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Heterogeneous electron transfer reorganization energy at the inner Helmholtz plane in a polybromide redox-active ionic liquid[†]

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In ionic liquids (ILs), the electric double layer (EDL) is where heterogeneous electron transfer (ET) occurs. Nevertheless, the relationship between the EDL structure and its kinetics has been rarely studied, especially for ET taking place in the inner Helmholtz plane (IHP). This is largely because of the lack of an appropriate model system for experiments. In this work, we determined the reorganization energy (λ) of Br_2 reduction in a redox-active IL 1-ethyl-1-methylpyrrolidinium polybromide (MEPBr_{2n+1}) based on the Marcus–Hush–Chidsey model. Exceptionally fast mass transport of Br_2 in MEPBr_{2n+1} allows voltammograms to be obtained in which the current plateau is regulated by electron-transfer kinetics. This enables investigation of the microscopic environment in the IHP of the IL affecting electrocatalytic reactions through reorganization energy. As a demonstration, TiO_2 -modified Pt was employed to show pH-dependent reorganization energy, which suggests the switch of major ions at the IHP as a function of surface charges of electrodes.

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Introduction

The electric double layer (EDL) structure is critical to understand and predict the activities of electrochemical reactions.^{1–3} For example, the enhanced HER/HOR activities using a $\text{Ni}(\text{OH})_2$ -modified Pt(111) electrode were ascribed to the shift in the potential of zero charge (PZC) of Pt *via* structural changes in the EDL.^{1,2} Applying a strong electric field (10^8 to 10^{10} V m^{-1}) gives rise to the rearrangement and reorientation of electrolytes in the vicinity of the electrode surface, creating a unique medium for heterogeneous electron transfer. The dielectric constant (ϵ) of solvents closer to the electrode is smaller than that of bulk,^{4–6} making reorganization energy decrease.⁷ This phenomenon highlights in the inner Helmholtz plane (IHP) where the electric field is the strongest across the EDLs.

The impact of the interfacial electric field on electrochemical activities is more prominent in room-temperature ionic liquids (RT-ILs). RT-ILs have attracted considerable interest in the fields of electrocatalysis,^{8,9} batteries¹⁰ and supercapacitors¹¹ because of their high ionic conductivity, wide electrochemical window, and high thermal stability.¹² The extremely high ionic

strength of RT-ILs leads to a unique EDL structure that is different from those of conventional electrolytes in solution. The EDL of a RT-IL is compact and consists of densely packed ions.¹³ An ion in the EDL of a RT-IL strongly interacts with the surrounding ions, and ionic composition at the surface is likely to respond to the electrode charge.¹⁴ This is where heterogeneous electron transfer takes place so that quantitative analysis should be essential to understand the origins of many phenomena in RT-ILs, including PZC-sensitive electrochemical reactions. Nevertheless, few studies have addressed the relationship between the EDL structure and electrochemical activities in RT-ILs.^{15,16} This is primarily because it is hard to probe the properties of the RT-IL near the electrode, especially in the IHP.

The Marcus–Hush–Chidsey (MHC) model for heterogeneous electron transfer kinetics has been successfully applied to associate electron transfer kinetics with the microscopic properties of electrode–electrolyte interfaces.^{3,7,17–20} According to the Marcus theory, the reorganization energy (λ) reflects the properties of a medium and can be utilized to evaluate the local ϵ value in an EDL.⁶ λ can be calculated by fitting the electron transfer rate *vs.* electrode potential curve to the corresponding equation based on the MHC model. Yet experimental quantitation of λ for electrochemical systems is challenging. Because mass transport (MT) is slower than electron transfer (ET) in most electrochemical systems, one can hardly assume that a given voltammogram reflects the ET kinetics predominantly. Most of the previous studies to measure λ were performed for redox species immobilized on electrodes where the number of redox-active species was fixed.^{17,18} Once the molecules are

