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Fabrication of binary g-C₃N₄/UU-200 composites with enhanced visible-light-driven photocatalytic performance toward organic pollutant eliminations†

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In this study, g-C₃N₄/UU-200 heterojunction photocatalysts displaying superior photocatalytic activity for organic pollutant elimination under white LED light irradiation were fabricated *via* an *in situ* solvothermal method. The successful construction of a heterojunction between g-C₃N₄ and UU-200 was evidenced by X-ray diffraction, Fourier transform infrared spectroscopy, X-ray photoelectron spectroscopy, scanning electron microscopy, and transmission electron microscopy. The improved photocatalytic degradation of rhodamine B (RhB) and tetracycline hydrochloride (TCH) over g-C₃N₄/UU-200 compared with that over the individual components can be attributed to the anchoring of the g-C₃N₄ layered structure on the UU-200 surface promoting the decrease of the bandgap of UU-200, as confirmed by ultraviolet–visible diffuse reflectance spectroscopy, and to the light-induced charge separation efficiency stemming from a suitable heterojunction structure, which was revealed by photoluminescence spectroscopy and electrochemical analyses. Specifically, the 40% g-C₃N₄/UU-200 composite showed the highest photocatalytic activity toward the degradation of RhB (97.5%) within 90 min and TCH (72.6%) within 180 min. Furthermore, this catalyst can be recycled four runs, which demonstrates the potential of the g-C₃N₄/UU-200 composite as an alternative visible-light-sensitive catalyst for organic pollutant elimination.

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1 Introduction

Semiconductor-based photocatalysis has excellent potential in recent decades for environmental remediation. It has been recognized as a promising method to address the degradation

of many organic pollutants such as organic dyes and antibiotics due to its simplicity, economy, ambient conditions and usability.¹ The overuse of antibiotics and industrial production activities such as textiles has resulted in severe water pollution in many developing countries. Those antibiotics and organic dyes can cause irreversible harm to public health and ecological balance.² In this method, when a semiconductor photocatalyst is exposed to photons with energy equal to or higher than its band gap, the excitation of an electron from the valence band to the conduction band takes place to produce electron–hole pair (e⁻/h⁺).³ These electron–hole pairs can migrate to form reactive species such as hydroxyl radical ([•]OH) or superoxide radical anions ([•]O₂⁻). These reactive species are strong and nonselective oxidants that react with most organic pollutants contamination to break them down into smaller pieces and eventually into water and carbon dioxide.⁴ However, the limitation of this method is that the recombination rate of electron–hole pairs is fast, thus significantly reducing the photodegradation efficiency. Coupling two or more semiconductors to form heterojunction structures has been considered an effective method to overcome this drawback.^{5–7}

Metal–organic frameworks (MOFs) are attractive materials due to their potential applications in the fields of gas

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adsorption/separation,⁸ sensor technology,⁹ drug delivery,¹⁰ and especially in photocatalysis.^{11,12} Therefore, the development of new MOF structures with versatile properties is attracting increasing research attention. The most common metal ions used in the synthesis of MOFs include Fe(III), Zn(II), Co(II), Cu(II), and Zr(IV);¹³ however, bismuth-based MOFs (Bi-MOFs) remain less explored. Bismuth is nontoxic and possesses a flexible coordination geometry, which renders it a suitable metal for building MOF structures.¹⁴ Besides, bismuth compounds have catalytic activity and are often used as green catalysts. Recently, several Bi-MOFs with novel structures and interesting optical and photocatalytic properties have been synthesized.^{15–18} In particular, trimesate-based Bi-MOFs have been shown to possess diverse structures, such as CAU-17 (CAU: Christian-Albrechts-Universität), which bears nine Bi(III) ions, nine BTC³⁻ anions, and nine H₂O molecules in the asymmetric unit,¹⁹ UU-200 (UU: Uppsala University), which crystallizes in the space group *Pnmm* (No. 58) with three Bi(III) ions and four BTC³⁻ anions in the asymmetric unit,²⁰ and Bi-BTC having the *P21/n* space group, in which two Bi atoms form a {Bi₂O₁₄} dimer with six carboxyl groups from six different BTC³⁻ anions and the dimers are interconnected with BTC³⁻.¹⁵ Besides their diverse structures, trimesate-based Bi-MOFs have other advantages for photocatalytic applications, including high stability and easy preparation. However, their performance as photocatalysts is poor owing to their low visible-light response and low conductivity compared with traditional inorganic semiconductors.^{21,22} Consequently, further modification processes are required to boost the photocatalytic activity of trimesate-based Bi-MOFs, such as doping, formation of nanostructures, crystal regulation, and surface modification.²³ Among these strategies, the combination of trimesate-based Bi-MOFs with different materials to construct heterostructured composites has recently emerged as an effective method.²⁴ However, composites of trimesate-based Bi-MOFs with metal-free semiconductors having photocatalytic application are still rare.

Among the numerous metal-free semiconductor materials, graphitic carbon nitrides (g-C₃N₄) are considered as effective catalysts for the photodegradation of organic pollutants because of their raw material abundance, narrow bandgap (*ca.* 2.7 eV), facile fabrication methods, chemical stability, low economic cost, and versatility for structural modifications.²⁵ Unfortunately, their photocatalytic performance is rather low owing to the poor separation efficiency of photogenerated charges. Nevertheless, the separation of photogenerated electron/hole (e⁻/h⁺) pairs and, in turn, the photocatalytic activity, can be improved by combining g-C₃N₄ with other semiconductor photocatalysts to form heterostructured composite structures.²⁶ Moreover, g-C₃N₄ possesses a stacked two-dimensional structure with triazine or tri-*s*-triazine building units that facilitates the construction of their composites with MOFs due to the formation of π - π stacking interactions between the aromatic rings of the organic ligands in the MOFs and the triazine rings of g-C₃N₄.²⁷ Furthermore, g-C₃N₄ interacts and can be easily encapsulated on the MOF surfaces due to large surface electrostatic interactions,²⁸ promoting the separation of photoinduced charges. A series of MOF/g-C₃N₄ heterostructured composite structures have been constructed by combining g-C₃N₄

semiconductors with various MOFs, such as MIL-101(Fe),²⁹ NH₂-MIL-101(Fe),³⁰ UiO-66,¹⁹ MIL-53(Al),^{31,32} MIL-100(Fe),²² MIL-125(Ti),³³ MIL-88B(Fe),³⁴ and NH₂-MIL-88B(Fe),³⁵ which exhibit enhanced photocatalytic activity as a result of the improvement in pore volume, high-efficiency separation and transfer of photoinduced charge carriers, and visible-light responsiveness. However, despite these advances, the construction of binary g-C₃N₄/UU-200 composites for the removal of organic pollutants has not been reported to date.

Herein, g-C₃N₄/UU-200 composites were fabricated *via an in situ* solvothermal method and applied to the removal of rhodamine B (RhB) and tetracycline hydrochloride (TCH) for the first time. UU-200, a trimesate-based Bi-MOF, was selected as the representative MOF because it is nontoxic, relatively inexpensive, and photostable and has great potential in catalysis.³⁶ The g-C₃N₄ component was prepared according to a calcination method using a mixture of melamine and thiourea as nitrogen-rich organic precursors and NH₄Cl as a gas bubble template. The obtained g-C₃N₄ was then mixed with the UU-200 precursors for the *in situ* synthesis of g-C₃N₄/UU-200 composites. The obtained composites were characterized systematically and tested for RhB and TCH elimination under white LED light irradiation to evaluate the effect of the heterojunction structure. According to the experimental results, the g-C₃N₄/UU-200 composites show better performance in the photocatalytic degradation of RhB and TCH than the individual components. A plausible photocatalytic reaction mechanism is discussed in detail.

2 Experimental

2.1 Preparation of the g-C₃N₄/UU-200 heterojunction

A single-step solvothermal reaction was performed to synthesize g-C₃N₄/UU-200 heterojunctions with various g-C₃N₄ mass ratios. First, a certain amount of pristine g-C₃N₄ powder (Text S2, ESI[†]) was suspended in 60 mL DMF/MeOH mixture (1 : 1, v/v), and the mixture was then subjected to ultrasonication for 60 min to achieve a homogeneous aqueous suspension. Subsequently, a mixture of bismuth(III) nitrate pentahydrate (0.150 g, 0.303 mmol) and 1,3,5-benzenetricarboxylic acid (1.340 g, 6.060 mmol) was added to the above suspension with stirring for 30 min and was then transferred to a teflon-lined stainless steel autoclave, which was heated to 120 °C for 24 h. The resultant light-yellow g-C₃N₄/UU-200 product was isolated by centrifugation and then washed with DMF and MeOH solvents three times sequentially. The solvent was then removed at 80 °C for 12 h in a vacuum oven. The obtained composites were named *x*% g-C₃N₄/UU-200, where *x* denotes the weight percentage of g-C₃N₄ (*x* = 20%, 40%, and 60% w/w). For comparison, the synthesis of bare UU-200 and g-C₃N₄ was conducted as described in the ESI (Text S2 and S3, respectively). All the catalysts were characterized as shown in the ESI (Text S4).

