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Selective and rapid water transportation across a self-assembled peptide-diol channel *via* the formation of a dual water array†

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Achieving superfast water transport by using synthetically designed molecular artifacts, which exclude salts and protons, is a challenging task in separation science today, as it requires the concomitant presence of a proper water-binding site and necessary selectivity filter for transporting water. Here, we demonstrate the water channel behavior of two configurationally different peptide diol isomers that mimic the natural water channel system, *i.e.*, aquaporins. The solid-state morphology studies showed the formation of a self-assembled aggregated structure, and X-ray crystal structure analysis confirmed the formation of a nanotubular assembly that comprises two distinct water channels. The water permeabilities of all six compounds were evaluated and are found to transport water by excluding salts and protons with a water permeability rate of 5.05×10^8 water molecules per s per channel, which is around one order of magnitude less than the water permeability rate of aquaporins. MD simulation studies showed that the system forms a stable water channel inside the bilayer membrane under ambient conditions, with a 2 \times 8 layered assembly, and efficiently transports water molecules by forming two distinct water arrays within the channel.

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Introduction

Even though 71% of the Earth's surface is covered with water,¹ only a tiny percentage is available as freshwater.^{2–4} In consequence, more than one-third of the world's population faces severe freshwater scarcity that limits access to safe water for drinking and maintaining basic hygiene.^{5–8} To solve these issues, scientists have devised various water purification techniques, and in 90% of these cases, the reverse osmosis (RO) method is applied for the desalination of seawater.^{9–11} In this regard, reverse osmosis membranes or, more precisely, artificial water channels (AWCs), which have the ability to transport water selectively by excluding the salts from the seawater, have an essential role.¹² However, one disadvantage of this technique is that the desalination process is highly energy-consuming.^{13,14} Therefore, an extremely sophisticated membrane material is required, which can complete the process by consuming less

energy and making the process economical and available to all.^{15,16}

In nature, aquaporin-based protein channels are available, which are known to transport water across lipid bilayer membranes with high permeability and permselectivity.^{17–20} However, the lack of stability and compatibility of aquaporins outside the cellular environment delimit its widespread use for water purification applications.²¹ Therefore, the development of biomimetic artificial water channels (AWCs) is fundamentally essential for studying their water transport behavior and can be useful for medical and water purification purposes.^{12,22}

Over the past decade, various AWCs based on unimolecular and supramolecular systems have been reported to achieve superfast water transport across the bilayer membranes.^{23–30} Some of them have even shown ultrafast water transport rates. For example, carbon nanotube porins (CNTPs) transport water across the bilayer membrane with a transport rate 6 times³¹ that of AQP1, a natural water channel protein.^{32,33} However, it also transports monovalent cations such as Na^+ , K^+ , and protons.³⁴ Nevertheless, the exclusion of salts and protons by the channel forming substrate is an indispensable criterion for water purification application.³⁵ In this regard, Barboiu, Hou, and their co-workers made tremendous efforts to develop synthetic channels for selective water transport, and later, the studies were further extended by Kumar, Zeng, and their co-workers. The reports from Barboiu and co-workers demonstrated the I-quartet and the hydroxyl-based water channels for selective and fast water

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permeation across the bilayer membrane.^{36–38} Contemporarily, Hou and co-workers reported the hydrazide-appended pillar[5] arene derivatives which showed proton excluding water transport behavior.²⁵ Later, Kumar and co-workers reported peptide-appended hybrid[4]arene (PAH[4]) based water transport systems that transport water (water permeability $> 10^9$ H₂O per s per channel) with a comparable transport rate of AQPs while excluding salts.³⁹ The studies were further extended in recent reports by Zeng and co-workers, where they have reported several foldamer-based water channels that have shown excellent water permeability rates, which exclude salts and protons during the water transport process.^{40–42}

As mentioned earlier, most artificial water channels have shown water permeability slightly less or close to that of aquaporins, but very few have demonstrated good water-salt permselectivity.⁴³ The lack of a water selectivity mechanism (such as that of aquaporins) in most of the synthetic channels makes them less selective for water over the ions.³⁷ Therefore, the proper structural design of the channel, which can interact with the water molecules selectively and softly, can show excellent water transport behavior with high salt rejection. Moreover, the absence of intermolecular hydrogen bonding interactions between two water molecules in the same water array can restrict the Grotthuss mechanism for proton transport, which is essential for showing good water channel behavior.^{44–46} However, obtaining permselectivity while achieving superfast permeability is still an exciting and challenging task in water purification science.

Herein, we report the peptide backbone-based 1,2-diol systems **1S–4S** and **1R–4R** (Fig. 1A) that form a dimeric rosette assembly and allow water transport across bilayer membranes by forming two distinct water channels. The diol group is used in our previous report to develop an anion-selective self-assembled ion channel,⁴⁷ where the functional group acts as a hydrogen bond donor to interact with the anion. In contrast, it is also found in the literature that the hydroxyl groups (from the diol unit) have the ability to form the self-assembled water channel while participating either as a hydrogen bond donor or acceptor site (*via* either O–H···OH₂ or O···H–OH interactions).³⁷

