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## Aromatic sensitizers in luminescent hybrid films†

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Atomic layer deposition offers a unique set of design possibilities due to the vast range of metal and organic precursors that can be used and combined. In this work, we have combined lanthanides with aromatic aids as strongly absorbing sensitizers to form highly luminescent thin films. Terephthalic acid is used as a base sensitizer, absorbing shorter wavelengths than 300 nm. The absorption range is extended towards the near-UV and blue range by increasing the aromatic system and adding functional groups that have strong red-shifting effects. While terbium and europium provide green and red emission, yttrium allows emission from the sensitizer itself spanning the whole color range from purple, blue and green to red. Many organic dye molecules show very high luminescence quantum yields and several of the molecules and materials investigated in this work show bright luminescence.

### Introduction

Atomic layer deposition (ALD) and its organic counterpart molecular layer deposition (MLD) is a remarkable synthesis tool capable of synthesizing a vast array of materials ranging from inorganic<sup>1</sup> including quaternary<sup>2</sup> and even quinary compounds,<sup>3</sup> nanostructures,<sup>4</sup> over to organic–inorganic hybrids,<sup>5</sup> crystalline MOF films<sup>6</sup> and pure organics.<sup>7</sup> The layer-by-layer chemistry and low deposition temperatures prevents interdiffusion between layers and enables sub-nanometer scale designs, while the vast array of deposition materials allows combinations and layered nanocomposites of different material classes. We have recently shown how combining ALD and MLD allows designing nanocomposites that combines luminescence properties of inorganic fluorides and optical absorption of organic molecules,<sup>8</sup> in addition to controlling and confining movement of excited states and thus reducing concentration quenching effects.<sup>9</sup> In addition, deposition by ALD allows on high aspect ration 3D nanostructures like photonic crystals or waveguides, and the (generally) low deposition temperature can be compatible with temperature sensitive substrate materials like textiles and polymers.

The inclusion of organic species into hybrid or nano-composite thin film is an exciting field as the organic component adds very strong optical absorption and an exceptional flexibility in designing the molecules towards specific absorption and emission ranges. Unfortunately, the choice of organic components in MLD is quite limited. The need to get these molecules into the vapour phase while also having sufficient

reactivity reduces the practical range to fairly small molecules. This puts strong limits on the range of optical and electronic properties to choose from.

Reviews on ALD and MLD chemistries<sup>7,10</sup> reveals that the organic precursors utilized so far are quite limited beyond a single aromatic ring or a short conjugated chain. There are examples of larger conjugated systems like naphthalene and biphenyl<sup>6</sup> showing near-UV activity and even visible absorption.<sup>11</sup> These optically absorbing dye precursors require sublimation temperatures of around 250 °C or more, approaching the temperature where such organic molecules start to decompose or polymerize. This illustrates well the difficulty in balancing Vis and NIR active optical properties with ALD/MLD compatibility. An exception to this is our recent work on quinizarin molecules, which sublimates efficiently at 130 °C and produces deep pink films with TMA as cation precursor.<sup>12</sup>

In this work, we explore a range of mono- and bi-aromatic dicarboxylic acids deposited with lanthanides as cations, as shown in Fig. 1. All three lanthanides have very similar deposition chemistries.<sup>13</sup> Terbium and europium was chosen due to their well-known and easily identified sharp emission peaks which are easily distinguishable from organic molecule emissions. In addition, these two lanthanides have large energy gaps from the emitting energy level(s) to the next lower energy level. This allows them to be highly luminescent even in high-phonon matrices such as hybrid materials. Yttrium was chosen as an optically inactive ion that allows characterization of the organic molecules emission properties in the same lanthanide-organic hybrid material environment. In previous works, we have investigated the growth and material chemistry of optically active hybrid materials.<sup>5,6,12,14</sup> Here, we focus on the hybrid materials optical properties, absorption and luminescence colour.

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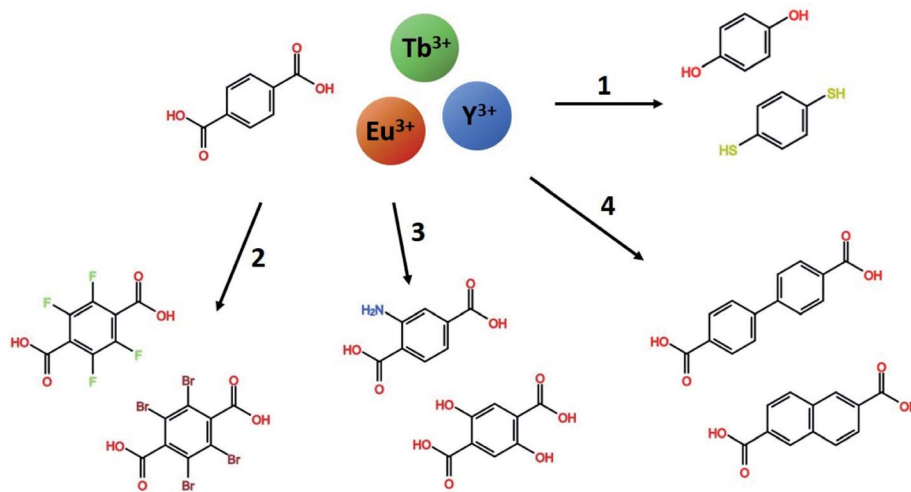


Fig. 1 The range of mono- and bi-aromatic di-carboxylic acids investigated in this work, combined with the three lanthanides Tb, Eu and Y. Using terephthalic acid and Tb as a base system, four paths of molecule variations are explored. (1) Other binding group in H<sub>2</sub>bdo and H<sub>2</sub>bdt, (2) halogenated H<sub>2</sub>bdc-F and H<sub>2</sub>bdc-Br, (3) functionalized H<sub>2</sub>bdc-NH<sub>2</sub> and H<sub>2</sub>bdc-2OH, and (4) the larger conjugated system H<sub>2</sub>bpdc and H<sub>2</sub>ndc.