2.2 Photocatalytic experimentals

The photocatalytic degradation of RhB and TCH was performed in a double-layer interbed glass beaker photoreactor with a capacity of 250 mL.^{32,36} In the photocatalytic activity test, an



aqueous suspension containing the photocatalyst (10 mg) and the organic compound (100 mL of RhB or TCH) was stirred magnetically without irradiation for 1 h to achieve the adsorption–desorption equilibrium. The initial concentration of RhB and TCH was 3×10^{-5} and 5 mg L^{-1} , respectively. The experiments were conducted at pH 5. Then, the photocatalytic process was started under visible-light irradiation using four white light Cree Xlamp XM-L2 LEDs ($4 \times 10 \text{ W}$) for a certain period of time at room temperature. During the photodegradation of organic pollutants, 5 mL of suspension was collected from the reaction mixture at a given interval of time and centrifuged for 30 min at 6000 rpm to remove the catalyst particles. Afterward, the decrease in the concentration of the organic pollutant was recorded using a Cary 60 ultraviolet-visible (UV-vis) spectrophotometer (Agilent, USA). The rate of photodegradation of organic pollutants was calculated as C/C_0 , where C is the concentration of the organic pollutant solution collected after a specified time interval and C_0 is the initial concentration. A trapping test was conducted to analyze the reactive species during the photocatalytic degradation of TCH as described in the ESI (Text S5).^{32,36} Moreover, a recycling test was performed as described in the ESI (Text S6).^{32,36} The electrochemical characterization of UU-200, $g\text{-C}_3\text{N}_4$, and 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ was conducted as described in the ESI (Text S7).

3 Results and discussion

3.1 Fabrication and characterization of precursors and composites

The crystalline structure of the UU-200 and $g\text{-C}_3\text{N}_4$ precursors and the $g\text{-C}_3\text{N}_4/\text{UU-200}$ composites was analyzed by means of X-ray diffraction (XRD) (Fig. 1). The XRD pattern of UU-200 matched well with those previously reported (CCDC 2103784),³⁷ and that of $g\text{-C}_3\text{N}_4$ exhibited peaks at 2θ of 13.1° and 27.5° corresponding to the (100) and (002) diffraction planes of $g\text{-C}_3\text{N}_4$ (JCPDS 87-1526), respectively.^{38,39} Meanwhile, the XRD patterns of the $g\text{-C}_3\text{N}_4/\text{UU-200}$ samples were similar to that of UU-200 and showed a peak corresponding to the (002) plane of $g\text{-C}_3\text{N}_4$, whose position was slightly right-shifted, suggesting that $g\text{-C}_3\text{N}_4$ was effectively anchored on the UU-200 surface.

The interactions between UU-200 and $g\text{-C}_3\text{N}_4$ were further studied using Fourier transform infrared (FT-IR) spectroscopy (Fig. 2). In the spectrum of the UU-200 catalyst, the stretching vibrations of the carboxylate groups were observed at 1360 and 1549 cm^{-1} , and the Bi–O bond vibration produced a band at 522 cm^{-1} .³⁹ Compared with H_3BTC , the stretching vibrations of the carboxylate groups were right-shifted due to the interaction between the Bi(III) ions and the carboxylate groups of the BTC^{3-} linker.²² The FT-IR spectrum of pure $g\text{-C}_3\text{N}_4$ exhibited peaks at 810 , $1200 - 1700$, and $3100 - 3300 \text{ cm}^{-1}$, which can be ascribed to the characteristic breathing modes of the *s*-triazine units and $sp^2 \text{ C}=\text{N}$,⁴⁰ the stretching vibrational modes of $\text{C}=\text{N}$ heterocycles,⁴¹ and the N–H stretching vibration,^{42–44} respectively. The spectra of the $g\text{-C}_3\text{N}_4/\text{UU-200}$ composites showed the typical absorption bands of UU-200 and $g\text{-C}_3\text{N}_4$, although a shift in the vibration peaks at 810 and 1240 cm^{-1} was observed compared

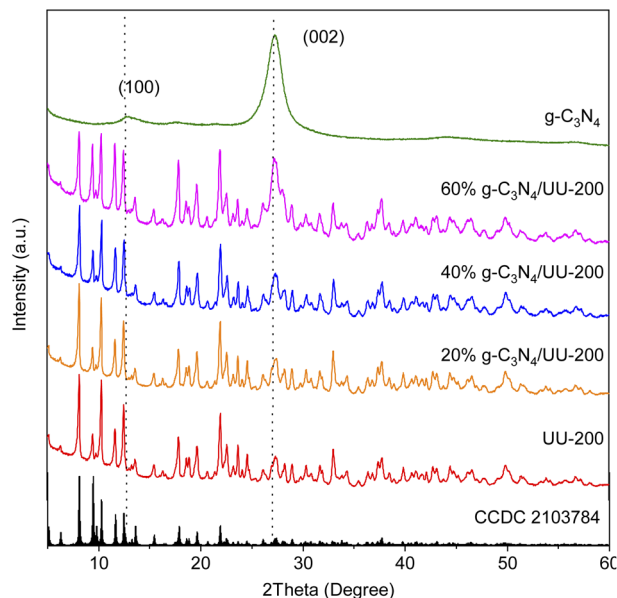


Fig. 1 XRD patterns of UU-200, $g\text{-C}_3\text{N}_4$ and UU-200 $g^{-1}\text{-C}_3\text{N}_4$ composites.

with pure $g\text{-C}_3\text{N}_4$, which can be attributed to the interactions between $g\text{-C}_3\text{N}_4$ and UU-200.

The morphologies of the binary $g\text{-C}_3\text{N}_4/\text{UU-200}$ composites and their individual components were investigated using scanning electron microscopy (SEM) and transmission electron microscopy (TEM) techniques (Fig. 3). It was found that the UU-200 catalyst exhibited condensed bundles with a rod-like shape having a rod thickness of around $200 - 400 \text{ nm}$ and smooth surfaces (Fig. 3 A and D), which is in agreement with previously reported experimental data for UU-200.12 A layered structure

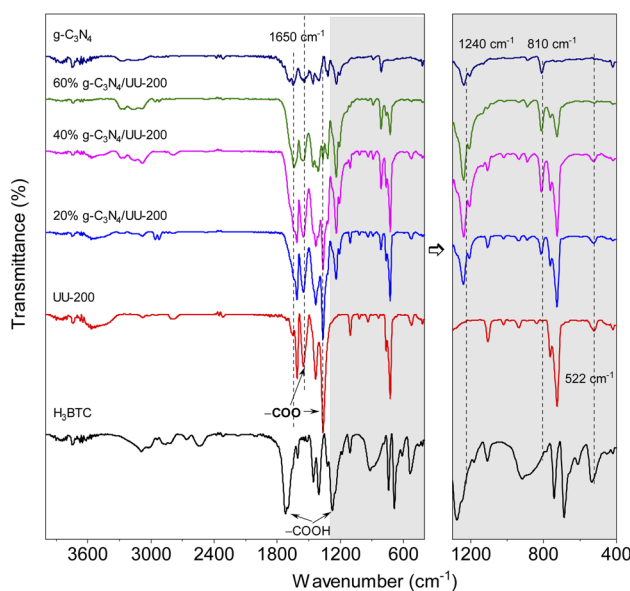


Fig. 2 FT-IR spectra of UU-200, $g\text{-C}_3\text{N}_4$ and UU-200/ $g\text{-C}_3\text{N}_4$ composites.



