sodium silicate labyrinths during the first  $\sim 10$  hours with an error of  $1 \mu\text{m}$ ; as well as of the (C) rate of growth of the labyrinth through the percentage of the cell area covered by the  $\text{Cu}(\text{NO}_3)_2$  + 2 M (red), and 3 M (blue) sodium silicate chemical gardens. The coefficients of determination for the fit lines in (a) are  $R^2 = 0.819$  (red) and  $R^2 = 0.795$  (blue), and in (b)  $R^2 = 0.812$  (red) and  $R^2 = 0.891$  (blue). The error in the width in 5a is approximately 3%, not great enough to be represented. Note that the lines in 5c do not start at 0% because they represent the area in the cell with material, which could include the initial pill and the cell borders.

diminishes over time until the labyrinth becomes static. This agrees with what we expected: the pressure from the material in the pill decreases as said material transforms into the labyrinth or laminar precipitate, finally stopping the growth when the pill is completely dissolved. However,  $\text{CuCl}_2$  experiments at different silicate concentrations vary in their area growth over time, with less concentration causing a more rapid growth than higher ones. Also, note that these curves do not start at the zero origin because there were already some black pixels before filming; that is the pill as well as the borders of the Hele-Shaw cell.

For  $\text{Cu}(\text{NO}_3)_2$ , the behaviour of the tube widths for 4 M and 5 M silicate also increased with time (Fig. 5A); however, in this case the 5 M experiment had wider channels than the 4 M experiment, contrary to Fig. 4B. The growth velocities of the labyrinths in the 4 M and 5 M silicate  $\text{Cu}(\text{NO}_3)_2$  experiments followed a similar exponentially decreasing tendency, and again in this case the higher concentration experiment (5 M silicate) was slower (Fig. 5B). However, the initial velocities for  $\text{Cu}(\text{NO}_3)_2$  were much higher than for  $\text{CuCl}_2$ . Finally, for the

