

Cite this: *J. Mater. Chem. C*, 2022, 10, 3006

Effect of the nanoparticle size on thermometric properties of a single-band ratiometric luminescent thermometer in $\text{NaYF}_4:\text{Nd}^{3+}$ †

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Due to the fact that temperature is one of the key physical quantities determining the occurrence of physical phenomena, chemical reactions or biological processes, one of the important issues to be solved is the precise determination of temperature. In order to achieve this even under demanding conditions, a lot of attention has recently been focused on methods of remote temperature readout, especially on luminescence thermometry. One of the recent and very interesting methods, the so-called single-band ratiometric luminescence thermometry, is based on the phenomenon of absorption from an excited state and allows for temperature determination based on thermally induced changes in the intensity of a single emission band under two excitation conditions, *i.e.*, by absorption from the ground state and from the excited state. However, the complexity of the latter process in particular requires a number of studies leading to an understanding of the factors affecting the excited level occupancy, which will enable the development of highly sensitive luminescent thermometers based on this method. One of the important aspects is to investigate the influence of the size of the material and the associated occurrence of surface effects, as considered in this work. $\text{NaYF}_4:2\%\text{Nd}^{3+}$ nanocrystals with grain sizes varying from 15 nm to 200 nm were analyzed. A regular trend of widening the useful temperature range with increasing nanocrystallite size was observed. Very high relative sensitivities were obtained, which in the nanocrystals of the largest size analyzed were as high as $19.1\% \text{ K}^{-1}$ at 198 K.

Received 21st December 2021,
Accepted 18th January 2022

DOI: 10.1039/d1tc06069d

rsc.li/materials-c

Introduction

Luminescence thermometry is a remote temperature readout method based on thermally induced changes in the spectroscopic parameters of a given phosphor.^{1–9} Due to its several advantages, it has attracted the attention of many researchers over the past years. One of such features is a remote operating mode, which enables temperature readout even in severe conditions that prevail for instance in high-temperature smelting furnaces or in case of rotating elements of equipment without disturbing them or stopping the machine.^{10–12} So far, for remote temperature sensing, the thermovision method has been mainly used, however, it suffers from some limitations, which include the possibility of imaging the temperature of only the first object encountered in the optical path and the dependence of the outcomes on a predetermined value of the camera input parameter, which is the emissivity.¹³ An erroneous

estimation of the emissivity can result in a temperature misalignment that can reach several hundred degrees. Therefore, luminescence thermometry can be an accurate alternative to thermovision cameras, which, owing to the possibility of imaging not only the surface but also deeper layers, in addition to applications in the field of technology, can also be suitable for applications in biomedicine, *e.g.* in cancer therapy.^{14–16} In luminescence thermometry aimed at biological applications, the most suitable optically active ions are Nd^{3+} , which is due to several factors. First of all, Nd^{3+} ions exhibit the ability to both be excited and to emit in the near-infrared (NIR) region,^{17,18} which covers the region of biological optical windows.^{19–25} This is highly desirable because in such a spectral region, light absorption and scattering by cells and tissues are minimized, so that the depth of penetration of the radiation beams can be deeper and the reliability of the results improved. Furthermore, the 808 nm excitation line, which is one of the most suitable wavelengths for Nd^{3+} ions excitation, is about 20 times less absorbed by the tissues compared to the 980 nm beam frequently used for excitation of up-conversion emission (used to excite Yb^{3+} ions).²⁶ Hence, the risk of side effects due to overheating of tissues during phosphor excitation is minimized.

As shown recently, unique configuration of the energy level diagram of Nd^{3+} ions, including energy difference between

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/d1tc06069d



ground $^4I_{9/2}$ and first excited $^4I_{11/2}$ states of around 2000 cm^{-1} , and the possibility of the occurrence of cross-relaxation process, makes them perfect candidate for the single band ratiometric (SBR) approach of luminescence thermometry.^{9,27–34} Temperature determination by this method is based on the analysis of the intensity ratio of a single emission band being excited by two wavelengths, where one of them is matched to the conventionally used ground state absorption (GSA) and the other to the excited state absorption (ESA). This approach was developed to overcome the limitations of the conventional dual-band ratiometric thermometry, in which the modulation of the emission spectra shape by the dispersive dependence of the extinction of the medium surrounding the phosphor leads to unreliable temperature readout (in the SBR shape of the emission band in modified in the same way regardless of the excitation wavelength used). The intensity of the ESA-excited luminescence depends on many factors. If only the thermal factor is considered, according to Boltzmann's distribution it would be proportional to the energy distance between the ground and the excited multiplet states.³⁵ However, the population of an excited level from which the ESA process occurs is affected by many other factors in addition to the thermal energy supplied to the system. One of the factors influencing the occupancy of the $^4I_{11/2}$ level from which absorption occurs is the interionic interactions through the cross-relaxation (CR) pathway.^{36–38} The energy match between pairs of $^4F_{3/2}$ – $^4I_{15/2}$ and $^4I_{9/2}$ – $^4I_{15/2}$ levels makes $\{^4F_{3/2}; ^4I_{9/2}\} \rightarrow \{^4I_{15/2}; ^4I_{15/2}\}$ cross-relaxation very probable. This process occurs particularly efficiently for high concentrations of dopant ions, however, although it is beneficial for the ESA process, on the other hand, it leads to the quenching of the Nd^{3+} emission intensity. Furthermore, as a result of efficient nonradiative depopulation of excited states for heavily Nd^{3+} ion doped nanomaterials, light to heat conversion can occur, which may further quench the emission intensity and reduce the reliability of measurements.^{39,40} Therefore, it is necessary to optimize a dopant ion concentration that balances those two effects. Moreover, one of the key drivers affecting the occupation of the excited level is the probability of nonradiative transitions, which depends heavily on the energy of the host phonons and the quenching associated with the surface of the phosphor.