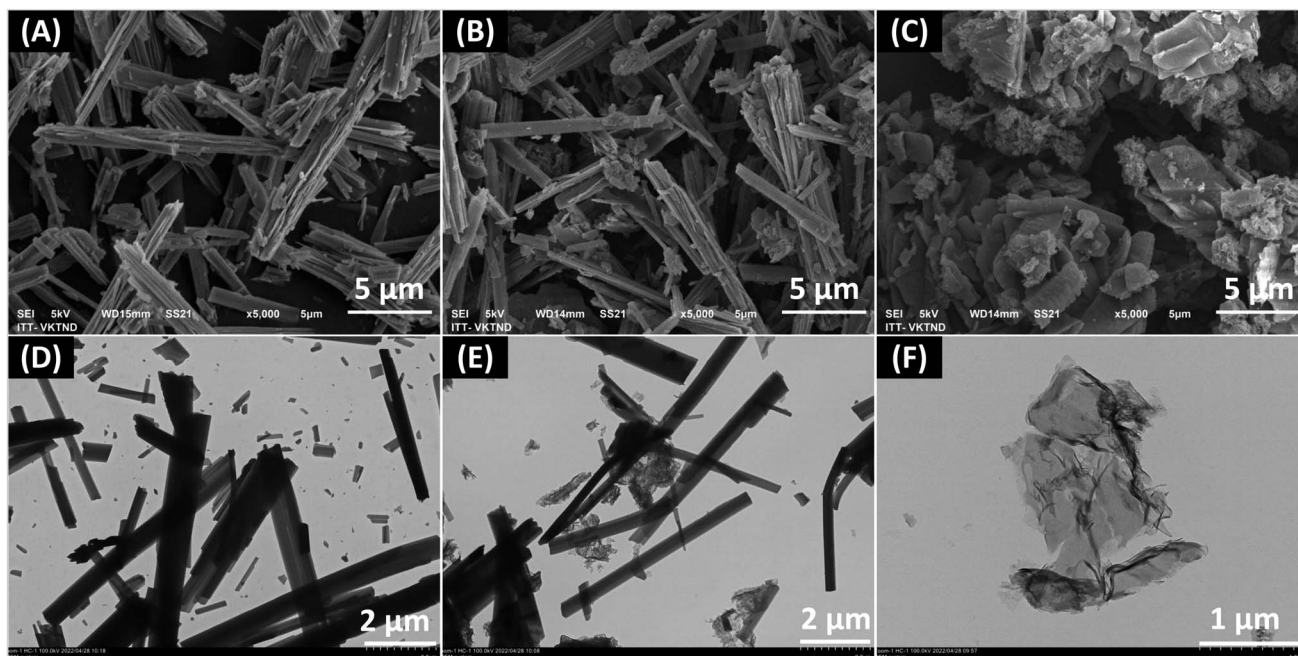


Fig. 3 SEM images (A–C) and TEM images (D–F) of UU-200 (A and D), $g\text{-C}_3\text{N}_4/\text{UU-200}$ composite (B and E) and $g\text{-C}_3\text{N}_4$ (C and F).

containing irregular pieces was observed for $g\text{-C}_3\text{N}_4$ (Fig. 3C). Meanwhile, the TEM images of $g\text{-C}_3\text{N}_4$ showed relatively thin and transparent features with layered wrinkles (Fig. 3 F). The surface of the $g\text{-C}_3\text{N}_4/\text{UU-200}$ samples was rougher than that of pure UU-200 (Fig. 3B and S1[†]), indicating that the UU-200 surface was covered by $g\text{-C}_3\text{N}_4$. The binary composite structure could be observed clearly in the TEM images of the $g\text{-C}_3\text{N}_4/\text{UU-200}$ samples (Fig. 3E and S2[†]). According to these observations, the $g\text{-C}_3\text{N}_4/\text{UU-200}$ heterostructure could be expected to be suitable for photocatalysis under visible-light irradiation.

The surface chemical composition and the chemical states of the surface elements present in UU-200, 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$, and $g\text{-C}_3\text{N}_4$ were analyzed using X-ray photoelectron spectroscopy (XPS). The XPS survey spectrum of pristine UU-200 showed the presence of Bi, O, and C elements, while that of bare $g\text{-C}_3\text{N}_4$ revealed C, N, and O elements (Fig. S3[†]). The presence of C, N, O, and Bi elements was detected in the XPS survey spectrum of 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ (Fig. S3[†]).

The binding energy (BE) values of the Bi 4f spectra of UU-200 and 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ (Fig. 4A) were located at around 159.5 (Bi 4f_{7/2}) and 164.8 eV (Bi 4f_{5/2}) with a ΔE (Bi 4f_{5/2}–Bi 4f_{7/2}) of 5.3 eV, indicating that Bi was mainly in the form of Bi(III).⁴⁵ The C 1s peak of 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ can be divided into five main peaks at 284.8, 286.3, 287.9, 288.9, 289.5, and 291.9 eV (Fig. 4C). The peak at 284.8 eV can be attributed to the C–C/C=C bonds of UU-200 and $g\text{-C}_3\text{N}_4$.^{40,46} The BEs at 286.3 and 288.9 eV correspond to the C–O and C=O bonds, respectively, of UU-200 and $g\text{-C}_3\text{N}_4$.^{22,47} The peak at 287.9 eV is characteristic of the C–N/C=N bonds of $g\text{-C}_3\text{N}_4$.⁴⁸ The peak at 289.5 eV can be assigned to the O–C=O bonds of UU-200, and that at 291.9 eV is ascribable to $\pi\text{-}\pi^*$ satellite bands in $g\text{-C}_3\text{N}_4$.⁴⁹ Fig. 4D shows the O 1s spectra of UU-200, 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$, and $g\text{-C}_3\text{N}_4$. For $g\text{-C}_3\text{N}_4$, the

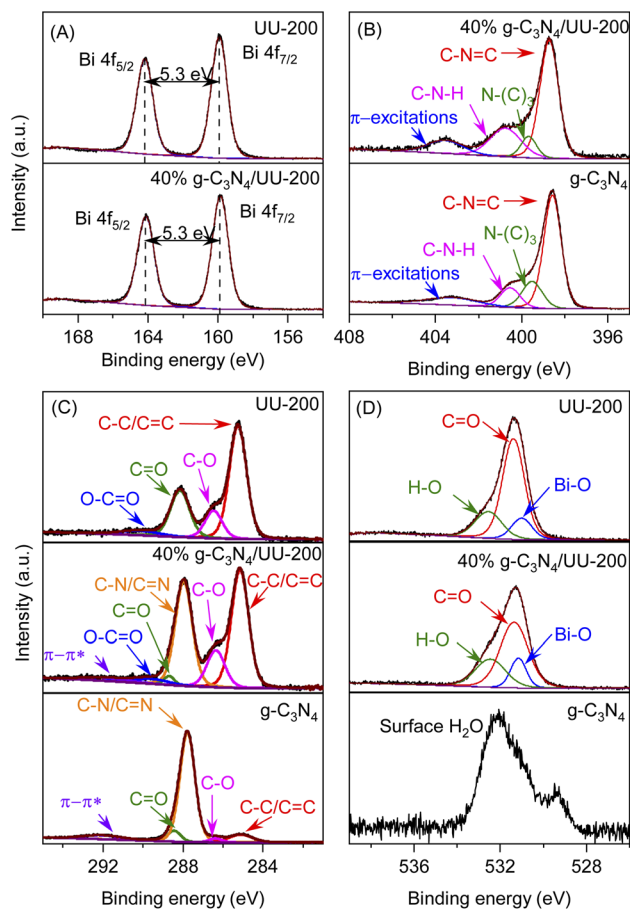


Fig. 4 High-resolution Bi 4f XPS spectra of UU-200 and 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ (A); high-resolution N 1s XPS spectra of $g\text{-C}_3\text{N}_4$ and 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ (B); high-resolution C 1s (C) and O 1s (D) XPS spectra of $g\text{-C}_3\text{N}_4$, UU-200 and 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$.



deconvoluted peak at 532.0 eV can be ascribed to the surface-adsorbed H₂O in the three samples.^{50,51} The O 1s spectra of UU-200 and 40% g-C₃N₄/UU-200 exhibited three peaks at 530.9, 531.4, and 532.5 eV, which correspond to Bi–O bands, the carboxylate groups of the BTC linkers, and surface hydroxyl groups, respectively.^{52,53} The N 1s peak of the 40% g-C₃N₄/UU-200 composite can be separated into four peaks at 398.7, 399.7, 400.8, and 403.6 eV (Fig. 4 B). These peaks stem from the C–N=C bonds of the triazine rings, the tertiary nitrogen N–(C)₃ groups, the C–N–H bonds of the free amino groups, and π -excitations, respectively.⁴⁶ Meanwhile, pristine g-C₃N₄ gave rise to a C–N=C peak at 398.5 eV. Compared with pristine g-C₃N₄, the BE of the C–N=C bonds of the 40% g-C₃N₄/UU-200 composite shifted to higher binding energy by 0.2 eV, revealing an interaction between the Bi and N atoms in the g-C₃N₄/UU-200 composite. Therefore, the XPS spectra further demonstrated the formation of the binary g-C₃N₄/UU-200 composites and the interaction between UU-200 and g-C₃N₄ components. The BEs at 286.3 and 288.9 eV correspond to the C–O and C=O bonds, respectively, of UU-200 and g-C₃N₄.^{22,47} The peak at 287.9 eV is characteristic of the C–N/C=N bonds of g-C₃N₄.⁴⁸ The peak at 289.5 eV can be assigned to the O–C=O bonds of UU-200, and that at 291.9 eV is ascribable to π - π^* satellite bands in g-C₃N₄.⁴⁹ Fig. 4D shows the O 1s spectra of UU-200, 40% g-C₃N₄/UU-200, and g-C₃N₄. For g-C₃N₄, the deconvoluted peak at 532.0 eV can be ascribed to the surface-adsorbed H₂O in the three samples.^{50,51} The O 1s spectra of UU-200 and