## Experimental

The films were deposited in an F-120 research-type ALD-reactor (ASM Microchemistry Ltd). The deposition temperature were 250 °C except when otherwise noted. Ln-β-diketonate Ln(thd)<sub>3</sub> (Strem Chemicals, >99.9%, Ln = Y, Eu, Tb) were used as lanthanide precursors, while a range of different organic precursors were used as optical sensitizers. The full names, abbreviations used in this work and the sublimation temperatures used for all precursors are listed in Table 1. Nitrogen was used as carrier and purge gas, supplied from gas bottles (AGA, 99.999%). p-Type Si(100) substrates were used for all depositions. In addition, 0.5 × 4 cm<sup>2</sup> Si(100) strips were placed some 8 cm apart in the gas inlet and exhaust sides of the deposition chamber to monitor thickness gradients. Silicon substrates were used with native oxide intact and its thickness was measured before each deposition. All substrates were dry wiped and dust was removed using pressurized air. Pulse times for Ln(thd)<sub>3</sub> and organics were 1.5 and 2 seconds respectively. Purge times were between 1 and 2 seconds. Exceptions to the

standard parameters are that H<sub>2</sub>bpdc was deposited at 275 °C due to its high precursor temperature, H<sub>2</sub>bdt was deposited at 200, 250 and 300 °C, and H<sub>2</sub>bto depositions used 3 second lanthanide and organic pulses. In addition, Tb<sub>2</sub>bdc<sub>3</sub> was used as a basis material and was deposited in the 225–350 °C range in steps of 25 °C.

*In situ* quartz crystal microbalance (QCM) analyses were conducted using a 6 MHz AT-cut quartz crystal. The crystal was mounted in a home-made holder and was used to monitor the mass increase, proportional to the change in frequency during the deposition to determine saturation conditions for pulse and purge parameters. The signal was recorded using a Colnatec Eon-LT and processed by averaging over 16 consecutive ALD cycles. The temperature was stabilized for 60 to 90 minutes before any experiments were conducted to ensure a stable temperature and response from the QCM-crystals. Note that the QCM investigations were conducted over an extended time-frame where modifications to the setup have occurred. Thus, the scale of the frequency changes is not necessarily comparable between different materials in this work.

Table 1 Precursor names and sublimation temperatures

	Compound name	Short name	T <sub>sub</sub> (°C)
	Ln(2,2,6,6-tetramethyl-3,5-heptanedione) <sub>3</sub>	Y(thd) <sub>3</sub>	130
		Eu(thd) <sub>3</sub>	150
		Tb(thd) <sub>3</sub>	145
(1)	1,4-Benzenedicarboxylic acid	H <sub>2</sub> bdc	200
	2,3,5,6-Tetrafluoro-1,4-benzenedicarboxylic acid	H <sub>2</sub> bdc-F	145
	2,3,5,6-Tetrabromo-1,4-benzenedicarboxylic acid	H <sub>2</sub> bdc-Br	230
(2)	2-Amino-1,4-benzenedicarboxylic acid	H <sub>2</sub> bdc-NH <sub>2</sub>	210
	2,5-Dihydroxy-1,4-benzenedicarboxylic acid	H <sub>2</sub> bdc-2OH	225
(3)	Biphenyl-4,4'-dicarboxylic acid	H <sub>2</sub> bpdc	250
	2,6-Naphthalenedicarboxylic acid	H <sub>2</sub> ndc	225
(4)	Benzenediol	H <sub>2</sub> bdo	130
	Benzenedithiol	H <sub>2</sub> bdt	70



Film thickness and refractive index  $n(\lambda)$  were determined with a J. A. Woollam alpha-SE spectroscopic ellipsometer in the 380–900 nm range. This ellipsometry data was modelled using a Cauchy model. A VASE spectroscopic ellipsometer with range of 280–1000 nm, also from J. A. Woollam, was used to model  $n(\lambda)$  and the extinction coefficient  $k(\lambda)$ . These films were modelled with a general oscillator model consisting of several Gaussian peaks. Silicon samples were used for ellipsometry. In both instruments, the sample is modelled by a single layer on top of a thin native silicon oxide layer. This native oxide was measured before each deposition and was generally in the 2–4 nm range. Interface and surface roughnesses were not included in the model as these did not improve the fit. The relation between absorption coefficient  $\alpha(\lambda)$  and  $k(\lambda)$  is given in eqn (1). For readers more accustomed to  $\alpha(\lambda)$ , the practical difference between these is a constant scaling factor ( $2\pi$ ) a simple wavelength dependent factor ( $1/\lambda$ ). FTIR characterization was done using a Bruker VERTEX 70 FTIR spectrometer in transmission mode on single-side polished silicon substrates. Luminescence (PL) and excitation (PLE) measurements was done using an Edinburgh Instruments FLS920 fluorescence spectrometer with a 450 W Xe lamp as excitation source and a Hamamatsu R928 PMT for detection. PL decay measurements utilized an optical parametric oscillator (OPO) system (Opotek HE 355 II) pumped by the third harmonic of a Nd:YAG laser as excitation source. The OPO system was set to 355 nm and a repetition rate of 20 Hz. The detector for decay measurements was the same equipment used for the PLE.

$$\alpha(\lambda) = \frac{2\pi}{\lambda} k \quad (1)$$

## Results and discussion

This work uses  $\text{Tb}_2\text{bdc}_3$  as basis material, modifying the luminescence and absorption properties by replacing  $\text{Tb}^{3+}$  with other lanthanides and bdc with other aromatic molecules. Thus, the growth and properties of  $\text{Tb}_2\text{bdc}_3$  was investigated in more detail than the other materials, while the other materials focus on their differences compared to  $\text{Tb}_2\text{bdc}_3$ .

### $\text{Tb}_2\text{bdc}_3$ – growth and optical properties

The saturation conditions for  $\text{Tb}_2\text{bdc}_3$  based on QCM experiments is given in Fig. 2. Both lanthanide and acid precursor show clear saturation at around 2 seconds pulses. Purge durations have little effect apart from a gradual growth rate reduction upon long acid purges.

The QCM response during a cycle with long pulse and purge times is shown in Fig. 3. The two ideal half cycle reactions are given below. For comparison, the masses of the precursors and leaving groups are:  $\text{Tb}(\text{thd})_3 = 709$ ,  $\text{H}_2\text{bdc} = 166$ ,  $\text{Hthd} = 184$   $\text{g mol}^{-1}$ . Both the lanthanide and acid pulses show initial rapid frequency change of around 2 seconds followed by a slower evolution. The mass decrease during the acid pulse is due to  $\text{H}_2\text{bdc}$  being slightly lighter than the leaving molecule  $\text{Hthd}$ . The ratio between the ideal reactions are  $\Delta m_1/\Delta m_2 = 16$ , while