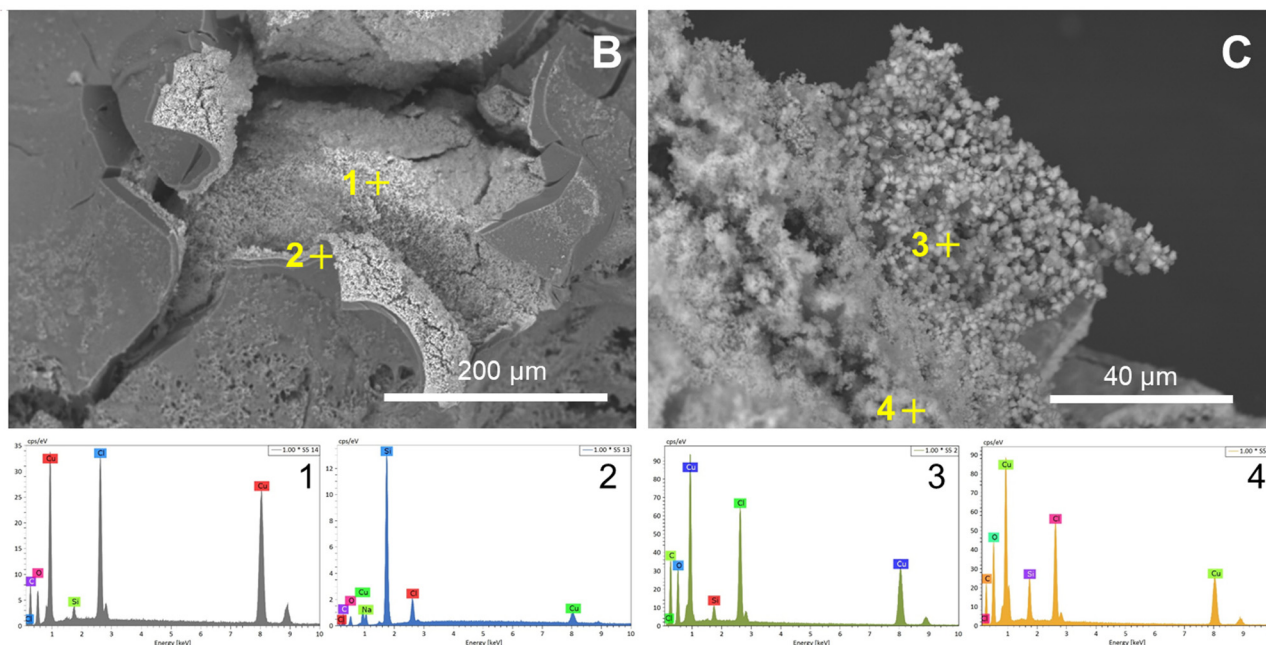
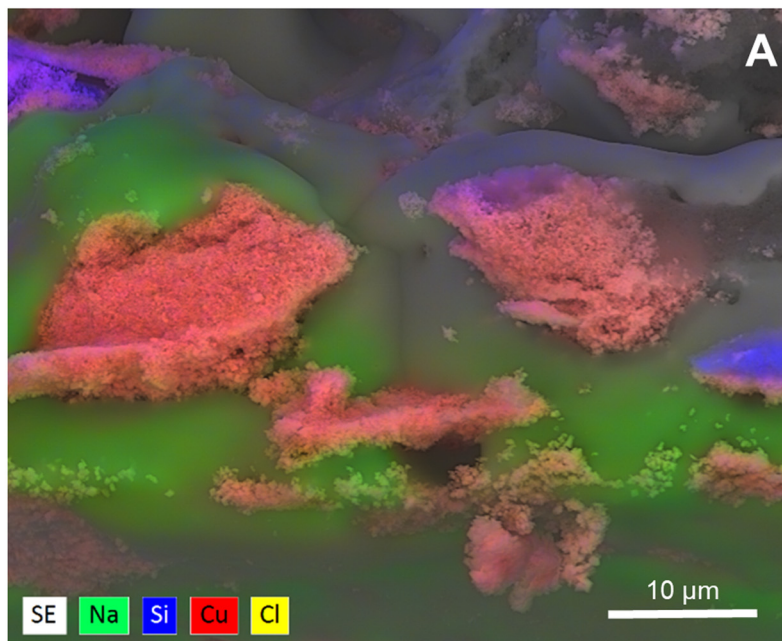
silicate concentrations that created laminar structures with  $\text{Cu}(\text{NO}_3)_2$  (2 M, and 3 M); we can see that the garden growth rate decreases and tends to become constant (Fig. 5C). There was a bigger difference between the growth curves for the  $\text{Cu}(\text{NO}_3)_2$  experiments than for the  $\text{CuCl}_2$ ; and in the case of  $\text{Cu}(\text{NO}_3)_2$  the higher silicate concentration (3 M) led to faster growth.

### Chemical microanalysis

We conducted further chemical microanalysis on two samples of precipitate from Cu salt experiments that produced the most well-defined labyrinths (3 M silicate +  $\text{CuCl}_2$ , and 5 M silicate +  $\text{Cu}(\text{NO}_3)_2$ ).

**SEM/EDX.** In both samples the Si and Na from the external solution were heterogeneously distributed across the precipitate surface (Fig. 6A and 7A). The elements from the Cu salt anions (Cl and N) were also observed on the surface. In labyrinth precipitates from  $\text{CuCl}_2$  and silicate, Cu appeared together with Cl in separate clusters of crystals (Fig. 6A); however, labyrinth precipitates from  $\text{Cu}(\text{NO}_3)_2$  and silicate showed that Cu





**Fig. 6** (A) EDX elemental analysis map of a  $\text{CuCl}_2$  3 M silicate labyrinth showing Na, Si, Cu, and Cl. (B) SEM of the peeled-off Si-layer of a  $\text{CuCl}_2$  3 M silicate labyrinth with EDX analysis of interior and exterior. (C) SEM of the crystals formed inside of a  $\text{CuCl}_2$  3 M silicate labyrinth with EDX analysis of the crystals.