40% g-C₃N₄/UU-200 exhibited three peaks at 530.9, 531.4, and 532.5 eV, which correspond to Bi–O bands, the carboxylate groups of the BTC linkers, and surface hydroxyl groups, respectively.^{52,53} The N 1s peak of the 40% g-C₃N₄/UU-200 composite can be separated into four peaks at 398.7, 399.7, 400.8, and 403.6 eV (Fig. 4 B). These peaks stem from the C–N=C bonds of the triazine rings, the tertiary nitrogen N–(C)₃ groups, the C–N–H bonds of the free amino groups, and π -excitations, respectively.⁵⁴ Meanwhile, pristine g-C₃N₄ gave rise to a C–N=C peak at 398.5 eV. Compared with pristine g-C₃N₄, the BE of the C–N=C bonds of the 40% g-C₃N₄/UU-200 composite shifted to higher binding energy by 0.2 eV, revealing an interaction between the Bi and N atoms in the g-C₃N₄/UU-200 composite. Therefore, the XPS spectra further demonstrated the formation of the binary g-C₃N₄/UU-200 composites and the interaction between UU-200 and g-C₃N₄ components.

The physicochemical properties of the samples were investigated *via* N₂ adsorption–desorption isotherms. Fig. 5 displays the N₂ adsorption–desorption isotherms and the pore size distribution curves of UU-200, g-C₃N₄, and the g-C₃N₄/UU-200 composites. As shown in Fig. 5, all samples exhibited type IV sorption isotherms according to IUPAC classification, indicating their mesoporous characteristics.⁵⁵ The Brunauer–Emmett–Teller (BET) specific surface area (S_{BET}) values of the 20% g-C₃N₄/UU-200, 40% g-C₃N₄/UU-200, and 60% g-C₃N₄/UU-200 composites were 44.6, 42.8 and 40.1 m² g^{−1}, respectively, which were higher than that of UU-200 (37.6 m² g^{−1}) (Table 1). As shown in the SEM results, the presence of g-C₃N₄ during the synthesis of UU-200 may affect the formation of the condensed bundles of the rod-like crystals of UU-200, resulting in a decrease in their size. The S_{BET} of the composites decreased with the increase of the g-C₃N₄ loading. Although the S_{BET} values of the g-C₃N₄/UU-200 composites were lower than that of g-C₃N₄, they were much higher than that of pristine UU-200, demonstrating that the g-C₃N₄/UU-200 composites could provide abundant surface active sites and accelerate the transfer of photogenerated charges.⁵⁶ In addition, the large pore size of 40% g-C₃N₄/UU-200 can be expected to enhance the adsorption capacity for organic molecules on the surface active sites of the composite, thereby accelerating the photocatalytic activity.

The optical adsorption properties of the g-C₃N₄/UU-200 composites and their individual components were studied by

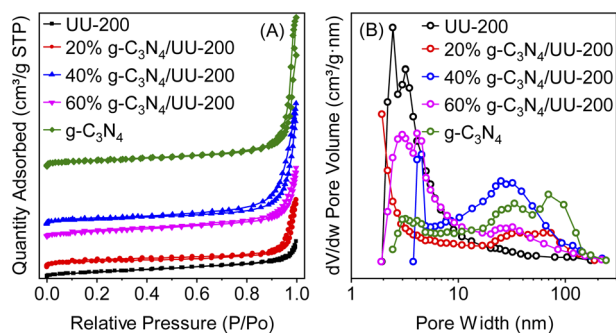


Fig. 5 (A) N₂ adsorption–desorption isotherms and (B) pore size distributions of g-C₃N₄. UU-200 and 40% g-C₃N₄/UU-200 composites.

Table 1 BET specific surface area value (S_{BET}), pore volume (V), average pore width (D), bandgap energy (E_g), and reaction rate constant (k) of UU-200, g-C₃N₄, and the g-C₃N₄/UU-200 composites

Samples	S_{BET}	V	D	E_g	RhB		TCH	
	(m ² g ^{−1})	(10 ^{−3} cm ³ g ^{−1})	(nm)	(eV)	k (10 ^{−3} min ^{−1})	R^2	k (10 ^{−3} min ^{−1})	R^2
UU-200	37.6	95.9	2.45	3.77	17.4	0.987	4.33	0.983
20% g-C ₃ N ₄ /UU-200	44.6	164	1.92	2.68	31.3	0.993	5.00	0.958
40% g-C ₃ N ₄ /UU-200	42.8	311	4.53	2.66	38.2	0.993	5.87	0.973
60% g-C ₃ N ₄ /UU-200	40.1	178	4.15	2.64	23.6	0.983	5.63	0.974
gC ₃ N ₄	47.1	399	3.24	2.54	16.1	0.993	5.11	0.998



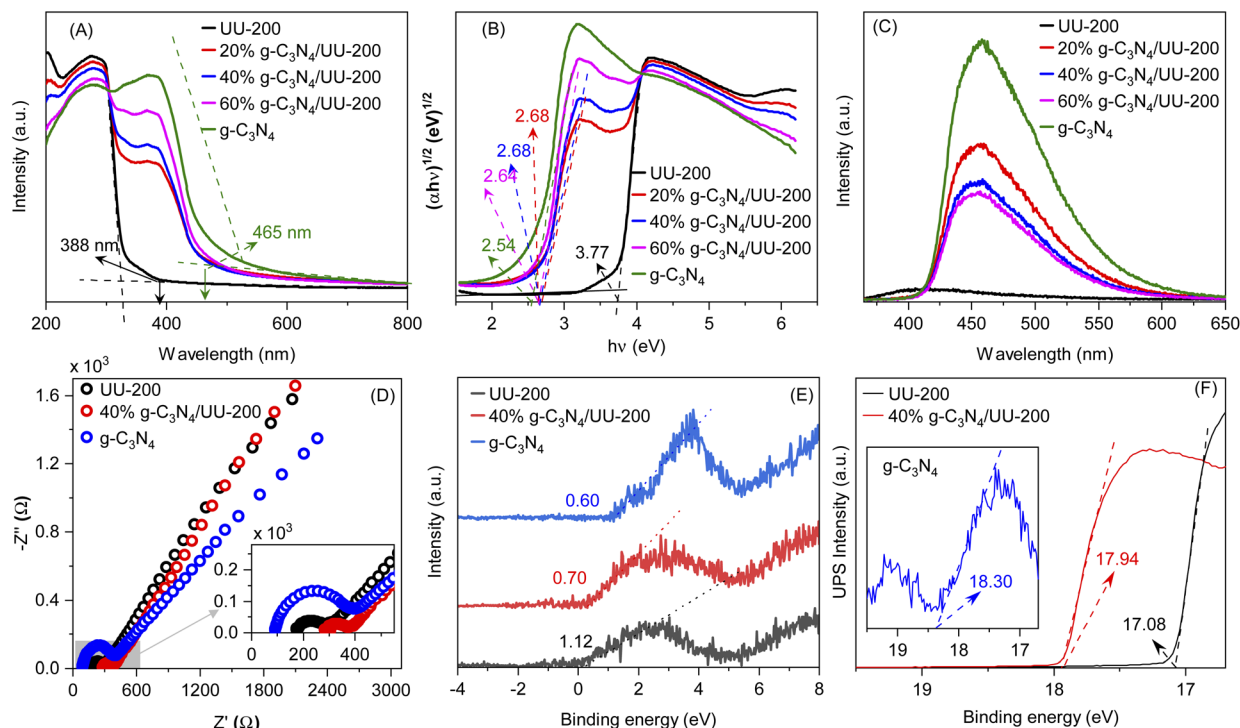


Fig. 6 UV-vis diffuse reflectance spectra (A), band-gap energies (B), photoluminescence spectra (C), EIS Nyquist plot (D), XPS-VB spectra (E) and UP spectra (F) of UU-200, g-C₃N₄ and g-C₃N₄/UU-200 composites.