In their proposed design, eight aggregated stacking units come closer to form a cylindrical channel through which the water transport phenomenon was observed. Here, in our design, we attached one alkyl group connected peptide unit with the diol group, which is expected to provide an additional hydrogen bond acceptor site to interact with the water molecules through C=O···H–OH interaction. It is anticipated that this additional hydrogen bonding interaction site will support the formation of a water channel structure utilizing a minimum number of aggregated stacking units compared to the previous report. Overall, the peptide-diol molecules were designed to have multiple hydrogen bond donor and acceptor sites such as O–H, N–H, C=O, *etc.*, for efficient noncovalent interactions. We expected that these systems would self-aggregate into a stacked sheet by a parallel one-dimensional array of monomers utilizing the intermolecular C=O···H–N bonding interactions. Two such one-dimensional sheets would bundle up *via* intermolecular hydrogen bonding interaction through hydroxyl groups (from the diol subunit) and carbonyl groups (from the peptide subunit) with the water molecules to form a water channel (Fig. 1B). The chirality of the phenylalanine α -carbon was varied for generating two different series. The octyl and alkyl chains were introduced onto the monomers to tune the membrane permeability,⁴³ which was estimated by calculating the log *P* values of each compound using the Marvin Sketch program (Fig. 1A).⁴⁸ Moreover, we expected that the dimeric rosette assembly would offer adequate channel pore and strong noncovalent interactions for water recognition by excluding the salts and hence increasing the water–salt permselectivity.

Results and discussion

Synthesis of peptide 1,2-diols

The synthesis of the channel forming molecules was started from L-phenylalanine and D-mannitol. At first, the L-phenylalanine was converted to its amide analogues **8a–8d** by reacting with different aliphatic carboxylic acids following a reported literature protocol (Scheme S1†).⁴⁹ On the other hand, D-mannitol was converted to acetonide protected dibenzyl-amine derivative **12** in three steps following our previously reported protocol (Scheme S2†).^{50,51} The hydrogenation of dibenzyl derivative **12** in the presence of Pd–C led to the debenzylation and reduction of the triple bond for the formation of acetonide-protected amine **13** in quantitative yield. Then, amine **13** reacted with different L-phenylalanine derivatives **8a–8d** *via* an EDC·HCl assisted coupling reaction to form amide derivatives

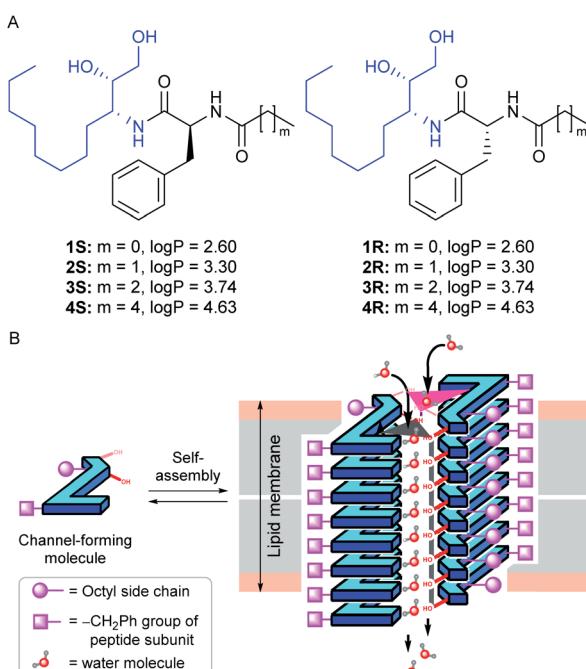
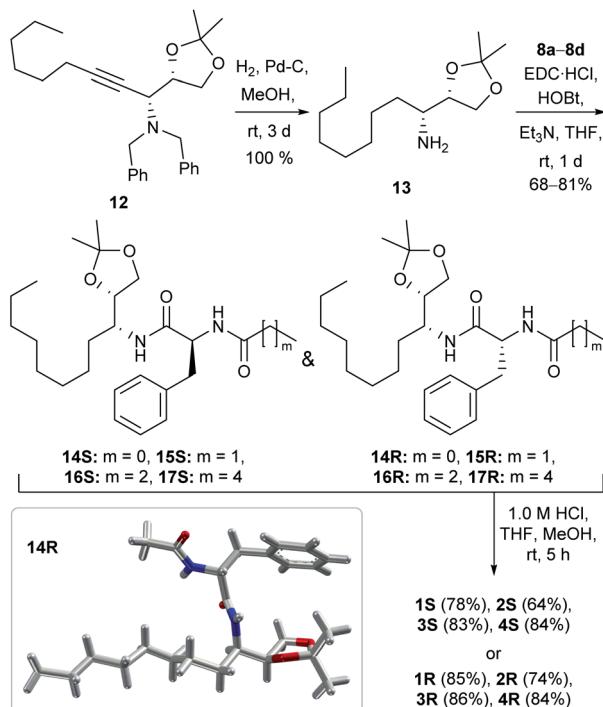


Fig. 1 Chemical structures and the calculated logP values of peptide 1,2-diols **1S–4S** and **1R–4R** (A). Schematic representation of the self-assembled water channel formation by transporter molecules in the lipid bilayer membrane (B).





Scheme 1 The synthetic scheme of peptide 1,2-diols **1S–4S** and **1R–4R**. The X-ray crystal structure of **14R**, obtained from a nitrobenzene solution of the protected compound (bottom left position, inside the box).

14S–17S. Surprisingly, the diastereomeric compounds **14R–17R** were also formed almost in equal amounts due to the inversion of configuration at phenylalanine α -carbon. *S*- and *R*-configuration products were isolated separately by column chromatography. The formation of the *R*-configuration product was confirmed from the crystal structure of **14R** (Scheme 1). Subsequently, all the acetonide-protected amides were separately treated with dilute HCl for the deprotection of the acetonide group to furnish the desired diol compounds **1S–4S** and **1R–4R** in 64–86% yield (Scheme 1).