was mixed with N and Si through the surface of the precipitate (Fig. 7A). Fig. 6B shows a peeled-off Si-layer in the  $\text{CuCl}_2$  silicate labyrinth; the external wall contained a large amount of Si and Na, while the inside was made from Cu and Cl crystals. Fig. 7B also shows a rupture in the  $\text{Cu}(\text{NO}_3)_2$  silicate labyrinth, where a similar general structure can be seen. Although the crystals' geometry and organization are not the same, we can still see a Si and Na layer around the crystal formation. Analysis of those crystals shows a high percentage of Cu, although in this case O and some Si also

appear. Fig. 6C enlarges a section of the  $\text{CuCl}_2$  silicate labyrinth where very tiny Cu and Cl crystals appear; this contrasts with the much larger crystals that form flower-like structures in  $\text{Cu}(\text{NO}_3)_2$  silicate labyrinths as seen in Fig. 7C.

**XRD.** In the case of the  $\text{CuCl}_2$  labyrinth at 3 M silicate, the XRD patterns indicated both halite [ $\text{NaCl}$ ] and paratacamite [ $\text{Cu}_4(\text{OH})_6\text{Cl}_2$ ], two halide minerals (Fig. 8a). Halite shows the combination from the chlorine ions and the sodium from the silicate solution, indicating that the Na cations cross the



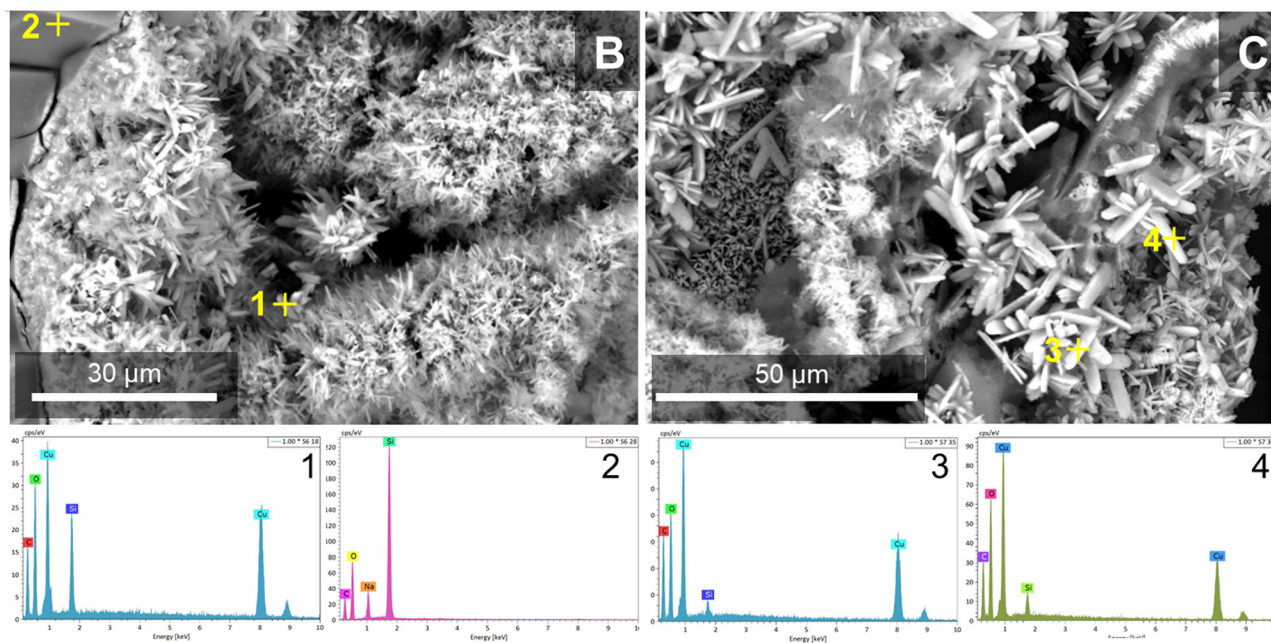
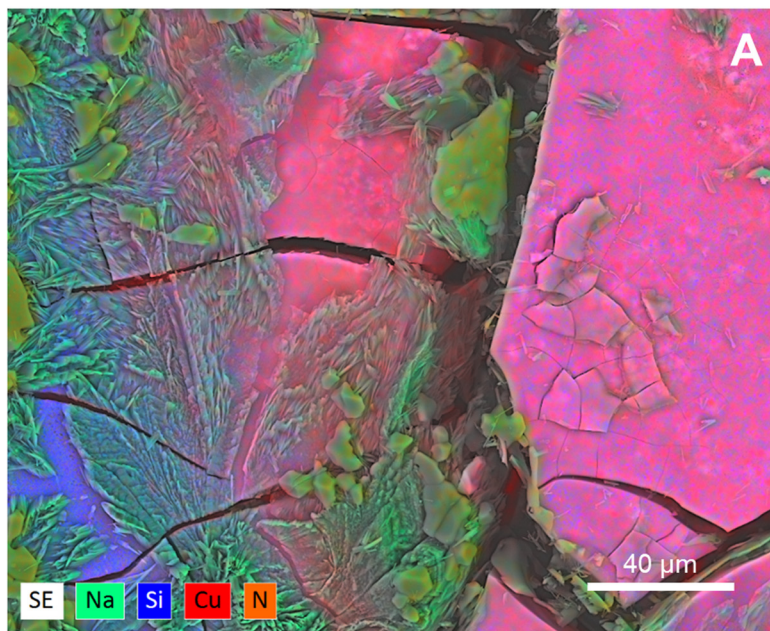