UV-vis diffuse reflectance spectroscopy (UV-vis DRS) (Fig. 6 A). The photoabsorption edges of UU-200 and g-C₃N₄ were observed at 330 and 465 nm, respectively. Meanwhile, the UV-vis DRS spectra of the g-C₃N₄/UU-200 photocatalysts displayed a red shift compared with that of UU-200 because of the distribution of g-C₃N₄ layers on the UU-200 surface, which suggests that g-C₃N₄ can extend the absorption of UU-200 into the visible-light region. The energy gaps (E_g) of UU-200, g-C₃N₄, 20% g-C₃N₄/UU-200, 40% g-C₃N₄/UU-200, and 60% g-C₃N₄/UU-200 were obtained from the plots of $(\alpha h\nu)^{1/2}$ versus photon energy (Fig. 6 B), and the estimated values are shown in Table 1.^{57,58}

The recombination of photogenerated charges can create photons, which can be detected with high sensitivity by photoluminescence (PL) emission spectroscopy. Specifically, a lower PL intensity indicates less recombination of photoexcited charge carriers, which is related to higher photocatalytic activity.⁵⁹ The PL spectra of UU-200, g-C₃N₄, and 40% g-C₃N₄/UU-200 composites are presented in Fig. 6C. The highest PL emission peak intensity was observed for g-C₃N₄ at around 459 nm, indicating its highest recombination of photoexcited charge carriers. However, after the incorporation of UU-200 to construct the g-C₃N₄/UU-200 composites, the intensity of the PL emission peak decreased, which means that the recombination of charge carriers was comparatively suppressed in the composites. The PL emission analyses revealed that the combination of UU-200 and g-C₃N₄ efficiently enhanced the separation of charge carriers, thereby improving the photocatalytic activity. Especially, UU-200 exhibited the lowest

intensity in the emission peak at around 400 nm, implying the lowest recombination rate of charge carriers. However, the photocatalytic activity of UU-200 was much lower than that of the 40% g-C₃N₄/UU-200 composite. The highest photocatalytic efficiency of the 40% g-C₃N₄/UU-200 composite could be ascribed to its low recombination rate of charge carriers. In fact, the strong migration and separation of photoexcited charge carriers was confirmed by analyzing the electrochemical impedance spectra (EIS). Two common methods have applied for treating the impedance data, the Nyquist and the Bode plots.⁶⁰ Fig. 6D shows the Nyquist impedance plots of the UU-200, g-C₃N₄, and 40% g-C₃N₄/UU-200 samples. The corresponding Bodes' plots used to determine the equivalent electrochemical circuit are shown in Fig. S4 (ESI)†. The arc radius on the EIS Nyquist plot shows a clear trend of charge transfer behavior in the three different photocatalysts, which correlates with interfacial charge transport resistance; generally, a small arc radius is indicative of a fast charge transfer rate on the catalyst surface.^{60,61} As expected, the EIS Nyquist plot of the 40% g-C₃N₄/UU-200 composite showed a smaller arc radius than that of UU-200, which means that the separation and migration of charge carriers was enhanced in the 40% g-C₃N₄/UU-200 composite.

To reveal band position of g-C₃N₄/UU-200, UPS and VB-XPS were investigated. Fig. 6E and F show the XPS-VB spectra and UP spectra of UU-200, g-C₃N₄, and 40% g-C₃N₄/UU-200, respectively. The energy difference between the Fermi level and valence band maximum ($E_F - E_{VBM}$) of +0.60, +1.12, and +0.70 eV for UU-200, g-C₃N₄, and 40% g-C₃N₄/UU-200,



respectively. The work function (ϕ) of UU-200 (-4.14 eV vs. vacuum), $g\text{-C}_3\text{N}_4$ (-2.92 eV vs. vacuum), and 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ (-3.28 eV vs. vacuum) was calculated from eqn (1). ΔE is the cut-off energy of secondary electron emission spectra (Fig. 6F). Using the above bandgap energies and the eqn (2),^{62,63} the conduction band (CB) potential of UU-200, $g\text{-C}_3\text{N}_4$, and 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ were estimated, showing that the CB level of UU-200 was more negative than that of $g\text{-C}_3\text{N}_4$. The bandgap structures and charge migration of 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ are described in Fig. S5.†

$$\phi = 21.22 - \Delta E \quad (1)$$

$$E_{\text{VB}} = E_{\text{g}} + E_{\text{CB}} \quad (2)$$

3.2 Photocatalytic performances

Next, the photocatalytic performance of the $g\text{-C}_3\text{N}_4/\text{UU-200}$ composites and their individual components in the degradation of RhB and TCH under visible-light illumination was investigated (Fig. 7A and D). The RhB and TCH degradation hardly occurred in the absence of photocatalysts. In darkness,

all the photocatalysts could adsorb certain amounts of RhB (32.5% for UU-200, 26.8% for 20% $g\text{-C}_3\text{N}_4/\text{UU-200}$, 20.9% for 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$, 11.0% for 60% $g\text{-C}_3\text{N}_4/\text{UU-200}$, and 1.59% for $g\text{-C}_3\text{N}_4$) and TCH (28.5% for UU-200, 27.5% for 20% $g\text{-C}_3\text{N}_4/\text{UU-200}$, 24.9% for 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$, 17.6% for 60% $g\text{-C}_3\text{N}_4/\text{UU-200}$, and 4.23% for $g\text{-C}_3\text{N}_4$) and reached the adsorption-desorption equilibrium within 30 min. In the presence of the photocatalysts, the photocatalytic reaction proceeded efficiently under illumination of white LED light. The photocatalytic degradation efficiency for RhB and TCH over UU-200 was 85.0% and 67.6%, respectively. Meanwhile, the addition of $g\text{-C}_3\text{N}_4$ to UU-200 improved the photocatalytic degradation efficiency. Thus, 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ exhibited the highest degradation efficiencies for RhB (97.5%) and TCH (72.6%). The Hinshelwood plot used to study the kinetic of the RhB and TCH photodegradation process and the slopes of the plot show the rate constant for the photodegradation.^{64,65} The reaction rate constant for the degradation of RhB followed the order of $g\text{-C}_3\text{N}_4$ (0.0161 min^{-1}) < UU-200 (0.0174 min^{-1}) < 60% $g\text{-C}_3\text{N}_4/\text{UU-200}$ (0.0236 min^{-1}) < 20% $g\text{-C}_3\text{N}_4/\text{UU-200}$ (0.0313 min^{-1}) < 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ (0.0382 min^{-1}) (Fig. 7B). The 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$ composite also showed the highest photocatalytic