Solid-state morphology analysis

After synthesizing all the desired derivatives, we performed the experiments for morphological studies. At first, field emission scanning electron microscopy (FESEM) images for **4S** and **4R** were recorded (while preparing the samples in chloroform, Fig. S1A and S2A \dagger), where no specific self-aggregated pattern was observed. However, when we performed the same experiments for these compounds by preparing the sample in moist methanol (2% water in methanol solvent), self-aggregated left-handed helical fibril morphology was observed for each compound (Fig. 2A and S1B \dagger for **4S** and Fig. S2B \dagger for **4R**). These observations indicate that the designed systems can form a self-aggregated assembly, where a polar protic solvent assists the assembled states. Such a fibril formation can be rationalized by considering the intermolecular hydrogen bonding interactions between the peptide backbones to form sheet structures followed by the self-aggregation *via* water molecules bridging to

create a continuous chain, where multiple chains of this type bundled up for the formation of a left-handed helical fibril (Fig. 2A and C). 52,53 The fibril formation was further confirmed by transmission electron microscopy (TEM) studies for **4S** and **4R**, where the samples were prepared in moist methanol (2% water in methanol solvent, Fig. 2B and S3 \dagger).

Solid-state crystal structure analysis

A good quality crystal of **2R**, suitable for solid-state single crystal X-ray analysis, was obtained by dissolving the compound in moist methanol (2% water in methanol solvent) followed by slow evaporation. In the solid-state, multiple monomer units participate in the parallel self-aggregation process *via* intermolecular amide C=O \cdots H-N hydrogen bonding interactions to form one-dimensional stacked sheets. Further, two such one-dimensional sheets bundle up antiparallelly through water-bridged hydrogen bonds. In this assembly, each water molecule forms one C=O \cdots H-OH hydrogen bond with the phenylalanine carbonyl of one monomer and a second C=O \cdots H-OH hydrogen bond with the terminal ethylamide carbonyl of the next monomer of the same sheet (Fig. 2D). Moreover, it also forms two C–O \cdots H-OH hydrogen bonds, one with the secondary hydroxyl carbon of the first monomer and the other with a monomer from the second sheet C–O \cdots H-OH hydrogen bond with the primary hydroxyl carbon of a monomer from the second sheet. Interestingly, the presence of two distinct water channels was evident within the cavity created by the two sheets. Moreover, the water molecules of each channel (*i.e.*, in the same water array) are well separated (4.93 Å) to restrict the proton transport as per our proposed design. Finally, two anti-parallel aggregated sheets are ensembled in a way that allows all the polar groups to either face inwardly or interact with each other to favour the assembly, whereas the hydrophobic phenyl and alkyl units are outwardly directed, which can help it to be inserted inside the lipid bilayer easily.

IR studies of channel-forming molecules in the lipid membrane

To check the channel formation behaviour of the channel-forming molecules, IR spectra were recorded after incorporating the compound in the lipid membrane. 54,55 The change in the IR spectrum of the vesicles in the presence of the channel-forming molecule (5 mol% with respect to the lipid) compared to the free vesicles indicated the insertion of the channel-forming compound inside the bilayer membrane (Fig. S6 \dagger). Additionally, the broadening of the diol O–H peak (3268 cm $^{-1}$) and shifting of the amide C=O peak (1628 cm $^{-1}$ to 1640 cm $^{-1}$) of the compound in the presence of the lipid indicated the formation of a channel inside the bilayer membrane.

Water permeability studies across the lipid bilayer membrane

The surprising results from the solvent dependent morphology studies and water channel formation by **2R** in the solid-state encouraged us to assess the water transport permeability and selectivity by the transporter molecules across the phospholipid bilayer membrane. For this, a stopped-flow light scattering