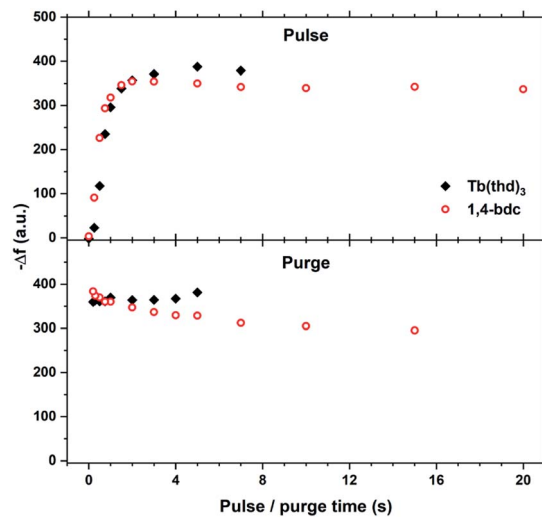
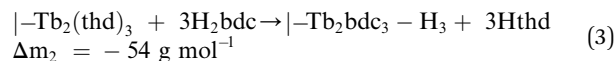
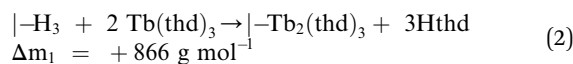


Fig. 2 QCM characterization of growth saturation as a function pulse durations for  $\text{Tb}_2\text{bdc}_3$ .

the measured  $\Delta m_{\text{Tb}}/\Delta m_{\text{bdc}} = 3.5$ . This indicates that either less lanthanide or more acid is deposited during each half cycle than expected from the ideal reactions.



The growth rate as a function of deposition temperature is shown in Fig. 4. The growth decreases linearly with temperature up until 300 °C, and becomes almost stable up to 350 °C. The films above 300 °C had a brown-ish colour and was non-luminescent indicating thermal decomposition of either  $\text{H}_2\text{bdc}$  or  $\text{Hthd}$ , or both. The luminescence intensity of the films deposited in the 225–275 °C range did not vary significantly. Thus it was decided to keep the deposition temperature at 250 °C for all films, unless otherwise stated.

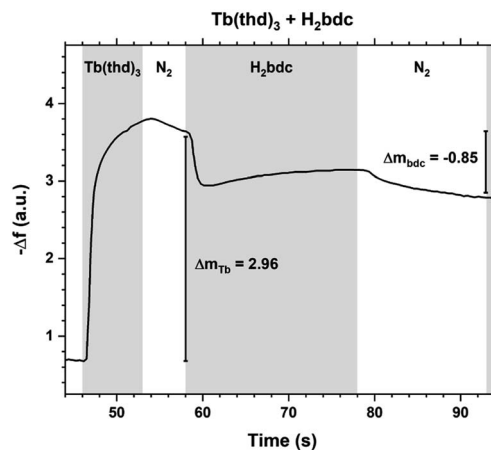


Fig. 3 QCM response during a  $\text{Tb}_2\text{bdc}_3$  cycle with long pulse and purge times.  $\text{Tb}(\text{thd})_3/\text{purge}/1,4\text{-bdc}/\text{purge} = 7/5/20/15$ .



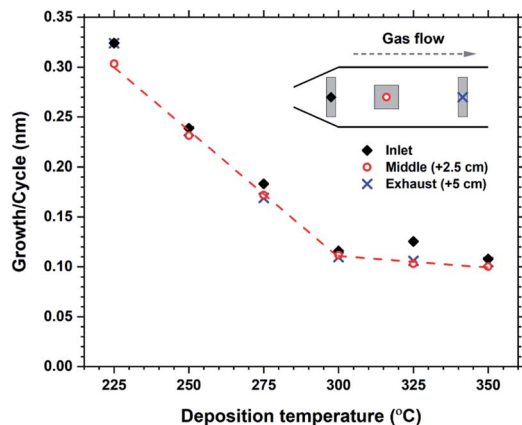


Fig. 4 Growth rate as a function of deposition temperature. Measurements from 3 different locations over a 5 cm distance indicates the evenness at each temperature.

Some hybrid materials are not long-term stable in air, either through absorption and release of for example moisture or oxidation reactions.  $Tb_2bdc_3$  films were characterized by ellipsometry immediately after deposition and 4 additional times up to 3 days to monitor any changes to the films thickness. The results are shown in Fig. 5. The films show a swelling of about 4% after 3 days with no other apparent changes to the films structure or properties. In fact, films deposited over 4 years ago show identical optical and luminescent properties as freshly deposited films.

Fig. 6 shows FTIR spectra from a  $Tb_2bdc_3$  film taken during the day of deposition and 16 days later. The spectra show that there are no significant differences in number, position or intensities of peaks, indicating no change in bonding during this time. The films show no sign of protonated carboxylic acid groups, which should be clearly visible around  $3000\text{ cm}^{-1}$ . FTIR is useful to investigate the type of bonding between the metal and acid group based on the splitting ( $\Delta_{acid}$ ) between the asymmetric and symmetric carboxylate stretching peaks.<sup>15</sup> A  $\Delta_{acid}$  in the  $50\text{--}150\text{ cm}^{-1}$  range indicate bidentate binding,  $130\text{--}200\text{ cm}^{-1}$  indicate bridging and  $>200\text{ cm}^{-1}$  indicate unidentate

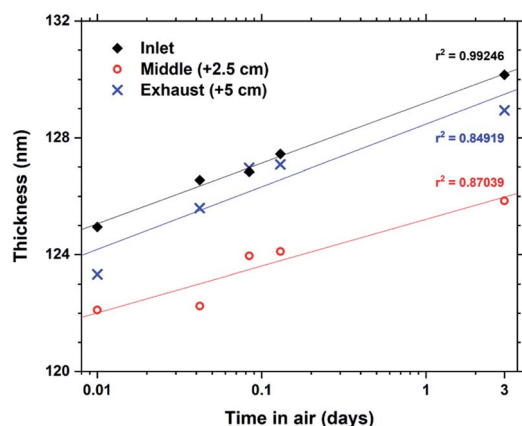


Fig. 5 Evolution of film thickness upon air exposure for  $Tb_2bdc_3$  films.

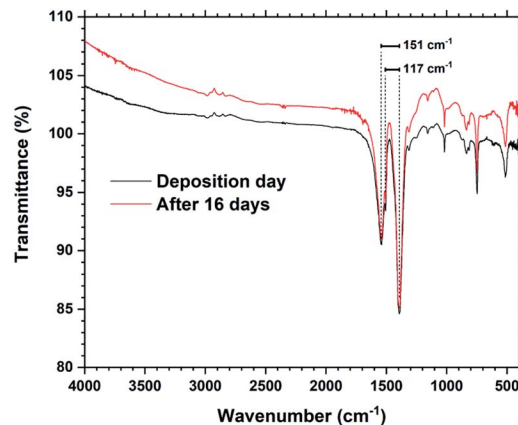


Fig. 6 FTIR spectra of a  $Tb_2bdc_3$  film taken on the day of deposition and after 16 days. The splitting between the asymmetric and symmetric peaks are shown.

binding. The asymmetric carboxylic stretching peak is also split into a double peak, giving two values for  $\Delta_{acid}$ , 151 and  $117\text{ cm}^{-1}$ . This indicates that the  $Tb_2bdc_3$  films contains a mixture of bidentate and bridging binding. An enlarged version zoomed in on these peaks is shown in ESI Fig. 1,<sup>†</sup> where also the exact wavenumber position of these three peaks are numbered.

