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**Strong fluorescence is a general feature of dipyrrolonaphthalimidinediones bearing two nitrophenyl substituents. Methyl groups simultaneously being weakly electron-donating and inducing steric hindrance appear to be a key structural parameter that allows for significant emission enhancement, whereas Et<sub>2</sub>N groups cause fluorescence quenching. The magnitude of two-photon absorption increases if 4-nitrophenyl substituents are present while the contribution of Et<sub>2</sub>N groups is detrimental.**

Recent developments in molecular photonics<sup>1</sup> and biochemistry<sup>2</sup> have increased the demand for novel organic materials with improved emission characteristics and enhanced resistance towards ambient conditions. Bearing in mind the stability of  $\pi$ -conjugated organic materials, one of the most common strategies is to decrease electron density, usually enabled by introducing an electron-deficient moiety into an aromatic backbone.<sup>3</sup> The nitro group ( $-\text{NO}_2$ ) is one of the strongest electron-withdrawing moieties, and by meeting this criterion, potentially sets the stage for novel materials with intriguing optical characteristics.

Nitroaromatics have, however, long been considered as non-fluorescent, and highly emissive examples of  $\text{NO}_2$ -containing chromophores have been sparsely reported in the literature (Fig. 1A).<sup>4</sup> Yet, they are promising candidates for constructing probes<sup>5</sup> tracking important biologically relevant species and as efficient two-photon absorbers.<sup>6</sup>

In recent years, some strategies have been developed to improve the luminescence performance of nitroaromatics.<sup>4b</sup>

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<sup>1</sup>H NMR and <sup>13</sup>C NMR spectra, and steady-state and time-resolved spectra. See  
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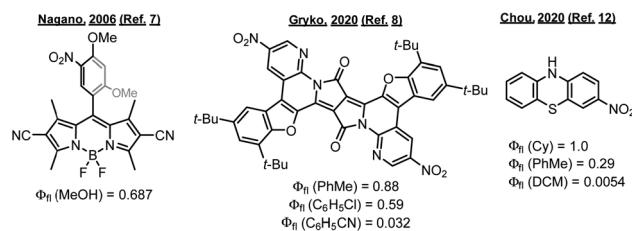
## Realization of nitroaromatic chromophores with intense two-photon brightness<sup>†</sup>

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From the mechanistic viewpoint, fluorescence can be successfully induced by blocking the  $\text{d-PeT}^7$  process or when the nitro group does not participate in the  $S_0 \rightarrow S_1$  transition.<sup>8</sup> Applying a weaker electron-donating moiety for the construction of nitro-containing conjugates may also give rise to nitroaromatics with enhanced emission. Dyes containing nitro-functionalized thiophene,<sup>9</sup> fluorene,<sup>10</sup> tetrabenzofluorenes,<sup>11</sup> or phenothiazine<sup>12</sup> featuring enhanced emission have been described, but these approaches contain some drawbacks, such as a tendency for deexcitation *via* ISC (heavy atom effect of sulphur) or lack of significant spectral red-shifts of both the absorption and emission bands.

Due to its intriguing properties, the electron-deficient dipyrrolonaphthalimidinedione (DPND) core<sup>13</sup> has recently been successfully employed in the construction of dyes with red/NIR emission,<sup>14</sup> OFETs,<sup>15</sup> stable singlet fission (SF) materials<sup>16</sup> and singlet oxygen ( $^1\text{O}_2$ ) photosensitisers.<sup>17</sup> Studying quadrupolar, centrosymmetric DPNDs possessing two nitroaryl substituents at positions 3 and 9, we discovered that creating nitroaromatics with

### A. Exemplary, highly fluorescent nitroaromatics:



### B. Nitroaryl-decorated DPNDs (this work):

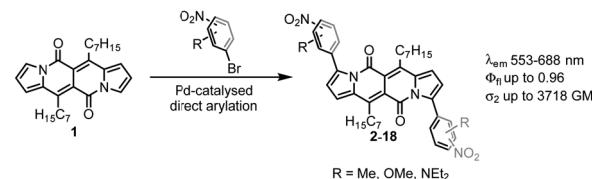


Fig. 1 (A) Exemplary highly emissive nitroaromatics. (B) The idea of the present work.