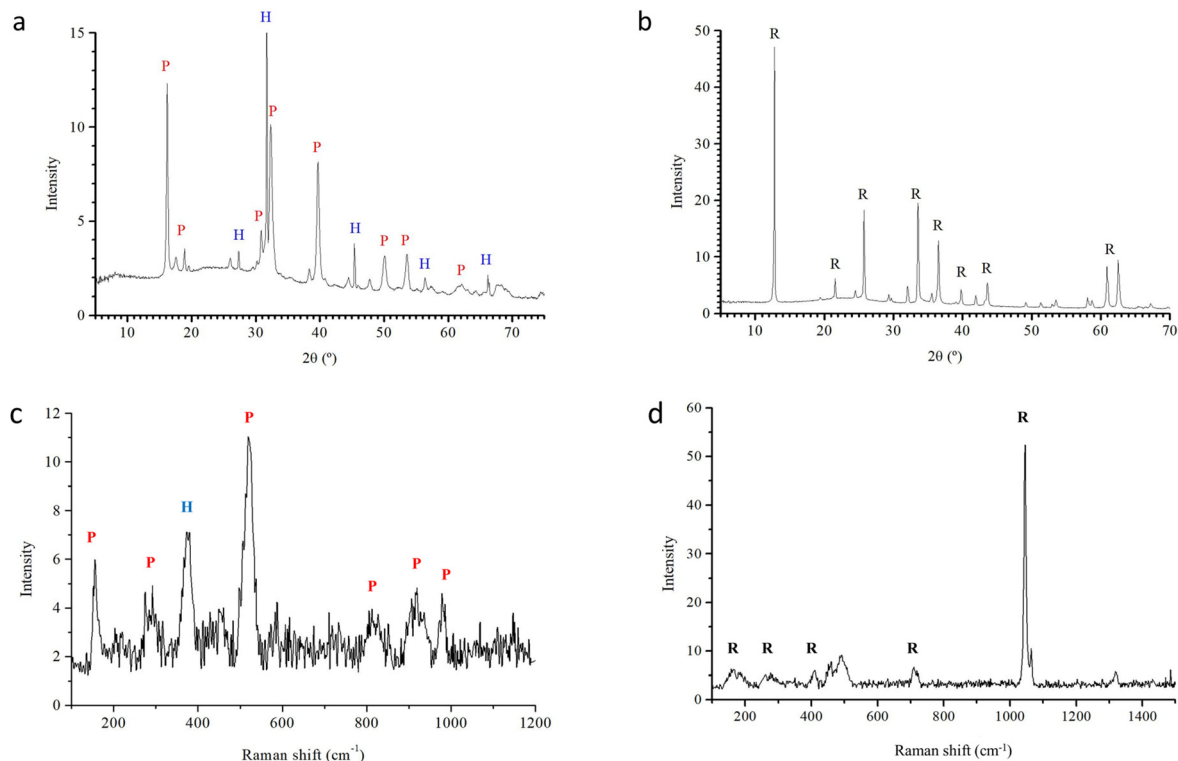
Fig. 7 (A) EDX elemental analysis map of a  $\text{Cu}(\text{NO}_3)_2$  5 M silicate sample showing Na, Si, Cu, and N on the precipitate surface. (B) SEM of the broken surface of a  $\text{Cu}(\text{NO}_3)_2$  5 M silicate labyrinth with EDX analysis of interior and exterior. (C) SEM and EDX of the crystals formed inside of a  $\text{Cu}(\text{NO}_3)_2$  5 M silicate labyrinth.