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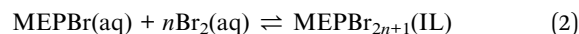
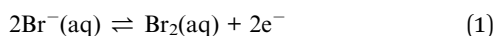
anchored on the electrode, the EDL structure should get perturbed. Moreover, only λ of the outer Helmholtz plane (OHP) could be obtained for the redox species that reside at a large distance from the electrode beyond the IHP. Considering that majority of the important catalytic reactions take place in the IHP accompanying adsorptive processes, it is crucial to evaluate λ of the IHP and look into the microenvironment involved. This requires a proper model system that allows experimental measurements varying with reaction conditions. This is more significant in RT-ILs because the high viscosity slows down the diffusion of the redox-active species, resulting in obscure interpretation of the kinetics based on currents.^{21,22}

In this study, we suggest a Br_2 reduction in 1-ethyl-1-methylpyrrolidinium polybromide (MEPBr_{2n+1}), which is a $\text{Br}^-/\text{Br}_{2n+1}^-$ -based RT-IL, as an appropriate system to investigate the effects of the ion environment and PZC involved in heterogeneous ET kinetics. The mass transport of Br^-/Br_2 in MEPBr_{2n+1} is even faster than proton hopping,²³ and this substantially expands the ET kinetics-governed potential window. In the MEPBr_{2n+1} system, the voltammograms are well fitted to the MHC model for heterogeneous ET kinetics, including clear ET-limited steady-state current. As a demonstration, the λ values of Pt and TiO_2 -modified Pt electrodes were compared to determine the influence of electrode surface charges.

Results and discussion

Ultrafast mass transport system of MEPBr_{2n+1}

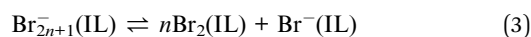
A MEPBr_{2n+1} droplet was electrochemically synthesized in a 250 mM solution of MEPBr in 1 M aqueous potassium phosphate buffer at a Pt macroelectrode (1.2 V vs. Ag/AgBr (3 M KBr)).²³ Under these conditions, Br_2 is generated at the Pt electrode, following which it is captured by MEPBr in the solution to produce MEPBr_{2n+1} , which is immiscible with water to form a droplet:²⁴



The MEPBr_{2n+1} droplet was immobilized on the Pt surface. The potential was held constant at the Pt electrode for several hours until the droplet became large enough to be observed by the naked eye.

MEPBr_{2n+1} is composed of polybromides, Br_{2n+1}^- , *i.e.* Br_3^- , Br_5^- , and Br_7^- , that can be confirmed by Raman spectroscopy.²³ In MEPBr_{2n+1} , the redox centers are highly concentrated, and MEP^+ and Br_{2n+1}^- can act as both the electrolyte and solvent. MEPBr_{2n+1} stores a very high concentration (7.5 M) of Br^- .²⁵ The apparent diffusion coefficients of redox species in MEPBr_{2n+1} measured by electrochemical impedance spectroscopy (EIS) are surprisingly high, 6×10^{-4} to $3 \times 10^{-3} \text{ cm}^2 \text{ s}^{-1}$.²³ This is two to three orders of magnitude higher than those of the species in aqueous solutions.

Br_{2n+1}^- has weak coordinate bonds between Br^- and Br_2 , which is called halogen bonding (dashed line in Fig. 1a), and Br^- , Br_2 , and Br_{2n+1}^- are at equilibrium in polybromide ILs:²⁶



The facile dissociation and reconstitution of halogen bonding facilitate MT of Br_2 , Br^- , and Br_{2n+1}^- . Fig. 1b–d show the mechanism proposed for the fast MT of Br^- and Br_2 in MEPBr_{2n+1} when Br_2 is electrochemically reduced. The equilibrium among Br^- , Br_2 , and Br_{2n+1}^- buffers the Br_2 concentration gradient near the electrode (Fig. 1b). Br_2 elimination from Br_{2n+1}^- , with a higher n , needs lower energy.²⁷ Since MEPBr_{2n+1} contains high concentrations of Br_{2n+1}^- and Br^- , it can minimize the depletion of redox species at the electrode surface caused by the faradaic reaction. In addition, Br_2 is deemed to transport through the Grotthuss-like hopping mechanism *via* the bromine network within the polybromide IL (Fig. 1c).^{28–31} Owing to the hopping transport, the species appears to move a long distance, although the actual displacement is small. The apparent diffusion coefficient of the redox