Recently, the critical role of surface processes on the ESA process probability and thus on the thermometric parameters of SBR thermometers, such as relative sensitivity, temperature resolution, and useful temperature range, has been highlighted using $\text{YAG}:\text{Nd}^{3+}$ as an example. Unfortunately those studies were limited only to three types of samples of drastically varying sizes, *i.e.*, nanocrystals, microcrystals, and ceramics.²⁹ The main insight coming from this research is that the growing ratio of ions on the surface of the phosphor relative to those in the volume part, observed with decreasing particle size, makes the luminescence more sensitive to the interactions with functional groups of the medium or with ligands attached to the nanocrystal surface.⁴¹ The interactions at the surface of the structure can strongly contribute to the depopulation of the excited $^4I_{11/2}$ level of Nd^{3+} ions, since the energy difference

of less than 2000 cm^{-1} occurring between the $^4I_{9/2}$ and $^4I_{11/2}$ states can be effectively overcome by at least one vibrational mod of the OH^- groups.

For biomedical applications, the size of thermometers must be within the nanometer scale, so the consideration of the impact of size and surface effects on the thermometric properties of SBR luminescent thermometers should be very carefully investigated and evaluated. Therefore, in this work, the detailed studies of the influence of the particle size on the thermometric properties of the Nd^{3+} based luminescent thermometer operating in the SBR mode are presented. Since the probability of nonradiative depopulation of the excited states is affected by the host phonons, the studies presented here were conducted on the low-phonon host material, namely $\text{NaYF}_4:2\%\text{Nd}^{3+}$ nanocrystals with sizes ranging from 10 to 200 nm.

Experimental

Materials preparation

The materials were synthesized by thermal decomposition method in boiling oleic acid and octadecene acting as solvents.⁴²

Neodymium(III) acetate hydrate ($(\text{CH}_3\text{CO}_2)_3\text{Nd}\cdot 3\text{H}_2\text{O}$ with 99.9% purity), yttrium(III) acetate hydrate ($(\text{CH}_3\text{CO}_2)_3\text{Y}\cdot 3\text{H}_2\text{O}$ with 99.9% purity), gadolinium(III) acetate hydrate ($(\text{CH}_3\text{CO}_2)_3\text{Gd}\cdot 3\text{H}_2\text{O}$ with 99.9% purity), ammonium fluoride (NH_4F of 98% purity), acetic acid ($\text{CH}_3\text{CO}_2\text{H}$ of 99% purity), oleic acid ($\text{CH}_3(\text{CH}_2)_7\text{CH}=\text{CH}(\text{CH}_2)_7\text{COOH}$ of 90% purity) and 1-octadecene ($\text{CH}_3(\text{CH}_2)_{15}\text{CH}=\text{CH}_2$ with 90% purity) were purchased from Sigma Aldrich. Sodium hydroxide (NaOH with 99.8% purity), ethanol ($\text{C}_2\text{H}_5\text{OH}$, 96% pure p.a.), methanol (CH_3OH , 99.8%), *n*-hexane (C_6H_{14} , pure p.a.) and chloroform (CHCl_3 , 98.5%) were purchased from POCH S.A. (Poland). All of the chemical reagents were used as received without future purification.

The synthesis was performed for 1.50 mmol of final product. Depending on the dopant ion concentration, appropriate amounts of neodymium acetate and either yttrium acetate or gadolinium acetate depending on the matrix synthesized (NaYF_4 , NaGdF_4) were placed in a 250 mL three-neck round-bottom flask along with 22.5 mL of octadecene and 9 mL of oleic acid. The solution was then magnetically stirred and heated slowly to $140\text{ }^\circ\text{C}$ under vacuum conditions and further stirred at this temperature for 30 min to form $\text{Y}(\text{oleate})_3$ and to remove residual water and oxygen. At the same time, 0.22222 g (6 mmol) of ammonium fluoride and 0.14999 g (3.75 mmol) of sodium hydroxide were weighed into another vessel and 10 mL of methanol was added and magnetically stirred together. In the next step the temperature of oleates was reduced to $50\text{ }^\circ\text{C}$ and the atmosphere was changed from vacuum to a gentle flow of nitrogen. When the temperature of the solution reached $50\text{ }^\circ\text{C}$, the methanol solution of NaOH and NH_4F was added quickly to the flask through the side neck and stirred under these conditions for 30 minutes. After this time, the temperature was increased to $85\text{ }^\circ\text{C}$ and the conditions were changed to vacuum to completely evaporate the methanol from the