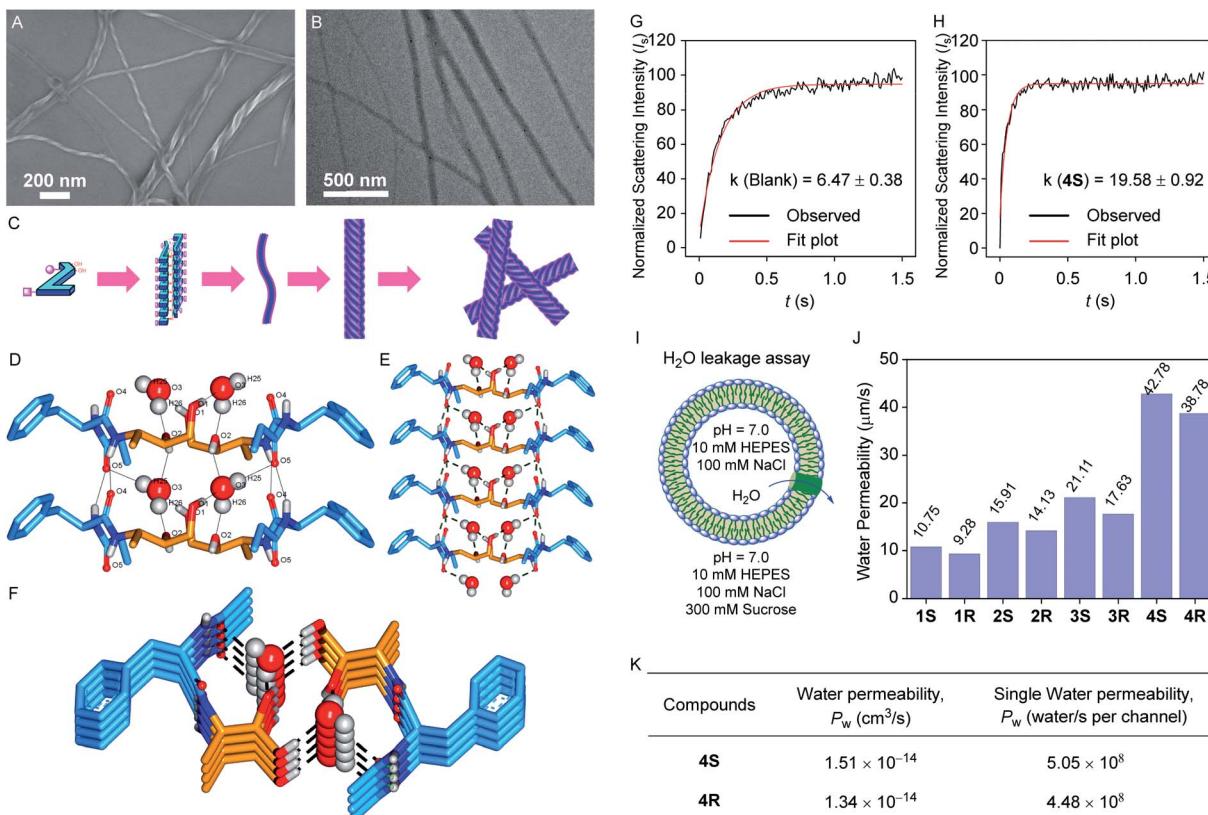


Fig. 2 The FESEM (A) and TEM (B) images of the self-aggregated **4S** compound (prepared from 100 μM solution in moist methanol) with the cartoon representation of the probable self-aggregating pattern for the obtained morphology (C). The crystal structure of **2R** with water molecules forming the self-assembled water channel: 2 \times 2 layered top view H-bonding interaction distances (D), 2 \times 4 layered lateral view (E), and 2 \times 4 layered top view (F). The dashed lines in the crystal structure are the representation of the hydrogen bonding interactions between the channel-forming molecules and water molecules. The alkyl chains are omitted for clarity. The light scattering intensity vs. time plot of blank DOPC vesicles (G) and DOPC vesicles with **4S** (H; lipid to transporter ratio = 50 : 1) recorded on a stopped-flow instrument for measuring water permeability. The representation of vesicles for water leakage assay (I), the comparison of corrected water permeability (corrected from blank) of all transporters (J), and single channel permeability of **4S** and **4R** (K) at a 50 : 1 lipid to transporter ratio.

instrument was used for measuring the light scattering intensity of the vesicle suspension at a millisecond scale. At first, large unilamellar vesicles (LUVs) with a diameter of \sim 150 nm were prepared using the DOPC lipid by incorporating each channel forming compound (molar channel-to-lipid ratio, mCLRs = 1 : 50) in 100 mM NaCl solution buffered to 7.0 using HEPES (Fig. 2I).^{40,41} When such a vesicle suspension is exposed suddenly to an equal volume of 300 mM sucrose solution, it leads to the shrinkage of the vesicles by outwardly directed water leakage through the water channels. This event subsequently leads to the change in light scattering intensity of the vesicle suspension. For each compound, a significant change in light scattering intensity compared to the blank sample was observed, and the rate constant value was also calculated (Fig. 2G, H and S7A–S7K†). Further, the rate constant values were used to calculate the water permeability values.^{39,56}

All our molecules **1S–4S** and **1R–4R** were found to transport the water across the bilayer membrane (Fig. 2H, J, and K), and the permeability values calculated using the conventional model are 85.7 (\pm 2.8%), 100.3 (\pm 3.5%), 115.1 (\pm 7.3%), 176.5 (\pm 4.7%), 81.5 (\pm 0.6%), 95.2 (\pm 5.8%), 105.2 (\pm 5.9%), and 164.8

(\pm 6.9%) $\mu\text{m s}^{-1}$, respectively. We have also recalculated the permeability values for **1S–4S** and **1R–4R** using a recently updated model,^{39,56,57} considering the osmolyte concentration correction factor, and the values are 30.2 (\pm 2.8%), 35.4 (\pm 3.5%), 40.6 (\pm 7.3%), 62.3 (\pm 4.7%), 28.8 (\pm 0.6%), 33.6 (\pm 5.8%), 37.1 (\pm 5.9%), and 58.17 (\pm 6.9%) $\mu\text{m s}^{-1}$, respectively. We assume that the 8 layers of the self-assembled system (2 \times 8 layers) form a single channel structure and further considering the channel insertion efficiency and the corrected mCLRs values (Fig. S8 and S9†), the water permeability (unit = $\text{cm}^3 \text{s}^{-1}$) and single channel water permeability (unit = water molecules per s per channel) values were calculated for **4S** and **4R** channels. The **4S** channel exhibited the highest water transport rate with a permeability value of $1.51 \times 10^{-14} \text{ cm}^3 \text{s}^{-1}$ compared to the **4R** channel, for which the estimated permeability value is $1.34 \times 10^{-14} \text{ cm}^3 \text{s}^{-1}$. Further, these water permeability values were used to calculate the single channel water permeability, and the values are 5.05×10^8 and 4.48×10^8 water molecules per s per channel for **4S** and **4R**, respectively, indicating **4S** as the most efficient water channel forming molecule (Fig. 2J, S7 and Table S2†). The

formation of two distinct water channels within a single aggregate could be a prominent reason for the excellent single channel water permeability of **4S** and **4R**. Moreover, the data indicate that the efficiency of the water channels increases when a longer hydrophobic alkyl tail is present, indicating that the higher $\log P$ value contributes to better membrane permeability of the channel. It is also apparent that the activity of an *S*-configuration product is higher than that of the corresponding *R*-configuration product. This may be attributed to the better self-assembly of the *S*-configuration molecules.

