



Cite this: *J. Mater. Chem. C*, 2022, 10, 2390

Methylamine gas healing of perovskite films: a short review and perspective

Lianzheng Hao,^{ab} Qiangqiang Zhao,^c Xiao Wang,^a Shuping Pang^{id}*^a and Guanglei Cui^{id}*^a

The first challenge in the commercialization of perovskite solar cells (PSCs) is how to easily fabricate large-scale, high-quality perovskite films. Although some methods have been developed for the fabrication of large-scale films, the film quality still needs to be improved. The MA⁰ healing method, as a post-processing technique, can be perfectly combined with other large-scale methods. However, determining the mechanism of the whole process is crucial for commercial production and subsequent research. Herein, we summarize the research progress on the MA⁰ healing method. Firstly, we introduce the interaction during the liquification process including the interaction of amines with Pb²⁺, the generation of hydrogen bonds, the influence of H₂O, and the reactions between different amines. Then, we systematically discuss the dynamic mechanism of adsorption and desorption. Finally, several aspects that need to be further clarified are proposed.

Received 9th September 2021,
Accepted 4th November 2021

DOI: 10.1039/d1tc04302a

rsc.li/materials-c

Introduction

Perovskite solar cells (PSCs) have attracted extensive attention from both the research community and industry due to their advantage of facile fabrication *via* solution method. Recently, the efficiency of solution-processed small-area PSCs has reached 25.5% certified by NREL,¹ while the certified module (802 cm²) just reached 17.9%.² It remains a great challenge to simply fabricate large-scale, high-quality perovskite films and

PSCs with high consistency. Various methods have been developed to fabricate large-scale perovskite films, including but not limited to D-bar and blade coating,^{3,4} slot-die coating,^{5–8} spray coating,^{9–14} dip coating,^{6,15} and air-knife-assisted coating.¹⁶ Generally, solution-processed perovskite polycrystalline films have large roughness and pin hole structures, which is one of the important reasons why the efficiency of the modules is far below that of the small-sized devices.

Considering the uniformity of the perovskite film, methylamine gas (MA⁰) healing is one of the most promising methods due to its easy processing and high consistency.^{17–22} The discovery of MA⁰ healing method was inspired by the reversible transformation of MAPbI₃ perovskite phase during the MA⁰ absorbing and degassing process.²¹ Generally, the rough MAPbI₃ perovskite film with low coverage is visually translucent

^a Qingdao Institute of Bioenergy and Bioprocess Technology, Chinese Academy of Sciences, Qingdao 266101, P. R. China. E-mail: pangsp@qibebt.ac.cn

^b Center of Materials Science and Optoelectronics Engineering, University of Chinese Academy of Sciences, Beijing 100049, P. R. China

^c Qingdao University of Science and Technology, Qingdao 266042, P. R. China



Lianzheng Hao

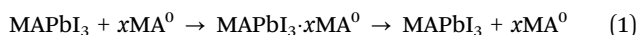
Lianzheng Hao received his BEng Degree from the School of Materials Science and Engineering, Qingdao University of Science and Technology in 2019. He is currently a Master's student in Prof. Guanglei Cui's group at Qingdao Institute of Bioenergy and Bioprocess Technology, Chinese Academy of Sciences. His current research focuses on perovskite materials and solar cells.



Shuping Pang

Shuping Pang joined the group of Prof. K. Müllen at the Max Planck Institute for Polymer Research (MPIP) and obtained his PhD Degree in September 2011. Then he moved back to China and worked as a Professor at the Qingdao Institute of Bioenergy and Bioprocess Technology (QIBEBT), Chinese Academy of Sciences (CAS). His current research focus is on photoelectric devices.

and matte. After quick exposure to MA⁰, the film bleached, forming an intermediate liquid phase of MAPbI₃·xMA⁰. Subsequently, MAPbI₃·xMA⁰ could quickly return to the highly uniform dark perovskite phase when the MA⁰ environment was removed.¹⁷ A schematic diagram of this process is shown in Fig. 1a, and scanning electron microscope (SEM) images of the surface morphology before and after MA⁰ healing are shown in Fig. 1b. This process is generally represented by the following equation:



This method essentially belongs to post-processing technique and it hardly has high requirements for the initial film. Therefore, the MA⁰ healing method can perfectly combine the above-mentioned large-scale methods to easily prepare large-sized, highly uniform and lattice-oriented MAPbI₃ films. Fig. 1c shows the MAPbI₃ module fabricated in the laboratory by this method, where the film has a mirror-smooth surface. Han *et al.* utilized the complex precursors of MA⁰·MAI and MA⁰·PbI₂ combined with the pressure processing method to fabricate a 36.1 cm² MAPbI₃ module, reaching a PCE of 12.1%.²³ Park *et al.* applied D-bar coating to fabricate MAPbI₃ modules on an area of over 100 cm², reaching a PCE of 17.82% based on the MA⁰-assisted MAPbI₃ solution.³ The company Suzhou GCL Nano Co. Ltd successfully applied slot-die printing technology combined with MA⁰ healing method to fabricate large-scale (45 × 65 cm²) highly uniform MAPbI₃ films (Fig. 1d). In the initial stage, researches were mainly focused on how to fabricate high-quality perovskite films through this method. Qi *et al.* reported that introducing MA⁰ in the annealing process can reduce impurities at the perovskite grain boundaries (GBs) and promote continuity between adjacent grains, and thus reduce defects at the GBs, reaching a high PCE of 18.4%.²⁴ Increasing the degassing temperature of the liquid–solid phase transition from MAPbI₃·xMA⁰ to MAPbI₃ could also enhance the quality of the

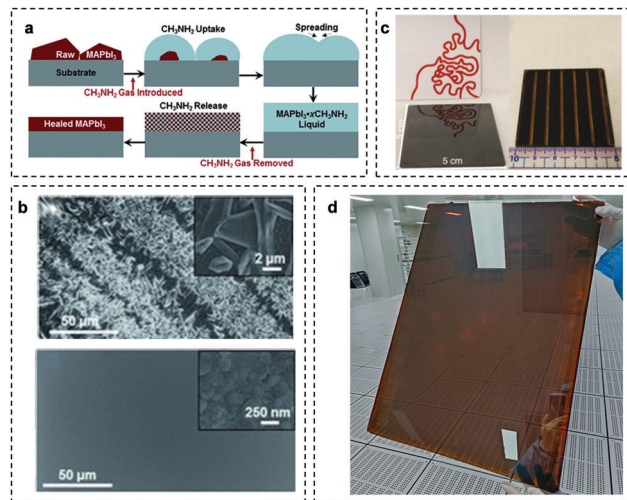


Fig. 1 (a) Schematic diagram of the MA⁰ healing process. (b) SEM micrographs of MAPbI₃ film before (top) and after (bottom) the MA⁰ healing process. Copyright 2015 Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim. (c) The MAPbI₃ module fabricated in the laboratory through MA⁰ healing method. Copyright 2018 Springer Nature. (d) Large-sized MAPbI₃ film (45 × 65 cm²) prepared by MA⁰ healing method (provided by Suzhou GCL Nano Co. Ltd).