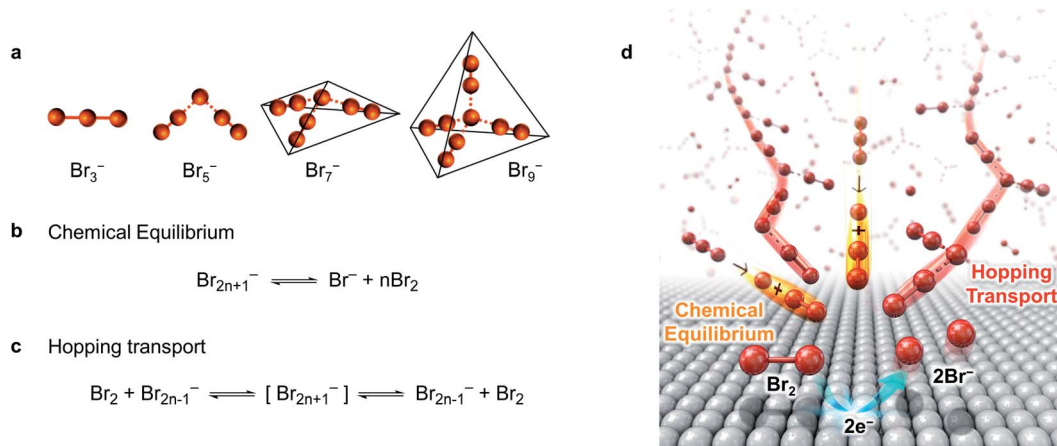


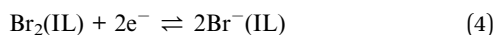
Fig. 1 Ultrafast mass transport system of Br_2 within MEPBr_{2n+1} . (a) Molecular structure of Br_{2n+1}^- . Dashed lines represent halogen-bonding. (b) Chemical equilibrium between Br_{2n+1}^- , Br^- and Br_2 . (c) Grotthuss-like mechanism of Br_2 hopping transport. (d) Schematic representation for the ultrafast MT system of Br_2 within MEPBr_{2n+1} .



species in MEPBr_{2n+1} is actually higher than that of proton hopping so that the bromine reduction in MEPBr_{2n+1} should be kinetically controlled even when a large overpotential, η , is applied to the electrode. Overall, the literature strongly implies that the MT in the polybromide IL is extraordinarily fast.

Verifications of electron transfer kinetic controlled current

A polycrystalline Pt ultramicroelectrode (UME) or carbon UME was dipped in the electrochemically synthesized MEPBr_{2n+1} droplet (Fig. S1a†). Fig. 2a shows the CVs of Br_2 reduction at the Pt UME and carbon UME. The reduction in MEPBr_{2n+1} involves an inner-sphere electrochemical reaction of Br_2 to generate Br^- :



Since Br_2 reduction proceeds at the UME in a polybromide droplet having excess Br_2 for only a short time, the electrochemical measurements do not cause significant perturbation to the composition of the polybromide IL. There was no side reaction such as the hydrogen evolution reaction in this potential range (Fig. S2†). Once the droplet was sufficiently large, the CVs were independent of its size. The CVs were also identical regardless of whether the reference electrode was located in the droplet or in the aqueous solution outside the droplet (Fig. S1b†). This shows that the impedance of the interface at MEPBr_{2n+1} droplet/aqueous solution as well as that of aqueous solution was negligible.

The CV of Br_2 reduction shows a steady-state current (i_{ss}) at a high η regardless of the electrode material (Fig. 2a). In common electrochemical systems, i_{ss} usually originates from the restricted supply of reactants: (i) spherical diffusion-limited steady-state current at the UME, (ii) slow adsorption of reactants on the electrode for adsorption-coupled electron transfer, or (iii) a slow homogeneous chemical reaction that produces reactants of an electrochemical reaction.