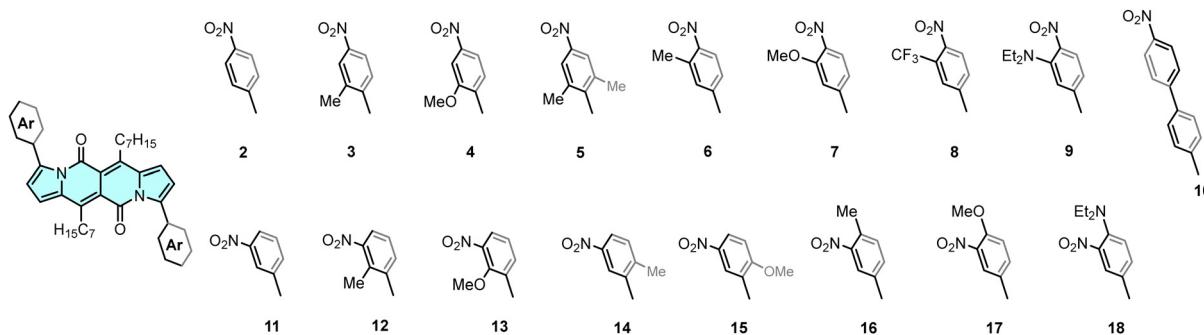


Fig. 2 Structure of nitroaromatics studied in this work.

acceptor–acceptor'-acceptor (A–A'-A) architecture constitutes a new viable strategy of preventing non-radiative deactivation.<sup>18</sup>

There is, however, still a huge gap in the knowledge of how positional isomerism of the nitro group relative to the chromophore core influences the emission characteristics of nitroaromatics.

To overcome this intrinsic limitation we now report how distinct structural modifications (different substitution patterns of nitroaryl moieties) alter the emissive properties and two-photon absorption characteristics<sup>19</sup> of these quadrupolar, A–A'-A dyes.

DPNDs possessing A–A'-A architecture, considered in this work, were assembled by employing a Pd-catalysed direct arylation<sup>14b</sup> between the DPND **1** and the corresponding aryl bromides (Fig. 1, see ESI† for synthetic details). As a result, a series of new quadrupolar nitroaromatics **4–10** and **13–18** were synthesized and fully described (Fig. 2). The same method was used for the synthesis of the previously described dyes **2**, **3**, **11** and **12**, which are included here for the complete assessment of substitution-dependent optical behaviour. The complete series of dyes differ in the location of the nitro groups (*para* or *meta*), the structure of the linker (phenyl or biphenyl), and the effect of

further substitution with electron-donating or -withdrawing substituents.

The effect of structural variations on the linear optical properties of all DPND derivatives in solution was subsequently examined by employing 1,2-dichlorobenzene (DCB), CH<sub>2</sub>Cl<sub>2</sub> (DCM), or MeCN (ACN) as solvents. The results are summarized in Table 1 and Fig. S2 (ESI†), and the spectral data are shown in Fig. S1–S26 (ESI†). As expected, nitroaromatics built around the DPND core display red-shifted absorption and emission bands compared with **1**, resulting from  $\pi$ -expansion of the chromophore and the presence of groups possessing distinct electronic character.