perovskite films with crystal grains in the order of tens of micrometers.²⁵ Additive engineering (such as MAcl) is also compatible with the MA⁰ healing method,²⁶ with the formation of a new intermediate phase, MAPbI₃·MAcl·xMA⁰ (PCE = 19.4%). The transformation of the as-prepared MAPbI₃ thin film from MAPbI₃·MAcl·xMA⁰ to MAPbI₃ is a two-step decomposition process, resulting in large, oriented grains spanning the film thickness and ultralow-density GB-network with only vertical GBs. Besides the gas healing of the rough MAPbI₃ perovskite films, MA⁰ was also employed as a cosolvent with the mixing of acetonitrile (ACN) or DMF (PCE > 18%).^{27–29} Recently, it was further discovered that extending the degassing time at an elevated temperature could increase the grain size, reaching millimeters (PCE = 21.36%).³⁰ Large-grain perovskite films possess a low trap-state density, long charge carrier lifetime, and excellent environmental stability. Through MA⁰ healing treatment, the device stability also has great enhancement due to the improvement in film quality. For example, Qi *et al.* reported that unencapsulated PSCs retained 90% of their initial efficiency after 800 h continuous light illumination under an N₂ atmosphere, while the control device decayed to 90% of its initial efficiency in just 300 h.³¹ Song *et al.* reported that MAPbI₃ devices fabricated using the optimized MA⁰ healing method with ultra-large grains retained 85% of their initial PCEs under illumination after 1000 h in an N₂ atmosphere, whereas the control devices decayed to about 60% of their original PCEs.³⁰ Due to the high uniformity of the perovskite films and the device consistency, this MA⁰ healing method represents a major advance in the scale-up fabrication of high-quality perovskite films and devices.^{3,22,23,31–34} **However, the limitation of the MA⁰ healing method is that it cannot be used for the FAPbI₃.** Recently, some research groups have started studying the related mechanisms of the whole process.^{35–37}



Guanglei Cui

Guanglei Cui received his PhD Degree from the Institute of Chemistry, CAS, in 2005. He then conducted postdoctoral research at the Max Planck Institute for Polymer Research and Solid States Research before joining QIBEBT, CAS, in 2009. He is currently a Professor and Leader of the Biomimetics for Energy Storage group, Director of the Applied Energy Technology Division, and Deputy Director of the Academic

Committee of the QIBEBT. His research topics include sustainable and highly efficient energy-storage materials, all-solid-state batteries, and novel energy devices.

Herein, we summarize the research progress of the MA⁰ healing method. This process can be divided into two segments, namely, liquification and recrystallization. In the case of the liquification process, we summarize the internal chemical interaction, which mainly include the interaction of amines with Pb²⁺, the generation of hydrogen bonds, the influence of H₂O and the reactions between different amines. For the recrystallization process, the dynamic mechanism of adsorption and desorption is systematically discussed. Finally, we propose several issues that need to be further investigated.

The internal chemical interaction

The liquification of perovskite materials in MA⁰ atmosphere is the basis of gas-healing technology. The understanding of underlying chemical reaction and intermolecular interaction is constantly being updated. Initially, Zhou *et al.*¹⁷ supposed that the lone-electron pair in the nitrogen (N) atom of MA⁰ molecules interacts with the PbI₆ octahedra, which leads to a liquefied state. Jen *et al.*³⁸ applied MAPbBr₃ and MASnI₃ as the raw film to determine whether the B-site metal ions and X-site halogen make a difference. The transformation phenomenon of MAPbBr₃ and MASnI₃ is very similar with MAPbI₃, which indicates that these two aspects are not determinant. Then, the interaction between MAPbI₃ and MA⁰ was considered as a neutral ligand coordinated with the Pb(II) changing the three-dimensional lead halide framework.^{39,40} Simultaneously, Rand *et al.*⁴¹ reported that proton transfer occurs from PbI₂ to alkylamines, which results in the formation of Pb–MA⁰ bonds and is responsible for the liquefaction. Also, the presence of an MAPbI₃·xMA⁰ complex has also been proved by witnessing the Tyndall effect in this complex.^{36,39,42,43} Cahoon *et al.*⁴⁴ supposed that the gas healing process should be interpreted as amino-deliqescence and amino-efflorescence, and the process of amino-deliqescence is driven by the highly exothermic dissolution of MAPbI₃ in MA⁰. Some reports have indicated that the existence of hydrogen bonding between NH₃⁺ in MAPbI₃ and NH₂ in MA⁰ is also very significant.^{35,36,39} The formation of MA⁰–MA⁺ dimers yields a situation comparable to ionic liquids, which

is also extremely important for the liquefaction of perovskite materials.³⁷ Thus, here we introduce the liquification process in detail.

(1) The reaction of amine and Pb²⁺

MA⁰ or other amine gases can not only liquify organic-inorganic perovskite structure, but also liquefy the PbI₂ structure.^{23,39,41,45} Qi *et al.*⁴⁵ reported that MA⁰ can penetrate the 2D layered structure of PbI₂ and react with the PbI₆⁴⁺ octahedra. Yan *et al.*³⁹ identified that the MA⁰ molecules act as a type of ligand coordinating with Pb²⁺. They supposed that the coordination was the filling of the empty lead 6p orbital (6s²6p²) by the lone electron pair from the nitrogen of MA⁰ (2s²). The Pb–N binding energy calculated by first principles calculation is about 80.04 kJ mol⁻¹. Recently, Hillebrecht *et al.*^{35,37} studied the structure of the intermediate phase and the change in chemical bonds in detail. They noted that there are actually two intermediate phases under the reversible reaction of MA⁰ and MAPbI₃. When a stream of MA⁰ passes over the PbI₂ or MAPbI₃ powder, a highly viscous liquid is formed, followed by pale-yellow platelets crystallized from it. Through the characterization of single-crystal X-ray diffraction, they obtained a neutral octahedral complex-(MA⁰)₄PbI₂. As shown in Fig. 2a, in this structure, it can be seen that four MA⁰ molecules act as the classical donor ligand to Pb²⁺, and the two iodine anions exist in a *cis*-position. If a stronger MA⁰ stream is injected from a cold trap, the PbI₂ or MAPbI₃ powder quickly liquefies and colorless crystals are precipitated from it. This colorless crystal is a type of new intermediate phase, which is stable at room temperature if it is kept in a closed vial with pure MA⁰ as an equilibrium gas phase. The direct measurement of the crystal indicated the existence of multiple twinning. Through further structure refinement of the cubic data, they found that its structure model is the well-known K₂PtCl₆ type⁴⁶ with octahedral [Pb(MA⁰)₆]²⁺ complexes and isolated I⁻ anions (Fig. 2b). In this structure, Pb²⁺ is only surrounded by the neutral MA⁰ ligand. When the MA⁰ is gradually evaporated, [Pb(MA⁰)₆]I₂ can be converted to (MA⁰)₄PbI₂. The above-mentioned crystal structures provide a crucial basis for