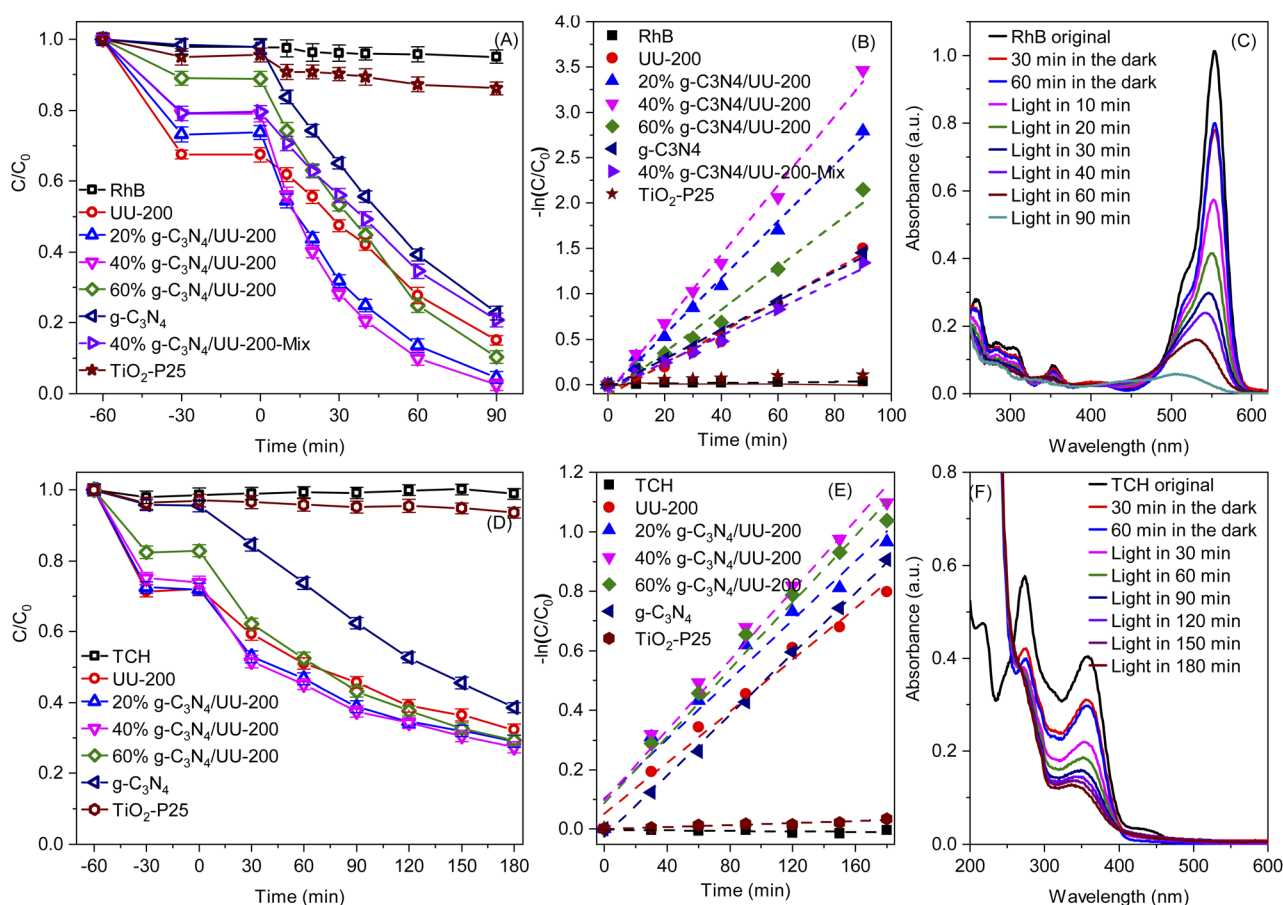


Fig. 7 Photocatalytic degradation of RhB (A) and TCH (D) over the as-prepared samples under visible-light irradiation; the first-order kinetics of RhB (B) and TCH (E) photocatalytic degradation; UV-vis absorption spectra of RhB (C) and TCH (F) solution separated from catalyst suspensions during illumination using 40% $g\text{-C}_3\text{N}_4/\text{UU-200}$. Conditions: [catalyst] = 10 mg, [RhB] = $3 \times 10^{-5} \text{ mg L}^{-1}$, [TC] = 5 mg L^{-1} , $V = 100 \text{ mL}$, Light source: White light LED ($4 \times 10 \text{ W}$).



reaction rate constant for the degradation of TCH (0.00587 min^{-1}), which was 1.36, 1.17, 1.04, and 1.15 times higher than that of UU-200 (0.00433 min^{-1}), 20% $\text{g-C}_3\text{N}_4/\text{UU-200}$ (0.00500 min^{-1}), 60% $\text{g-C}_3\text{N}_4/\text{UU-200}$ (0.00563 min^{-1}), and $\text{g-C}_3\text{N}_4$ (0.00511 min^{-1}), respectively (Fig. 7E). In addition, Fig. 7C and F reveals a change in the absorption spectra of RhB and TCH using 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$, respectively, owing to the photocatalytic removal of RhB and TCH during irradiation. Additionally, the photocatalytic performance of a physical mixture containing 40 wt% of $\text{g-C}_3\text{N}_4$ and UU-200 was lower than that of the 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$ composite, confirming that the improved photocatalytic activity of the composite stems from the construction of a suitable heterojunction. We also compared the photocatalytic degradation efficiency of the synthesized photocatalyst with commercial TiO_2 Degussa P25 ($\text{TiO}_2\text{-P25}$) photocatalyst. The experimental results indicated that $\text{g-C}_3\text{N}_4/\text{UU-200}$ composites exhibited an advantage of RhB (Fig. 7A) and TCH (Fig. 7D) elimination compared with $\text{TiO}_2\text{-P25}$ under LED light radiation. So, the $\text{g-C}_3\text{N}_4/\text{UU-200}$ composite has been considered a potential catalyst for practical application in polluted water treatment.

3.3 Photocatalytic mechanism

Next, the photocatalytic mechanism of 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$ was investigated by determining the main oxidative species involved in the photocatalytic reaction, such as hydroxyl radicals ($\cdot\text{OH}$), active e^- , active h^+ , and superoxide anion radical ($\cdot\text{O}_2^-$) (Fig. 8A, B). In the visible-light photodegradation of RhB, the photocatalytic performance decreased remarkably after adding $\text{Na}_2\text{C}_2\text{O}_4$ and benzoquinone (BQ), suggesting that both h^+ and $\cdot\text{O}_2^-$ are the major reactive species. Meanwhile, the addition of $\text{K}_2\text{Cr}_2\text{O}_7$ inhibited slightly the photocatalytic process, revealing e^- as the secondary active species. In the case of TCH removal over 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$, the photocatalytic degradation decreased significantly in the presence of $\text{Na}_2\text{C}_2\text{O}_4$, $\text{K}_2\text{Cr}_2\text{O}_7$, and BQ, which indicates that h^+ , e^- , and $\cdot\text{O}_2^-$ play more significant roles in the photocatalytic process than $\cdot\text{OH}$.⁶⁶

On the basis of these results, the photocatalytic mechanism depicted in Fig. 8C can be proposed for the photocatalytic

degradation of TCH and RhB over 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$. Upon exposure to LED light, $\text{g-C}_3\text{N}_4$ in the heterostructure is excited to produce photogenerated e^-/h^+ pairs. Then, the e^- in the CB of $\text{g-C}_3\text{N}_4$ can diffuse to the VB of UU-200 driven by the internal electric field, thus promoting the separation of photogenerated charge carriers. The e^- on $\text{g-C}_3\text{N}_4$ can capture surface-absorbed O_2 to generate $\cdot\text{O}_2^-$ because the CB level of $\text{g-C}_3\text{N}_4$ is higher than the redox potential of $\text{O}_2/\cdot\text{O}_2^-$ (-0.33 eV vs. NHE),^{67,68} which is consistent with the results of the trapping test. Meanwhile, the transferred h^+ on UU-200 and the remained h^+ in the CB of $\text{g-C}_3\text{N}_4$ cannot oxidize the adsorbed H_2O to form $\cdot\text{OH}$ because the E_{VB} of $\text{g-C}_3\text{N}_4$ and UU-200 are lower than the $\cdot\text{OH}/\text{H}_2\text{O}$ potential ($+2.68 \text{ eV vs. NHE}$).^{69,70} Then, these reactive species (h^+ , e^- , and $\cdot\text{O}_2^-$) can attack the $\text{C}=\text{C}$ bond of RhB and TCH to generate degradation intermediates. The main byproducts of the visible-light catalytic degradation of TCH by the 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$ composite were identified by performing a liquid chromatography–mass spectrometry analysis (Fig. S6†). According to the results, Fig. 9 depicts the possible pathways for the photocatalytic TCH degradation. The characteristic ion signal observed at m/z 445 corresponds to TCH, and its intensity decreased gradually with increasing the irradiation time. Various intermediates were also detected, namely, P_1 (m/z 427), P_2 (m/z 410), and P_3 (m/z 288), which were generated *via* dihydroxylation, deamination, and ring-opening reactions, respectively.^{71,72} These intermediates further decomposed into smaller fragments, including P_4 (m/z 85), P_5 (m/z 74), and P_6 (m/z 57). Upon prolonging the irradiation time, all intermediates were converted into H_2O and CO_2 .

3.4 Reusability and stability of the photocatalyst

To examine the photostability of the binary 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$ composite, reusability and stability tests were performed. As shown in Fig. S7,† after four cycling tests, the photodegradation efficiency of TCH did not decrease significantly, and 62.8% of TCH was degraded after four cycles of operation, demonstrating the high stability of 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$. The slight decrease in the photocatalytic activity of the reused 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$ composite can be attributed to the centrifugation process

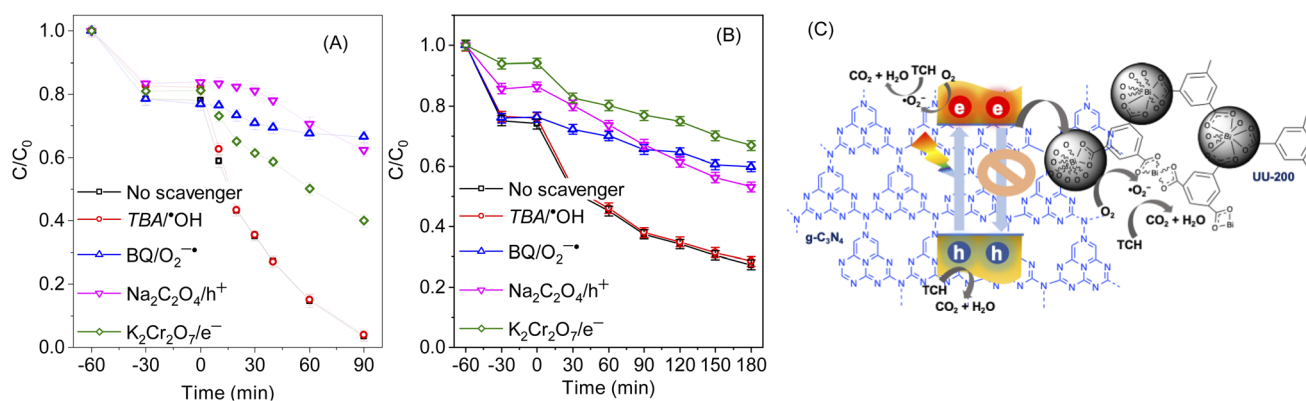


Fig. 8 Effects of different scavengers on the degradation of RhB (A) and TCH (B) on 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$; photodegradation mechanism of TCH over 40% $\text{g-C}_3\text{N}_4/\text{UU-200}$ under visible light irradiation (C). Conditions: [catalyst] = 10 mg, [RhB] = $3 \times 10^{-5} \text{ mg L}^{-1}$, [TCH] = 5 mg L^{-1} , $V = 100 \text{ mL}$, Light source: White light LED ($4 \times 10 \text{ W}$), [scavenger] = $2 \times 10^{-4} \text{ mg L}^{-1}$.