reaction mixture. After the methanol evaporation the reaction temperature was increased to 300 °C as quickly as possible and maintained at this temperature for 60 minutes under the nitrogen flow. The final transparent solution was then cooled to room temperature. The NPs were precipitated by addition of ethanol and isolated by centrifugation at 10 000 rpm for 10 min. For purification, the resulting pellet was dispersed in a minimal amount of *n*-hexane and again precipitated with excess ethanol. The UCNPs were isolated by centrifugation at 14000 rpm for 10 min. The final product stabilized with OA ligands was dispersed in 3.75 cm³ of chloroform (CHCl₃).

Samples for spectroscopic measurements were prepared by dropping 0.75 μL onto a glass slide and evaporating the chloroform.

Characterization

Powder diffraction data were obtained using a PANalytical X'Pert Pro diffractometer equipped with an Anton Paar TCU 1000 N Temperature Control Unit using Ni-filtered Cu K α radiation ($V = 40$ kV, $I = 30$ mA). Transmission electron microscope (TEM) images were performed with the Philips CM-20 SuperTwin transmission electron microscope, operating at 160 kV. A drop of the suspension was put on a copper microscope grid covered with carbon. Before the measurement, the sample was dried and purified in a H₂/O₂ plasma cleaner for 1 min. The hydrodynamic size of nanoparticles was determined by dynamic light scattering (DLS) conducted by Malvern Zeta-Sizer Nano ZS90 at room temperature in quartz cuvette using chloroform as a dispersant. The excitation spectra and luminescence decay profiles were measured using the FLS1000 Fluorescence Spectrometer from Edinburgh Instruments equipped with 450 W Xenon lamp and μ Flash lamp as an excitation sources and R5509-72 photomultiplier tube from Hamamatsu in nitrogen-flow cooled housing as a detector. To carry out the temperature measurement, the temperature of the sample was controlled using a THMS 600 heating-cooling stage from Linkam (0.1 °C temperature stability and 0.1 °C set point resolution). The emission spectra were measured using the 808 nm and 1060 nm excitation lines from laser diodes (LD) and a Silver-Nova Super Range TEC Spectrometer from Stellar-net (1 nm spectral resolution) as a detector.

Results and discussion

Although there are two known polymorphs of the NaYF₄ structure, *i.e.*, crystallographic α -cubic and β -hexagonal phases, for single-band ratiometric thermometry applications, tetrafluorides crystallizing into a β -polymorph are more suitable. This is due to the fact that interionic interactions *via* cross-relaxation which have a beneficial impact on the intensity of ESA induced luminescence are at least an order of magnitude stronger in the hexagonal phase than in its cubic counterpart and that much more efficient luminescence is observed in the hexagonal phase,^{43,44} which is one of the crucial requirements of luminescence thermometry. In the crystallographic structure

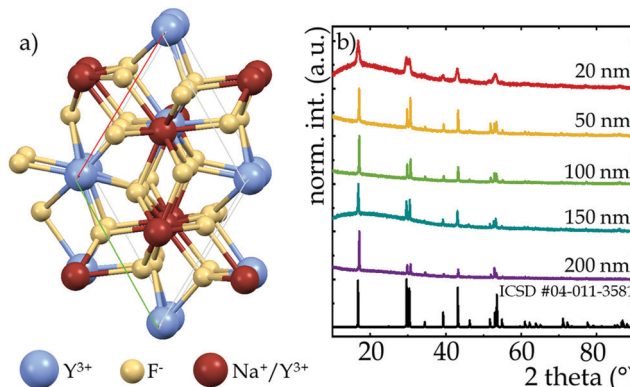


Fig. 1 Visualization of Na⁺ and Y³⁺ sites in the hexagonal NaYF₄ structure (a); X-ray diffraction patterns of NaYF₄:2%Nd³⁺ nanocrystals with different sizes (b).

of β -phase of the NaYF₄, an ordered arrangement of F⁻ ions, which neighbor two cationic positions can be found. One of them is completely filled with Y³⁺ ions while the other has a mixed occupancy with Na⁺ and Y³⁺ ions (Fig. 1a). The NaYF₄ host adopts the $P\bar{6}$ symmetry group⁴⁵ and the unit cell parameters in the hexagonal phase take the values $a = b = 5.96$ Å and $c = 3.53$ Å. X-Ray diffraction patterns of β -NaYF₄:2%Nd³⁺ nanocrystals are presented in Fig. 1b. The diffraction reflections of synthesized materials correlate well with the reference pattern ICSD 04-011-3581, which confirms that the phase purity of samples was not affected by particles size. Nevertheless, one can see the broadening of the Bragg reflections relative to the standard pattern, which is due to the small size of the nanocrystallites.^{46–48}