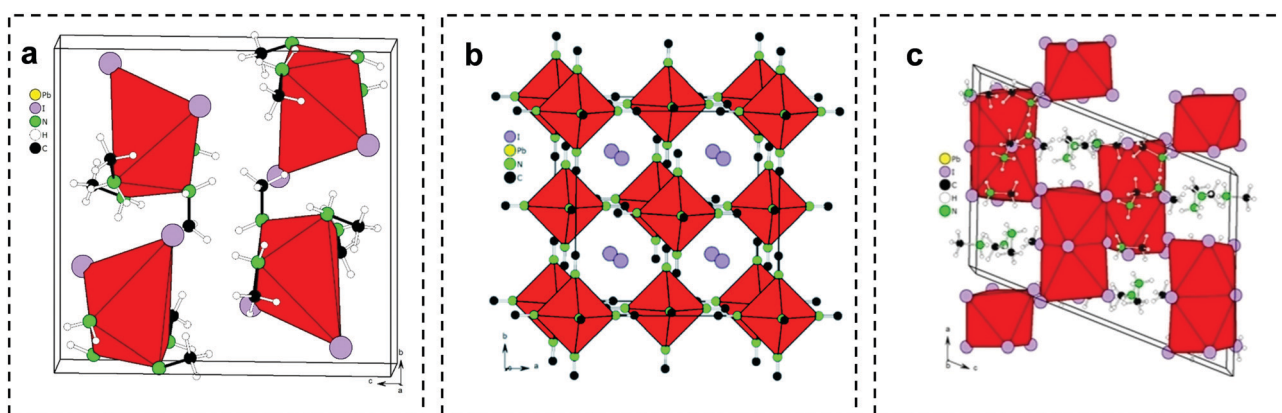


Fig. 2 (a) Crystal structure of (MA⁰)₄PbI₂. Copyright 2021 Wiley-VCH GmbH. (b) Crystal structure of [Pb(MA⁰)₆]I₂. Copyright The Royal Society of Chemistry 2020. (c) Crystal structure of (MA⁰)₅(MA⁰)₂Pb₂I₉ in (010) direction. Copyright 2021 Wiley-VCH GmbH.

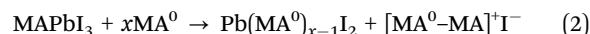
understanding the interaction between MA⁰ and the inorganic framework in perovskite materials. The formation of the liquid state is due to the fact that the MA⁰ molecules coordinate with Pb²⁺ as the ligand destroys the initial perovskite crystal structure, and then change its physical state, but there is no evidence that the MA⁰ healing of perovskite films pass through [Pb(MA⁰)₆]I₂ and (MA⁰)₄PbI₂ to MAPbI₃ finally. Therefore, other liquid intermediates should exist with a disordered structure, which coordinate with a different number of MA⁰ molecules.

(2) The interaction of MA⁰ and MA⁺ (generation of hydrogen bonds)

The crystal structure of MAPbI₃ can be regarded as [PbI₆]⁴⁻ octahedra in three-dimensional (3D) space, where the MA⁺ ions fill the gaps among the octahedra.^{47–49} Some experiments found that organic iodides such as MAI and MACl could also absorb MA⁰ to form a liquid phase, which is much faster than that in perovskite materials.^{26,43,50} Thus, the interaction between MA⁰ and MA⁺ cannot be ignored. Zheng *et al.*³⁶ confirmed the existence of MA⁰-MA⁺ dimers by applying ¹H NMR. They dispersed MAPbI₃ powder, MA⁰-MAPbI₃ crystals and MA⁰ in DMSO-d₆ for room-temperature ¹H NMR measurements. As shown in Fig. 3a, the ¹H NMR signals of MAPbI₃ appeared at 7.47 and 2.38 ppm, corresponding to the resonance of the

protons adjacent to nitrogen and carbon, respectively, while the ¹H NMR signal of the MA⁰-MAPbI₃ sample appeared at 4.51 ppm, which is halfway between 7.47 ppm for MAPbI₃ and 1.06 ppm for MA⁰. They attributed this chemical shift to the hydrogen bonding between NH₃⁺ in MAPbI₃ and NH₂ in MA⁰. The formation of hydrogen bonding was also verified by temperature-dependent ¹H NMR, as shown in Fig. 3b. The negative correlation between hydrogen bonds and temperature agrees well with the fact that the strength of hydrogen bonding is diminished at elevated temperature.^{51,52} They also found a similar phenomenon in the MA⁰-MAPbCl₃ and MA⁰-MAPbBr₃ systems. The diameter of the MA⁰-MA⁺ dimers is 445 pm, which is about twice that of MA (217 pm). This will easily cause the collapse of the 3D perovskite structures.⁵³

Around the same time, Andreas Hinsch *et al.* also confirmed the existence of MA⁰-MA⁺ dimers through Raman spectra.³⁵ They exposed the MA⁰-MAPbI₃ complex in the air and measured the change in peaks in the Raman spectra. The only band at 290 cm⁻¹ in the liquid state under high pressure (Fig. 3c) is related to the rotational mode of MA⁺, which is in agreement with the result from Park *et al.*³ This rotational mode is normally inactive, and the detectable rotational mode of MA⁺ can be attributed to the lower local symmetry of the PbI₃ framework or the formation of MA⁰-MA⁺. During the process of MA⁰ evaporation, the MA⁺ rotational mode gradually disappeared.⁵⁴ In addition, the bands representing the CH₃-NH₃ rocking and C-N stretching vibrations of the free MA⁺ under vacuum⁵⁵ gradually vanished during crystallization, as shown in Fig. 3c. These results demonstrate the existence of MA⁰-MA⁺ dimers. According to the above-mentioned two types of reaction, the liquification of MAPbI₃ may be driven by two effects, MA⁰ bonding to Pb²⁺ atom *via* the lone electron pair along with the formation of MA⁰-MA⁺ dimers. Thus, the chemical reaction for liquification can be described as follows:



(3) The influence of H₂O

MA⁰ healing is normally performed in a dry environment, which is due to the negative effect of H₂O on the film quality. Qi *et al.*⁴⁵ reported that PbI₂ can be converted to MAPbI₃ in the presence of H₂O during the reaction of PbI₂ and MA⁰. Hillebrecht *et al.*³⁵ also studied this reaction, and they reported that H₂O could react with MA⁰ to produce MA⁺ cations and OH⁻ anions, generating a new compound, [(MA)₅(MA⁰)₂Pb₂I₉], in the MA⁰ healing process. This compound contains [Pb₂I₉]⁵⁺ units with two face-sharing octahedra and MA⁰-MA⁺ dimers bridged by a hydrogen bond (Fig. 2c). The charge balance of the system is achieved by MA⁺ and the dimeric (MA⁰-MA)⁺ cations. (MA)₅(MA⁰)₂Pb₂I₉ will convert to MAPbI₃ by further degassing, but Pb(OH)I remains on the film as a by-product. Walsh *et al.* analyzed the effect of H₂O on the decomposition of MAPbI₃. They found that the kinetic barrier of I-ion migration was greatly reduced and all types of vacancy defects were easier to create deep transition levels in the presence of water. These effects could aggravate the decomposition of the perovskite.⁵⁶

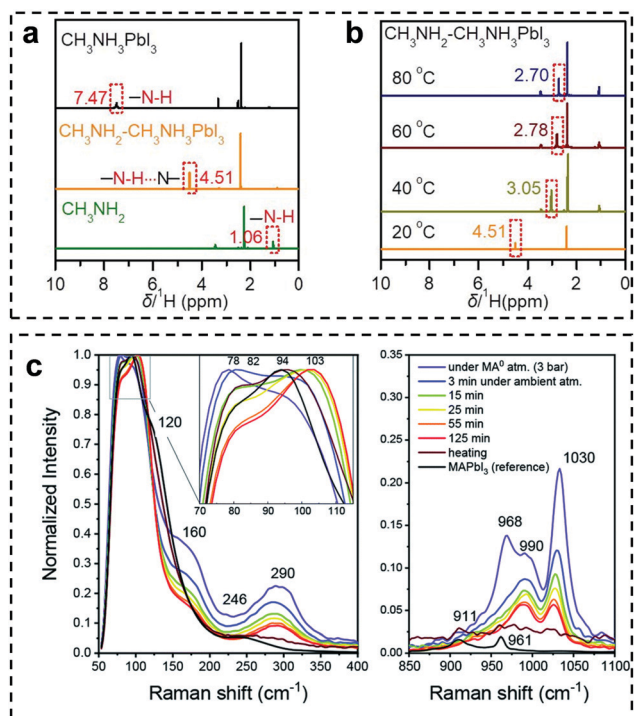
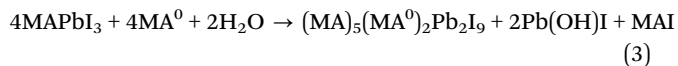