First, the current coming from Br_2 reduction is not governed by MT. The MT-limited voltammogram of the UME is sigmoidal, indicating the current-voltage relationship for a cathodic η :

$$i/i_0 = (1 - i/i_{\text{c}})\text{e}^{-\alpha f\eta} \quad (5)$$

Here, i_0 is the exchange current, i_{c} is the MT-limited steady-state current for the cathodic reaction, α is the transfer coefficient, and $f = F/RT$. For reduction, the anodic contribution of the cathodic branch is negligible at high η . Fig. 2b shows a comparison between the experimental CV and a simulated CV at the carbon UME based on eqn (5); i_0 , and α are calculated from the Tafel plot (Table S1†). Significant differences between these voltammograms indicate that Br_2 reduction on the carbon UME cannot be explained by the MT limiting current. Moreover, comparison of the CVs of Br_2 reduction at Pt UMEs of different diameters in Fig. 2c suggests that Br_2 reduction on Pt is not governed by MT. If the current of UME is limited by MT, one should observe different η values at which the current reaches the plateau, i_{ss} , according to the geometric radius of the electrodes.³² This is because the time for transition from planar

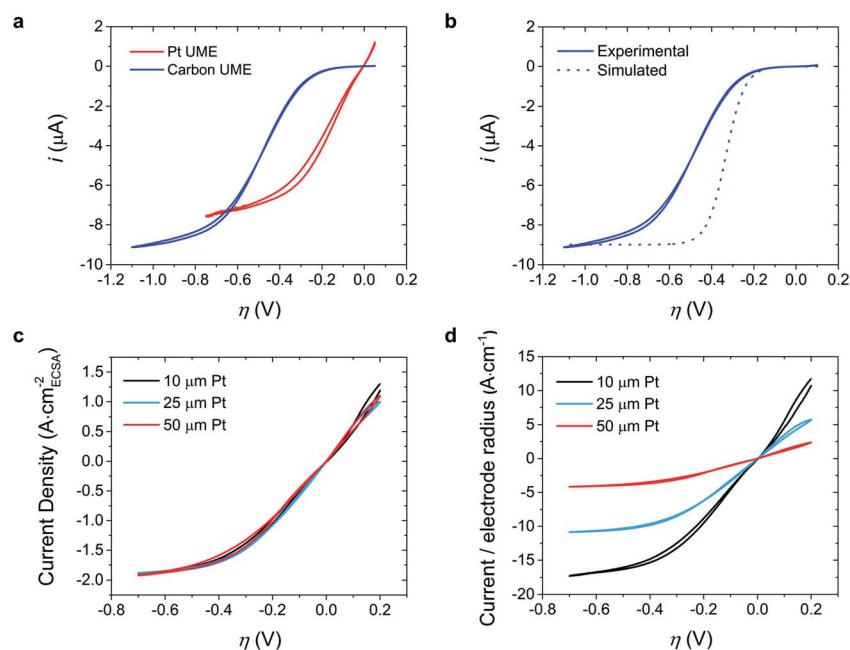


Fig. 2 Voltammograms of MEPBr_{2n+1} . (a) CVs at the Pt UME and carbon UME of diameters 10 μm and 11 μm , respectively, in MEPBr_{2n+1} (pH 3, phosphate buffer). The scan rate is 10 mV s^{-1} . (b) Comparison of the CV of the carbon UME in MEPBr_{2n+1} (pH 3, phosphate buffer) and the voltammogram predicted from eqn (5). (c and d) CVs at Pt UMEs of diameters 10, 25 and 50 μm in MEPBr_{2n+1} (pH 3, phosphate buffer). The scan rate is 10 mV s^{-1} . Current is divided by the (c) ECSA of Pt electrodes calculated from the H_{upd} voltammogram, and the (d) geometric radius of UME.



diffusion to spherical diffusion depends on the electrode radius when a MT limiting current flows at the UME. However, the shapes of the voltammograms for the Pt UMEs are identical regardless of the geometric diameters of the UMEs. In addition, i_{ss} at the Pt UME is regardless of its radius (Fig. 2d), which is not in agreement with that the MT-limited i_{ss} should be proportional to the geometrical radius of the electrode.³² Instead, the currents in Fig. 2c are proportional to the electrochemically active surface areas (ECSAs) of the electrodes, which are calculated from the charges of the hydrogen underpotential deposition (H_{upd}) in potassium phosphate buffer. These experimental results confirm that the rates of Br_2 reduction in the MEPBr_{2n+1} droplet at the carbon and Pt UMEs are not restricted by MT, even over a wide range of η (-0.7 V or more negative).