Water permselectivity studies across the lipid bilayer membrane

As discussed earlier, the salt exclusion is the essential criteria for water purification application of the water channel system. The salt exclusion behavior of the channel molecules while transporting the water was confirmed by performing the water leakage assay in different osmolytes other than sucrose.^{40,58} Here, we have calculated the reflection coefficient value for our system, where the reflection coefficient is the water permeability ratio of the channel in NaCl and sucrose solutions. We observed the reflection coefficient value 0.98 (Fig. S10†), which is close to 1, indicating that the water permeability value remains unaltered while shifting from sucrose solution to NaCl solution. This observation also confirmed the salt exclusion capability of our water channel system.

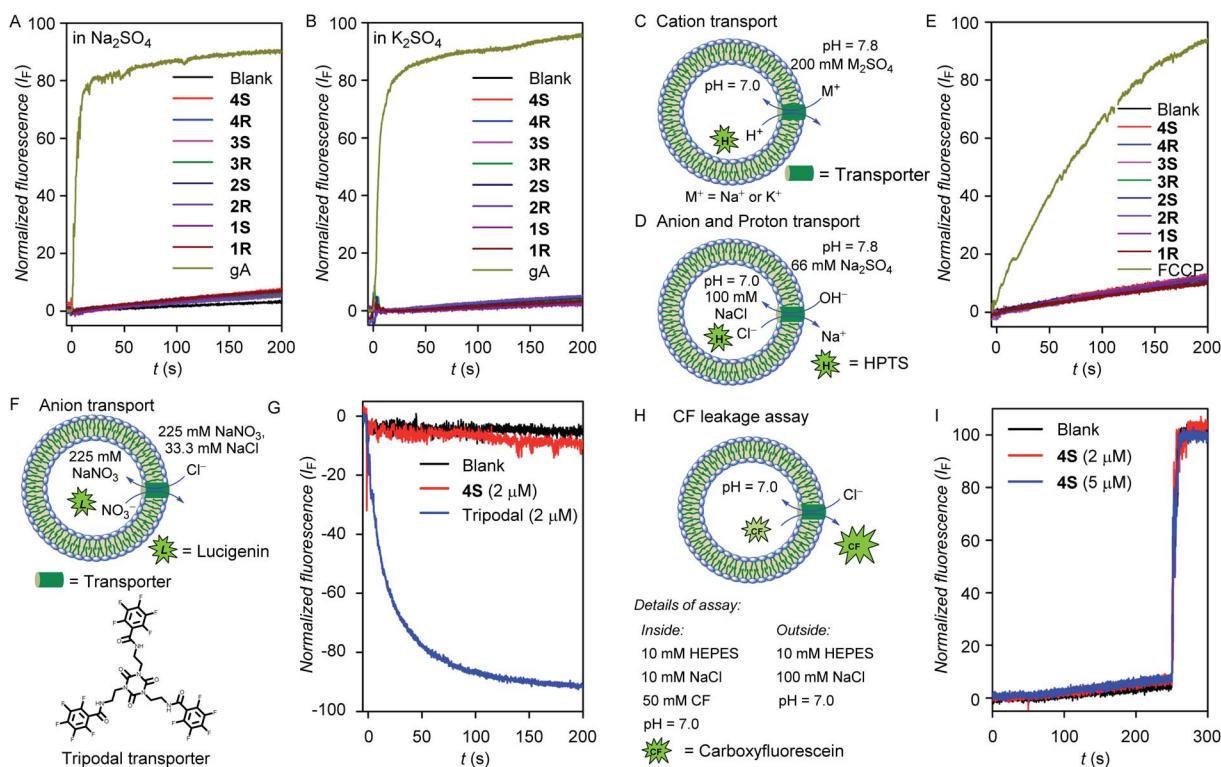


Fig. 3 The Na^+ (A) and K^+ (B) transport activities of gramicidin (0.1 μM) and channel molecules (2 μM) in HPTS assay across EYPC-LUVs \supset HPTS. The vesicular representations and assay details of cation transport (C) and anion transport (D) experiments. The H^+ transport activities of FCCP (0.2 μM) and water channel-forming molecules (2 μM) in HPTS assay across the EYPC-LUVs \supset HPTS (E). The vesicular representation (F) and anion transport activities in lucigenin assay across the EYPC-LUVs \supset Lucigenin with a tripodal transporter (2 μM) and water channel molecules (2 μM) (G). The assay details (H) and normalized fluorescence intensity plot using **4S** (2 and 5 μM concentrations) for CF leakage assay (I).