osmotic membranes to the inner solution and precipitate on the labyrinth walls. Paratacamite is a hydroxylated form of  $\text{CuCl}_2$ , indicating the entrance of the  $\text{OH}^-$  anions from the strongly basic external solution to the acidic internal solution through the osmotic membrane. This result correlates with halite and paratacamite XRD patterns found in the literature.<sup>15,16</sup> When studying the other sample [ $\text{Cu}(\text{NO}_3)_2$  labyrinth at 5 M silicate], only one compound was found: rouaite [ $\text{Cu}_2(\text{NO}_3)(\text{OH})_3$ ] (Fig. 8b) which is a hydroxylated form of  $\text{Cu}(\text{NO}_3)_2$ .<sup>17</sup> This result was verified when looking at the

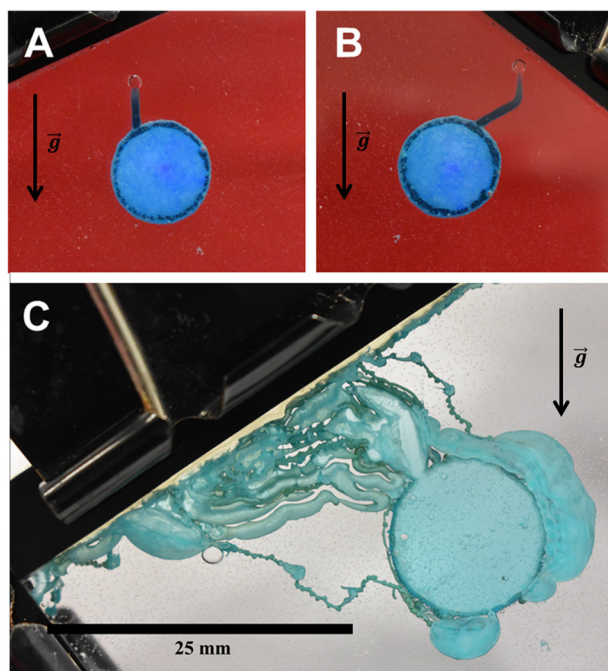
sample's crystals with SEM (Fig. 7C) and comparing with crystals of  $\text{Cu}_2(\text{NO}_3)(\text{OH})_3$  reported previously,<sup>18</sup> confirming they have the same slice-like hexagon structures. A large background of amorphous material assigned to silicate materials was also observed in the baseline of the diffractogram.

**Raman spectroscopy.** Raman spectroscopy showed that the labyrinth samples produce spectra similar to paratacamite and rouaite, including the hydroxyl deformation regions;<sup>15,19</sup> this indicates that both samples contain hydroxylated compounds (Fig. 9), agreeing with the XRD results.





**Fig. 8** (a) and (b) XRD patterns of (a)  $\text{CuCl}_2$  + 3 M silicate labyrinth sample where P are paratamite peaks and H are halite peaks, and (b)  $\text{Cu}(\text{NO}_3)_2$  + 5 M silicate labyrinth sample where R are rouaite peaks. (c) and (d) Micro-Raman analysis spectra of (c)  $\text{CuCl}_2$  + 3 M silicate labyrinth sample, and (d)  $\text{Cu}(\text{NO}_3)_2$  + 5 M silicate labyrinth sample. P are paratamite peaks, H are halite peaks, and R are rouaite peaks.