Fig. 3 (a) ¹H NMR spectra of MAPbI₃ powder, MA⁰-MAPbI₃ crystals and MA⁰ dispersed in DMSO-d₆. (b) Temperature-dependent ¹H NMR spectra of MA⁰-MAPbI₃ crystals dispersed in DMSO-d₆. The NMR tube with the precursor sample was heated from 20 °C to 80 °C and tested at the corresponding temperatures *in situ*. Copyright 2020, the American Chemical Society. (c) Raman spectra of a liquified MAPbI₃-xMA⁰ complex during liquid-solid conversion as the MA⁰ leaves the complex. Copyright The Royal Society of Chemistry, 2020.

These results emphasize the importance of avoiding humidity during the MA⁰ healing process, and the liquification process can be represented by eqn (3) with the existence of H₂O.

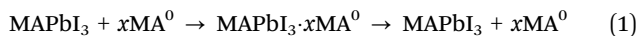


(4) The reactions between different amines

(i) **Substitution reaction between different amines.** Initially, Zhou *et al.* reported that ethylamine (C₂H₅NH₂) or *n*-butylamine (CH₃(CH₂)₃NH₂) can liquefy MAPbI₃ films, but the film cannot completely recover to the MAPbI₃ perovskite phase after the degassing process.¹⁷ Thus, they gave the following equation:

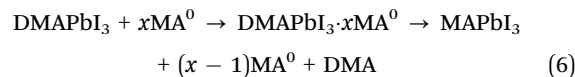


Then Zong *et al.*⁵⁷ transformed a rough, polycrystalline NH₄PbI₃ non-perovskite thin film into a dense, ultrasmooth, textured MAPbI₃ perovskite thin film through the MA⁰ healing process. It is interesting that the yellow NH₄PbI₃ film initially converts to a black MAPbI₃ phase rather than being liquefied directly, and then transforms to MAPbI₃·xMA⁰ under the MA⁰ environment (Fig. 4a). The whole reaction process can be described by the following two equations:



Simultaneously, Pang *et al.*⁵⁸ replaced NH₄PbI₃ with “HPbI₃” ((CH₃)₂NH₂PbI₃)^{59,60} to fabricate high-quality MAPbI₃ films.

It is clear that (CH₃)₂NH₂PbI₃ (DMAPbI₃) is directly liquefied (Fig. 4b), which is different from the behavior of the NH₄PbI₃ crystals in the MA⁰ environment. Thus, the reaction can be described as the following equation:



Actually, the conversion of organic–inorganic hybrid perovskites depends on the environment they are kept in. If a type of amine (monoamine) is adequate in the environment, the perovskite will be converted to the corresponding amine structure. The transformation time depends on the volatilization of the amine in the initial perovskite structure. For example, Wang *et al.*⁶¹ confirmed that MAPbI₃, BA₂PbI₄, and OA₂PbI₄ can transform into each other in the corresponding A site cation gas environment. Chen *et al.*⁶² also reported the mutual transformation of MAPbI₃ and EAPbI₃. Raga *et al.*⁴⁵ gave the transformation of different halide perovskites, as shown in Fig. 4c.

(ii) **Addition–elimination reaction.** Formamidinium lead iodide (FAPbI₃) has become the most widely studied light-absorbing perovskite material, which is ascribed to its suitable bandgap. However, the FAPbI₃ perovskite phase cannot be obtained by the MA⁰ healing method, and its underlying mechanism remained unclear for a long time. Zhou *et al.*⁶³ directly replaced MA⁺ with FA gas at 150 °C, where FA⁰ needed to be *in situ* prepared due to its instability. Moreover, the FA⁰ molecules will form the *s*-triazine structure spontaneously by the addition–elimination reaction.⁶⁴ This high-temperature gas treatment is only a transformation

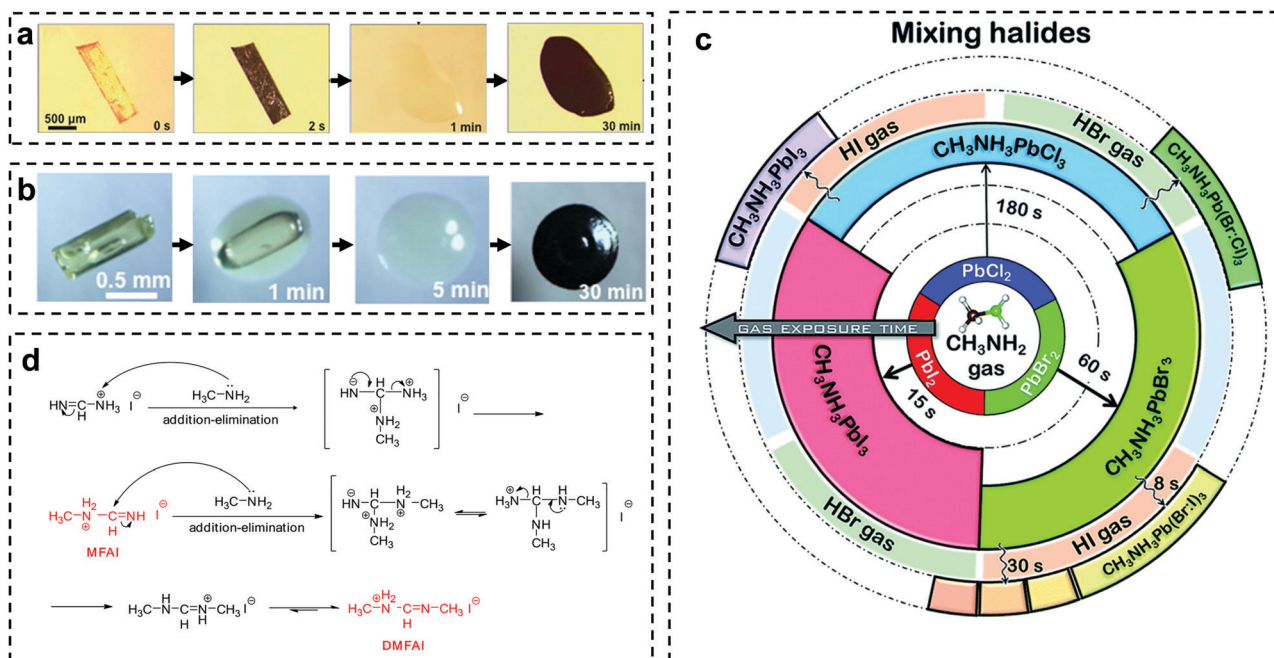


Fig. 4 (a) Optical photos of the morphological changes of NH₄PbI₃ crystal during the MA⁰ healing process. Copyright 2016, the American Chemical Society. (b) Optical photos of the morphological changes of “HPbI₃” (DMAPbI₃) crystal during the MA⁰ healing process. Copyright 2016, Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim. (c) Perovskite conversion from PbX₂ (X = I, Br, and Cl) films exposed to MA gas and to HI and HBr gases. Copyright, The Royal Society of Chemistry, 2016. (d) Addition–elimination reactions between MA⁰ and FAI. Copyright 2020, Elsevier Inc.

process without any modification of the film morphology. Zhang *et al.*⁶⁵ performed MA⁰ healing on a mixed organic cation MA_xFA_{1-x}PbI₃ perovskite film, forming MAPbI₃ perovskite films with a second non-perovskite phase similar to the δ-FAPbI₃ structure. Recently,⁶⁶ Wang *et al.* discovered that FAI will undergo addition–elimination reactions with MA⁰. The reaction mechanism of the reaction is shown in Fig. 4d. The lone-pair electrons of the N atom in MA⁰ have nucleophilicity and the imine bond in FAI is an active electrophilic group, which leads to an addition–elimination reaction to form *N*-methyl formamidinium iodide (MFAI) and NH₃. Similarly, the formed MFAI also has an imine bond, which can undergo a second addition–elimination reaction with MA⁰ to form *N,N'*-dimethyl formamidinium iodide (DMFAI). Therefore, the gas healing of the rough FAPbI₃- or FA-containing perovskite films to a highly uniform one is still a big challenge.