The optical absorption and luminescence properties of  $Tb_2bdc_3$  is shown in Fig. 7, alongside a photograph of a sample under a 280 nm diode. Transmittance data is obtained with a 28 nm film on silica substrate, where the multiplet absorption from the benzene ring of terephthalic acid is clearly seen at around 250 and 300 nm. The PLE spectrum is obtained by monitoring the  $^5D_4 \rightarrow ^7F_5$  emission. The two first aromatic absorption peaks are clearly seen in the PLE spectrum. The spectrum is not corrected for the decrease in the Xe lamps light intensity towards deep UV, which causes the apparent red-shift of the high energy peak.

Fig. 7 forms the basis for exploring the other organic sensitizers and lanthanides with the aim to obtain absorption towards

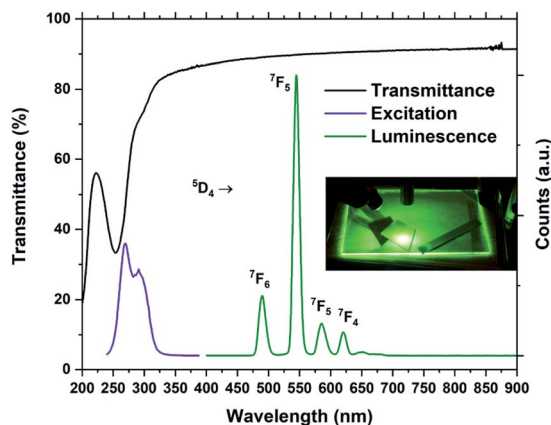


Fig. 7 Transmittance, PL and PLE spectra of  $Tb_2bdc_3$ . The inset shows a glass plate and silicon substrates coated with  $Tb_2bdc_3$  under 280 nm illumination from a diode.



Table 2 Summary of growth rates with Tb(thd)<sub>3</sub>

	Organic	$T_{\text{dep}}$ (°C)	$g$ (nm/cycle)	Parameters (s)
	bdc	250	0.23	1.5/1.5/2/1
(1)	bdo	175	0.12	3/1/3/1
	bdt	250	0	1.5/1.5/2/1
(2)	bdc-F	250	0.24	1.5/1.5/2/1
	bdc-Br	250	0.27	1.5/1.5/2/1
(3)	bdc-NH <sub>2</sub>	250	0.27	1.5/1.5/2/1
	bdc-2OH	250	~1 <sup>a</sup>	1.5/1.5/2/2
(4)	bpdc	275	0.26	1.5/1/2/1
	ndc	250	0.27	1.5/1.5/2/2

<sup>a</sup> All films showed strong gradients.

the near-UV and blue range and different emission colours. Each of the 4 groups of molecules illustrated in Fig. 1 is compared to Tb<sub>2</sub>bdc<sub>3</sub> individually below, while the deposition parameters and obtained growth rates are summarized in Table 2.

### 1,4-Benzenediol and 1,4-benzenedithiol

The aim of these two R-OH and R-SH molecules is to investigate if other binding groups that carboxylic acid can be used to obtain hybrid films with Ln(thd)<sub>3</sub>, in addition to how these bonds affect the materials optical properties. In particular the sulfur atom can participate in low energy charge transfer transitions and lower the d-shell and create low energy f → d transitions with lanthanides.<sup>16</sup>

Films deposited with Tb(thd)<sub>3</sub> and H<sub>2</sub>bdo showed a growth rate of 0.13 nm per cycle, which is about half of the growth rate of Tb<sub>2</sub>bdc<sub>3</sub>. The films were visually transparent like Tb<sub>2</sub>bdc<sub>3</sub>, however these films did not luminesce under a 254 nm UV lamp. As enhancing the optical and luminescent properties of lanthanide-hybrid materials is the aim of this work, the H<sub>2</sub>bdo precursor was not further explored apart from noting that H<sub>2</sub>bdo and Tb(thd)<sub>3</sub> do react and form ALD films. Thus, hydroxy groups can be suitable reaction groups to bind to Ln(thd)<sub>3</sub>. In fact, H<sub>2</sub>bdo sublimates at 70 °C lower than H<sub>2</sub>bdc, which can be a desired property.

Tb(thd)<sub>3</sub> and H<sub>2</sub>bdt did not result in film deposition. Film growth was attempted at 200, 250 and 300 °C. This was somewhat surprising as H<sub>2</sub>S can be used with Ln(thd)<sub>3</sub> to produce lanthanide sulfide films.<sup>17</sup>

### Tb<sub>2</sub>bdc<sub>3</sub>: Ln<sup>3+</sup> and halogenated bdc

Halogenation of the aromatic centre of H<sub>2</sub>bdc can affect both the growth and optical properties in addition the physical properties of the precursor. From Table 1 it is seen that H<sub>2</sub>bdc-F and H<sub>2</sub>bdc-Br sublimates at 70 °C lower and 30 °C higher than H<sub>2</sub>bdc, respectively. Table 2 show that both precursors result in similar growth rates as H<sub>2</sub>bdc. Regarding growth, the main difference between H<sub>2</sub>bdc, H<sub>2</sub>bdc-F and H<sub>2</sub>bdc-Br is a substantial difference in sublimation temperature.