Upon inspection of Fig. 2, two series of dyes can be recognized: **2–10** bearing nitro groups at the *para* positions relative to the DPND core, and **11–18** where the  $-\text{NO}_2$  occupies *meta* positions. Regardless of the location, some trends are common to both groups. First of all, the magnitude of the response strongly depends on the presence of a substituent (Me, OMe) at the *ortho* position relative to the DPND core. Functionalization at this position increases the dihedral angle between the chromophore and the nitroaryl moiety, therefore decreasing

Table 1 Comparison of spectral data for the chosen compounds in three different solvents: 1,2-dichlorobenzene (DCB), CH<sub>2</sub>Cl<sub>2</sub> (DCM), and MeCN (ACN)

Compd	Solvent <sup>a</sup>	$\lambda_{\text{abs}}^b$ /nm	$\lambda_{\text{fl}}^b$ /nm	Stokes' shift <sup>c</sup> /cm <sup>-1</sup>	$\Phi_{\text{fl}}$	$\tau^d$ /ns	$k_{\text{r}}^e \times 10^{-8}$ /s <sup>-1</sup>	$k_{\text{nr}}^e \times 10^{-8}$ /s <sup>-1</sup>
<b>1</b> <sup>f</sup>	DCB	508	528	750	0.81	5.1	1.59	0.373
	DCM	504	523	720	0.73	5.7	1.28	0.474
	ACN	498	525	1030	0.65	5.7	1.14	0.614
<b>2</b> <sup>f</sup>	DCB	569	605	1050	0.45	3.6	1.25	1.53
	DCM	562	601	1150	0.41	3.4	1.21	1.74
	ACN	559	599	1190	0.07	—	—	—
<b>5</b>	DCB	524	560	1230	0.96	4.7	2.04	0.085
	DCM	521	560	1340	0.34	1.9	1.79	3.47
	ACN	518	564	1570	0.005	—	—	—
<b>9</b>	DCB	555	602	1390	0.006	—	—	—
	DCM	549	608	1760	0.003	—	—	—
	ACN	542	584	1330	0.003	—	—	—
<b>14</b>	DCB	525	568	1440	0.76	4.9	1.55	0.490
	DCM	522	566	1490	0.58	3.2	1.81	1.31
	ACN	520	569	1660	0.075	—	—	—
<b>18</b>	DCB	575	669	2440	0.12	2.6	0.462	3.38
	DCM	565	670	2770	0.02	—	—	—
	ACN	557	688	3420	0.04	2.8	0.143	3.43

<sup>a</sup> DCB = 1,2-dichlorobenzene; DCM = dichloromethane; ACN = acetonitrile. <sup>b</sup> Absorption and fluorescence maxima. <sup>c</sup> Stokes' shifts. <sup>d</sup> Lifetimes of the emissive excited states obtained from time-correlated single photon counting (for  $\tau \gtrsim 1.5$  ns). <sup>e</sup> Radiative and non-radiative decay rate constants:  $k_{\text{r}} = \Phi_{\text{fl}} \tau^{-1}$  and  $k_{\text{nr}} = (1 - \Phi_{\text{fl}}) \tau^{-1}$ . <sup>f</sup> The data were taken from ref. 18.



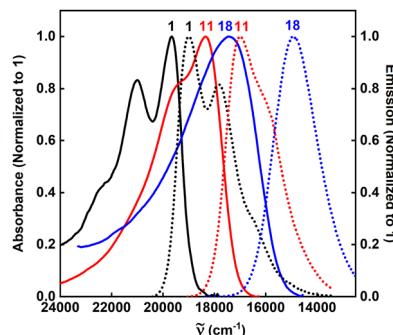


Fig. 3 Absorption and emission spectra of **1** (black solid and dotted lines), **11** (red solid and dotted lines), and **18** (blue solid and dotted lines) dyes in DCB. The **1**, **11**, and **18** emission spectra were excited at 480, 515, and 535 nm, respectively.

conjugation, which results in a blue-shift of both the absorption and emission bands (2 *vs.* 3–4 and 11 *vs.* 12–15). Conversely, a reduced degree of rotation around the C–C single bond generally leads to an enhanced fluorescence response (Table 1 and Table S1 (ESI<sup>†</sup>), Fig. 3).