The size modifications were made by changing the ratio of NH₄F to RE³⁺ ions. As shown by Qiyue *et al.*,⁴² the concentration of F⁻ can control the nucleation rate,⁴⁹ so that the size of NaYF₄ nanoparticles can be effectively controlled by changing the NH₄F amount. For the stoichiometric F⁻/RE³⁺ ratio, tiny nanoparticle sizes on the order of 20 nm were obtained, however, by reducing the amount of NH₄F from 4 mmoles (stoichiometric) to 2 mmoles, *i.e.*, causing a shortage of F⁻ ions relative to RE³⁺ ions, the size of the nanocrystals increased up to 200 nm. However, changes in nanocrystal size due to changes in F⁻ ion concentration do not occur linearly but exponentially (see Fig. S1, ESI[†]). Size alterations induced by changes in the NH₄F/RE³⁺ ratio were confirmed by TEM images (Fig. 2) and the corresponding hydrodynamic size distributions determined from DLS measurements (Fig. S2, ESI[†]). The morphological studies proved that the average sizes of the nanocrystals were 15, 60, 100, 150 and 200 nm for the NH₄F/RE³⁺ ratios equal 4, 3, 2.5, 2.25, 2, respectively (Fig. 2). Furthermore, it showed that non-aggregated nanocrystals were obtained, however, the greater the F⁻ ions deficiency, the wider the size distribution was recorded.

The spectroscopic properties of all the synthesized materials were investigated. As presented in the simplified energy level diagram of Nd³⁺ ions (Fig. 3a), two wavelengths were used to excite the NaYF₄:2%Nd³⁺ nanocrystals. Both wavelengths were



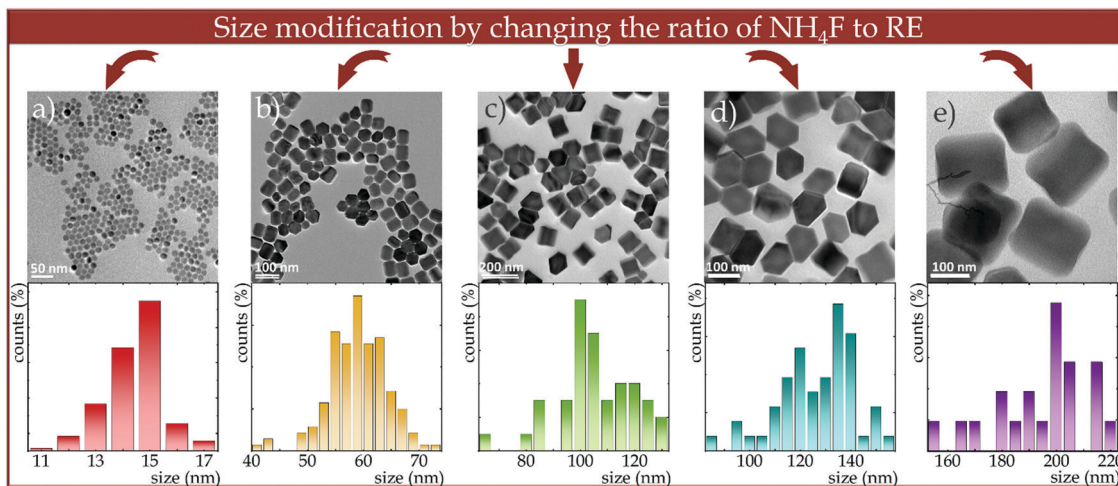


Fig. 2 Effect of changing the NH_4F to RE^{3+} ratio on $\text{NaYF}_4:2\%\text{Nd}^{3+}$ nanocrystals size and morphology and corresponding size distributions determined from TEM images using Feret's method.