Films produced with these three precursors and Tb(thd)<sub>3</sub> all produced luminescent with the same green colour and with similar intensity under a 254 nm UV lamp. To further investigate differences in optical properties, the films optical

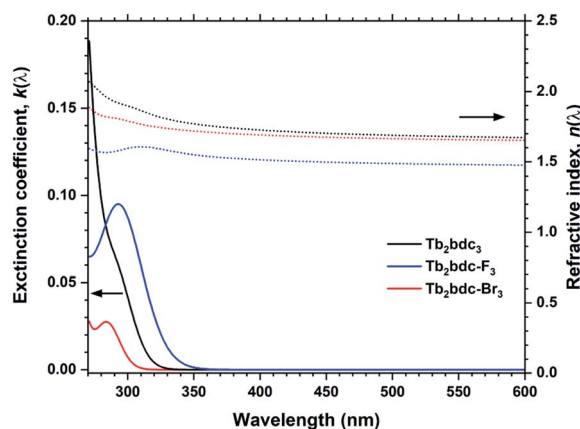


Fig. 8  $k(\lambda)$  and  $n(\lambda)$  for Tb<sub>2</sub>bdc<sub>3</sub>, Tb<sub>2</sub>bdc-F<sub>3</sub> and Tb<sub>2</sub>bdc-Br<sub>3</sub> films.

absorption and luminescence decays were characterized.  $k(\lambda)$  and  $n(\lambda)$  of these films is shown in Fig. 8. Compared to Tb<sub>2</sub>bdc<sub>3</sub>, fluorination causes a red-shift in absorption and a lower  $n(\lambda)$  while bromination causes a blue-shift with unchanged  $n(\lambda)$ . Fig. 9 show decays of the <sup>5</sup>D<sub>4</sub> → <sup>7</sup>F<sub>5</sub> emissions of the three materials. All three can be fitted with a single exponential model. The brominated sample show a lower lifetime than Tb<sub>2</sub>bdc<sub>3</sub> and Tb<sub>2</sub>bdc-F<sub>3</sub>. This could indicate higher quenching rates in Tb<sub>2</sub>bdc-Br<sub>3</sub>. In summary, bromination offers neither improved growth (high sublimation temperature), more near-UV absorption (weaker and blue-shifted) nor improved luminescence yield (lower Tb<sup>3+</sup> lifetime). On the other hand, the main advantage of fluorination is the substantially lower sublimation temperature and a small red-shift of the absorption while the growth and material properties are otherwise very similar to the un-halogenated acid.

Replacing Tb<sup>3+</sup> with Eu<sup>3+</sup> and Y<sup>3+</sup> results in Ln<sub>2</sub>bdc<sub>3</sub> films with very different emissions, shown in Fig. 10. The PLE spectra are very similar to each other, although the lighter Y<sup>3+</sup> shows a small red-shift compared to the heavier Tb<sup>3+</sup> and Eu<sup>3+</sup>. Both Tb<sub>2</sub>bdc<sub>3</sub> and Eu<sub>2</sub>bdc<sub>3</sub> produce their regular Tb<sup>3+</sup> and Eu<sup>3+</sup> emission spectra, showing that bdc is an efficient sensitizer for these two lanthanides. Y<sup>3+</sup> doesn't have a partially filled d or f shell, allowing the sensitizer molecule to emit its own photon.

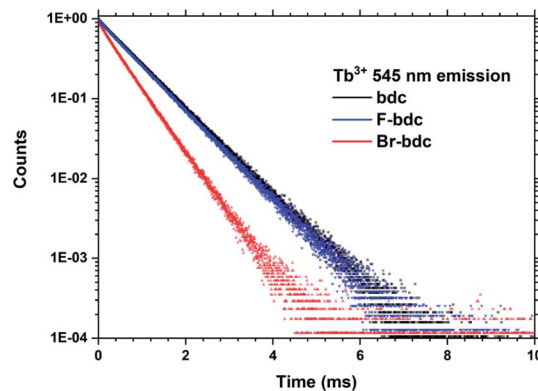


Fig. 9 PL decay data from Tb<sub>2</sub>bdc<sub>3</sub>, Tb<sub>2</sub>bdc-F<sub>3</sub> and Tb<sub>2</sub>bdc-Br<sub>3</sub> films.



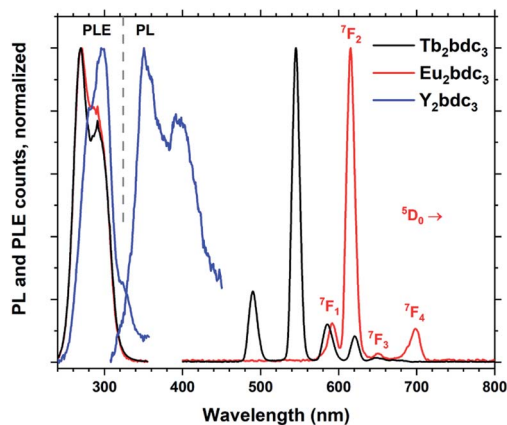


Fig. 10 Normalized PLE and PL spectra of  $\text{Ln}_2\text{bdc}_3$ ,  $\text{Ln} = \text{Tb}, \text{Eu}$  and  $\text{Y}$ . The emission lines of  $\text{Eu}^{3+}$  are named.

$\text{Y}_2\text{bdc}_3$  emits in the near-UV, partly extending into the blue range. This gives the material a violet emission colour, relatively weak compared to  $\text{Tb}_2\text{bdc}_3$  and  $\text{Eu}_2\text{bdc}_3$ .

### $\text{H}_2\text{bdc-NH}_2$ and $\text{H}_2\text{bdc-OH}$

These two molecules are interesting from both a chemical and optical perspective. In addition to the two carboxylic acid groups, they have one amine and two hydroxy groups, respectively, which can take part in bonding. Depositions with  $\text{Tb}(\text{thd})_3$  and  $\text{H}_2\text{bdc-NH}_2$  does indeed show a somewhat higher growth rate, while  $\text{H}_2\text{bdc-OH}$  resulted a fourfold increase in growth rate and strong gradients throughout the gradients. The large increase in growth rate and gradients were reproducible and varying deposition parameters did not result in more homogeneous films. This indicates that the growth chemistry is more complex than for the other sensitizer acids, in particular for  $\text{H}_2\text{bdc-OH}$ .

These films show large red-shifts compared to  $\text{Tb}_2\text{bdc}_3$ , covering the near-UV and even blue range for  $\text{Tb}_2(\text{bdc-OH})_3$  (Fig. 11). However, all depositions with  $\text{Tb}, \text{Eu}$  and  $\text{Y}$  with these two acids produced non-luminescent samples. The fact that the  $\text{Y}$  depositions were quenched indicate that the sensitizer acids are

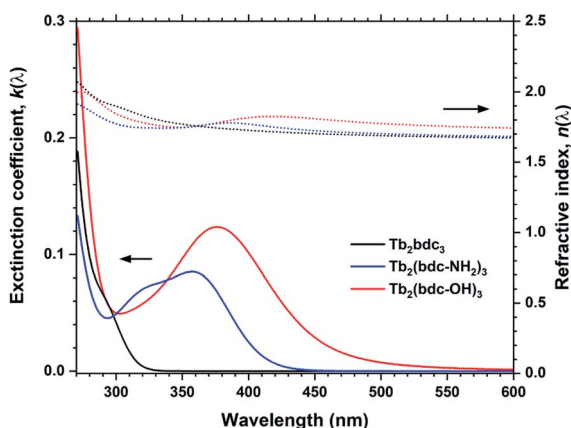


Fig. 11  $k(\lambda)$  and  $n(\lambda)$  for  $\text{Tb}_2\text{bdc}_3$ ,  $\text{Tb}_2\text{bdc-NH}_2$  and  $\text{Tb}_2\text{bdc-OH}$  films.