The dynamics of the recrystallization

During the MA⁰ healing process, the next step is desorption-induced recrystallization of the perovskite film. Control of this step is also very important in the quality of perovskite films. Firstly, the absorption capacity determines the state of the intermediate and affects the subsequent recrystallization process. Then, the desorption of MA⁰ affects the generation of a crystal nucleus and its growth process. Thus, it is necessary to determine the dynamic mechanism of these processes.

(1) Adsorption of MA⁰ gas

The first factor affecting the adsorption capacity of MAPbI₃ to absorb MA⁰ is the pressure. Jen *et al.*³⁸ observed that MAPbI₃ films preferred to completely liquefy under a high MA⁰ pressure compared with low pressure. Under high MA⁰ vapor pressure, the diffusion of MA⁰ molecules could quickly break the physical bonds of the initial material (Fig. 5a). In contrast, under a low MA⁰ vapor pressure, the diffusion of MA⁰ is constrained to the initial physical extent (Fig. 5b). This is consistent with the conclusion of Hillebrecht *et al.*,³⁵ who found that the intermediate phase is [Pb(CH₃NH₂)₆]I₂ under high MA⁰ vapor, and that under low MA⁰ vapor pressure is (CH₃NH₂)₄PbI₂.

Temperature is another factor affecting the adsorbing capacity of MAPbI₃ films to absorb MA⁰. Zang *et al.*²⁵ discovered the phenomenon that the liquid MAPbI₃·*x*MA⁰ intermediate phase under steady MA⁰ atmosphere still transform to solid state at elevated temperature. As shown in Fig. 5c, all the MAPbI₃ samples are exposed to nitrogen or MA⁰ atmosphere maintained from 25 °C to 100 °C respectively. It is clear that high temperature is conducive to the degassing of perovskite films, appearing a black perovskite phase. When the temperature exceeds 55 °C, the liquid phase starts transitioning back to the solid state even under MA atmosphere. At 100 °C, no liquid intermediate phase exists.

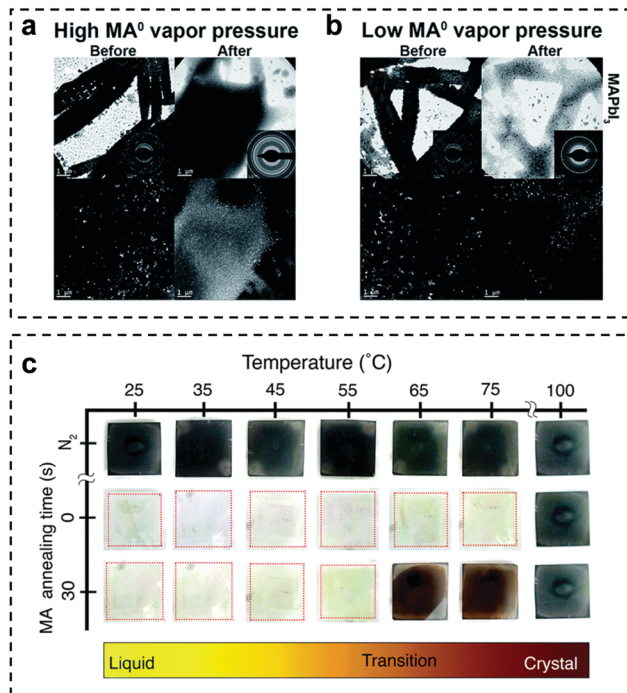


Fig. 5 TEM images of MAPbI₃ before and after exposure to MA⁰ vapor. Each pane contains four pictures, the top line are the bright field images and select area electron diffraction (inset). The bottom line are the dark field images. (a) High MA⁰ vapor pressure and (b) low MA⁰ vapor pressure. Copyright, The Royal Society of Chemistry, 2016. (c) Temperature-dependent annealing of MAPbI₃ under MA gas atmosphere. First line: images of the films under an N₂ flow after 5 minutes. Second line: images of the films immediately after exposure to MA atmosphere. Third line: images of the samples after 30 seconds under a continual flowing MA atmosphere. Copyright, The Royal Society of Chemistry 2016.

(2) Desorption of MA⁰ gas

Temperature also has a great effect on the degassing process. Zang *et al.*²⁵ heated the liquid phase at 100 °C under an MA⁰ atmosphere and the grain size of the film reached tens of microns. Besides, the release speed of MA⁰ is directly related to crystallization. Song *et al.*³⁰ studied the influence of the MA⁰ release speed on the crystallization process from MAPbI₃·*x*MA⁰ to MAPbI₃. In the traditional MA⁰ gas healing process, the pressure of MA⁰ quickly decreases (close to 0) by removing the MA⁰ environment as soon as possible. The quick formation of supersaturation leads to the formation of massive nuclei and fine-grained (~100 nm in size) thin films. Thus, high supersaturation is detrimental to the formation of large grains. Accordingly, slowing the release speed of MA⁰ molecules from the liquid intermediate could lead to a low nucleation density and long growth time of the grains. As shown in Fig. 6a, the grain size is strongly dependent on the release time. The degassing process was conducted at 120 °C, and the grain size increased to millimetres when the degassing time was 30 s. With a further increase in the degassing time, the grain size seemed to no longer expand. Besides, it has been proven that the as-prepared perovskite films possessed a (110)-uniaxial crystallographic orientation and small full width of half