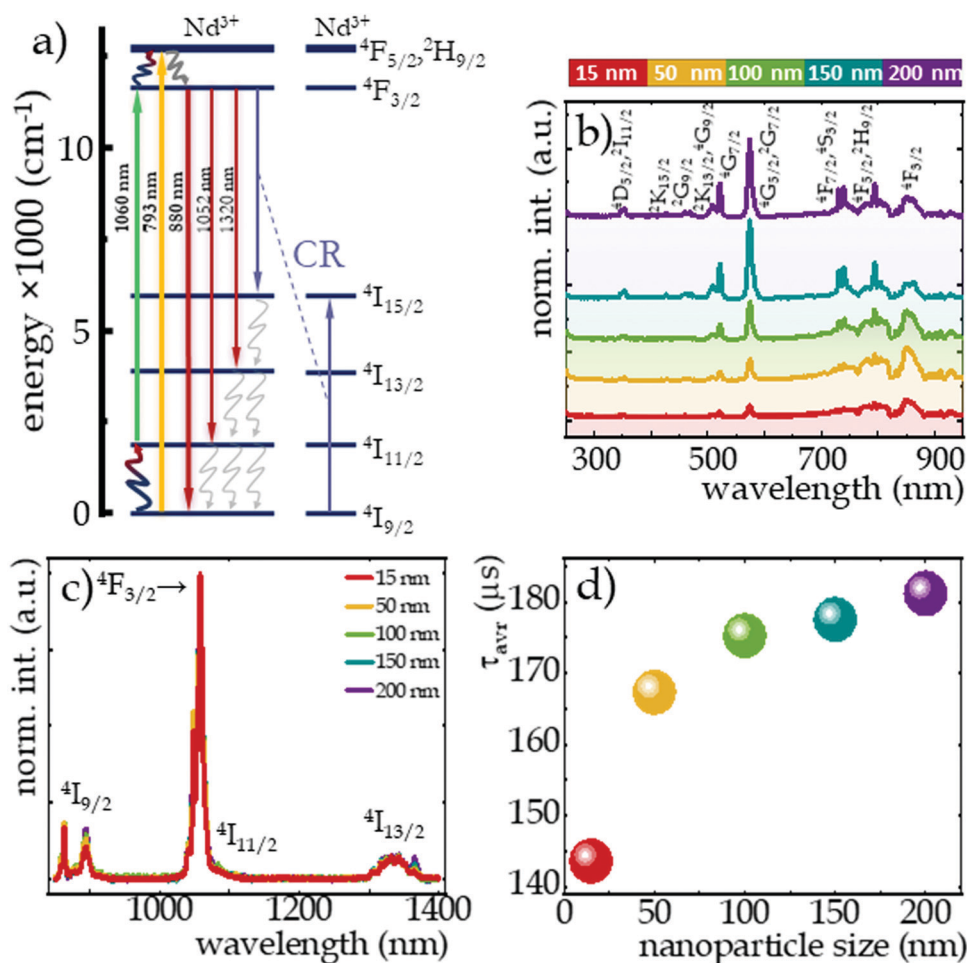


Fig. 3 Simplified energy diagram of Nd^{3+} ions (a); impact of the $\text{NaYF}_4:2\%\text{Nd}^{3+}$ nanoparticles size on the excitation spectra monitored at 1058 nm (b), on the emission spectra upon 793 nm excitation (c) and on the luminescence decay profiles of the excited ${}^4\text{F}_{3/2}$ state measured at 123 K (d). The average decay times of ${}^4\text{F}_{3/2}$ level at 123 K are presented in inset of (d).

selected so that the excitation of the nanocrystals occurred in the near-infrared region. One of them is the GSA-matched

793 nm wavelength, which allows ${}^4\text{I}_{9/2} \rightarrow {}^4\text{F}_{5/2}, {}^2\text{H}_{9/2}$ electron transfer traced by immediate nonradiative relaxation to the



$^4F_{3/2}$ level and the other is the ESA-matched 1060 nm wavelength, which results in electron transfer from the excited $^4I_{11/2}$ state to the $^4F_{3/2}$ level. Although in general probability of the absorption from the ground state exceeds ESA, with a sufficiently large filling of the $^4I_{11/2}$ excited level, this second type of excitation can also occur. Several mechanisms may contribute to the population of this level, such as: (1) $\{^4F_{3/2}; ^4I_{9/2}\} \rightarrow \{^4I_{15/2}; ^4I_{15/2}\}$ cross-relaxation followed by a nonradiative transition to the $^4I_{11/2}$ state (purplish and gray wavy arrows, respectively, in Fig. 3a); (2) as a result of $^4F_{3/2} \rightarrow ^4I_{11/2}$ emission transitions, which are characterized by a nearly 60% branching ratio; or (3) by thermalization of the excited level with electrons from the ground state according to the Boltzmann distribution (navy blue-red wavy arrow). On the other hand, the $^4I_{11/2}$ state may be also depopulated by the surface-related quenching processes. This is due to the small energy distance between the $^4I_{11/2}$ and $^4I_{9/2}$ levels, which is smaller than *e.g.* one vibrational mod of the OH^- groups of water or COOH of the surface ligands. It is also worth noting that the increase of the particle size leads to the increase of the crystallinity of the particles. Surface-related nonradiative quenching caused by structural defects associated with the modified coordination of surface ions and by quenchers attached to the nanoparticle surface, is particularly highlighted in nanosized phosphors.⁴¹ An increase in the intensity of the excitation bands monitored at 1058 nm with growing size of the nanocrystals was noted (Fig. 3b). This is presumably due to an increase with nanocrystal size in the ratio of ions present in the volume relative to those located on the nanocrystal surface, so that surface-related quenching effects are attenuated. Furthermore, the ratio of the intensity of the excitation band at 350 nm ($^4I_{9/2} \rightarrow ^4D_{5/2}, ^2I_{11/2}$) to the band at 580 nm ($^4I_{9/2} \rightarrow ^4G_{5/2}, ^2G_{7/2}$) was observed to be size dependent (Fig. S3, ESI[†]). Its value increased with increasing nanocrystal size from 0.043 in 15 nm nanocrystals to 0.133 in 200 nm nanocrystals. It may indicate a change in the symmetry of the crystal field acting on Nd^{3+} ions. Although the largest absorption cross section is evident for the band at 580 nm associated with the $^4I_{9/2} \rightarrow ^4G_{3/2}, ^2G_{7/2}$ transition, as mentioned above, the 793 nm wavelength associated with the $^4I_{9/2} \rightarrow ^4F_{5/2}, ^2H_{9/2}$ transition was used to provide excitation in the NIR region. As can be seen in Fig. 3c, independently on the excitation wavelength used three main emission bands, associated with the radiative depopulation of the $^4F_{3/2}$ state can be observed in the emission spectra at 880 nm ($^4F_{3/2} \rightarrow ^4F_{9/2}$), 1058 nm ($^4F_{3/2} \rightarrow ^4F_{11/2}$) and 1335 nm ($^4F_{3/2} \rightarrow ^4F_{13/2}$). In addition, the effect of size on the intensity of individual components of the emission band at 880 nm associated with electronic transitions from R_1 and R_2 of the Stark sublevels to the Stark Z_1 sublevel of the $^4I_{9/2}$ state was noted (Fig. S4a, ESI[†]). The intensity of the short-wavelength component decreases with size, which can be explained in terms of the self-filtering effect. When the size of the particle increases, the average optical path that the photons emitted by the Nd^{3+} ions must travel is longer, and thus the probability of its reabsorption is higher. An effect of nanocrystals size changes was also noted in subtle changes in the kinetics of the excited $^4F_{3/2}$ level measured at 123 K (Fig. S5, ESI[†]). Due to