quenched due to how they are packed in the structure. Both acids produce strongly luminescent solutions when dissolved. Crystalline MOF structures with  $\text{H}_2\text{bdc-NH}_2$  and  $\text{Tb}$  and  $\text{Eu}$  has been shown to produce sensitized  $\text{Ln}^{3+}$  emission.<sup>18</sup> A study by Vogel *et al.* on MOF structures of  $\text{Y}$  and  $\text{H}_2\text{bdc-OH}$  show that the exact bonding scheme between  $\text{Ln}^{3+}$  and the sensitizer acid will have major impacts on the materials optical and luminescent properties.<sup>19</sup> Thus, although the films produced in our work are non-luminescent, it is likely that control of the resulting structure and molecule packing can produce luminescent films. We have previously shown that control of crystallinity can be achieved both through *in situ* modulation by pulsing other species to modify growth mechanics and ex-situ treatments.<sup>20,21</sup> However, exploration of how crystallinity and control of growth mechanics can enable luminescent variants is beyond the scope of the current work.

### $\text{H}_2\text{bpdc}$ and $\text{H}_2\text{ndc}$

These two biaromatic acids are simpler in terms of not having other functional groups that affect growth. They both produce homogeneous and gradient free films with  $\text{Y}, \text{Tb}$  and  $\text{Eu}$ , with a slightly larger growth rate than  $\text{Tb}_2\text{bdc}_3$  due to their larger size. Interestingly,  $\text{Y}_2\text{pbdc}_3$  and  $\text{Y}_2\text{ndc}_3$  films show bright violet and blue luminescence from the sensitizer molecule, while absorption covers the near UV range with a tail extending until 400 nm. Thus, due to their good optical properties and even

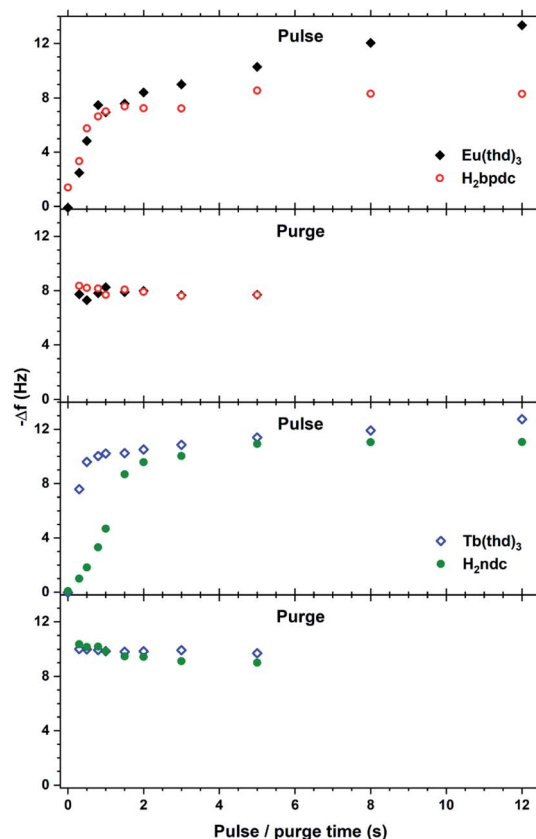


Fig. 12 QCM characterization of growth saturation as a function pulse durations for  $\text{Eu}_2\text{bpdc}_3$  and  $\text{Tb}_2\text{ndc}_3$ .



growth, these two sensitizer acids were investigated in more detail.

The QCM response per cycle  $\text{Eu}_2\text{bpdc}_3$  and  $\text{Tb}_2\text{ndc}_3$  is shown in Fig. 12.  $\text{Tb}_2\text{ndc}_3$  show clear saturation after 1 and 1.5 s  $\text{Tb}(\text{thd})_3$  and  $\text{H}_2\text{ndc}$  pulses, respectively.  $\text{Eu}_2\text{bpdc}_3$  show strong indication of the same saturation values. However,  $\text{Eu}_2\text{bpdc}_3$  also show a steady increase with increasing  $\text{Eu}(\text{thd})_3$  pulse times longer than 1 s, and some growth at 0 s acid pulses. This is likely due to the higher precursor temperature used for  $\text{H}_2\text{bpdc}$ . We based the sublimation temperatures on prior sublimations temperatures in ALD-like vacuum conditions. Likely, 250 °C is slightly too warm for this precursor, leading to some precursor vapours seeping into the reaction chamber. This CVD component is quite small for the 2 and 3 s  $\text{Ln}(\text{thd})_3$  and  $\text{H}_2\text{bpdc}$  pulses used for samples, which was also seen in negligible growth gradients.

The QCM response during a single cycle of  $\text{Eu}_2\text{bpdc}_3$  and  $\text{Tb}_2\text{ndc}_3$  using long pulse and purge times is shown in Fig. 13. As for  $\text{Tb}_2\text{bdc}_3$  in Fig. 3, both  $\text{Eu}_2\text{bpdc}_3$  and  $\text{Tb}_2\text{ndc}_3$  show an initial rapid growth for both  $\text{Ln}(\text{thd})_3$  and acid pulse, indicating surface saturation. Contrary to  $\text{Tb}_2\text{bdc}_3$  which experience a mass decrease during the acid pulse, these two depositions experience a mass gain. This is due to the large  $\text{H}_2\text{bpdc}$  and  $\text{H}_2\text{ndc}$  being heavier than  $\text{Hthd}$  while  $\text{H}_2\text{bdc}$  is lighter, 242 and 216  $\text{g mol}^{-1}$  respectively.

The perhaps most interesting aspects of  $\text{Ln}_2\text{bpdc}_3$  and  $\text{Ln}_2\text{ndc}_3$  films are their optical properties. These films have absorptions covering most of near UV while being almost an order of magnitude stronger than  $\text{Ln}_2\text{bdc}_3$ ,  $\text{Ln}_2(\text{bdc-NH}_2)_3$  and  $\text{Ln}_2(\text{bdc-OH})_3$ . Unlike  $\text{Ln}_2(\text{bdc-NH}_2)_3$  and  $\text{Ln}_2(\text{bdc-OH})_3$ , these films also result in highly luminescent films. Fig. 14 summarizes  $k(\lambda)$ ,  $n(\lambda)$ , PLE and optical transmittance of selected films, in addition to the PL spectra originating from the aromatic acids.