Br_2 reduction is composed of elementary steps: the Heyrovsky step, Volmer step, and Tafel step. Its mechanism differs depending on the reaction conditions, such as the reactants, solvents, and electrodes.^{33–35} If the rate-determining step (rds) is the Tafel step (ex. $Br_2 \rightarrow 2Br_{ad}$), the adsorption rate of redox species at the electrode surface should govern the overall Br_2 reduction rate, resulting in i_{ss} .³⁶ However, the current is not likely to be limited by the adsorption because Br_2 adsorption on Pt is fast enough.^{37–39}

It is widely accepted that the dissociation of Br_{2n+1}^- to Br_2 and Br^- is fast enough not to restrict the Br_2 reduction current.^{33,40} In the experiments, the voltammograms of Br_2 reduction vary sensitively with changes in the surface structures

of the electrodes. This can hardly occur when a homogeneous chemical reaction is the rds.

Application of the MHC model to electron transfer kinetics at Pt and carbon electrodes

The evidence consistently shows that i_{ss} in MEPBr_{2n+1} is not controlled by mass transport, adsorption, or a chemical reaction. On the other hand, the Br_2 reduction CVs on the Pt and carbon UME are fitted well to the MHC model (Fig. 3b and c), suggesting that the current is governed by the ET kinetics.

The MHC model is an ET kinetics model that incorporates energy distributions of electrons in electrodes and electrolytes into the Marcus theory of heterogeneous ET.^{17,41} Fig. 3a illustrates the principle of the MHC model. The distribution of occupied electronic states in the electrode, $g(E)$, follows the Fermi–Dirac equation, and the energy levels of electrons in the redox-active species of electrolytes vary with the degree of stabilization by solvation. The electron energy distribution of these species can be represented by a probability density function $W(E)$, which is a function of λ . In the case of reduction, an ET can occur from an occupied state in the electrode to the LUMO that has the corresponding energy. Thus, the rate constant of an electrochemical reduction is proportional to the integral of the product of the number of occupied states in the electrode and $W(E)$ of LUMOs of redox-active species along the electron energy level:^{41,42}