Dye **5** constitutes a special case among all the investigated nitroaromatics as it bears four *ortho*-located additional methyl groups that diminish the electronic coupling between substituents and the core, and effectively prevent the free rotation of aryl rings. This structural manipulation leads to a record-breaking value of  $\varphi_{\text{fl}}$  among all nitro-DPNs while the maxima of both  $\lambda_{\text{abs}}$  and  $\lambda_{\text{em}}$  are further hypsochromically shifted compared to those measured for **3** and **4** and just 15–30 nm bathochromically shifted *vs.* parent DPN **1**.

The presence of an additional weak electron-donating substituent (Me or OMe) at the *ortho* position relative to the  $\text{NO}_2$  group also enhances emission in all tested solvents (Table 1 and Table S1, ESI<sup>†</sup>). It marginally influences the position of both the absorption and emission bands in the case of **6** and **7** within the *para* series, while for the *meta* one, a small red-shift can be noted, especially for  $\lambda_{\text{em}}$  in **17**. The potent effect of these structural factors, namely the presence of steric hindrance as well as an additional *ortho* substituent relative to the nitro group accumulates within **12**, giving rise to an enormous  $\varphi_{\text{fl}}$  value in relatively polar  $\text{CH}_2\text{Cl}_2$  (0.92).<sup>18</sup>

The incorporation of a substituent with a weak electron-accepting effect ( $-\text{CF}_3$ , dye **8**) or longer biaryl linkage (dye **10**) does not really alter the absorption/emission characteristics as compared with **2** (Table 1 and Table S1 and Fig. S6 and S8, ESI<sup>†</sup>). At the same time, however, the presence of  $\text{CF}_3$  groups leads to a sharp decrease of  $\varphi_{\text{fl}}$  in polar solvents.

Finally, we investigated the influence of an auxochrome with a positive mesomeric effect ( $-\text{NET}_2$ ). It was found that groups featuring a strong mesomeric effect ( $-\text{NO}_2$ ,  $-\text{NET}_2$ ) placed at the *para* positions relative to the DPN core mainly govern both the absorption and emission properties of the studied DPNs. This is particularly true for the diethylamino group in **18**, giving rise to a significantly red-shifted emission (*ca.* 670 nm) but very weak fluorescence response (Fig. 3). The emission quenching is

Table 2 Two-photon absorption data of DPNs **1–18** in DCM

Dye	$\lambda_{\text{2PA}}^{\text{max1}}/\text{nm}$	$\sigma_2^{\text{max1}}/\text{GM}$	$\lambda_{\text{2PA}}^{\text{max2,max3}}/\text{nm}$	$\sigma_2^{\text{max2,max3}}/\text{GM}$
<b>1</b>	830	27	750	44
<b>2</b>	820 (sh)	707	755(sh) < 690	1204 > 3718
<b>3</b>	820	102	755/690	176/332
<b>4</b>	840	270	755(sh)/720(sh)	396/557
<b>5</b>	820	47	765/720	152/157
<b>6</b>	820	275	765(sh)/720(sh)	496/809
<b>7</b>	810	401	770/720	569/959
<b>8</b>	820(sh)	288	755 < 690	607 > 1134
<b>9</b>	860(sh)	0.7	755/720	4.3/5.6
<b>10</b>	840(sh)	319	765(sh)/720	856/1384
<b>11</b>	840	52	770(sh)/730	115/190
<b>12</b>	820	47	760(sh)/720(sh)	89/113
<b>13</b>	820	43	760(sh)/720	95/138
<b>14</b>	820	25	760(sh)/720(sh)	47/60
<b>15</b>	880	31	820(sh)/700	46/162
<b>16</b>	890	84	790/730	272/310
<b>17</b>	890	89	770(sh)/700	279/513
<b>18</b>	940	27	810/755	86/104

plausibly assigned to photo-induced electron transfer (PeT) occurring in the excited state.