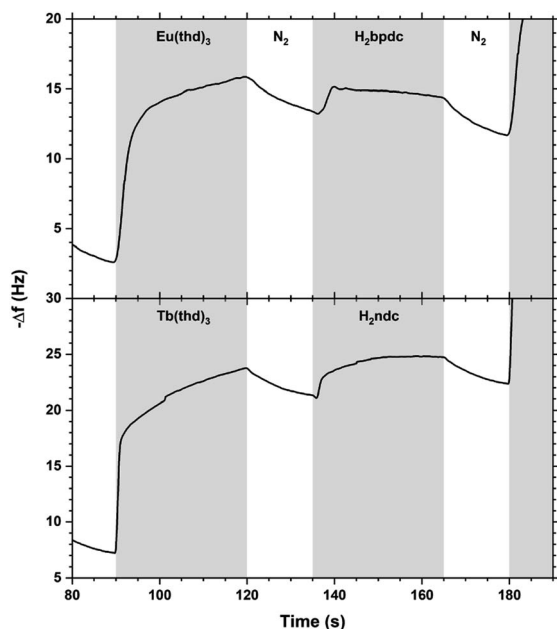


Fig. 13 QCM response during a  $\text{Eu}_2\text{bpdc}_3$  and  $\text{Tb}_2\text{ndc}_3$  cycle with long pulse and purge times.  $\text{Ln}(\text{thd})_3/\text{purge}/\text{acid}/\text{purge} = 30/15/30/15$ .

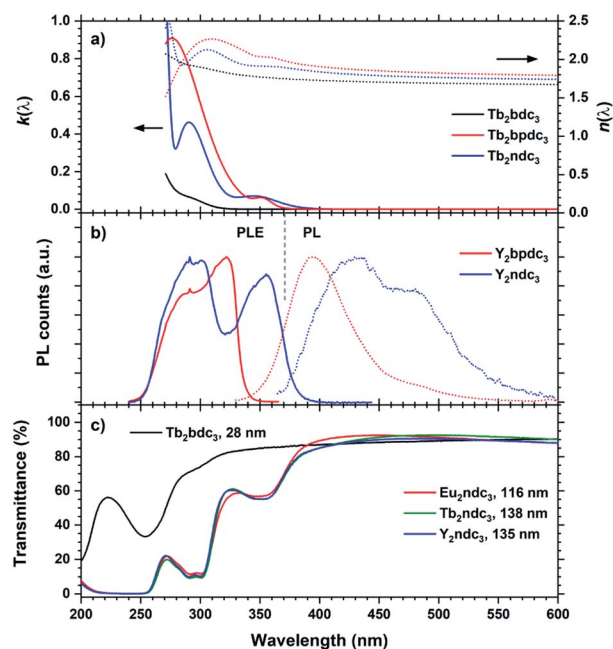


Fig. 14 (a)  $k(\lambda)$  and  $n(\lambda)$  of  $\text{Tb}_2\text{bpdc}_3$  and  $\text{Tb}_2\text{ndc}_3$ .  $\text{Tb}_2\text{bdc}_3$  is included for comparison. (b) PLE and PL of  $\text{Y}_2\text{bpdc}_3$  and  $\text{Y}_2\text{ndc}_3$ .  $\lambda_{\text{ex}}/\lambda_{\text{em}} = 320/420$  nm and  $355/475$  nm for  $\text{Y}_2\text{bpdc}_3$  and  $\text{Y}_2\text{ndc}_3$ , respectively. (c) Transmittance spectra of a selection of thin films from this work, illustrating the strong and near-UV absorption from the organic molecules.

$\text{Y}_2\text{bpdc}_3$  and  $\text{Y}_2\text{ndc}_3$  show strong violet and blue emissions. Their emission spectra are shown in Fig. 14(b). Both consist of broad emission peaks which are normal for aromatic emission. The emission from  $\text{Y}_2\text{ndc}_3$  consist of two distinguishable but overlapping peaks at 430 and 480 nm, while  $\text{Y}_2\text{bpdc}_3$  consist of one major peak at 395 nm and a smaller shoulder at 475 nm. Replacing Y with Tb or Eu fully removed the aromatic emissions, indicating complete energy transfer to these lanthanides in  $\text{Ln}_2\text{bpdc}_3$  and  $\text{Ln}_2\text{ndc}_3$ . However, while both Tb and Eu were luminescent in  $\text{Ln}_2\text{bpdc}_3$ , only  $\text{Eu}^{3+}$  was luminescent in  $\text{Ln}_2\text{ndc}_3$ .  $\text{Tb}_2\text{ndc}_3$  showed neither aromatic nor Tb emission. The lack of emission from  $\text{bpdc}^{2-}$  and  $\text{ndc}^{2-}$  means that there is a complete transfer from these two to something else. ESI Fig. 3† show the emission spectra of  $\text{Y}_2\text{bpdc}_3$  and  $\text{Y}_2\text{ndc}_3$  along with the ground state f-f absorption transitions of  $\text{Tb}^{3+}$  and  $\text{Eu}^{3+}$ . Both lanthanides show overlapping absorption lines with the emission of both sensitizers. Note that for wavelengths shorter than 400 nm, both lanthanides have dense absorption lines that are excluded in the figure for clarity. The lack of  $\text{Tb}^{3+}$  emission in  $\text{Tb}_2\text{ndc}_3$  indicated that either  $\text{Tb}^{3+}$  is fully quenched or that  $\text{Tb}^{3+}$  never received the energy. In both cases, a possible explanation is a low energy intervalence charge transfer (IVCT) state that can accept the energy from either the sensitizer, lanthanide or both. This IVCT and its quenching on  $\text{Tb}^{3+}$  is well known in  $d^0$  metal oxides.<sup>22</sup> Both  $\text{Tb}^{3+}$  and  $\text{Eu}^{3+}$  are redox active (to  $\text{Tb}^{4+}$  and  $\text{Eu}^{2+}$ , respectively), and so is polyaromatic molecules.<sup>23</sup> However, a detailed investigation into the quenching mechanisms is beyond the scope of this work.