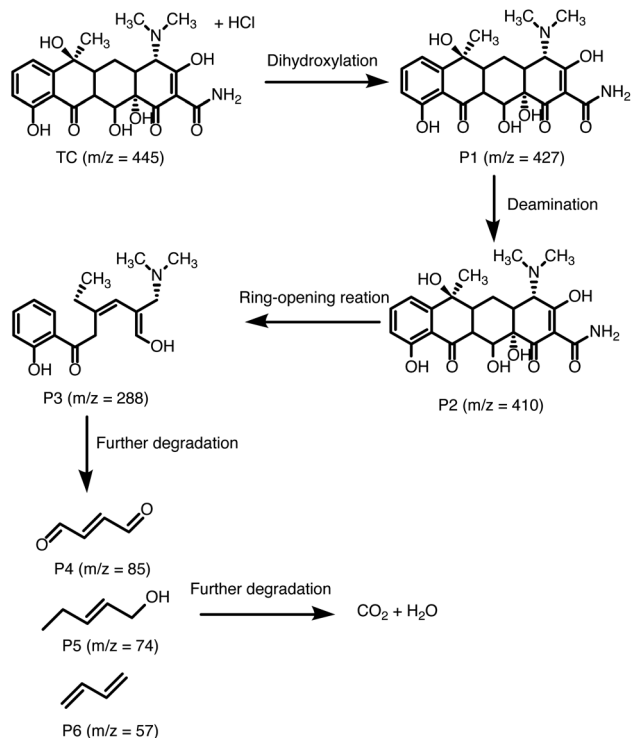


Fig. 9 Proposed TCH degradation pathways in the 40% g-C₃N₄/UU-200 system.

causing a significant loss of photocatalyst during the recovery. Moreover, the stability of the 40% g-C₃N₄/UU-200 composite upon light irradiation was investigated by performing SEM, XRD, and FT-IR analyses before and after the photoreaction. No noticeable difference was found between the XRD patterns (Fig. S8†) and FT-IR spectra (Fig. S9†) of the 40% g-C₃N₄/UU-200 catalyst measured before and after four runs, although the intensities of the corresponding peaks decreased significantly, suggesting that the structure and the functional groups remained unaltered. In addition, no obvious change in the SEM images (Fig. S10†) was detected. Taken together, these results demonstrate the good stability and recyclability of the 40% g-C₃N₄/UU-200 photocatalyst under the experimental conditions.

4 Conclusions

In summary, g-C₃N₄/UU-200 heterojunction photocatalysts were successfully produced *via* a solvothermal process. RhB and TCH were selected as model organic pollutants to investigate the photodegradation efficiency of the prepared photocatalysts under white LED light irradiation. It was found that the optimal 40% g-C₃N₄/UU-200 composite exhibited higher photodegradation efficiency for RhB (97.5%) within 90 min and for TCH (72.6%) within 180 min than bare UU-200 and g-C₃N₄. PL and EIS analyses revealed that the introduction of g-C₃N₄ on the UU-200 surface resulted in effective charge separation and enhanced photoabsorption efficiency. Therefore, the g-C₃N₄/UU-200 composite exhibited outstanding photocatalytic behavior in the degradation of RhB and TCH. The

photoinduced h⁺, e⁻, and ·O₂⁻ are the major contributors to the TCH decomposition as confirmed by a trapping test. According to these results, the present study provides new insights into the construction of heterojunction photocatalysts for organic pollutant elimination from water.

Author contributions

Trinh Duy Nguyen: writing-original draft, conceptualization, writing – review & editing, supervision. Vinh Huu Nguyen: writing-original draft, methodology, data curation, visualization. Ai Le Hoang Pham: methodology, data curation. Tuyen Van Nguyen: methodology, writing – review & editing. Taeyoon Lee: writing – review & editing, methodology, formal analysis.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 1 A. Nezamzadeh-Ejhih and E. Shahriari, *J. Ind. Eng. Chem.*, 2014, **20**, 2719–2726.
- 2 N. Arabpour and A. Nezamzadeh-Ejhih, *Mater. Sci. Semicond. Process.*, 2015, **31**, 684–692.
- 3 H. Derikvandi and A. Nezamzadeh-Ejhih, *J. Hazard. Mater.*, 2017, **321**, 629–638.
- 4 P. Mohammadyari and A. Nezamzadeh-Ejhih, *RSC Adv.*, 2015, **5**, 75300–75310.
- 5 S. A. Mirsalari, A. Nezamzadeh-Ejhih and A. R. Massah, *Environ. Sci. Pollut. Res.*, 2022, **29**, 33013–33032.
- 6 Q. Xu, S. Wageh, A. A. Al-Ghamdi and X. Li, *J. Mater. Sci. Technol.*, 2022, **124**, 171–173.
- 7 L. Zhang, J. Zhang, H. Yu and J. Yu, *Adv. Mater.*, 2022, **34**, 2107668.
- 8 H. Li, K. Wang, Y. Sun, C. T. Lollar, J. Li and H.-C. Zhou, *Mater. Today*, 2018, **21**, 108–121.
- 9 J. Tang, X. Ma, J. Yang, D.-D. Feng and X.-Q. Wang, *Dalton Trans.*, 2020, **49**, 14361–14372.
- 10 J. Yang, H. Wang, J. Liu, M. Ding, X. Xie, X. Yang, Y. Peng, S. Zhou, R. Ouyang and Y. Miao, *RSC Adv.*, 2021, **11**, 3241–3263.
- 11 A. Bavykina, N. Kolobov, I. S. Khan, J. A. Bau, A. Ramirez and J. Gascon, *Chem. Rev.*, 2020, **120**, 8468–8535.
- 12 S. Zhou, L. Lu, D. Liu, J. Wang, H. Sakiyama, M. Muddassir, A. Nezamzadeh-Ejhih and J. Liu, *CrystEngComm*, 2021, **23**, 8043–8052.
- 13 P. Sharafi-Badr, P. Hayati and G. Mahmoudi, *Micro and Nano Technologies*, Elsevier, 2022, pp. 81–103.
- 14 Q.-X. Wang and G. Li, *Inorg. Chem. Front.*, 2021, **8**, 572–589.
- 15 G. Wang, Y. Liu, B. Huang, X. Qin, X. Zhang and Y. Dai, *Dalton Trans.*, 2015, **44**, 16238–16241.