Inspired by the large two-photon absorption (2PA) responses of nitroaromatics bearing arylethynyl or arylethynyl  $\pi$ -linkers between the nitro groups and the DPN core,<sup>19</sup> we investigated the 2PA properties of dyes **1–18** using the two-photon excited fluorescence (TPEF) method (see ESI<sup>†</sup>). The results are collected in Table 2 and Table S2, S3 (ESI<sup>†</sup>). No 2PA or very low response is observed in the spectral region corresponding to the lowest one-photon allowed excited state (*i.e.* above 1000 nm), due to symmetry reasons.<sup>20</sup> In contrast, significant 2PA broad bands peaking (or shouldering) in the 800–850 nm range and even more intense 2PA bands closer to 700 nm are observed at higher energy (Fig. 4). The 2PA magnitude in the NIR1 region strongly depends on the structure of the dyes and the nature and position of the substituents (Table 2). The presence of  $\text{NO}_2$  substituents at the *para* position is a key structural factor: analogous dyes from the *meta* series show much smaller 2PA responses, as noted from the comparison of dyes **2** and **1**; **3** and **12**; and **4** and **13** (Fig. S27, ESI<sup>†</sup>). For the derivatives having a *para*  $\text{NO}_2$  group (**2–10**), the presence of an additional substituent generates a marked decrease of the 2PA both in the case of substituents positioned in the *ortho* (Fig. 4) or *meta* (Fig. S28, ESI<sup>†</sup>)

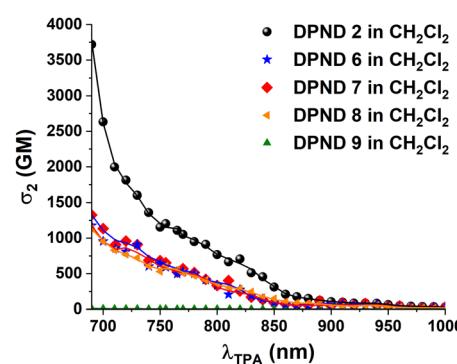


Fig. 4 Two-photon absorption spectrum of DPN **2**, **6–9** in  $\text{CH}_2\text{Cl}_2$ .



position. For the derivatives having a *meta* NO<sub>2</sub> group (**11–18**), the influence of additional substituents is different depending on their position (Fig. S28, ESI<sup>†</sup>). Indeed, the presence of electron-donating MeO in the *para* position leads to a larger 2PA response compared to derivative **11**. We yet note that DPNDs **9** and **18** possessing Et<sub>2</sub>N substituents have markedly weaker two-photon absorption (Fig. 4 and Fig. S3, ESI<sup>†</sup>) most probably due to steric hindrance. Finally, adding an additional 1,4-phenylene moiety in *para* NO<sub>2</sub> derivatives does not lead to an increase of the 2PA (Table 2 and Fig. S29, ESI<sup>†</sup>), indicating that the biphenyl twist may prevent extended delocalization favorable to 2PA enhancement.

To conclude, we found that strong fluorescence is a general feature in quadrupolar, centrosymmetric dipyrrolonaphthyridinediones possessing nitrophenyl substituents. Two-photon allowed transition to higher excited states is responsible for the main two-photon absorption band located around 700 nm. Combining the steric and electronic effect of various substituents enables the fine-tuning of both one-photon and two-photon absorption reaching emission at *ca.* 670 nm and a strong 2PA response (two-photon brightness >1500 GM). Careful selection of the substitution pattern allows for effective control of the optical response, which builds the foundation for designing efficient one- or two-photon absorbers based on other cross-conjugated chromophores.

Conceptualization: B. S.; investigation: B. S., G. C., M. K., P. P.; supervision: D. T. G., M. B.-D., P. P.; visualization: B. S., Y. M. P.; writing – original draft: B. S. D. T. G.; writing – review & editing: D. T. G., M. B. D., B. S.

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## Conflicts of interest

There are no conflicts to declare.

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