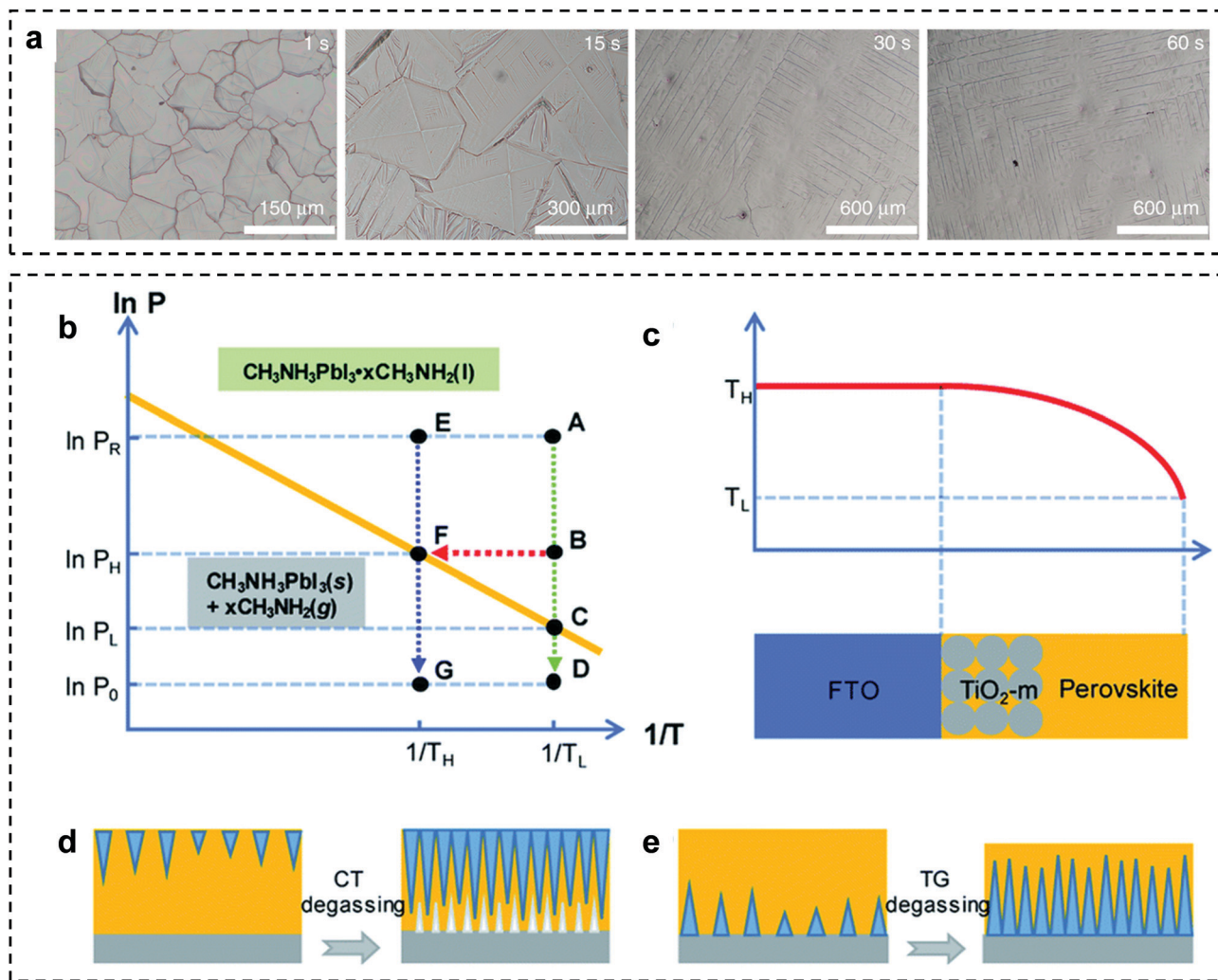


Fig. 6 (a) Optical microscopic images of the MAPbI₃ perovskite films prepared from the liquid intermediate with different degassing times (1, 15, 30, and 60 s). Copyright 2020, Springer Nature. (b) Thermodynamic model of the MA⁰ desorption process. P_R is the pressure of MA⁰ reacting with the perovskite film. P_0 is the actual pressure of the gas desorption process. (c) Diagram of the temperature gradient in the device. Schematic of the crystal growth model with (d) CT-MA (conventional MA⁰ healing) method and (e) TG-MA (temperature gradient-induced MA⁰ healing) method. Copyright The Royal Society of Chemistry, 2020.

maximum (FWHM), exhibiting a very low trap density, long carrier lifetime, and excellent environmental stability.

It is clear that the fast removal of the solvent can highly enhance the uniformity of perovskite films.^{24,27,67,68} Considering this, in solar cell devices, the carrier transport is mainly in the vertical direction. The directional growth of the perovskite layer to form a grain structure in the through-thickness direction is necessary. According to the Clausius-Clapeyron equation^{69,70} (Fig. 6b), the point on the oblique line represents that MAPbI₃· x MA⁰ is in an equilibrium state. When the MA⁰ pressure is lower than the value on the line, MAPbI₃· x MA⁰ begins to degas. In the conventional MA⁰ gas healing process at room temperature, the process from A to D happens instantaneously. Supersaturation is the driving force behind the crystallization, corresponding to the C–D process in Fig. 6b. After MA⁰ is desorbed from the surface, the perovskite film starts the nucleation process from the surface and continues to nucleate

throughout the whole film, which will cause the formation of pinhole structures in the perovskite film (Fig. 6d). It is clear that the higher the temperature, the easier the desorption reaction will occur, as shown in Fig. 6b with the corresponding gas desorption process from E to G. In this case, a longitudinal temperature gradient (TG degassing) was formed to induce the directional growth of perovskite film.⁷¹ The raw FTO/TiO₂/MAPbI₃ substrate was preheated to 60 °C, and then subjected to the MA⁰ healing process with cold MA⁰. The surface temperature of the perovskite film quickly decreased, while the bottom of the film remained at a relatively higher temperature due to the large enthalpy of the glass substrate, as shown in Fig. 6c. It was proven that the crystallization started from the bottom of the film, and then grew upward, as shown in Fig. 6e, resulting in a reduction in defects in the perovskite film and improved carrier injection efficiency in the electron selective layer.

Conclusion and outlook

In this review, we summarized the knowledge on the MA⁰ healing method. Firstly, we introduced the internal chemical interaction. In the current perception, MAPbI₃ liquefaction may be driven by two effects, MA⁰ bonding to the Pb²⁺ atom *via* its lone electron pair along with the formation of MA⁰-MA⁺. (Moreover, humidity should be avoided during the reaction, given that H₂O will participate in the reaction and produce Pb(OH)I as the by-product. Also different types of R-NH₂PbI₃ can transform mutually if the corresponding amine is adequate.) Then, we introduced the dynamics of recrystallization. The temperature, MA⁰ pressure, and release speed of MA⁰ all have a great effect on the recrystallization. Although, the chemical basis of the reaction between MA⁰ and MAPbI₃ is relatively clear, there are still some areas that are undefined and need further investigation.

1. In the process of industrial production, adequate fluidity is needed for the intermediate phase to repair the hole defect in the initial film. We mentioned above that the MA⁰ molecule, acting as a ligand, coordinates with Pb₂⁺. However, the influence of the number of MA⁰ molecules coordinated with Pb₂⁺ on the viscosity and surface tension of the intermediate is still unclear. Thus, a detailed investigation on this will be helpful for industrial production.

2. It has been reported that some MA⁰ molecules can remain at the grain boundary of the MAPbI₃ layers after the MA⁰ healing treatment.³² It is still unclear whether the MA₀ molecules are embedded in the lattice and their impact on the long-term stability of the encapsulated devices.

3. There are some pathways to obtain large grain films such as increasing the temperature, retarding the degassing speed, and adding additives. The obtained perovskite film normally has large grains and excellent uniaxial orientation. However, whether each grain is a single crystal is still under debate. Thus, the defect density in the large gains should be further investigated.

4. Recently, FA-rich perovskite materials have been shown to be more stable and highly efficient photovoltaic materials. However, direct MA⁰ gas healing to repair FAPbI₃ materials is not feasible due to the reaction between FAI and MA⁰, and there are also difficulties in repairing FAPbI₃ materials through direct FA⁰ healing due to the instability of FA⁰. Therefore, a suitable gas healing method for FA-containing perovskite films is still lacking.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the funding of Shandong Energy Institute (SEI I202129), Youth Innovation Promotion Association CAS (Y201944), Natural Science Foundation of Shandong Province (ZR2020KB001 & ZR2018ZB0315). We also

thank Suzhou GCL Nano Co. Ltd for the fabrication of large sized perovskite films, and the Qingdao Key Lab of Solar Energy Utilization & Energy Storage Technology.