- 16 G. Wang, Q. Sun, Y. Liu, B. Huang, Y. Dai, X. Zhang and X. Qin, *Chem. - Eur. J.*, 2015, **21**, 2364–2367.
- 17 M. Savage, S. Yang, M. Suyetin, E. Bichoutskaia, W. Lewis, A. J. Blake, S. A. Barnett and M. Schröder, *Chem. - Eur. J.*, 2014, **20**, 8024–8029.
- 18 M. Feyand, M. Köppen, G. Friedrichs and N. Stock, *Chem. - Eur. J.*, 2013, **19**, 12537–12546.
- 19 A. K. Inge, M. Köppen, J. Su, M. Feyand, H. Xu, X. Zou, M. O'Keeffe and N. Stock, *J. Am. Chem. Soc.*, 2016, **138**, 1970–1976.
- 20 M. Åhlén, E. Kapaca, D. Hedbom, T. Willhammar, M. Strømme and O. Cheung, *Microporous Mesoporous Mater.*, 2022, **329**, 111548.
- 21 J. Schneider, M. Matsuoka, M. Takeuchi, J. Zhang, Y. Horiuchi, M. Anpo and D. W. Bahnemann, *Chem. Rev.*, 2014, **114**, 9919–9986.
- 22 V. H. Nguyen, A. L. H. Pham, V.-H. Nguyen, T. Lee and T. D. Nguyen, *Chem. Eng. Res. Des.*, 2022, **177**, 321–330.
- 23 D. Jiang, P. Xu, H. Wang, G. Zeng, D. Huang, M. Chen, C. Lai, C. Zhang, J. Wan and W. Xue, *Coord. Chem. Rev.*, 2018, **376**, 449–466.
- 24 S.-M. Zhou, D.-K. Ma, P. Cai, W. Chen and S.-M. Huang, *Mater. Res. Bull.*, 2014, **60**, 64–71.
- 25 Y. Deng, L. Tang, G. Zeng, Z. Zhu, M. Yan, Y. Zhou, J. Wang, Y. Liu and J. Wang, *Appl. Catal., B*, 2017, **203**, 343–354.
- 26 Y. He, J. Cai, L. Zhang, X. Wang, H. Lin, B. Teng, L. Zhao, W. Weng, H. Wan and M. Fan, *Ind. Eng. Chem. Res.*, 2014, **53**, 5905–5915.
- 27 C.-C. Wang, X.-H. Yi and P. Wang, *Appl. Catal., B*, 2019, **247**, 24–48.
- 28 J. Wen, J. Xie, Z. Yang, R. Shen, H. Li, X. Luo, X. Chen and X. Li, *ACS Sustainable Chem. Eng.*, 2017, **5**, 2224–2236.
- 29 F. Zhao, Y. Liu, S. B. Hammouda, B. Doshi, N. Guijarro, X. Min, C.-J. Tang, M. Sillanpää, K. Sivula and S. Wang, *Appl. Catal., B*, 2020, **272**, 119033.
- 30 B. Liu, Y. Wu, X. Han, J. Lv, J. Zhang and H. Shi, *J. Mater. Sci.: Mater. Electron.*, 2018, **29**, 17591–17601.
- 31 D. Guo, R. Wen, M. Liu, H. Guo, J. Chen and W. Weng, *Appl. Organomet. Chem.*, 2015, **29**, 690–697.
- 32 X. Zhang, Y. Yang, W. Huang, Y. Yang, Y. Wang, C. He, N. Liu, M. Wu and L. Tang, *Mater. Res. Bull.*, 2018, **99**, 349–358.
- 33 H. Wang, X. Yuan, Y. Wu, G. Zeng, X. Chen, L. Leng and H. Li, *Appl. Catal., B*, 2015, **174–175**, 445–454.
- 34 Z.-D. Lei, Y.-C. Xue, W.-Q. Chen, L. Li, W.-H. Qiu, Y. Zhang and L. Tang, *Small*, 2018, **14**, 1802045.
- 35 X. Li, Y. Pi, L. Wu, Q. Xia, J. Wu, Z. Li and J. Xiao, *Appl. Catal., B*, 2017, **202**, 653–663.
- 36 V. H. Nguyen, L. Van Tan, T. Lee and T. D. Nguyen, *Sustainable Chem. Pharm.*, 2021, **20**, 100385.
- 37 M. Åhlén, E. Kapaca, D. Hedbom, T. Willhammar, M. Strømme and O. Cheung, *Microporous Mesoporous Mater.*, 2022, **329**, 111548.
- 38 Y. Bai, P.-Q. Wang, J.-Y. Liu and X.-J. Liu, *RSC Adv.*, 2014, **4**, 19456–19461.
- 39 X. Liu, Y. Su, Q. Zhao, C. Du and Z. Liu, *Sci. Rep.*, 2016, **6**, 28689.
- 40 M. Solehudin, U. Sirimahachai, G. A. M. Ali, K. F. Chong and S. Wongnawa, *Adv. Powder Technol.*, 2020, **31**, 1891–1902.
- 41 K. Bhunia, M. Chandra, S. Khilari and D. Pradhan, *ACS Appl. Mater. Interfaces*, 2019, **11**, 478–488.
- 42 Y. Zhou, W. Lv, B. Zhu, F. Tong, J. Pan, J. Bai, Q. Zhou and H. Qin, *ACS Sustainable Chem. Eng.*, 2019, **7**, 5801–5807.
- 43 S. Ghattavi and A. Nezamzadeh-Ejhieh, *Int. J. Hydrogen Energy*, 2020, **45**, 24636–24656.
- 44 N. Raeisi-Kheirabadi and A. Nezamzadeh-Ejhieh, *Int. J. Hydrogen Energy*, 2020, **45**, 33381–33395.
- 45 M. Guo, Y. Wang, Q. He, W. Wang, W. Wang, Z. Fu and H. Wang, *RSC Adv.*, 2015, **5**, 58633–58639.
- 46 V. H. Nguyen, T. D. Nguyen and T. Van Nguyen, *Top. Catal.*, 2020, **63**, 1109–1120.
- 47 P. Wang, Z. Guan, Q. Li and J. Yang, *J. Mater. Sci.*, 2018, **53**, 774–786.
- 48 Q. Liang, X. Liu, J. Wang, Y. Liu, Z. Liu, L. Tang, B. Shao, W. Zhang, S. Gong, M. Cheng, Q. He and C. Feng, *J. Hazard. Mater.*, 2021, **401**, 123355.
- 49 J. Zou, Y. Yu, W. Yan, J. Meng, S. Zhang and J. Wang, *J. Mater. Sci.*, 2019, **54**, 6867–6881.
- 50 C. Liu, X. Dong, Y. Hao, X. Wang, H. Ma and X. Zhang, *New J. Chem.*, 2017, **41**, 11872–11880.
- 51 P. Wang, S. Sun, X. Zhang, X. Ge and W. Lü, *RSC Adv.*, 2016, **6**, 33589–33598.
- 52 D. Xie, Y. Ma, Y. Gu, H. Zhou, H. Zhang, G. Wang, Y. Zhang and H. Zhao, *J. Mater. Chem. A*, 2017, **5**, 23794–23804.
- 53 R. R. Solís, A. Gómez-Avilés, C. Belver, J. J. Rodriguez and J. Bedia, *J. Environ. Chem. Eng.*, 2021, **9**, 106230.
- 54 O. Elbanna, M. Fujitsuka and T. Majima, *ACS Appl. Mater. Interfaces*, 2017, **9**, 34844–34854.
- 55 K. S. W. Sing, *Pure Appl. Chem.*, 1985, **57**, 603–619.
- 56 J. Huang, Q. Tian, H. Feng, C. Xue, J. Li and Q. Xu, *Chem. Eng. J.*, 2022, **447**, 137568.
- 57 D. Pradhan and K. T. Leung, *Langmuir*, 2008, **24**, 9707–9716.
- 58 H. Derikvandi and A. Nezamzadeh-Ejhieh, *J. Photochem. Photobiol., A*, 2017, **348**, 68–78.
- 59 H. Shi, G. Chen, C. Zhang and Z. Zou, *ACS Catal.*, 2014, **4**, 3637–3643.
- 60 Z. Amani-Beni and A. Nezamzadeh-Ejhieh, *Anal. Chim. Acta*, 2018, **1031**, 47–59.
- 61 Y. Liu, F. Le Formal, F. Boudoire and N. Guijarro, *ACS Appl. Energy Mater.*, 2019, **2**, 6825–6833.
- 62 S. Samanta, S. Khilari, D. Pradhan and R. Srivastava, *ACS Sustainable Chem. Eng.*, 2017, **5**, 2562–2577.
- 63 S. A. Mirsalari and A. Nezamzadeh-Ejhieh, *Sep. Purif. Technol.*, 2020, **250**, 117235.
- 64 M. Boudart and G. Djega-Mariadassou, *Kinetics of Heterogeneous Catalytic Reactions*, Princeton University Press, 2014.
- 65 M. Mehrali-Afjani, A. Nezamzadeh-Ejhieh and H. Aghaei, *Chem. Phys. Lett.*, 2020, **759**, 137873.
- 66 A. Nezamzadeh-Ejhieh and M. Karimi-Shamsabadi, *Chem. Eng. J.*, 2013, **228**, 631–641.
- 67 H. Wang, L. Zhang, Z. Chen, J. Hu, S. Li, Z. Wang, J. Liu and X. Wang, *Chem. Soc. Rev.*, 2014, **43**, 5234–5244.



Paper

- 68 S. Ghattavi and A. Nezamzadeh-Ejehieh, *Composites, Part B*, 2020, **183**, 107712.
- 69 R. Zhao, X. Sun, Y. Jin, J. Han, L. Wang and F. Liu, *J. Mater. Sci.*, 2019, **54**, 5445–5456.
- 70 N. Omrani and A. Nezamzadeh-Ejehieh, *Sep. Purif. Technol.*, 2020, **235**, 116228.
- 71 W. Shi, S. Yang, H. Sun, J. Wang, X. Lin, F. Guo and J. Shi, *J. Mater. Sci.*, 2021, **56**, 2226–2240.
- 72 W. Wang, K. Xiao, L. Zhu, Y. Yin and Z. Wang, *RSC Adv.*, 2017, **7**, 21287–21297.