**Fig. 9** Labyrinth experiment using a  $\text{Cu}(\text{NO}_3)_2$  pill inside a Hele-Shaw cell and submerged in a 5 M sodium silicate solution (A) ~20 minutes after the start of the reaction rotated  $45^\circ$  to the left; (B) ~40 minutes after the start of the reaction rotated  $45^\circ$  to the right; and (C) its finished structure after ~20 hours. The bubble guides the tube to the top of the cell even with the experiment rotated.

### Controlling the labyrinth

In an experiment with 5 M silicate +  $\text{Cu}(\text{NO}_3)_2$ , when the cell was rotated  $45^\circ$  to the left at the start of the experiment, the initial bubble-guided tube grew vertically upward (perpendicular to the ground) (Fig. 9a) rather than parallel to the cell walls as in previous experiments. Moreover, after changing the rotation to  $45^\circ$  to the right, the tube changed direction to similarly grow perpendicular to the ground (Fig. 9b). This is expected since bubble-guided tube growth is buoyancy directed and therefore changing the orientation of the cell relative to gravity will allow the ability to direct the direction of tube growth as long as a bubble is present. As we can see in Fig. 9c, after some time the final labyrinth also stayed parallel to the ground instead of following the rotation of the cell. Although this is a simple test, it was enough to affirm that the labyrinth's geometry can be changed at will through external modifications. For example, changing the direction of the cell so that the labyrinth is always growing downwards may increase the velocity of its growth.

### Conclusions

Our experiments studying of the formation of chemical gardens in 2D reactors uncovered a curious phenomenon: the formation of labyrinthine structures. The conditions for this formation are a balance of concentration of the silicate solution and the type of Cu salt of the seed used. One necessary



condition needed for this phenomenon is the existence of a gas bubble. The balance between the buoyancy forces of this gas bubble and the osmotic pressure of the internal solution, along with the mechanical properties of the precipitate that forms the labyrinth walls, allows the formation of these structures. One property that varies between the Cu salts is their solubility. Lower solubilities can decrease the reactivity of the pellet or increase the mechanical resistance to the bubble movement. The Cu salts used have high solubility allowing the formation of these labyrinths. However, solubility cannot be too high, as then a higher concentration of silicate would be needed to reach the balance necessary for the labyrinth formation (as seen with the highly soluble  $\text{Cu}(\text{NO}_3)_2$ ).

Even though the labyrinths in this study were based on only two Cu metal salts ( $\text{CuCl}_2$  and  $\text{Cu}(\text{NO}_3)_2$ ), chemical gardens can be formed with a wide variety of metal salt seed crystals,<sup>19</sup> and so it would be worthwhile to explore which other chemistries could also produce labyrinth structures. As commented previously, the formation of labyrinths (as opposed to laminar precipitates) in other chemistries would be dependent upon the details of the system including the metal salt cation and anion, and silicate concentration. The growth stages we present for these 2D chemical garden systems are generally useful since for any metal salt plus silicate chemistry that is able to reach stage 3 (labyrinth growth from bubble-guided tube formation), it would be theoretically possible to control the geometry of the labyrinth by rotating the experiment as shown in Fig. 9. The walls of the two labyrinth structures studied here exhibited crystals related to the metal salt coating the inner walls and silicate on the exterior, as has been observed in previous chemical garden studies. Chemical garden systems such as these serve as an interesting example of inorganic self-organization, and could be used to create microfluidic labyrinths which could facilitate reactivity on the interior crystalline walls, the composition and orientation of which could be externally controlled.

## Author contributions

S.T.M., T.H.M., and P.K. conducted the lab experiments. J.H.E.C. and C.I.S.D. developed the 2D reactor technique. S.T.M. and C.I.S.D. conducted microscopy analysis. S.T.M., L.M.B., C.I.S.D., and J.H.E.C. wrote the paper.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

L. M. B. was supported by JPL Topical Research & Technology Development and a NASA PECASE award; L. M. B.'s research was carried out at the Jet Propulsion Laboratory, California Institute of Technology, under a contract with the National Aeronautics and Space Administration (80NM0018D0004). S. T. and T. H. thank the CSIC and Spanish Andalusian 'Garantía Juvenil'

project AND21\_IACT\_M2\_058. The authors would like to acknowledge the contributions of the European COST Action CA17120 supported by the EU Framework Programme Horizon 2020.

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