the nonexponential behavior of the decay curves, the average decay times (according to eqn (S1) and Fig. S6, ESI[†]) were determined. As shown in Fig. 3d, with decreasing size of nanoparticles, a gentle reduction in decay times from $\tau_{\text{avr}} = 181 \mu\text{s}$ for 200 nm nanoparticles to 144 μs recorded in 15 nm $\text{NaYF}_4:2\%\text{Nd}^{3+}$ was evident (see Table S1, ESI[†]). This may be a confirmation of the stronger surface-related luminescence quenching when the smaller the size of the structure analyzed.

The influence of temperature on the emission intensity of Nd^{3+} ions and thermometric parameters of the SBR-based luminescence thermometer is shown in Fig. 4. The clearly opposite direction of thermally induced changes in luminescence intensity of $^4F_{3/2} \rightarrow ^4I_{9/2}$ emission band at 862 nm upon excitation of ESA *versus* GSA has been evidenced (Fig. 4a and b for ESA and GSA, respectively). A qualitative comparison between nanocrystals of different sizes is provided by the integral emission intensities under ESA (Fig. 4a) and GSA (Fig. 4b) excitations. As expected, at low temperatures the occupancy of the $^4I_{11/2}$ level was negligibly small and ESA-excited luminescence was not observed, however, beyond a certain threshold temperature the process occurred and an intense emission band was visible at high temperatures. The threshold temperature for which the process occurred shifted toward lower temperatures as the size of the nanocrystals increased (Fig. 4c). This is due to the fact that the smaller the nanocrystals, the stronger the surface-related quenching interactions that cause depopulation of the $^4I_{11/2}$ state. In smaller sized nanocrystals, the number of ions on the surface may dominate over that in the bulk part, thus stronger surface-related quenching is noted. Consequently, the smaller the nanocrystals, the higher the thermal energy which competes with the depopulation processes must be delivered to the system to effectively populate the level from which the ESA process occurs. Regardless of size, the trend of change was similar for all nanocrystals, but the rate of change increased with the size. That is, at 473 K, a 200-fold increase in the intensity of ESA-excited luminescence relative to the initial intensity was seen in 20 nm nanocrystals, while a 4-fold enhancement of this effect, and thus an 800-fold increase in intensity relative to the initial, was recorded for 200 nm nanocrystals. Depending on the size, surface effects may have different impact on the nonradiative depopulation of the excited level, resulting in weaker intensities in the smaller nanocrystals. It is worth noting here that at high temperatures, two additional bands were visible in the emission spectrum at 800 nm (above 288 K) and 750 nm (above 400 K), which are due to the thermalization of the higher $^4F_{5/2}, ^2H_{9/2}$ and $^4F_{7/2}, ^4S_{3/2}$ levels and their radiative depopulation to the $^4I_{9/2}$ level (inset on Fig. 4a). In the case of GSA-excited luminescence, the intensity of the same $^4F_{3/2} \rightarrow ^4I_{9/2}$ band manifests a gradual decrease at elevated temperatures, however, no clear size trend of the changes was observed. The observed decrease in intensity is the result of multiphonon relaxation. However, in this case the energy difference between emitting $^4F_{3/2}$ state and the next lower-lying $^4I_{15/2}$ state, equals around 6000 cm^{-1} , thus