these compounds were observed, whereas, carbonyl cyanide-4-(trifluoromethoxy)phenylhydrazone (FCCP, a proton carrier) showed a significant transport activity (~92%) even at the one-tenth concentration (Fig. 3E). These observations indicate that the 1,2-diol compounds are incapable of transporting anions and protons. Additionally, the anion influx by **4S** was also evaluated across large unilamellar EYPC vesicles by entrapping lucigenin dye.^{47,62} Under intravesicular NaNO_3 (225 mM) and extravesicular NaNO_3 (225 mM) and NaCl (33.3 mM) conditions, no destruction of the chloride gradient in the presence of **4S** was observed (Fig. 3F and G). In contrast, a reported tripodal chloride transporter⁶⁰ showed excellent chloride influx activity (~88%) under comparable conditions. Therefore, the inability of **4S** for transporting anions was further established. In the above-mentioned anion, cation, and proton transport experiments, the intra- and extra-vesicular osmolyte concentration balance has not been maintained. To check for any ion transport in the iso-osmolar salt concentration, the ion transport activity was evaluated using standard HPTS assay, where 100 mM NaCl solution was used as intra- and extra-vesicular salt. In this assay, no transport activity was observed for the **4S** channel (Fig. S14†), which ruled out the possibility of ion transport by the channel-forming molecule.

The carboxyfluorescein (CF) leakage assay,⁶³ using 2 and 5 μM concentrations of **4S** resulted in no release of CF dye (Fig. 3H and I), which indicates that the membrane integrity is preserved in the presence of the compounds and also rules out any possibility of pore formation.

Molecular dynamics simulation studies of the water channel in the lipid bilayer model

The excellent water channel properties of the 1,2-diol system encouraged us to carry out computational studies to get atomistic insight into the channel architecture and water transport mechanism. The comparable water transport properties between **4S** and **4R** and, most importantly, the availability of crystallographic coordinates of the *R*-configuration product with water, prompted us to focus on simulating the **4R** superstructure and water permeation through it. The self-assembly pattern obtained from the crystal structure of **2R** with water was used to generate the 2×8 layered (*i.e.*, 8 layers of 2 molecules) assemblies of **4R**. The assembly of **4R** was generated by elongating each ethyl chain with a pentyl one.

Fig. 4A and B, respectively, demonstrate the representative snapshot of the membrane-embedded 2×8 layered assembly of **4R** at the initial point and at completion of the computer

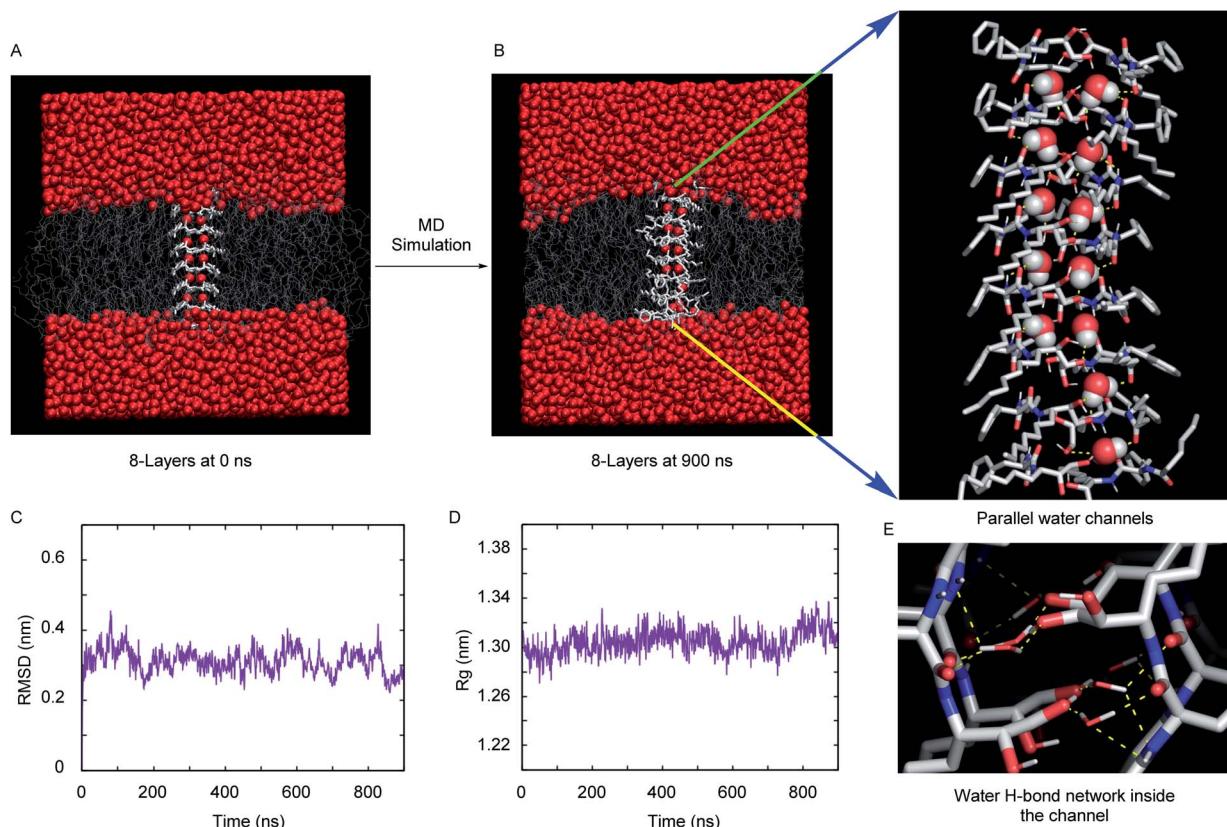


Fig. 4 Snapshot from MD simulation of the 2×8 layered **4R** synthetic channel at the initial point *i.e.*, 0 ns (A) and at the completion of the simulation, *i.e.* 900 ns (B). Also shown is an enlarged view image of the simulated synthetic channel. The presence of two distinct water files inside the layer is evident. The RMSD (C) and radius of gyration (D) of the 2×8 layered **4R** channel. The snapshot highlighting the water hydrogen bond network inside the channel (E). Color code: membrane hydrophobic cores are shown in thin white lines, and the synthetic channel are represented via white licorice representation. In (B), the water oxygen and hydrogen atoms are shown by red and white space-filling representation, respectively. In (E), the same color code is maintained for water, except a licorice representation is rendered.