References

- N. R. E. Laboratory, Best research-cell efficiency chart <https://www.nrel.gov/pv/cell-efficiency.html>.
- N. R. E. Laboratory, Champion Photovoltaic Module Efficiency Chart, <https://www.nrel.gov/pv/module-efficiency.html>.
- D.-N. Jeong, D.-K. Lee, S. Seo, S. Y. Lim, Y. Zhang, H. Shin, H. Cheong and N.-G. Park, *ACS Energy Lett.*, 2019, **4**, 1189–1195.
- Y. Deng, X. Zheng, Y. Bai, Q. Wang, J. Zhao and J. Huang, *Nat. Energy*, 2018, **3**, 560–566.
- J.-E. Kim, Y.-S. Jung, Y.-J. Heo, K. Hwang, T. Qin, D.-Y. Kim and D. Vak, *Sol. Energy Mater. Sol. Cells*, 2018, **179**, 80–86.
- L. Huang, C. Li, X. Sun, R. Xu, Y. Du, J. Ni, H. Cai, J. Li, Z. Hu and J. Zhang, *Org. Electron.*, 2017, **40**, 13–23.
- M. He, B. Li, X. Cui, B. Jiang, Y. He, Y. Chen, D. O'Neil, P. Szymanski, M. A. El-Sayed, J. Huang and Z. Lin, *Nat. Commun.*, 2017, **8**, 16045.
- F. Ye, H. Chen, F. Xie, W. Tang, M. Yin, J. He, E. Bi, Y. Wang, X. Yang and L. Han, *Energy Environ. Sci.*, 2016, **9**, 2295–2301.
- M. Park, W. Cho, G. Lee, S. C. Hong, M. C. Kim, J. Yoon, N. Ahn and M. Choi, *Small*, 2019, **15**, e1804005.
- W. C. Chang, D. H. Lan, K. M. Lee, X. F. Wang and C. L. Liu, *ChemSusChem*, 2017, **10**, 1405–1412.
- S. Das, B. Yang, G. Gu, P. C. Joshi, I. N. Ivanov, C. M. Rouleau, T. Aytug, D. B. Geohegan and K. Xiao, *ACS Photonics*, 2015, **2**, 680–686.
- Z. Liang, S. Zhang, X. Xu, N. Wang, J. Wang, X. Wang, Z. Bi, G. Xu, N. Yuan and J. Ding, *RSC Adv.*, 2015, **5**, 60562–60569.
- J. H. Heo, M. H. Lee, M. H. Jang and S. H. Im, *J. Mater. Chem. A*, 2016, **4**, 17636–17642.
- D. K. Mohamad, J. Griffin, C. Bracher, A. T. Barrows and D. G. Lidzey, *Adv. Energy Mater.*, 2016, **6**, 1066994.
- H. C. Liao, P. Guo, C. P. Hsu, M. Lin, B. Wang, L. Zeng, W. Huang, C. M. M. Soe, W. F. Su, M. J. Bedzyk, M. R. Wasielewski, A. Facchetti, R. P. H. Chang, M. G. Kanatzidis and T. J. Marks, *Adv. Energy Mater.*, 2016, **7**, 1601660.
- L.-L. Gao, C.-X. Li, C.-J. Li and G.-J. Yang, *J. Mater. Chem. A*, 2017, **5**, 1548–1557.
- Z. Zhou, Z. Wang, Y. Zhou, S. Pang, D. Wang, H. Xu, Z. Liu, N. P. Padture and G. Cui, *Angew. Chem., Int. Ed.*, 2015, **54**, 9705–9709.
- S. R. Raga, Y. Jiang, L. K. Ono and Y. Qi, *Energy Technol.*, 2017, **5**, 1750–1761.
- C. Li, S. Pang, H. Xu and G. Cui, *Solar RRL*, 2017, **1**, 1700076.
- J.-a. Yang, T. Qin, L. Xie, K. Liao, T. Li and F. Hao, *J. Mater. Chem. C*, 2019, **7**, 10724–10742.
- Y. Zhao and K. Zhu, *Chem. Commun.*, 2014, **50**, 1605–1607.
- Y. Zhou and N. P. Padture, *ACS Energy Lett.*, 2017, **2**, 2166–2176.