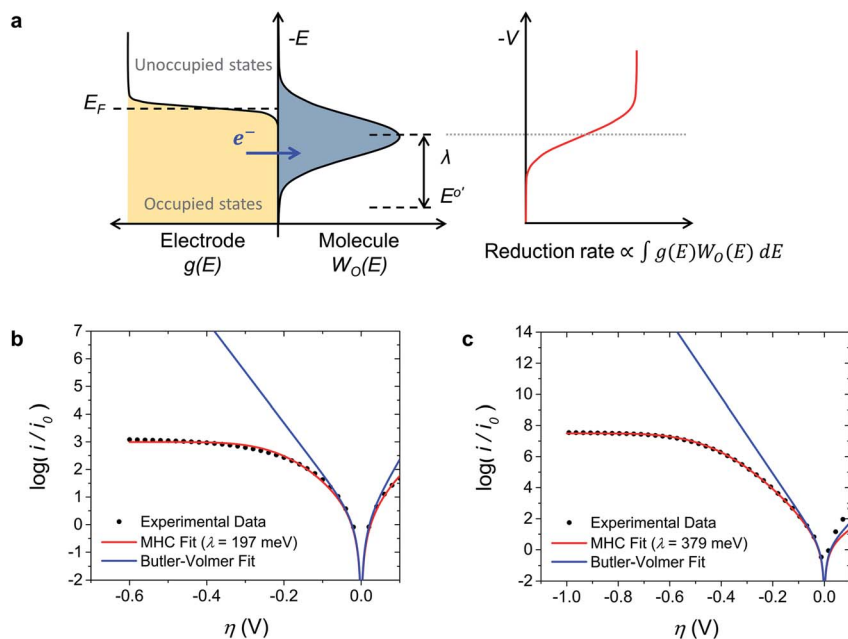


Fig. 3 Fitting the voltammograms to the Marcus–Hush–Chidsey model. (a) Electronic level diagram of the electrode–electrolyte interface in the MHC model. The relationship between the heterogeneous electron transfer rate and electronic states at an electrode–electrolyte interface in the case of reduction is shown. The y-axis is the energy, E . E_F , E^0 and V are the Fermi level, formal potential, and electrode potential, respectively. The bell curve represents $W_0(E)$ of LUMOs of redox-active species. k_{red}^{MHC} is proportional to the overlapped area of the number of occupied states in the electrode (yellow) and unoccupied states in the redox-active species (blue). k_{red}^{MHC} reaches half of the maximum when $-\Delta G^0$ for electron transfer is equal to λ . (b and c) Normalized Tafel plots for Br_2 reduction in the MEPBr_{2n+1} with pH 3 phosphate buffer at the (b) Pt UME and (c) carbon UME. The experimental data (black dots) are compared to the fitting results of the MHC model (red lines) and the Butler–Volmer model (blue lines).



$$k_{\text{red/ox}}^{\text{MHC}}(\eta) = Z \int_{-\infty}^{\infty} \exp\left(-\frac{(x - \lambda \pm e(E - E^0))^2}{4\lambda k_B T}\right) \frac{1}{1 + \exp(x/k_B T)} dx \quad (6)$$

Here, $k_{\text{red/ox}}^{\text{MHC}}$ is the rate constant of heterogeneous ET in the MHC model, Z is the pre-exponential factor accounting for the electronic coupling and the electronic density of states of the electrode, x is the energy, e is the elementary charge, and k_B is the Boltzmann constant. According to the MHC model, when η reaches a certain value, $k_{\text{red/ox}}^{\text{MHC}}$ does not increase with η because there is no corresponding $W(E)$ at high E . This can cause ET kinetics-limited i_{ss} . ET-limited i_{ss} normalized to i_0 depends on λ .⁴² η reaching ET-limited i_{ss} increases with λ .

Fig. 3b and c show the Tafel plots for Br_2 reduction in MEPBr_{2n+1} at the Pt UME and carbon UME (black dots) and the fits of these plots to the Butler-Volmer model (blue line) and MHC model (red line). The current-overpotential equation of the MHC kinetics derived from eqn (6) was utilized to fit the experimental data to the MHC model. The details on the derivation of the current-overpotential equation are described in the ESI.† The experimental data agree well with the MHC model, revealing $\lambda = 197$ and 379 meV for the Pt UME and carbon UME, respectively. The Butler-Volmer model can explain the experimental data only for small values of η . A carbon UME has remarkably higher λ than that of a Pt UME. Correspondingly, not only i_{ss}/i_0 but also η reaching i_{ss} are larger at carbon than at Pt.

Table S2† lists the λ values of Pt UMEs in buffers of different pHs on which λ was not dependent noticeably. In the microscopic ET theory, λ includes reorganization not only of solvents (outer-sphere reorganization energy, λ_o) but also of redox species (inner-sphere reorganization energy, λ_i). Since Br_2 reduction involves bromide adsorption on the electrode, λ may be associated with λ_i through changes in the bond length between bromide and the electrode during ET. The binding strength of Pt to bromide should be influenced by the solution pH,⁴³ and the MEPBr_{2n+1} droplet has a significant water content because even hydrophobic ILs are known to contain water when surrounded by an aqueous solution.¹² But the experimental λ of

Pt was not sensitive to pH. This shows that variation in Pt-Br binding strength with pH change affects reorganization energy within a limited range. Hence, the contribution of λ_i to λ is deemed to be negligible in this system.

Considering that Br_2 reduction is an inner-sphere reaction occurring in the IHP, λ_o should reflect the structure of the IHP. It is widely accepted that ϵ of the solvating medium is the most crucial factor for λ_o not only in dilute electrolytes,^{3,6,7,44,45} but also in RT-IL.⁴⁶⁻⁴⁸ Several theoretical studies have proposed that λ_o is inversely proportional to ϵ in RT-ILs.⁴⁶⁻⁴⁸ In ionic liquids, the redox species are mainly surrounded by ions rather than solvents.^{49,50} Hence, polarizabilities of ions should be important for ϵ . Table S3† compares reorganization energies of MEPBr_{2n+1} and 1-butyl-1-methylpyrrolidinium polybromide (MBPBr_{2n+1}). As predicted, the reorganization energy is smaller in MBPBr_{2n+1} which has higher cation polarizability.⁵¹ It is worth noting that interpretation of the reorganization energy of the IHP as a part of the EDL structure needs further study. Current reorganization energy models of RT-ILs are derived from the Debye-Hückel theory,⁴⁶⁻⁴⁸ which is based on the polarization of diluted electrolytes.^{52,53} But very strong Coulomb interactions among ions make the dielectric contribution of ILs complex. For example, hysteresis of the potential-dependent EDL structure of ILs has been reported,⁵⁴⁻⁵⁶ which may be associated with the ultraslow capacitive process of ILs.^{57,58} Changes of the effective dielectric constant of the IHP medium may alter the potential gradient between the nominal plane of the electrode and IHP.⁵⁹