simulation after about microsecond (900 ns) long time-evolution. The visual inspection of the snapshot indicates that the simulations maintain a stable superstructure during the time period at ambient pressure (1 bar) and temperature (298 K). Unlike previous investigations, which required introduction of an order of magnitude larger lateral pressure or restraining forces for maintaining a stable synthetic channel,^{36,37} the current simulation was found to produce a structurally integrated channel inside the phospholipid bilayer under the ambient conditions of 1 bar pressure. This is further demonstrated by a steady time profile of root-mean-squared deviation (RMSD) of the 2 × 8 layered assembly relative to the crystal structure (Fig. 4C). A low RMSD of 0.3 nm from the crystal structure and a steady trend radius of gyration (R_g) profile (Fig. 4D) with an average R_g fluctuating around 1.3 nm further provide credence to the stability of the 2 × 8 layered superstructure.

We had also explored the effect of the number of layers on the overall integrity of the superstructure. Fig. S15A and S15B† compare the representative snapshot of superstructures across 2 × 7 layers and 2 × 9 layers, respectively. As depicted from the steady RMSD and R_g profiles (Fig. S15C and S15D†), the overall assembly remains stable in these two cases as well. However, when compared across all three cases (2 × 7 layers, 2 × 8 layers and 2 × 9 layers), the registry and overall orientation of the 2 × 8 layered assembly inside the membrane are found to be relatively optimal. Hence, we proceeded with simulation trajectories of 2 × 8 layers for subsequent analysis of water permeation dynamics inside the channel.

A close inspection of the synthetic channel (see the enlarged snapshot of Fig. 4B) highlighted the distinct presence of two

parallel, non-intersecting water wires inside the synthetic channel. A molecular-level characterization of the inner architecture of water inside the simulated channel indicated that the water molecule forms an intricate hydrogen bond network (Fig. 4E) with the phenylalanine carbonyl oxygen, the diol (or hydroxyl) moiety and also with the amide nitrogen of the channel-forming molecules. The organization of inner water molecules is mainly guided by the integrity of the superstructure. However, there is no hydrogen bond between the water molecules present inside the channel. The visual inspection of the MD simulation trajectory indicated significant water leakage inside the phospholipid bilayer, as induced by the synthetic channel. A quantitative analysis of water leakage events (see method) provided further insights into the dynamical behavior of water molecules in a lipid bilayer environment. Fig. 5A shows the time profile of the number of water molecules leaked inside the membrane, suggesting significant water transport *via* the channel at ambient pressure and temperature. Very interestingly, tracking the movement of leaked water molecules revealed multiple water permeation events during the period of simulation. Fig. 5B and C depict two representative water-permeation events in which the water molecules make a full pass across the membrane from one of the leaflets to the other leaflet. The characterization of the simulation trajectory indicated a regular occurrence of water permeation events during the simulation period (Fig. 5D). However, the durations of such water permeation events are not uniform. In fact, as shown in Fig. 5E, the duration of water permeation is quite heterogeneous, ranging from as low as 2 ns to as large as 14 ns. Overall, the extent of water leakage and water permeation *via* maintaining two distinct files of water wires inside a single