- 23 H. Chen, F. Ye, W. Tang, J. He, M. Yin, Y. Wang, F. Xie, E. Bi, X. Yang, M. Gratzel and L. Han, *Nature*, 2017, **550**, 92–95.
- 24 Y. Jiang, E. J. Juarez-Perez, Q. Ge, S. Wang, M. R. Leyden, L. K. Ono, S. R. Raga, J. Hu and Y. Qi, *Mater. Horiz.*, 2016, **3**, 548–555.
- 25 D. L. Jacobs and L. Zang, *Chem. Commun.*, 2016, **52**, 10743–10746.
- 26 F. Ji, S. Pang, L. Zhang, Y. Zong, G. Cui, N. P. Padture and Y. Zhou, *ACS Energy Lett.*, 2017, **2**, 2727–2733.
- 27 N. K. Noel, S. N. Habisreutinger, B. Wenger, M. T. Klug, M. T. Hörantner, M. B. Johnston, R. J. Nicholas, D. T. Moore and H. J. Snaith, *Energy Environ. Sci.*, 2017, **10**, 145–152.
- 28 A. J. Ramadan, N. K. Noel, S. Fearn, N. Young, M. Walker, L. A. Rochford and H. J. Snaith, *Chem. Mater.*, 2018, **30**, 7737–7743.
- 29 Z. Liu, J. Hu, H. Jiao, L. Li, G. Zheng, Y. Chen, Y. Huang, Q. Zhang, C. Shen, Q. Chen and H. Zhou, *Adv. Mater.*, 2017, **29**, 1606774.
- 30 H. Fan, F. Li, P. Wang, Z. Gu, J. H. Huang, K. J. Jiang, B. Guan, L. M. Yang, X. Zhou and Y. Song, *Nat. Commun.*, 2020, **11**, 5402.
- 31 Z. Liu, L. Qiu, E. J. Juarez-Perez, Z. Hawash, T. Kim, Y. Jiang, Z. Wu, S. R. Raga, L. K. Ono, S. Liu and Y. Qi, *Nat. Commun.*, 2018, **9**, 1–11.
- 32 K. Huang, C. Wang, C. Zhang, S. Tong, H. Li, B. Liu, Y. Gao, Y. Dong, Y. Gao, Y. Peng and J. Yang, *Org. Electron.*, 2018, **55**, 140–145.
- 33 L. Hong, Y. Hu, A. Mei, Y. Sheng, P. Jiang, C. Tian, Y. Rong and H. Han, *Adv. Funct. Mater.*, 2017, **27**, 1703060.
- 34 Z. Shao, Z. Wang, Z. Li, Y. Fan, H. Meng, R. Liu, Y. Wang, A. Hagfeldt, G. Cui and S. Pang, *Angew. Chem., Int. Ed.*, 2019, **58**, 5587–5591.
- 35 D. Bogachuk, L. Wagner, S. Mastroianni, M. Daub, H. Hillebrecht and A. Hinsch, *J. Mater. Chem. A*, 2020, **8**, 9788–9796.
- 36 X. Huang, R. Chen, G. Deng, F. Han, P. Ruan, F. Cheng, J. Yin, B. Wu and N. Zheng, *J. Am. Chem. Soc.*, 2020, **142**, 6149–6157.
- 37 M. Daub and H. Hillebrecht, *Eur. J. Inorg. Chem.*, 2021, 1490–1497.
- 38 T. Zhao, S. T. Williams, C.-C. Chueh, D. W. deQuilettes, P.-W. Liang, D. S. Ginger and A. K. Y. Jen, *RSC Adv.*, 2016, **6**, 27475–27484.
- 39 M. Long, T. Zhang, H. Zhu, G. Li, F. Wang, W. Guo, Y. Chai, W. Chen, Q. Li, K. S. Wong, J. Xu and K. Yan, *Nano Energy*, 2017, **33**, 485–496.
- 40 X. Y. Feng, K. W. Ng, S. P. Wang, W. Z. Chen, Z. Z. Zhang, W. Chen, Y. Y. Zhao, B. Tu, Z. K. Tang, H. Pan and Z. B. He, *J. Mater. Chem. A*, 2020, **8**, 13585–13593.
- 41 R. A. Kerner, T. H. Schloemer, P. Schulz, J. J. Berry, J. Schwartz, A. Sellinger and B. P. Rand, *J. Mater. Chem. C*, 2019, **7**, 5251–5259.
- 42 C. Wu, H. Li, Y. Yan, B. Chi, J. Pu, J. Li, M. Sanghadasa and S. Priya, *Nano Energy*, 2017, **36**, 295–302.
- 43 R. G. Niemann, A. G. Kontos, D. Palles, E. I. Kamitsos, A. Kaltzoglou, F. Brivio, P. Falaras and P. J. Cameron, *J. Phys. Chem. C*, 2016, **120**, 2509–2519.
- 44 J. K. Meyers, L. Y. Serafin, A. D. Orr and J. F. Cahoon, *Chem. Mater.*, 2021, **33**, 3814–3822.
- 45 S. R. Raga, L. K. Ono and Y. Qi, *J. Mater. Chem. A*, 2016, **4**, 2494–2500.
- 46 R. Eßmann, G. Kreiner, A. Niemann, D. Rechenbach, A. Schmieding, T. Sichla, U. Zachwieja and H. Jacobs, *Z. Anorg. Allg. Chem.*, 1996, **622**, 1161–1166.
- 47 S. Luo and W. A. Daoud, *Materials*, 2016, **9**, 123.
- 48 X. Guo, C. McCleese, C. Kolodziej, A. C. Samia, Y. Zhao and C. Burda, *Dalton Trans.*, 2016, **45**, 3806–3813.
- 49 M. W. Lufaso and P. M. Woodward, *Acta Crystallogr., Sect. B: Struct. Sci.*, 2001, **57**, 725–738.
- 50 Y. Chang, L. Wang, J. Zhang, Z. Zhou, C. Li, B. Chen, L. Etgar, G. Cui and S. Pang, *J. Mater. Chem. A*, 2017, **5**, 4803–4808.
- 51 S. H. Gellman, G. P. Dado, G. B. Liang and B. R. Adams, *J. Am. Chem. Soc.*, 1991, **113**, 1164–1173.
- 52 E. D. Walter, L. Qi, A. Chamas, H. S. Mehta, J. A. Sears, S. L. Scott and D. W. Hoyt, *J. Phys. Chem. C*, 2018, **122**, 8209–8215.
- 53 W.-J. Yin, J.-H. Yang, J. Kang, Y. Yan and S.-H. Wei, *J. Mater. Chem. A*, 2015, **3**, 8926–8942.
- 54 C. Quarti, G. Grancini, E. Mosconi, P. Bruno, J. M. Ball, M. M. Lee, H. J. Snaith, A. Petrozza and F. D. Angelis, *J. Phys. Chem. Lett.*, 2014, **5**, 279–284.
- 55 T. Glaser, C. Muller, M. Sendner, C. Krekeler, O. E. Semonin, T. D. Hull, O. Yaffe, J. S. Owen, W. Kowalsky, A. Pucci and R. Lovrincic, *J. Phys. Chem. Lett.*, 2015, **6**, 2913–2918.
- 56 Y. H. Kye, C. J. Yu, U. G. Jong, Y. Chen and A. Walsh, *J. Phys. Chem. Lett.*, 2018, **9**, 2196–2201.
- 57 Y. Zong, Y. Zhou, M. Ju, H. F. Garces, A. R. Krause, F. Ji, G. Cui, X. C. Zeng, N. P. Padture and S. Pang, *Angew. Chem., Int. Ed.*, 2016, **55**, 14723–14727.
- 58 S. Pang, Y. Zhou, Z. Wang, M. Yang, A. R. Krause, Z. Zhou, K. Zhu, N. P. Padture and G. Cui, *J. Am. Chem. Soc.*, 2016, **138**, 750–753.
- 59 W. Ke, I. Spanopoulos, C. C. Stoumpos and M. G. Kanatzidis, *Nat. Commun.*, 2018, **9**, 4785.
- 60 M. Daub and H. Hillebrecht, *Z. Anorg. Allg. Chem.*, 2018, **644**, 1393–1400.
- 61 J. Wang, X. Yao, W.-J. Xiao, S. Wang, G. Xu, X.-Q. Chen, S.-C. Wu, I. Visoly-Fisher, E. A. Katz, Y. Li, J. Lin, W.-S. Li and Y. Li, *Solar RRL*, 2018, **2**, 1800125.
- 62 F. Wang, W. Li, H. Liu, L. Zhu and H. Chen, *Nanoscale*, 2019, **11**, 14465–14471.
- 63 Y. Zhou, M. Yang, S. Pang, K. Zhu and N. P. Padture, *J. Am. Chem. Soc.*, 2016, **138**, 5535–5538.
- 64 C. Chen, Y. Rao, Z. Li, X. Wang, G. Cui, W. Wang and S. Pang, *Solar RRL*, 2021, **5**, 2000715.
- 65 Y. Zhang, Z. Zhou, F. Ji, Z. Li, G. Cui, P. Gao, E. Oveisi, M. K. Nazeeruddin and S. Pang, *Adv. Mater.*, 2018, **30**, e1707143.

Review

- 66 X. Wang, Y. Fan, L. Wang, C. Chen, Z. Li, R. Liu, H. Meng, Z. Shao, X. Du, H. Zhang, G. Cui and S. Pang, *Chemistry*, 2020, **6**, 1369–1378.
- 67 M. Long, T. Zhang, Y. Chai, C.-F. Ng, T. C. W. Mak, J. Xu and K. Yan, *Nat. Commun.*, 2016, **7**, 1–11.
- 68 T. Zhang, N. Guo, G. Li, X. Qian, L. Li and Y. Zhao, *J. Mater. Chem. A*, 2016, **4**, 3245–3248.
- 69 L. M. Wheeler, D. T. Moore, R. Ihly, N. J. Stanton, E. M. Miller, R. C. Tenent, J. L. Blackburn and N. R. Neale, *Nat. Commun.*, 2017, **8**, 1722.
- 70 V. Goetz and A. Marty, *Chem. Eng. Sci.*, 1992, **47**, 4445–4454.
- 71 L. Hao, Z. Li, L. Wang, R. Liu, Z. Shao, Z. Zhou, X. Guo, G. Cui, S. F. Liu and S. Pang, *J. Mater. Chem. A*, 2020, **8**, 17019–17024.