Effect of electrode surface charges on the reorganization energy

To examine the intrinsic electrode properties affecting λ , bare Pt was compared to surface-modified Pt. TiO_2 is a suitable material for surface modification because it acts as a passive layer that minimizes the effects on the electrochemical reaction; moreover, it is chemically inert at the present experimental pH and potential. TiO_2 was electrodeposited on the Pt UME using a previously reported method.⁶⁰ The TiO_2 coverage (θ_{TiO_2}) of the TiO_2 -modified Pt electrode ($\text{TiO}_2@\text{Pt}$) was calculated by comparing the H_{upd} values before and after electrodeposition (Fig. 4a).

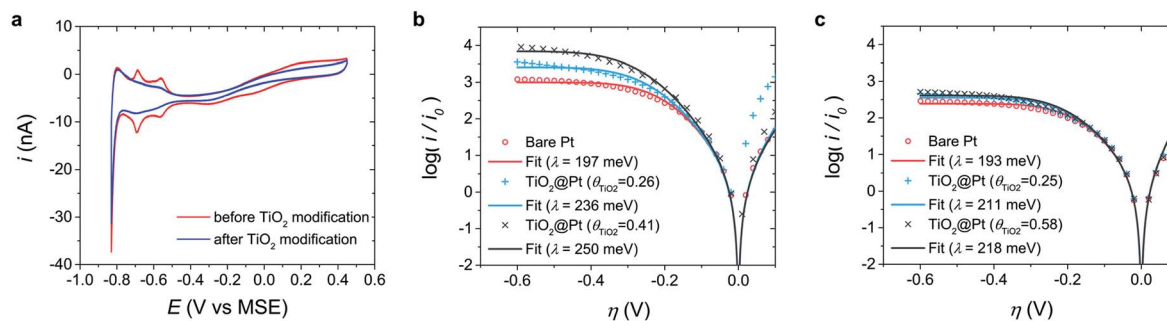


Fig. 4 Tafel analyses of $\text{TiO}_2@\text{Pt}$ electrodes in MEPBr_{2n+1} . (a) H_{upd} voltammogram at bare Pt and $\text{TiO}_2@\text{Pt}$ UMEs ($\theta_{\text{TiO}_2} = 0.25$) in 1 M aqueous potassium phosphate buffer (pH 3) recorded at a scan rate of 250 mV s^{-1} . The reference electrode was a mercury-mercurous sulfate electrode (MSE). (b and c) Normalized Tafel plots of $\text{TiO}_2@\text{Pt}$ with various θ_{TiO_2} values in MEPBr_{2n+1} at (b) pH 3 and (c) pH 4. The experimental data (markers) are compared to the fitting results of the MHC model (lines).