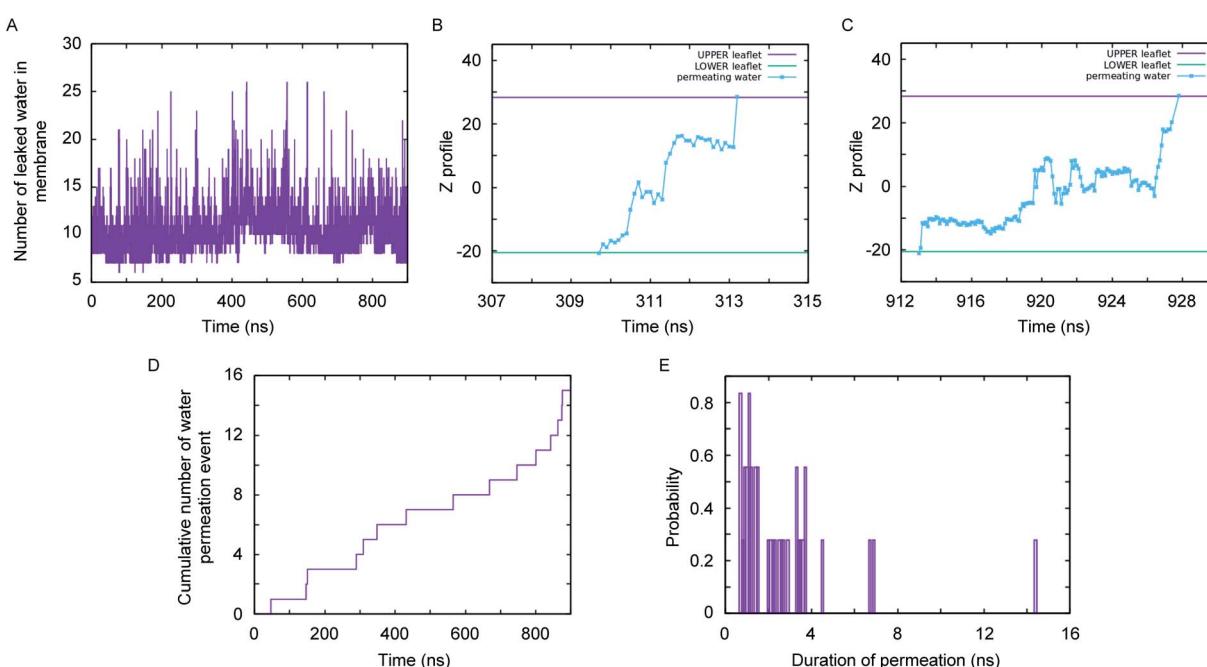


Fig. 5 Water permeation analysis across the 2 × 8 layered 4R superstructure: the time profile of the number of leaked water molecules inside the membrane (A). The time profiles of two representative water permeation events through channels (B and C), the cumulative time profiles of water permeation events (D), and the distribution of time-span for water permeation events (E).



superstructure dictates a highly efficient synthetic water channel. The observation of two coexisting parallel water wires inside the channel formed by **4R** is considerably distinct, when compared with previous investigations, which have reported either a single water wire or more disordered water structure inside the channel.

The quantitative value of water/salt selectivity was estimated by MD simulation studies⁴⁰ in the presence of NaCl salt, where 7247 water molecules in total were present on both sides of the lipid bilayer, and the salt concentration was maintained at 0.10 M by adding 12 atoms of Na⁺ and 12 atoms of Cl⁻ ions. In the transport simulation, a total of 52 223 223 water entry events up to the 1/8th (the first layer; till 4.79 Å) of the water channel was recorded on a time scale of 900 ns. Out of this, 16 molecules of water were found to be successfully transported across the bilayer membrane within this time scale. Additionally, we have not observed any transport of either of the Na⁺ and Cl⁻ ions across the bilayer membrane. The transport selectivity (water/Na⁺ or water/Cl⁻) was calculated from the data for Na⁺ and Cl⁻ ions, and is 5 223 223/(7247/12), *i.e.* 8649.

The **4R** channel has shown efficient experimental water permeability. Therefore, we have chosen this derivative for MD simulation studies as mentioned above, where the initial coordinates were generated from the solid-state crystal structure of the **2R** channel by elongating the alkyl chain length. However, similar MD simulation studies (Fig. S16A and S16B†) were also performed for the **2R** channel as well, where the initial channel structure was generated directly from the crystal structure. Here, we observed the maintenance of a stable channel structure similar to the **4R** channel, with a slight deviation of RMSD (Fig. S16C†) and R_g values (Fig. S16D†) without applying any lateral pressures and restrain. Moreover, the salt-exclusion water permeation events were observed in these experiments (Fig. S16E and S16F†), and it was seen that for a particular configuration product, the elongation of alkyl chain length (facing outside) does not make any difference in the channel-forming mode.

In the quest for the channel structure of **4S**, the 2 × 8 layered **4S** channel assembly was generated by inverting the stereochemistry of the chiral phenylalanine centre and elongating the peptide attached methyl group to pentyl on each molecule of the 2 × 8 layered **2R** channel, and subsequently, MD simulation studies were performed. However, stable channel formation was not observed in the case of **4S** after applying similar simulation conditions (Fig. S17†). Perhaps some other type of channel assembly is formed in the case of the **4S** channel for the transport of water across the bilayer membrane. However, as the single-channel permeabilities of **4R** (4.48×10^8 water molecules per s) and **4S** (5.05×10^8 water molecules per s) do not differ significantly, the channel structures of **2R** and **4R** are sufficient to represent the peptide 1,2-diol-based water channel systems.

Conclusions

In conclusion, we have designed and synthesized peptide diol-based water channel systems, which form two distinct water

channels in a single self-assembled structure by employing the intermolecular hydrogen bonding interactions of water molecules with the hydroxyl groups and amide side chains. The experimental analysis revealed that the molecules are able to transport water across the lipid bilayer membrane with the highest permeability rate of 5.05×10^8 water molecules per s per channel, observed for the **4S** derivative. Most importantly, a precise single-channel diameter of 2.5 Å was expected to allow only water molecules to pass selectively through the pore by excluding Na⁺, Cl⁻, and H⁺ ions, and this was evidenced from experimental and computational studies. The MD simulation studies for **4R** and **2R** superstructures confirmed the formation of a stable water channel assembly inside the bilayer membrane at ambient temperature and pressure, and more importantly, they show the transport of water molecules through its channel pore by forming two distinct water arrays. The superfast water transport rate and salt exclusion behaviour of the self-assembled channels can be applied in various applications, including the development of biomimetic RO membranes.

Data availability

The detailed experimental procedures and characterization data are provided in the ESI.† The solid state single crystal X-ray analysis data have been deposited to Cambridge Crystallographic Data Center (CCDC) with the deposition codes: 2122390 (for **2R**) and 2122389 (for **14R**).

Author contributions

P. T. conceived the project and directed the experimental studies. J. M. directed the computational studies. D. M. synthesized and characterized the compounds. B. R. D. performed the computational studies. D. M., M. A., and A. M. performed the transport studies. D. M., B. R. D., J. M., and P. T. wrote the paper. All authors approved the final version of the manuscript.

Conflicts of interest

There are no conflicts to declare.

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