The Tb and Eu emission spectra were identical to those from  $\text{Ln}_2\text{bdc}_3$  in Fig. 10. In addition, the PLE spectra of Tb and Eu

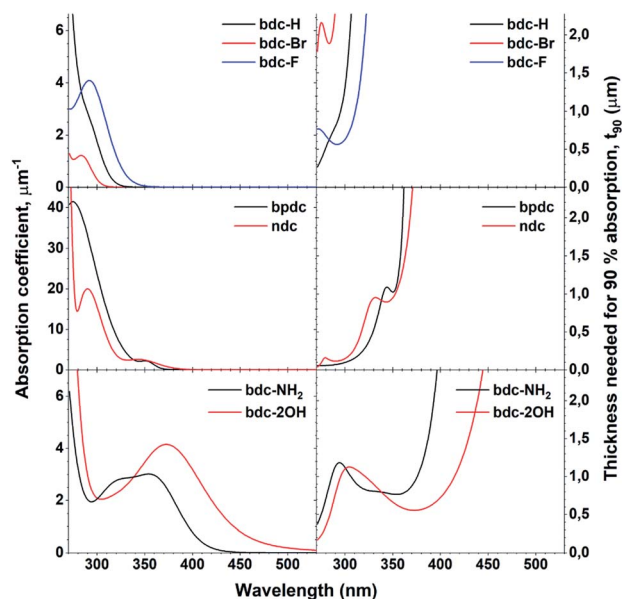


Fig. 15 Absorption coefficient of Tb films with all acids, obtained through ellipsometry and the calculated thickness needed for 90% absorption,  $t_{90}$ . Note the different y-scale on the absorption coefficients for bpdc and ndc.

were identical to the PLE spectra of the aromatic emission in  $Y_2\text{bpdc}_3$  and  $Y_2\text{ndc}_3$ , thus only the PL and PLE spectra of the Y variants are shown in Fig. 14. The modelled  $k(\lambda)$  and  $n(\lambda)$  were

also identical between lanthanides. Comparing the modelled  $k(\lambda)$  in (a) and PLE in (b) show that the absorption and excitation peaks matches well for  $\text{Ln}_2\text{ndc}_3$ . However, even though both  $\text{Ln}_2\text{bpdc}_3$  and  $\text{Ln}_2\text{ndc}_3$  show similar absorption at 350 nm, this absorption only produces an excitation band in  $\text{Ln}_2\text{ndc}_3$ .

In (c) the transmittance spectra on all three  $\text{Ln}_2\text{ndc}_3$  materials are shown alongside  $\text{Tb}_2\text{bdc}_3$  for comparison. All samples are on silica substrates. The absorption bands 350 and 290 nm match well with the modelled  $k(\lambda)$  and PLE spectre.

## Summary of absorption properties

One of the defining properties of these mono- and bi-aromatic hybrid films is their strong near-UV absorption. Fig. 15 summarizes the absorption coefficient for all terbium films (note the different y-scale for the two bi-aromatic films). To help visualize how  $k(\lambda)$  and absorption coefficients manifests as absorption in thin films, the film thickness required for 90% absorption according to Beer-Lambert law is plotted as  $t_{90}$ . Plotting  $t_{90}$  helps understanding how thick films are required for a given absorption. For example, for 90% absorption at 300 nm, films of 100 and 150 nm are sufficient for  $\text{Ln}_2\text{bpdc}_3$  and  $\text{Ln}_2\text{ndc}_3$ , while 90% absorption at 375 nm requires 550 nm and 1  $\mu\text{m}$  of  $\text{Ln}_2(\text{bdc-2OH})_3$  and  $\text{Ln}_2(\text{bdc-NH}_2)_3$ , respectively. The luminescence from these films span the full range from violet to blue, green and red as shown in Fig. 16 along with the films colour coordinates.

## Conclusion

In this work we have deposited hybrid films of lanthanides with 9 different mono- and bi-aromatic di-carboxylic acids, and explored their optical and luminescent properties. The terbium-terephthalic acid system has been used as a base and is explored in more detail, including film growth and air stability. Yttrium has been used as an optically passive cation so that the luminescence properties of the aromatic molecule itself can be investigated. The yttrium films show that the two bi-aromatic acids produce films with bright violet and blue luminescence. Yet with Tb and Eu all bi-aromatic films are non-luminescent except  $\text{Eu}_2\text{ndc}_3$ . The bi-aromatic films show by far the strongest absorption strength of all the acids. The two acids with added functional groups on the other hand, namely bdc-2OH and bdc-NH<sub>2</sub>, show the largest red-shift in absorption and even reaches slightly into the blue range for bdc-2OH. In conclusion, the optical properties of hybrid materials made from lanthanides and aromatic di-carboxylic acids show a large variation in both absorption range and strength while their luminescence spans violet, blue, green and red.

## Author contributions

Hansen: conceptualization, data curation, formal analysis, supervision, writing – original draft. Svendsen: investigation, formal analysis, writing – review & editing. Nesteng: investigation, formal analysis, writing – review & editing. Nilsen: data

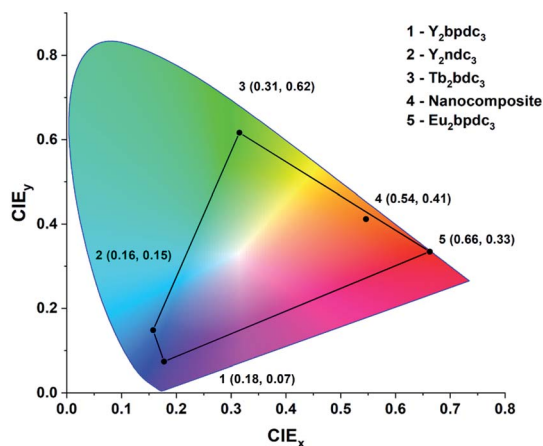
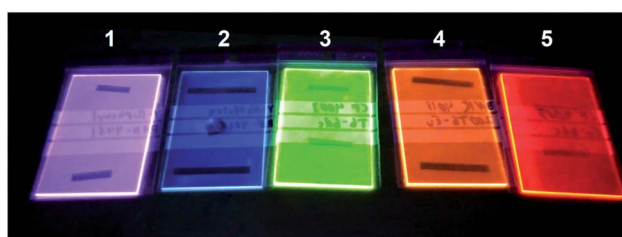


Fig. 16 Photo and CIE coordinates of samples from this work under a 254 nm UV lamp. The violet, blue, green and red films as  $Y_2\text{bpdc}_3$ ,  $Y_2\text{ndc}_3$ ,  $\text{Tb}_2\text{bdc}_3$  and  $\text{Eu}_2\text{bdc}_3$ . The orange film is a hybrid-fluoride nanocomposite from another recent work,<sup>8</sup> showing orange  $\text{Eu}^{3+}$  emission sensitized by terephthalic acid layers.



curation, formal analysis, supervision, writing – review & editing, project administration, funding acquisition.

## Conflicts of interest

There is no conflict of interest to report.

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