Fig. 4b and c show the normalized Tafel plots of bare Pt and TiO₂@Pt UMEs in the MEPBr_{2n+1} droplet synthesized at pH 3 and pH 4, respectively. The λ of TiO₂@Pt increases with θ_{TiO_2} at pH 3, while no remarkable changes were observed at pH 4. The pH dependency of TiO₂@Pt should be ascribed to the surface charge of TiO₂ because the isoelectric point of TiO₂ is pH 4–5;^{61,62} thus, the net surface charge of TiO₂ is positive at pH 3 and almost zero at pH 4. Accordingly, the surface charge of Pt does not change significantly with pH because halides suppress the formation of Pt oxide. EIS analysis shown in the ESI (Fig. S3†) supports the pH dependent PZC of Pt and TiO₂@Pt UMEs in MEPBr_{2n+1}. Introducing charges on the electrode surface will alter the ionic composition of the IHP in the IL. The EDL of an IL is very compact because of its extremely high ionic strength; most of the charges at the electrode surface are compensated in the IHP. When the electrode surface is positively charged, Br[−] anions, which have a higher charge density than other anions in MEPBr_{2n+1}, are likely to occupy the IHP to compensate for the surface charges. When the electrode is less positively charged, the MEP cations compensate for these charges proportionately. The polarizabilities of quaternary ammonium cations (10–20) are higher than that of Br[−] (5.8).⁶³ Accordingly, the IHP medium of TiO₂@Pt should be less polarizable at pH 3. Such a trend should manifest more at higher θ_{TiO_2} . The λ value of TiO₂@Pt at pH 4 is independent of θ_{TiO_2} , and almost zero charge of TiO₂ accounts for this. It is difficult to evaluate the λ value of TiO₂@Pt at pH 2 or less because TiO₂ is not stable at such low pH. In addition, the synthesis of the polybromide IL from an aqueous solution of MEPBr above pH 5 is accompanied by bromate formation, which can significantly change the composition of MEPBr_{2n+1}. The experimental results from TiO₂@Pt corroborate how electrode surface charge influences λ , which can explain the high λ of carbon compared to that of Pt. Being consistent with the work function and nonspecific adsorption of Br[−], the PZC of carbon is more negative than that of Pt.⁶⁴ Hence Br[−] is more probable in the IHP so that the medium in it should be less polarizable. It suggests that the net dielectric constant of the electrolytes around the carbon would be lower.

To date, majority of research has addressed electrocatalytic activity mostly in terms of the adsorption on the electrode materials based on the Sabatier principle.^{65–67} However, the community is increasingly realizing that it is insufficient to understand the electrocatalysis relying on the adsorption descriptor only.¹ In line with a few recent approaches, our report here supports that the properties of electrode materials give rise to characteristic change in the reaction environment of the liquid phase in the vicinity of the electrode surface, significantly influencing the electrocatalytic current.

Conclusions

RT-ILs have emerged as promising electrolytes for electrocatalytic reactions.^{8,9} Water-in-salt electrolytes, which are attracting significant interest owing to their potential applications in batteries⁶⁸ and electrocatalysis,⁶⁹ have EDL structures similar to those of RT-ILs.⁷⁰ In this study, the reorganization energy of Br₂ reduction was evaluated in a RT-IL, MEPBr_{2n+1},

which has an incredibly high diffusion coefficient for Br[−] and Br₂. A distinct kinetically controlled steady-state current was observed at a high overpotential, and the Br₂ reduction voltammogram could be successfully fitted to the MHC model. The reorganization energy as obtained for Br₂ reduction is sensitive to the surface charges of the electrode. This can be explained by the polarizability of the IHP, which dramatically varies with the surface charges. As demonstrated in this work, polybromide RT-ILs can serve as a model system not only for RT-ILs but also for water-in-salt electrolytes, in order to understand the electrocatalytic activities related to the EDL structure. Employing the MEPBr_{2n+1} RT-IL, one could investigate how a variety of electrode surfaces, including electrode materials and surface modifications other than TiO₂, affect the microscopic environment within the IHP so as to understand the given heterogeneous electron transfer kinetics. We believe that the findings of this work will provide new insights into the design of electrocatalysts.

In future research, a wide range of cations with different polarizabilities will be employed in the polybromide IL to collect more pieces of evidence for the dependence of λ on the polarizabilities of the solution. In addition, computational studies on the polybromide–electrode interface could help with evaluating quantitative contribution of λ_0 and λ_i , respectively.

Data availability

The data that supports the findings of this study are included in this article and ESI.† Further information is available from the corresponding author upon reasonable request.

Author contributions

M. Kim and S. Park were responsible for the conceptualization, performing the experiments, data analysis, writing the original draft and review, and editing. These authors contributed equally. T. D. Chung was responsible for the funding acquisition, writing and editing, and supervision of the work.

Conflicts of interest

There are no conflicts to declare.

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