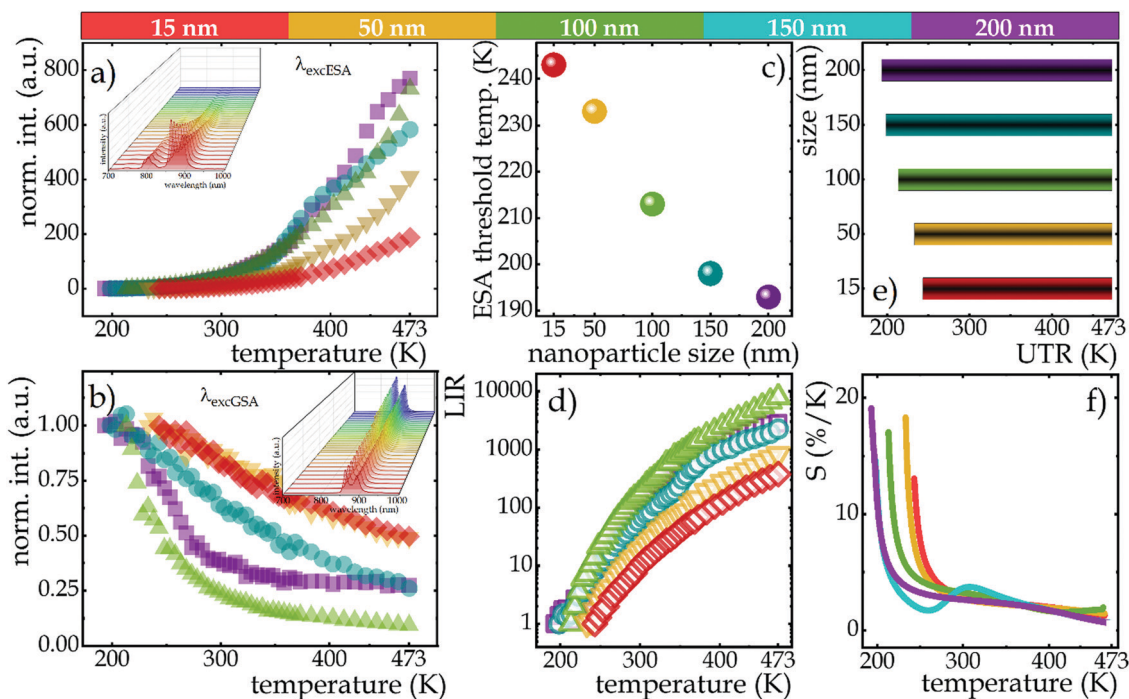


Fig. 4 The impact of the size on thermal evolution of normalized integral emission intensities of $\text{NaYF}_4:2\%\text{Nd}^{3+}$ nanocrystals upon ESA (a) and GSA (b) excitation, the insets show the thermal evolution of ${}^4\text{F}_{3/2} \rightarrow {}^4\text{I}_{9/2}$ emission band at 862 nm of representative $\text{NaYF}_4:2\%\text{Nd}^{3+}$ nanocrystals with 50 nm diameter upon 1060 nm and 793 nm excitations; the influence of nanocrystal size on the threshold temperature for the occurrence of ESA-induced luminescence (c); the temperature dependent LIR values of $\text{NaYF}_4:2\%\text{Nd}^{3+}$ nanocrystals with different size (d); the influence of nanocrystals size on the useful temperature range (e); the relative sensitivities of Nd^{3+} -doped NaYF_4 based SBR LTs (f).

resulting in significantly lower efficiency of the nonradiative depopulation process related to the surface effects.

This opposite nature of the emission intensity variation with excitation wavelength is a great basis for developing a strongly temperature-dependent LIR parameter, according to the following formula:

$$\text{LIR} = \frac{\int {}^4\text{F}_{3/2} \rightarrow {}^4\text{I}_{9/2}(\text{ESA})d\lambda}{\int {}^4\text{F}_{3/2} \rightarrow {}^4\text{I}_{9/2}(\text{GSA})d\lambda} \quad (1)$$

Due to the greater susceptibility of ESA-excited luminescence to temperature changes, the nature of the LIR changes is mainly determined by the trend of the integral emission intensities upon ESA excitation (Fig. 4d). Regardless of the size, the observed trends were very similar, however, the degree of LIR enhancement at 473 K relative to its initial value strongly depended on the size. For 20 and 50 nm nanocrystals, increasing the temperature to 473 K resulted in 400-fold and 800-fold improvement in LIR values, respectively. However, for 100 nm nanocrystals, the increase in LIR value relative to the initial one reached more than 8000-fold improvement. A threshold temperature above which excited state absorption begins to occur, resulting from the competing processes of thermal population and surface-related depopulation of the excited level manifests as a narrowing of the useful temperature range (UTR) for smaller nanocrystals (Fig. 4e).

The ability of SBR-based luminescent thermometers can be quantified by the relative sensitivity to temperature changes.

This is a standardized parameter, determined as follows:

$$S_R = \frac{1}{\text{LIR}} \cdot \frac{\Delta\text{LIR}}{\Delta T} \times 100\% \quad (2)$$

where ΔLIR represents the change of LIR value in response the change of temperature ΔT . Regardless of the size, the maximum sensitivity was obtained at low temperatures of the UTR corresponding to a particular sample and for 200 nm $\text{NaYF}_4:2\%\text{Nd}^{3+}$ nanoparticles was as high as $19.1\% \text{K}^{-1}$ (Fig. 4f). However, in this temperature range, the signal-to-noise ratio of ESA-excited luminescence was small, resulting in a poor reliability of the results. As the temperature increased, the relative sensitivity decreased, reaching close to $5\% \text{K}^{-1}$ at 273 K and despite the continuing decrease maintaining a value above $0.7\% \text{K}^{-1}$ throughout the entire temperature range analyzed. The ability of the analyzed nanocrystals to determine temperature can be established by calculating the temperature resolution, according to the following formula

$$\delta T = \frac{1}{S_R} \cdot \frac{\delta\text{LIR}}{\text{LIR}} \quad (3)$$

As expected, in the low temperature range as a result of the small signal-to-noise ratio, the temperature resolution exceeded 1 K, however in the temperature range above 300 K a significant improvement was seen and values close to 0.1 K were observed, with no critical effect of size on these values (Fig. S7, ESI†). These results indicate that size changes within the nanometer scale do not drastically affect the relative



sensitivities and temperature resolutions of single-band ratiometric luminescence thermometers. Therefore, it is possible to use nanostructures in SBR-based luminescence thermometry without worrying about falsification of results due to reduction in nanocrystal size by release of ions in the medium. The influence of the particle size for NaYF₄:Nd³⁺ nanocrystal is less spectacular in respect to that observed for YAG:Nd³⁺.²⁹ This results from the fact that in the previous studies much significant difference in the particle size were considered (nanocrystals, microcrystals and optical ceramic). It is worth noting that the negative influence of the surface quenching processes on the thermometric performance of the SBR thermometers can be also eliminated by the passivation of the nanoparticle's surface using *e.g.* core shell architecture. This strategy will be evaluated in further studies.

Conclusions

In this paper, the effect of size changes in nanometer-scale on the thermometric properties of single-band ratiometric luminescent thermometers based on NaYF₄:2%Nd³⁺ nanocrystals was analyzed. Five sizes of NaYF₄:2%Nd³⁺ nanoparticles in the range from 20 to 200 nm were analyzed. Thermometric studies were performed using two excitation wavelengths matching absorption from the ground state ($\lambda_{\text{EXC}} = 793 \text{ nm}$) and from the excited state ($\lambda_{\text{EXC}} = 1060 \text{ nm}$) in order to investigate the suitability of analyzed materials for single-band ratiometric luminescence thermometry. The emission intensity of the band at 880 nm associated with the ⁴F_{3/2} → ⁴I_{9/2} electronic transition upon the mentioned excitation wavelengths was used to determine the temperature remotely. Due to the divergent effects of two different physical processes, *i.e.* the thermal population of the excited ⁴I_{11/2} state and the nonradiative multiphonon quenching of the ⁴F_{3/2} level, two opposite trends were observed on the thermally induced intensity changes depending on the excitation wavelength. As a consequence of nonradiative multiphonon quenching, a decrease in the intensity of GSA-excited luminescence was observed, while the thermal population of the ⁴I_{11/2} state resulted in a strong and dynamic increase in the intensity of ESA-excited luminescence. The determined temperature-dependent LIR parameter showed very pronounced changes with increasing temperature and the trend was similar regardless of nanocrystals size, however, the rate of change already depended on it. The least pronounced increase in LIR value at 473 K relative to the initial value was observed in the smallest nanocrystals analyzed, which is due to the existence of a higher number of ions affected by the surface-related quenching of the ⁴I_{11/2} state population, which is due to the fact that the smaller the nanocrystals are, the higher the ratio of ions on the surface to those in the volume. Moreover, for smaller nanocrystals the threshold temperature for which the ESA process occurs is shifted towards higher temperatures. Consequently, the smaller the nanocrystals, the narrower the useful temperature range was observed. Furthermore, although the size of the nanocrystals affected the probability of ESA and

the UTR, it does not critically affect the sensitivity values over a wide temperature range from 273 to 473 K. The high relative sensitivities are also reflected in terms of temperature resolutions that at temperatures above 273 K reach values of 0.1–0.2 K. Therefore, although a significant effect of crucially changing size at the nanoparticle–microparticle–ceramic scale has been previously demonstrated, no drastic changes in the values of thermometric parameters such as relative sensitivity or temperature resolution were observed for changes at the nanometer scale. Consequently, SBR-based luminescent thermometers can be successfully applied without fear that size reduction, *e.g.* due to possible ion release, could generate some falsification of results.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by National Science Center Poland (NCN) under project no. UMO-2019/33/N/ST5/00011.

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